

METHODS OF PREPARATION OF AN OLEFIN OLIGOMERIZATION CATALYST

FIELD OF THE INVENTION

[0001] The present invention relates to preparation of catalysts for use in a process for producing an olefin oligomer. More particularly, the present invention relates to preparing oligomerization catalysts comprising a chromium-containing compound, a nitrogen-containing compound, a metal alkyl, and an optional halide-containing compound for use in a process for producing an alpha-olefin oligomer comprising 1-hexene or 1-octene from ethylene.

BACKGROUND OF THE INVENTION

[0002] Olefin oligomerization catalysts are known in the art, but sometimes lack selectivity to a desired product and also have a low product yield. Enhancements in preparation methods for oligomerization catalysts to improve productivity and selectivity to the desired product can reduce catalyst cost and improve economics.

SUMMARY OF THE INVENTION

[0003] Disclosed herein is a method of making an olefin oligomerization catalyst, comprising contacting a chromium-containing compound, a heteroatomic ligand, and a metal alkyl, wherein the chromium-containing compound comprises less than about 5 weight percent chromium oligomers.

[0004] Further disclosed herein is method of making an olefin oligomerization catalyst comprising a chromium-containing compound, a nitrogen-containing compound, and a metal alkyl, the method comprising adding a composition comprising the chromium-containing compound to a composition comprising the metal alkyl.

[0005] Further disclosed herein is a method of making an olefin oligomerization catalyst comprising a chromium-containing compound, a nitrogen-containing compound, and a metal alkyl, the method comprising abating all or a portion of water, acidic protons, or both from a composition comprising the chromium-containing compound, a composition comprising the nitrogen-containing compound, or combinations thereof prior to or during the preparation of the catalyst.

[0006] Further disclosed herein is a method of making an olefin oligomerization catalyst comprising a chromium-containing compound, a heteroatomic ligand, a metal alkyl, the method comprising abating all or a portion of water, acidic protons, or both from a composition comprising the chromium-containing compound, a composition comprising the heteroatomic ligand, or combinations thereof, and wherein the heteroatomic ligand is described by the general formula $(R)_nA-B-C(R)_m$ wherein A and C are

independently selected from a group consisting of phosphorus arsenic, antimony, oxygen, bismuth, sulfur, selenium, and nitrogen; B is a linking group between A and C; each R is independently selected from any homo or hetero hydrocarbyl group; and n and m are determined by the respective valence and oxidation state of A and C.

[0007] Further disclosed herein are methods of oligomerizing olefins by contacting such catalysts with an alpha olefin.

BRIEF DESCRIPTION OF THE DRAWINGS

[0008] Figs. 1A through 1D illustrate various embodiments of a method of preparing an oligomerization catalyst comprising bulk addition of catalyst components.

[0009] Figs. 2A through 2D illustrate various embodiments of a method for abating water in the preparing of an oligomerization catalyst.

[0010] Figs. 3A through 3B illustrate various embodiments of a method for abating water in the preparing of an oligomerization catalyst.

[0011] Figs. 4A through 4E illustrate various embodiments of a method of preparing an oligomerization catalyst comprising simultaneous addition of catalyst components.

[0012] Fig. 5 is a graph of the average catalyst residence time (i.e. catalyst age) versus the purity of hexene produced.

[0013] Fig. 6 is the spectra of chromium(III) 2-ethylhexanoate.

[0014] Fig. 7 is the spectra of chromium(III) 2-ethylhexanoate, sample 17-1.

[0015] Fig. 8 is also the spectra of chromium(III) 2-ethylhexanoate, sample 17-2.

[0016] Fig. 9 is also the spectra of chromium(III) 2-ethylhexanoate, sample 17-3.

[0017] Fig. 10 is the spectra of chromium(III) 2-ethylhexanoate, sample 17-4a.

[0018] Fig. 11 is the spectra of azeotroped chromium(III) 2-ethylhexanoate, sample 17-4b.

[0019] Fig. 12 is the spectra of azeotroped chromium(III) 2-ethylhexanoate, sample 17-4c.

[0020] Fig. 13 is the spectra heated of chromium(III) 2-ethylhexanoate, sample 17-4d.

[0021] Fig. 14 is the spectra of heated chromium(III) 2-ethylhexanoate, sample 17-4e.

[0022] Fig. 15 is the spectra of chromium(III) 2-ethylhexanoate sample 17-4f used for oligomerization catalyst preparation.

DETAILED DESCRIPTION

[0023] The present invention relates to a method of making a catalyst for use in oligomerizing an olefin. Generally, the catalyst comprises a chromium-containing compound, a heteroatomic ligand, a metal alkyl, an optional halide-containing compound and an optional solvent. In an aspect, the method of making the catalyst comprises adding a composition comprising the chromium-containing compound to composition comprising the metal alkyl. In an aspect, the method of making the catalyst comprises abating all or a portion of water, acidic protons, or both from a composition comprising the chromium-containing compound, a composition comprising a non-metal halide-containing compound, a composition comprising a solvent, or combinations thereof prior to contact thereof with a composition comprising a metal halide-containing compound. Additional aspects relating to the method of making the catalyst are described herein. Additionally, the applicable heteroatomic ligands are described herein and are generally applicable to the method of making the catalyst. In embodiments, the heteroatomic ligand is a nitrogen-containing ligand.

[0024] The present invention also relates to a method of oligomerizing an olefin utilizing the method of making the catalyst as disclosed herein. In an aspect, the method of oligomerizing olefin can be a method of trimerizing an olefin and/or tetramerizing an olefin. In embodiments, the method of oligomerizing an olefin utilizing is a method for trimerizing ethylene; alternatively a method for tetramerizing ethylene. In an embodiment, a method for trimerizing an olefin, e.g., ethylene to 1-hexene, is disclosed wherein the catalyst does not include the optional halide-containing compound. Alternatively, a method for trimerizing an olefin, e.g., ethylene to 1-hexene, is disclosed wherein the catalyst does include the optional halide-containing compound. In an embodiment, a method for tetramerizing an olefin, e.g., ethylene to 1-octene, is disclosed wherein the catalyst does not include the optional halide-containing compound. Alternatively, a method for tetramerizing an olefin, e.g., ethylene to 1-octene, is disclosed wherein the catalyst does include the optional halide-containing compound. As used herein, the term optional halide-containing compound is intended to cover embodiments where the halide-containing compound is present as well as embodiments where the halide-containing compound is not present.

[0025] The present invention also relates to aspects of a chromium-containing compound utilized in the method of making the catalyst disclosed herein and the method of oligomerizing an olefin utilizing the

method of making the catalyst disclosed herein. In embodiments, the chromium-containing compound can comprise a chromium carboxylate. In some embodiments, the chromium carboxylate can be chromium(III) tris-2-ethylhexanoate. In an aspect, the composition comprising a chromium-containing compound contains less than 10 weight percent chromium oligomers based upon the total weight of all chromium species within the composition the chromium-containing compound. Other aspects of the chromium-containing compound and the composition comprising the chromium-containing compound are described herein. Additionally, the aspects of the chromium containing compound are also applicable to the chromium-containing compound utilized in the method of making the catalyst and the method of oligomerizing an olefin.

[0026] As used herein, a catalyst component includes a chromium-containing compound, a heteroatomic ligand, a metal alkyl, an optional halide-containing compound, an optional solvent, or combinations thereof. In an embodiment, a catalyst component includes a chromium-containing compound, a nitrogen-containing compound, a metal alkyl, an optional halide-containing compound, an optional solvent, or combinations thereof. In the various embodiments disclosed herein, contacting of catalyst components may occur in one or more contact zones. A contact zone is a zone in which the components are commingled and/or combined, and thereby contacted. The contact zone may be disposed in a vessel, e.g. a storage tank, tote, container, mixing vessel, reactor, etc.; a length of pipe, e.g. a tee, inlet, injection port, or header for combining component feed lines into a common line; or any other suitable apparatus for bringing the components into contact. As used herein, the terms contacted and combined refer to any addition sequence, order, or concentration for contacting or combining two or more catalyst components. The term added to refers to a first catalyst component added, e.g., poured, into a second catalyst component. Where a first catalyst component is added to a second catalyst component, the initial concentration, or molar ratio, of the first catalyst component compared to the second catalyst component typically is relatively small and increases over the duration of the addition. In some embodiments, contacting of components may occur in one or more upstream contact zone(s) prior to further contacting with other catalyst component(s) in one or more downstream contact zone(s). Where a plurality of contact zones are employed, contacting may occur simultaneously across the contact zones, sequentially across the contact zones, or both, as is suitable for a given embodiment. Contacting may be carried out in a batch or continuous process, as is suitable for a given embodiment.

[0027] In embodiments utilizing a vessel for contacting the components, the components may be optionally mixed by a mixer disposed in the vessel and the formed mixture may then be removed for subsequent processing. In embodiments utilizing a tee or other means for combining lines such as a header, an optional in-line mixer may be placed in the commingled catalyst feed line to ensure that adequate contacting of the combined components takes place, and the mixture is thus formed as it passes through

the commingled feed line. Where a method of making a catalyst recites contact or combination of catalyst components, such may be carried out by contacting or combining all or a portion of such components in various embodiments.

[0028] As used herein, a composition comprising a catalyst component includes the catalyst component alone or in combination with one or more additional compounds, solvents, or both. None, some, or all of the contacting steps may be carried out in the presence of a solvent (sometimes referred to as an optional solvent), which may be introduced to a contact zone via inclusion with one or more compositions comprising a catalyst component or may be introduced separately to a contact zone, for example in a solvent line or as an initial charge to a contact zone.

[0029] Disclosed herein is a method of making a catalyst comprising a chromium-containing compound, a nitrogen-containing compound, a metal alkyl, an optional halide-containing compound, and optionally a solvent for use in oligomerizing an olefin, wherein a composition comprising the chromium-containing compound is contacted in a contact zone with a composition comprising the metal alkyl. In Fig. 1, four embodiments for contacting the composition comprising the chromium-containing compound with the composition comprising the metal alkyl in a contact zone are illustrated. Figures 1A through 1D are included as illustrative representations of embodiments of the present disclosure and do not limit the disclosure. For example, Figures 1A-1D are described in the context of the use of a pyrrole-containing compound as the nitrogen-containing compound, but it should be understood that other nitrogen-containing compounds as described herein may be used in the embodiments shown in Figures 1A-1D or any other embodiments disclosed herein. Furthermore, Figures 1A-1D are described in the context of the use of a halide-containing compound, but it should be understood that in alternative embodiments, the halide-containing compound may be omitted from the embodiments shown in Figures 1A-1D or from any other embodiments disclosed herein.

[0030] In an embodiment as illustrated in Fig. 1A, the composition comprising the metal alkyl may be disposed in contact zone 115 and the composition comprising the chromium-containing compound may be contacted with or added to the composition comprising the metal alkyl present in contact zone 115 via line 110. The final catalyst composition may be recovered as a product via line 170. The composition comprising the chromium-containing compound in line 110 may further comprise a pyrrole-containing compound, a non-metal halide-containing compound, the solvent, or combinations thereof. The composition comprising the chromium-containing compound may also comprise an amount of non-halide metal alkyl to abate undesired water, acidic protons, or both, as disclosed in more detail herein. The final catalyst composition may be further dilute with a solvent (which may not be identical to the catalyst preparation solvent) prior to use in the oligomerization reaction.

[0031] The composition comprising the metal alkyl present in contact zone 115, may comprise the pyrrole-containing compound, the halide-containing compound, the solvent, or combinations thereof. The halide-containing compound may be a metal halide, non-metal halide, or combinations thereof. The composition comprising the metal alkyl may also comprise a metal alkyl halide, a non-halide metal alkyl, a non-metal halide, a metal halide, or combinations thereof. The metal alkyl halide in this and other embodiments may comprise diethylaluminum chloride (DEAC) and the non-halide metal alkyl may comprise triethyl aluminum (TEA). In an embodiment the metal alkyl may be the halide-containing compound, e.g. DEAC is the halide-containing compound and the metal alkyl.

[0032] In an embodiment as illustrated in Fig. 1B, a pyrrole-chromium mixture may be formed in contact zone 225 by contacting a composition comprising the pyrrole-containing compound fed to contact zone 225 via line 220 and the composition comprising the chromium-containing compound fed to contact zone 225 via line 210, which may occur about instantaneously or over a first period of time of from about 1 minute to about 12 hours, alternatively from about 1 minute to about 6 hours, alternatively from about 1 minute to about 3 hours, alternatively from about 1 hour to about 2 hours. Introduction of the composition comprising the chromium-containing compound and the composition comprising the pyrrole-containing compound to contact zone 225 may be sequential (e.g. chromium followed by pyrrole or vice-versa) or simultaneous. Once the pyrrole-chromium mixture has been contacted in contact zone 225 the pyrrole-chromium mixture from contact zone 225 may be contacted with or added to the composition comprising the metal alkyl present in contact zone 215 via line 240, which may occur about instantaneously or over a second period of time of from about 1 minute to about 12 hours, alternatively from about 1 minute to about 6 hours, alternatively from about 1 minute to about 3 hours, to form the final catalyst product in contact zone 215. The final catalyst product may be withdrawn from contact zone 215 via line 270. The final catalyst composition may be further dilute with a solvent (which may not be identical to the catalyst preparation solvent) prior to use in the oligomerization reaction.

[0033] The composition comprising the pyrrole-containing compound in line 220 and the composition comprising the chromium-containing compound in line 210 may be contacted, e.g., over the first period of time, at an about constant pyrrole to chromium (Py:Cr) molar ratio or alternatively at a variable Py:Cr molar ratio to form the pyrrole-chromium mixture in contact zone 225. The pyrrole-chromium mixture in contact zone 225 may then be contacted with or added to, e.g., over the second period of time, the metal alkyl present in contact zone 215 via line 240, or alternatively already present in contact zone 215, at an about constant Py:Cr molar ratio, for example in the range of from about 1.0:1 to about 4.0:1. Alternatively, the pyrrole-chromium mixture in contact zone 225 may then be contacted with or added to, e.g., over the second period of time, the metal alkyl present in contact zone 215 via line 240 at a variable Py:Cr molar ratio. In an embodiment the variable Py:Cr molar ratio is decreasing over the

second period of time where a decreasing Py:Cr molar ratio refers to a general decreasing trend in the molar ratio from the start of the addition sequence to the finish and occasional increases in the ratio within the overall decreasing trend are acceptable. In an embodiment a decreasing trend of the Py:Cr refers to the specific situation where the ending Py:Cr ratio is less than the beginning Py:Cr ratio. In an embodiment, an initial Py:Cr molar ratio at the start of the addition may be greater than the final Py:Cr molar ratio of the catalyst; and an ending Py:Cr molar ratio at the end of the addition may be less than the final Py:Cr molar ratio of the catalyst. In an embodiment, the final Py:Cr molar ratio of the catalyst may be in a range of from about 1.0:1 to about 4.0:1; the initial Py:Cr molar ratio may be greater than about 6:1, alternatively greater than about 20:1, alternatively greater than about 40:1, alternatively greater than about 60:1; and the ending Py:Cr molar ratio may be greater than or equal to about 0, alternatively greater than or equal to about 0.1:1, alternatively greater than or equal to about 0.3:1, and alternatively greater than or equal to about 0.6:1. In an embodiment, the initial Py:Cr molar ratio is about twice the final Py:Cr molar ratio of the catalyst during a first about one-half of the addition and the ending Py:Cr molar ratio is about 0 during a second about one-half of the addition, wherein the final Py:Cr molar ratio of the catalyst is in a range of from about 1.0:1 to about 4.0:1. Introduction of a pyrrole-containing compound and a chromium-containing compound in a contact zone (e.g., formation of a Py:Cr mixture) as disclosed in various embodiments may be carried out as disclosed in this paragraph, including but not limited to the embodiments shown in Figs. 1D, 2C, 2D, 3B, and 4A-E.

[0034] The composition comprising the chromium-containing compound in line 210 may comprise a non-metal halide-containing compound, the solvent, or combinations thereof. The composition comprising the pyrrole-containing compound in line 220 may comprise a non-metal halide-containing compound, the solvent, or combinations thereof. The composition comprising the chromium-containing compound in line 210, the composition comprising the pyrrole-containing compound in line 220, or both may also comprise an amount of non-halide metal alkyl to abate undesired water, acidic protons, or both as disclosed herein. Alternatively, the non-halide metal alkyl may be contacted with or added to the pyrrole-chromium mixture, for example in line 240 via line 230, in contact zone 225 (not shown), or both, to abate undesired water, acidic protons, or both. The composition comprising the metal alkyl present in contact zone 215, may comprise the halide-containing compound, the solvent, or combinations thereof. The composition comprising the metal alkyl may also comprise a metal alkyl halide, a non-halide metal alkyl, a metal halide, non-metal halide, or combinations thereof.

[0035] In an embodiment as shown in Fig. 1C, a pyrrole-metal alkyl mixture may be formed in contact zone 325 by contacting the composition comprising the pyrrole-containing compound fed to contact zone 325 via line 320 with the composition comprising the metal alkyl fed to contact zone 325 via line 315 which may occur about instantaneously or over a first period of time. Addition of the

composition comprising the pyrrole-containing compound and the composition comprising the metal alkyl to contact zone 325 may be sequential (e.g. pyrrole followed by metal alkyl or vice-versa) or simultaneous. Once the pyrrole-metal alkyl mixture has been contacted in contact zone 325 the pyrrole-metal alkyl mixture from contact zone 325 may be disposed via line 360 in contact zone 335. The composition comprising the chromium-containing compound may then be contacted with or added to contact zone 335 via line 310, which may occur about instantaneously or over a second period of time. The composition comprising the chromium-containing compound is thus contacted with or added to the pyrrole-metal alkyl mixture present in contact zone 335, to form the final catalyst product in contact zone 335. Addition of the composition comprising the pyrrole-metal alkyl mixture and the composition comprising the chromium-containing compound to contact zone 335 may be sequential (e.g. pyrrole-metal alkyl followed by the chromium containing compound or vice-versa) or simultaneous. The final catalyst product may be withdrawn from contact zone 335 via line 370. The final catalyst composition may be further diluted with a solvent (which may not be identical to the catalyst preparation solvent) prior to use in the oligomerization reaction.

[0036] Although the embodiment shown in Fig. 1C shows two contact zones being used to perform the addition sequences, the addition sequences could alternatively be performed in a single contact zone, for example, in contact zone 325. In this embodiment, the composition comprising the metal alkyl may first be placed in the contact zone. In a second step the composition comprising the pyrrole-containing compound may be contacted with or added to the composition comprising the metal alkyl present in the contact zone (or visa-versa) to adequately contact and form the pyrrole-metal alkyl mixture. In a third step, the composition containing the chromium-containing compound may be contacted with or added to the pyrrole-metal alkyl mixture to form the final catalyst product.

[0037] The composition comprising the chromium-containing compound in line 310 may comprise a non-metal halide-containing compound, the solvent, or combinations thereof. The composition comprising the pyrrole-containing compound in line 320 may comprise a non-metal halide-containing compound, the solvent, or combinations thereof. The composition comprising the chromium-containing compound in line 310, the composition comprising the pyrrole-containing compound in line 320, or both may comprise an amount of non-halide metal alkyl to abate undesired water, acidic protons, or both. The composition comprising the metal alkyl in line 315, may comprise the halide-containing compound, the solvent, or combinations thereof. The composition comprising the metal alkyl may also comprise a metal alkyl halide, a non-halide metal alkyl, a metal halide, non-metal halide, or combinations thereof.

[0038] In an embodiment as shown in Fig. 1D, a composition comprising the pyrrole-containing compound in line 420 and a composition comprising the chromium-containing compound in line 410 may be simultaneously contacted with or added to, which may occur about instantaneously or over a period of

time, with a composition comprising the metal alkyl present in contact zone 415, and a final catalyst product may be withdrawn from contact zone 415 via line 470. The final catalyst composition may be further diluted with a solvent (which may not be identical to the catalyst preparation solvent) prior to use in the oligomerization reaction. The composition comprising the chromium-containing compound and the composition comprising the pyrrole-containing compound may be contacted with or added to the composition comprising the metal alkyl at Py:Cr molar ratios described previously.

[0039] The composition comprising the chromium-containing compound in line 410 may comprise a non-metal halide-containing compound, the solvent, or combinations thereof. The composition comprising the pyrrole-containing compound in line 420 may comprise a non-metal halide-containing compound, the solvent, or combinations thereof. In the embodiment shown in Fig. 1D, the composition comprising the metal alkyl in contact zone 415, may comprise the halide-containing compound, the solvent, or combinations thereof, each added to contact zone 415 through various input lines not shown in Fig. 1D. The composition comprising the metal alkyl may also comprise a metal alkyl halide, a non-halide metal alkyl, a metal halide, non-metal halide, or combinations thereof. The composition comprising the chromium-containing compound in line 410, the composition comprising the pyrrole-containing compound in line 420, or both may comprise an amount of non-halide metal alkyl to abate undesired water, acidic protons, or both.

[0040] Further disclosed herein is a method of making a catalyst comprising abating all or a portion of water, acidic protons, or both from a composition comprising the chromium-containing compound, a composition comprising the nitrogen-containing compound, a composition comprising the optional non-metal halide-containing compound, a composition comprising the solvent, or combinations thereof prior to contact thereof with a composition comprising the metal halide-containing compound. Abating water, acidic protons, or both may include neutralizing acidic protons; physically removing water; physically removing acidic protons; chemically binding or reacting free water such that the water is no longer free; or combinations thereof. The amount of water, acid protons, or both removed from the catalyst component may be determined using known methods, for example infrared analysis to determine water content.

[0041] In embodiments to prepare a catalyst, one or more of the catalyst components may contain water, for example the composition comprising the chromium-containing compound. Water may be present in a catalyst compound, for example as a contaminant or as a co-product produced during the preparation of the catalyst compound. For example, water may be co-produced during preparation of the chromium-containing compound, and such water may complex with the chromium. Acidic protons may also be present, for example carboxylic acid (e.g., ethylhexanoic acid) remaining from production of the chromium-containing compound (e.g., chromium tris(2-ethylhexanoate)). This free water as well as acid

present in the chromium source can subsequently react with a metal halide present in the catalyst, for example the metal alkyl halide such as DEAC, to form corrosive compounds, e.g. hydrogen halide compound (e.g. hydrochloric acid). Such compounds may cause corrosion in downstream equipment over time, in particular when heated, for example in downstream fractionation facilities. Accordingly, it may be desirable to abate water, acidic protons, or both, when making the catalyst to prevent downstream formation of potentially corrosive by-products. The presence of water may also reduce the effectiveness of the metal alkyl.

[0042] Furthermore, in embodiments of a method of preparing a catalyst, impurities in the catalyst components can participate in unwanted side reactions leading to the formation of precipitates. These precipitates may lead to further unwanted reactions, for example polymer formation in the oligomerization of ethylene to an olefin composition comprising 1-hexene or 1-octene. Water may be an initiator of the precipitation reactions and therefore may be desirably abated from the catalyst components to improve selectivity to 1-hexene or 1-octene. Abating water, acidic protons, or both may also have beneficial impact on catalyst efficiency, even where corrosive compounds are produced. For example, in an embodiment, water is abated from one or more catalyst components by contact thereof with a corrosive abatement compound such as a halide-containing compound, which reacts with and abates the water. Reactions of water with a corrosive abatement compound such as a halide-containing compound may produce a corrosive compound, e.g., HCl, and such should be taken into account in the overall design of the system. Examples of suitable halide-containing compounds for reaction with water include a metal halide, a metal alkyl halide, a non-halide metal alkyl and a metal halide, a non-metal halide, or combinations thereof. The use of a halide-containing compound to abate water may be used in place of or in addition to other water abatement embodiments disclosed herein such as the use of a non-halide metal alkyl to abate water.

[0043] In an embodiment, water, acidic protons, or both may be abated by pre-contacting one or more catalyst components with a non-corrosive abatement compound, which is a compound that does not form a corrosive compound such as a hydrogen halide compound upon contact with the water, acidic protons, or both. Non-corrosive abatement compounds include, for example, a non-halide metal alkyl such as TEA. Corrosive abatement compounds are compounds that can form a corrosive compound upon contact with water, acidic protons or both such as (i) a metal alkyl halide, (ii) a metal halide and a metal alkyl, and (iii) a non-metal halide and a metal alkyl. The corrosive abatement compounds also include any other combination of compounds that form a corrosive compound upon contact with water, acidic proton, or both.

