

Description

Contact lenses are commonly placed into two categories, soft lenses and hard lenses. Soft lenses are rubbery and pliable, and they are fitted so as to closely match the shape of the cornea. Because of their softness they are more comfortable for the wearer than are hard lenses and because of their tight fit they have to be oxygen permeable in an amount sufficient to meet the needs of the cornea if worn for extended periods.

Hard lenses on the other hand have to possess excellent dimensional stability and are fitted so as to sit on the cornea only with the center of the lens, leaving enough space between their edges and the eye to allow good tear liquid exchange and thereby oxygen supply to the cornea. Although hard lenses are less comfortable to wear, they can be fabricated with greater precision and are preferred to correct, for instance, astigmatism. For both types of lenses excellent wettability by the tear liquid is important. This is especially true for soft lenses; soft lenses which cannot freely rotate on the eye and which are not easily wettable can become stuck to the cornea - which also is hydrophobic - and difficult to remove, sometimes only with injury to the eye, due to the suction-cup effect. For this reason commercially successful soft lenses are primarily of the hydrogel type, most commonly made from poly-2-hydroxyethyl methacrylate (HEMA), from HEMA copolymers, or from vinylpyrrolidone copolymers. These contact lenses generally contain from 38-75 % water and this water content is primarily responsible for their softness, their oxygen permeability (DK) of 8 to 30 barrers and their good physiological response.

On the negative side, the high water content makes a rather weak polymer, and oxygen permeability cannot therefore be further increased by raising water content or cutting thinner lenses. Oxygen permeability of 20-50 barrers for hard lenses and up to 100 barrers for soft lenses are desirable and required for extended wear contact lenses. In addition, hydrogel contact lenses have to be regularly sterilized.

These drawbacks could be overcome if silicone rubber could be used as contact lens material; it is strong, comfortably soft and polysiloxane is the polymer with one of the highest known oxygen permeability. Unfortunately polysiloxane is also very hydrophobic and imparts this hydrophobicity to polymers which contain a substantial amount of it. This hydrophobic character also leads to excessive deposition of protein and/or lipids on a contact-lens surface.

A number of methods have been used to make polymethylmethacrylate (PMMA) and polysiloxane based contact lenses more wettable. U.S. 4,131,696 describes the treatment of PMMA contact lenses with aqueous, colloidal suspensions of inorganic hydroxides. Plasma-treatment in a glow discharge chamber using water, hydrazine, ammonia, formamide, ethylene diamine and similar compounds is used to hydrophilize PMMA and silicone rubber and is described in U.S. 3,952,178. A similar process using gases such as oxygen, nitrogen, NO₂, N₂O, N₂O₃, NO, SO₂ and CO is described in FR 2,483,310. U.S. 4,139,660 describes the treatment of hydrophobic surfaces with halides of elements from group IV of the Periodic Table, followed by an aqueous rinse. U.S. 4,055,378 describes a treatment by activated oxygen ions of an SiO₂-filled silicone lens, by which the SiO₂ filler becomes exposed to the surface and/or organic silicone groups are oxidized to SiO₂ to thereby render the surface more wettable. U.S. 4,062,627 and U.S. 4,099,859 describe free-radical initiated grafting reactions of N-vinylpyrrolidone or hydroxyalkyl(meth)acrylates on the surfaces of silicone rubber contact lenses.

Every one of these methods suffers from one or several drawbacks. For example, plasma treatments and oxidizing treatment of PMMA and silicone surfaces produce only very thin modified surface regions and therefore these treatments do not provide sufficiently permanent wettability. If the effect wears off, it can only be restored by a new and expensive repetition of the original treatment. Grafting methods involve the use of a swelling monomer on a precision-cut contact lens, which can easily lead to permanent distortions of its optical properties. In addition, by the very nature of free-radical initiated grafting reactions quality control is very difficult and the whole process is cumbersome. The grafting polymerization must be conducted at temperature levels and periods of time under which the vinyl monomer is capable of diffusing into and swelling previously grafted polymer chains, thereby producing irregular regions of proliferating grafting by a positive feedback effect. In other words, after the very first grafting to the hydrophilic silicone surface has been accomplished the new surface is much more absorptive to the monomer. The monomer is thermodynamically compelled to swell the grafted hydrophilic polymer and further irradiation will produce additional grafting in regions already reached in the polymer. Thus, the process is one of positive feedback and can lead to gross differences in amount of grafted hydrophilic polymer per unit area to silicone substrate. Upon subsequent immersion in water the more grafted regions swell to the greater degree, thereby potentially distorting the interface and producing undesirable optical dispersion.

It has not unexpectedly been discovered that if glycidyl groups containing contact lenses are treated under mild aqueous conditions with a variety of watersoluble nucleophiles, such as an organic amine,

alcohol, thiol, urea, thiourea, sulfite, bisulfite or thiosulfate, a permanently wettable surface is obtained without impairing the clarity or the optical quality of the lens. It is a further advantage of the process of this invention that it is exceedingly simple and very easy to control and, should the effect every wear off, very safely and easily repeated by the wearer of the lens at home. It is a further objective of this invention to use the described surface-treatment of glycidyl groups containing contact lenses during the routine contact lens cleaning procedures carried out by the contact lens wearer. We speculate that the high quality and distortion-free nature of the final contact lens is a result of the single-step nature of the reaction. Step-reaction, of which polycondensation is a prime example, are much easier to control than are chain-reactions, where a very rapid growth reaction generally leads unpredictably to very high molecular weights. Single-step reactions also assure excellent compatibility of the treated surface with the underlying bulk polymer whereas the high molecular weight of free-radical grafted chains leads to phase separation and thereby optical haziness.

Glycidyl acrylate and glycidyl methacrylate as components of silicone-free hydrogel contact lenses have been specifically described in U.S. 3,787,380, 3,758,448, 3,772,235, 3,947,401; in several other patents it is disclosed as a possible component of the monomer mixture. In all instances the purpose of using this monomer is to influence water uptake and crosslinking reactions. U.S. 4,182,725, 4,182,723 and 4,169,119 describe silicone rubber lenses made wettable by inclusion of small amounts of glycidyl-(meth)-acrylate polymers, without additional surface treatment. Although an increase in wettability has been shown in these patents, this increase is insufficient to fulfill the requirements of a silicone-based soft contact lens having a surface of sufficient wettability to provide for adequate patient comfort.

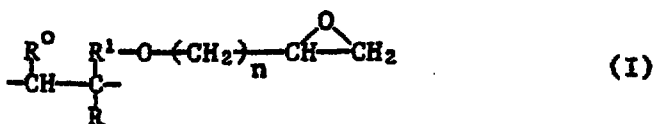
In Canadian Patent No. 1,099,546 glycidyl methacrylate is used as part of a hydrogel-contact lens formulation to introduce rigidity. U.S. 4,338,419, in relevant part, discloses the complete removal of residual uncrosslinked oxirane units present as a minor impurity in a crosslinked hydrophilic hydrogel material formed of a polymerized comonomer mixture initially containing not more than 5 % oxirane monomer as a crosslinking agent, and exemplifying less than 1 % by weight thereof in the initial monomer mixture, by reaction with a thiol. No increase in the hydrophilicity of the surface of the lens materials is reported.

