

**A NOVEL AGROPOLYMER USED FOR PURIFICATION OF POLLUTED
WATER OR CONTAMINATED WATER CONTAINING METAL OR IONS
AND A METHOD OF PRODUCING THE AGROPOLYMER**



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FIELD OF THE INVENTION

This invention in general relates to the filed of agricultural bio-technology. In particular, this invention pertains to a novel agropolymer, a method of manufacturing the agropolymer from any plant parts such as seed coats, hulls, husks or seed covers of plants including agricultural crops (*Oryza sativa*, *Panicum miliaceum*, *Setaria italica*, *Cajanus cajan*, *Vigna mungo*, *Vigna radiata*, *Triticum sp.*, *Ricinus communis*, *Helianthus annus*, *Gossypium sp.* *Arachis sp.*). These novel agropolymers have extensive industrial applications and also useful for water pollution control.

BACKGROUND OF THE INVENTION

Many biological origin, metal sequestering substances were known such as Tannins, Humic acid, whole cell biomass, Chitin and Chitin derivatives, Metallothioneins, Microbial Polysaccharides, Melannins, Ployphenolic bipoigments, Bacterial cell wall polymers, Microbially produced chelating agents (Siderophores). However, the above materials are either costly, not available in sufficient quantity and / or less effective. On the other hand, the present research work is aimed to produce agropolymers from cheaply available plant materials, preferably, agricultural raw materials such as seed coats or hulls of agricultural crops (*Oryza sativa*, *Panicum miliaceum*, *Setaria italica*, *Cajanus cajan*, *Cigna mungo*, *Vigna radiata*, *Triticum sp.*, *Ricinus communis*, *Helianthus annus*, *Gossypium sp.* *Arachis sp.*) These agropolymers have extensive industrial applications.

This invention has its application in the development of production of novel agropolymers preferably from seed coats or hulls of agricultural crops (*Oryza sativa*, *Panicum miliaceum*, *Setaria italica*, *Cajanus cajan*, *Vigna mungo*, *Vigna radiata*,



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Triticum sp., *Ricinus communis*, *Helianthus annus*, *Gossypium sp.* *Arachis sp.*) and their applications in metal sequestration and removal of ions.

Therefore, the objective of the present invention is to invent novel agropolymers from plant materials including seed coats of agricultural crops (*Oryza sativa*, *Panicum miliaceum*, *Setaria italica*, *Cajanus cajan*, *Vigna mungo*, *Vigna radiata*, *Triticum sp.*, *Ricinus communis*, *Helianthus annus*, *Gossypium sp.* *Arachis sp.*). It is also the intention of invention to develop viable method for production agropolymers and use of such agropolymers to sequester or remove metals and / or ions. Thus, preferably, these agropolymers are produced from the seed coats or hulls of agricultural crops.

OBJECTS OF THE INVENTION

Accordingly, it is the primary object of the invention is to invent bio-origin, non-toxic, biodegradable, cheaply available, very effective molecules/ polymers from agricultural sources, for industrial applications and water pollution control.

Another object of the invention is to develop production method for agropolymers from plant materials such as seed coats, husks, hulls, seed covers of agricultural crops (*Oryza sativa*, *Panicum miliaceum*, *Setaria italica*, *Cajanus cajan*, *Vigna mungo*, *Vigna radiata*, *Triticum sp.*, *Ricinus communis*, *Helianthus annus*, *Gossypium sp.* *Arachis sp.*).

Yet, another object of the invention is to invent a polymer which will reduce environment pollution by using agriculturally derived metal and ion sequestering agropolymers from seed coats or hulls of agricultural crops.

Still another object of the invention is to provide a method of removal of heavy metals and ions from aqueous solutions using agropolymers of the invention.

One more object of the invention is to provide a method / technique to reduce metal and ion pollution from water to protect the environment.

Another object of the invention is to provide a method to convert the toxic metal and/or ion containing water to non-toxic, metal and/or ion free water by using agricultural crops derived non-toxic substances, which are biodegradable substances.

Another object of the invention is to provide a novel method to reduce environmental pollution by using agriculturally derived metal and ion sequestering agropolymers obtained from seed coats or hulls of agricultural crops.

SUMMARY OF INVENTION

To meet the above objects and others, present invention provides a novel agropolymer comprising carbohydrate and / or silica matrix substantially devoid of atleast proteins, tannins, pigments and polyphenols, and having metal binding reactive sites, a method of producing the agropolymers and use of such agropolymers to isolate or remove metals or ions from aqueous solutions.

DETAILED DESCRIPTION OF THE INVENTION

Now, the invention will be described in detail so as to illustrate and explain various salient features of the invention.

One embodiment of the invention is to provide a novel agropolymer having metal binding sites, which are incorporated, in a matrix of the agropolymer either by alkali treatment or hydrogen peroxide treatment or alkaline hydrogen peroxide treatment.

The matrix of the agropolymer is obtained from any plant parts such as seed coats, seed cover, hulls and husks. The metal binding reactive sites (organo-metallic bonds) which were created by reacting agropolymers with metal as observed under Infra Red (IR) Spectroscopy.

Another embodiment of the present invention relates to a method of production of agropolymers from plant materials selected from seed coats or hulls of agricultural crops (such as seed coats or hulls, from (*Oryza sativa*, *Panicum miliaceum*, *Setaria italica*, *Cajanus cajan*, *Vigna mungo*, *Vigna radiata*, *Triticum sp.*, *Ricinus communis*, *Helianthus annus*, *Gossypium sp.* *Arachis sp.*) which have metal and ion sequestering nature.

