

TITLE

5 MULTI-COMPONENT LIQUID AZODINITRILE MIXTURES

RELATED APPLICATIONS

10 This application claims the benefit of U.S. Provisional Application No. 60/003,270, filed August 29, 1995 by the same inventor.

FIELD OF THE INVENTION

15 The present invention relates to multicomponent azodinitrile mixtures which are liquid below 25°C for use as free radical polymerization initiators in various polymerization reactions.

BACKGROUND OF THE INVENTION

20 Azodinitrile compounds represent an important class of free-radical initiating agents which are used in a variety of industrial applications. Examples of such applications include: vinyl polymerizations, graft polymerizations, halogenations, and blowing agents. With respect to peroxide based initiators, azodinitriles display consistent decomposition
25 behavior, less color formation, and a higher degree of safety.

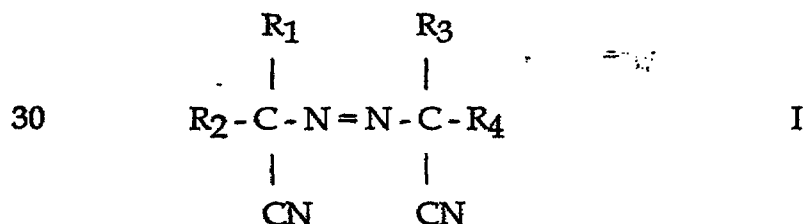
 Azodinitrile compounds can be either symmetrical or asymmetrical about the azo linkage. Symmetrical azo initiators of commercial interest are generally solid and have low to medium solubility in organic solvents.
30 Asymmetrical azo initiators, by contrast, are often low melting solids with higher solubilities in organic solvents. Polymer manufacturers have long sought liquid azodinitrile initiators to eliminate the problems associated with the commercially available, solid azodinitriles such as dusting, ergonomics, and inability to add via screw feeders. A liquid azodinitrile compound or
35 mixture of compounds with a melting point of 0°C or lower, preferably -15°C, would be ideal for many commercial applications. In addition, a liquid azo mixture which facilitates polymerizations using little or no solvent for low volatile organic compound coatings is desirable.

U.S. Patent 3,987,025 of Moore issued October 19, 1976 discloses liquid mixtures of symmetric and asymmetric azodinitrile compounds with a maximum freezing point of 25°C. The starting aminonitriles are limited to two and to those having unbranched hydrocarbon groups. This patent also suggests that it is possible to prepare liquid azodinitrile mixtures when more than two aminonitrile starting materials are used, and that liquid azodinitrile mixtures will be highly favored as the number of aminonitriles is increased above two. However, not all mixtures prepared using more than two aminonitriles are liquids at 25°C, and Moore does not teach any such mixtures.

Many synthetic routes are known for preparation of azodinitriles via reaction of a ketone with a hydrazine compound to generate a hydrazo compound which is then oxidized to azodinitrile. A second approach reacts a cyanohydrin of a ketone with ammonia to form an aminonitrile and oxidatively coupling the aminonitrile to form an azodinitrile. Aliphatic azodinitriles can be prepared by reacting a metal hypochlorite with an aminonitrile in water containing a surfactant. These syntheses do not specifically address how to obtain mixtures prepared from more than two aminonitriles which are liquid at low temperatures. Such mixtures are desirable for use as free-radical initiating agents in industrial applications.

SUMMARY OF THE INVENTION

The present invention comprises an azodinitrile composition and a process for the preparation comprising a mixture of at least six different azodinitriles of formula I

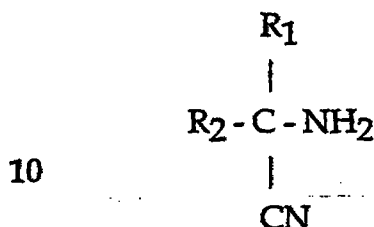


wherein

R_1 , R_2 , R_3 , and R_4 are each independently selected from the group consisting of acyclic aliphatic hydrocarbon radicals of 1 - 9 carbon atoms, preferably 4 - 9 carbon atoms, said mixture having a freezing point of a

maximum of 25°C. Preferably the compositions are liquid at and have a maximum freezing point of 0°C to -15°C.

The present invention further comprises a process for the preparation of azodinitrile mixtures comprising reacting three or more different
 5 aminonitriles of formula



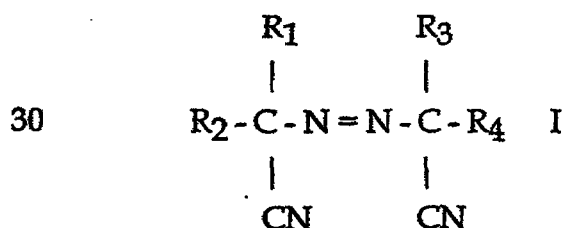
wherein

15 R_1 and R_2 are each independently selected from the group consisting of acyclic aliphatic hydrocarbon radicals of 1 - 9 carbon atoms,

with 5% to 15% by weight based on the reaction mixture of a metal hypochlorite, $M(OCl)_x$, wherein M is selected from the group consisting of sodium, potassium, or calcium, and x is the valence of M ,

20 in an aqueous medium in the presence of 0.25 to 10% by weight based on the weight of aminonitrile of a surface active compound or mixtures thereof having a hydrophilic-lipophilic balance of 8 to 35 at a temperature of -10° C to 30 °C,

said metal hypochlorite and alpha-aminonitrile being present in an equivalent ratio of from 1:1 to 2:1 of hypochlorite to aminonitrile and
 25 recovering from the reaction product a mixture of aliphatic azodinitrile compounds of the formula



35 wherein R_1 , R_2 , R_3 , and R_4 are each independently selected from the group consisting of acyclic aliphatic hydrocarbon radicals of 1-9 carbon atoms.

BRIEF DESCRIPTION OF THE DRAWING

Figure 1 depicts a phase diagram showing the concentration ranges of aminonitrile reactants as detailed in Example 1 which yield liquid mixtures at -15°C for a mixture of .

- Butanenitrile, 2,2'-azobis(2-methyl-
- 5 Pentanenitrile, 2,2'-azobis(2-methyl-
- Heptanenitrile, 2,2'-azobis(2-methyl-
- Heptanenitrile, 2-[(1-cyano-1-methylpropyl)azo]-2-methyl-
- Heptanenitrile, 2-[(1-cyano-1-methylbutyl)azo]-2-methyl-
- Pentanenitrile, 2-[(1-cyano-1-methylpropyl)azo]-2-methyl-

- 10 Figure 2 depicts a phase diagram showing the concentration of aminonitrile reactants which yield liquid mixtures of azodinitriles at 20°C for the same product mixture as defined for Figure 1.

Figure 3 depicts a phase diagram showing the concentration ranges of aminonitrile reactants as detailed in Example 2 which yield liquid mixtures at

- 15 -15°C for a mixture of

- Butanenitrile, 2,2'-azobis(2-methyl-
- Pentanenitrile, 2,2'-azobis(2-methyl-
- Octanenitrile, 2,2'-azobis(2-methyl-
- Octanenitrile, 2-[(1-cyano-1-methylpropyl)azo]-2-methyl-
- 20 Octanenitrile, 2-[(1-cyano-1-methylbutyl)azo]-2-methyl-
- Pentanenitrile, 2-[(1-cyano-1-methylpropyl)azo]-2-methyl-

- 25 Figure 4 depicts a phase diagram showing the concentration ranges of aminonitrile reactants as detailed in Example 3 which yield liquid mixtures at 7°C for a mixture of

- Butanenitrile, 2,2'-azobis(2-methyl-
- Heptanenitrile, 2,2'-azobis(2-methyl-
- Octanenitrile, 2,2'-azobis(2-methyl-
- Octanenitrile, 2-[(1-cyano-1-methylpropyl)azo]-2-methyl-
- 30 Octanenitrile, 2-[(1-cyano-1-methylhexyl)azo]-2-methyl-
- Heptanenitrile, 2-[(1-cyano-1-methylpropyl)azo]-2-methyl-

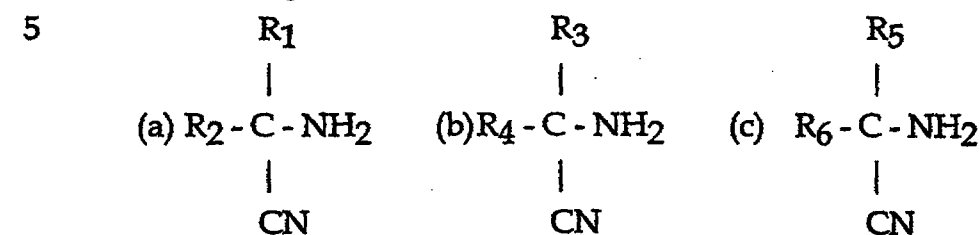
DETAILED DESCRIPTION OF THE INVENTION

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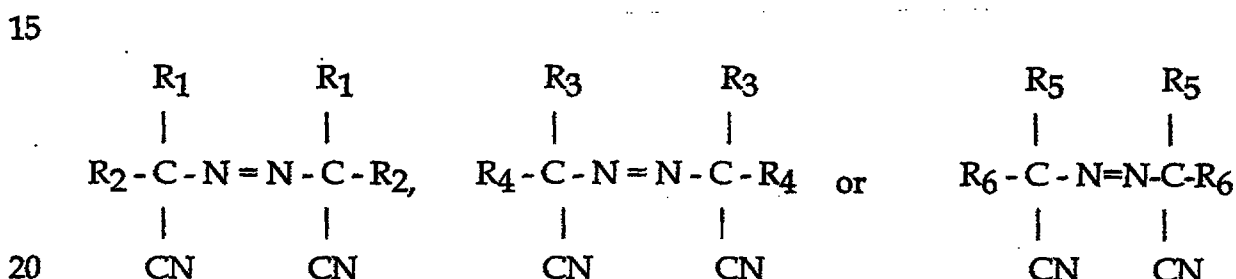
- The present invention comprises a mixture of azodinitrile compounds such that liquid behavior is obtained at temperatures at or below 25°C. Preferably the mixtures have a maximum freezing point of 0°C, and more preferably -15°C. The compositions contain both symmetrical and
- 40 asymmetrical azodinitrile compounds.