It is an object of the present invention, in clear contradistinction to the prior art, to provide normally hydrophobic oxirane containing contact lenses with a substantially increased hydrophilic surface, i.e. to increase the wettability thereof, by incorporating a hydrophilicity inducing amount of reactive water soluble nucleophile selected from an organic amine, alcohol, thiol, urea or thiourea, or a sulfite, bisulfite, or thiosulfate - oxirane reaction product on the surface of such lens.

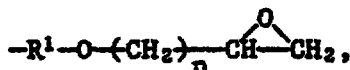
It is further object of the present invention to provide a method to correcting vision defects of a refractive nature by placing in the eye of a patient in need of the same, a contact lens made of such material containing the aforementioned surface modification.

These other object of the instant invention are more fully set forth in the following detailed description.

One embodiment of the present invention relates to a contact lens possessing a wettable surface in its ocular aqueous environment of use fabricated from a hydrophobic addition polymer incorporating between 0.5 and 30 % by weight oxirane substituted units in the backbone thereof, wherein the outer surfaces of said lens contain an effective hydrophilicity inducing amount of a reaction product of said oxirane substituted units of the formula I,



wherein R is hydrogen or methyl, R¹ is carbonyl or lower alkylene; n is 1 to 4; and R⁰ is hydrogen or



with a reactive water soluble nucleophile selected from an organic amine, alcohol, thiol, urea or thiourea, or sulfite, bisulfite or thiosulfate.

By hydrophobic addition polymers is meant that class of addition polymers which exhibits a contact angle by the captive bubble technique of less than 150°, preferably less than 145°, and most preferably

less than 140°.

Alternatively, the hydrophobicity of such addition polymers can be characterized in terms of the wettability of the polymer surface, generally expressed as the rate at which a water film on the surface of the polymer retracts or beads up. Thus, in the context of the present invention, useful oxirane containing hydrophobic addition polymers are those wherein a sample, having a vertical length of 40 mm, a width of 20 mm and a thickness of 1 mm is vertically removed from distilled water and the water film retracts, or beads up, at a rate such that about 90 % of the surface is visually water free in less than about three seconds, preferably in about two seconds or less.

By a hydrophilicity inducing amount of the aforementioned reaction product of a reactive water soluble nucleophile with the oxirane containing units of formula I is meant an amount sufficient to increase the contact angle, as measured by the captive bubble technique by at least 5°, more preferably by at least 10°, most preferably by at least 15°.

In terms of wettability, as defined in accordance with the retraction time of a 40-mm by 20-mm by 1-mm sample, a hydrophilicity inducing amount can be characterized as an amount sufficient to increase the 90 % retraction time of water to greater than about 5 seconds, preferably greater than about 10 seconds.

Hydrophobic addition polymers containing units of formula I are known, per se, or can be manufactured by polymerization techniques well known in the art, by copolymerizing between 0.5 and 30 percent by weight of an oxirane containing monomer of the formula II,



with 99.5 to 70 percent by weight of one or more conventional mono- or polyvinyllic comonomers employed in the contact lens field.

Suitable comonomers include alkyl acrylates and methacrylates, especially C₁-C₂₀ alkyl acrylates and C₁-C₂₀ alkyl methacrylates, such as methyl methacrylate, ethyl methacrylate, methyl acrylate, butyl methacrylate, butyl acrylate and 2-ethylhexyl acrylate; alkanolic vinyl esters, especially C₁-C₆ alkanolic vinyl esters such as vinyl acetate and vinyl butyrate; alkenes, especially C₂-C₈ alkenes, including ethylene, 1-butene and 1-hexene; styrenes, especially styrene and alpha-methyl styrene; vinyl ethers, especially C₁-C₆ alkyl vinyl ethers, including methyl vinyl ether, ethyl vinyl ether and butyl vinyl ether; dialkyl maleates, fumarates or itaconates, especially C₁-C₆ dialkyl maleates, fumarates or itaconates, including dimethyl maleate, dimethyl fumarate, diethyl maleate and dimethyl itaconate; allyl esters and esters, especially allyl C₁-C₆ alkyl ethers and allyl C₂-C₆ alkanolate esters, including allyl methyl ether, allyl ethyl ether and allyl acetate; perfluoro C₃-C₆ alkyl acrylates or methacrylates; perfluoroalkoxylated bis-acrylates or methacrylates; poly- or oligo-alkylsiloxane acrylates or meth-acrylates.

Also, minor amounts of crosslinking agent, to increase hardness, stability and the mechanical characteristics of the polymer are generally employed. Suitable crosslinking agents include, for example, C₂-C₆ alkylene ether di-methacrylates and acrylates, e.g. ethylene glycol dimethacrylate, diethylene glycol dimethacrylate, triethylene glycol dimethacrylate, glycerine trimethacrylate; allyl acrylate or methacrylate, divinyl benzene, poly- or oligo-alkyl-siloxane di-acrylate or -methacrylate.

Also, as a minor constituent, there may be present one or more conventional hydrophilic monomers. The amount of such hydrophilic monomers will, for the purpose of this invention, be less than the amount needed to raise the aforementioned contact angle of the polymer to greater than about 150°, preferably less than 145°, most preferably less than about 140°, and in an amount insufficient to retard the 90 % retraction time to greater than about three seconds, preferably insufficient to retard the 90 % retraction time to greater than about two seconds or less. Suitable hydrophilic monomers are those monomers conventionally employed in the contact lens field and include hydroxylated C₂-C₆ alkyl acrylates and methacrylates, including 2-hydroxyethyl methacrylate, 2-hydroxyethyl acrylate, 2,3-dihydroxypropyl acrylate, and 2,3-dihydroxypropyl methacrylate; N-vinyl lactams, including N-vinyl pyrrolidone; acrylamide; C₁-C₂ alkyl acrylamides, such as methyl acrylamide or dimethyl acrylamide.

Preferably, the hydrophobic addition polymer is an addition polymer containing from 1 to 20 weight percent of the units of formula I, i.e. formulated from a corresponding amount of the monomer of formula II. Most preferably, the polymer contains from 5 to 12 weight percent of the units of formula I.

Preferred comonomers are C₁-C₆ acrylates and methacrylates, including methyl methacrylate, butyl acrylate and 2-ethylhexyl acrylate; ethylene glycol dimethacrylate; and poly- or oligo-alkyl-siloxane mono-

or bi-acrylates or methacrylates; or mixtures thereof.

Preferred monomers of formula II include glycidyl acrylate, glycidyl methacrylate and allyl glycidyl ether. Most preferred is glycidyl methacrylate.

The polymer can be obtained from the monomer mixture by methods known per se and generally under conditions which do not affect the oxirane ring. Thus, the monomer mixture can be polymerized at a temperature between 0°C and 100°C, in the optional presence of a free radical catalyst and/or in the presence of actinic radiation, such as UV radiation.

The polymer may be cast in the form of a contact lens by the use of a conventional mold, or by spin casting polymerization techniques or may be cast as a rod, or as buttons which are then machined into the desired shape of a contact lens.

For example, contact lenses suitable for carrying out surface reactions can be made by either molding or casting, or by machining from clear polymers containing 0.5 to 30 weight percent of units of formula I in the polymer backbone. They can be hard contact lens materials as exemplified by polymethylmethacrylate (PMMA), such as PMMA copolymers with glycidyl methacrylate; or they can be silicone containing hard lenses, which contain in addition to PMMA or other hard (meth)-acrylic polymers oxygen-permeable poly- or oligosiloxane units, and which are for instance described in U.S. 3,808,178; 4,120,570; 4,139,513; 4,152,508; 4,189,546; and 4,157,508. Since glycidyl methacrylate is a monomer which imparts high hardness to an acrylic copolymer, its use as an additional comonomer in the compositions covered in these patents does not negatively influence their properties.