[0044] In an embodiment, one or more catalyst components such as a composition comprising the chromium-containing compound, a composition comprising the nitrogen-containing compound, an

optional non-metal halide-containing compound, a solvent, or combinations thereof, are contacted with a non-halide metal alkyl to abate water, acidic protons, or both. The non-halide metal alkyl can react with free water, acid protons, or both contained in the catalyst component(s) when pre-contacted to abate water, acidic protons, or both. The non-halide metal alkyl may be pre-mixed in a contact zone with the one or more catalyst components. The pre-mix may be made by either adding the non-halide metal alkyl to the catalyst component(s) or vice versa, and in an embodiment, the pre-mix may be made by adding the non-halide metal alkyl to the catalyst component(s). These additions can be made in various ratios as described below.

[0045] In an embodiment, the non-halide metal alkyl is added to or contacted with a composition comprising the chromium-containing compound. Given that the chromium may react with the non-halide metal alkyl to form a gel, it may be desirable to maintain a low concentration of non-halide metal alkyl by adding it to the composition comprising the chromium-containing compound, so that there may only be an amount available to react with the water and acid. Conversely, with a high concentration of non-halide metal alkyl, such as can occur when adding the composition comprising the chromium-containing compound to the non-halide metal alkyl, more non-halide metal alkyl would be available to react with the chromium (and thereby form a gel) after the water and acid were removed.

[0046] In each embodiment, the water or acid abating substance (e.g., a non-halide metal alkyl) may be contacted with or added to one or more catalyst components in an amount effective to abate substantially all free/available water, acidic protons, or both from some or all of the components contacted with the non-halide metal alkyl. In an embodiment, the amount of non-halide metal alkyl contacted with or added to such components is small relative to the amount of the catalyst components to which it is being contacted with or added to. In an embodiment, the portion of the non-halide metal alkyl contacted with or added to a catalyst component(s) may be less than or equal to about 30 weight percent of the catalyst component(s) to which it is contacted with or added to; alternatively less than about 20 weight percent of the catalyst component(s) to which it is contacted with or added to; alternatively less than about 10 weight percent of the catalyst component(s) to which it is contacted with or added to; alternatively less than about 5 weight percent of the catalyst component(s) to which it is contacted with or added to. In an embodiment, the portion of the non-halide metal alkyl contacted with or added to a catalyst component(s) may be less than or equal to about 120 mole percent of the catalyst component(s) to which it is contacted with or added to; alternatively less than about 80 mole percent of the catalyst component(s) to which it is contacted with or added to; alternatively less than about 40 mole percent of the catalyst component(s) to which it is contacted with or added to; alternatively less than about 20 mole percent of the catalyst component(s) to which it is contacted with or added to. The non-halide metal alkyl may be contacted with or added to a catalyst component(s) in an amount such that the non-halide metal alkyl to catalyst

component(s) molar ratio may be less than about 1.5:1, alternatively less than about 1.2:1, alternatively less than about 1:1. The non-halide metal alkyl may be contacted with or added to a catalyst component(s) in a molar ratio sufficient to abate at least about 25% of the water, acidic protons, or both associated with the catalyst component(s) present in the pre-contacting contact zone; alternatively at least about 90% of the water, acidic protons, or both associated with the catalyst component(s) present in the pre-contacting contact zone; alternatively at least about 100% of the water, acidic protons, or both associated with the catalyst component(s) present in the pre-contacting contact zone; alternatively in an amount that may be at least about 10% in excess of an amount sufficient to abate at least about 100% of the water, acidic protons, or both associated with the catalyst component(s) present in the pre-contacting contact zone; alternatively in an amount that may be at least about 20% in excess of an amount sufficient to abate at least about 100% of the water, acidic protons, or both associated with the catalyst component(s) present in the pre-contacting contact zone; alternatively in an amount that may be at least about 30% in excess of an amount sufficient to abate at least about 100% of the water, acidic protons, or both associated with the catalyst component(s) present in the pre-contacting contact zone; alternatively in an amount that may be at least about 100% in excess of an amount sufficient to abate at least about 100% of the water, acidic protons, or both associated with the catalyst component(s) present in the pre-contacting contact zone; or alternatively in an amount that may be at least about 200% in excess of an amount sufficient to abate at least about 100% of the water, acidic protons, or both associated with the catalyst component(s) present in the pre-contacting contact zone.

[0047] Upon abatement of water, acidic protons, or both from one or more catalyst components, such abated catalyst components may be stored until needed for preparation of a catalyst composition. Such storage may or may not be in the presence of a solvent. The pre-mix comprising a portion of non-halide metal alkyl and one or more abated catalyst component(s) may then be contacted with the remaining catalyst components including the metal alkyl halide to form the final catalyst product. The remaining catalyst components may also comprise additional non-halide metal alkyl to comprise the total non-halide metal alkyl composition in the final catalyst. In an embodiment, the additional non-halide metal alkyl may be the same as that used in the pre-mix. Alternatively, the additional non-halide metal alkyl may be different from that used in the pre-mix.

[0048] Figures 2A - 2D represent various embodiments for abating water, acidic protons, or both in the composition comprising the chromium-containing compound, the composition comprising the nitrogen-containing compound, or both prior to contact with the composition comprising a metal halide-containing compound. Figures 2A through 2D are included as illustrative representations of embodiments of the present disclosure and do not limit the disclosure. For example, Figures 2A-2D are described in the context of the use of a pyrrole-containing compound as the nitrogen-containing compound, but it should

be understood that other nitrogen-containing compounds as described herein may be used in the embodiments shown in Figures 2A-2D or any other embodiments disclosed herein. Furthermore, Figures 2A-2D are described in the context of the use of a halide-containing compound, but it should be understood that in alternative embodiments, the halide-containing compound may be omitted from the embodiments shown in Figures 1A-1D or from any other embodiments disclosed herein. Various embodiments for abating water, acidic protons, or both may be combined to increase overall effectiveness.

[0049] The composition comprising a chromium-containing compound may be contacted with the non-halide metal alkyl to form a mixture prior to contacting the mixture with the remaining catalyst components. In an embodiment shown in Fig. 2A a composition containing the chromium-containing compound may be disposed in contact zone 510, the placement of which may take place via input line 505. The composition in contact zone 510 may optionally contain solvent, other catalyst components, or combinations thereof, provided that contact zone 510 does not comprise (i) a metal alkyl halide, (ii) a metal halide and a metal alkyl, or (iii) a non-metal halide and a metal alkyl. Non-halide metal alkyl, optionally in solvent, may be added to the composition containing a chromium-containing compound in contact zone 510 via line 530. The non-halide metal alkyl may be added in an amount less than or equal to about 30 weight percent of the composition containing the chromium-containing compound to which it is added or in other amounts as disclosed herein.

[0050] The resultant mixture in contact zone 510 may then be passed from contact zone 510 via line 511 and optionally fed into a filter 512, comprising dry (free of any water) filter medium, for filtering any precipitate that may have formed from the mixture. The precipitate may be filtered and the filtrate may be passed via line 513 into contact zone 515 for contacting with the remaining catalyst components including a composition comprising the metal alkyl, the pyrrole-containing compound, the halide-containing compound (e.g., a metal halide or non-metal halide), the solvent, any remaining non-halide metal alkyl, metal alkyl halide, or combinations thereof, which may be placed into contact zone 515 via various input lines not shown in Fig. 2A. A catalyst product may then be withdrawn from contact zone 515 via line 570. Where filtering is omitted, the remaining catalyst components may be alternatively contacted in contact zone 510.

[0051] The composition comprising a pyrrole-containing compound may be contacted with the non-halide metal alkyl to form a mixture prior to contacting the mixture with the remaining catalyst components. In an embodiment shown in Fig. 2B a composition comprising a pyrrole-containing compound may be disposed in contact zone 620 via input line 607. The composition in contact zone 620 may optionally contain solvent, other catalyst components, or combinations thereof, provided that contact zone 620 does not comprise (i) a metal alkyl halide, (ii) a metal halide and a metal alkyl, or (iii) a non-metal halide and a metal alkyl. Non-halide metal alkyl, which may be in solvent, may be added to the

composition containing a nitrogen-containing compound in contact zone 620 via line 630. The non-halide metal alkyl may be added in an amount less than or equal to about 10 weight percent of the composition containing the pyrrole-containing compound to which it is added or in other amounts as disclosed herein.

[0052] The resultant mixture in contact zone 620 may then be passed from contact zone 620 via line 621 and optionally filtered (not shown) to remove any precipitate that may have formed in the mixture. The resultant mixture may then be fed into contact zone 615 for contacting with the remaining catalyst components including a composition comprising the metal alkyl, the chromium-containing compound, the halide-containing compound (e.g., a metal halide or non-metal halide), the solvent, any remaining non-halide metal alkyl, metal alkyl halide, or combinations thereof, which may be placed into contact zone 615 via various input lines not shown in Fig. 2B. A catalyst product may then be withdrawn from contact zone 615 via line 670. Where filtering is omitted, the remaining catalyst components may be alternatively contacted in contact zone 620 via various input lines not shown in Fig. 2B.

[0053] The composition comprising the chromium containing compound may be contacted with the composition comprising pyrrole-containing compound to form a mixture prior to contacting the mixture with the non-halide metal alkyl. In an embodiment as illustrated in Fig. 2C, a pyrrole-chromium mixture may be formed in contact zone 725 by contacting a composition comprising the pyrrole-containing compound fed to contact zone 725 via line 720 and the composition comprising the chromium-containing compound fed to contact zone 725 via line 710, which may occur about instantaneously or over a first period of time. Feeding of the composition comprising the chromium-containing compound and the composition comprising the pyrrole-containing compound to contact zone 725 may be sequential (e.g. chromium followed by pyrrole or vice-versa) or simultaneous and at constant or varying Py:Cr ratios as disclosed previously. Once the pyrrole-chromium mixture has been contacted in contact zone 725 the pyrrole-chromium mixture from contact zone 725 may be placed in contact zone 731 via line 740. The pyrrole-chromium mixture may optionally contain solvent, other catalyst components, or combinations thereof, but does not comprise (i) a metal alkyl halide, (ii) a metal halide and a metal alkyl, or (iii) a non-metal halide and a metal alkyl. Non-halide metal alkyl, which may be in solvent, may be added to the pyrrole-chromium mixture in contact zone 731 via line 730. The non-halide metal alkyl may be added in an amount less than or equal to about 10 weight percent of the pyrrole-chromium mixture to which it is added or in other amounts as disclosed herein. Although not shown in Fig. 2C, contact zone 725 and contact zone 731 may be the same contact zone providing that the addition sequence as described above remains the same.

[0054] The resultant mixture in contact zone 731 may then be passed from contact zone 731 via line 732 and may optionally be filtered (not shown) to remove any precipitate that may have formed in the mixture. The mixture may be fed into contact zone 715 for contacting with the remaining catalyst

components including a composition comprising the metal alkyl, the halide-containing compound (e.g., a metal halide or non-metal halide), the solvent, any remaining non-halide metal alkyl, metal alkyl halide, or combinations thereof, which may be placed into contact zone 715 via various input lines not shown in Fig. 2C. A catalyst product may then be withdrawn from contact zone 715 via line 770 and may optionally be filtered in a filter (not shown). Where filtering is omitted, remaining catalyst components may be alternatively contacted in contact zone 725 or 731.

[0055] The composition comprising a chromium-containing compound may be contacted with the non-halide metal alkyl to form a first mixture; the composition comprising a pyrrole-containing compound may be contacted with the non-halide metal alkyl to form a second mixture; and the first and second mixtures may be contacted with the remaining catalyst components. In an embodiment shown in Fig. 2D a composition containing a chromium-containing compound may be disposed in contact zone 810, the placement of which takes place via input line 805. The composition in contact zone 810 may optionally contain solvent, other catalyst components, or combinations thereof, but contact zone 810 does not comprise (i) a metal alkyl halide, (ii) a metal halide and a metal alkyl, or (iii) a non-metal halide and a metal alkyl. Non-halide metal alkyl, which may be in solvent, may be added to the composition containing a chromium-containing compound in contact zone 810 via line 830 forming a first mixture. The non-halide metal alkyl may be added in an amount less than or equal to about 10 weight percent of the composition containing the chromium-containing compound to which it is added or in other amounts as disclosed herein.

[0056] A second mixture can be formed in contact zone 820. The composition comprising a pyrrole-containing compound may be disposed in contact zone 820, the placement of which takes place via input line 807. The composition comprising a pyrrole-containing compound in contact zone 820 may optionally contain solvent, other catalyst components, or combinations thereof, but does not comprise (i) a metal alkyl halide, (ii) a metal halide and a metal alkyl, or (iii) a non-metal halide and a metal alkyl. Non-halide metal alkyl, which may be in solvent, may be added to the composition containing a pyrrole-containing compound in contact zone 820 via line 831 forming the second mixture. The non-halide metal alkyl may be added in an amount less than or equal to about 10 weight percent of the composition containing the nitrogen-containing compound to which it is added or in other amounts as disclosed herein.

[0057] The first mixture, second mixture, or both may optionally be filtered (not shown) to remove any precipitate that may have formed in the mixtures. Optionally, either the first, second, or both mixtures may be stored. The first and second mixtures may then be fed into contact zone 815 via lines 811 and 821, respectively for contacting with the remaining catalyst components including the composition comprising metal halide. Alternatively, although not shown in Fig. 2D the first and second mixtures may be contacted separately in another contact zone prior to being fed via a commingled feed line into contact

zone 815, and such commingled feed line may be optionally filtered to remove any precipitate that may have formed. Contact zone 815 initially may be comprised of a composition comprising the metal alkyl, a halide-containing compound (e.g., a metal halide or non-metal halide), a solvent, the remaining non-halide metal alkyl, metal alkyl halide, or combinations thereof, all of which have been placed into contact zone 815 via various input lines not shown in Fig. 2D. A catalyst product may then be withdrawn from contact zone 815 via line 870 and optionally filtered (filter not shown). In alternative embodiments, remaining catalyst components may be contacted in contact zone 810 or 820.

[0058] The addition of the composition comprising the pyrrole-containing compound and the composition comprising the chromium-containing compound as shown in Figs. 2C and 2D may be made in constant or varying Py:Cr ratios as disclosed previously.

[0059] Water may be removed from the chromium-containing compound prior to contact with the metal halide-containing compound according to various water abatement embodiment disclosed herein. In an embodiment, the chromium-containing catalyst feedstock may be contacted with an azeotropic solvent such as an aromatic compound, paraffin solvent, chlorinated solvent, other solvent, or mixture of solvents capable of forming an azeotrope with water. The azeotropic solvent, the chromium-containing compound, and any water present form a solution and the solution may be subjected to an azeotropic distillation to remove the water, wherein the solvent-water azeotrope is a lower boiling component. Optionally, the solvent used to remove water by azeotropic distillation may be recovered after the azeotropic distillation. In an embodiment, the azeotropic solvent used to remove water using azeotropic distillation may comprise ethylbenzene, benzene, meta-xylene, ortho-xylene, para-xylene, mixed xylenes, toluene, octane, nonane, heptane, hexane, mixed hexanes, cyclohexane, carbon tetrachloride, chloroform, dichloromethane, 1,1,2 trichloroethane, or combinations thereof. The amount of water removed from a catalyst component by various abatement methods may be monitored using known analytical methods such as infrared analysis.

[0060] In an embodiment shown in Fig. 3A a composition containing a chromium-containing compound may be disposed in contact zone 910, the placement of which takes place via an input line 905. The composition in contact zone 910 may optionally contain solvent, other catalyst components, or combinations thereof, but contact zone 910 does not comprise (i) a metal alkyl halide, (ii) a metal halide and a metal alkyl, or (iii) a non-metal halide and a metal alkyl. An azeotropic solvent, e.g., a composition comprising an aromatic compound such as ethylbenzene, may be added to the composition containing a chromium-containing compound in contact zone 910 via line 902 or directly added to separator 900. The azeotropic solvent may be added in an amount effective to form an azeotropic solution with the chromium-containing compound. In an embodiment, the azeotropic solvent may be added in an amount from about 0.5 to about 1000 times the weight of the composition containing the chromium-containing

compound to which it is added, alternatively from about 0.5 to about 500 times the weight, alternatively from about 0.5 to about 100 times the weight, alternatively from about 0.5 to about 50 times the weight, alternatively from about 0.5 to about 25 times the weight, alternatively from about 0.5 to about 15 times the weight. The resultant azeotropic solution in contact zone 910 may then be passed from contact zone 910 via line 911 and fed into a separator 900 for the azeotropic distillation of the solution to remove the water. Operating temperature of separator 900 will depend on the azeotropic solvent used and the pressure maintained on the separator. The water may be removed from separator 900 through overhead line 912, optionally the aromatic compound recovered, and the remaining abated components may be fed via line 913 into contact zone 915 for contacting with the remaining catalyst components including the composition comprising the metal alkyl, the nitrogen-containing compound (e.g. the pyrrole-containing compound or any other nitrogen-containing compound described herein), the halide-containing compound (e.g., a metal halide or non-metal halide), the catalyst solvent, any remaining non-halide metal alkyl, metal alkyl halide, or combinations thereof, which may be placed into contact zone 915 via various input lines not shown in Fig. 3A. A catalyst product may then be withdrawn from contact zone 915 via line 970 and optionally filtered (not shown). Alternatively the water abated material comprising the chromium-containing compound may be stored prior to contact with the remaining catalyst components. Optionally, the abated components from line 913 may be subjected to further water abatement as described herein, for example contact with a non-halide metal alkyl, adsorbent, or both prior to contact with the remaining catalyst components.

[0061] In an embodiment, one or more catalyst components other than (i) a metal alkyl halide, (ii) a metal halide and a metal alkyl, or (iii) a non-metal halide and a metal alkyl, for example the composition comprising the chromium-containing compound, the composition comprising the nitrogen-containing compound, the non-metal halide-containing compound, the solvent, or combinations thereof are contacted with an adsorbent to abate water. The contacting may occur prior to contacting with (i) a metal alkyl halide, (ii) a metal halide and a metal alkyl, or (iii) a non-metal halide and a metal alkyl. In some embodiments, contacting the chromium-containing compound with the nitrogen-containing compound may enhance the solubility of the chromium-containing compound in a solvent (e.g. ethylbenzene) as well as reduce the solution viscosity. Thus, the reduced viscosity and more soluble solution may enhance the suitability of the solution to water abatement by means of passing it through an adsorbent such as molecular sieves, to remove all or a portion of any water present. In an embodiment, the nitrogen-containing compound added may constitute substantially all or only a portion of the nitrogen-containing compound required to make the catalyst composition. Other known means for reducing viscosity, enhancing solubility, or both may be employed such that a catalyst component becomes suitable for contact with an adsorbent to remove water.

[0062] Adsorption as used herein refers to the separation operation in which one component of a gas or liquid mixture is selectively retained in the pores of a resinous or microcrystalline solid. A gas or liquid mixture contacts a solid (the adsorbent) and a mixture component (the adsorbate, which is typically water) adheres to the surface of the solid. In an embodiment, an adsorbent may be used to abate water by adding the adsorbent to catalyst component(s) in a vessel and mixing thoroughly for adequate contacting of the adsorbent with the catalyst component(s). The mixture may then be allowed to stand and after a period of time, the adsorbent settles to the bottom of the vessel. Separation can be completed by decanting or filtration (e.g., suction filtration). Alternatively, water may be abated by passing the catalyst component(s) through a fixed adsorption bed comprised of an adsorbent, allowing the mixture adequate contact time for the adsorbate to sufficiently adhere to the adsorbent, and then removing the abated catalyst component(s) from the adsorption bed. The adsorbent may then be replaced or regenerated for the next use. The original adsorption capacity of the saturated bed may be recovered by any suitable regeneration method, for example, thermal regeneration, regeneration by pressure swing, or regeneration by purging.

[0063] In the embodiments, any suitable adsorbent may be used. Examples of suitable adsorbents include 3-Angstrom molecular sieves, 5-Angstrom molecular sieves, 13X molecular sieves, alumina, silica, or combinations thereof. 3-Angstrom (3A) and 5-Angstrom (5A) refers to the size of the molecule the material can adsorb, for example, the 3A molecular sieve can adsorb molecules less than 3 angstrom and the 5A molecular sieve can adsorb molecules less than 5 angstrom. Molecular sieves are crystalline structures not unlike sponges on a molecular scale. They have a solid framework defining large internal cavities where molecules can be adsorbed. These cavities are interconnected by pore openings through which molecules can pass. Because of their crystalline nature, the pores and cavities are the same size, and depending on the size of the openings, they can adsorb molecules readily, slowly, or not at all, thus functioning as molecular sieves -- adsorbing molecules of certain sizes while rejecting larger ones.

[0064] In an embodiment wherein the nitrogen-containing compound is a pyrrole-containing compound as illustrated in Fig. 3B, a pyrrole-chromium mixture may be formed in contact zone 1025 by contacting a composition comprising the pyrrole-containing compound fed to contact zone 1025 via line 1020 and the composition comprising the chromium-containing compound fed to contact zone 1025 via line 1010, which may occur about instantaneously or over a first period of time. Feeding of the chromium-containing composition and the pyrrole-containing composition to contact zone 1025 may be sequential (e.g. chromium followed by pyrrole or vice-versa) or simultaneous and at constant or varying Py:Cr ratios as disclosed previously. Once the pyrrole-chromium mixture has been contacted in contact zone 1025 the pyrrole-chromium mixture from contact zone 1025 may be passed to contact zone 1000 via line 1040. The pyrrole-chromium mixture may optionally contain solvent, other catalyst components, e.g.

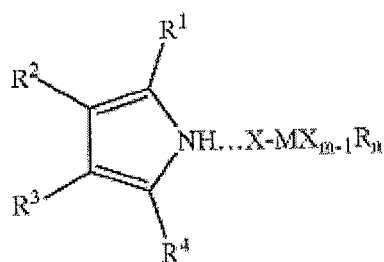
a non-metal halide, or combinations thereof, but does not comprise (i) a metal alkyl halide, (ii) a metal halide and a metal alkyl, or (iii) a non-metal halide and a metal alkyl. The pyrrole-chromium mixture is contacted with an adsorbent disposed in contact zone 1000. Contact zone 1000 may be a fixed adsorption bed as described in a previous embodiment, sized accordingly to the volumes of materials being adsorbed. The pyrrole-chromium mixture may be passed through the adsorption bed comprised of an adsorbent, e.g. 3A molecular sieve, allowing for the adsorption process to occur over a second period of time to adsorb essentially all of the free water from the pyrrole-chromium mixture. Contact with the adsorbent in contact zone 1000 may be carried out according to various known methods. One skilled in the art will understand that other nitrogen-containing compounds as described herein may be used in the embodiments shown in Figures 3B.

[0065] The water abated mixture in contact zone 1000 may then be passed from contact zone 1000 via line 1018 and contacted with the remaining catalyst components in contact zone 1015 including the composition comprising the metal alkyl, a halide-containing compound (e.g., a metal halide or non-metal halide), the solvent, any remaining non-halide metal alkyl, metal alkyl halide, or combinations thereof, which may be placed into contact zone 1015 via various input lines not shown in Fig. 3B. A catalyst product may then be withdrawn from contact zone 1015 via line 1070 and optionally filtered (not shown). Alternatively the water abated material comprising the chromium-containing compound may be stored prior to contact with the remaining catalyst components. Optionally, the water abated compounds from contact zone 1000 may be subjected to further water abatement as described herein, for example contact with a non-halide metal alkyl, azeotropic distillation, or both prior to contact with the remaining catalyst components.