The compositions of the present invention are prepared from three or more aminonitriles. The particular azodinitrile product mixture obtained depends upon the aminonitrile starting materials employed.

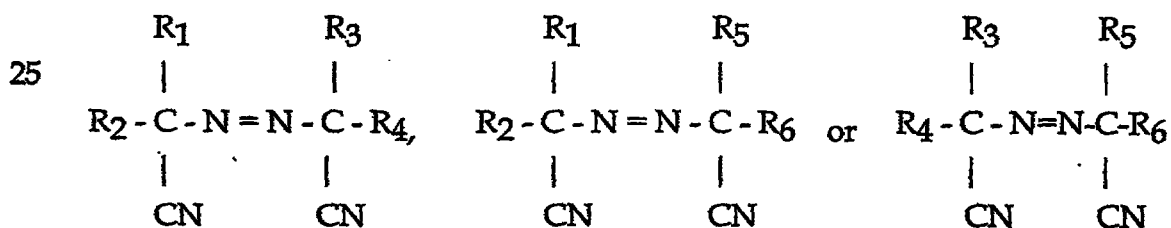
Starting with three different aminonitriles



wherein R₁, R₂, R₃, R₄, R₅, and R₆ are each independently selected from the group consisting of acyclic aliphatic hydrocarbon radicals of 1 - 9 carbon atoms results in an azodinitrile mixture comprising symmetric products such as



and asymmetric products such as



It can easily be determined that when starting with three different aminonitriles that there will be three different symmetric azodinitriles, and three different asymmetrical azodinitriles. When starting with four different aminonitriles there will be four different symmetric azodinitriles, and six different asymmetrical azodinitriles. In the general case when starting with n different aminonitriles there will be n different symmetric azodinitriles, and $n!/2(n-2)!$ different asymmetrical azodinitriles in the resulting product.

Suitable acyclic aliphatic hydrocarbon radicals for R₁, R₂, R₃, R₄, R₅ and R₆ include, for example, methylbutyro, methylpentano, methylheptano, methyloctano, ethylbutano, cyanomethylpropyl, cyanomethylbutyl, and cyanodimethylbutyl.

5 Preferred azodinitrile compositions of the present invention are liquid at and have a maximum freezing point of 25°C and include mixture A as follows:

10 A. Butanenitrile, 2,2'-azobis(2-methyl-
Pentanenitrile, 2,2'-azobis(2-methyl-
Butanenitrile, 2,2'-azobis(2-ethyl-
Pentanenitrile, 2-[(1-cyano-1-methylpropyl)azo]-2-methyl-
Pentanenitrile, 2-[(1-cyano-1-ethylpropyl)azo]-2-methyl-
Butanenitrile, 2-[(1-cyano-1-methylpropyl)azo]-2-ethyl-

15 Also preferred is the following mixture B of the present invention having a maximum freezing point of 7°C.

20 B. Butanenitrile, 2,2'-azobis(2-methyl-
Heptanenitrile, 2,2'-azobis(2-methyl-
Octanenitrile, 2,2'-azobis(2-methyl-
Octanenitrile, 2-[(1-cyano-1-methylpropyl)azo]-2-methyl-
Octanenitrile, 2-[(1-cyano-1-methylhexyl)azo]-2-methyl-
Heptanenitrile, 2-[(1-cyano-1-methylpropyl)azo]-2-methyl-

25 More preferred azodinitrile compositions of the present invention are liquid at and have a maximum freezing point of 0°C and include mixtures C and D as follows:

30 C. Pentanenitrile, 2,2'-azobis(2-methyl-
Butanenitrile, 2,2'-azobis(2-ethyl-
Heptanenitrile, 2,2'-azobis(2-methyl-
Heptanenitrile, 2-[(1-cyano-1-methylbutyl)azo]-2-methyl-
Heptanenitrile, 2-[(1-cyano-1-ethylpropyl)azo]-2-methyl-
35 Pentanenitrile, 2-[(1-cyano-1-ethylpropyl)azo]-2-methyl-

D. Butanenitrile, 2,2'-azobis(2-methyl-
Pentanenitrile, 2,2'-azobis(2-methyl-
Pentanenitrile, 2,2'-azobis(2,4-dimethyl-
40 Heptanenitrile, 2,2'-azobis(2-methyl-
Heptanenitrile, 2-[(1-cyano-1-methylpropyl)azo]-2-methyl-
Heptanenitrile, 2-[(1-cyano-1-methylbutyl)azo]-2-methyl-
Heptanenitrile, 2-[(1-cyano-1,3-dimethylbutyl)azo]-2-methyl-
Pentanenitrile, 2-[(1-cyano-1-methylbutyl)azo]-2,4-dimethyl-

Pentanenitrile, 2-[(1-cyano-1-methylpropyl)azo]-2-methyl-
 Pentanenitrile, 2-[(1-cyano-1-methylpropyl)azo]-2,4-dimethyl-

More particularly preferred are azodinitrile compositions which are
 5 liquid at and have a maximum freezing point of -15°C and include the
 following mixtures E, F and G:

- E. Butanenitrile, 2,2'-azobis(2-methyl-
 Pentanenitrile, 2,2'-azobis(2-methyl-
 10 Heptanenitrile, 2,2'-azobis(2-methyl-
 Heptanenitrile, 2-[(1-cyano-1-methylpropyl)azo]-2-methyl-
 Heptanenitrile, 2-[(1-cyano-1-methylbutyl)azo]-2-methyl-
 Pentanenitrile, 2-[(1-cyano-1-methylpropyl)azo]-2-methyl-
- 15 F. Butanenitrile, 2,2'-azobis(2-methyl-
 Pentanenitrile, 2,2'-azobis(2-methyl-
 Octanenitrile, 2,2'-azobis(2-methyl-
 Octanenitrile, 2-[(1-cyano-1-methylpropyl)azo]-2-methyl-
 Octanenitrile, 2-[(1-cyano-1-methylbutyl)azo]-2-methyl-
 20 Pentanenitrile, 2-[(1-cyano-1-methylpropyl)azo]-2-methyl-
- G. Butanenitrile, 2,2'-azobis(2-methyl-
 Pentanenitrile, 2,2'-azobis(2-methyl-
 Butanenitrile, 2,2'-azobis(2-ethyl-
 25 Heptanenitrile, 2,2'-azobis(2-methyl-
 Heptanenitrile, 2-[(1-cyano-1-methylpropyl)azo]-2-methyl-
 Heptanenitrile, 2-[(1-cyano-1-methylbutyl)azo]-2-methyl-
 Heptanenitrile, 2-[(1-cyano-1-ethylpropyl)azo]-2-methyl-
 Pentanenitrile, 2-[(1-cyano-1-methylpropyl)azo]-2-methyl-
 30 Pentanenitrile, 2-[(1-cyano-1-ethylpropyl)azo]-2-methyl-
 Butanenitrile, 2-[(1-cyano-1-methylpropyl)azo]-2-ethyl-

The ratio of starting aminonitriles can affect the physical phase of the
 resulting product mixture. Figure 1 depicts a phase diagram at -15°C for the
 above particularly preferred composition E. As detailed in Example 1, this
 35 composition is prepared from 2-amino-2methylbutyronitrile (Component A),
 2-amino-2methylpentanonitrile (Component B) and 2-amino-
 2methylheptanonitrile (Component C). The phase diagram of Figure 1 shows
 the percentages of each starting aminonitrile, which when reacted, results in a
 product composition which is liquid at -15°C . The liquid compositions
 40 roughly correspond to the lined area. This corresponds to 30-50% by weight

of component A, 5-25% by weight of component B, and 30-55% by weight of component C.

Figure 2 depicts a phase diagram at 20°C for these same components A, B, and C. The liquid compositions roughly correspond to the lined area which correlates with 20-55% by weight of component A, 5-30% by weight of component B, and 30-55% by weight of component C. Also shown are three data points (Ø) which display partial liquid behavior. For these compositions removal of solids, such as by filtration, would result in the desired liquid composition.

Figure 3 depicts a phase diagram at -15°C for the above defined composition F. As detailed in Example 2 this composition is prepared from 2-amino-2-methylbutyronitrile (Component A), 2-amino-2-methylpentanonitrile (Component B), and 2-amino-2-methyloctanonitrile (Component D). The phase diagram of Figure 3 shows the percentages of each starting aminonitrile, which when reacted, results in a product composition which is 100% liquid, or about 50% liquid, or solid at -15°C. The circles represent 100% liquid, circles with a line through them represent about 50% liquid, and the solid black circles represent 100% solid.