Another class of contact lenses especially useful for this invention are silicone containing lenses, such as hydrophobic, silicone containing lenses, as for instance disclosed in U.S. 4,182,725 or 4,245,069.

Inclusion of 0.5 to 30 weight percent of a compound of formula II, e.g. glycidyl methacrylate into the above disclosed polymer compositions gives lenses which can be modified in the manner described hereinafter.

Most preferred silicone contact lens materials are:

Hydrophobic, silicone containing copolymers of 1-30 % glycidyl (meth)acrylate, 50-90 % of the reaction product of a diisocyanate capped polydimethylsiloxanylalkyl-polyol of a MW from 800-5000 with 2-hydroxyethyl (meth)acrylate, or N-tertbutylaminoethyl meth-acrylate, or of said polydimethylsiloxane polyol with 2-isocyanato-ethyl methacrylate, and 0-40 % of a C₁-C₂₀ alkyl (meth)acrylate.

The reactive nucleophile may be a water soluble aliphatic, cycloaliphatic or aromatic primary, secondary or tertiary amine or polyamine, especially those having up to 12 carbon atoms, and preferably amines which contain an additional functional group as part of the molecule, such as a conventional hydrophilic moiety, or several such moieties. Primary and secondary amines are the preferred compounds.

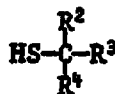
Conventional hydrophilic moieties which may be present as substituents include, for example, hydroxy, amido, polyethoxy, lower alkoxy, carboxy, sulfato, sulfonamido, sulfo, phosphato and phosphonato. Preferred polyethoxy groups include polyethyleneoxy of 2 to 100 units.

Other reactive nucleophiles are water soluble sulfite, bisulfite and thiosulfate salts, water soluble mono- or dithiols, either aliphatic or aromatic in nature, and preferably water soluble thiols of up to 12 carbon atoms optionally containing one or more hydrophilic moieties, e.g. as enumerated in the preceding paragraph:

Exemplary of suitable water soluble compounds are:

methylamine, dimethylamine, ethylamine, isopropylamine, diisopropylamine, poly(ethyleneimine), 2-aminopropylether, polyethylene or polypropylene oxide diamine (Jeffamine) and those in which the amine is contained in a ring such as pyrrolidone, piperidine, piperazine, morpholine, thiomorpholine, ethylenediamine, bis-aminopropylpiperazine, aniline, p-methoxyethoxybenzylamine, aminoethanol, diethanolamine, n-propanolamine, di-n-propanolamine, isopropanolamine, diisopropanolamine and other lower branched alkanols which can be polyhydroxy such as 2-amino-2-methyl-1-propanol, tris(hydroxymethyl) aminomethane, 1,3 bis [tris(hydroxymethyl)methylamino] propane, N-(3-aminopropyl)diethanol, D-glucosamine, N-methyl-D-glucamine, N-(β-hydroxyethyl)piperazine, 2-piperidineethanol, p-aminophenol, aminocresol, aminobenzylalcohol, glycine, serine, aspartic acid, sarcosine, tyrosine, tricine, 3-aminobutyric acid, 6-aminoheptanoic acid, 4-amino-3-hydroxybutyric acid, 2-aminoethanesulfonic acid, N-methyltaurine, 2-aminoethyl hydrogen sulfate, N-(2-acetamido)-2-aminoethanesulfonic acid, aminobenzoic acid, p-aminocinnamic acid, 4-amino-3-hydroxy-1-naphthalene sulfonic acid, aminosalicylic acid, 6-amino-m-toluenesulfonic acid, 3-amino-L-tyrosine dihydrochloride, purine, pyrimidine, adenine, 2-amino-purine, cytosine, thymine, uracil, adenosine, cytidine, guanosine, deoxyadenosine, thymidine, adenylic acid, adenosine triphosphate, thymidylic acid, deoxycytidine diphosphate, thymidinediphosphate, aminopyridine, imidazole, 2-aminoimidazole, oxazole, thiazole, (4-aminobutyl)guanidine, S-(2-aminoethyl)isothiuronium hydrobromide, 3-aminocrotonamide and phosphatidylethanolamine.

Preferred water soluble thiols include compounds of the formula



wherein R^2 is hydrogen or R^5 and R^3 is hydrogen or R^6 where R^5 , R^6 and R^4 are independently straight or branched chain alkyl, cycloalkyl, aryl, aralkyl, alkaryl of up to 12 carbon atoms which may be unsaturated or interrupted by heteroatoms such as oxygen or nitrogen, or R^2 and R^3 together is an aliphatic divalent radical of up to 6 carbon atoms, and R^5 , R^6 and R^4 may be unsubstituted or substituted with one or more of the aforementioned hydrophilic moieties.

Useful preferred compounds include water soluble amines of the general structure:



wherein R^7 is hydrogen or R^9 where R^9 and R^8 may be independently a straight or branched chain alkyl, cycloalkyl, aryl, aralkyl or alkaryl of up to 12 carbon atoms that may include unsaturation or be interrupted by heteroatoms such as oxygen and nitrogen, or R^7 and R^8 together form a divalent aliphatic radical of up to 6 carbon atoms which is unsubstituted or substituted by one or more hydrophilic groups. R^9 and R^8 can also together represent a heterocyclic moiety. Suitable hydrophilic moieties are those recited above.

Exemplary of suitable water soluble thiols are: ethanethiol, 1,4-butanedithiol, 2-mercaptoethylether, pentaerythritol-tetrathio-propionate and acetate; polyethyleneglycoldimercaptoacetate and methylthioglycolate, allyl mercaptan, 2-mercaptoethanol, 3-mercaptoopropanol, 4-mercaptobutanol, 1-thioglycerol, thioerythritol, 2,3-dimercaptoopropanol, pentaerythritolmono (or di or tri)thio-propionate or acetate, thioglycolic acid, thioacetic acid, 3-mercaptopropionic acid, thiolactic acid, thiomalic acid, thiosuccinic acid, thiosalicylic acid, thiobenzic acid and their respective water soluble salts, glutathione, furfuryl mercaptan, 2-mercaptobenzimidazole, 2-mercaptobenzoxazole, 2-mercapto-3-pyridinol, dimethylaminopropanethiol, 2-mercaptoethylamine and 2-n-butylaminoethanethiol.

Also, especially preferred are the water soluble salts of sulfite, bisulfite, and thiosulfate, especially the pharmaceutically acceptable salts thereof, such as alkaline or alkaline earth metal salts, especially the sodium salts.

Exemplary of suitable water soluble ureas and thioureas are: urea, methylurea, ethylurea, butylurea, hydantoin, 5-hydantoinacetic acid, 2-amino-3-ureido-propionic acid, diureidoacetic acid, pyrimidine-tetrone, 6-amino-1, 3-dimethyluracil, 5-aminoorotic acid, 2-amino-5-ureidovaleric acid, 2-imidazolidinone; thiourea, 1-methyl-2-thio-urea, 5-methyl-2-thiouracil and 2-imidazolidinethione.