[0066] Embodiments for abating water, acidic protons, or both as disclosed herein, for example the embodiments shown in Figs. 2A - 2D and 3A - 3B, may be applied alone or in combination to other processes and catalyst compositions known in the art, for example, water, acidic protons, or both may be abated from the catalyst compositions or components disclosed in reference U.S. Patent No. 6,133,495, U.S. App. No. 2002/0035029, WO 01/83447, WO 03/053890, WO 03/053891, WO 04/056477, WO 04/056478, WO 04/056479, and WO 04/056480, each of which is incorporated herein in its entirety. Likewise, embodiments for preparing catalysts, for example embodiments shown in Figs. 1A-1D and 4A-4E, may be applied alone or in combination to other processes and catalyst compositions known in the art, for example those set forth in 6,133,495, 2002/0035029, WO 01/83447, WO 03/053890, WO 03/053891, WO 04/056477, WO 04/056478, WO 04/056479, and WO 04/056480. When applying the water abatement and catalyst preparation embodiments to these catalyst compositions or components disclosed in reference 6,133,495, 2002/0035029, WO 01/83447, WO 03/053890, WO 03/053891, WO 04/056477, WO 04/056478, WO 04/056479, and WO 04/056480, the appropriate substitutions and adjustment should

be made for components that have a similar function; e.g. substitution of the heteroatomic ligands of 6,133,495, 2002/0035029, WO 01/83447, WO 03/053890, WO 03/053891, WO 04/056477, WO 04/056478, WO 04/056479, and WO 04/056480 for the nitrogen-containing compound disclosed herein and adjustments of the ligand:Cr (e.g., pyrrole:chromium) molar ratios to account for the number of equivalents of ligand(s) per mole of the ligand. Furthermore, catalyst compositions or components disclosed in reference U.S. Patent No. 6,133,495, U.S. App. No. 2002/0035029, WO 01/83447, WO 03/053890, WO 03/053891, WO 04/056477, WO 04/056478, WO 04/05477, WO 04/056478, WO 04/056479, and WO 04/056480 may be combined with other catalyst compositions or components as set forth herein to make various final catalysts according to various embodiments described herein, and water may be abated from any one or more of such compositions or components by any one or more abatement method disclosed herein.

[0067] In an embodiment, water, acidic protons, or both may be abated from the catalyst composition for producing an alpha-olefin oligomer disclosed in US Patent No. 6,133,495. In an embodiment, the nitrogen-containing compound as described herein is replaced with a pyrrole ring-containing compound as described in 6,133,495. A chromium-based catalyst is prepared by bringing a pyrrole ring-containing compound, an alkyl aluminum compound, and a halogen-containing compound into contact with each other in a hydrocarbon solvent, halogenated hydrocarbon solvent or mixture thereof, and then bringing the mixed resultant solution into contact with the chromium compound, wherein water, acidic protons, or both are abated from the catalyst or a component thereof prior to or during preparation of the catalyst. In an embodiment, the chromium-based catalyst is prepared by bringing the chromium compound, the pyrrole ring-containing compound, the alkyl aluminum compound, and the halogen-containing compound into contact with each other in a hydrocarbon solvent, halogenated hydrocarbon solvent or mixture thereof in the absence of alpha-olefin under such a condition that the concentration of the chromium compound in the resultant mixed solution is about 1×10^{-7} to 1 mol/liter, alternatively about 1×10^{-5} to 3×10^{-2} mol/liter, alternatively adjusted to not more than about 8×10^{-3} mol/liter, alternatively, not more than about 0.416 mg Cr/mL, wherein water, acidic protons, or both are abated from the catalyst or a component thereof prior to or during preparation of the catalyst. In an embodiment, water, acidic protons, or both are abated from a catalyst component comprising a pyrrole derivative represented by the general formula (I):



(I)

wherein R¹ to R⁴ are a hydrogen atom or a linear or branched hydrocarbon group having 1 to 20 carbon atoms, in which R³ and R⁴ may integrally form a ring; X is a halogen atom; M is an element selected from the group consisting of those belonging to 3-Group, 4-Group, 6-Group (exclusive of chromium), 13-Group, 14-Group and 15-Group of the Periodic Table; m and n are numbers satisfying the relationships of $1 \leq m \leq 6$, $0 \leq n \leq 5$ and $2 \leq m+n \leq 6$ with the proviso that the sum of m and n is identical to the valence of the element M; n represents the number of Rs; and R is a hydrogen atom or a linear or branched hydrocarbon group having 1 to 20 carbon atoms and when n is not less than 2, and Rs may be the same or different.

[0068] In an embodiment, water, acidic protons, or both may be abated from the catalyst composition disclosed in US Patent No. 2002/0035029. In an embodiment, the nitrogen-containing compound as described herein is replaced with a neutral multidentate ligand as described in 2002/0035029. In an embodiment, a catalyst for oligomerization of ethylene comprises:

(i) an organometallic complex having a neutral multidentate ligand having a tripod structure, represented by the following formula (1):

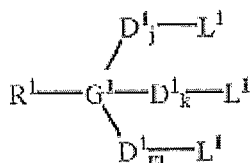


wherein A may be a neutral multidentate ligand having a tripod structure, M may be a transition metal atom of group 3 to group 10 of the periodic table, each Q may be independently selected from the group consisting of a hydrogen atom, a halogen atom, a straight chain or branched alkyl group having 1 to 10 carbon atoms which may have a substituent, an aryl group having 6 to 10 carbon atoms which may have a substituent, and n is an integer equal to a formal oxidation valence of M, and

(ii) an alkylaluminoxane;

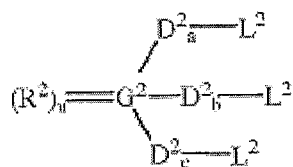
said neutral multidentate ligand A in formula (1) being a tridentate ligand represented by the following formula (2) or formula (3):

(2)



wherein j, k and m independently represent an integer of 0 to 6, each D^1 independently represents a divalent hydrocarbon group which may have a substituent, each L^1 independently represents a substituent containing an element of group 14, 15, 16 or 17 of the periodic table, with the proviso that all of the three L^1 's are not concurrently a substituent containing an element of group 14 or 17, G^1 represents a carbon or silicon atom, and R^1 represents a hydrogen atom, an alkyl group having 1 to 10 carbon atoms which may have a substituent, or an aryl group having 6 to 10 carbon atoms which may have a substituent;

(3)



wherein a, b and c independently represent an integer of 0 to 6; u represents an integer of 0 or 1; each D^2 independently represents a divalent hydrocarbon group which may have a substituent; each L^2 independently represents a substituent containing an element of group 14, 15, 16 or 17 of the periodic table, with the proviso that all of the three L^2 's are not concurrently a substituent containing an element of group 14 or 17, G^2 represents a nitrogen or phosphorus atom when u is 0, or a phosphorus atom when u is 1, and R^2 represents an oxygen or sulfur atom. Water, acidic protons, or both may be abated from the catalyst or a component thereof prior to or during preparation of the catalyst.

[0069] In an embodiment, a catalyst for the oligomerization of ethylene comprises:

- (i) an organometallic complex having a neutral multidentate ligand having a tripod structure, represented by the following formula (1):

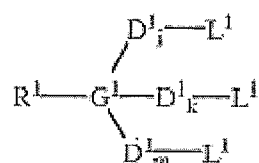


wherein A is a neutral multidentate ligand having a tripod structure, M is a transition metal atom of group 3 to group 10 of the periodic table, each Q is independently selected from the group consisting of a hydrogen atom, a halogen atom, a straight chain or branched alkyl group having 1 to 10 carbon atoms

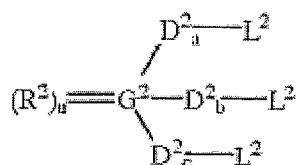
which may have a substituent, an aryl group having 6 to 10 carbon atoms which may have a substituent, and n is an integer equal to a formal oxidation valence of M, and

- (ii) an alkylaluminoxane, and
- (iii) a halogenated inorganic compound;

said neutral multidentate ligand A in formula (1) being a tridentate ligand represented by the following formula (2) or formula (3):



wherein j, k and m independently represent an integer of 0 to 6, each D^1 independently represents a divalent hydrocarbon group which may have a substituent, each L^1 independently represents a substituent containing an element of group 14, 15, 16 or 17 of the periodic table, with the proviso that all of the three L^1 's are not concurrently a substituent containing an element of group 14 or 17, G^1 represents a carbon or silicon atom, and R^1 represents a hydrogen atom, an alkyl group having 1 to 10 carbon atoms which may have a substituent, or an aryl group having 6 to 10 carbon atoms which may have a substituent;



wherein a, b and c independently represent an integer of 0 to 6; u represents an integer of 0 or 1; each D^2 independently represents a divalent hydrocarbon group which may have a substituent; each L^2 independently represents a substituent containing an element of group 14, 15, 16 or 17 of the periodic table, with the proviso that all of the three L^2 's are not concurrently a substituent containing an element of group 14 or 17, G^2 represents a nitrogen or phosphorus atom when u is 0, or a phosphorus atom when u is 1, and R^2 represents an oxygen or sulfur atom. Water, acidic protons, or both may be abated from the catalyst or a component thereof prior to or during preparation of the catalyst.

[0070] In an embodiment, a catalyst for the oligomerization of ethylene comprises:

(i) an organometallic complex having a neutral multidentate ligand having a tripod structure, represented by the following formula (1):



wherein A is a neutral multidentate ligand having a tripod structure, M is a transition metal atom of group 3 to group 10 of the periodic table, each Q is independently selected from the group consisting of a hydrogen atom, a halogen atom, a straight chain or branched alkyl group having 1 to 10 carbon atoms which may have a substituent, an aryl group having 6 to 10 carbon atoms which may have a substituent, and n is an integer equal to a formal oxidation valence of M,

(ii) an alkylaluminoxane,

(iii) a halogenated inorganic compound, and

(iv) an alkyl group-containing compound represented by the following formula (4):



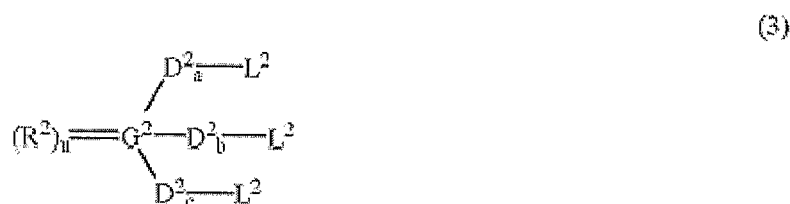
wherein p and q are numbers satisfying the formulae: $0 < p \leq 3$ and $0 \leq q < 3$, provided that (P+q) is in the range of 1 to 3, E represents an atom, other than a hydrogen atom, of group 1, 2, 3, 11, 12 or 13 of the periodic table, each R independently represents an alkyl group having 1 to 10 carbon atoms, and each J independently represents a hydrogen atom, an alkoxide group having 1 to 10 carbon atoms, an aryloxy group having 6 to 10 carbon atoms, an aryl group having 6 to 10 carbon atoms or a halogen atom;

said neutral multidentate ligand A in formula (1) being a tridentate ligand represented by the following formula (2) or formula (3):



wherein j, k and m independently represent an integer of 0 to 6, each D^I independently represents a divalent hydrocarbon group which may have a substituent, each L^I independently represents a substituent containing an element of group 14, 15, 16 or 17 of the periodic table, with the proviso that all of the three L^I s are not concurrently a substituent containing an element of group 14 or 17, G^I represents a carbon or

silicon atom, and R^1 represents a hydrogen atom, an alkyl group having 1 to 10 carbon atoms which may have a substituent, or an aryl group having 6 to 10 carbon atoms which may have a substituent;



wherein a, b and c independently represent an integer of 0 to 6; u represents an integer of 0 or 1; each D^2 independently represents a divalent hydrocarbon group which may have a substituent; each L^2 independently represents a substituent containing an element of group 14, 15, 16 or 17 of the periodic table, with the proviso that all of the three L^2 's are not concurrently a substituent containing an element of group 14 or 17, G^2 represents a nitrogen or phosphorus atom when u is 0, or a phosphorus atom when u is 1, and R^2 represents an oxygen or sulfur atom. Water, acidic protons, or both may be abated from the catalyst or a component thereof prior to or during preparation of the catalyst.

[0071] In an embodiment, a catalyst for the oligomerization of ethylene comprises:

(i) an organometallic complex having a neutral multidentate ligand having a tripod structure, represented by the following formula (1):



wherein A is a neutral multidentate ligand having a tripod structure, M is a transition metal atom of group 3 to group 10 of the periodic table, each Q is independently selected from the group consisting of a hydrogen atom, a halogen atom, a straight chain or branched alkyl group having 1 to 10 carbon atoms which may have a substituent, an aryl group having 6 to 10 carbon atoms which may have a substituent, and n is an integer equal to a formal oxidation valence of M,

(ii) an alkylaluminoxane, and

(iii) an alkyl group-containing compound represented by the following formula (4):



wherein p and q are numbers satisfying the formulae: $0 < p \leq 3$ and $0 \leq q < 3$, provided that (p+q) is in the range of 1 to 3, E represents an atom, other than a hydrogen atom, of group 1, 2, 3, 11, 12 or 13 of the periodic table, each R independently represents an alkyl group having 1 to 10 carbon atoms, and each J

independently represents a hydrogen atom, an alkoxide group having 1 to 10 carbon atoms, an aryloxy group having 6 to 10 carbon atoms, an aryl group having 6 to 10 carbon atoms or a halogen atom;

said neutral multidentate ligand A in formula (1) being a tridentate ligand represented by the following formula (2) or formula (3):



wherein j, k and m independently represent an integer of 0 to 6, each D^1 independently represents a divalent hydrocarbon group which may have a substituent, each L^1 independently represents a substituent containing an element of group 14, 15, 16 or 17 of the periodic table, with the proviso that all of the three L^1 s are not concurrently a substituent containing an element of group 14 or 17, G^1 represents a carbon or silicon atom, and R^1 represents a hydrogen atom, an alkyl group having 1 to 10 carbon atoms which may have a substituent, or an aryl group having 6 to 10 carbon atoms which may have a substituent;



wherein a, b and c independently represent an integer of 0 to 6; u represents an integer of 0 or 1; each D^2 independently represents a divalent hydrocarbon group which may have a substituent; each L^2 independently represents a substituent containing an element of group 14, 15, 16 or 17 of the periodic table, with the proviso that all of the three L^2 s are not concurrently a substituent containing an element of group 14 or 17, G^2 represents a nitrogen or phosphorus atom when u is 0, or a phosphorus atom when u is 1, and R^2 represents an oxygen or sulfur atom. Water, acidic protons, or both may be abated from the catalyst or a component thereof prior to or during preparation of the catalyst.

[0072] In an embodiment, a catalyst for oligomerization of ethylene comprises:

(i) an organometallic complex having a neutral multidentate ligand having a tripod structure, represented by the following formula (1):



wherein A is a neutral multidentate ligand having a tripod structure, M is a transition metal atom of group 3 to group 10 of the periodic table, each Q is independently selected from the group consisting of a hydrogen atom, a halogen atom, a straight chain or branched alkyl group having 1 to 10 carbon atoms which may have a substituent, an aryl group having 6 to 10 carbon atoms which may have a substituent, and n is an integer equal to a formal oxidation valence of M,

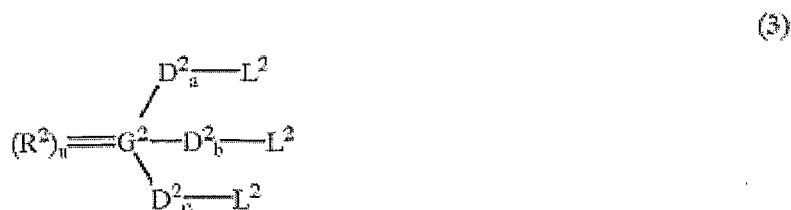
(ii) an alkylaluminoxane, and

(iii) at least one compound selected from the group consisting of an amine compound and an amide compound;

said neutral multidentate ligand A in formula (1) being a tridentate ligand represented by the following formula (2) or formula (3):



wherein j, k and m independently represent an integer of 0 to 6, each D^1 independently represents a divalent hydrocarbon group which may have a substituent, each L^1 independently represents a substituent containing an element of group 14, 15, 16 or 17 of the periodic table, with the proviso that all of the three L^1 's are not concurrently a substituent containing an element of group 14 or 17, G^1 represents a carbon or silicon atom, and R^1 represents a hydrogen atom, an alkyl group having 1 to 10 carbon atoms which may have a substituent, or an aryl group having 6 to 10 carbon atoms which may have a substituent;



wherein a, b and c independently represent an integer of 0 to 6; u represents an integer of 0 or 1; each D^2 independently represents a divalent hydrocarbon group which may have a substituent; each L^2 independently represents a substituent containing an element of group 14, 15, 16 or 17 of the periodic table, with the proviso that all of the three L^2 's are not concurrently a substituent containing an element of group 14 or 17, G^2 represents a nitrogen or phosphorus atom when u is 0, or a phosphorus atom when u is

1, and R^2 represents an oxygen or sulfur atom. Water, acidic protons, or both may be abated from the catalyst or a component thereof prior to or during preparation of the catalyst.

[0073] In an embodiment, a catalyst for the oligomerization of ethylene comprises:

(i) an organometallic complex having a neutral multidentate ligand having a tripod structure, represented by the following formula (1):

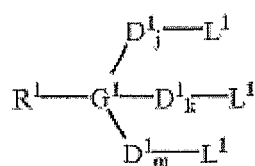


wherein A is a neutral multidentate ligand having a tripod structure, M is a transition metal atom of group 3 to group 10 of the periodic table, each Q is independently selected from the group consisting of a hydrogen atom, a halogen atom, a straight chain or branched alkyl group having 1 to 10 carbon atoms which may have a substituent, an aryl group having 6 to 10 carbon atoms which may have a substituent, and n is an integer equal to a formal oxidation valence of M,

- (ii) an alkylaluminoxane,
- (iii) at least one compound selected from the group consisting of an amine compound and an amide compound, and
- (iv) an alkyl group-containing compound represented by the following formula (4):

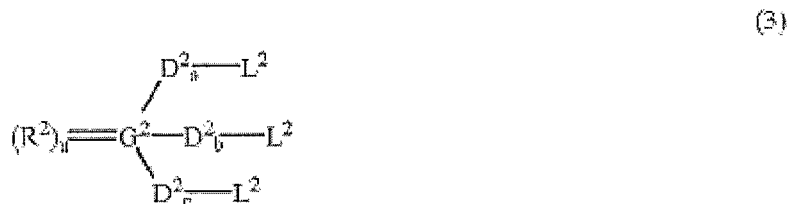


wherein p and q are numbers satisfying the formulae: $0 < p \leq 3$ and $0 \leq q < 3$, provided that (P+q) is in the range of 1 to 3, E represents an atom, other than a hydrogen atom, of group 1, 2, 3, 11, 12 or 13 of the periodic table, each R independently represents an alkyl group having 1 to 10 carbon atoms, and each J independently represents a hydrogen atom, an alkoxide group having 1 to 10 carbon atoms, an aryloxy group having 6 to 10 carbon atoms, an aryl group having 6 to 10 carbon atoms or a halogen atom; said neutral multidentate ligand A in formula (1) being a tridentate ligand represented by the following formula (2) or formula (3):



(2)

wherein j, k and m independently represent an integer of 0 to 6, each D^1 independently represents a divalent hydrocarbon group which may have a substituent, each L^1 independently represents a substituent containing an element of group 14, 15, 16 or 17 of the periodic table, with the proviso that all of the three L^1 's are not concurrently a substituent containing an element of group 14 or 17, G^1 represents a carbon or silicon atom, and R^1 represents a hydrogen atom, an alkyl group having 1 to 10 carbon atoms which may have a substituent, or an aryl group having 6 to 10 carbon atoms which may have a substituent;



wherein a, b and c independently represent an integer of 0 to 6; u represents an integer of 0 or 1; each D^2 independently represents a divalent hydrocarbon group which may have a substituent; each L^2 independently represents a substituent containing an element of group 14, 15, 16 or 17 of the periodic table, with the proviso that all of the three L^2 's are not concurrently a substituent containing an element of group 14 or 17, G^2 represents a nitrogen or phosphorus atom when u is 0, or a phosphorus atom when u is 1, and R^2 represents an oxygen or sulfur atom. Water, acidic protons, or both may be abated from the catalyst or a component thereof prior to or during preparation of the catalyst.

[0074] In an embodiment, an olefin oligomerization catalyst system incorporates a halogen source into a pyrrole ligand as disclosed in WO 01/83447, and water, acidic protons, or both may be abated from the catalyst system or a component thereof prior to or during preparation of the catalyst. In an embodiment, the nitrogen-containing compound as described herein is replaced with a halopyrrole ligand as described in WO 01/83447. In an embodiment, water, acidic protons, or both are abated from a catalyst component comprising a halopyrrole ligand. The catalyst system may comprise a chromium source, a metal alkyl, and the halopyrrole ligand and may be utilized for oligomerizing ethylene to an olefin composition comprising 1-hexene or 1-octene.

[0075] In an embodiment, an olefin oligomerization catalyst system incorporates a mixed heteroatomic ligand with at least three heteroatoms, of which at least one heteroatom is sulfur and at least 2 heteroatoms are not the same, as disclosed in WO 03/053890 and water, acidic protons, or both may be abated from the catalyst system or a component thereof prior to or during preparation of the catalyst. In an embodiment, the nitrogen-containing compound as described herein is replaced with a heteroatomic ligand as described in WO 03/053890. In an embodiment, water, acidic protons, or both are abated from the catalyst system or a catalyst component comprising a multidentate mixed heteroatomic ligand, which

includes at least three heteroatoms of which at least one is a sulfur atom. The catalyst system may comprise a chromium source, a metal alkyl, an aluminoxane, and the multidentate mixed heteroatomic ligand and may be utilized for oligomerizing ethylene to an olefin composition comprising 1-hexene or 1-octene.

[0076] In an embodiment, water, acidic protons, or both may be abated from the ligand and the ligand may be comprised of the following ligand types:

(a) $R^1A(R^2BR^3)(R^4CR^5)$ wherein R^1 , R^3 and R^5 may be hydrogen or independently be selected from the groups consisting of alkyl, aryl, aryloxy, halogen, nitro, alkoxycarbonyl, carbonyloxy, alkoxy, aminocarbonyl, carbonylamino, dialkylamino, or derivatives thereof, or aryl substituted with any of these substituents; R^2 and R^4 may be the same or different and are C_1 to about C_{15} hydrocarbyls; A is nitrogen or phosphorous; and B and C are sulfur; and

(b) $R^1A(R^2BR^3R^4)(R^5CR^6)$ wherein R^1 , R^3 , R^4 , and R^6 may be hydrogen or independently be selected from the groups consisting of alkyl, aryl, aryloxy, halogen, nitro, alkoxycarbonyl, carbonyloxy, alkoxy, aminocarbonyl, carbonylamino, dialkylamino, or derivatives thereof, or aryl substituted with any of these substituents; R^2 and R^5 may be the same or different and are C_1 to about C_{15} hydrocarbyls; A and B are individually nitrogen or phosphorous; and C is sulfur; and

(c) $A(R^1BR^2R^3)(R^4CR^5)$ wherein R^2 , R^3 , and R^5 may be hydrogen or independently be selected from the groups consisting of alkyl, aryl, aryloxy, halogen, nitro, alkoxycarbonyl, carbonyloxy, alkoxy, aminocarbonyl, carbonylamino, dialkylamino, or derivatives thereof, or aryl substituted with any of these substituents; R^1 and R^4 may be the same or different and are C_1 to about C_{15} hydrocarbyls; B is nitrogen or phosphorous; and A and C are sulfur; and

(d) $A(R^1BR^2R^3)(R^4CR^5R^6)$ wherein R^2 , R^3 , R^5 , and R^6 may be hydrogen or independently be selected from the groups consisting of alkyl, aryl, aryloxy, halogen, nitro, alkoxycarbonyl, carbonyloxy, alkoxy, aminocarbonyl, carbonylamino, dialkylamino, or derivatives thereof, or aryl substituted with any of these substituents; R^1 and R^4 may be the same or different and are C_1 to about C_{15} hydrocarbyls; B and C are individually nitrogen or phosphorous; and A is sulfur.