Figure 4 depicts a phase diagram at 7°C for the above defined composition B as detailed in Example 3. This composition is prepared from 2-amino-2-methylbutyronitrile (Component A), 2-amino-2-methylheptanonitrile (Component E) and 2-amino-2-methyloctanonitrile (Component D). This phase diagram shows the percentages of each starting aminonitrile, which when reacted, results in a product composition which is 100% liquid (circles), about 50% liquid (circles with lines through them) and solid (black circles).

The liquid azodinitrile mixtures of this invention are useful in producing ethylene copolymers. High pressure copolymerizations of ethylene and vinyl acetate, methyl methacrylate, ethyl acrylate, acyclic and methacrylic acids and salts, vinyl chloride, acrylonitrile, olefins such as propylene, butene-1 and butadiene, dibutyl maleate, carbon monoxide and the like can be carried out easily and efficiently using the azodinitrile mixtures of this invention as polymerization initiators.

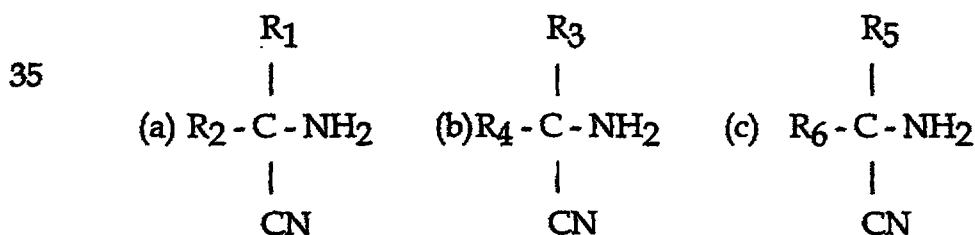
These new compositions are initiators for the polymerization or copolymerization of other unsaturated monomers such as alkenes, vinyl halides, vinyl esters, vinylidene halides, vinyl cyanides and alkenyl aromatic as well as curing agents for polyester resins, initiators for free radical initiated chemical reactions, blowing agents for producing foamed polymer and plastics and selective oxidizing agents.

Illustrative polymerizable monomers other than ethylene and ethylene-comonomers are vinyl chloride, vinylidene chloride, vinyl acetate, vinyl pyridine, vinyl pyrrolidone, vinyl carbazole, butadiene, isoprene, acrylonitrile, acyclic acid, acyclic acid esters, methacrylic acid, methacrylic acid esters, styrene, chlorostyrene and methyl styrenes.

The liquid azodinitrile mixtures of this invention are concentrated fluids and are useful full strength where desirable, convenient or necessary. Any organic solvent which is liquid under reaction conditions and inert with respect to the azodinitriles of this invention may be used as a diluent for the azodinitrile mixtures described herein. Some such suitable solvents include, for example, appropriate alcohols, aliphatic hydrocarbons, esters, glycol ethers, ketones, aliphatic petroleum naphthas, aromatic hydrocarbons, chlorinated aliphatic hydrocarbons, chlorinated aromatic hydrocarbons, amides, nitriles, or mixtures thereof.

The azodinitriles of the present invention are synthesized according to the procedure in U.S. Patent 4,028,345, the teachings of which are hereby incorporated by reference. At least three aminonitriles are reacted with sodium hypochlorite solution and quaternary amine salts. The excess hypochlorite is neutralized with sodium bisulfite and the resulting liquid azo mixture is separated from the aqueous phase, washed with bicarbonate, and dried under vacuum.

In the process of the present invention, three molecules of an amino compound selected from the group consisting of



wherein R_1, R_2, R_3, R_4, R_5 and R_6 are each independently an acyclic aliphatic hydrocarbon radical of 1-9 carbon atoms, are reacted to form the azodinitrile of the present invention. The reaction is accomplished in an aqueous medium with a metal hypochlorite represented by the formula $M(OCI)_x$, wherein M is selected from sodium; potassium and calcium and x is the valence of the M ion, and a surface active compound selected from the group consisting of an anionic, cationic, nonionic, amphoteric and mixed surface active agents or surfactants. The equivalent ratio of hypochlorite to aminonitrile is not critical. However, at equivalent ratios below 1:1 of hypochlorite to aminonitrile the yield is less desirable. At equivalent ratios above 2:1 there is no advantage. Generally, the equivalent ratio of 1:1 to 2:1 will give high yields although the ratio of 1.2:1 to 1.8:1 is preferred because of the especially high yields that result. The equivalent ratio referred to herein is defined as the equivalent of metal hypochlorite per mole of aminonitrile. An equivalent of metal hypochlorite is a mole of the hypochlorite divided by the valence of the metal. An equivalent of aminonitrile is the same as the molar amount of aminonitrile.

The amino compounds used as starting materials may be selected from the formulas given above denoted (a) (b) and (c) or combinations thereof. The amino compounds can be prepared by methods well known in the art, for example, by the method described by Anderson in U.S. Patent 2,711,405. A procedure that can be used to obtain the amino compounds involves charging an appropriate ketone to a platinum-lined pressure vessel and cooling this to dry ice-acetone temperature and then adding 5-10 grams of ammonia. Hydrogen cyanide is then introduced in portions in an amount equimolar to that of the ketone. The reaction vessel is warmed to room temperature and pressurized to 50 psig (34.5×10^4 Pa) with ammonia, heated to 40°C and held at 40°C and 50 psig (34.5×10^4 Pa) for 8 hours and finally cooled and the product is discharged from the vessel.

The hydrochlorite used in the present invention is a metal hydrochlorite represented by the formula $M(OCI)_x$ where M is selected from sodium, potassium and calcium and x is the valence of M . For reasons of convenience and economy, sodium hypochlorite is the preferred hypochlorite. Sodium hypochlorite can be prepared by passing chlorine gas

in an aqueous sodium hydroxide solution at about 9°C or it can be purchased commercially. Other hypochlorites can be prepared analogously.

5 The surface active compounds used in the present invention are defined as any compound or mixture of compounds that affects the surface tension when mixed with water and is not adversely affected in its properties if it reacts with the hypochlorite, aminonitrile, chloramine intermediate or final product of the present invention. The inclusion of a surfactant in the process of reacting aminonitriles to give an azodinitrile with hypochlorite
10 enables the reaction to proceed in strictly aqueous medium.

Surfactants for preparing emulsions are discussed by Paul Becher in "Emulsions, Theory and Practice" ACS Monograph No. 162, 1965. On pages 232-255. Becher discusses the importance of the Hydrophilic-Lipophilic
15 Balance of a surfactant (HLB) on its ability to serve as an emulsifying agent in a particular application. The HLB numbers which have been assigned to many surfactants indicate balance in their affinity for water (hydrophilic) or non-polar organic liquids (lipophilics). A high HLB number indicates high water solubility and low organic solubility, a low number indicates a high
20 organic solubility and low water solubility.

The azodinitriles of the present invention can generally be produced with surfactants or mixtures thereof within the range of about 8.0 to about 35.0 HLB range. The surfactants or surface active agents useful in the present
25 invention may be a mixture of said surfactants or surface active agents. Thus, in surfactant mixtures, one component of the mixture may have a HLB number outside the range described herein as long as the HLB of the mixture is within said range.

30 The atomospheric pressure system is entirely aqueous, requiring no organic solvent to be present as a promoter or co-solvent with water. The surfactant is mixed with the water as is the sodium hypochlorite or other metal hypochlorite and the aminonitrile is added with sufficient cooling to handle the heat load.

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During the reaction period, it is necessary to add an agent to destroy residual sodium hypochlorite, chloramines and other oxidizing impurities and thereby give a cleaner product. Sodium bisulfite serves this purpose, as

well as gaseous or liquid SO₂. Acid such as hydrochloric or sulfuric is introduced during the product workup to enhance the activity of sodium bisulfite by converting it in part or whole to sulfur dioxide, a powerful reducing agent. Other inorganic salts may be employed.

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Any concentration of sodium hypochlorite less than about 16% can be used, but 5-15% is preferred. With less than 5%, yields of product tend to drop off. However, calcium hypochlorite, which is available as a 100% active material, is diluted to attain the preferred 5-15% range. Potassium
10 hypochlorite solutions of the above concentration can also be preferred.

The time required to complete the reaction is dependent on temperature. A typical temperature for the reaction is -10°C to 30°C. At the preferred temperature range of 5°-15°C, the reaction takes from about 10
15 minutes to 30 minutes. At a temperature of -5°C, the reaction will take well over 1 hour. At 30°C, the reaction can be complete in 5 minutes. The time required for the reaction for a specific product at a specific temperature and batch size can readily be determined.

20 Azodinitriles compositions which exist in solid and liquid phases simultaneously at a particular temperature can be filtered to remove the solids. The resulting filtrate would be suitable for use as a polymerization initiator in the same manner as the compositions of the present invention.

25 The present invention is further illustrated by the examples below.

EXAMPLES

For all examples the percent solids were determined as follows. Samples were placed in glass ampoules and stored at the recorded
30 temperatures for at least 72 hours. Solids were visually estimated relative to the entire amount of sample present in the ampoule (i.e. and layer of solid occupying one tenth of the ampoule was designated 10% solids).