Exemplary of suitable water soluble alcohols are: methanol, ethanol, glycol, glycerol, polyvinyl alcohol, polyethylene glycol, dextrose, glucose, mannose, fructose, sorbitol, maltose, dextran, dextrin, agarose, gum arabic, phenol, acetamidophenol, pyrogallol and resorcinol.

The reaction can be carried out in any liquid medium which is a solvent for the reactant but is preferably an aqueous medium, consisting of water or water-solvent mixtures. The addition of solvent to the reaction medium is sometimes necessary to ensure penetration of the reactant into the substrate, especially if this substrate is hard and hydrophobic. Solvents may also be required to solubilize otherwise insoluble reactants. It is in some applications such as contact lenses desirable to cause as little swelling of the substrate as possible during the reaction in order to avoid later optical distortion.

In many cases it is also advisable to use a buffered aqueous solution to insure reproducibility and avoid unfavorably high or low pH range. Preferably the reaction is carried out at a pH between 3 and 10, more preferably between 7 and 9. Since at the very low pH necessary to effect reaction between glycidyl groups and tertiary amines, hydrolysis of the glycidyl group occurs also, tertiary amines are generally unsuited to prepare the products of this invention. Any of the commonly used buffers can be used as long as they do not interfere with the reaction, for instance phosphates, borate, citrates. Some reactants, such as TRIS (trihydroxymethyl-aminomethane) are themselves buffers.

However, to react alcohols, a higher pH is needed. Thus, simple heating in e.g. ethanol normally is not effective. A reaction with alcohols according to the invention is therefore carried out in the presence of strong bases, e.g. alkali or alkaline earth metal hydroxides, such as sodium hydroxide. If for example amine

and ethanol are present and the pH is not high enough, only amine reacts. If pH is high enough, amine and alcohol mixture results in some product involving each. However, there is a preferential reacting of the species with the greatest nucleophilic character. The key to reacting an alcohol is deprotonating the alcohol.

The reaction temperature can vary over a wide range, typically it is between 0 and 100 °C, preferably between 20 and 60 °C, most preferably between 25 and 35 °C. The reaction time can also vary as a matter of choice of reaction temperature and reactant concentration. Dependent on substrate and reactant, reaction times from 1 minute to 1 day are useful and most commonly they range from 10 minutes to 5 hours.

Reactant concentrations can vary from 0.1 % to 100 %, preferably they range from 5 % to 50 %, most preferably they are between 10 and 30 %.

When the reaction is carried out in primarily aqueous or in organic medium with hydrophobic contact lenses, especially silicone containing materials, the reaction and resulting modifications can be restricted to a surface region whose depth depends on a variety of parameters such as molecular size of reactant, nature and amount of cosolvent, reaction time and temperature. By determining the rate of diffusion and of reaction of a given reactant the depth of the modified region can accurately be predicted. Generally, a depth of up to about 10 percent of the thickness of the lens can be reacted. Excessive penetration may result in reduced oxygen permeability and should generally be avoided. Also, excessive penetration of the reactant may in some instances, unduly distort the lens parameters, e.g. shape. Preferably a depth of up to about 5 percent of the thickness if reacted.

In the following examples, wettability is determined by contact-angle measurements on the underside of samples immersed in distilled water, floating either an air bubble or a drop of n-octane to the lower surface. In these measurements a high contact angle corresponds to high wettability. In some cases wettability is denoted in terms of retraction time. We define the retraction time as follows: If a sample (40 x 20 x 1 mm) is vertically removed from water, the water film of the surface retracts (or beads up). The number of seconds elapsing until 90 % of the surface is free of water is defined as the retraction time (t_R). Low wettability would correspond to a t_R of less than three seconds, preferably two seconds or less.

In the following examples, temperatures are given in degrees Centigrade and compositions in percents by weight if not defined explicitly otherwise.

The polymers listed in Table A hereinafter are prepared by UV-initiated free-radical polymerization of vinylic monomers with a suitable dimethacrylate as crosslinking agent. The short-chain crosslinking agent is ethylene glycol dimethacrylate (EGDA). Polymeric crosslinking agents are either the trimethylacrylate (PSM) ~6000 MW (fluid "1248" from Dow Corning) with 3 moles isophorone diisocyanate followed by capping with 2-hydroxyethyl methacrylate (= PSM) or the dimethacrylate (PEM) obtained by reacting 1 mol of poly-n-butyleneoxide diol of ~2000 MW (Polymeg 2000, Wyandotte Chem. Corp.) with 2 moles isophorone diisocyanate followed by capping with 2-hydroxyethyl methacrylate (= PEM).

Example A: 80 g PSM, 30 g methyl methacrylate (MMA) and 10 g glycidyl methacrylate (GMA) are mixed in a bottle together with 0.2 % 1-hydroxycyclohexyl phenyl ketone (IRGACURE-184), a UV-initiator. The sample is degassed under nitrogen and filled into MYLAR lined glass molds (10 cm x 10 cm) using 1 mm rubber cord or cut-out 0.1 mm MYLAR sheet spacers. The molds are irradiated by Black-Light Blue Sylvania lamps for 3 hours, after which time the sample is removed and extracted in absolute ethanol for two days. After drying the polymer can be used for surface modification.

The following table A lists the compositions synthesized in the manner described above.

Table A

Polymer Number	Composition % (w/w)								
	PEM*	PSM*	HEMA	DMA	MMA	EHA	BA	GMA	EGDA
1	-	50	10	-	20	-	-	20	-
2	-	50	10	-	-	20	-	20	-
3	-	60	-	-	-	20	-	20	-
4	19	-	-	60	-	-	-	20	1
5	15	-	-	60	25	-	-	-	-
6	-	-	-	-	39.5	39.5	-	20	1
7	-	-	-	-	40	40	-	20	-
8	-	-	-	-	39.5	-	39.5	20	1
9	-	-	-	-	40	-	40	20	-
10	-	-	-	-	79	-	-	20	1
11	-	-	-	-	-	-	79	20	1
12	50	-	-	-	30	-	-	20	-
13	-	50	-	-	30	-	-	20	-

BA : butyl acrylate

DMA : dimethyl acrylamide

EGDA: ethylene glycol dimethacrylate

EHA : 2-ethylhexyl acrylate

GMA : glycidyl methacrylate

HEMA: 2-hydroxyethyl methacrylate

MMA : methyl methacrylate

* : see explanation before example A

If hereinafter reference is made to a polymer 1 or 2 or 3 etc., the polymers having the compositions given in said table A are intended.

The following examples show the effect of solvent on the yield of reaction and the results are summarized in Table B.

Example 1: Three sections (19x7x1 mm) of polymer 1 of combined weight of 0.3787 g (5.33×10^{-4} moles GMA) are stirred for 19 hours in a solution of 1.72 g (0.016 moles; this is 30 molar excess of amine to epoxy groups) diethanolamine in 4.4 g absolute ethanol (final volume of 6 ml). Then the samples are washed by stirring in absolute ethanol (25 ml, 3x1 hour). The samples are then dried in a vacuum oven at 70°C for 2 1/2 hours. From the weight of modified polymer (0.4108 g), the yield based on GMA is calculated to be 57 %.

Example 2: This example is performed in a similar fashion to that of example 1 except the solvent used is 70 % ethanol/30 % water (w/w). The final washing step (total of 4 washings) is followed by an additional wash in absolute ethanol before vacuum drying. The calculated yield based on GMA content is 24 %.