In a preferred embodiment, the present invention provides a method of producing agropolymers from a plant material such as seed coats or hulls of agricultural crops (seed coats or hulls, from *Oryza sativa*, *Panicum miliaceum*, *Setaria italica*, *Cajanus cajan*, *Vigna mungo*, *Vigna radiata*, *Triticum sp.*, *Ricinus communis*, *Helianthus annus*, *Gossypium sp.* *Arachis sp.*), said method comprising the following steps:-

- a. powdering the seed coat or hull material,
- b. micronization of seed coat or hull powder to required microns size,
- c. treating the said micronized seed coat or hull powder in alkali or alkaline hydrogen peroxide or hydrogen peroxide,
- d. treating the said material with repeated washings with water or with acid to remove alkali and or hydrogen peroxide residue,
- e. treating the said material with acid solution to remove bound metals,
- f. neutralizing the said molecules by removal of acid by way of water washings or by addition of diluted alkali solution and
- g. drying the resultant agropolymers at room temperature or with a dryer (70-80°C).

In another embodiment, powdering of the seed coat or hull material was obtained by grinders and micronization of seed coat or hull powder is effected using micronizer to obtain desired particle size (in microns).

In one more embodiment, the alkali treatment is effected with sodium hydroxide or potassium hydroxide or sodium carbonate or adding any alkali solution to the plant materials and or adding hydrogen peroxide to the reaction mixture.

Hydrogen peroxide (5 to 30% concentration) treatment along with alkali (1 to 10% w/w) resulted in agropolymers with efficient metal sequestration. Alkali and or Hydrogen peroxide residues left after treatment were removed by repeated water washings or by addition of diluted acid (using such as H_2SO_4 , HCL or HNO_3) separately.

In this method, if any bound metals are present in agropolymers, they are eluted with 1-3% acid solutions (using such as H_2SO_4 or HCL or HNO_3) and any acid residues left behind are removed by repeated water washings or by addition of diluted alkali (using such as NaOH or KOH) solution, and moisture content left behind, which is being removed by decantation of supernatant followed by drying in a dryer (70-80°C) and / or drying at room temperature.

The alkali treatment without hydrogen peroxide also results agropolymers with metal sequestration particularly for the production of agropolymers from cereals and millets. The micronized seed coats or hulls are mixed with alkali (such as sodium hydroxide or potassium hydroxide). Alkali treatment is carried either by adding alkali solution to micronized seed coat or hull power or by adding directly alkali powder/flakes to micronized seed coat or hull powder later followed by addition of water. Alkali treatment releases dark brownish yellow colour substances from seed coats or hulls which are water soluble in nature. Lower percentage of alkali (5 to 7.5%) takes more retention time to remove these dark brownish yellow substances from seed coats or hulls where as higher percentage of alkali (20-25% solution) removes the dark brownish yellow substances within 3-4 hours. Alkali content is removed by repeated water washings or by addition of acid solution (such as HCL or H_2SO_4). The bound metals from resulting substances were removed by treating with mineral acids including sulfuric, nitric or hydrochloric (1-3%). The resultant material was neutralized with repeated water washings or by addition of required amount of 1M alkali (such as NaOH, KOH) solution. The material after removal of supernatant is dried at room temperature or in a drier (70-80°C). Lower the size, better the metal or ion sequestering property.

The yield of agropolymers depends on the size and reaction procedure adopted. Lower size agropolymers yield is less compared to big size agropolymers yield. Process losses are more to produce lower sizes agropolymers. The source of the material also decides the yield of the agropolymer. In general, about 30 to 40% yield of agropolymers was obtained in case of producing lower sizes agropolymers. About 75 to 80% yield of agropolymers is obtained in case of producing higher size (above 150 micron size) agropolymers.

The present method involves treating the plant material with alkali and / or Hydrogen peroxide. The plant material can be treated by alkaline hydrogen peroxide and in this case, the alkaline treatment can be performed before hydrogen peroxide treatment or along with hydrogen peroxide or after treatment with hydrogen peroxide. The most preferred embodiment involves treating the plant material with alkaline hydrogen peroxide combindly at one time.

The flow chart illustrates various stages / steps in the production of agropolymers.

The micronized seed coat or hulls powder is, treated with sulfuric acid or hydrochloric acid, (3-5%) for a period of 5-6 hours and the resultant material after removal of acid traces exhibited metal or ion sequestering property, but with less sequestering ability compared to agropolymers produced by alkali treatment method as described earlier.

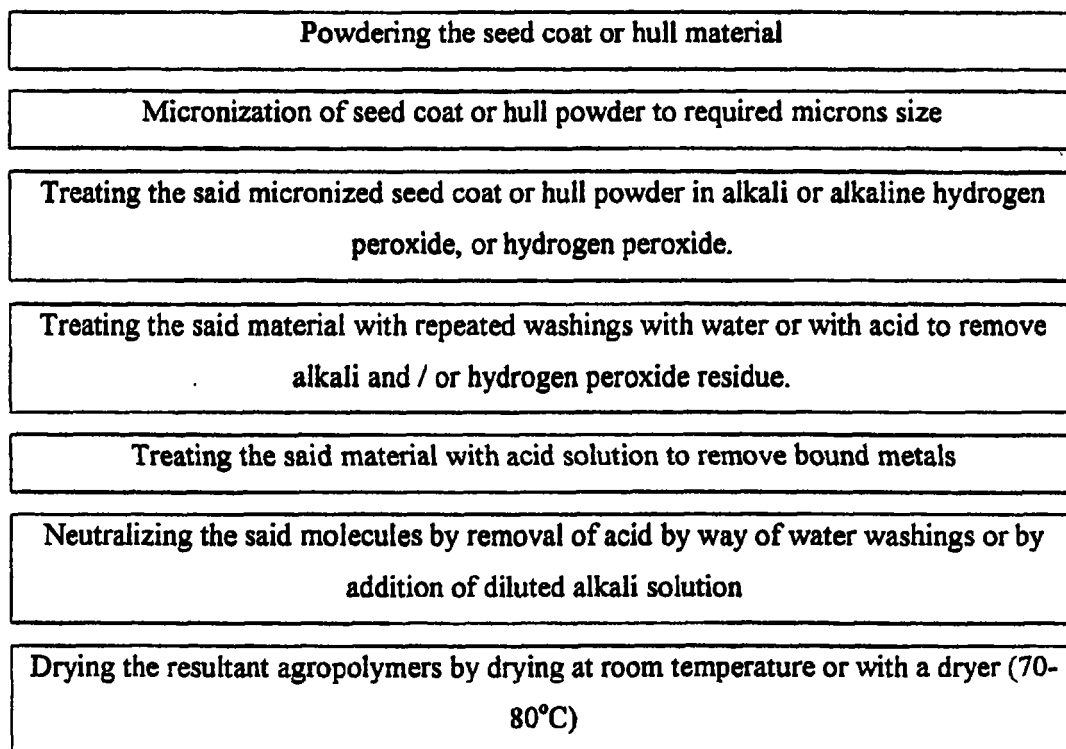
One of the salient features of the invention is the identify agropolymers present in seed coats or hulls of agricultural crops (*Oryza sativa*, *Panicum miliaceum*, *Setaria italica*, *Cajanus cajan*, *Vigna mungo*, *Vigna radiata*, *Triticum sp.*, *Ricinus communis*, *Helianthus annus*, *Gossypium sp.* *Arachis sp.*).