EXAMPLE 1

To a 1.0-liter beaker equipped with a motor-driven Teflon-coated
35 stirrer, thermocouple, 100-ml addition funnel, and situated in a dry ice-acetone cooling bath, 286.7 g (0.5665 moles) of 14.71% sodium hypochlorite were added along with 1.54 g of dioctyldimethyl ammonium chloride. The temperature was lowered to 7 °C and a mixture of aminonitriles consisting of

20g (0.20389 moles) 2-amino-2-methylbutanonitrile (MEK), 10g (0.0891 moles) amino-2-methylpentanonitrile (MPK), and 20g (0.1429 moles) 2-amino-2-methylheptanonitrile (MAK) which had been sparged of excess ammonia for one hour at 35 °C, were added over a 50 minute period. After the addition, the temperature was held at 7 °C for 30 minutes followed by neutralization of excess hypochlorite with 20 g of sodium bisulfite while the pH was maintained above 9 with the addition of 3.1g of 10N sodium hydroxide. The liquid organic layer was then separated from the more dense aqueous phase and subsequently washed with 5% sodium bicarbonate. After 2 additional washes and separations with deionized water, the crude product was dried under vacuum (10 mm Hg, 13.3×10^{-2} Pa) for 2 hours. Final purity by gas evolution and gas chromatography, 98%. Final yield 91%. Liquid at -15°C.

The product mixture contained:

Butanenitrile, 2,2'-azobis(2-methyl-
Pentanenitrile, 2,2'-azobis(2-methyl-
Heptanenitrile, 2,2'-azobis(2-methyl-
Heptanenitrile, 2-[(1-cyano-1-methylpropyl)azo]-2-methyl-
Heptanenitrile, 2-[(1-cyano-1-methylbutyl)azo]-2-methyl-
Pentanenitrile, 2-[(1-cyano-1-methylpropyl)azo]-2-methyl-

Reactants, %

MEK	70	65	55	54	50	45	40	40	37	35	33	25	25	15
MPK	15	10	20	12	12.5	10	40	20	13	5	33	25	20	20
MAK	15	25	25	34	37.5	45	20	40	50	60	33	50	55	65
%Yield	82	76	82	90	89	91	87	91	93	91	90	87	88	93
%Purity	94	97	97	98	93	95	96	96	91	90	96	95	94	88
Form	S	S	S	S	L	L	S	L	L	S	S	S	S	S

This data is represented on Figure 1. S indicates solid, L indicates liquid, and 50% L indicates 50% liquid.

The following examples were completed following the procedure given in Example 1, using the aminonitriles indicated on a percentage by weight basis. The form is indicated using the above defined abbreviations.

EXAMPLE 2

Reactant, %

2-Amino-2-methylbutyronitrile	33	40	20	40	50	25	25
2-Amino-2-methylpentanonitrile	33	20	40	40	25	50	25
2-Amino-2-methyloctanonitrile	33	40	40	20	25	25	50

% Yield	94.5	87.5	86.9	87.8	73	86.8	87
% Purity	96.3	96.0	99.6	92.0	97.3	96.9	99.2
Form	S	50%L	L	S	50%L	S	S

5 This data is represented on Figure 3.

Product Mixture:

10 Butanenitrile, 2,2'-azobis(2-methyl-
Pentanenitrile, 2,2'-azobis(2-methyl-
Octanenitrile, 2,2'-azobis(2-methyl-
Octanenitrile, 2-[(1-cyano-1-methylpropyl)azo]-2-methyl-
Octanenitrile, 2-[(1-cyano-1-methylbutyl)azo]-2-methyl-
Pentanenitrile, 2-[(1-cyano-1-methylpropyl)azo]-2-methyl-

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EXAMPLE 3

Reactants, %

2-Amino-2-methylbutyronitrile	33	20	40	40	50	25	25
2-Amino-2-methylheptanonitrile	33	40	20	40	25	50	25
2-Amino-2-methyloctanonitrile	33	40	40	20	25	25	50

20

% Yield	79	62.8	87.3	89.7	82.5	54.0	90.9
% Purity	94.8	98.13	89.4	99.9	99.9	98.1	98.3
Form	L	L	L	50%L	50%L	S	S

This data is represented in Figure 4

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Product Mixture:

30 Butanenitrile, 2,2'-azobis(2-methyl-
Heptanenitrile, 2,2'-azobis(2-methyl-
Octanenitrile, 2,2'-azobis(2-methyl-
Octanenitrile, 2-[(1-cyano-1-methylpropyl)azo]-2-methyl-
Octanenitrile, 2-[(1-cyano-1-methylhexyl)azo]-2-methyl-
Heptanenitrile, 2-[(1-cyano-1-methylpropyl)azo]-2-methyl-

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Reactants:

2-Amino-2-methylbutyronitrile	33%
2-Amino-2-methylpentanonitrile	33%
2-Amino-2-ethylbutyronitrile	33%

40

Yield = 88.3%

Purity = 93.1%

33% solid at 25°C

100% solid at 0 °C

Product Mixture:

- 5 Butanenitrile, 2,2'-azobis(2-methyl-
 Pentanenitrile, 2,2'-azobis(2-methyl-
 Butanenitrile, 2,2'-azobis(2-ethyl-
 Pentanenitrile, 2-[(1-cyano-1-methylpropyl)azo]-2-methyl-
 Pentanenitrile, 2-[(1-cyano-1-ethylpropyl)azo]-2-methyl-
 10 Butanenitrile, 2-[(1-cyano-1-methylpropyl)azo]-2-ethyl-

EXAMPLE 5Reactants:

- | | | |
|----|------------------------------------|-----|
| | 2-Amino-2-methylbutyronitrile | 21% |
| 15 | 2-Amino-2,4-dimethylpentanonitrile | 26% |
| | 2-Amino-2-methylpentanonitrile | 24% |
| | 2-Amino-2-methylheptanonitrile | 29% |

Yield 95%,

20 Purity 93%,

Liquid at 0 °C

100% solids at -15°C

Product Mixture:

- 25 Butanenitrile, 2,2'-azobis(2-methyl-
 Pentanenitrile, 2,2'-azobis(2-methyl-
 Pentanenitrile, 2,2'-azobis(2,4-dimethyl-
 Heptanenitrile, 2,2'-azobis(2-methyl-
 Heptanenitrile, 2-[(1-cyano-1-methylpropyl)azo]-2-methyl-
 30 Heptanenitrile, 2-[(1-cyano-1-methylbutyl)azo]-2-methyl-
 Heptanenitrile, 2-[(1-cyano-1,3-dimethylbutyl)azo]-2-methyl-
 Pentanenitrile, 2-[(1-cyano-1-methylbutyl)azo]-2,4-dimethyl-
 Pentanenitrile, 2-[(1-cyano-1-methylpropyl)azo]-2-methyl-
 35 Pentanenitrile, 2-[(1-cyano-1-methylpropyl)azo]-2,4-dimethyl-

EXAMPLE 6Reactants:

- | | | |
|----|--------------------------------|-----|
| | 2-Amino-2-methylpentanonitrile | 33% |
| | 2-Amino-2-methylheptanonitrile | 33% |
| 40 | 2-Amino-2-methyloctanonitrile | 33% |

Yield = 99%

Purity = 93%
 Liquid at 25°C
 100% Solid at 0°C

Product Mixture:

- 5 Pentanenitrile, 2,2'-azobis(2-methyl-
 Heptanenitrile, 2,2'-azobis(2-methyl-
 Octanenitrile, 2,2'-azobis(2-methyl-
 Octanenitrile, 2-[(1-cyano-1-methylbutyl)azo]-2-methyl-
 Octanenitrile, 2-[(1-cyano-1-methylhexyl)azo]-2-methyl-
 10 Heptanenitrile, 2-[(1-cyano-1-methylbutyl)azo]-2-methyl-

EXAMPLE 7

- | | | |
|----|--------------------------------|-----|
| | 2-Amino-2-methylbutanonitrile | 25% |
| 15 | 2-Amino-2-ethylbutanonitrile | 25% |
| | 2-Amino-2-methylpentanonitrile | 25% |
| | 2-Amino-2-methylheptanonitrile | 25% |

Yield = 85%

Purity = 97%

- 20 Liquid at 25°C
 Liquid at -15°C

Product Mixture:

- Butanenitrile, 2,2'-azobis(2-methyl-
 Pentanenitrile, 2,2'-azobis(2-methyl-
 25 Butanenitrile, 2,2'-azobis(2-ethyl-
 Heptanenitrile, 2,2'-azobis(2-methyl-
 Heptanenitrile, 2-[(1-cyano-1-methylpropyl)azo]-2-methyl-
 Heptanenitrile, 2-[(1-cyano-1-methylbutyl)azo]-2-methyl-
 Heptanenitrile, 2-[(1-cyano-1-ethylpropyl)azo]-2-methyl-
 30 Pentanenitrile, 2-[(1-cyano-1-methylpropyl)azo]-2-methyl-
 Pentanenitrile, 2-[(1-cyano-1-ethylpropyl)azo]-2-methyl-
 Butanenitrile, 2-[(1-cyano-1-methylpropyl)azo]-2-ethyl-

COMPARATIVE EXAMPLE A

Reactants:

- | | | |
|----|--------------------------------|-----|
| 35 | 2-Amino-2-methylbutyronitrile | 25% |
| | 2-Amino-2-methylpropanonitrile | 25% |
| | 2-Amino-2-methylpentanonitrile | 25% |
| | 2-Amino-2-methylheptanonitrile | 25% |

Yield 56%,
 Purity 96%, 10% solids at 25 °C
 75% solids at 0 °C
 100% solids at -15 °C

5

Product Mixture:

- Propanenitrile, 2,2'-azobis(2-methyl-
 Butanenitrile, 2,2'-azobis(2-methyl-
 Pentanenitrile, 2,2'-azobis(2-methyl-
 10 Heptanenitrile, 2,2'-azobis(2-methyl-
 Heptanenitrile, 2-[(1-cyano-1-methylethyl)azo]-2-methyl-
 Heptanenitrile, 2-[(1-cyano-1-methylpropyl)azo]-2-methyl-
 Heptanenitrile, 2-[(1-cyano-1-methylbutyl)azo]-2-methyl-
 Pentanenitrile, 2-[(1-cyano-1-methylethyl)azo]-2-methyl-
 15 Pentanenitrile, 2-[(1-cyano-1-methylpropyl)azo]-2-methyl-
 Butanenitrile, 2-[(1-cyano-1-methylethyl)azo]-2-methyl-

COMPARATIVE EXAMPLE B20 Reactants:

2-Amino-2-methylbutyronitrile	33%
1-Amino cyclohexano carbonitrile	33%
2-Amino-2-methylheptanonitrile	33%

25 Yield 80%,
 Purity 96%, 20% solids at 25 °C
 60% solids at -15°C

Product Mixture:

- 30 Butanenitrile, 2,2'-azobis(2-methyl-
 Cyclohexanecarbonitrile, 1,1'-azobis-
 Heptanenitrile, 2,2'-azobis(2-methyl-
 Heptanenitrile, 2-[(cyanocyclohexane)azo]-2-methyl-
 Heptanenitrile, 2-[(1-cyano-1-methylpropyl)azo]-2-methyl-
 35 Cyclohexanecarbonitrile, 2-[(1-cyano-1-methylpropyl)azo]-

COMPARATIVE EXAMPLE CReactants:

2-Amino-2-methylbutyronitrile	33%
40 1-Amino cyclohexano carbonitrile	33%
2-Amino-2-methylpentanonitrile	33%

Yield 70%,
Purity 94%, 30% solids at 25 °C
100% solids at -15°C

5

Product Mixture:

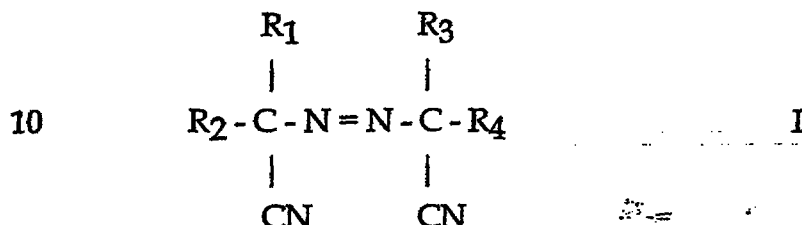
- Butanenitrile, 2,2'-azobis(2-methyl-
Cyclohexanecarbonitrile, 1,1'-azobis-
Pentanenitrile, 2,2'-azobis(2-methyl-
10 Cyclohexanecarbonitrile, 2-[(1-cyano-1-methylpropyl)azo]-
Cyclohexanecarbonitrile, 2-[(1-cyano-1-methylbutyl)azo]-
Pentanenitrile, 2-[(1-cyano-1-methylpropyl)azo]-2-methyl-

CH-2533



What is claimed is:

- 5 1. An azodinitrile composition comprising a mixture of at least six different azodinitriles of formula I



wherein

- 15 R₁, R₂, R₃, and R₄ are each independently selected from the group consisting of acyclic aliphatic hydrocarbon radicals of 1 - 9 carbon atoms, said mixture having a freezing point of a maximum of 25°C.

- 20 2. The composition of Claim 1 having a freezing point of a maximum of 0 °C.

3. The composition of Claim 1 having a freezing point of a maximum of -15 °C.

- 25 4. The composition of Claim 1 wherein R₁, R₂, R₃, and R₄ are each independently selected from the group consisting of methylbutyro, methylpentano, methylheptano, methyloctano, ethylbutano, cyanomethylpropyl, cyanomethylbutyl, and cyanodimethylbutyl.

- 30 5. The composition of Claim 1 selected from the group consisting of mixtures A and B:

- A. Butanenitrile, 2,2'-azobis(2-methyl-
 Pentanenitrile, 2,2'-azobis(2-methyl-
 35 Butanenitrile, 2,2'-azobis(2-ethyl-
 Pentanenitrile, 2-[(1-cyano-1-methylpropyl)azo]-2-methyl-
 Pentanenitrile, 2-[(1-cyano-1-ethylpropyl)azo]-2-methyl-
 Butanenitrile, 2-[(1-cyano-1-methylpropyl)azo]-2-ethyl-

- 40 B. Butanenitrile, 2,2'-azobis(2-methyl-





- 5 Heptanenitrile, 2,2'-azobis(2-methyl-
Octanenitrile, 2,2'-azobis(2-methyl-
Octanenitrile, 2-[(1-cyano-1-methylpropyl)azo]-2-methyl-
Octanenitrile, 2-[(1-cyano-1-methylhexyl)azo]-2-methyl-
Heptanenitrile, 2-[(1-cyano-1-methylpropyl)azo]-2-methyl-

6. The composition of Claim 2 selected from the group consisting of mixtures C and D

- 10 C. Pentanenitrile, 2,2'-azobis(2-methyl-
Butanenitrile, 2,2'-azobis(2-ethyl-
Heptanenitrile, 2,2'-azobis(2-methyl-
Heptanenitrile, 2-[(1-cyano-1-methylbutyl)azo]-2-methyl-
Heptanenitrile, 2-[(1-cyano-1-ethylpropyl)azo]-2-methyl-
15 Pentanenitrile, 2-[(1-cyano-1-ethylpropyl)azo]-2-methyl-
- D. Butanenitrile, 2,2'-azobis(2-methyl-
Pentanenitrile, 2,2'-azobis(2-methyl-
Pentanenitrile, 2,2'-azobis(2,4-dimethyl-
Heptanenitrile, 2,2'-azobis(2-methyl-
20 Heptanenitrile, 2-[(1-cyano-1-methylpropyl)azo]-2-methyl-
Heptanenitrile, 2-[(1-cyano-1-methylbutyl)azo]-2-methyl-
Heptanenitrile, 2-[(1-cyano-1,3-dimethylbutyl)azo]-2-methyl-
Pentanenitrile, 2-[(1-cyano-1-methylbutyl)azo]-2,4-dimethyl-
Pentanenitrile, 2-[(1-cyano-1-methylpropyl)azo]-2-methyl-
25 Pentanenitrile, 2-[(1-cyano-1-methylpropyl)azo]-2,4-dimethyl-

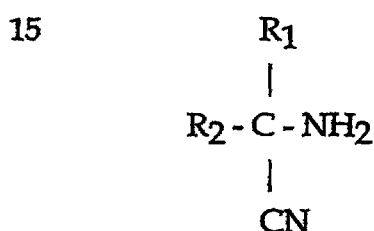
7. The composition of Claim 3 selected from the group consisting of mixtures E, F and G.

- 30 E. Butanenitrile, 2,2'-azobis(2-methyl-
Pentanenitrile, 2,2'-azobis(2-methyl-
Heptanenitrile, 2,2'-azobis(2-methyl-
Heptanenitrile, 2-[(1-cyano-1-methylpropyl)azo]-2-methyl-
Heptanenitrile, 2-[(1-cyano-1-methylbutyl)azo]-2-methyl-
35 Pentanenitrile, 2-[(1-cyano-1-methylpropyl)azo]-2-methyl-
- F. Butanenitrile, 2,2'-azobis(2-methyl-
Pentanenitrile, 2,2'-azobis(2-methyl-
Octanenitrile, 2,2'-azobis(2-methyl-
Octanenitrile, 2-[(1-cyano-1-methylpropyl)azo]-2-methyl-
40 Octanenitrile, 2-[(1-cyano-1-methylbutyl)azo]-2-methyl-
Pentanenitrile, 2-[(1-cyano-1-methylpropyl)azo]-2-methyl-



- G. Butanenitrile, 2,2'-azobis(2-methyl-
 Pentanenitrile, 2,2'-azobis(2-methyl-
 Butanenitrile, 2,2'-azobis(2-ethyl-
 Heptanenitrile, 2,2'-azobis(2-methyl-
 5 Heptanenitrile, 2-[(1-cyano-1-methylpropyl)azo]-2-methyl-
 Heptanenitrile, 2-[(1-cyano-1-methylbutyl)azo]-2-methyl-
 Heptanenitrile, 2-[(1-cyano-1-ethylpropyl)azo]-2-methyl-
 Pentanenitrile, 2-[(1-cyano-1-methylpropyl)azo]-2-methyl-
 Pentanenitrile, 2-[(1-cyano-1-ethylpropyl)azo]-2-methyl-
 10 Butanenitrile, 2-[(1-cyano-1-methylpropyl)azo]-2-ethyl-

8. A process for the preparation of azodinitrile mixtures comprising reacting three or more different aminonitriles of formula



20

wherein

R_1 and R_2 are each independently selected from the group consisting of acyclic aliphatic hydrocarbon radicals of 1 - 9 carbon atoms, with 5% to 15% by weight based on the reaction mixture of a metal hypochlorite, $M(OCl)_x$, wherein M is selected from the group consisting of

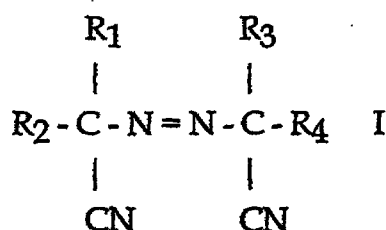
25 sodium, potassium, or calcium, and x is the valence of M ,

in an aqueous medium in the presence of 0.25 to 10% by weight based on the weight of aminonitrile of a surface active compound or mixtures thereof having a hydrophilic-lipophilic balance of 8 to 35 at a temperature of