Example 3: This example is performed in a similar fashion to that of example 2 except the solvent used is 60 % ethanol/40 % aqueous buffer (pH 9). Since the concentration of buffer is 1 % (potassium phosphate and sodium borate) and that of reactant approximately 3.4 Molar, the solvent behaves as 60 % ethanol/40

% water. The calculated yield based on GMA content is 10 %.

Example 4: This example is performed in a similar fashion to that of example 2 except the solvent used is deionized water. The washing step is 2 x 1 hour in water, 2 x 1 hour in 60 % ethanol/40 % water and then absolute ethanol. The calculated yield based on GMA content is 0 %.

5 The results are summarized in Table B and show that the yield of the reaction decreases as the percentage of water in the reaction solvent increases. Since the samples swell to a greater extent in ethanol than in water (length of sample in ethanol, water and dry state is 2.1, 2.0 and 1.9 cm respectively) more epoxy groups would be exposed in ethanol and available for reaction with the amine. Also we note that as the reaction proceeds the modified samples swell to an even greater extent. For example, a dry sample 1.9
10 cm in length increases to 2.1 cm after swelling in absolute ethanol; then after a reaction with diethanol amine in ethanol the length increases further to 2.5 cm. This can account for the yield dropping from ~60 % when the reaction takes place in absolute ethanol to a negligible value when the reaction takes place in deionized water (or buffer, pH 9). Note that by stating the yield is zero means that the weight gained after the drying step is not detectable. A 0.1 % yield, which would not be detected by our weighing procedure,
15 can result in a depth of several thousand Angstroms having been modified. We also note from Table B that the introduction of the alcohol groups from diethanolamine increases the hydrophilicity of the material.

Table B
Reaction of Polymer 1 with Excess Diethanolamine in Ethanol
and Aqueous Ethanol

Example	Solvent	Yield (%)	H ₂ O-Content (%)
1	Absolute Ethanol ^{a)}	57	11.2
2	70 % Ethanol 30 % Water ^{b)}	24	8.1
3	60 % Ethanol 40 % Buffer (pH 9) ^{c)}	10	6.1
4	Deionized Water ^{c)}	0	3.2

a) 19 hours at room temperature in a volume of 6 ml.

b) 16 1/2 hours at room temperature in a volume of 6 ml.

c) 16 1/2 hours at room temperature in a volume of 5 ml.

Examples 5-8: The following examples show the dramatically improved wettability of polymer 2 after treatment with amine and sulfite reactants in 15 % ethanol/85 % water. The reaction procedure used is essentially that mentioned earlier except the last washing step is not with ethanol but with deionized water. The contact angles are measured after approximately 7 days equilibration in water.

The monomer EHA is included in the polymer 2 composition to give the final product a soft, draping quality. In addition to rendering the final polymerized samples soft, the addition of EHA also leads to a material that is more hydrophobic. Thus in reaction solvents that contain a high proportion of water the polymerized sample swells to an even lower degree resulting in an even more surface specific reaction (i.e. less depth).

Example	Reactant	Contact Angle	
		Air/Water	Octane/Water
5	TRIS ^{a)}	168	167
6	D-glucosamine ^{b)} (pH 8)	167	168
7	Na ₂ SO ₃ ^{c)}	156	160
8	- (Control) ^{d)}	146	142

a) 1.53 g TRIS (= Tris (hydroxymethyl)aminomethane) in 10 g solvent.

b) 2.73 g D-Glucosamine·HCl in 10 g solvent, half-neutralized.

c) 1.6 g Na₂SO₃ in 18 g solvent.

d) Polymer 2 in 10 g solvent without reactant.

The results in the above table indicate, in most cases, more than a 20° increase in contact angle. This wettability change can be observed more dramatically by eye. Water on the control sample beads up within seconds but clings for several minutes on the reacted samples.

Example 9-12: The following examples show an improved wettability of polymer 1 after treatment with amine and sulfite reactants in 15 % ethanol/85 % water. The reaction procedure is similar to that for example 5-8. In this case the replacement of EHA with MMA leads to a polymer that is a harder and stiffer material than polymer 2.

Example	Reactant	Contact Angle	
		Air/Water	Octane/Water
9	TRIS	152	146
10	D-Glucosamine (pH 8)	148	139
11	Na ₂ SO ₃	-	147
12	- (Control)	-	138

For this polymer the contact angle increase is significant but much less than that for the more flexible polymer 2.

Example 13 and 14: The following example with polymer 3 shows the high wettability of the siloxane macromer compositions after a posttreatment. Reaction is in 15 % ethanol/85 % water.

Example	Reactant	Contact Angle	
		Air/Water	Octane/Water
13	TRIS	161	156
14	- (Control)	144	141

Example 15: The following example shows the wettability change as a function of reaction time for polymer 3 reacted with TRIS in 15 % ethanol/85 % water. After the designated time, the sample is subjected to the washing procedure as stated in example 5.

Reaction Time	Contact Angle	
	Air/Water	Octane/Water
No Tris, Control	146	142
10 minutes	155	143
1.5 hours	159	159
17 hours	168	167

The results show that although at 10 minutes a 9° air/water contact angle change occurs, only a 1° change occurs in the octane/water angle. However a significant change is observed in both angles after 1.5 hours although still not of the same magnitude as that of an overnight reaction.

Examples 16-32: The following examples demonstrate the universality of the posttreatment reaction on non-silicone copolymers containing 20 % GMA in 15 % ethanol/85 % water. In these materials the rigidity is increased with MMA and the flexibility increased with EHA. Also shown is that the presence of crosslinking agent does not disturb the effect of the reaction. The reaction procedure is similar to that of examples 5-8.

Example	Polymer Number	Reactant	Contact Angle	
			Air/Water	Octane/Water
16	6	TRIS	143	130
17	6	Na ₂ SO ₃	136	132
18	6	Control	121	102
19	7	TRIS	119	83
20	7	Na ₂ SO ₃	140	136
21	7	Control	119	97
22	8	TRIS	128	118
23	8	Na ₂ SO ₃	137	136
24	8	Control	113	97
25	9	TRIS	122	105
26	9	Na ₂ SO ₃	115	106
27	9	Control	114	93
28	10	TRIS	120	111
29	10	Na ₂ SO ₃	134	123
30	10	Control	111	95
31	11	Na ₂ SO ₃	142	139
32	11	Control	109	89

If one averages the contact angle increases on different copolymers one finds, for sodium sulfite, a 17° air/water and 30° octane/water angle increase and for TRIS a 9° air/water and 7° octane/water angle increase over that of the unreacted control.

Example 33-40: The following examples show that the effectiveness of the posttreatment on GMA containing SIMAC compositions (those are e.g. the compositions PEM and PSM as defined before example A) extends to levels of GMA much less than 20 %. The polymers 3a-3g are synthesized according to the procedure given in example A by replacing GMA with MMA in polymer 3. The reaction procedure is similar to that in examples 5-8 using TRIS in 15 % ethanol/85 % water.