Another salient feature of the invention is regarding the production of agropolymers from seed coats or hulls or agricultural crops. For production of agropolymers from seed coats or hulls (such as seed coats or hulls, from *Oryza sativa*,

Panicum miliaceum, *Setaria italica*, *Cajanus cajan*, *Vigna mungo*, *Vigna radiata*, *Triticum sp.*, *Ricinus communis*, *Helianthus annuus*, *Gossypium sp.* *Arachis sp.*) alkaline hydrogen peroxide treatment method was developed.

The agropolymers have extensive industrial applications. They can be effectively used for pollution control to protect environment from metal or ion contamination. The wide application of the said substance is an important aspect of the invention. Accordingly, the invention pertains to a method of producing agropolymers, preferably from seed coats or hulls of agricultural crops.

FLOW DIAGRAM OF AGROPOLYMES MANUFACTURING PROCESS



In another embodiment of the invention, the applicant provides method of treating metals / ions using the novel agropolymers of the present invention.

In particular, this invention pertains to a method of preventing environmental degradation by metal and ion pollution using agropolymers produced from seed coats or hulls of agricultural crops. More particularly, this invention relates to a method of purifying metal or ion polluted aqueous solution including metal or ion polluted industrial effluent ground water and drinking water. The method of purification is carried using a novel agropolymer which is capable of sequestering metals including metals like organic, mercury, etc. and/or ions from polluted water. The ultimate objective of the invention is to protect environment from metal and ion pollution. This objective is achieved not by synthetic route but by using natural biological, agricultural resources, which are abundant in nature. This invention has its salient feature resides in inventing substances and effective utilisation of this substances by using in novel technique to purify / sequester metals and or ions from polluted water.

In embodiment this present invention, a method of removal of metals and ions from aqueous solutions using agropolymers produced from seed coats or hulls of agricultural crops (from seed coats or hulls from *Oryza sativa*, *Panicum miliaceum*, *Setaria italica*, *Cajanus cajan*, *Vigna mungo*, *Vigna radiata*, *Triticum sp.*, *Ricinus communis*, *Helianthus annus*, *Gossypium sp.* *Arachis sp.*) is described. Many metals such as iron, copper, aluminium, arsenic, mercury, lead, zinc set and ions can be removed from an aqueous solution using the present novel agropolymer.

This invention further relates effective utilization of metal and ion sequestering agropolymers from the seed coats or hulls of agricultural crops. This invention is useful for metal and ion pollution control using biotechnological approach for environmental protection. Although many plant substances are known for metal sequestration but production on large scale for industrial usage was not achieved so far due to many factors such as availability of raw materials and production costs. Many biological origin, metal sequestering substances were known such as Tannins, Humic acid, whole cell biomass, Chitin and Chitin derivatives, Metallothioneins, Microbial Polysaccharides, Melannins, Ployphenolic bipoigments, Bacterial cell wall polymers, Microbially produced chelating agents (Siderophores). Very few studies have been targeted to found utility in actual situations. As it is a

growing area, experimental studies have been carried out in order to generate / produce / develop an utility method for this new kind of agropolymers from agricultural resources to find utility in actual situations.

The applicant has now devised a method for utilisation of metal and/or ion sequestering agropolymers obtained from seed coats or hulls of agricultural crops. Thus, according to this invention the agropolymers produced from the seed coats or hulls of agricultural crops are responsible for metal, and/or ion sequestering properly. For utility of metal and ions sequestering agropolymers from seed coats of agricultural crops, technically feasible and economically viable method are developed.

The present invention relates to a method for purification of a metal and ion polluted / contaminated aqueous solutions including metal or ion polluted drinking water or ground water, by treating the polluted water with agropolymers and or metal impregnated agropolymers, the said agropolymers being manufactured from plant materials such as seed coats, husks or hulls of the various agricultural crops (*Oryza sativa*, *Panicum miliaceum*, *Setaria italica*, *Cajanus cajan*, *Vigna mungo*, *Vigna radiata*, *Triticum sp.*, *Ricinus communis*, *Helianthus annus*, *Gossypium sp.* *Arachis sp.*), the said method comprising the contacting the ion or metal polluted water with agropolymers and or metal impregnated agopolymer using a column or batch mode, effecting the ion or metal sequestration resulting in pollution free water, the said sequestration being carried out under optimum conditions (such as pH) for maximum sequestration efficiency.

