

SOLID ORGANIC PEROXIDE COMPOSITION

The invention relates to a solid organic peroxide composition that can be suitably stored and transported in stainless steel containers.

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Liquid organic peroxide compositions, generally containing organic peroxide diluted with an organic solvent or suspended or emulsified in water, are conventionally stored and transported in HDPE-based containers or stainless steel containers. Corrosion of the so-filled stainless steel containers is not an issue.

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Solid organic peroxides are conventionally stored and transported in HDPE-based containers or LDPE-based bags. However, handling of small LDPE bags is labor intensive and may cause dusting, with consequential safety issues, whereas transportation of solid peroxydicarbonates and diacyl peroxides in big bags (super sacks or flexible IBCs) is not allowed by UN regulations. Furthermore, safety requirements on industrial plants and factories may require a container to have an electrical resistivity that can only be obtained with stainless steel containers. In addition, the durability of HDPE containers is limited to about 5 years.

20 In storing and transporting certain solid organic peroxides in stainless steel containers, it has turned out that the containers tend to corrode. This corrosion not only limits the lifetime of the container, but also releases iron ions (and potentially other metal ions) in the peroxide. Transition metals, such as iron, are known to catalyse peroxide decomposition, which is evidently undesired.

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It is therefore an object of the present invention to provide a solid organic peroxide composition that can be suitably stored and transported in stainless steel containers, without corrosion of the container.

30 Investigations towards the cause of the corrosion have resulted in the conclusion that it is the residual chloroformate or acid chloride in the organic peroxide which, in combination with moisture or residual water in the product, leads to the formation of HCl fumes. These HCl fumes lead to corrosion of the stainless steel container.

35 Solid organic peroxides that are prepared from chloroformates are peroxydicarbonates. Peroxydicarbonates are conventionally prepared by reacting a chloroformate with

hydrogen peroxide in alkaline medium.

Solid organic peroxides that are prepared from acid chlorides are diacyl peroxides. Diacyl peroxides are conventionally prepared by reacting an acid chloride with
5 hydrogen peroxide in alkaline medium.

Upon transport and storage of these solid organic peroxides in stainless steel containers, the HCl fumes tend to corrode the stainless steel in the presence of oxygen and moisture.

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It has now been found that solid peroxydicarbonate and diacylperoxide compositions that contain a small amount of an HCl scavenger can be safely stored and transported in stainless steel containers without leading to corrosion.

15 At the same time, these compositions are safe and stable, meaning that the HCl scavenger does not promote the decomposition of the organic peroxide, can be mixed with the organic peroxide without significant segregation, and remains flowable.

The present invention therefore relates to an organic peroxide composition comprising:

- 20 - at least 40 wt% - based on the weight of the entire composition - of an organic peroxide that is solid at room temperature, said organic peroxide being selected from peroxydicarbonates and diacylperoxides,
- 0.001-5 wt% - based on the weight of organic peroxide in the composition - of an HCl scavenger that is solid at room temperature.

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The invention also relates to a packaged organic peroxide composition comprising said organic peroxide composition in a stainless steel container.

Suitable solid HCl scavengers include

- 30 - metal carboxylates, which include the metal salts of mono-, di- and tri-carboxylic acids. Examples of metal salts of mono-carboxylic acids are metal stearates, metal lactates, and metal lactylates, such as Ca, Mg, Zn, Al, Na, K, and Li stearates, lactates, or lactylates; more preferably calcium stearate and calcium stearyl-2-lactylate. Examples of metal salts of dicarboxylic acids are metal hexahydrophthalates or metal
35 bicyclic[2.2.1]heptane dicarboxylates, such as Ca, Mg, Zn, Al, Na, K, and Li hexahydrophthalates or bicyclic[2.2.1]heptane dicarboxylates. Examples of metal salts

of tricarboxylic acids are metal citrates, such as Ca, Mg, Zn, Al, Na, K, and Li mono-, di-, and tricitrates, more in particular sodium monocitrate or potassium monocitrate.