Ex.	Poly- mer	MMA (%)	GMA (%)	Contact Angle			
				Treated		Untreated Control	
				Air/H ₂ O	Octane/H ₂ O	Air/H ₂ O	Octane/H ₂ O
33	3	-	20	158	154	147	142
34	3a	5	15	158	154	143	131
35	3b	10	10	157	154	147	139
36	3c	15	5	154	152	144	136
37	3d	17.5	2.5	160	152	-	-
38	3e	19	1	153	151	-	-
39	3f	19.5	0.5	151	147	-	-
40	3g	20	-	147	141	144	141

Examples 41-43: The following examples show that the effectiveness of the posttreatment on GMA containing non-silicone copolymers covers a wide range of GMA content. Polymers 6a and 6b are synthesized according to the procedure given in example A by replacing GMA with MMA. Also synthesized is a homopolymer of GMA, crosslinked with 1 % EGDA (Polymer 6c). In these cases a lowering of GMA content is balanced by an increase in MMA content. The reaction procedure is similar to that noted in examples 5-8 this time using TRIS in 15 % ethanol/85 % water.

Ex.	Pol.	MMA (%)	GMA (%)	Contact Angle			
				Treated		Untreated	
				Air/H ₂ O	Octane/H ₂ O	Air/H ₂ O	Octane/H ₂ O
41	6a	54.5	5	151	134	134	117
42	6b	49.5	10	160	143	135	127
43	6c	-	99+ 1 % EGDA	143	131	127	112

The results show an approximately 17° octane/H₂O contact angle increase after reaction with TRIS over that of the unreacted control.

Example 44-47: The following examples show the effectiveness of the posttreatment in 15 % ethanol, on polymer 3 with a variety of primary amines. The reaction procedure is similar to that noted in examples 5-8.

Example	Reactant	$t_R^a)$	Contact Angle	
			Air/H ₂ O	Octane/H ₂ O
44	H ₂ NCH ₂ CH ₂ OH	64	157	152
45	H ₂ NC(CH ₃) ₂ CH ₂ OH	>120	154	151
46	H ₂ NC(CH ₂ OH) ₃	>120	155	152
47	-(Control)	1	141	134

a) t_R = retraction time in seconds evaluated 4 days after reaction with amine.

The results show a dramatic improvement in wettability using primary amines containing the hydrophilic hydroxyl function.

Examples 48-55: The following examples show the effectiveness of the posttreatment in 15 % ethanol on polymer 3 with a variety of secondary amines. The reaction procedure is similar to that noted in examples 5-8.

Example	Reactant	$t_R^a)$	Contact Angle	
			Air/H ₂ O	Octane/H ₂ O
48	HN[CH(CH ₃) ₂] ₂	5	154	150
49	NH[CH ₂ CH(OH)CH ₃] ₂	4	155	152
50	(HOCH ₂) ₃ CNH(CH ₂) ₃ NHC(CH ₂ OH) ₃	3	154	150
51	HN(CH ₃)CH ₂ [CH(OH)] ₄ CH ₂ OH	30	155	144
52	HN(CH ₃)CH ₂ CH ₂ SO ₃ Na	~100	148	146
53	morpholine	10	157	153
54	HN(CH ₃)CH ₂ COOH	~100	153	150
55	(HOCH ₂) ₃ CNHCH ₂ COOH	3	150	147
47	-(Control)	1	141	134

a) t_R = retraction time in seconds.

The results show that even a hydrophobic amine (Ex. 48) can lead to a wettability increase presumably because the hydroxyl group produced on opening the epoxy ring counteracts the hydrophobic methyl groups. Examples 52 and 54 point out that ionic groups strongly affect the retraction time.

Examples 56-59: The following examples show the effectiveness of the posttreatment in 15 % ethanol on polymer 3 with a variety of mercaptans. The reaction procedure is similar to that noted in example 5-8 except that the solutions are 10 % in mercaptan and 1 M in sodium hydroxide. The final pH of the resultant solution is indicated in the Table below:

Example	Reactant	pH	t_R ^{a)}	Contact Angle	
				Air/H ₂ O	Octane/H ₂ O
56	HSCH ₂ CH ₂ OH	10	28	158	154
57	HSCH ₂ CH(OH)CH ₂ OH	10	58	156	153
58	HSCH ₂ COOH	5.5	31	156	152
59	<u>o</u> -HSC ₆ H ₄ COOH	7.6	>100	157	153
47	-(Control)	---	1	141	134

a) t_R = retraction time in seconds.

The results are similar to the secondary amine examples in that ionic groups (Ex. 59) lead to especially long retraction times.

Examples 60-63: The following examples show the effectiveness of the posttreatment in 15 % ethanol on polymer 3 with thiourea, a hydrazide and thiosulfate. The reaction procedure is similar to that noted in examples 5-8 except that the solutions are 1.5 g of reactant in 15 g solvent.

Example	Reactant	Contact Angle	
		Air/H ₂ O	Octane/H ₂ O
60	$\text{H}_2\text{N}-\overset{\text{S}}{\underset{\text{ }}{\text{C}}}-\text{NH}_2$	147	143
61	$\text{H}_2\text{N}-\text{NH}-\text{CO}-\overset{\text{+}}{\text{N}}(\text{CH}_3)_3\text{Cl}^-$	149	144
62	Na ₂ S ₂ O ₃	151	144
63	-(Control)	142	139

The results show an improvement in wettability over that of the untreated control.

Other Evidence of Reactions on Epoxy Containing Polymers

We use two methods to more fully characterize the reaction of GMA containing polymers with various reactants in 15 % ethanol i.e. (A) infrared absorption changes and (B) tinting.

(A) In this method we find it necessary to use thin (0.1-mm) films of a polymer that swell in this solvent and then monitor the decrease of the infrared absorption bands, corresponding to the epoxy group, signaling its reaction. The bands monitored using polymer 5 are 850 and 910 cm⁻¹.

(B) A novel way of providing evidence for the posttreatment reaction is by dyeing thick (1-mm) films with fluoresceinamine. The amino group of this dye reacts with epoxy groups and therefore films still containing epoxy groups are tintable. The posttreated films are placed in a 0.04 % solution of fluoresceinamine, pH ~12. The films are then extracted overnight in methanol and then further extracted by soxhlet extraction (25 hours in absolute ethanol).

Examples 64-68: The following examples show reactions on polymer 4 using a hydrophilic primary amine and secondary amine. Where a reaction occurs (by disappearance of epoxy bands) tinting fails and where a reaction fails (by non-disappearance of epoxy bands) tinting occurs. Included is a control polymer (Polymer 5) that does not contain epoxy groups (Ex. 68).

Example	Reactant	I.R.	Intensity of Fluoresceinamine Emission ^{a)}
64	H ₂ NCH ₂ CH ₂ OH	+	1
65	N-methyl-D-glucamine	+	1
66	Tricine	-	6
67	no reactant	-	6
68	Control, no epoxy		1

+ = reacted, - = unreacted

1 = colorless, 8 = deep colored

a) Intensity determined while sample illuminated with lamp of 360-nm emission maximum.

The results show that simple primary and secondary amines react but a hindered amine such as tricine fails to react. The control example (No. 68) shows that the amino dye is not absorbed strongly by non-covalent action and that non-epoxy sites in the polymer are not reactive to the amine.