In fact, the agropolymers being produced from seed husks, hulls or coats of agricultural crops such as *Oryza sativa*, *Panicum miliaceum*, *Setaria italica*, *Cajanus cajan*, *Vigna mungo*, *Vigna radiata*, *Triticum sp.*, *Ricinus communis*, *Helianthus annus*, *Gossypium sp.* *Arachis sp.*) these agropolymers are capable of removal of metals or ions from aqueous solutions from PPM to PPB level.

Accordingly, the invention also envisages a method of treating metal or ion contaminated aqueous solution, the said method comprises by contacting the above

defined polluted water with agropolymers (by column or batch mode), the said substances obtained from biological / agricultural plant sources (seed coats or hulls from plant material selected from agricultural plants such as *Oryza sativa*, *Panicum miliaceum*, *Setaria italica*, *Cajanus cajan*, *Vigna mungo*, *Vigna radiata*, *Triticum sp.*, *Ricinus communis*, *Helianthus annus*, *Gossypium sp.* *Arachis sp.*) which have ability to reduce metal and ion load from PPM and PPB level from polluted water, the said metal or ion comprising water by contacting with agropolymers on column or batch mode.

Further, the invention envisages a method of treating metal and/or ion polluted water where in the agropolymers are obtained from seed coats or hulls of agricultural crops (seed coats or hulls, from *Oryza sativa*, *Panicum miliaceum*, *Setaria italica*, *Cajanus cajan*, *Vigna mungo*, *Vigna radiata*, *Triticum sp.*, *Ricinus communis*, *Helianthus annus*, *Gossypium sp.* *Arachis sp.*). The agropolymers are also used in the possible sites of ground water contamination of toxic metals. In fact, the natural ground water which is rich in arsenic is treated with the agropolymers of the present invention and the results revealed that the arsenic content has been much reduced so that the arsenic rich water indeed become a potable water.

BRIEF DESCRIPTION OF THE ACCOMPANYING DRAWINGS

In the enclosed drawings;

Raw husk, hull or coat was micronized and IR Spectroscopy was done with KBr pellet. Samples were treated with ferric chloride and dried before taking IR Spectroscopy with KBr pellet. IR Spectroscopy of each sample figure was done in three patterns; (1) total scan: 4000 to 500 wave numbers (cm⁻¹), (2) scan from 4000 to 2200 wavenumbers (cm⁻¹), and (3) scan from 2000 to 600 wavenumbers (cm⁻¹).

Figure 1 relates to IR Spectra of raw rice husk.

Figure 2 relates to IR Spectra of rice husk treated with alkaline hydrogen peroxide.

Figure 3 relates to IR Spectra of raw rice husk treated with ferric chloride.

Figure 4 relates to IR Spectra of alkaline hydrogen peroxide treated rice husk which in turn treated with ferric chloride.

As shown in IR Spectra of various samples of rice husk revealed that the alkaline hydrogen peroxide treatment to rice husk resulted in more reactive bonds i.e. organo-metallic bonds.

Figure 5 relates to IR Spectra of raw *Setaria italica* husk.

Figure 6 relates to IR Spectra of alkaline hydrogen peroxide treated *Setaria italica* husk.

Figure 7 relates to IR Spectra of raw *Setaria italica* husk treated with ferric chloride.

Figure 8 relates to IR Spectra of alkaline hydrogen peroxide treated *Setaria italica* husk which in turn treated with ferric chloride.

As shown in the above IR Spectroscopy of figures 5 to 8, the more reactive organo-metallic bonds were formed with iron.

Figure 9 relates to IR Spectra of alkaline hydrogen peroxide treated wheat husk.

Figure 10 relates to IR Spectra of alkaline hydrogen peroxide treated wheat husk which in turn treated with ferric chloride.

As shown in IR Spectroscopy of figures 9 and 10, alkaline hydrogen peroxide treated wheat (*Triticum sp.*) husk resulted in many organo-metallic bonds in particular significantly at 2360 ± 10 and 2340 ± 10 wavenumbers(cm^{-1})

Figure 11 relates to IR Spectra of *Panicum miliaceum* husk.

Figure 12 relates to IR Spectra of alkaline hydrogen peroxide treated *Panicum miliaceum* husk which in turn treated with ferric chloride.

The IR Spectroscopy of figures 11 and 12 reveals that organo-metallic bonds were more predominant from 1600 to 600 wavenumbers (cm^{-1}).

Figure 13 relates to IR Spectra of alkaline hydrogen peroxide treated cotton seed (*Gossypium sp.*) hulls.

Figure 14 relates to IR Spectra of alkaline hydrogen peroxide treated cotton seed (*Gossypium sp.*) hulls which in turn treated with ferric chloride.

As shown in IR Spectroscopy of figures 13 and 14, alkaline hydrogen peroxide treated cotton seed (*Gossypium sp.*) hulls resulted in many organo-metallic bonds in particular significantly at 2360 ± 10 and 2340 ± 10 wavenumbers(cm^{-1})

Figure 15 relates to IR Spectra of alkaline hydrogen peroxide treated castor (*Ricinus communis*) seed coats.

Figure 16 relates to IR Spectra of alkaline hydrogen peroxide treated castor (*Ricinus communis*) seed coats which in turn treated with ferric chloride

As shown in IR Spectroscopy of figures 15 and 16, alkaline hydrogen peroxide treated castor seed coats (*Ricinus communis*) resulted in many organo-metallic bonds.

Figure 17 relates to IR Spectra of alkaline hydrogen peroxide treated sunflower (*Helianthus annuus*) seed coats.

Figure 18 relates to IR Spectra of alkaline hydrogen peroxide treated sunflower (*Helianthus annuus*) seed coats which in turn treated with ferric chloride.