30 $-10^\circ C$ to $30^\circ C$,

said metal hypochlorite and alpha-aminonitrile being present in an equivalent ratio of from 1:1 to 2:1 of hypochlorite to aminonitrile and recovering from the reaction product a mixture of aliphatic azodinitrile compounds of the formula

35



40

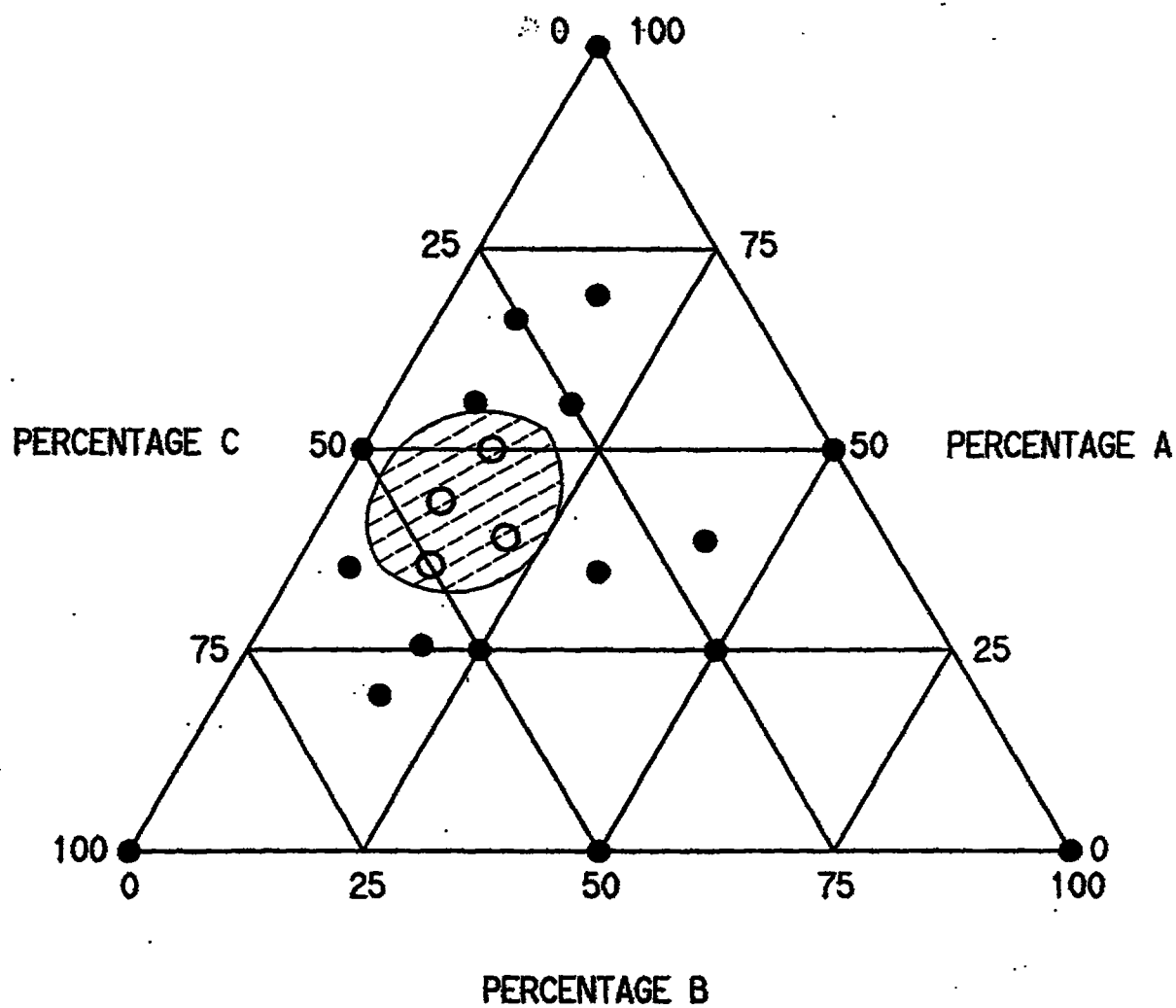




wherein R₁, R₂, R₃, and R₄ are each independently selected from the group consisting of acyclic aliphatic hydrocarbon radicals of 1-9 carbon atoms.