Tintability of reactive dyes of posttreated polymers

Example 69: A sample of polymer 3 and of polymer 3 treated with 1-thioglycerol according to the procedure described in example 57 (~ 1-mm thick) are soaked for one half hour in 4 ml of 10 % Na₂CO₃ pH 11.6 and then 0.1 ml of dyestuff stock solution (0.1 g Royal Blue in 10 ml of 1 mM hydrochloric acid) is added. The samples are kept overnight in the dyestuff solution, rinsed in buffer pH 7.4 (50 mM potassium phosphate) and extracted for 24 h in 20 ml of methanol. The blue color of the thioglycerol treated sample does not change on subjecting the sample to soxhlet extraction (66 hours, distilled water), whereas the control becomes completely colorless.

Example 70: Polymer 4 which has been treated with diisopropanolamine in a procedure described in example 49 is tinted with Royal Blue. The tinting procedure is the same as that described in Example 69. The resultant blue color is of deeper intensity than that described in example 69.

Example 71: Polymer 4 which has been treated with TRIS in aqueous solution in a procedure described in example 5 is tinted with Brilliant Yellow G. The dyestuff stock solution contains 0.1 g of dye in 10 ml of 1 mM hydrochloric acid. The tinting procedure (using a 0.1-mm film) is similar to that described in example 69 except the wash procedure is completed with the methanol extraction. The resultant tint is a deep yellow color.

Depth of modification as a function of reactant, solvent, temperature of time

Example 72: This example shows a time course of reaction of a fluorinated thiol C₈F₁₇C₂H₄SH (= R₁SH) with the epoxy groups of a dry sample of polymer 12.

The solution, composed of 94 g R₁SH, 15 g triethylamine and 114 g absolute ethanol is maintained at 48° C. Dry samples with dimensions 15x15x1 mm are placed in the solution and at specified times removed and immediately extracted in 300 g of absolute ethanol (maintained at 50° C) for at least 4 hours with a minimum of one change of fresh ethanol. The samples are dried overnight in vacuo at 80° C. From the weight gained the yield of reaction is calculated. An SEM-electron microprobe scan for sulfur reveals the depth of the modifications as well as the reaction yield. The following table compares the yield calculated from the weight gain and the line scan.

Reaction Time (h)	Modification	
	% of cross-section by SEM - Line Scan	% by Weight
1	27	22
2	53	42
3	69	62
6	100	100
19 1/2	100	100

The results show that the weight gain is restricted to a small penetration depth and that even after prolonged reaction time no other functional group (i.e. non-epoxy) of the polymer react. The scans are rectangular in shape suggesting that the modification is diffusion controlled.

Example 73: This example shows the yield of reaction of an 18 hour posttreatment on polymer 12 with 1-thioglycerol under varying solvent and temperature conditions.

The solutions used are 0.3 M thioglycerol in aqueous ethanol containing sodium hydroxide to ensure all thiols existed as the thiolate anion. For example to make a solution in 45 % ethanol, 0.5 g 1-thioglycerol is dissolved in 4.16 g 1N NaOH, 3.74 ml water and 8.1 g absolute ethanol. Dry samples (15-x15-x1-mm) are placed in the solutions overnight. They are then extracted in about 20 ml of the same solvent composition for about 5 hours with three changes of fresh solvent and dried overnight in vacuo at 80° C. From the weight gain and SEM line scan the yield of reaction is determined and presented in the table shown below.

Temperature, °C	% EtOH	Modification	
		% of cross-section by SEM - Line Scan	% by Weight
25	15	0	0
25	45	10	15
25	70	32	33
50	15	0	0
50	45	32	29
50	70	69	57

The results show that the yield increases as the temperature is raised and the solvent swelling power (i.e. increased ethanol content) is increased. The modification proceeds as a sharply moving boundary from the surface into the interior of the polymer.

Effect of posttreatment on protein adsorption (Examples 74-80):

The protein of choice is albumin because of its abundance in body fluids. A 20x15x1-mm sample of polymer 12 or 13 is immersed in 1/8 ml of a ¹⁴C-albumin cocktail (10 mM potassium phosphate buffer pH 7.4, 144 mM NaCl, 0.48 mg/ml "cold" (non-radioactive) bovine serum albumin, 0.02 mg/ml ¹⁴C-bovine

serum albumin) for 2 hours at room temperature. The sample is then removed from the cocktail and then subjected to four sequential rinses of 15 ml of water to remove adhering droplets of cocktail. The sample is then stirred in 10 ml 0.1 M Tris-Cl pH 7 containing 1 % octylphenoxy polyethoxy ethanol (Triton X-100) for 2 hours at room temperature to elute unreacted but adsorbed protein and then rinsed in 15 ml water. Protein bound to the polymer (after Triton wash) is determined by scintillation counting (15 ml Packard Instagel). The amount of protein eluted by Triton is determined from a 1-ml (10 %) sample of the Triton wash counted in 15 ml of Instagel. The total protein bound to the polymer (i.e. after the original 4 water rinses) is the sum of the counts on the polymer after the Triton wash and the counts removed in the Triton wash.

Ex.	Polymer	Reactant	Counts Above Background (cpm)	
			Total	On Polymer After Triton Wash
74	12	1-thioglycerol	378	78
75	12	sarcosine	569	160
76	12	Na ₂ SO ₃	487	177
77	12	TRIS	1150	630
78	12	-(Control)	1314	974
79	13	mercaptoethanol	1220	530
80	13	-(Control)	1556	956

The examples shown above indicated dramatic changes in albumin adsorption and desorption (using Triton X-100) of treated samples compared to untreated control samples. The treatments on the above polymers inhibit protein adsorption and facilitate its removal with a nonionic surfactant.

Example 81: The example shows an improvement in wettability of a polymer sheet (PSM = 25 %, GMA = 15 %, neopentyl glycol dimethacrylate = 16 % isobornyl methacrylate = 44 %, prepared as in Example A) after room temperature treatment with an alcohol, polyvinyl alcohol (PVA), and a further improvement in this wettability with the crosslinking of this surface-bound PVA with 1,4-butanediol diglycidyl ether (BUDGE). Where indicated in the reaction sequence below, 2 g of aqueous 10 % (w/w) PVA (Elvanol 71-30, DuPont, 3.17 g 10 % (w/w) aqueous NaOH or 0.51 g BUDGE is added to the reaction bottle that contains the polymer sheet. The duration of the PVA, NaOH and BUDGE steps are 1 day, 2 1/4 hours and 2 hours respectively. Exhaustive water washes follow the treatments. The H₂O only (control) bottle is filled with H₂O at the beginning of the experiment.

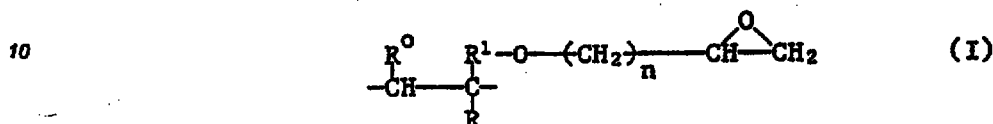
Retraction times are measured one week after the reaction.

Reaction Sequence	Retraction time, t_R (s)
a) PVA → NaOH → BUDGE	42
b) PVA → NaOH	23
c) H ₂ O only (control)	< 1



Claims

1. A contact lens possessing a wettable surface in its ocular aqueous environment of use fabricated from a hydrophobic addition polymer incorporating between 0.5 and 30 % by weight oxirane substituted units in the backbone thereof, wherein the outer surfaces of said lens contain an amount sufficient to increase the contact angle by at least 5° of a reaction product of said oxirane substituted units of the formula I



wherein R is hydrogen or methyl; R¹ is carbonyl or lower alkylene; n is 1 to 4; and R⁰ is hydrogen or

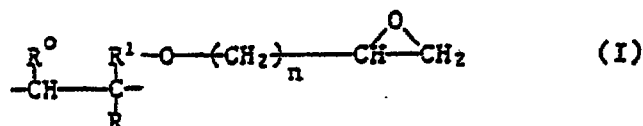


with a reactive water soluble nucleophile selected from an organic amine, alcohol, thiol, urea or thiourea, or sulfite, bisulfite or thiosulfate.