As shown in IR Spectroscopy of figures 17 and 18, alkaline hydrogen peroxide treated sunflower seed coats (*Helianthus annuus*) resulted in many organo-metallic bonds.

Figure 19 relates to IR Spectra of alkaline hydrogen peroxide treated redgram (*Cajanus cajan*) seed coats.

Figure 20 relates to IR Spectra of alkaline hydrogen peroxide treated redgram (*Cajanus cajan*) seed coats which in turn treated with ferric chloride

As shown in IR Spectroscopy of figures 19 and 20, alkaline hydrogen peroxide treated redgram seed coats (*Cajanus cajan*) resulted in many organo-metallic bonds.

Figure 21 relates to IR Spectra of alkaline hydrogen peroxide treated greengram (*Vigna radiata*) seed coats.

Figure 22 relates to IR Spectra of alkaline hydrogen peroxide treated greengram (*Vigna radiata*) seed coats which in turn treated with ferric chloride.

As shown in IR Spectroscopy of figures 21 and 22, alkaline hydrogen peroxide treated greengram (*Vigna radiata*) seed coats resulted in many organo-metallic bonds in particular significantly at 2360 ± 10 and 2340 ± 10 wavenumbers(cm^{-1})

Figure 23 relates to IR Spectra of alkaline hydrogen peroxide treated blackgram (*Vigna mungo*) seed coats.

Figure 24 relates to IR Spectra of alkaline hydrogen peroxide treated blackgram (*Vigna mungo*) seed coats which in turn treated with ferric chloride. As shown in IR Spectroscopy of figures 23 and 24, alkaline hydrogen peroxide treated blackgram (*Vigna mungo*) seed coats resulted in many organo-metallic bonds.

Now, the applicant provide the following specific description by way of examples and illustrations of the invention and this should not be construed to limit the scope of the invention in any manner.

The metal sequestering property of agropolymers derived from seed coats of agricultural crops (seed coats or hulls, from *Oryza sativa*, *Panicum miliaceum*, *Setaria italica*). When a gram agropolymer sample is taken in a 1000 ml volumetric flask and made upto the mark with 20.0 PPM of standard copper and silver solutions separately and then the solution in the flasks are kept for 2 hours with regular shaking in between, then the solution as filtered and the copper and silver present in solution was estimated Spectrophotometrically. The difference between the metal content present in solution before and after addition of agropolymer indicate the metal absorbing / sequestering ability of particular agropolymer. As shown in table – 2, agropolymers sequestered metals such as copper and silver. A gram of the *Setaria italica* agropolymers absorbed / sequestered 6.0, 4.1 milligrams of copper and silver respectively. A gram of *Panicum miliaceum* agropolymers sequestered 1.6 and 2.5 milligrams of copper and silver respectively. A gram of the *Oryza sativa* agropolymers sequestered 4.5 and 4.7 milligrams of copper and silver respectively.

Agropolymers exhibits a nature of absorbing more metal content from a solution containing high metal concentration. Using this property agropolymers were allowed to absorb more metal content in high metal concentration solutions with longer retention time (12 – 24) hours. Agropolymers were added to metal (such as iron or aluminium containing solutions) and unbound metal was removed by through washings with water and or neutralising with alkali in case of reaction medium is in acidic pH later the material was dried. Metal impregnated Agropolymer was taken in a beaker and to it 250 ml. Of water was added and stirred well, then the agropolymer material was filled on a column mode. The agropolymer material present in column was washed with 50 ml of 1 N acid to obtain bound metal. The metal content in acid washings were estimated by Spectrophotometrically. Table – 3 denotes the amount of metal content bound on agropolymer. A gram of *Setaria italica* agropolymer absorbed (when treated with aluminium chloride) an amount of 14.4 milligrams of aluminium. A gram of *Oryza sativa* agropolymers absorbed (when treated with aluminium sulphate) an amount of 8.6 milligrams of aluminium. A gram of *Setaria italica* Agropolymer absorbed (when treated with Ferric chloride) an amount of 4.7 milligrams of iron.

Now, the following description will illustrate the arsenic sequestering property of agropolymer. In this experiment agropolymers bound with copper, zinc and iron were taken separately for measuring the arsenic sequestering nature.

One gram of agropolymer was added to 100 ml. Solution containing 6.6 PPM sodium arsenate and stirred will for a period of 3 to 4 hours. The arsenic content present in supernatant was estimated by Spectrophotometrically. As shown in Table 4, metal impregnated agropolymers, absorbed arsenic from aqueous solutions. Copper, iron and zinc impregnated *Setaria italica* agropolymers absorbed about 73 – 75 percent of arsenic from an initial arsenic amount of 6.6. PPM.

Agropolymers without metal impregnation also absorb arsenic. As shown in Table 5, agropolymers reduced significantly arsenic content from natural waters containing arsenic.

In the following description the fluoride removal property of agropolymer will be described. As it is agropolymer can not absorb any fluoride ion but when agropolymer was bound with metals such as aluminium, the metal bound agropolymer can absorb fluoride ions significantly. Agropolymers bound with aluminium sulphate were added to a solution containing 5 PPM sodium fluoride (1000 milligrams of metal impregnated agropolymer was added to 50 ml of 5 PPM sodium fluoride) these metal bound agropolymers absorbed fluoride ion, 77.4 and 90.87 % respectively. Similarly, these metal bound agropolymers removed fluoride ion from natural water containing fluoride. When a natural water having 4.15 PPM fluoride content was mixed with (1) Aluminium chloride impregnated *Setaria Italica*. Agropolymer and (2) Aluminium Sulphate impregnated *Oryza Sativa* Agropolymer at a dose of 1 gm. Per liter removed fluoride ion significantly.