[0077] In an embodiment the ligand may comprise bis(2-ethylsulfanyl-ethyl)-amine, bis(2-methylsulfanyl-ethyl)-amine, bis(2-butylsulfanyl-ethyl)-amine, bis(2-decylsulfanyl-ethyl)-amine, bis(2-butylsulfanyl-ethyl)-amine, bis(2-decylsulfanyl-ethyl)-amine, bis(ethylsulfanylmethyl)-amine, bis(2-ethylsulfanyl-phenyl)-amine, bis(2-ethylsulfanyl-ethyl)phosphine, bis(2-ethylsulfanyl-ethyl)-ethylphosphine, bis(2-ethylsulfanylethyl)-phenylphosphine, N-methylbis(2-ethylsulfanyl-ethyl)-amine, (2ethylsulfanyl-ethyl)(3-ethylsulfanyl-propyl)-amine, (2-ethylsulfanyl-ethyl)(2diethylphosphino-ethyl)-

amine, (2-ethylsulfanyl-ethyl)(2-diethylphosphinoethyl)-sulfide, (2-ethylsulfanyl-ethyl)(2-diethylamino-ethyl)-amine and (ethylsulfanyl-ethyl)(2-diethylamino-ethyl)-sulfide, (2-ethylsulfanyl-ethyl)(2-diethylphosphino-ethyl)-phosphine, (2-ethylsulfanyl-ethyl)(2-diethylaminoethyl)-ethylphosphine, bis-(2-diethylphosphino-ethyl)-sulfide, bis-(2-diethylamino-ethyl)-sulfide, (2-diethylphosphino-ethyl)(2-diethylamino-ethyl)sulfide and derivatives thereof, wherein water, acidic protons, or both may be abated from the ligand.

[0078] In an embodiment, an olefin oligomerization catalyst system incorporates a mixed heteroatomic ligand with at least three heteroatoms, of which at least heteroatom is nitrogen and at least two heteroatoms are not the same, as disclosed in WO 03/053891, and water, acidic protons, or both may be abated from the catalyst system or a component thereof prior to or during preparation of the catalyst. In an embodiment, the nitrogen-containing compound as described herein is replaced with a mixed heteroatomic ligand as described in WO 03/053891. In an embodiment, the ligand may be a multidentate mixed heteroatomic ligand for an oligomerization of olefins catalyst, which ligand includes at least three heteroatoms. At least one heteroatom may be nitrogen and at least two heteroatoms may not be the same. The ligand may contain, in addition to nitrogen, at least one phosphorous heteroatom. In an embodiment, the ligand may be selected such that none of the non-carbon based heteroatoms are directly bonded to any of the other non-carbon based heteroatoms. In an embodiment, the ligand may not include a sulfur heteroatom. In an embodiment, water, acidic protons, or both may be abated from a ligand having the structure $R^1A(R^2BR^3R^4)(R^5CR^6R^7)$ wherein R^1 , R^3 , R^4 , R^6 and R^7 may be hydrogen or independently be selected from the groups consisting of alkyl, aryl, aryloxy, halogen, nitro, alkoxycarbonyl, carbonyloxy, alkoxy, aminocarbonyl, carbonylamino, dialkylamino, or derivatives thereof, or aryl substituted with any of these substituents; R^2 and R^5 are the same or different and are C_1 to about C_{15} hydrocarbyls; and at least A, B or C is nitrogen with the remainder of A, B and C being individually nitrogen or phosphorous.

[0079] In an embodiment the ligand may comprise bis-(2-diethylphosphino-ethyl)-amine, bis-(diethylphosphino-methyl)-amine, bis-(2-diethylphosphino-phenyl)-amine, N-methylbis-(2-diethylphosphino-ethyl)-amine, bis-(2-diphenylphosphino-ethyl)-amine, (2-diethylphosphino-ethyl)(3-diethylphosphino-propyl)-amine, bis-(2-dicyclohexylphosphino-ethyl)-amine, N-benzylbis-(2-diethylphosphino-ethyl)-amine, N-methyl-(2-diethylphosphino-ethyl)(3-diethylphosphino-propyl)-amine, (2-diethylphosphinoethyl)(2-diethylamino-ethyl)-amine, N-methyl-(2-diethylphosphino-ethyl)(2-diethylamino-ethyl)-amine and bis-(2-diethylamino-ethyl)ethylphosphine. A suitable multidentate mixed heteroatomic ligand is bis-(2-diethylphosphino-ethyl)-amine and derivatives thereof, wherein water, acidic protons, or both may be abated from the ligand.

[0080] In an embodiment, an olefin oligomerization catalyst system incorporates a heteroatomic ligand, as disclosed in WO 04/056477, WO 04/056478, WO 04/056479 or WO 04/056480, and water,

acidic protons, or both may be abated from the catalyst system or a component thereof prior to or during preparation of the catalyst. In an embodiment, the nitrogen-containing compound as described herein is replaced with a heteroatomic ligand as described in WO 04/056477, WO 04/056478, WO 04/056479 or WO 04/056480. In an embodiment, water, acidic protons or both may be abated from any heteroatomic ligand described in WO 04/056477, WO 04/056478, WO 04/056479 or WO 04/056480.

[0081] In an aspect, the heteroatomic ligand can be described by the formula $(R)_nA-B-C(R)_m$ wherein A and C are independently selected from a group which comprises phosphorus, arsenic, antimony, oxygen, bismuth, sulfur, selenium, and nitrogen, and B is a linking group between A and C, and R is independently selected from any homo or hetero hydrocarbyl group and n and m is determined by the respective valence and oxidation state of A and/or C. In embodiments, A and/or C may be a potential electron donor for coordination with the transition metal. An electron donor is defined as that entity that donates electrons used in chemical, including dative covalent, bond, formation. In some embodiments, at least one R group is substituted with a polar substituent.

[0082] In an aspect, the heteroatomic ligand can be described by the formula $(R^1)(R^2)A-B-C(R^3)(R^4)$ where A and C are independently selected from a group which comprises phosphorus, arsenic, antimony, bismuth and nitrogen and B is a linking group between A and C. In embodiments, R^1 , R^2 , R^3 and R^4 can independently be hydrocarbyl, substituted hydrocarbyl, heterohydrocarbyl or substituted heterohydrocarbyl groups. In some embodiments, R^1 , R^2 , R^3 and R^4 can be non-aromatic and aromatic, including heteroaromatic. In other embodiments, R^1 , R^2 , R^3 and R^4 can independently be substituted aromatic or substituted heteroaromatic groups. In further embodiments, R^1 , R^2 , R^3 and R^4 can independently be linked to one or more of each other or to the linking group B to form a cyclic structure together with A and C, A and B or B and C. In an aspect, the substituents on R^1 , R^2 , R^3 and R^4 can be polar; alternatively, non-polar; alternatively, electron donating; or alternatively, not electron donating. In embodiments, a non-electron donating substituents can be non-polar. In embodiments, a polar substituent can be electron donating.

[0083] IUPAC defines non-polar as an entity without a permanent electric dipole moment. Suitable non-polar substituents may be a methyl, ethyl, propyl, butyl, isopropyl, isobutyl, tert-butyl, pentyl, hexyl, cyclopentyl, 2-methylcyclohexyl, cyclohexyl, cyclopentadienyl, phenyl, bi-phenyl, naphthyl, tolyl, xylyl, mesityl, ethenyl, propenyl and benzyl group, or the like. Polar is defined by IUPAC as an entity with a permanent electron dipole moment. Any polar substituents on R^1 , R^2 , R^3 , and R^4 may be electron donating. Examples of polar substituents include without limitation methoxy, ethoxy, isopropoxy, C_3 - C_{20} alkoxy, phenoxy, pentafluorophenoxy, trimethylsiloxy, dimethylamino, methylsulfanyl, tosyl, methoxymethyl, methylthiomethyl, 1,3-oxazolyl, methoxymethoxy, hydroxyl, amino, phosphino, arsino, stibino, sulfate, nitro and the like.

[0084] In embodiments, A and/or C may be independently oxidized by S, Se, N or O where the valence of A and/or C allows for such oxidation. In other embodiments, A and C may be independently phosphorus or phosphorus oxidised by S or Se or N or O.

[0085] In an aspect, R^1 , R^2 , R^3 and R^4 can independently be hydrocarbyl, substituted hydrocarbyl, heterohydrocarbyl or substituted heterohydrocarbyl groups where any substituents are non-electron donating; alternatively, substituted aromatic or substituted heteroaromatic groups containing non-electron donating substituents on the atom adjacent to the atom bound to A or C; or alternatively, substituted aromatic or substituted hetero-aromatic groups containing non-polar substituents on the atom adjacent to the atom bound to A or C. In some embodiments, two or more of R^1 , R^2 , R^3 and R^4 can be aromatic or heteroaromatic containing at least one non-electronic donating substituent on the atom adjacent to the atom bound to A or C. In other embodiments, R^1 , R^2 , R^3 and R^4 can be aromatic or heteroaromatic containing at least one non-polar substituent on the atom adjacent to the atom bound to A or C. In other embodiments, R^1 , R^2 , R^3 and R^4 can independently be aromatic or substituted aromatic groups where the substituent on the atom adjacent to the atom bound to A or C is non-electron donating; or alternatively, aromatic or substituted aromatic groups where the substituent on the atom adjacent to the atom bound to A or C is not a polar group. In further embodiments, R^1 , R^2 , R^3 and R^4 can be aromatic or heteroaromatic and each of R^1 , R^2 , R^3 and R^4 can be substituted on at least one of the atoms adjacent to the atom bound to A or C by a non-electron donating group; or alternatively, aromatic or heteroaromatic and each of R^1 , R^2 , R^3 and R^4 can be substituted on at least one of the atoms adjacent to the atom bound to A or C by a non-polar group.

[0086] In an aspect, R^1 , R^2 , R^3 and R^4 are independently selected from non-aromatic and aromatic, including heteroaromatic, groups where at least one of R^1 , R^2 , R^3 and R^4 is substituted with a polar group. In some embodiments, up to four of R^1 , R^2 , R^3 and R^4 can have substituents on the atom adjacent to the atom bound to A or C. In embodiments when at least one of R^1 , R^2 , R^3 and R^4 is substituted with a polar group, each of R^1 , R^2 , R^3 and R^4 can be aromatic, including heteroaromatic, but not all of R^1 , R^2 , R^3 and R^4 , if the all are aromatic, are substituted on the atom adjacent to the atom bound to A or C. In some embodiments when at least one of R^1 , R^2 , R^3 and R^4 is substituted with a polar group, not more than two of R^1 , R^2 , R^3 and R^4 , if they are aromatic, can have substituent on the atom adjacent to the atom bound to A or C. In other embodiments, any polar substituents on R^1 , R^2 , R^3 and R^4 if they are aromatic, may not be on the atom adjacent to the atom bound to A or C. In yet other embodiments, at least one of R^1 , R^2 , R^3 and R^4 , if aromatic, can be substituted with a polar substituent on the 2nd or further atom from the atom bound to A or C.

[0087] In an aspect, R^1 , R^2 , R^3 and R^4 are independently non-aromatic or aromatic, including heteroaromatic, groups. In embodiments, one or more of R^1 , R^2 , R^3 and R^4 may be not electron donating.

In some embodiments, R^1 , R^2 , R^3 and R^4 can independently be non aromatic or aromatic, including hetero aromatic, groups and not all the groups R^1 , R^2 , R^3 and R^4 , if aromatic, have a substituent on the atom adjacent to the atom bound to A or C. In other embodiments, each non electron donating substituent on one or more of R^1 , R^2 , R^3 and R^4 may be non-polar. In embodiments, each R^1 , R^2 , R^3 and R^4 can be independently selected from the group comprising a benzyl, phenyl, tolyl, xylyl, mesityl, biphenyl, naphthyl, anthracenyl, methoxy, ethoxy, phenoxy, tolyloxy, dimethylamino, diethylamino, methylethylamino, thiophenyl, pyridyl, thioethyl, thiophenoxy, trimethylsilyl, dimethylhydrazyl, methyl, ethyl, ethenyl, propyl, butyl, propenyl, propynyl, cyclopentyl, cyclohexyl, ferrocenyl and tetrahydrofuranyl group. Alternatively, each R^1 , R^2 , R^3 and R^4 can be independently selected from a group comprising a phenyl, tolyl, biphenyl, naphthyl, thiophenyl and ethyl group.

[0088] In some embodiments, R^1 , R^2 , R^3 and R^4 can be independently selected from a group comprising a methyl, ethyl, ethylenyl, propyl, propenyl, propynyl, butyl, cyclohexyl, 2-methylcyclohexyl, 2-ethylcyclohexyl, 2-isopropylcyclohexyl, benzyl, phenyl, tolyl, xylyl, o-methylphenyl, o-ethylphenyl, o-isopropylphenyl, o-t-butylphenyl, cumyl, mesityl, biphenyl, naphthyl, anthracenyl, methoxy, ethoxy, phenoxy, tolyloxy, dimethylamino, thiomethyl, thiophenyl, trimethylsilyl, dimethylhydrazyl group.; alternatively a benzyl, phenyl, tolyl, xylyl, mesityl, biphenyl, naphthyl, anthracenyl, methoxy, ethoxy, phenoxy, tolyloxy, dimethylamino, diethylamino, methylethylamino, thiophenyl, pyridyl, thioethyl, thiophenoxy, trimethylsilyl, dimethylhydrazyl, methyl, ethyl, ethenyl, propyl, butyl, propenyl, propynyl, cyclopentyl, cyclohexyl, ferrocenyl and tetrahydrofuranyl group; or alternatively, R^1 , R^2 , R^3 and R^4 can independently be selected from a group comprising a phenyl, tolyl, biphenyl, naphthyl, thiophenyl and ethyl group.

[0089] B may be selected from any one of a group comprising: organic linking groups comprising a hydrocarbyl, substituted hydrocarbyl, heterohydrocarbyl and a substituted heterohydrocarbyl; inorganic linking groups comprising single atom links; ionic links; and a group comprising methylene, dimethylmethylene, 1,2-ethane, 1,2-phenylene, 1,2-propane, 1,2-catechol, 1,2-dimethylhydrazine, $-B(R^5)-$, $-Si(R^5)_2-$, $-P(R^5)-$ and $-N(R^5)-$ where R^5 is hydrogen, a hydrocarbyl or substituted hydrocarbyl, a substituted heteroatom or a halogen. Alternatively, B may be $-N(R^5)-$ and R^5 is a hydrocarbyl or a substituted hydrocarbyl group.

[0090] R^5 may be hydrogen or may be selected from the groups consisting of alkyl, substituted alkyl, aryl, substituted aryl, aryloxy, substituted aryloxy, halogen, nitro, alkoxycarbonyl, carbonyloxy, alkoxy, aminocarbonyl, carbonylamino, dialkylamino, silyl groups or derivatives thereof, and aryl substituted with any of these substituents. Alternatively, R^5 may be an isopropyl, a 1-cyclohexyl-ethyl, a 2-methylcyclohexyl or a 2-octyl group. B may be selected to be a single atom spacer. A single atom linking spacer is defined as a substituted or non-substituted atom that is bound directly to A and C. A and/or C

may be independently oxidized by S, Se, N or O. A and C may be independently phosphorus or phosphorus oxidized by S or Se or N or O.

[0091] In another embodiment, any of the groups in the ligand R^1 , R^2 , R^3 , R^4 or R^5 may include any cyclic heteroatomic group such as cyclopentadienyl-dimethylsilyl-t-butyl group or a cyclic homoatomic group such as cyclopentadienyl, indenyl or fluorene group.

[0092] The ligand may also contain multiple $(R)_nA-B-C(R)_m$ units. Not limiting examples of such ligands include dendrimeric ligands as well as ligands where the individual units are coupled either *via* one or more of the R groups or *via* the linking group B. In non-limiting embodiments, the dendrimeric ligands can include 1,2-di-(N(P(*o*-ethylphenyl)₂)₂)-benzene, 1,4-di-(N(P(*o*-ethylphenyl)₂)₂)-benzene, N(CH₂CH₂N(P(*o*-ethylphenyl)₂)₂)₃ and 1,4-di-(P(*o*-ethylphenyl)-N(methyl)P(*o*-ethylphenyl)₂)-benzene; alternatively, 1,2-di-(N(P(4-methoxyphenyl)₂)₂)-benzene, 1,4-di-(N(P(4-methoxyphenyl)₂)₂)-benzene, N(CH₂CH₂N(P(4-methoxyphenyl)₂)₂)₃ and 1,4-di-(P(4-methoxyphenyl)-N(methyl)P(4-methoxyphenyl)₂)-benzene, 1,2-di-(N(P(phenyl)₂)₂)-benzene, 1,4-di-(N(P(phenyl)₂)₂)-benzene, N(CH₂CH₂N(P(phenyl)₂)₂)₃ and 1,4-di-(P(phenyl)-N(methyl)P(phenyl)₂)-benzene; alternatively, 1,2-di-(N(P(4-methoxyphenyl)₂)₂)-benzene, 1,4-di-(N(P(4-methoxyphenyl)₂)₂)-benzene, N(CH₂CH₂N(P(4-methoxyphenyl)₂)₂)₃ and 1,4-di-(P(4-methoxyphenyl)-N(methyl)P(4-methoxyphenyl)₂)-benzene; or alternatively, 1,2-di-(N(P(phenyl)₂)₂)-benzene, 1,4-di-(N(P(phenyl)₂)₂)-benzene, N(CH₂CH₂N(P(phenyl)₂)₂)₃ and 1,4-di-(P(phenyl)-N(methyl)P(phenyl)₂)-benzene.

[0093] The ligands can be prepared using procedures known to one skilled in the art and procedures disclosed in published literature.

[0094] In non-limiting embodiments, the ligands can include (*o*-ethylphenyl)₂PN(methyl)P(*o*-ethylphenyl)₂, (*o*-isopropylphenyl)₂PN(methyl)P(*o*-isopropylphenyl)₂, (*o*-methylphenyl)₂PN(methyl)P(*o*-methylphenyl)₂, (*o*-ethylphenyl)₂PN(methyl)P(*o*-ethylphenyl)(phenyl), (*o*-ethylphenyl)₂PN(isopropyl)P(*o*-ethylphenyl)₂, (*o*-isopropyl)₂PN(isopropyl)P(*o*-isopropyl)₂, (*o*-methyl)₂PN(isopropyl)P(*o*-methyl)₂, (*o*-*t*-butylphenyl)₂PN(methyl)P(*o*-*t*-butylphenyl)₂, (*o*-*t*-butylphenyl)₂PN(isopropyl)P(*o*-*t*-butylphenyl)₂, (*o*-ethylphenyl)₂PN(pentyl)P(*o*-ethylphenyl)₂, (*o*-ethylphenyl)₂PN(phenyl)P(*o*-ethylphenyl)₂, (*o*-ethylphenyl)₂PN(*p*-methoxyphenyl)P(*o*-ethylphenyl)₂, (*o*-ethylphenyl)₂PN(benzyl)P(*o*-ethylphenyl)₂, (*o*-ethylphenyl)₂PN(1-cyclohexylethyl)P(*o*-ethylphenyl)₂, (*o*-ethylphenyl)₂PN(2-methylcyclohexyl)P(*o*-ethylphenyl)₂, (*o*-ethylphenyl)₂PN(cyclohexyl)P(*o*-ethylphenyl)₂, (*o*-ethylphenyl)₂PN(allyl)P(*o*-ethylphenyl)₂, (3-ethyl-2-thiophenyl)₂PN(methyl)P(3-ethyl-2-thiophenyl)₂, (2-ethyl-3-thiophenyl)₂PN(methyl)P(2-ethyl-3-thiophenyl)₂ and (2-ethyl-4-pyridyl)₂PN(methyl)P(2-ethyl-4-pyridyl)₂.

[0095] In other non-limiting embodiments, the ligands can include (3-methoxyphenyl)₂PN(methyl)P(3-methoxyphenyl)₂, (4-methoxyphenyl)₂PN(methyl)P(4-methoxyphenyl)₂, (3-methoxyphenyl)₂PN(isopropyl)P(3-methoxyphenyl)₂, (4-methoxyphenyl)₂PN(isopropyl)P(4-methoxyphenyl)₂, (4-methoxyphenyl)₂PN(2-ethylhexyl)P(4-methoxyphenyl)₂, (3-methoxyphenyl)(phenyl)PN(methyl)P(phenyl)₂ and (4-methoxyphenyl)(phenyl)PN(methyl)P(phenyl)₂, (3-methoxyphenyl)(phenyl)PN(methyl)P(3-methoxyphenyl)(phenyl), (4-methoxyphenyl)(phenyl)PN(methyl)P(4-methoxyphenyl)(phenyl), (3-methoxyphenyl)₂PN(methyl)P(phenyl)₂ and (4-methoxyphenyl)₂PN(methyl)P(phenyl)₂, (4-methoxyphenyl)₂PN(1-cyclohexylethyl)P(4-methoxyphenyl)₂, (4-methoxyphenyl)₂PN(2-methylcyclohexyl)P(4-methoxyphenyl)₂, (4-methoxyphenyl)₂PN(decyl)P(4-methoxyphenyl)₂, (4-methoxyphenyl)₂PN(pentyl)P(4-methoxyphenyl)₂, (4-methoxyphenyl)₂PN(benzyl)P(4-methoxyphenyl)₂, (4-methoxyphenyl)₂PN(phenyl)P(4-methoxyphenyl)₂, (4-fluorophenyl)₂PN(methyl)P(4-fluorophenyl)₂, (2-fluorophenyl)₂PN(methyl)P(2-fluorophenyl)₂, (4-dimethylaminophenyl)₂PN(methyl)P(4-dimethylamino-phenyl)₂, (4-methoxyphenyl)₂PN(allyl)P(4-methoxyphenyl)₂, (phenyl)₂PN(isopropyl)P(2-methoxyphenyl)₂, (4-(4-methoxyphenyl)-phenyl)₂PN(isopropyl)P(4-(4-methoxyphenyl)-phenyl)₂ and (4-methoxyphenyl)(phenyl)PN(isopropyl)P(phenyl)₂.

[0096] In yet other non-limiting embodiments, the ligands can include (phenyl)₂PN(methyl)P(phenyl)₂, (phenyl)₂PN(pentyl)P(phenyl)₂, (phenyl)₂PN(phenyl)P(phenyl)₂, (phenyl)₂PN(*p*-methoxyphenyl)P(phenyl)₂, (phenyl)₂PN(*p*-*t*-butylphenyl)P(phenyl)₂, (phenyl)₂PN((CH₂)₃-N-morpholine)P(phenyl)₂, (phenyl)₂PN(Si(CH₃)₃)P(phenyl)₂, (((phenyl)₂P)₂NCH₂CH₂N, (ethyl)₂PN(methyl)P(ethyl)₂, (ethyl)₂PN(isopropyl)P(phenyl)₂, (ethyl)(phenyl)PN(methyl)P(ethyl)(phenyl), (ethyl)(phenyl)PN(isopropyl)P(phenyl)₂, (phenyl)₂P(=Se)N(isopropyl)P(phenyl)₂, (phenyl)₂P-CH₂CH₂P(phenyl)₂, (*o*-ethylphenyl)(phenyl)PN(isopropyl)P(phenyl)₂, (*o*-methylphenyl)₂PN(isopropyl)P(*o*-methylphenyl)(phenyl), (phenyl)₂PN(benzyl)P(phenyl)₂, (phenyl)₂PN(1-cyclohexylethyl)P(phenyl)₂, (phenyl)₂PN[CH₂CH₂CH₂Si(OMe)₃]P(phenyl)₂, (phenyl)₂PN(cyclohexyl)P(phenyl)₂, (phenyl)₂PN(2-methylcyclohexyl)P(phenyl)₂, (phenyl)₂PN(allyl)P(phenyl)₂, (2-naphthyl)₂PN(methyl)P(2-naphthyl)₂, (*p*-biphenyl)₂PN(methyl)P(*p*-biphenyl)₂, (*p*-methylphenyl)₂PN(methyl)P(*p*-methylphenyl)₂, (2-thiophenyl)₂PN(methyl)P(2-thiophenyl)₂, (phenyl)₂PN(methyl)N(methyl)P(phenyl)₂, (*m*-methylphenyl)₂PN(methyl)P(*m*-methylphenyl)₂, (phenyl)₂PN(isopropyl)P(phenyl)₂, and (phenyl)₂P(=S)N(isopropyl)P(phenyl)₂.