- metal carbonates and metal bicarbonates, such as Ca, Mg, Zn, Na, K, or Li carbonate, or Na, K, or Li bicarbonate. A preferred (bi)carbonate is calcium carbonate;

5 - metal oxides, such as CaO, MgO, and ZnO, with CaO being the preferred metal oxide;

- metal silicates, such as calcium silicate or magnesium silicate

- anionic clays;

- and combinations thereof.

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Anionic clays - also called hydrotalcites or layered double hydroxides - have a crystal structure consisting of positively charged layers built up of specific combinations of divalent and trivalent metal hydroxides between which there are anions and water molecules. Hydrotalcite is an example of a naturally occurring anionic clay, in which
15 carbonate is the predominant anion and the layers contain Mg and Al. Various natural, synthetic, and modified forms with other di- and/or trivalent metals and/or anions (including nitrate anions, organic anions and pillaring anions) are known. The anionic clay should be able to exchange the anions in its interlayer for chloride anions. Preferred anionic clays are anionic clays with carbonate or hydroxide anions in the
20 interlayers.

The metal carboxylates, metal oxides, metal silicates, and metal (bi)carbonates neutralize the HCl. The anionic clays exchange anions (e.g. carbonate) for chloride ions, thereby encapsulating the chloride ions in the clay structure.

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Suitable solid peroxydicarbonates are dicetyl peroxydicarbonate, dimyristyl peroxydicarbonate, dicyclohexyl peroxydicarbonate, and di(tert-butylcyclohexyl) peroxydicarbonate.

30 Suitable solid diacylperoxides include aliphatic and aromatic diacylperoxides. Preferred aliphatic diacylperoxides are those with aliphatic chains of at least 10 carbon atoms, such as dilauroyl peroxide.

Preferred aromatic diacylperoxides include dibenzoyl peroxide and substituted
35 dibenzoyl peroxides, such as di(2-methylbenzoyl)peroxide, di(4-methylbenzoyl)peroxide, and di(2,4-dichlorobenzoyl)peroxide.

Peroxydicarbonates and diacylperoxides find use in processes for enhancing the melt strength of polypropylene (towards high melt strength polypropylene; HMS-PP) by extruding said polypropylene in the presence of the peroxydicarbonate or diacylperoxide. The present invention therefore also relates to the use of the composition according to the present invention for the modification of polypropylene, such as the production of HMS-PP.

Acid scavengers like calcium stearate, calcium stearyl-2-lactylate, calcium lactate, and/or anionic clays are already used as additives in a polypropylene (modification) process for the purpose of trapping acidic catalyst residues that may deteriorate the polypropylene. This means that the introduction of calcium stearate, calcium stearyl-2-lactylate, calcium lactate, and/or anionic clays, but also calcium carbonate (the carbonate is released as CO₂ during polypropylene modification) via the composition according to the present invention, does not introduce any new materials in the modified polypropylene. Therefore, calcium stearate, calcium stearyl-2-lactylate, calcium lactate, anionic clays, and/or calcium carbonate are preferred HCl scavengers in the composition of the present invention.

Peroxydicarbonates specifically suitable for this polypropylene modification are dicetyl peroxydicarbonate, dimyristyl peroxydicarbonate, and di(tert-butylcyclohexyl) peroxydicarbonate. A diacyl peroxides specifically suitable for this polypropylene modification is di(4-methylbenzoyl)peroxide. Therefore, in a preferred embodiment, the composition according to the present invention contains one or more of these peroxydicarbonates and diacyl peroxides.

Even more preferred are compositions containing a peroxydicarbonate or diacyl peroxide selected from the group consisting of dicetyl peroxydicarbonate, dimyristyl peroxydicarbonate, di(tert-butylcyclohexyl) peroxydicarbonate, and di(4-methylbenzoyl)peroxide in combination with an HCl scavenger selected from the group consisting of calcium stearate, calcium stearyl-2-lactylate, anionic clay, and/or calcium carbonate.

The most preferred organic peroxides are dicetyl peroxydicarbonate and di(4-methylbenzoyl)peroxide.