Now, this invention will describe the important parameters for metal sequestration and bound metal elution. The said substances effective use is dependant on suitable pH for metal sequestration. The bound metals in the said substances can be eluted with mineral acids including sulfuric, nitric or hydrochloric at 0.8 – 1.0 pH.

One of the salient features of the invention is the identification of metal or ion sequestering property of agropolymers present in seed coats or hulls of agricultural crops and utilisation of agropolymers in water treatment.

Another salient feature of the invention is that the agropolymers are useful on column mode or batch mode and suitable for repeated usage on many cycles to remove or sequester various metals from PPM to PPB level.

Metal impregnated agropolymers were found useful in fluoride ion sequestration and metal absorption such as arsenic. The said substance sequestering

ability is maximum at neutral pH range for most of the metals tested. The most of the bound metals present in said substances (agropolymers) can be eluted with mineral acids including sulfuric, nitric and hydrochloric at 0.8 – 1.0 pH. The agropolymers are suitable for repeated usage on column or batch mode offers an economic advantage.

The agropolymers have extensive industrial applications. They can be effectively used for pollution control to protect environment from metal or ion contamination. The wide application of the said substance is the subject matter of invention. Accordingly, this invention pertains to a method of removing metals and ions from contaminated water by either column method or batch method, so that the toxicity causing metals and ions are removed from water.

Now, it is essential to summarise the findings of the investigation carried by the applicant and the findings are given in the following Tables 1 to 5:-

Table – 1 denotes the metal (iron) sequestering property of agropolymers derived from *Oryza sativa*, *Panicum miliaceum*, *Setaria italica*, *Cajanus cajan*, *Vigna mungo*, *Vigna radiata*, *Triticum sp.*, *Ricinus communis*, *Helianthus annus*, *Gossypium sp.* *Arachis sp.* by alkaline hydrogen peroxide treatment. Ferric chloride solution was passed through a gram of agropolymer in a column at a flow rate of 2ml per minute and bound metal content was estimated by eluting with 2 – 5% hydrochloric acid. Before elution, the unbound excess metal from column was removed by washing with 2.5pH hydrochloric acid solution. The bound metal was estimated using par reagent at 535n.m. spectrophotometrically.

Table -2 denotes the metals sequestering property of agropolymers derived form seed coats or hulls, by alkaline treatment of *Setaria italica*, *panicum miliaceum*, and *Oryza sativa*, sequestered copper and silver.

Table – 3 denotes the metal impregnation (aluminium and iron) on agropolymers derived from seed coats or hulls from *Setaria italica*, and *Oryza sativa*.

Table – 4 denotes the arsenic sequestering property of metal impregnated Agropolymer (copper, iron, zinc impregnated on Agropolymer derived from seed coats or hulls form *Setaria italica*).

Table – 5 denotes the arsenic sequestering property of agropolymers from natural waters containing arsenic.

Present invention explains the ion (fluoride) removal or sequestering property of metal impregnated agropolymers (aluminium impregnated on agropolymers derived from seed coats or hulls from *Setaria italica* and *Oryza sativa*).

The results derived from the present invention gives a scope of using agropolymers and metal impregnated agropolymers using in affinity columns to purify or bind or remove or sequester or react with reactive substances useful for various industrial applications (such as using these agropolymers in manufacturing of various derivatives such as Biodegradable plastics, resins) including using agropolymers for reducing ground water contamination by industrial waste water containing metals and ions.

TABLE – 1
METAL (IRON) SEQUESTERING PROPERTY OF AGROPOLYMERS
PRODUCED BY ALKALINE PEROXIDE TREATMENT METHOD

S.No.	AGROPOLYMER PRODUCED FROM SEED COATS OR HULLS OF:	IRON CONTENT (MILLIGRAM) SEQUESTERED BY A GRAM OF AGROPOLYMER
1.	<i>Oryza sativa</i>	5.15
2.	<i>Panicum miliaceum</i>	3.125
3.	<i>Setaria italica</i>	5.0
4.	<i>Cajanus cajan</i>	11.75
5.	<i>Vigna mungo</i>	6.875
6.	<i>Vigna radiata</i>	17.18
7.	<i>Triticum sp.</i>	3.125
8.	<i>Ricinus communis</i>	2.3
9.	<i>Helianthus annus</i>	11.25
10.	<i>Gossypium sp.</i>	6.25
11.	<i>Arachis sp. (Red seed coat or cover of groundnut)</i>	20.56

TABLE – 2
METAL SEQUESTERING PROPERTY OF AGROPOLYMERS PRODUCED
BY ALKALINE TREATMENT METHOD

S.NO	AGROPOLYMER PRODUCED FROM SEED COATS OR HULLS OF:	COPPER CONTENT (MILLIGRAM) SEQUESTERED BY A GRAM OF AGROPOLYMER	SILVER CONTENT (MILLIGRAM) SEQUESTERED BY A GRAM OF AGROPOLYMER
1	<i>Setaria italica</i>	6.0	4.1
2.	<i>Panicum miliaceum</i>	1.6	2.5
3.	<i>Oryza sativa</i>	4.5	4.7

TABLE - 3
BOUND METAL CONTENT PRESENT IN METAL IMPREGNATED
AGROPOLYMERS PRODUCED BY ALAKALINE TREATMENT METHOD

S.NO	AGROPOLYMER PRODUCED FROM SEED COATS OR HULLS + METAL	METAL CONTENT (MILLIGRAM) SEQUESTERED BY A GRAM OF AGROPOLYMER
1	<i>Setaria italica</i> + ALUMINIUM CHLORIDE	14.4
2.	<i>Oryza sativa</i> + ALUMINIUM SULPHATE	8.6
3.	<i>Setaria italica</i> + FERRIC CHLORIDE	4.7