2. A contact lens according to claim 1, wherein the contact angle is increased by at least 10°.
3. A contact lens according to claim 1, wherein the contact angle is increased by at least 15°.
4. A contact lens according to claim 1, wherein the reaction product is a reaction product between said oxirane and a water soluble aliphatic, cycloaliphatic or aromatic primary or secondary amine having up to 12 carbon atoms.
5. A contact lens according to claim 1, wherein the reaction product is a reaction product between said oxirane and a water soluble pharmaceutically acceptable salt of a sulfite, bisulfite or thiosulfate.
6. A contact lens according to claim 1, wherein the hydrophobic polymer has a contact angle less than about 150°.
7. A contact lens according to claim 6, wherein said contact angle is less than 145°.
8. A contact lens according to claim 1, wherein the contact lens material is a silicone containing copolymer.
9. A contact lens according to claim 1 wherein the reaction product is prepared by conducting the reaction in aqueous medium at a pH between 3 and 10 and a temperature between 0 and 100° C for a period of time sufficient such that the surface region of the lens, up to 10 percent of the thickness thereof is reacted to form reaction product.

Revendications

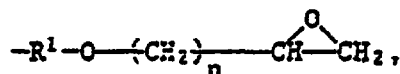
1. Lentille de contact ayant une surface mouillable dans son environnement oculaire aqueux d'utilisation, fabriquée à partir d'un polymère d'addition hydrophobe incorporant entre 0,5 et 30% en poids d'unités substituées par un oxiranne dans son tronc, de telle sorte que les surfaces extérieures de ladite lentille contiennent une quantité suffisante pour augmenter l'angle de contact d'au moins 5°, d'un produit de réaction desdites unités, substituées par un oxiranne, de formule I



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où R est un hydrogène ou un méthyle; R¹ est un carbonyle ou un alkylène inférieur; n est de 1 à 4; et R⁰ est un hydrogène ou

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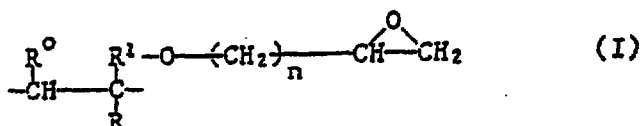
avec un nucléophile réactif soluble dans l'eau choisi parmi une amine, un alcool, un thiol, une urée ou thiourée, organiques, ou un sulfite, bisulfite ou thiosulfate.

2. Lentille de contact selon la revendication 1, dans laquelle l'angle de contact est augmenté d'au moins 10°.
- 20 3. Lentille de contact selon la revendication 1, dans laquelle l'angle de contact est augmenté d'au moins 15°.
4. Lentille de contact selon la revendication 1, dans laquelle le produit de réaction est un produit de réaction entre ledit oxiranne et une amine primaire ou secondaire, aliphatique, cycloaliphatique ou aromatique, soluble dans l'eau, ayant au maximum 12 atomes de carbone.
- 25 5. Lentille de contact selon la revendication 1, dans laquelle le produit de réaction est un produit de réaction entre ledit oxiranne et un sel soluble dans l'eau, pharmaceutiquement acceptable, d'un sulfite, bisulfite ou thiosulfate.
- 30 6. Lentille de contact selon la revendication 1, dans laquelle le polymère hydrophobe a un angle de contact inférieur à environ 150°.
7. Lentille de contact selon la revendication 6, dans laquelle ledit angle de contact est inférieur à 145°.
- 35 8. Lentille de contact selon la revendication 1, dans laquelle la matière de la lentille de contact est un copolymère contenant un silicone.
9. Lentille de contact selon la revendication 1, dans laquelle le produit de réaction est préparé en effectuant la réaction dans un milieu aqueux à un pH entre 3 et 10 et à une température entre 0 et 100° C, pendant une période de temps suffisante pour que la région superficielle de la lentille, jusqu'à 10 pourcents de son épaisseur, soit soumise à la réaction de façon à former le produit de réaction.
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45 Ansprüche

1. Kontaktlinse mit einer benetzbaren Oberfläche im wässrigen Augenmilieu der Anwendung, hergestellt aus einem hydrophoben Additionspolymer, das zwischen 0,5 und 30 Gew.-% Oxiran-substituierte Einheiten in der Hauptkette enthält, worin die äußeren Oberflächen dieser Linse eine für eine Vergrößerung des Kontaktwinkels um wenigstens 5° ausreichende Menge eines Reaktionsproduktes dieser Oxiran-substituierten Einheiten der Formel I
- 50

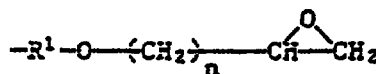
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worin R Wasserstoff oder Methyl ist; R¹ Carbonyl oder niederes Alkyl ist; n 1 bis 4 ist; und R⁰ Wasserstoff oder

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ist, mit einem reaktiven wasserlöslichen Nucleophil, das ausgewählt wird aus einem organischen Amin, Alkohol, Thiol, Harnstoff oder Thioharnstoff, oder Sulfit, Bisulfit oder Thiosulfat, enthalten.

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2. Kontaktlinse nach Anspruch 1, worin der Kontaktwinkel um wenigstens 10° vergrößert wird.
3. Kontaktlinse nach Anspruch 1, worin der Kontaktwinkel um wenigstens 15° vergrößert wird.
- 15 4. Kontaktlinse nach Anspruch 1, worin das Reaktionsprodukt ein Reaktionsprodukt zwischen diesem Oxiran und einem wasserlöslichen aliphatischen, cycloaliphatischen oder aromatischen primären oder sekundären Amin mit bis zu 12 Kohlenstoffatomen ist.
- 20 5. Kontaktlinse nach Anspruch 1, worin das Reaktionsprodukt ein Reaktionsprodukt zwischen diesem Oxiran und einem wasserlöslichen pharmazeutisch annehmbaren Salz eines Sulfits, Bisulfits oder Thiosulfats ist.
6. Kontaktlinse nach Anspruch 1, worin das hydrophobe Polymer einen Kontaktwinkel von weniger als etwa 150° hat.
- 25 7. Kontaktlinse nach Anspruch 6, worin dieser Kontaktwinkel weniger als 145° beträgt.
8. Kontaktlinse nach Anspruch 1, worin das Kontaktlinsenmaterial ein Silicon enthaltendes Copolymer ist.
- 30 9. Kontaktlinse nach Anspruch 1, worin das Reaktionsprodukt hergestellt wird, indem man die Reaktion in wäßrigem Medium bei einem pH-Wert zwischen 3 und 10 und bei einer Temperatur zwischen 0 und 100°C für eine Zeitdauer durchführt, die ausreichend ist, daß der Oberflächenbereich der Linse in einem Ausmaß von bis zu 10 Prozent der Dicke umgesetzt wird, um Reaktionsprodukt zu bilden.

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