The organic peroxide composition according to the present invention comprises at least 40 wt%, preferably at least 50 wt%, more preferably at least 60 wt%, even more preferably at least 85 wt%, and most preferably at least 90 wt% - based on the weight of the entire composition - of an organic peroxide that is solid at room temperature,
5 said organic peroxide being selected from peroxydicarbonates and diacylperoxides.

The organic peroxide composition according to the present invention comprises 0.001-5 wt%, preferably 0.005-2.5 wt%, and most preferably 0.01-0.5 wt% - based on the weight of organic peroxide in the composition - of an HCl scavenger that is solid at
10 room temperature.

Other components that can be present in the organic peroxide composition are water, phlegmatizers or diluents, or additives, provided that the organic peroxide composition remains a solid material.

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The composition according to the present invention can have the form of a physical mixture of organic peroxide particles and HCl scavenger particles.

Alternatively, it has the form of particles (powder, flakes, granules, pills, pellets, or other solid particle form) containing both the peroxydicarbonate and HCl scavenger in one
20 particle.

The organic peroxide composition according to the present invention can be prepared in several ways.

25 In one embodiment, the organic peroxide and the HCl scavenger are physically mixed. This mixing can be performed before addition of the composition to the stainless steel container for storage and transport. Alternatively, the organic peroxide and the HCl scavenger can be added to the stainless steel container individually, after which the ingredients are mixed inside the container.

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Suitable mixing equipment includes tumble mixers and rotating screw mixers (e.g. Nauta mixers). Alternatively, the solids can be dosed by in-line mixing, e.g. by dosing to a rotating screw mixer or a conveying screw, under continuous mixing.

35 In a second embodiment, the HCl scavenger is added to the reactor in which the organic peroxide is produced. According to this embodiment, the HCl scavenger can be

present during the reaction of the chloroformate or acid chloride with hydrogen peroxide in alkaline medium, or can be added to the reactor after the reaction has been completed.

- 5 In a third embodiment, the HCl scavenger is added to molten organic peroxide. The resulting product can then be transformed into solid particles by, e.g. extrusion or granulation.

10 The present invention also relates to a packaged organic peroxide composition comprising the organic peroxide composition as described above in a stainless steel container.

The present invention also relates to a stainless steel container comprising:

- 15 - a solid composition comprising at least 40 wt%, preferably at least 50 wt%, more preferably at least 60 wt%, even more preferably at least 85 wt%, and most preferably at least 90 wt% - based on the weight of the entire solid composition - of an organic peroxide that is solid at room temperature and is selected from peroxydicarbonates and diacylperoxides,
- an HCl scavenger that is solid at room temperature, in an amount of at least 0.001 wt%, preferably 0.001-5 wt%, more preferably 0.005-2.5 wt%, and most preferably 0.01-0.5 wt% - based on the weight of organic peroxide present in the container.

According to one embodiment, the HCl scavenger and the solid composition can be admixed according to one of the preparation embodiments described above in order to form an organic peroxide composition according to the present invention. In another
25 embodiment, the HCl scavenger is not admixed with the solid composition, but is present in a separate compartment within the container. Said separate compartment should be permeable to HCl fumes.

As explained above, the filled stainless steel container will not be subject to corrosion
30 by its contents. The term "container" refers to any type of packaging that can be closed and can be used to store and transport solid organic peroxide compositions.

The container preferably has a size of 200 to 4000 liters, more preferably 500-1500 liters. This includes stainless steel intermediate bulk containers (IBC's). A specifically preferred type of container is a Cone Valve IBC, which is specifically suited for powders
35 as it prevents typical powder flow problems such as bridging, blockages, segregation, flushing, and core-flow. Such containers are available from Matcon®.

As mentioned above, the organic peroxide composition according to the present invention finds use in the modification of polypropylene, in particular the production of high melt strength polypropylene (HMS-PP). In this modification, the composition is added to the polypropylene prior to or during extrusion. The composition may be added to the polypropylene after suspending it in water, dispersing it in an inert solvent such as isododecane, or in any solid physical form, e.g. flakes or powder. The quantity of organic peroxide to be used will depend on the desired degree of modification and on the type of polypropylene employed. Preferably, use is made of organic peroxide concentrations in the range of 0.1 to 3.0 g of peroxide per 100 g polypropylene, more preferably in the range of 0.5 to 2.0 g per 100 g polypropylene; all calculated as pure and dry organic peroxide.