TABLE - 4
ARSENIC SEQUESTERING PROPERTY OF METAL IMPREGNATED
AGROPOLYMERS PRODUCED BY ALAKALINE TREATMENT METHOD

S.NO	AGROPOLYMER PRODUCED FROM SEED COATS OR HULLS + METAL	INITIAL ARSENIC CONTENT (PPM)	ARESENIC SEQUESTERED BY A GRAM OF AGROPOLYME R (PER CENTAGE)
1.	<u><i>Setaria italica</i></u> + COPPER SULPHATE + IMPREGNATED	6.6	73.18
2.	<u><i>Setaria italica</i></u> + FERRIC CHLORIDE IMPREGNATED	6.6	73.3
3.	<i>Setaria italica</i> + ZINC CHLORIDE IMPREGNATED	6.6	75

TABLE – 5

ARSENIC SEQUESTERING PROPERTY OF AGROPOLYMERS FROM
NATURAL WATERS CONTAINING ARSENIC

Sl.No.	AGROPOLYMER PRODUCED FROM SEED COATS OR HULLS	INITIAL ARSENIC CONTENT PRESENT IN NATURAL WATERS	ARSENIC CONTENT PRESENT IN NATURAL WATERS AFTER TREATMENT WITH AGROPOLYMERS FOR 12 HOURS
1.	<i>Setaria italica</i>	325 PPB	40PPB
2.	<i>Oryza sativa</i>	325 PPB	50PPB

* The natural waters containing arsenic samples were collected from West Bengal State in India.

Advantages:

The novel agropolymers obtained from plant materials preferably from seed coats or hulls of agricultural crops (such as seed coats or hulls, from agricultural plants such as *Oryza sativa*, *Panicum miliaceum*, *Setaria italica*, *Cajanus cajan*, *Vigna mungo*, *Vigna radiata*, *Triticum sp.*, *Ricinus communis*, *Helianthus annuus*, *Gossypium sp.* *Arachis sp.*) are non-toxic biodegradable plant materials suitable for easy disposal which can be produced at relatively cheaper cost. The properties of agropolymers offers scope of their utility in affinity chromatography systems, as materials to remove or bind or purify or react through immobilisation with reactive molecules through column or batch mode and scope of utility in manufacturing of biodegradable plastics, resins, carrier materials etc. This economically viable technically feasible production method of agropolymers is useful for environmental protection and has extensive industrial applications and further helps farmers to get more income as the raw materials are from agricultural sources.

The novel metal and ion sequestered agropolymers preferably form seed coats of agricultural crops are capable of metal load removal from PPM to PPB level and these substances be produced at relatively cheaper cost compared to other plant derived metal sequestering substances.

Accordingly, a few of the advantages of this invention are summarised as follows:-

1. The invented agropolymers are from natural origin.
2. The process of pollution control using agropolymer is extremely simple.
3. The method or process to prevent environmental degradation by these agropolymers is effective as they can reduce metal or ions load from PPM to PPB level. Therefore, pollution free water without toxic substances is obtained.
4. Agropolymers are environmental friendly and biodegradable.
5. The process of purification of polluted water is economically viable and technically feasible.

It is to be noted that an embodiment of the invention is described in detail in the complete specification so as to illustrate salient features of the invention, further it is envisaged that within the scope of the invention various modification of the invention are permissible. It is evident that description will in no way delimit the scope of the invention. The scope of the invention is described in the ensuing pages.

CLAIMS

1. A novel agropolymer comprising carbohydrate and/or silica matrix substantially devoid of at least proteins, tannins and polyphenols and having metal binding reactive sites.
2. A novel agropolymer as claimed in claim 1 wherein the carbohydrate matrix is obtained from any plant parts such as seed coats, seed covers, hulls and husk.
3. A novel agropolymer as claimed in claim 1 wherein the carbohydrate matrix is obtained from various agricultural crops such as seed coats or hulls from *Oryza sativa*, *Panicum miliaceum*, *Setaria italica*, *Cajanus cajan*, *Vigna mungo*, *Vigna radiata*, *Triticum sp.*, *Ricinus communis*, *Helianthus annus*, *Gossypium sp.* *Arachis sp.*
4. A novel agropolymer as claimed in claim 1 wherein the raw husk is micronized and IR Spectroscopy is done with KBr pellet and thereafter the samples are treated with ferric chloride and dried, the ferric chloride treated sample when subjected to IR Spectroscopy reveals that the alkaline hydrogen peroxide treatment to rice husk resulted in more reactive bonds i.e. organo-metallic bonds.
5. A novel agropolymer as claimed in claim 1 wherein IR Spectroscopy is done with KBr pellet and the alkaline hydrogen peroxide treated wheat (*Triticum sp.*) husk resulted in many organo-metallic bonds in particular significantly at 2360 ± 10 and 2340 ± 10 wave numbers(cm^{-1}), when alkaline hydrogen peroxide treated *Triticum sp.* husk were treated with ferric chloride.
6. A novel agropolymer as claimed in claim 1 wherein IR Spectroscopy is done with KBr pellet and the alkaline hydrogen peroxide treated cotton seed (*Gossypium sp.*) coat resulted in many organo-metallic bonds in particular significantly at 2360 ± 10 and 2340 ± 10 wave numbers(cm^{-1}), when alkaline hydrogen peroxide treated *Gossypium sp.* husk were treated with ferric chloride.
7. A novel agropolymer as claimed in claim 1 wherein IR Spectroscopy is done with KBr pellet and the alkaline hydrogen peroxide treated green gram (*Vigna radiata*) coat resulted in many organo-metallic bonds in particular significantly at 2360 ± 10 and 2340 ± 10 wave numbers(cm^{-1}), when alkaline hydrogen peroxide treated *Vigna radiata* coat were treated with ferric chloride.