[0097] In an aspect, the catalyst comprising the heteroatomic ligand described by the formula (R¹)(R²)A-B-C(R³)(R⁴) can oligomerize ethylene to olefin composition comprising 1-hexene, 1-octene, or mixtures thereof. In other embodiments, the heteroatomic ligand described by the formula (R¹)(R²)A-B-

$C(R^3)(R^4)$ can trimerize ethylene to 1-hexene; alternatively, tetramerize ethylene to 1-octene; or alternatively, trimerize and tetramerize ethylene to mixtures of 1-hexene and 1-octene.

[0098] The heteroatomic ligand can be modified to be attached to a polymer chain so that the resulting heteroatomic coordination complex of the transition metal is soluble at elevated temperatures, but becomes insoluble at 25°C. This approach would enable the recovery of the complex from the reaction mixture for reuse and has been used for other catalyst as described by D. E. Bergbreiter *et al.*, *J. Am. Chem. Soc.*, 1987, 109, 177-179. In a similar vein these transition metal complexes can also be immobilized by binding the heteroatomic ligands to silica, silica gel, polysiloxane or alumina or the like backbone as, for example, demonstrated by C. Yuanyin *et al.*, *Chinese J. React Pol*, 1992, 1(2), 152-159 for immobilizing platinum complexes.

[0099] In an embodiment using a pyrrole-containing compound, a nitrogen-containing compound may be contacted with the metal alkyl prior to contacting the metal alkyl with the chromium-containing compound, the pyrrole-containing compound, the optional halide-containing compound, the solvent, or combinations thereof, to make a catalyst for use in oligomerizing an olefin. Typically, preparation of catalyst can result in undesirable reaction products of metal alkyls, e.g., aluminum alkyls, with water impurities. Water present in the catalyst components at the time they are added to the metal alkyl compound may be a source of precipitates that can lead to polymer formation in the oligomerization reaction. Such precipitates may be abated by the addition of a nitrogen compound to the metal alkyl, thereby enhancing the solubility of the undesirable reaction products and preventing them from precipitating out, and further minimizing polymer production in the oligomerization reaction.

[0100] The nitrogen-containing compound may be comprised of amines, pyrroles, pyridines, substituted pyrroles such as indoles, di and tri nitrogen heterocycles, or combinations thereof. In an embodiment, the nitrogen-compound may be 2,5-dimethylpyrrole, which in this case the nitrogen compound can serve in two different functions: one, in the formation of the active site in the catalyst system; and two, in preventing the precipitation of the product of the water and metal alkyl reaction (as a solubility enhancer). In an embodiment, the nitrogen-containing compound is tributyl amine. In an embodiment, the final catalyst product is comprised of from about 0.01 to about 10 moles nitrogen per mole metal; alternatively the final catalyst product is comprised of from about 0.05 to about 5 moles nitrogen to mole metal; or alternatively the final catalyst product is comprised of from about 0.1 to about 0.5 moles nitrogen to mole metal.

[0101] In an embodiment for making a catalyst comprising a chromium-containing compound, a nitrogen-containing compound, a metal alkyl, an optional halide-containing compound, and optionally a solvent for use in oligomerizing an olefin, the chromium-containing compound, the nitrogen-containing

compound, and the metal alkyl may be simultaneously contacted. In an embodiment the simultaneous contact of the catalyst components occur via addition to a single contact zone. The simultaneous contacting may occur over a period of time of from about 1 minute to about 12 hours; alternatively from about 1 minute to about 6 hours; or alternatively from about 1 minute to about 3 hours. In an embodiment, the simultaneous contacting may occur over a period of less than or equal to about 120 minutes to form a catalyst product. In an embodiment, one or more of the catalyst components may be fed to the contact zone at mass flow rates of from about 0.1 Kg/hr to about 500 Kgs/hr, alternatively from about 5 Kg/hr to about 250 Kgs/hr; alternatively from about 10 Kg/hr to about 150 Kgs/hr; alternatively from about 0.1 Kg/hr to about 100 Kgs/hr; alternatively from about 0.1 Kg/hr to 50 Kgs/hr; alternatively from about 0.5 Kg/hr to 25 Kgs/hr; or alternatively from about 1.0 Kg/hr to 10 Kgs/hr. Such mass flow rates may also be employed with other embodiments described herein. In an embodiment, the simultaneous contacting is performed in a continuous process (wherein the period of time may be an extended period of time), or alternatively in a batch process. In an embodiment, the metal alkyl may be in a solution comprising a non-metal halide and a metal alkyl, a metal alkyl halide, a metal halide and a metal alkyl, or combinations thereof. In an embodiment, the halide-containing compound may also be simultaneously contacted with the chromium-containing compound, the nitrogen-containing compound, and the metal alkyl, for example by simultaneous addition to the hydrocarbon solvent.

[0102] In an embodiment as shown in Fig. 4A, the composition comprising the chromium-containing compound may be fed into contact zone 1100 via line 1110, the composition comprising the nitrogen-containing compound may be fed into contact zone 1100 via line 1120, the composition comprising the metal alkyl may be fed into contact zone 1100 via line 1115, and the composition comprising the optional halide-containing compound may be fed into contact zone 1100 via line 1180, all compositions being fed into contact zone 1100 simultaneously over a period of time. In an embodiment as shown in Fig. 4B, the composition comprising the chromium-containing compound may be fed into contact zone 1100 via line 1110, the composition comprising the nitrogen-containing compound may be fed into contact zone 1100 via line 1120, the compositions comprising the metal alkyl and the optional halide-containing compound may be pre-contacted and fed into contact zone 1100 via line 1117, the final compositions being fed into contact zone 1100 simultaneously over a period of time. In an embodiment as shown in Fig. 4C, the compositions comprising the chromium-containing compound and the nitrogen-containing compound may be pre-contacted and fed into contact zone 1100 via line 1122, the composition comprising the metal alkyl may be fed into contact zone 1100 via line 1115, and the compositions comprising the optional halide-containing compound may be fed into contact zone 1100 via line 1180, the final compositions being fed into contact zone 1100 simultaneously over a period of time. In an embodiment as shown in Fig. 4D, the compositions comprising the chromium-containing compound and

the nitrogen-containing compound may be pre-contacted and fed into contact zone 1100 via line 1122 and the compositions comprising the metal alkyl and the optional halide-containing compound may be pre-contacted and fed into contact zone 1100 via line 1117, the final compositions being fed into contact zone 1100 simultaneously over a period of time. In the embodiments shown in Figs. 4A-4D, a hydrocarbon solvent may be placed in contact zone 1100 before, after, or concurrently with addition of the various catalyst components. Contact zone 1100 may comprise a single vessel, for example a storage tank, tote, container, mixing vessel, etc. A catalyst product may be withdrawn from contact zone 1100 via line 1170 and optionally filtered (filter not shown). In the embodiments shown in Figs. 4A-4D, the addition of the composition comprising the nitrogen-containing compound (e.g. the pyrrole-containing compound or any other nitrogen containing ligand described herein) and the composition comprising the chromium-containing compound may be made in constant or varying Py:Cr ratios as disclosed previously. Additionally, the water, acidic protons, or both abatement embodiments set forth in Figs. 2A-2D and 3A-3B may be combined with the simultaneous addition embodiments of Figs. 4A-4D.

[0103] In an embodiment for making a catalyst comprising a chromium-containing compound, a nitrogen-containing compound, a metal alkyl, an optional halide-containing compound, and optionally a solvent for use in oligomerizing an olefin, the compositions comprising the chromium-containing compound, the nitrogen-containing compound, the metal alkyl, the optional halide-containing compound, or combinations thereof may be contacted with a previously prepared oligomerization catalyst composition. The previously prepared oligomerization catalyst solution may comprise the same or different chromium-containing compound, nitrogen-containing compound, metal alkyl, and optional halide-containing compound. The optional halide-containing compound may comprise a metal halide, a metal alkyl halide, or combinations thereof.

[0104] Any of the embodiments disclosed herein for making catalysts may be carried out wherein the new catalyst may be prepared in one or more contact zones comprising existing, previously prepared active catalyst. For example, in the embodiments shown in Figs. 4A-D, contact zone 1100 may be a holding tank for active catalyst to be fed to an oligomerization reactor and be comprised of previously prepared oligomerization catalyst. The various catalyst compounds in lines 1110, 1115, 1117, 1120, 1122, and 1180 may be simultaneously combined with the previously prepared oligomerization catalyst composition in contact zone 1100.

[0105] In an embodiment as shown in Fig. 4E, contact zone 1200 may be a holding tank for active catalyst to be fed to an oligomerization reactor and comprises previously made oligomerization catalyst. The chromium-containing compound in line 1210 may be combined with a hydrocarbon solvent in line 1250 forming a first solution in contact zone 1225. The nitrogen-containing compound in line 1220, the metal alkyl in line 1215, and the optional halide-containing compound in line 1280 may be combined with

the hydrocarbon solvent in line 1251 forming a second solution in contact zone 1235. The hydrocarbon solvent in line 1250 may be the same or different hydrocarbon solvent in line 1251. The first solution in contact zone 1225 and the second solution in contact zone 1235 may then be contacted (e.g., simultaneously or sequentially, including a plurality of iterative addition sequences) with the previously made oligomerization catalyst composition in contact zone 1200 via lines 1216 and 1218, respectively, to make the new catalyst composition. Optionally, a mixer may be disposed in contact zone 1200 to thoroughly mix the new and existing catalyst components. Again, the contacting of the composition comprising the nitrogen-containing compound (e.g. the pyrrole-containing compound or any other nitrogen containing ligand described herein) and the composition comprising the chromium-containing compound may be made in constant or varying Py:Cr ratios as disclosed previously. Additionally, the water, acidic protons, or both abatement embodiments set forth in Figs. 2A-2D and 3A-3B may be combined with the simultaneous addition embodiment of Fig. 4E.

[0106] Contacting of the catalyst components can be done under any conditions sufficient to thoroughly contact the components. Typically, contacting is performed in an inert atmosphere, such as, for example, nitrogen and/or argon. The reaction temperature for the disclosed methods of making a catalyst for use in oligomerizing an olefin can be any temperature. For ease of operation, ambient temperature may be employed. In order to effectuate a more efficient reaction, temperatures which maintain the reaction mixture in a liquid state are desirable. In an embodiment, reaction temperature is maintained at less than about 120 °C; alternatively less than about 100 °C; alternatively less than about 75 °C; alternatively less than about 50 °C; or alternatively less than about 25 °C when contacting the compositions comprising the chromium-containing compound, the nitrogen-containing compound, the metal alkyl, the optional halide-containing compound, or combinations thereof to make the catalyst. The preparation of the catalyst system at a low temperature may increase catalyst activity and reduce levels of undesirable co-product polymer.

[0107] The reaction pressure for the disclosed methods of making a catalyst for use in oligomerizing an olefin can be any pressure that does not adversely effect the reaction. Generally, pressures within the range of from about atmospheric pressure to about three atmospheres are acceptable. For ease of operation atmospheric pressure may be employed.

[0108] The reaction time for the disclosed methods of making a catalyst for use in oligomerizing an olefin can be any amount of time that can react substantially all reactants (i.e., catalyst components). Depending on the reactants, as well as the reaction temperature and pressure, reaction time can vary. Usually, times of less than about 1 day can be sufficient, for example from about 1 minute to about 12 hours. In an embodiment, reaction time is from about 1 minute to about 6 hours, alternatively from about

1 minute to about 3 hours. Longer times usually provide no additional benefit and shorter times may not allow sufficient time for complete reaction.

[0109] The resultant olefin oligomerization catalyst system prepared as described above in any of the embodiments can be collected and kept under a dry, inert atmosphere to maintain chemical stability and reactivity. In an embodiment, it may be desirable to contact the catalyst with the olefin within about 1000 hours of preparation of the catalyst; alternatively the catalyst may be contacted with the olefin within about 800 hours of preparation of the catalyst; alternatively the catalyst may be contacted with the olefin within about 600 hours of preparation of the catalyst; alternatively the catalyst may be contacted with the olefin within about 400 hours of preparation of the catalyst; or alternatively the catalyst may be contacted with the olefin within about 200 hours of preparation of the catalyst. In an embodiment, the olefin oligomerization catalyst comprising the chromium-containing compound, the nitrogen-containing compound, the metal alkyl, the optional halide-containing compound, and optionally the solvent may product a product (e.g., hexane) having a purity of at least 99.4 at a time within about 200 hours after preparation of the catalyst; alternatively the product may have a purity of at least about 99.3 at a time within about 400 hours after preparation of the catalyst; alternatively the product may have a purity of at least about 99.1 at a time within about 600 hours after preparation of the catalyst; alternatively the product may have a purity of at least about 98.8 at a time within about 800 hours after preparation of the catalyst; or alternatively the product may have a purity of less than about 98.8 at a time greater than about 900 hours after preparation catalyst.

[0110] The chromium-containing compound may be one or more organic or inorganic chromium compounds, with a chromium oxidation state of from about 0 to about 6. As used in this disclosure, chromium metal may be included in this definition of a chromium compound. Generally, the chromium-containing compound will have a formula of CrX_n , wherein X can be the same or different and can be any organic or inorganic radical, and n may be an integer from 0 to 6. Suitable organic radicals can have from about 1 to about 20 carbon atoms per radical, and are selected from alkyl, alkoxy, ester, ketone, amino radicals, or combinations thereof. The organic radicals can be straight-chained or branched, cyclic or acyclic, aromatic or aliphatic, and can be made of mixed aliphatic, aromatic, and/or cycloaliphatic groups. Suitable inorganic radicals include, but are not limited to halides, sulfates, oxides, or combinations thereof.

[0111] The chromium-containing compound may be a chromium(II) compound, chromium(III) compound, or combinations thereof. Suitable chromium(II) compounds include, but are not limited to, chromous fluoride, chromous chloride, chromous bromide, chromous iodide, chromium(II) bis(2-ethylhexanoate), chromium(II) acetate, chromium(II) butyrate, chromium(II) neopentanoate, chromium(II) laurate, chromium(II) stearate, chromium(II) oxalate, chromium(II) benzoate, chromium(II)

pyrrolide(s), or combinations thereof. Suitable chromium(III) compounds include, but are not limited to, chromium carboxylates, chromium naphthenates, chromium halides, chromium pyrrolides, chromium benzoates, chromium dionates, or combinations thereof. Specific chromium(III) compounds include, but are not limited to, chromium(III) trichloride tris-tetrahydrofuran complex, (benzene)-tricarbonyl chromium(III), chromium(III) 2,2,6,6-tetramethylheptanedionate, chromium(III) naphthenate, chromium(III) chloride, chromic bromide, chromic chloride, chromic fluoride, chromium(III) hexacarbonyl, chromium(III) acetylacetonate, chromium(III) pyrrolide(s), or combinations thereof.

[0112] In an embodiment, the chromium-containing compound is a chromium(III) carboxylate. Without limitation, examples of chromium(III) carboxylates include chromium(III) isooctanoate, chromium(III) tris(2-ethylhexanoate), chromium(III) oxy-2-ethylhexanoate, chromium(III) dichloroethylhexanoate, chromium(III) acetate, chromium(III) butyrate, chromium(III) neopentanoate, chromium(III) laurate, chromium(III) stearate, chromium(III) oxalate, chromium(III) benzoate, chromium(III) octanoate, chromium(III) propionate, or combinations thereof. In some embodiments, the chromium-containing compound can be chromium(III) tris(2-ethylhexanoate).

[0113] A typical chromium carboxylate preparation may contain a mixture of compounds, referred to generally as the chromium carboxylate preparation mixture and referred to more specifically as the Cr-mix when the chromium carboxylate is chromium(III) 2-ethylhexanoate. Compounds in the preparation mixture may include the desired chromium carboxylate, denoted hereafter as $\text{Cr}(\text{COOR})_x$, a hydrated chromium species, denoted hereafter as $\text{Cr}(\text{H}_2\text{O})_x$, chromium oligomers, denoted hereafter as Cr_x , free water and free acid. In an embodiment, the weight percent chromium in the preparation mixture ranges from about 10.3 wt% to 12.8 wt%; alternatively from 10.4 wt% to 11.8 wt%; alternatively from 10.5 wt% to 11.2 wt% based on the total weight of the preparation mixture. In such an embodiment, the chromium content may be decreased to within the wt.% ranges disclosed. Methods for decreasing the chromium content such as ligand titration and heating are known to one of ordinary skill in the art. The wt% chromium is relative to the carboxyl group used.

[0114] The hydrated chromium species, $\text{Cr}(\text{H}_2\text{O})_x$, include any species having water complexed to the chromium atom, for example including but not limited to hydrated chromium carboxylates. The chromium oligomers, Cr_x , include any chromium containing species containing more than one chromium atom per chromium-containing species. The chromium oligomer and hydrated chromium species can have a negative impact upon catalyst performance in the oligomerization of ethylene to 1-hexene or 1-octene. Thus, the presence of the chromium oligomers and hydrated chromium species should be controlled to achieve the desired catalytic performance in the oligomerization of ethylene to 1-hexene or 1-octene.

[0115] Herein the weight percent chromium carboxylate refers to the amount of the desired $\text{Cr}(\text{COOR})_x$ based on the total amount of chromium-containing species in the mixture as indicated by equation 1:

$$\text{wt. \% Cr carboxylate} = \frac{\text{wt. Cr carboxylate}}{\sum \text{wt. Cr species}} \quad (1)$$

where $\sum$ wt. Cr species includes all chromium containing species present in the composition including chromium oligomer, hydrated chromium species, and the desired chromium carboxylate. In embodiments, the wt. % chromium carboxylate can be greater than 90 wt. %; alternatively greater than 92.5 weight percent; alternatively, greater than 95 wt. %; alternatively, greater than 97.5 weight percent; or alternatively, greater than 99.0 weight percent.

[0116] In an embodiment, the monomeric chromium content and the residual (excess) radicals are optimized. This value is designated by the ratio moles $\text{Cr} : ((\text{moles ligand} \times \text{number of coordination equivalents of the ligand}) / \text{mole of ligand}) / \text{Cr oxidation number}$. In an embodiment the ratio moles $\text{Cr} : ((\text{moles ligand} \times \text{number of coordination equivalents of the ligand}) / \text{mole of ligand}) / \text{Cr oxidation number}$ is from about 0.9:1 to about 1.1:1, alternatively from about 0.94:1 to about 1.08:1, alternatively from about 0.97:1 to about 1.05:1.

[0117] In an embodiment the chromium carboxylate is chromium(III) 2-ethylhexanoate, which may also be referred to as $\text{Cr}(\text{EH})_3$. During the manufacture of chromium(III) 2-ethylhexanoate a mixture of compounds may be obtained, which includes the desired $\text{Cr}(\text{EH})_3$, hydrated chromium species, free and coordinated 2-ethylhexanoic acid, chromium oligomers and free water. Hereafter, the chromium(III) 2-ethylhexanoate mixture will be referred to as Cr-mix.

[0118] In an embodiment, the oligomerization catalyst produced using the Cr-mix disclosed herein provides a selective catalyst with a high conversion rate. Without wishing to be limited by theory, the relationship between the components of the Cr-mix and catalyst selectivity, purity, conversion and productivity may be described by Table 1A:

Table 1A

	Selectivity	Purity	C ₂ conversion	Productivity
chromium oligomers	-	-	+	
hydrated chromium species	+		-	-
Free acid	-			-

In Table 1A, the impact of the particular species in the Cr-mix on catalyst selectivity, purity, conversion and productivity are indicated as a positive effect (improved performance) by a plus sign (+) or a negative effect (decreased performance) by a minus sign (-). In Table 1A, selectivity to 1-hexene relates to the selectivity of the catalyst in producing 1-hexene versus all olefin oligomer produced and purity relates to the production of 1-hexene versus all C₆ product produced. Table 1A indicates that the presence of chromium oligomers in the Cr-mix may have a positive impact on conversion while negatively affecting selectivity and purity. In contrast, the presence of hydrated chromium species may have a positive impact on selectivity while negatively impacting conversion and productivity. Finally, the presence of free acid in the Cr-mix may negatively impact selectivity and productivity. Without wishing to be limited by theory, a Cr-mix having reduced quantities of chromium oligomers, hydrated chromium species and free acid amounts in the ranges disclosed herein may produce an oligomerization catalyst having increased selectivity, purity, conversion and productivity when compared to a catalyst produced with a Cr-mix having those components in amounts outside of the disclosed ranges.

[0119] In an embodiment, the amount of chromium oligomers present in the chromium carboxylate preparation mixture (e.g. Cr-mix) can be less than about 5 wt.%; alternatively, less than about 2.5 wt.%; alternatively, less than about 1.0 wt. %; or alternatively, less than 0.5 wt. % based on the total weight of chromium in the compound. Generally, the chromium carboxylate is soluble in methanol while the chromium oligomers are insoluble in methanol. Thus, one can determine the quantity of chromium oligomers by, placing a known quantity of chromium preparation in methanol, filtering the solution to collect the chromium oligomers, drying the chromium oligomers and calculating the weight percentage of chromium oligomers in based upon the total quantity of chromium species in the chromium carboxylate preparation. In an aspect, the chromium carboxylate preparation is a chromium(III) tris-2-ethylhexanote preparation (i.e., Cr-mix) and the chromium carboxylate is chromium(III) tris-2-ethylhexanote.

[0120] In an embodiment, the amount of hydrated chromium species present in the chromium carboxylate preparation (e.g., the Cr-mix) is less than about 5 wt.%; alternatively, less than to about 2.5 wt.%; alternatively, less than about 1 wt. %; alternatively, less than 0.5 wt. %; or alternatively, less than 0.25 weight percent. In an embodiment, the amount of hydrated chromium is determined via an acetone solubility test, wherein hydrated chromium(III) carboxylates are not soluble in acetone while non-hydrated chromium(III) carboxylates are freely soluble in acetone. Alternatively, the amount of hydrated chromium may be determined via UV-Vis, wherein there is a color difference between the hydrated and non hydrated chromium(III) carboxylates. The amount of hydrated chromium species present in the chromium carboxylate preparation may be adjusted by methods known to one of ordinary skill in the art, for example titration.

[0121] In an embodiment, the amount of free acid in the chromium carboxylate preparation (e.g., the Cr-mix) is below 50 wt.%; alternatively below 30 wt.%; alternatively below 20 wt.%; alternatively, less than 10 wt. %. The amount of free acid may be adjusted to the disclosed ranges by the acid abatement methods disclosed herein. Alternatively, the amount of free acid may be adjusted to the disclosed ranges by any method suitable for the adjustment of free acid chemically compatible with the compounds described. Such methods are known to one of ordinary skill in the art.

[0122] In an embodiment, the particulates, insoluble in hexane, in the chromium carboxylate preparation (e.g., the Cr-mix) are below 1 weight percent; alternatively below 0.5 wt.% ; alternatively below 0.2 wt.%. The amount of particles present in the Cr-mix may be adjusted by methods known to one of ordinary skill in the art, for example filtration.

[0123] In an embodiment, the water content in the chromium carboxylate preparation (e.g., the Cr-mix) is below 1 wt.%; alternatively below 0.5 wt.%; alternatively below 0.2 wt.%. Alternatively, the water content is less than or equal to 1000 ppmw. The water content in Cr-mix may be determined by any suitable analytical method for the determination of water content. Such methods for determination of water content are known to one of ordinary skill in the art and include techniques such as Karl Fischer and infrared analysis. In an embodiment, the water content in Cr-mix may be reduced further through the use of water abatement techniques such as molecular sieves or azeotropic distillation or metal alkyl addition as described previously.

[0124] In an embodiment, the nitrogen-containing ligand coordinates the chromium in the chromium-containing compound thus forming a Cr-N bond that is at least a portion of the active site of the oligomerization catalyst. Without limitation, examples of such nitrogen-containing ligands have been disclosed herein and include pyrrole and pyrrole-containing compounds, the pyrrole ring-containing compounds and pyrrole derivatives disclosed in U.S. Patent No. 6,133,495, the neutral multidentate ligands disclosed in U.S. Patent No. 2002/0035029, a pyrrole ligand as disclosed in WO 01/83447, a mixed heteroatomic ligand as disclosed in WO 03/053890 or WO 03/053891 or a heteroatomic ligand as disclosed in WO 04/056477, WO 04/056478, WO 04/056479 or WO 04/056480.