EXAMPLES

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Example 1

A composition was prepared by tumble mixing a HDPE flask containing dicetyl peroxydicarbonate flakes (containing 97.5 wt% peroxide and 0.3 wt% cetylchloroformate) with different amounts (in wt% based on the organic peroxide weight) of different acid scavengers.

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Experiments with an additional amount of cetylchloroformate (0.9 wt%) were also performed.

The following acid scavengers were used: calcium stearate, hydrotalcite (DHT-4A, ex-Kisuma), calcium carbonate.

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Pre-cleaned smooth stainless steel (SS) 316L coupons (20x10x1 mm) were weighed and subsequently placed in a glass vial with approximately 1 g organic peroxide composition.

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The open glass vial was placed into a 250 ml HDPE flask containing approximately 20 g organic peroxide composition. The HDPE flask was closed with a screw-cap and was stored for at least 12 days at 20°C.

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After the storage period, the coupons were visually inspected for corrosion and, after

the below cleaning procedure, re-weighed to determine the weight loss. According to the cleaning procedure, any rust was removed from the corroded coupons with a wetted sponge (water/abrasive crème). The coupons were rinsed with distilled water and then with acetone, dried in an oven at 60°C for 1 hour, and cooled to room
5 temperature.

Table 1

Exp.	HCl scavenger	Stainless steel type	days at 20°C	Corrosion?	Weight loss (%)	Observations
No acid scavenger						
1	-	316L	12	++	0.108	clear corrosion, brown droplets and deterioration of coupon
2	-	254 SMO	12	+	0.027	Corrosion, but less than 316L
Ca-stearate						
3	0.05% wt% Ca-stearate	316L	14	--		no corrosion
4	0.25% wt% Ca-stearate	316L	60	--		no corrosion
DHT-4A						
5	0.01 wt% DHT-4A	316L	19	--		no corrosion
6	0.025 wt% DHT-4A	316L	42	--		no corrosion
7	0.2 wt% DHT-4A	316L	19	--		no corrosion
CaCO₃						
8	0.016 wt% CaCO ₃	316L	19	--		no corrosion
9	0.04 wt% CaCO ₃	316L	19	--		no corrosion
10	0.1 wt% CaCO ₃	316L	19	--		no corrosion
Dicetyl peroxydicarbonate flakes with additional 0.9 wt% cetylchloroformate						
11	-	316L	14	++	0.220	clear corrosion
12	2.5 wt% Ca stearate	316L	19	--		no corrosion
Influence of temperature on corrosion; no HCl scavenger						
13	10°C	316L	12	++	n.d.	clear corrosion
14	20°C	316L	12	++	n.d.	clear corrosion

The results show that the organic peroxide composition, without HCl scavenger, corrodes 316L stainless steel. The results also show that it corrodes a higher quality stainless steel 254 SMO; a stainless steel that is resistant to dilute HCl solutions.

5 Example 2

Several compositions prepared in Example 1 were tested for their corroding effects using the following test set-up:

Pre-cleaned smooth 316L stainless steel coupons (20x10x1 mm) were weighed and subsequently placed in a 250 ml HDPE flask: directly on top, or below a 21 g sample of
10 the dicetyl peroxydicarbonate-containing composition.

The HDPE flask was closed with a screw-cap and was stored for at least 11 days at 20°C.

After the test, the coupons were inspected, cleaned and weighed as explained in Example 1.

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Table 2

Exp.	HCl scavenger	days at 20°C	Corrosion?	Weight loss (%)	Observations
Coupon below the composition					
15	-	13	++	0.097	Clear corrosion
16	0.05 wt% Ca stearate	11	--		No corrosion
17	0.1 wt% Ca stearate	11	--		No corrosion
18	0.5 wt% Ca stearate	13	--		No corrosion
Coupon on top of the composition					
19	-	13	++	0.187	Clear corrosion
20	0.05 wt% Ca stearate	11	--		No corrosion
21	0.1 wt% Ca stearate	11	--		No corrosion

These results again confirm that the HCl scavenger is able to prevent corrosion of all inner walls of a stainless steel container.