1/4

FIG. 1



2/4

FIG. 2

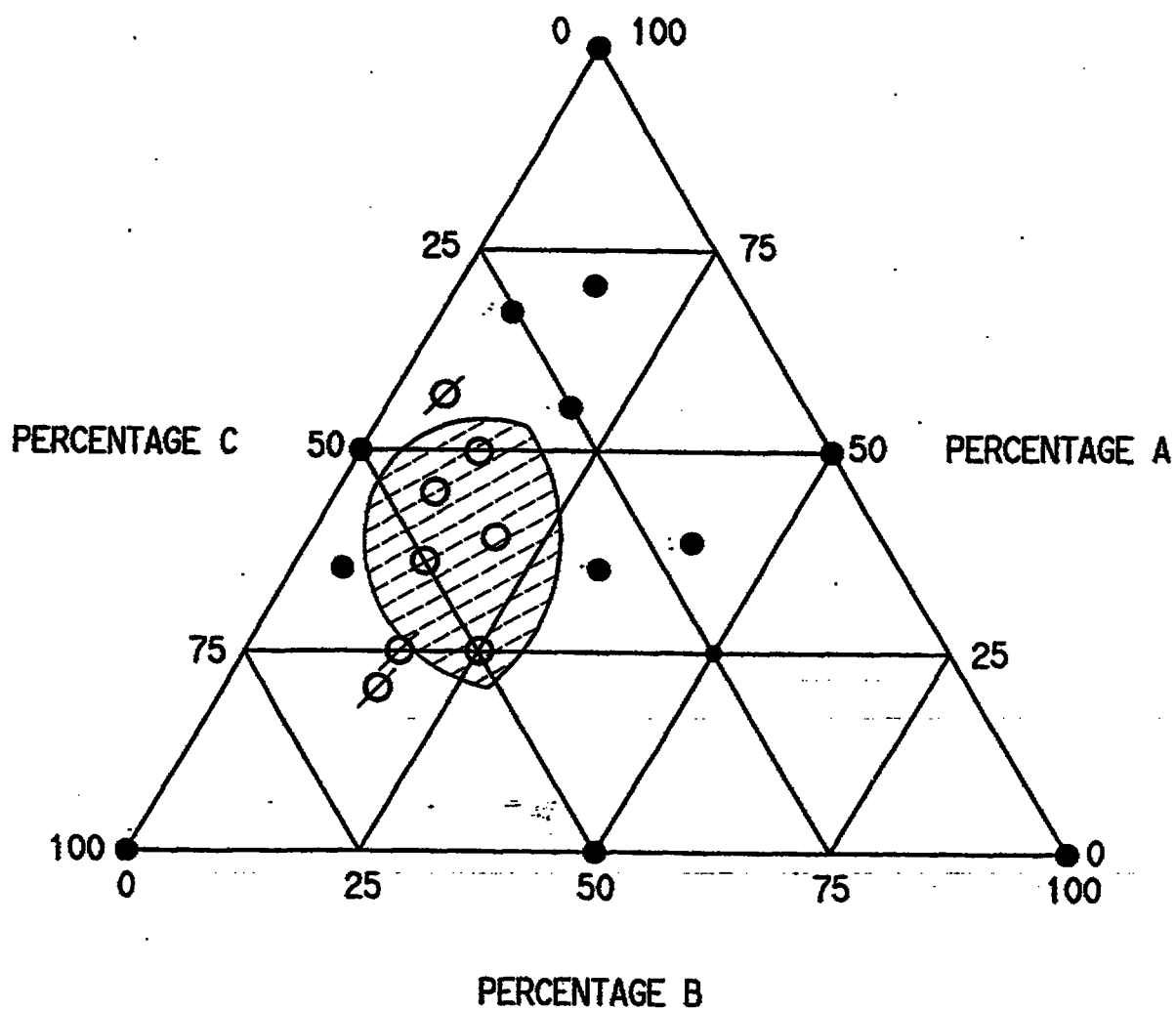


FIG. 3

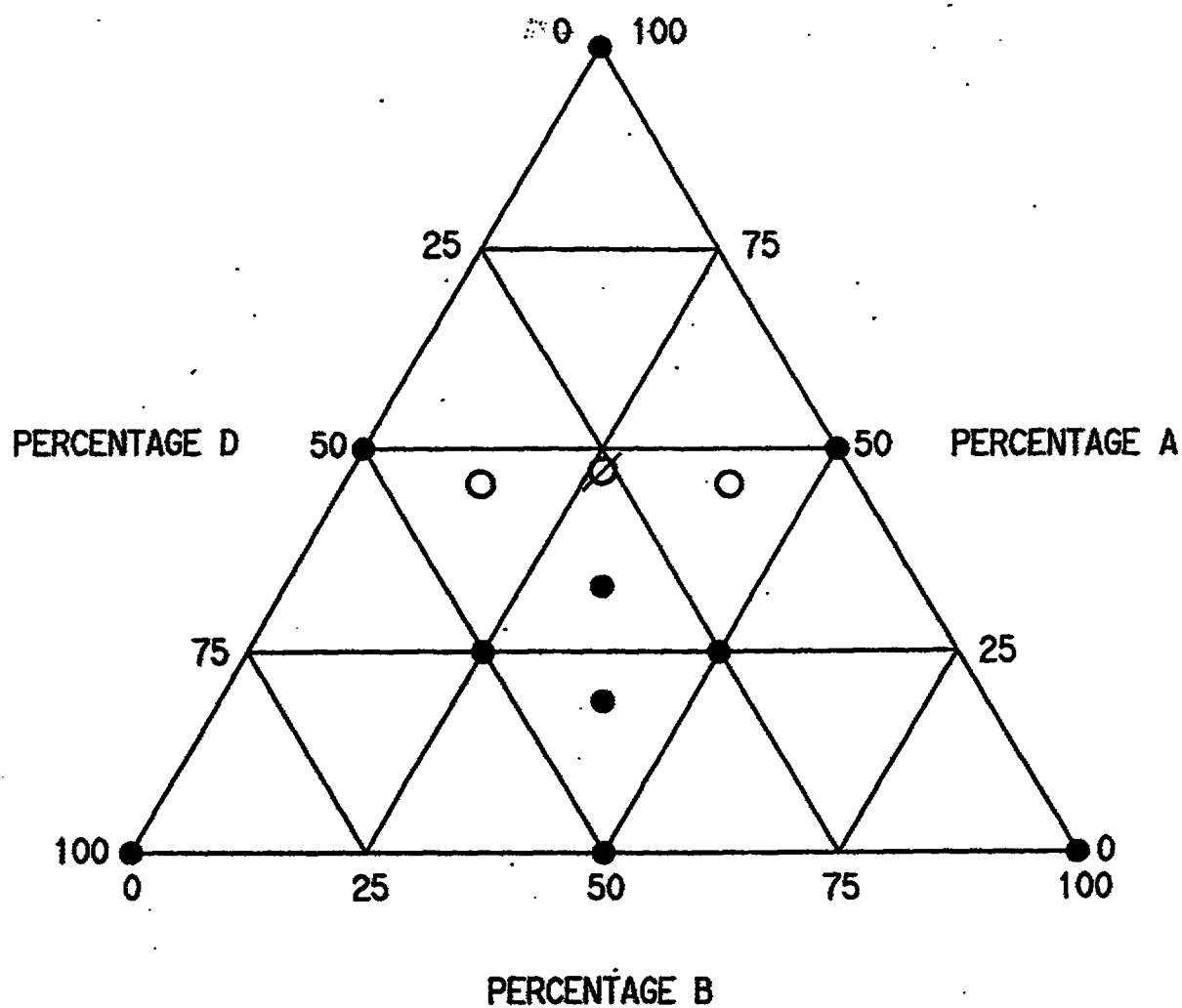
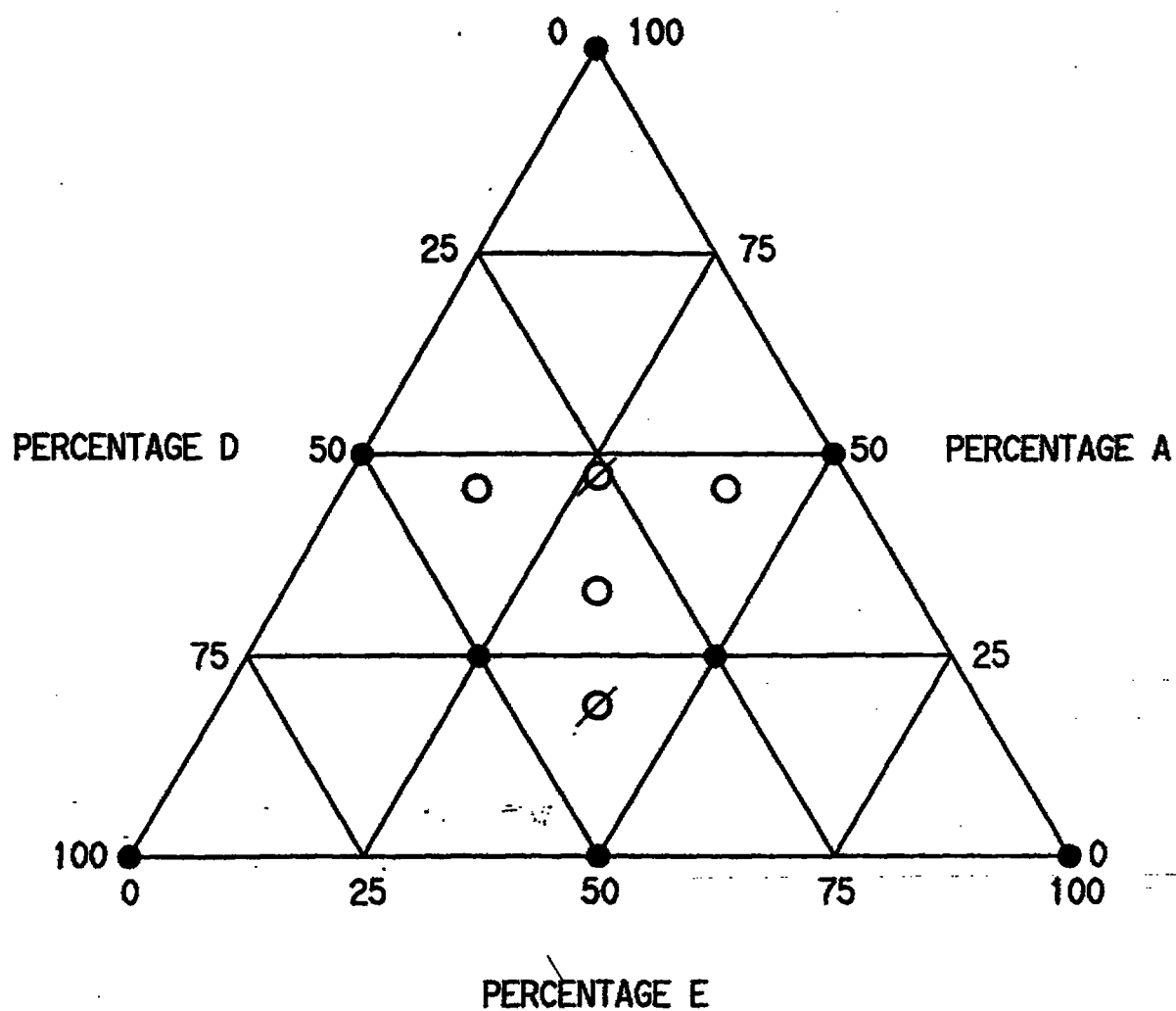


FIG. 4

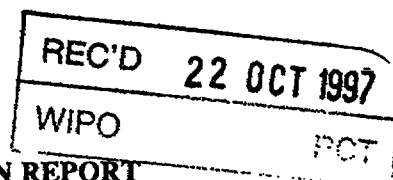


PATENT COOPERATION TREATY

PCT

INTERNATIONAL PRELIMINARY EXAMINATION REPORT

(PCT Article 36 and Rule 70)



Applicant's or agent's file reference CH2533A	FOR FURTHER ACTION See Notification of Transmittal of International Preliminary Examination Report (Form PCT/IPEA/416)	
International application No. PCT/US 96/ 13876	International filing date (day/month/year) 29/08/1996	Priority date (day/month/year) 29/08/1995
International Patent Classification (IPC) or national classification and IPC C07C255/65		
Applicant E. I. DU PONT DE NEMOURS AND COMPANY		

1. This international preliminary examination report has been prepared by this International Preliminary Examining Authority and is transmitted to the applicant according to Article 36.

2. This REPORT consists of a total of 7 sheets, including this cover sheet.

☒ This report is also accompanied by ANNEXES, i.e., sheets of the description, claims and/or drawings which have been amended and are the basis for this report and/or sheets containing rectifications made before this Authority (see Rule 70.16 and Section 607 of the Administrative Instructions under the PCT).

These annexes consists of a total of 2 sheets.

3. This report contains indications and corresponding pages relating to the following items:

- I ☒ Basis of the report
- II ☐ Priority
- III ☐ Non-establishment of opinion with regard to novelty, inventive step and industrial applicability
- IV ☐ Lack of unity of invention
- V ☒ Reasoned statement under Article 35(2) with regard to novelty, inventive step or industrial applicability; citations and explanations supporting such statement
- VI ☐ Certain documents cited
- VII ☐ Certain defects in the international application
- VIII ☒ Certain observations on the international application

Date of submission of the demand 24/02/1997	Date of completion of this report 22.09.97
Name and mailing address of the IPEA/  European Patent Office D-80298 Munich Tel. (+49-89) 2399-0, Tx: 523656 epmu d Fax: (+49-89) 2399-4465	Authorized officer  Telephone No.

INTERNATIONAL PRELIMINARY EXAMINATION REPORT

Intern. application No.

PCT/US96/13876

I. Basis of the report

1. This report has been drawn up on the basis of (Replacement sheets which have been furnished to the receiving Office in response to an invitation under Article 14 are referred to in this report as "originally filed" and are not annexed to the report since they do not contain amendments.):

☐ the international application as originally filed.

☒ the description, pages 1-18 _____, as originally filed,
pages _____, filed with the demand,
pages _____, filed with the letter of _____,
pages _____, filed with the letter of _____.

☒ the claims, Nos. 6, 7(part), 8(part) _____, as originally filed,
Nos. _____, as amended under Article 19,
Nos. _____, filed with the demand,
Nos. 1-4, 5(part), 7(part), 8(part) _____, filed with the letter of 22/07/97,
Nos. _____, filed with the letter of _____.