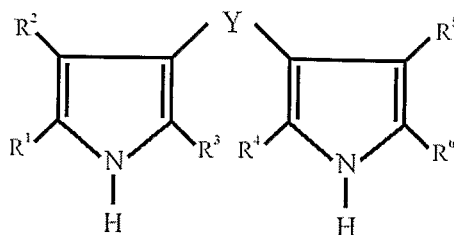
[0125] In an embodiment, the nitrogen-containing compound is a pyrrole-containing compound. The pyrrole-containing compound can be any pyrrole-containing compound that will react with a chromium salt to form a chromium pyrrolide complex. The pyrrole-containing compound includes hydrogen pyrrolide, e.g., pyrrole (C_4H_5N), derivatives of pyrrole, as well as metal pyrrolide complexes, alkali metal pyrrolides, salts of alkali metal pyrrolides, or combinations thereof. A pyrrolide (or a pyrrole) can be any compound comprising a 5-membered, nitrogen-containing heterocycle, such as pyrrole, derivatives of pyrrole, substituted pyrrole, and mixtures thereof. Broadly, the pyrrole-containing

compound can be pyrrole, any heteroleptic or homoleptic metal complex or salt containing a pyrrolide radical or ligand, or combinations thereof.

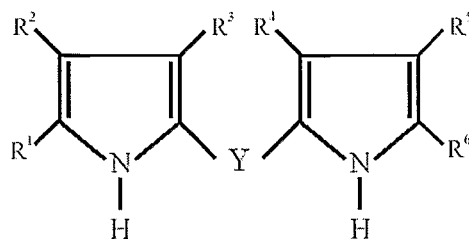
[0126] Generally, the pyrrole-containing compound will have from about 4 to about 20 carbon atoms per molecule. Pyrrolides (or pyrroles) include hydrogen pyrrolide (pyrrole), derivatives of pyrrole, substituted pyrrolides (or pyrroles), lithium pyrrolide, sodium pyrrolide, potassium pyrrolide, cesium pyrrolide, the salts of substituted pyrrolides, or combinations thereof. Examples of substituted pyrrolides (or pyrroles) include, but are not limited to, pyrrole-2-carboxylic acid, 2-acetylpyrrole, pyrrole-2-carboxaldehyde, tetrahydroindole, 2,5-dimethylpyrrole, 2,4-dimethyl-3-ethylpyrrole, 3-acetyl-2,4-dimethylpyrrole, ethyl-2,4-dimethyl-5-(ethoxycarbonyl)-3-pyrrole-propionate, ethyl-3,5-dimethyl-2-pyrrole-carboxylate

[0127] In an embodiment the pyrrole-containing compound is 2,5-dimethylpyrrole. The content of 2,5-dimethylpyrrole is greater than 98 weight percent; alternatively greater than 99.0 weight percent; alternatively greater than 99.5 weight percent. The water content of the pyrrole containing compound is below 1 weight percent; alternatively below 0.5 weight percent; alternatively below 0.01 weight percent. The color of the pyrrole containing compound (Platinum-Cobalt Number) is below 200; alternatively below 120; alternatively below 80.

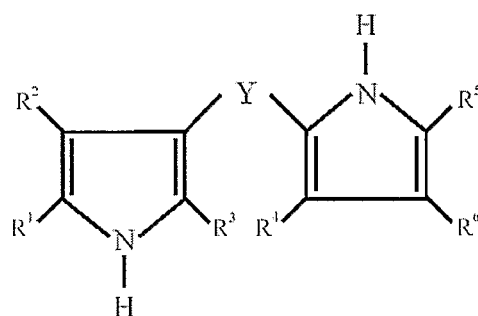
[0128] In an embodiment, the pyrrole-containing compound used in an oligomerization catalyst system comprises a dimeric pyrrole compound, for example one or more compounds represented by the following general structures:



Structure I



Structure II



Structure III

wherein, each R^1 - R^6 may independently be H, or a C_1 - C_{20} aromatic group, or any two vicinal to each other, taken together with the carbon atom to which they are bonded may form an aromatic or non-aromatic ring. Y is a structural bridge having 1 to 20 carbon atoms and may include linear, branched, or cyclic paraffinic or aromatic or contain cyclic paraffinic or aromatic structures and may include hetero atoms such as oxygen or sulfur in the form of linear, branched, or cyclic ether, silyl, sulfide, sulfone, sulfoxide functionality.

[0129] In an embodiment shown as Structure (I), R^1 , R^3 , R^4 , and R^6 are methyl group, R^2 and R^5 are hydrogens, and $Y=(CH_2)_n$ wherein $n=1-10$. In an embodiment shown as Structure (II), R^1 and R^6 are methyl groups, R^2 - R^5 are hydrogens, and $Y=(CH_2)_n$ wherein $n = 1-10$. In an embodiment shown as Structure (III), R^1 , R^3 , and R^5 are methyl groups, R^2 , R^4 , and R^6 are hydrogen, and $Y=(CH_2)_n$ wherein $n=1-10$.

[0130] Use of the dimeric pyrroles may produce a catalyst system with activity and selectivity to a desired oligomerized product, such as, for example comprising 1-hexene or 1-octene as well as low polymer production.

[0131] The metal alkyl, sometimes referred to as an activating compound, may be a heteroleptic or homoleptic metal alkyl compound of any of the metals aluminum, boron, lithium, magnesium, or zinc. The metal alkyl may be a metal alkyl halide such as DEAC; a non-halide metal alkyl such as TEA; or

combinations thereof. One or more metal alkyls can be used. The ligand(s) on the metal can be aliphatic, aromatic, or combinations thereof. For example, the ligand(s) may be any saturated or unsaturated aliphatic radical. The metal alkyl may be a compound that can be considered both a Lewis acid and a metal alkyl. As used in this disclosure, a Lewis acid may be defined as any compound that may be an electron acceptor. Activating compounds which are both a metal alkyl and a Lewis acid include alkylaluminum compounds, alkylmagnesium, alkylzinc, alkyllithium compounds, or combinations thereof. The metal alkyl can have any number of carbon atoms. However, due to commercial availability and ease of use, the metal alkyl will usually comprise less than about 70 carbon atoms per metal alkyl molecule and alternatively less than about 20 carbon atoms per molecule. In an embodiment, the metal alkyls are non-hydrolyzed, i.e., not pre-contacted with water, such as alkylaluminum compounds, derivatives of alkylaluminum compounds, halogenated alkylaluminum compounds, and mixtures thereof for improved product selectivity, as well as improved catalyst system reactivity, activity, productivity, or combinations thereof. In an embodiment the metal alkyl may be non-halide metal alkyl, a metal alkyl halide, a non-hydrolyzed alkylaluminum compound, a hydrolyzed alkylaluminum compound, or combinations thereof.

[0132] Suitable non-halide metal alkyls include, but are not limited to, alkylaluminum compounds, alkyl boron compounds, alkylmagnesium compounds, alkylzinc compounds, alkyllithium compounds, or combinations thereof. Suitable non-halide metal alkyls include, but are not limited to, n-butyllithium, s-butyllithium, t-butyllithium, diethylmagnesium, dibutylmagnesium, diethylzinc, triethylaluminum, trimethylaluminum, tripropylaluminum, tributylaluminum, triisobutylaluminum, tri-n-hexylaluminum, tri-n-octylaluminum, diethylaluminum ethoxide, diethylaluminum phenoxide, and mixtures thereof. Suitable metal alkyl halide compounds include, but are not limited to, ethylaluminum dichloride, diethylaluminum chloride, diethylaluminum bromide, diethylaluminum sesquichloride, diisobutylaluminum chloride, ethylaluminum sesquichloride, diethylaluminum bromide, diethylaluminum iodide, ethylaluminummethoxychloride, and mixtures thereof. In an embodiment, the alkylaluminum compound may be triethylaluminum.

[0133] When a oligomerization catalyst system may be the desired product, the metal alkyl may be at least one non-hydrolyzed alkylaluminum compound, expressed by the general formulae AlR_3 , AlR_2X , $AlRX_2$, AlR_2OR , $AlRXOR$, $Al_2R_3X_3$, or combinations thereof, wherein R may be an alkyl group and X may be a halogen atom. Suitable compounds include, but are not limited to, trimethylaluminum, triethylaluminum, tripropylaluminum, tri-n-butylaluminum, tri-iso-butylaluminum, tri-n-hexylaluminum, tri-n-octylaluminum, diethylaluminumchloride, diethylaluminumbromide, diethylaluminummethoxide, diethylaluminum phenoxide, ethylaluminummethoxychloride, ethylaluminum dichloride, diethylaluminum chloride, diethylaluminum bromide, ethylaluminum sesquichloride, or combinations thereof. In an

embodiment, the activating compound for an oligomerization catalyst system may be a trialkylaluminum compound, AlR_3 , for example triethylaluminum. Additionally, hydrolyzed alkylaluminum compounds, aluminoxanes, may be used. Aluminoxanes can be prepared as known in the art by reacting water or water containing materials with trialkylaluminum compounds. Suitable aluminoxanes are prepared from trialkylaluminum compounds such as trimethylaluminum, triethylaluminum, tripropylaluminum, tributylaluminum, trisobutylaluminum, trihexylaluminum or the like, and mixtures thereof. Mixtures of different aluminoxanes may also be used. Suitable hydrolyzed alkylaluminum compounds include, but are not limited to methylaluminoxane, modified methylaluminoxane, and ethylaluminoxanes, and mixtures thereof.

[0134] The olefin oligomerization catalyst systems can further comprise a catalyst support. A supported chromium catalyst system can be prepared with any support useful to support a chromium catalyst. Suitable catalyst supports include, but are not limited to, zeolites, inorganic oxides, either alone or in combination, phosphated inorganic oxides, and mixtures thereof, for example silica, silica-alumina, alumina, fluorided alumina, siled alumina, thoria, aluminophosphate, aluminum phosphate, phosphated silica, phosphated alumina, silica-titania, coprecipitated silica/titania, fluorided/siled alumina, and mixtures thereof. In an embodiment, the catalyst support is aluminophosphate.

[0135] The solvent may be a hydrocarbon solvent, a halogenated hydrocarbon solvent, or combinations thereof, usually having not more than 30 carbon atoms. Specific examples of the solvents may include aliphatic and alicyclic saturated hydrocarbons such as isobutane, pentane, n-hexane, hexanes, cyclohexane, n-heptane or n-octane, aliphatic and alicyclic unsaturated hydrocarbons such as 2-hexene, cyclohexene or cyclo-octene, aromatic hydrocarbons such as toluene, benzene or xylenes, ortho-xylene, meta-xylene, paraxylene, chlorobenzene, halogenated hydrocarbons such as carbon tetrachloride, chloroform, methylene chloride or chlorobenzene or dichlorobenzene, or the like. In an embodiment, the hydrocarbon solvent may be an aromatic or a halogenated aromatic compound having between about 6 to about 20 carbon atoms; a saturated or unsaturated hydrocarbon having from about 3 to about 14 carbon atoms; a halogenated saturated hydrocarbon having from about 1 to about 9 carbon atoms; or combinations thereof. The solvent may be a hydrocarbon such as cyclohexane, isobutane, n-hexane, hexanes, n-heptane, heptanes, pentane, or mixtures thereof. In an embodiment, the solvent is ethylbenzene. In an embodiment the solvent is tetradecene. In an embodiment, alpha-olefins may be used as the solvent, for example 1-hexene. In an embodiment, the solvent may comprise normal and/or isomeric mixtures of butene, hexene, octene, decene, dodecene, tetradecene, or combinations thereof.

[0136] In an embodiment, the hydrocarbon compound used as a solvent can be any combination of one or more aromatic or aliphatic unsaturated hydrocarbon compounds. While not wishing to be bound by theory, it may be believed that an unsaturated hydrocarbon compound acts as more than a solvent, and

can be a reactant, a stabilizing component, or both, either during, subsequent, or both, to formation of an inventive catalyst system. Suitable unsaturated hydrocarbon compounds can be any unsaturated hydrocarbon compound that can solubilize the catalyst system. In an embodiment, aromatic compounds having from about 6 to about 20 carbon atoms per molecule as a solvent there can be used in combination with any unsaturated aliphatic hydrocarbon comprising less than about 20 carbon atoms per molecule. Specific unsaturated aliphatic compounds include ethylene, 1-hexene, 1,3-butadiene, and mixtures thereof. In an embodiment, the unsaturated aliphatic hydrocarbon compound may be ethylene, which may be both a solvent and a reactant. Specific unsaturated aromatic hydrocarbon compounds include, but are not limited to, toluene, benzene, ortho-xylene, metaxylene, para-xylene, ethylbenzene, xylene, mesitylene, hexamethylbenzene, and mixtures thereof.

[0137] The optional halide-containing compound can be any compound containing a halogen, for example organohalides (including those listed as suitable solvents); non-organohalides; metal halides (including metal alkyl halides such as those previously described and non-alkyl metal halides such as tin tetrachloride and magnesium chloride); non-metal halides; or combinations thereof. Suitable compounds include, but are not limited to, compounds with a general formula of R_mX_n , wherein R can be any organic radical, inorganic radical, or both, X can be a halide, selected from fluoride, chloride, bromide, iodide, or combinations thereof, and m and n each are numbers greater than 0. Where R is an organic radical, R may have from about 1 to about 70 carbon atoms per radical, alternatively from 1 to 20 carbon atoms per radical, for best compatibility and catalyst system activity. Where R is an inorganic radical, R may be selected from aluminum, silicon, germanium, hydrogen, boron, lithium, tin, gallium, indium, lead, and mixtures thereof. In an embodiment, the halide-containing compound is a chloride-containing compound such as DEAC or organochlorides. Specific organo halides compounds include, but are not limited to, methylene chloride, chloroform, benzylchloride chlorobenzene, carbon tetrachloride, chloroethane, 1,1-dichloroethane, 1,2-dichloroethane, tetrachloroethane, hexachloroethane, 1,4-di-bromobutane, 1-bromobutane, aryl chloride, carbon tetrabromide, bromoform, bromobenzene, iodomethane, diiodomethane, hexafluorobenzene trichloro-acetone, hexachloro-acetone, hexachloro-cyclohexane, 1,3,5-trichloro-benzene, hexachloro-benzene, trityl chloride, or mixtures thereof. Specific non-alkyl metal halides include but are not limited to silicon tetrachloride, tin (II) chloride, tin (IV) chloride, germanium tetrachloride, boron trichloride, scandium chloride, yttrium chloride, lanthanum chloride, titanium tetrachloride, zirconium tetrachloride, hafnium tetrachloride, aluminum chloride, gallium chloride, silicon tetrachloride, tin tetrachloride, phosphorus trichloride, antimony trichloride, trityl-hexachloro-antimonate, antimony pentachloride, bismuth trichloride, boron tribromide, silicon tetrabromide, , aluminum fluoride, molybdenum pentachloride, tungsten hexachloride, aluminum tribromide, aluminum trichloride, or combinations thereof. Specific metal alkyl halide compounds include, diethyl aluminum chloride, ethyl aluminum sesquichloride, ethyl aluminum dichloride, mixture

of non-halide metal alkyls and metal halides, trimethyl-chlorosilane, tributyl tin chloride, dibutyl tin dichloride, or combinations thereof.

[0138] Furthermore, the chromium-containing compound, the metal alkyl, or solvent can contain and provide a halide to the reaction mixture. For example, the halide source may be an alkylaluminum halide and may be used in conjunction with alkylaluminum compounds. Suitable alkylaluminum halides include, but are not limited to, diisobutylaluminum chloride, diethylaluminum chloride, ethylaluminum sesquichloride, ethylaluminum dichloride, diethylaluminum bromide, diethylaluminum iodide, and mixtures thereof.

[0139] The amount of each reactant used to prepare an oligomerization catalyst system can be any amount sufficient that, when combined to form the catalyst system, oligomerization occurs upon contact with one or more olefins. Generally, a molar excess of the metal alkyl is used. In an embodiment in which the nitrogen-containing compound is a pyrrole, expressed as a molar ratio, in terms of moles of nitrogen (N) in the pyrrole compound to moles of metal (M) in the metal alkyl, usually less than a 1:150 molar ratio is used. In an embodiment, the metal (M) is aluminum. In an embodiment, the N:M molar ratio is from about 1:1 to about 1:50, alternatively from about 1:1 to about 1:20, or alternatively from about 1:1 to about 1:10. Generally, the amount of metal alkyl/pyrrole solution used is determined based on the moles of chromium. In an embodiment, expressed as a molar ratio, in terms of moles of chromium (Cr) to moles of nitrogen (N) in the pyrrole compound to moles of metal (M) in the metal alkyl, i.e., Cr:N:M, the ratio of the chromium containing compound to the pyrrole-containing compound is at least about 1:15 and the ratio of the chromium containing compound to metal alkyl is at least about 1:150 such that Cr:N:M is at least about 1:15:150. In an embodiment, the Cr:N:M molar ratio is within a range of about 3:3:3 (also expressed as about 1:1:1) to about 1:3:100; alternatively, the Cr:N:M molar ratio is within a range of 1:3:9 to 1:3:21. In an embodiment, to prepare an oligomerization catalyst system, about one mole of chromium, as the element chromium (Cr), can be contacted with about 1 to about 50 moles of pyrrole-containing compound and about 1 to about 75 moles of aluminum, as the element, optionally in an excess of unsaturated hydrocarbon. The halide source may be present in an amount from about 1 to about 75 moles of halide, as the element. In an embodiment, about 1 mole of chromium, calculated as the element chromium (Cr), can be contacted with about 1 to about 15 moles of pyrrole-containing compound; about 5 to about 40 moles of aluminum, calculated as the element aluminum (Al); and about 1 to about 30 moles of the halide-containing compound, calculated as elemental halide (X); in an excess of unsaturated hydrocarbon. In an embodiment, about one mole of chromium, as the element (Cr), may be contacted with two to four moles of pyrrole-containing compound; 10 to 25 moles of aluminum, as the element (Al); and 2 to 15 moles of halide, as an element (X); in an excess of unsaturated hydrocarbon.

[0140] The ratio of pyrrole to chromium (Py:Cr) in the final catalyst composition recovered as product from the various embodiments disclosed herein is referred to as the final Py:Cr molar ratio. The final Py:Cr molar ratio of the catalyst may be in a range of from about 1.0:1 to about 4.0:1; alternatively from about 1.5:1 to about 3.7:1; alternatively from about 1.5:1 to about 2.5:1; alternatively from about 2.0:1 to about 3.7:1; alternatively from about 2.5:1 to about 3.5:1; or alternatively from about 2.9:1 to about 3.1:1.

[0141] The catalyst synthesis prepared in a hydrocarbon solvent may be referred to as a catalyst system solution. The resultant catalyst system, prior to introduction to any of the reactant, may have a chromium concentration of about less than about 50 mg Cr/ml catalyst system solution, for example from about 0.005g Cr/mL catalyst system solution to about 25 mg Cr/ml catalyst system solution, alternatively from about 0.1 mg Cr/ml catalyst system solution to about 25 mg Cr/ml catalyst system solution, alternatively from about 0.5 mg Cr/ml catalyst system solution to about 15 mg Cr/ml catalyst system solution, or alternatively from about 1 mg Cr/ml catalyst system solution to about 15 mg Cr/ml catalyst system solution

[0142] Catalysts prepared in accordance with the present disclosure may be used for the oligomerization of olefins, for example, alpha-olefins. The oligomerization of olefins may be conducted by any suitable oligomerization methods. In an embodiment, an oligomerization catalyst is contacted with one or more olefins in a reaction zone under suitable reaction conditions (e.g., temperature, pressure, etc.) to oligomerize the olefins. Linear or branched alpha-olefins having 2 to 30 carbon atoms can be used as the olefins raw material. Specific examples of the alpha-olefins may include ethylene, propylene, 1-butene, 1-hexene, 1-octene, 3-methyl-1-butene, 4-methyl-1-pentene or the like. When ethylene is used as the alpha-olefin, it is possible to produce an olefin composition comprising 1-hexene as a trimer or 1-octene as a tetramer of ethylene with a high yield and a high selectivity.

[0143] In the description above, like parts are marked throughout the specification and drawings with the same reference numerals, respectively. The drawing figures are not necessarily to scale. Certain features of the invention may be shown exaggerated in scale or in somewhat schematic form and some details of conventional elements may not be shown in the interest of clarity and conciseness. The present disclosure is susceptible to embodiments of different forms. There are shown in the drawings, and herein are described in detail, specific embodiments of the present disclosure with the understanding that the present disclosure is to be considered an exemplification of the principles of the invention, and is not intended to limit the invention to that illustrated and described herein. It is to be fully recognized that the different teachings of the embodiments discussed above may be employed separately or in any suitable combination to produce desired results. Specifically, the present disclosure for a method of making a catalyst by contacting of catalyst components should not be limited by any of the various embodiments

described. Various embodiments set forth in the figures may be combined. For example, the water, acidic protons, or both abatement embodiments set forth in Figs. 2A-2D and 3A-3B may be combined with the bulk addition embodiments of Figs. 1A-1D or the simultaneous addition embodiments of Figs. 4A-4E. Additionally, various embodiments for abating water may be combined in any desired number and sequence, for example azeotropic distillation followed by contact with a non-halide metal alkyl (e.g., TEA), contact with an adsorbent, or both in any order; contact with a non-metal halide followed by contact with an adsorbent (or vice-versa); azeotropic distillation before, after, or between contact with a non-metal halide followed by contact with an adsorbent; etc. The water, acidic protons, or both abatement, bulk addition, and simultaneous addition embodiments may be integrated in any desired and operable number and sequence in other embodiments. The method disclosed herein is for making an oligomerization catalyst that may be useful in any suitable reaction such that the reaction is an oligomerization reaction. In an embodiment, the method of the present disclosure is for an oligomerization catalyst for use in a trimerization reaction producing 1-hexene from ethylene or the tetramerization reaction to produce 1-octene and the detailed description above may be focused on these embodiments but with the understanding that the present invention may have broader applications.

EXAMPLES

[0144] Preparation of an oligomerization catalyst having been generally described, the following examples are given as particular embodiments of the catalyst disclosed and to demonstrate the practice and advantages thereof. It is understood that the examples are given by way of illustration and are not intended to limit the specification or the claims to follow in any manner.

[0145] Various embodiments for preparing the oligomerization catalyst are shown in examples 1 through 14. In example 1, selective 1-hexene catalyst is prepared at various temperatures and chromium concentrations. In example 2, selective 1-hexene catalyst is prepared by simultaneous addition of chromium/ethylbenzene and TEA/DEAC/pyrrole/ethylbenzene to the heel of previously prepared catalyst. In example 3, selective 1-hexene catalyst is prepared by using a pyrrole:chromium ratio of 6:1 for the first half of the chromium/pyrrole addition and a pyrrole:chromium ratio of 0 during the second half of the chromium/pyrrole addition. In example 4, selective 1-hexene catalyst is prepared by simultaneous addition of all catalyst components. In example 5, chromium compounds containing various amounts of water and chromium oligomers are used in the preparation of the selective 1-hexene catalyst. In example 6, selective 1-hexene catalyst is prepared by separate but simultaneous addition of the pyrrole and chromium components to a solution of TEA and DEAC. In example 7, selective 1-hexene catalyst is improved when a small amount of TEA is added to the chromium component and water, acidic protons, or both are abated. In example 8, water, acidic protons, or both are abated in the preparation of the selective 1-hexene catalyst by contacting a small amount of TEA with the chromium/pyrrole solution. In

example 9, preparation of the selective 1-hexene catalyst is made by varying the pyrrole:chromium ratio during the addition to TEA/DEAC.

[0146] In example 10, preparation of the selective 1-hexene catalyst is made using high initial pyrrole:chromium contact ratios when contacted with TEA/DEAC. In example 11, preparation of the selective 1-hexene catalyst is made using simultaneous separate addition of catalyst components to the heel of previously prepared catalyst. In example 12, preparation of the selective 1-hexene catalyst is made with the addition of a nitrogen compound to the alkylaluminum compound to solubilize products resulting from the reaction of water and aluminum alkyls. In example 13, water is abated when the pyrrole and chromium components are contacted to reduce the chromium component's viscosity, facilitating water removal using molecular sieves. In example 14, water is abated by azeotropic distillation to remove the water from the chromium catalyst component. In example 15, the impact of the catalyst age on 1-hexene purity is described. Several of the above examples also include the embodiment for the addition of chromium and/or pyrrole to the alkyl aluminums.