8. A novel agropolymer as claimed in claim 1 wherein IR Spectroscopy is done with KBr pellet and the alkaline hydrogen peroxide treated coats of *Panicum miliaceum*, *Setaria italica*, *Cajanus cajan*, *Vigna mungo*, *Ricinus communis*, *Helianthus annuus* coats resulted in typical organo metallic bonds characteristic of each material, when alkaline hydrogen peroxide treated coats of *Panicum miliaceum*, *Setaria italica*, *Cajanus cajan*, *Vigna mungo*, *Ricinus communis*, *Helianthus annuus* were treated with ferric chloride.
9. A method of producing agropolymers from a plant material such as seed coats, husks or hull of agricultural crops (seed coats or hulls, from *Oryza sativa*, *Panicum miliaceum*, *Setaria italica*, *Cajanus cajan*, *Vigna mungo*, *Vigna radiata*, *Triticum sp.*, *Ricinus communis*, *Helianthus annuus*, *Gossypium sp.*, *Arachis sp.*), said method comprising the following steps:-
 - a) powdering the seed coat or hull material,
 - b) micronization of seed coat or hull powder to required microns size,
 - c) treating the said micronized seed coat or hull powder in alkali or alkaline hydrogen peroxide or hydrogen peroxide,
 - d) treating the said material with repeated washings with water or with acid to remove alkali and or hydrogen peroxide residue,
 - e) treating the said material with acid solution to remove bound metals,
 - f) neutralizing the said molecules by removal of acid by way of water washings or by addition of diluted alkali solution, and
 - g) drying the resultant agropolymers at room temperature or with a dryer (70-80°C).
10. A method as claimed in claim 9 wherein, the powdering of the seed husk, coat or hull material is performed by grinding with grinders.
11. A method as claimed in claim 9 wherein, the micronization of seed, husk and coat or hull powder is effected using micronizer to obtain desired particle size (in microns).
12. A method as claimed in claim 9 wherein, the alkali treatment is effected with sodium hydroxide or potassium hydroxide.
13. A method as claimed in claim 9 wherein, the alkali treatment is effected with sodium carbonate.

14. A method as claimed in claim 9 wherein, the alkali treatment is effected by adding the alkali solution to the agropolymers.
15. A method as claimed in claim 9 wherein, the alkali treatment is effected by adding alkali powder/flakes directly in dry state to agropolymers followed by addition of water.
16. A method as claimed in claim 9 wherein, the alkali and/or peroxide residual amount after treatment is removed by repeated water washings.
17. A method as claimed in claim 9 wherein, the alkali and/or peroxide residual amount after treatment is removed by addition of diluted acid (using such as H_2SO_4 , HCl or HNO_3) separately.
18. A method as claimed in claim 9 wherein, the bound metals present in agropolymers is eluted with 1-3% acid solutions (using such as H_2SO_4 , HCl or HNO_3).
19. A method as claimed in claim 9 wherein, acid residues were removed by repeated water washings.
20. A method as claimed in claim 9 wherein, acid residues are removed by addition of diluted alkali (using such as $NaOH$ or KOH) solutions.
21. A method as claimed in claim 9 wherein, the moisture content is removed after decantation of supernatant by drying in a dryer (70-80°C).
22. A method as claimed in claim 9 wherein, the moisture content is removed after decantation of supernatant by drying at room temperature.
23. A method of purifying of metal and ion polluted / contaminated aqueous solutions including metal or ion polluted drinking water or ground water, by treating the polluted water with agropolymers and or metal impregnated agropolymers, the said agropolymers being manufactured from plant materials such as seed coats, husks or hulls of the various agricultural crops (*Oryza sativa*, *Panicum miliaceum*, *Setaria italica*, *Cajanus cajan*, *Vigna mungo*, *Vigna radiata*, *Triticum sp.*, *Ricinus communis*, *Helianthus annuus*, *Gossypium sp.* *Arachis sp.*), the said method comprising the contacting the ion or metal polluted water with agropolymers and or metal impregnated agropolymer using a column or batch mode, effecting the ion or metal sequestration resulting in pollution free water, the said sequestration

being carried out under optimum conditions (such as pH) for maximum sequestration efficiency.

24. A method as claimed in claim 23 wherein, the agropolymers being produced from seed husks, hulls or coats of agricultural crops such as *Oryza sativa*, *Panicum miliaceum*, *Setaria italica*, *Cajanus cajan*, *Vigna mungo*, *Vigna radiata*, *Triticum sp.*, *Ricinus communis*, *Helianthus annuus*, *Gossypium sp.* *Arachis sp.*) these agropolymers are capable of removal of metal from aqueous solutions from PPM to PPB level.
25. A method as claimed in claim 23 wherein, the agropolymers are capable of removal of metals such as arsenic from aqueous solutions from PPM to PPB level.
26. A method as claimed in claim 23 wherein, the agropolymers are capable of removal of ions from aqueous solutions from PPM to PPB level.
27. A method as claimed in claim 23 wherein, the said method is effected by using column mode by sequestering metal or ion load from contaminated aqueous solution uses column method where in the PPM to PPB metal or ion content is obtained, and the process conditions using suitable pH conditions of that particular metal binding and elution and these agropolymers have repeated usage on column mode.
28. A method as claimed in claim 23 wherein, the said method is effected by using batch mode by sequestering metal or ion load form contaminated aqueous solution uses batch method where in the PPM to PPB metal or ion content is obtained, and the process conditions using suitable pH conditions of that particular metal binding and elution and these agropolymers have repeated usage on batch mode.
29. A method as claimed in claim 23 wherein, the agropolymers are used for treating natural ground water rich in toxic metals such as arsenic or mercury to obtain substantially pure or portable water.
30. A method as claimed in claim 23 wherein, the agropolymers are used to prevent potential contamination of ground water by unwanted metals and/or ions in many industrial sites and affluent treatment plants.