☒ the drawings, sheets/fig 1/4-4/4 _____, as originally filed,
sheets/fig _____, filed with the demand,
sheets/fig _____, filed with the letter of _____,
sheets/fig _____, filed with the letter of _____.

2. The amendments have resulted in the cancellation of:

☐ the description, pages _____.
☐ the claims, Nos. _____.
☐ the drawings, sheets/fig _____.

3. ☐ This report has been established as if (some of) the amendments had not been made, since they have been considered to go beyond the disclosure as filed (Rule 70.2(c)):

4. Additional observations, if necessary:

INTERNATIONAL PRELIMINARY EXAMINATION REPORT

Intern. application No.

PCT/US96/13876

V. Reasoned statement under Article 35(2) with regard to novelty, inventive step and industrial applicability; citations and explanations supporting such statement

1. STATEMENT

Novelty (N)	Claims 1-8 _____	YES
	Claims _____	NO
Inventive Step (IS)	Claims _____	YES
	Claims 1-8 _____	NO
Industrial Applicability (IA)	Claims 1-8 _____	YES
	Claims _____	NO

2. CITATIONS AND EXPLANATIONS

1).

The closest prior art is considered to be US-A-3 987 025 (D3). This document is concerned with the production of liquid azonitrile mixtures. The mixtures specifically exemplified are obtained through the reaction of two different dinitriles.

1). The following documents have been considered for the purposes of this report:

D1 = US-A-3 783 148

D2 = US-A-4 028 345

D3 = US-A-3 987 025

2). The definition of the compositions given in claims 1-4 is not clear (see also section 8 of this written opinion). Because of this lack of clarity claim 1 is considered as not fulfilling the requirements of Art. 33(2) PCT for the following reasons. Document D3 discloses, in mixture A., a composition containing 2,2'-azobis-

(2-methylbutyronitrile), 2,2'-azobis(2-methylhexanonitrile) and 2-[(1-cyano-1methylpropyl)azo]-2-methylhexanonitrile. Each of these compounds can have the conformations E or Z about the N=N double bond. As such these mixtures contain six "different" azo-compounds. On top of this the carbon atoms connected to the azo group contains four different substituents and is therefore chiral (there are therefore also in this respect "different" compounds). This argument applies to all other compositions disclosed in D3..

- 3). The problem to be solved by the present application is to provide azodinitrile compositions, for use as free-radical initiating agents, which are liquid below 25°C (more preferred is below 0°C or even more preferred below -15°C).

The solution is to provide a mixture of six or more different azodinitriles.

- 4). D1 discloses the production of symmetrical azodinitriles by reacting an alpha-aminonitrile in the presence of metal hypochlorite.

- 5). D2 discloses the production of aliphatic azodinitriles by reacting two different alpha-aminonitrile in the presence of metalhypochlorite and a surface active agent at a temperature from -10 to 30°C.

- 6). D3 discloses liquid mixture of symmetrical and asymmetrical azonitriles which have a maximum freezing point of 25°C. Such a mixture can be prepared by reacting two different alpha-aminonitrile in the presence of hypochlorite (see example 1). This document mentions that these compounds "are generally liquid at from 25 to -30°C, the preferred mixtures being liquid at 0°C. With the most preferred being liquid at -20°C" (col.4, lines 32-35). This document further mentions that "it is also

INTERNATIONAL PRELIMINARY EXAMINATION REPORT

possible to prepare liquid azonitrile mixtures when more than two aminonitrile starting materials are used" (col.8, lines 17-21). The document furthermore states that "as the number of aminonitrile starting materials is increased, there is a numerical increase in the number of symmetrical azonitriles .. and an exponential increase in the number of un-symmetrical azonitriles" (col.8, lines 21-25). The document then continues to mention that "with the increase in the number of compounds it is evident that liquid azonitrile mixtures will be highly favoured as the number of aminonitriles is increased above two" (col.8, lines 25-30).

- 7). The closest prior art is considered to be D3. This document discloses mixtures containing three different azodinitriles. The mixtures disclosed having freezing points of 25°C to -25°C.
- 8). The only difference between claim 1 and D3 is that there are now six or more compounds in the mixture.
- 9), It is well known that mixtures of compounds have lower freezing points than pure compounds. There is furthermore a clear pointer in D3 that such mixtures are envisaged as being desirable (see paragraph 5 above).
The applicant himself has acknowledged on page 3 of his reply dated 22/07/97 that not all mixtures obtained from the reaction of 3-dinitriles produce a mixture possessing the desired property as claimed in claim 1 (see comparative example C on P17-18 of the description). He further goes on to say that the ratio of the three components is important and that only specific ranges provide the desired mixtures. It is considered obvious for a man skilled in the art that a number of compositions obtained from three dinitriles would have been liquid (see D3 as cited above). The present application is

INTERNATIONAL PRELIMINARY EXAMINATION REPORT

therefore considered not to satisfy the criterion set forth in Art 33(3) PCT because the subject-matter of Claims 1-7 does not involve an inventive step (Rule 65(1)(2) PCT).

- 10). The process of claim 8 is analogous to the process known from D3. It does not have any characterising features which distinguish it from the prior art. It is therefore considered that this claim can only be considered as satisfying Art 33(3) PCT when the compound claims 1-7 satisfy this criterion.

VIII. Certain observations on the international application

The following observations on the clarity of the claims, description, and drawings or on the question whether the claims are fully supported by the description, are made:

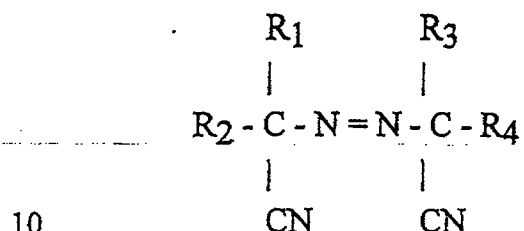
- 1). Claim 1 still defines the mixtures as containing "six azodinitriles of different formula I". This does not exclude E- and Z-isomers since these can also be considered "different". The same applies to the carbon atoms to which R1-R4 are attached since these can be stereoisomeric. These stereoisomers are also "different". Claim 1 is therefore considered not to fulfil the requirements of Art. 6 PCT.
- 2). Claims 1-3 furthermore do not meet the requirements of Article 6 PCT because the matter for which protection is sought is defined in terms of the result to be achieved [using the feature "having a freezing point of a maximum of 25°C" (0°C or -15°C) respectively]. To obtain such mixtures would require extensive experimentation (see also the examples in the description of which only a few of the examples produce mixtures fulfilling these criteria). In the present instance, therefore, such a formulation is not allowable because it seems possible to define the subject-matter in more concrete terms, viz. in terms of how the result is to be achieved.



CH-2533

What is claimed is:

1. An azodinitrile composition comprising a mixture of at least six
5 azodinitriles of different formula I



wherein

- 15 R_1 , R_2 , R_3 , and R_4 are each independently selected from the group consisting of acyclic aliphatic hydrocarbon radicals of 1 - 9 carbon atoms, cyanomethylpropyl, cyanomethylbutyl and cyanodimethylbutyl, said mixture having a freezing point of a maximum of 25°C.

2. The composition of Claim 1 having a freezing point of a maximum of 0°C.

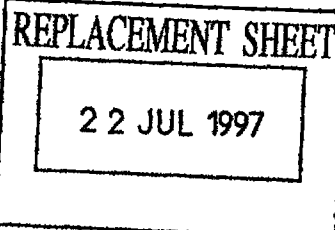
- 20 3. The composition of Claim 1 having a freezing point of a maximum of -15°C.

4. The composition of Claim 1 wherein R_1 , R_2 , R_3 , and R_4 are each independently selected from the group consisting of methylbutyl, methylpentyl, methylheptyl, methyloctyl, ethylbutyl, cyanomethylpropyl, cyanomethylbutyl, and cyanodimethylbutyl.

- 25 5. The composition of Claim 1 selected from the group consisting of mixtures A and B:

- A. Butanenitrile, 2,2'-azobis(2-methyl-
Pentanenitrile, 2,2'-azobis(2-methyl-
Butanenitrile, 2,2'-azobis(2-ethyl-
30 Pentanenitrile, 2-[(1-cyano-1-methylpropyl)azo]-2-methyl-
Pentanenitrile, 2-[(1-cyano-1-ethylpropyl)azo]-2-methyl-
Butanenitrile, 2-[(1-cyano-1-methylpropyl)azo]-2-ethyl-

- B. Butanenitrile, 2,2'-azobis(2-methyl-



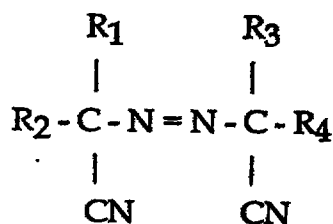


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International Bureau

INTERNATIONAL APPLICATION PUBLISHED UNDER THE PATENT COOPERATION TREATY (PCT)

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(54) Title: MULTI-COMPONENT LIQUID AZODINITRILE MIXTURES (57) Abstract An azodinitrile composition comprising a mixture of at least six different azodinitriles of formula (I), wherein R ₁ , R ₂ , R ₃ , and R ₄ are each independently acyclic aliphatic hydrocarbon radicals of 1 - 9 carbon atoms, said mixture having a freezing point of a maximum of 25 °C and a process for its preparation are disclosed.		



(I)