[0147] In the examples below, catalyst was prepared using one of two apparatus set-ups. One set-up is a lab scale set-up for preparing catalyst in small quantities, for example 100 ml, which are typically used for screening purposes. The other set-up is a pilot plant scale set-up typically designed for preparing larger quantities of catalyst, for example 3.5 gallons, which would be suitable for use in a pilot plant.

[0148] The lab scale set-up prepares catalyst in a dry box in which the atmosphere inside the box is controlled with an inert gas blanket to keep it free of oxygen and moisture, which may be detrimental to the catalyst components, the prepared catalysts, or both. All lab scale catalyst preparation procedures described in the examples below are performed in glassware in a dry box. Once the catalyst is prepared it is diluted with cyclohexane to the concentration desired for oligomerization reactor tests. The diluted catalyst solution is then transferred into a 300 cc metal cylinder to provide the means for transport of the catalyst to an oligomerization reactor under protected atmosphere. Note that any transfer of components via syringes described in the examples below is done in the dry box.

[0149] The pilot plant scale set-up prepares catalyst under a nitrogen blanket to control the atmosphere, keeping it free of oxygen and moisture. All pilot plant scale catalyst preparation procedures described in the examples below are performed in a 5 gallon reactor comprising a Hastelloy® autoclave. Once the catalyst is prepared it is filtered into a 5-10 gallon metal cylinder. About 150 grams of the prepared catalyst is then transferred from the large cylinder into a smaller, 300 cc, metal cylinder and transported to an inert gas blanketed dry box as described above. The prepared catalyst is transferred into glassware and is diluted with cyclohexane to the desired concentration for testing in the oligomerization

reactor. The diluted catalyst solution is then transferred into a 300 cc metal cylinder and transported to an oligomerization reactor.

[0150] In the examples below, the prepared catalyst is tested in either a batch or a continuous oligomerization reactor. The batch oligomerization reactor is a 1 liter autoclave that is sealed and is under a nitrogen blanket. It has a magnetic stirring device to stir the contents of the sealed container. Prepared catalyst solution transported to the oligomerization reactor in the 75 cc metal cylinder. Solvent, e.g., cyclohexane, is charged to the oligomerization reactor, and the catalyst is transferred to the reactor by connecting the cylinder to the reactor and pressurizing the cylinder with ethylene, which conveys the catalyst into the reactor. The oligomerization reactor is pressurized with 650 psig of ethylene and 50 psig of hydrogen, and is operated at a temperature of about 115 °C.

[0151] In some of the examples below, a continuous oligomerization reactor is used to test the prepared catalyst. The continuous oligomerization is performed by controlling of all the feeds to the reactor by using separate controls for each feed component. Hydrogen is fed to the reactor at a rate of about 0.5 L/hr, and ethylene is fed to the reactor at a rate of about 497 g/hr. The reactor is either a 1 liter or a 1 gallon autoclave, depending on the desired residence time in the reactor. Reaction temperature is about 115 °C, and pressure is about 800 psig.

[0152] Online samples of production from the continuous oligomerization reactor were collected via liquid sampling valves (manufactured by Valco) and fed to an online gas chromatograph (GC), a Hewlett Packard 6890, for analysis. The production samples were analyzed by the GC for the amount of ethylene, hexene, and C₆ isomers and higher oligomers. From this information the selectivity, purity, and conversion was calculated. Selectivity (1-C₆=) refers to the weight percent of ethylene converted into 1-hexene. Purity (1-C₆=/ C₆) refers to the weight percent of 1-hexene in the total of all C₆ isomers. Conversion (C₂=) refers to the weight percent of ethylene has been converted to oligomer product (e.g., hexene or decenes, etc.). Productivity refers how much 1-hexene the catalyst produced, and relates to the amount of catalyst used. Productivity is quantified in units of grams of 1-hexene per gram of chromium (g 1-C₆=/g Cr). In the batch processes, productivity is evaluated over a 30 minute time frame. Other evaluations made on the product include reactor polymer (Rx Polymer) and total polymer. At the end of each day, the reactor was opened and cleaned. Any polymer inside the reactor was collected, allowed to dry, and then weighed. This amount was then extrapolated to a commercial sized processing unit of 100,000,000 pounds/year and reported as reactor polymer, quantified in pounds per hour expected in a 100,000,000 pound per year plant (Lb/Hr 100MM/yr Plant). A filter comprising a stainless steel pad placed downstream of the reactor was also removed, dried and weighed at the end of each day for amounts of polymer. This amount of polymer was then scaled up to a 100,000,000 pounds/year plant and

added to the reactor polymer amount for reporting the total polymer, quantified in pounds per hour expected in a 100,000,000 pound per year plant (Lb/Hr 100MM/yr Plant).

[0153] To determine the presence of water of hydration in some of the samples an infrared analysis was done using a standard IR apparatus. The IR band for the complexed water, e.g., about 1450 cm^{-1} , of hydration is near the band for chromium oligomers, making it difficult to distinguish the two. Therefore, in some cases a methanol solution test for precipitation of chromium oligomers was performed to help in evaluating the online samples to determine the presence of water of hydration.

Example 1

[0154] **Catalyst 1-8:** Catalyst was prepared by adding 14.1 lbs of dry, nitrogen-purged toluene to a 5 gallon reactor. To the toluene was added 630.9 g chromium(III) 2-ethylhexanoate dissolved in 750 mL toluene followed by a 300 mL toluene rinse. 2,5-Dimethylpyrrole (388.9 mL) was added to the chromium solution in the reactor. The reactor was purged with nitrogen and brought to a temperature of $25\text{ }^{\circ}\text{C}$. A mixture of 1,600 g neat triethylaluminum (TEA) and 1,229 g neat diethylaluminum chloride (DEAC) was then added to the reactor followed by 0.2 lbs of toluene rinse. The temperature increased $18\text{ }^{\circ}\text{C}$ and was returned to $25\text{ }^{\circ}\text{C}$ with cooling. The contents of the reactor stood overnight and were then filtered, using a filter comprising a combination of a metal screen, filter paper, glass wool, diatomaceous earth, and another layer of glass wool. Additional catalysts were prepared in which the temperature and chromium concentration of the catalyst preparations were varied. The catalysts were tested for productivity in a 1 gallon continuous reactor and the results are shown in Table 1B.

Table 1B

Catalyst	Temp ($^{\circ}\text{C}$)	Concentration (mg Cr/mL)	Productivity (g 1-C6=/g Cr)	Rx Polymer (Lb/Hr 100MM/yr Plant)
Ratio Cr/pyrrole/TEA/DEAC (1/3/11/8)				
1	25	1	43,183	0.001
2	75	1	40,010	0.083
3	25	5	45,769	0.005
4	75	5	44,599	0.000
Ratio Cr/pyrrole/TEA/DEAC (1/1.8/6.5/5)				
5	25	1	41,961	0.015
6	75	1	38,008	0.005
7	25	5	43,373	0.016
8	75	5	27,127	0.906

[0155] The examples show that catalyst productivity increased with a reduction in catalyst preparation temperature. Additionally, the examples show the best catalyst productivity was observed in catalyst 3 and catalyst 7 with 45,769 g 1-C6=/g Cr and 43,373g 1-C6=/g Cr, respectively, when prepared

at low temperature (25 °C) and high chromium concentration (5 mg Cr/mL). Low reactor polymer was also observed under the best productivity conditions.

Example 2

[0156] Catalyst 9-10: An ethylbenzene solution containing 2.3 g chromium(III) 2-ethylhexanoate and 8.13 g ethylbenzene was prepared. A separate solution containing 6.05 g neat triethylaluminum (TEA), 4.63 g neat diethylaluminum chloride (DEAC), 1.37 g 2,5-dimethylpyrrole and 22.6 g ethylbenzene was also prepared. These two solutions were added to 30.98 g of active catalyst over a 40 minute period such that the addition time for both solutions started and ended at the same time. The catalyst was tested in a 1 L continuous reactor and the average results of two test runs are shown in Table 2 as Catalyst 10. The average of two test runs of a standard catalyst preparation is shown in Table 2 as Catalyst 9.

Table 2

Catalyst	Selectivity (1-C ₆ =)	Purity (1-C ₆ =)	Productivity (g 1-C ₆ =/g Cr)	Rx Polymer (Lb/Hr 100MM/yr Plant)	Total Polymer
9	89.3%	98.8%	82,575	0.00	13.33
10	89.1%	98.7%	82,989	0.00	7.18

[0157] The examples show that an acceptable catalyst can be prepared. The examples further indicate that a fewer number of tanks may be required to prepare catalyst.

Example 3

[0158] Catalyst 11: A solution was prepared by mixing 12.10 g neat triethylaluminum (TEA), 9.38 g neat diethylaluminum chloride (DEAC) and 20.02 g ethylbenzene. Two aliquots were added to this solution. The first contained 2.3 g chromium(III) 2-ethylhexanoate, 1.14 g ethylbenzene and 2.74 g 2,5-dimethylpyrrole. The second contained 2.3 g chromium(III) 2-ethylhexanoate and 1.14 g ethylbenzene. Ethylbenzene was added to obtain a total volume of 100 mL. The catalyst prepared by this method was tested in a 1 L continuous reactor. The average results of three test runs are shown in Table 3.

Table 3

Catalyst	Selectivity (1-C ₆ =)	Purity (1-C ₆ =/C ₆)	Catalyst Productivity (g 1-C ₆ =/g Cr)
11	91.2%	99.2%	80,759

[0159] The example shows high selectivity (91.2%), high purity (99.2%), and good catalyst productivity (80,759 g 1-C₆=/g Cr) for the catalyst preparation.

Example 4

[0160] **Catalyst 12:** Ethylbenzene (10.67 g) was added to a dry 100 mL volumetric flask. Individual chemicals were added to each of four separate 20 mL syringes. The chemicals added were 4.76 g chromium(III) 2-ethylhexanoate dissolved in 2.38 g ethylbenzene, 12.06 g neat triethylaluminum (TEA), 9.26 g neat diethylaluminum chloride (DEAC) and 2.74 g 2,5-dimethylpyrrole. To each of these syringes was added sufficient ethylbenzene to provide a total volume of 19-20 mL. The needles of the syringes were added to the 100 mL volumetric flask and the syringes emptied into the flask simultaneously at the same rate over 30 minutes. After the additions were complete, ethylbenzene was added to the flask to obtain a total volume of 100 mL. The catalyst (1 mL) prepared by this method was tested in a 1 L batch reactor at 116 °C and 680 psig. The results of this test are shown in Table 4.

Table 4

Catalyst	Selectivity (1-C ₆ =)	Purity (1-C ₆ =/C ₆)	Catalyst Productivity (g 1-C ₆ =/g Cr)
12	92.0%	98.7%	34,325

Example 5

[0161] **Catalyst 13-15:** Catalyst was prepared by adding 15.85 g ethylbenzene to a dry 100 mL volumetric flask. To this flask was added 12.09 g neat triethylaluminum (TEA), 9.26 g neat diethylaluminum chloride (DEAC) and 2.74 g 2,5-dimethylpyrrole. To this mixture was added 4.76 g chromium(III) 2-ethylhexanoate dissolved in 2.38 g ethylbenzene. The volume was brought to 100 mL with ethylbenzene. Different preparations of chromium(III) 2-ethylhexanoate were used to prepare the catalysts 13-15. In catalyst 13 the chromium content of the chromium(III) 2-ethylhexanoate was 10.5%. Infrared analysis and a methanol solubility test indicated that some water of hydration was present but no chromium oligomers. In catalyst 14 the chromium content was 10.9% and infrared analysis and methanol solubility indicated that neither water of hydration nor chromium oligomers were present. In catalyst 15 the analysis indicated the presence of chromium oligomers. The catalysts prepared were tested for activity in the continuous reactor (1 L) and the average results for two test runs of each preparation are shown in Table 5.

Table 5

Catalyst	Selectivity (1-C ₆ =)	Purity (1-C ₆ =/C ₆)	Conversion (C ₂ =)	Catalyst Productivity (g 1-C ₆ =/g Cr)
13	90.3%	99.1%	79.0%	83,642
14	88.7%	99.1%	84.5%	87,882
15	87.4%	98.1%	86.4%	88,460

[0162] The examples show that the best combination of purity and productivity are obtained when the water of hydration and chromium oligomers are not contained in the chromium(III) 2-ethylhexanoate in significant amounts.

Example 6

[0163] **Catalyst 16:** Ethylbenzene (20.01 g) was added to a dry 125 mL Erlenmeyer flask equipped with a magnetic stirrer. To the ethylbenzene was added 12.07 g neat triethylaluminum and 9.27 g neat diethylaluminum chloride. Into a 10 mL syringe was added 4.61 g chromium(III) 2-ethylhexanoate dissolved in 2.28 g ethylbenzene. Into a separate 10 mL syringe was added 2.73 g 2,5-dimethylpyrrole and 3.38 g ethylbenzene. Both of the syringes had an approximate volume of 7.5 mL. The syringe needles were put into opposite sides of the Erlenmeyer flask containing the diluted aluminum alkyls and the contents were added simultaneously over 30 minutes. After the addition was complete, the contents were transferred to a 100 mL volumetric flask and diluted to about 103 mL with ethylbenzene. This catalyst was tested in a continuous 1 L reactor and the results (average of three test runs) are shown in Table 6.

Table 6

Catalyst	Selectivity (1-C ₆ =)	Purity (1-C ₆ =/C ₆)	Conversion (C ₂ =)	Catalyst Productivity (g 1-C ₆ =/g Cr)
16	93.0%	98.9%	66.6%	72,691

Example 7

[0164] **Catalyst 17:** Neat triethylaluminum (TEA, 0.27 g) was added to 30.01 g of ethylbenzene. This solution was added slowly to 4.62 g chromium(III) 2-ethylhexanoate dissolved in 2.27 g ethylbenzene. This is an amount of TEA sufficient to react with water and excess acid present in the chromium(III) 2-ethylhexanoate. The chromium solution, after reaction with TEA, was added, over 50 minutes, to a solution containing TEA (11.81 g), diethylaluminum chloride (DEAC, 9.27 g), 2,5-dimethylpyrrole (2.75 g) and ethylbenzene (25.01 g). Ethylbenzene was subsequently added to provide a total volume of 100 mL.

[0165] **Catalyst 18:** A comparison catalyst was prepared by adding 30.02 g of ethylbenzene to 4.62 g chromium(III) 2-ethylhexanoate dissolved in 2.27 g ethylbenzene. The chromium solution was added, over 50 minutes, to a solution containing TEA (12.08 g), diethylaluminum chloride (DEAC, 9.28 g), 2,5-dimethylpyrrole (2.74 g) and ethylbenzene (25.00 g). Ethylbenzene was subsequently added to provide a total volume of 100 mL.

[0166] These catalysts were tested for productivity in a 1 L continuous reactor. The average of two separate runs for each catalyst is shown in Table 7.

Table 7

Catalyst	Selectivity (1-C ₆ =)	Purity (1-C ₆ =/C ₆)	Conversion (C ₂ =)	Catalyst Productivity (g 1-C ₆ =/g Cr)
17	90.0%	98.8%	88.3%	93,129
18	89.1%	98.8%	82.7%	86,306

[0167] The addition of TEA to a chromium(III) 2-ethylhexanoate solution provided a catalyst with increased activity. It will also reduce corrosion in equipment after the catalyst has been inactivated. The example further provides an example of TEA addition to chromium to abate water, acidic protons, or both.

Example 8

[0168] **Catalyst 19:** Neat triethylaluminum (TEA, 0.43 g) was added to 2.01 g of ethylbenzene. This solution was added slowly to 4.62 g chromium(III) 2-ethylhexanoate in 27.27 g ethylbenzene. This is a small excess of the amount of TEA sufficient to react with water and excess acid present in the chromium(III) 2-ethylhexanoate. To this chromium/TEA solution was added 2.73 g of 2,5-dimethylpyrrole. The chromium/TEA/dimethylpyrrole solution, was added, over 30-40 minutes, to a solution containing TEA (11.62 g), diethylaluminum chloride (DEAC, 9.25 g) and ethylbenzene (15.00 g). Ethylbenzene was then added to provide a total volume of 100 mL.

[0169] **Catalyst 20:** A comparison catalyst was prepared by adding 2.74 g 2,5-dimethylpyrrole to 4.61 g chromium(III) 2-ethylhexanoate dissolved in 2.27 g ethylbenzene. An immediate reduction in the viscosity of the chromium solution was observed. This chromium solution was added, over 30-40 minutes, to a solution containing TEA (12.08 g), diethylaluminum chloride (DEAC, 9.27 g) and ethylbenzene (20.00 g). Ethylbenzene was then added to provide a total volume of 100 mL.

[0170] These catalyst preparations were tested for productivity in a 1 L continuous reactor. The average of three separate test runs for each catalyst is shown in Table 8.

Table 8

Catalyst	Selectivity (1-C ₆ =)	Purity (1-C ₆ =/C ₆)	Conversion (C ₂ =)	Catalyst Productivity (g 1-C ₆ =/g Cr)
19	92.0%	98.9%	74.4%	80,252
20	92.5%	99.1%	71.8%	77,877

[0171] The addition of TEA provided a catalyst with increased activity. It can also reduce corrosion in downstream equipment after the catalyst is inactivated.

Example 9

[0172] Several catalysts, catalysts 21-23 were prepared in which the molar ratio of the 2,5-dimethylpyrrole/chromium was varied during the addition to the solution of aluminum alkyls.

[0173] **Catalyst 21:** A chromium solution of 4.61 g chromium(III) 2-ethylhexanoate dissolved in 2.27 g ethylbenzene was divided into four equal portions of 1.72 g each. To each of these portions was added a different amount of 2,5-dimethylpyrrole. To the first was added 1.52 g 2,5-dimethylpyrrole, to the second 0.84 g, to the third 0.27 g and to the fourth 0.12 g. The chromium/2,5-dimethylpyrrole portions were then added sequentially to a solution containing 12.07 g neat triethylaluminum (TEA), 9.29 g neat diethylaluminum chloride (DEAC) and 20.01 g ethylbenzene. The total addition time was approximately 50 minutes. The resulting catalyst solution was diluted to 100 mL with ethylbenzene. The results from testing of this catalyst, in a 1 L continuous reactor, are shown as Catalyst 21 in Table 9 below. The results shown are the average of four separate test runs.

[0174] **Catalyst 22:** A chromium solution of 4.61 g chromium(III) 2-ethylhexanoate dissolved in 2.27 g ethylbenzene was divided into four portions. To each of these portions was added a different amount of 2,5-dimethylpyrrole and a similar amount of ethylbenzene. The first portion contained 0.69 g chromium solution, 1.50 g 2,5-dimethylpyrrole and 7.51 g ethylbenzene. The second contained 1.38 g chromium solution, 0.81 g 2,5-dimethylpyrrole and 7.52 g ethylbenzene. The third portion contained 2.06 g chromium solution, 0.27 g 2,5-dimethylpyrrole and 7.50 g ethylbenzene. The fourth portion contained 2.75 g chromium solution, 0.16 g 2,5-dimethylpyrrole and 7.51 g ethylbenzene. The chromium/2,5-dimethylpyrrole/ethylbenzene portions were then added sequentially to a solution containing 12.07 g neat triethylaluminum (TEA), 9.27 g neat diethylaluminum chloride (DEAC) and 25.01 g ethylbenzene. The total addition time was approximately 60 minutes. The resulting catalyst solution was then diluted to 100 mL with ethylbenzene. The results from testing of this catalyst, in a 1 L continuous reactor, are shown as Catalyst 22 in Table 9 below. The results shown are the average of two separate test runs.

[0175] **Catalyst 23:** A chromium solution of 4.61 g chromium(III) 2-ethylhexanoate dissolved in 2.27 g ethylbenzene was divided into four portions. To each of these portions was added a different amount of 2,5-dimethylpyrrole and a similar amount of ethylbenzene. The first portion contained 0.35 g chromium solution, 1.53 g 2,5-dimethylpyrrole and 7.51 g ethylbenzene. The second contained 0.69 g chromium solution, 0.81 g 2,5-dimethylpyrrole and 7.49 g ethylbenzene. The third portion contained 2.06 g chromium solution, 0.27 g 2,5-dimethylpyrrole and 7.51 g ethylbenzene. The fourth portion contained 3.77 g chromium solution, 0.15 g 2,5-dimethylpyrrole and 7.50 g ethylbenzene. The chromium/2,5-dimethylpyrrole/ethylbenzene portions were then added sequentially to a solution containing 12.09 g neat triethylaluminum (TEA), 9.26 g neat diethylaluminum chloride (DEAC) and 25.02 g ethylbenzene. The total addition time was approximately 60 minutes. The resulting catalyst solution was then diluted to 100 mL with ethylbenzene. The results from testing of this catalyst, in a 1 L continuous reactor, are shown as Catalyst 23 in Table 9 below. The results shown are the average of two separate test runs.

Example 10

[0176] **Catalyst 24:** To a dry, nitrogen purged 5 gallon reactor was added 14.6 lbs of dry, nitrogen purged ethylbenzene. The reactor was purged with nitrogen and a mixture consisting of 1,592 g neat triethylaluminum (TEA) and 1,238 g neat diethylaluminum chloride (DEAC) was added to the reactor. The aluminum alkyl mix vessel was rinsed with 0.2 lbs of ethylbenzene and this rinse was added to the reactor. A chromium solution was prepared by adding 700 mL of ethylbenzene to 630.9 g chromium(III) 2-ethylhexanoate. The mixture was stirred until solution was obtained and was transferred to a 1 gallon cylinder followed by a 75 mL ethylbenzene rinse. The cylinder, containing the chromium solution, was pressured and depressured several times with nitrogen. Chromium/2,5-dimethylpyrrole (DMP) mixtures were added to the reactor in four batches from a chromium/DMP mix tank. For the first batch 65 g of chromium and 233 mL DMP were added to the mix tank and then this mixture was added to the reactor in 31-52 g increments with stirring and cooling so the temperature did not exceed 22 °C. For the second batch 130 g of chromium and 97 mL DMP were added to the mix tank and then this mixture was added to the reactor in 48-58 g increments with stirring and cooling so the temperature did not exceed 22 °C. For the third batch 326 g of chromium and 39 mL DMP were added to the mix tank and then this mixture was added to the reactor in 48-54 g increments with stirring and cooling so the temperature did not exceed 22 °C. For the fourth batch 789 g of chromium and 20 mL DMP were added to the mix tank and then this mixture was added to the reactor in 100-130 g increments with stirring and cooling so the temperature did not exceed 24 °C. Ethylbenzene (1 lb) was added to the chromium solution cylinder and used to rinse the chromium/DMP mix tank. The ethylbenzene rinse was then added to the reactor. The reactor was stirred for an additional 30 minutes. After standing overnight the catalyst solution was filtered, using a filter as

described above. The catalyst solution was tested for activity in a 1 L continuous reactor. The results are shown as Catalyst 24 in Table 9 below. The results shown are the average of two separate test runs.