CLAIMS

1. Organic peroxide composition that is solid at room temperature, said composition comprising:
- 5 - at least 40 wt% - based on the weight of the entire composition - of an organic peroxide that is solid at room temperature, said organic peroxide being selected from peroxydicarbonates and diacylperoxides,
- 0.001-5 wt% - based on the weight of organic peroxide in the composition - of an HCl scavenger that is solid at room temperature.
- 10 2. Organic peroxide composition according to claim 1 wherein the composition comprises 0.005-2.5 wt%, preferably 0.01-0.5 wt% - based on the weight of said organic peroxide in the composition - of said HCl scavenger.
- 15 3. Organic peroxide composition according to claim 1 wherein the composition comprises at least 85 wt%, preferably at least 90 wt% - based on the weight of the entire composition - of said organic peroxide.
4. Organic peroxide composition according to any one of the preceding claims wherein the organic peroxide is a peroxydicarbonate selected from the group consisting of dicetyl peroxydicarbonate, dimyristyl peroxydicarbonate, dicyclohexyl peroxydicarbonate, and di(tert-butylcyclohexyl) peroxydicarbonate.
- 20 5. Organic peroxide composition according to any one of claims 1-3 wherein the organic peroxide is a diacylperoxide selected from the group consisting of aliphatic diacylperoxides with aliphatic chains of at least 10 carbon atoms, dibenzoyl peroxide, and substituted dibenzoyl peroxides.
- 25 6. Organic peroxide composition according to any one of the preceding claims wherein the HCl scavenger is selected from the group consisting of metal carboxylates, metal oxides, metal (bi)carbonates, metal silicates, anionic clays, and combinations thereof.
- 30 7. Organic peroxide composition according to claim 6 wherein the HCl scavenger is selected from the group consisting of calcium stearate, calcium stearoyl-2-
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lactylate, CaO, calcium carbonate, and anionic clays, preferably selected from the group consisting of hydrotalcite, calcium stearate, and calcium carbonate.

- 5 8. Packaged organic peroxide composition comprising the organic peroxide composition according to any one of claims 1-7 in a stainless steel container.
9. Stainless steel container comprising the organic peroxide composition according to any one of claims 1-7.
- 10 10. Stainless steel container comprising:
- a solid composition comprising at least 40 wt% - based on the weight of said solid composition - of an organic peroxide that is solid at room temperature and is selected from peroxydicarbonates and diacylperoxides,
 - an HCl scavenger that is solid at room temperature, in an amount of at least 15 0.001 wt%, based on the weight of organic peroxide that is present in the stainless steel container.
11. Method for storing and/or transporting an organic peroxide composition that is solid at room temperature in a stainless steel container, said method involving 20 the addition to a stainless steel container of:
- (i) a solid composition comprising at least 40 wt% - based on the weight of said solid composition - of an organic peroxide that is solid at room temperature and is selected from peroxydicarbonates and diacylperoxides, and
 - (ii) at least 0.001 wt%, based on the weight of organic peroxide that is present 25 in the stainless steel container, of an HCl scavenger that is solid at room temperature.
12. Process for producing the organic peroxide composition according to any one of claims 1-7, comprising the step of physically mixing said organic peroxide and 30 said HCl scavenger.
13. Process for producing the organic peroxide composition according to any one of claims 1-7, wherein said HCl scavenger is added to the reactor in which the organic peroxide is produced.
- 35 14. Process for producing the organic peroxide composition according to any one of

claims 1-7, comprising the step of adding said HCl scavenger to molten organic peroxide.

- 5 15. Use of an organic peroxide composition according to any one of claims 1-7 for the modification of polypropylene, preferably a modification that results in increased melt strength of the polypropylene (HMS-PP).