Example 11

[0177] **Catalyst 25:** To a dry, nitrogen purged 5 gallon reactor was added 14.0 lbs of dry, nitrogen purged ethylbenzene. The reactor was purged with nitrogen and a mixture consisting of 1,283 g neat triethylaluminum (TEA) and 990 g neat diethylaluminum chloride (DEAC) was added to the reactor. The aluminum alkyl mix vessel was rinsed with 0.2 lbs of ethylbenzene and this rinse was added to the reactor. A chromium solution was prepared by adding 700 mL of ethylbenzene to 630.9 g chromium(III) 2-ethylhexanoate. The mixture was stirred until solution was obtained and was transferred to a 1 gallon cylinder followed by a 75 mL ethylbenzene rinse. The cylinder, containing the chromium solution, was pressured and depressured several times with nitrogen. Chromium/2,5-dimethylpyrrole (DMP) mixtures were added to the reactor in four batches from a chromium/DMP mix tank. For the first batch 52 g of chromium and 187 mL DMP were added to the mix tank and then this mixture was added to the reactor in 20-52 g increments with stirring and cooling so the temperature did not exceed 21 °C. For the second batch 104 g of chromium and 78 mL DMP were added to the mix tank and then this mixture was added to the reactor in 40-50 g increments with stirring and cooling so the temperature did not exceed 22 °C. For the third batch 261 g of chromium and 31 mL DMP were added to the mix tank and then this mixture was added to the reactor in 90-101 g increments with stirring and cooling so the temperature did not exceed 23 °C. For the fourth batch 625 g of chromium and 16 mL DMP were added to the mix tank and then this mixture was added to the reactor in 30-108 g increments with stirring and cooling so the temperature did not exceed 23 °C.

[0178] To the TEA/DEAC mix vessel was added 327 g neat TEA and 256 g neat DEAC. To the chromium/DMP mix tank was added 261 g of the chromium solution. To a separate cylinder connected to the reactor was added 78 mL of DMP. The reactor pressure was increased with nitrogen and the valves connecting each of the above cylinders to the reactor were opened. Reducing the reactor pressure transferred the contents of each of these vessels simultaneously to the reactor while the reactor was being stirred and cooled. An increase of 1 °C (20 °C to 21°C) was observed in the reactor temperature upon addition of the catalyst components.

[0179] Ethylbenzene (0.4 lb) was added to the chromium solution cylinder and used to rinse the chromium/DMP mix tank. The ethylbenzene rinse was then added to the reactor. Ethylbenzene (0.5 lb) was added to the DMP cylinder. This rinse of the DMP cylinder was added to the reactor. Ethylbenzene (0.2 lb) was added to the aluminum alkyl mix vessel and then pressured into the reactor. The reactor was stirred for an additional 30 minutes. After standing overnight the catalyst solution was filtered, using a

filter as described above. The catalyst solution was tested for activity in a 1 L continuous reactor. The results are shown as Catalyst 25 in Table 9. The results shown are the average of three separate test runs.

Table 9

Catalyst	Selectivity (1-C ₆ =)	Purity (1-C ₆ =/C ₆)	Conversion (C ₂ =)	Catalyst Productivity (g 1-C ₆ =/g Cr)
21	92.6%	98.8%	75.1%	81,432
22	91.2%	99.1%	77.6%	82,927
23	90.5%	99.1%	79.7%	84,536
24	91.0%	99.2%	87.5%	93,397
25	90.9%	99.0%	86.6%	92,297

[0180] Catalysts 21-24 show that varying the chromium to pyrrole ratio in a decreasing manner produces a catalyst which has increased selectivity, product purity, and productivity. Catalyst 25 demonstrates the separate simultaneous addition of catalyst components to a heel of active catalyst.

Example 12

[0181] Two catalysts were prepared, catalyst 26 and catalyst 27, with the addition of a nitrogen compound to the alkylaluminum compound to solubilize products resulting from the reaction of water and aluminum alkyls.

[0182] **Catalyst 26:** To a dry 100 mL volumetric flask was added 25.01 g ethylbenzene, 12.07 g neat triethylaluminum (TEA) and 9.27 g neat diethylaluminum chloride (DEAC) and 0.34 g tributylamine. To this was added a solution containing 4.61 g chromium(III) 2-ethylhexanoate, 2.27 g ethylbenzene and 2.74 g 2,5-dimethylpyrrole. Ethylbenzene was then added to provide a total volume of 100 mL. Upon standing overnight no film was observed in the neck of the flask and no precipitate was observed. When the amine was not added to the catalyst preparation a film was observed upon standing overnight. A film was observed in the neck of the flask after standing for an additional 24 hours. This catalyst was tested for activity in a 1 L continuous reactor. The results of two separate test runs are shown in Table 10 below as Catalyst 26.

[0183] **Catalyst 27:** To a dry 100 mL volumetric flask was added 25.01 g ethylbenzene, 12.07 g neat triethylaluminum (TEA) and 9.27 g neat diethylaluminum chloride (DEAC) and 0.34 g tributylamine. To this was added a solution containing 4.61 g chromium(III) 2-ethylhexanoate, 2.27 g ethylbenzene, 2.74 g 2,5-dimethylpyrrole and 1.06 g tributylamine. Ethylbenzene was then added to provide a total volume of 100 mL. Upon standing overnight no film was observed in the neck of the flask and no precipitate was observed. When the amine was not added to the catalyst preparation a film was

observed upon standing overnight. A film was observed in the neck of the flask after standing for an additional 24 hours. This catalyst was tested for activity in a 1 L continuous reactor. The results of two separate test runs are shown in Table 10 as Catalyst 27.

Table 10

Catalyst	Selectivity (1-C ₆ =)	Purity (1-C ₆ =/C ₆)	Conversion (C ₂ =)	Catalyst Productivity (g 1-C ₆ =/g Cr)
26	93.0%	99.2%	70.4%	76,697
27	92.8%	99.2%	69.5%	75,574

[0184] The example shows that the addition of an amine to the alkylaluminum compounds inhibits formation of detrimental precipitation from the catalyst solution.

Example 13

[0185] **Catalyst 28:** Chromium(III) 2-ethylhexanoate (18.44 g) dissolved in 9.1 g ethylbenzene produces a viscous solution. When 2,5-dimethylpyrrole (10.96 g) was added to this viscous solution a much thinner solution results. This thinner solution is much more adaptable to water removal by molecular sieves. Activated 3A molecular sieves (15.05 g) were added to the chromium/pyrrole/ethylbenzene solution and allowed to stand with periodic shaking for 22 days before catalyst preparation. A solution was prepared in a 100 mL volumetric flask consisting of ethylbenzene (25.00 g), neat triethylaluminum (12.07 g) and neat diethylaluminum chloride (9.26 g). To this aluminum alkyl solution was added 9.62 g of the dried chromium/pyrrole/ethylbenzene solution and the resulting catalyst was diluted to 100 mL with additional ethylbenzene. After standing overnight a film was observed in the neck of the flask but no precipitate was observed in the flask. This catalyst was tested in a 1 L continuous reactor and an average of two separate test runs is shown in Table 11 as Catalyst 28. A control using undried chromium/pyrrole/ethylbenzene solution was made at the same time. After standing overnight a film was observed in the neck of the flask and a precipitate was also observed. This catalyst was tested in a 1 L continuous reactor and an average of two separate test runs is shown in Table 11 as Catalyst 29.

Table 11

Catalyst	Selectivity (1-C ₆ =)	Purity (1-C ₆ =/C ₆)	Conversion (C ₂ =)	Catalyst Productivity (g 1-C ₆ =/g Cr)
28	93.2%	99.4%	76.0%	83,056
29	94.3%	99.3%	64.5%	71,312

[0186] In addition to the improved catalyst productivity as shown, reduced downstream corrosion could be obtained using the dried catalyst components.

Example 14

[0187] **Catalyst 30-31:** Chromium(III) 2-ethylhexanoate (222.10 g) was added to a round bottom flask equipped with a Dean-Stark tube. Ethylbenzene (147.39 g) was added and the flask was heated to reflux the contents. Reflux was continued until water no longer accumulated in the Dean-Stark tube. Ethylbenzene and water (27.13 g) were discarded from the Dean-Stark tube. This chromium solution was used to make catalyst by adding it to a 100 mL volumetric flask containing ethylbenzene (16.73 g), neat triethylaluminum (12.28 g), neat diethylaluminum chloride (9.26 g) and 2,5-dimethylpyrrole (2.74 g). Ethylbenzene was subsequently added to dilute the catalyst to a 100 mL volume. This catalyst was tested in a 1 L continuous reactor. The results of the test (two catalyst preparations and three separate test runs) are shown in Table 12 as Catalyst 30. A control catalyst prepared similarly but with chromium(III) 2-ethylhexanoate that had not been azeotrope dried was used. The results of testing the undried preparation are shown as Catalyst 31 in Table 12.

Table 12

Catalyst	Selectivity (1-C ₆ =)	Purity (1-C ₆ =/C ₆)	Conversion (C ₂ =)	Catalyst Productivity (g 1-C ₆ =/g Cr)
30	89.4%	98.7%	81.5%	84,462
31	88.5%	98.8%	82.3%	85,400

[0188] The example shows that drying the chromium component by azeotropic distillation prepares an effective catalyst and also will reduce equipment corrosion.

Example 15

[0189] An ethylene trimerization catalyst composition was prepared using methods known to those skilled in the art, placed in the catalyst feed tank (under inert conditions) of a continuous 1-hexene production process, and aged for approximately 900 hours. The continuous 1-hexene production process was then started using the aged catalyst in the feed tank for the trimerization of ethylene to 1-hexene. Periodically, additional fresh ethylene trimerization catalyst was prepared and added to the catalyst used in the continuous 1-hexene production process. The average age of the ethylene oligomerization catalyst composition periodically calculated to determine the average time the catalyst had resided in the catalyst feed tank based upon the average catalyst composition in the catalyst feed tank. Throughout the continuous 1-hexene production process, samples of the continuous 1-hexene

production process product we removed and analyzed for 1-hexene content. Fig. 5 shows the impact of the average catalyst residence time (i.e. catalyst age) on the purity of the hexene production produced by the continuous 1-hexene production process. Fig. 5 indicates that the purity of the 1-hexene product is negatively impacted by increasing age of the ethylene trimerization catalyst.

[0190] During the manufacture of chromium(III) 2-ethylhexanoate, also denoted as $\text{Cr}(\text{EH})_3$, a mixture of compounds is obtained. These compounds include the desired $\text{Cr}(\text{EH})_3$, hydrated chromium species, free and coordinated 2-ethylhexanoic acid, chromium oligomers and free water. In examples 16-27 the chemical composition of $\text{Cr}(\text{EH})_3$ and its effect on the oligomerization activity of the catalyst were investigated. $\text{Cr}(\text{EH})_3$ preparation is presented in Example 16. The impact of production variables in $\text{Cr}(\text{EH})_3$ preparation on oligomerization catalyst activity is investigated in Example 17 while in Example 18 the effect of chromium concentration and free acid addition on catalyst activity was determined. In Example 19, the effect of hydrated chromium species and chromium oligomers on catalyst activity was determined while the effect of water content on the catalyst activity was investigated in Example 20. The chemical composition of $\text{Cr}(\text{EH})_3$ provided by a second supplier was investigated using infrared analysis in Example 21. This material was further subjected to water abatement protocols as described in Examples 22 and 23. Infrared analysis of $\text{Cr}(\text{EH})_3$ from Supplier A was recorded in the absence of solvent with heating is presented in Example 24. Finally, the methodology for preparation of a 1-hexene catalyst is presented in Example 25 while testing of these catalysts in a batch trimerization or continuous trimerization run is detailed in Examples 26 and 27 respectively.

Example 16

[0191] A sample of chromium(III) 2-ethylhexanoate was prepared by dissolving 60.02 g (1.5 moles) of sodium hydroxide in 249.13 g of distilled water. 2-Ethylhexanoic acid (238.57 g, 1.65 moles) was added with stirring to form sodium 2-ethylhexanoate. In a separate container, 100.00 g (0.25 mole) of chromium nitrate nonahydrate (Alfa Aesar, 98.5%) was dissolved in 250 ml of distilled water. The chromium nitrate solution was slowly added to the sodium 2-ethylhexanoate solution with good stirring (stirring became difficult towards the end of the addition). When the addition was complete, 250 mL of hexanes (Fischer, H303-4) were added and stirring was continued for 15 minutes. The layers were separated and the hexane layer (and a 60 mL hexane wash) containing the chromium(III) 2-ethylhexanoate was washed with 100 mL 5% sodium hydroxide solution two times, water (100 mL), 10% sodium carbonate solution (100 mL) and finally with distilled water (100 mL) two times. Additional hexane was added and the hexane solution was then dried over anhydrous magnesium sulfate. The mixture was filtered with a Buchner funnel and a water aspirator. Frequent changes of the filter paper were required. The reduced pressure of a water aspirator removed most of the hexane. Additional hexane was added and the resulting solution was slowly added to 250 mL of acetone. A blue granular solid was

filtered from the acetone and air dried to yield trihydrated chromium(III) 2-ethylhexanoate. The infrared spectrum of a dilute ethylbenzene solution of this blue solid showed a strong absorption at 1540 cm^{-1} and no absorption peaks at either 1600 or 1700 cm^{-1} (Figure 6).

Example 17

[0192] The effect of chromium concentration and free acid addition on catalyst activity was determined using three lots of $\text{Cr}(\text{EH})_3$ prepared in the laboratory. Lot 1 contained 11.5 wt% chromium. Lot 2 contained 10.5 wt% and was the same as Lot 1 but the chromium concentration was adjusted by the addition of 2-ethylhexanoic acid. Lot 3 contained 10.5 wt% chromium and had been chemically dried with acetic anhydride to remove water. Catalysts were prepared from each of these lots and the catalysts were evaluated in the continuous reactor as described in Example 27. Catalyst preparations A and B differ in the ratios of the catalyst components. Both types of catalysts were prepared from these samples and tested. Low reactor chromium concentrations were used to magnify any differences. The results of the continuous reactor tests are shown in Table 13.

Table 13. Effect of Excess Acid and Drying Agent in the Chromium Component

Catalyst Lot	1-Hexene (% Selectivity)	1-Hexene (% Purity)	Catalyst Productivity (g 1-Hexene/g Cr)
<i>Catalyst A</i> ^a			
1	96.1	99.3	59,500
2	94.7	99.2	44,500
3	96.1	99.3	58,900
<i>Catalyst B</i> ^b			
1	96.5	99.3	88,700
2	95.9	na ^c	23,500
3	95.2	99.0	97,900

^a Mole ratio Cr/DMP/TEA/DEAC is 1/1.8/6.5/5, 2.0 ppm Cr in reactor.

^b Mole ratio Cr/DMP/TEA/DEAC is 1/3/11/8, 1.5 ppm Cr in reactor.

^c Not available as low productivity dropped amount of internal hexenes below detection limits.

The results demonstrate the detrimental effect of increased acid as seen in the lower catalyst productivity for Lot 2.

Example 18

[0193] Three samples (17-1, 17-2, and 17-3) of chromium(III) 2-ethylhexanoate were prepared from a chromium(III) 2-ethylhexanoate stock solution. The three samples were taken from the chromium(III) 2-ethylhexanoate stock solution at different heating and stripping times. These samples contain a mixture of compounds and are denoted herein as Cr-mix. These different compositions of Cr-mix were used to

prepare oligomerization catalysts as described in Example 25. In Sample 17-1 the chromium content of the Cr-mix was 10.5 wt%. The infrared spectrum of chromium(III) 2-ethylhexanoate sample 17-1 showed absorption peaks at 1540 and 1600 cm^{-1} (Figure 7). The absorption at 1600 cm^{-1} $\text{Cr}(\text{EH})_3$ was attributed to the nonhydrated $\text{Cr}(\text{EH})_3$ while the 1540 cm^{-1} absorption is expected for both the hydrated $\text{Cr}(\text{EH})_3$ and chromium oligomers. In order to estimate the amount of chromium oligomers in the mixture, a methanol solubility test was performed. The methanol solubility test can help in ascertaining the amount of chromium oligomers present, as they are insoluble in methanol. A compound that has greater than 90% of the material dissolve in methanol (i.e. a small amount of precipitate) has a low chromium oligomer content. The methanol solubility test showed little precipitate. In sample 17-2 the chromium content of the $\text{Cr}(\text{EH})_3$ was 10.9 wt%. The infrared spectrum of chromium(III) 2-ethylhexanoate sample 17-2 showed strong absorption at 1600 cm^{-1} with only a small peak at 1540 cm^{-1} (Figure 8). The methanol solubility test showed little precipitate. In sample 17-3 the chromium concentration was 12.7 wt%. The infrared spectrum of chromium(III) 2-ethylhexanoate sample 17-3 showed absorption peaks at 1540 and 1600 cm^{-1} (Figure 9). The methanol solubility test showed a large amount of precipitate. The oligomerization catalysts prepared using these chromium sources were tested in the continuous reactor as described in Example 27. The results (Table 14) demonstrate the effect of chromium oligomers and hydrated species in the chromium(III) 2-ethylhexanoate.

Example 19

[0194] The effect of hydrated chromium species and chromium oligomers on catalyst activity was determined. The hydrated chromium species and chromium oligomer amounts vary with the chromium preparation conditions. In order to investigate the effects of the chromium preparation variations, samples from a single $\text{Cr}(\text{EH})_3$ preparation that differed in the time of heating and vacuum stripping were obtained. The three different samples of $\text{Cr}(\text{EH})_3$ were used to prepare selective 1-hexene catalysts as described in Example 25. The amounts of hydrated chromium species and chromium oligomers were estimated by a combination of infrared analysis and methanol solubility. Both the hydrated chromium species and chromium oligomers show infrared absorption at 1540 cm^{-1} but the hydrated chromium species is soluble in the methanol test while chromium oligomers form precipitates. The rationale for the infrared absorption peak assignments is described in Example 24.

[0195] The chromium concentration in sample 17-1 was 10.5 wt%. This preparation has the shortest heating and vacuum stripping time. Infrared analysis of chromium(III) 2-ethylhexanoate sample 17-1 (Figure 7) and solubility in the methanol test indicated that some hydrated chromium species were present but only small amounts of chromium oligomers were present. In sample 17-2 additional heating and vacuum stripping resulted in a chromium concentration of 10.9 wt%. Infrared analysis of chromium(III) 2-ethylhexanoate sample 17-2 (Figure 8) and methanol solubility tests indicated that only

small amounts of either hydrated chromium species or chromium oligomers were present. In sample 17-3 the heating and vacuum stripping time was extended and a chromium concentration of 12.7 wt% was obtained. This sample would represent a high chromium concentration. Infrared analysis of chromium(III) 2-ethylhexanoate sample 17-3 (Figure 9) indicated the presence of either hydrated chromium species or chromium oligomers (1540 cm^{-1}). Since hydrated chromium species were not present in the earlier sample (sample 17-2), chromium oligomers were presumed to be present. A sizeable precipitate in the methanol solubility test confirmed the presence of chromium oligomers. The catalysts prepared were tested for activity in the continuous reactor (as described in Example 27) and the average results of two reactor runs for each preparation are shown in Table 14.

Table 14. Effect of Hydrated Chromium Species and Chromium Oligomers in Catalyst Preparation

Sample	Wt% Cr	% Selectivity (1-C ₆ =)	% Purity (1-C ₆ =/C ₆)	% Conversion (C ₂ =)	Productivity (g 1-C ₆ =/g Cr)
17-1	10.5	90.3	99.10	79.0	83,642
17-2	10.9	88.7	99.09	84.5	87,882
17-3	12.7	87.4	98.07	86.4	88,460

[0196] The results demonstrate that longer heating and vacuum stripping times resulted in higher chromium concentrations, lower hydrated chromium species, lower acid concentration and greatly increased viscosity. Table 14 shows that the best combination of product purity and catalyst productivity is obtained when hydrated chromium species and chromium oligomers are minimized in the chromium(III) 2-ethylhexanoate.

Example 20

[0197] The effect of water content in the chromium compound on catalyst activity was determined. An ethylbenzene azeotrope was formed in order to remove water from the Cr(EH)₃. The binary azeotrope of water with ethylbenzene actually contains a higher mole fraction of water (0.744) than the binary azeotrope of either benzene (0.295) or toluene (0.444). A base material on hand for 7 years, chromium(III) 2-ethylhexanoate (222.10 g), was added to a round bottom flask equipped with a Dean-Stark tube. Ethylbenzene (147.39 g) was added and the flask was heated to reflux the contents. Reflux was continued until water no longer accumulated in the Dean-Stark tube. A total 2.6 mL of water was collected in the Dean-Stark tube. Water analysis according to Karl Fisher titration ASTM E 1064-04a showed 96 ppm water to be present in the final Cr(EH)₃ ethylbenzene solution. This chromium solution was used to make catalyst as described in Example 25. Testing of a catalyst prepared from the dried Cr(EH)₃ was done in the continuous reactor (as described in Example 27) and showed similar productivity

and 1-hexene purity to the standard catalyst. The results demonstrated that water could be removed from $\text{Cr}(\text{EH})_3$ by azeotropic distillation.

[0198] A second set of samples of $\text{Cr}(\text{EH})_3$ in ethylbenzene were prepared for testing, and low moisture levels (500-700 ppm water) were observed. These low moisture levels were attributed to the ethylbenzene water azeotrope.

Example 21

[0199] Infrared analysis was used to characterize a second $\text{Cr}(\text{EH})_3$ sample (chromium(III) 2-ethylhexanoate Sample 17-4a). The sample was a blue-violet solid powder containing 14.6 wt% chromium. This material was found to be much less soluble in ethylbenzene than the $\text{Cr}(\text{EH})_3$ material described in example 17. Solubility and color indicated that this material contained hydrated chromium species. The presence of hydrated chromium species in chromium(III) 2-ethylhexanoate sample 17-4a was also consistent with the infrared analysis that showed a large peak at 1540 cm^{-1} and only a very small (trace) peak at 1600 cm^{-1} (Figure 10). There was no significant absorption peak at 1700 cm^{-1} indicating the low amount of free 2-ethylhexanoic acid.

Example 22

[0200] The blue violet solid described in Example 21, chromium(III) 2-ethylhexanoate Sample 17-4a, was azeotroped with ethylbenzene to see if the water of hydration could be removed in this manner. Heating 30.03 g of this material in 75.2 g ethylbenzene formed a gel that was difficult to handle. The gel had to be transferred twice to larger flasks. In the end a total of 367.7 g of ethylbenzene was added. Very little water was collected in the Dean Stark tube after a lengthy reflux time. Part of this water may have come from contact with the atmosphere during the difficult transfers to larger flasks. The color of the solution remained blue. At this low $\text{Cr}(\text{EH})_3$ concentration there is some interference in the infrared spectrum (Figure 11 – chromium(III) 2-ethylhexanoate Sample 17-4b) in the 1600 cm^{-1} peak with the ethylbenzene absorption peak at 1605 cm^{-1} . Color, solubility and the infrared analysis are consistent with the presence of hydrated chromium species remaining after attempted removal of the water of hydration by azeotroping. When 2-ethylhexanoic acid was added to this solution, water was observed accumulating in the Dean-Stark tube and the solution turned a green color. The infrared spectrum (Figure 12 - chromium(III) 2-ethylhexanoate Sample 17-4c) shows the peak at 1600 cm^{-1} becoming much larger than the peak at 1540 cm^{-1} . The color change, water accumulation and the shift in the infrared peaks are consistent with the removal of water from hydrated chromium species. A peak at 1700 cm^{-1} was observed indicating the presence of free 2-ethylhexanoic acid.

Example 23

[0201] The effect of acid addition on water abatement for the blue violet solid $\text{Cr}(\text{EH})_3$ Sample 17-4a was investigated. 2-Ethylhexanoic acid (3 g, 0.021 moles) from Aldrich was added to the solid blue-violet $\text{Cr}(\text{EH})_3$ Sample 17-4a (3 g, ca. 0.006 moles) in a 50 mL Erlenmeyer flask, this constituted a 3.5 molar excess of 2-ethylhexanoic acid. Some liquid acid remained after the solid was wet. A solution was not obtained. The mixture was heated to 192 °C for 2 hours without the presence of a solvent. The color changed to deep green. After cooling a green semisolid was obtained. When ethylbenzene was added to the green semisolid the solubility of the $\text{Cr}(\text{EH})_3$ was greatly increased. The infrared spectrum (Figure 13) of the resulting semisolid (chromium(III) 2-ethylhexanoate Sample 17-4d), dissolved in ethylbenzene, showed a strong absorption at 1600 cm^{-1} with only a small absorption at 1540 cm^{-1} . The color change, solubility and infrared analysis demonstrate the conversion of hydrated chromium species to the non-hydrated form by heating with acid.

[0202] To confirm the infrared absorption peak assignments, additional experiments were performed. A sample of trihydrated $\text{Cr}(\text{EH})_3$ was prepared in the laboratory from chromium nitrate and sodium 2-ethylhexanoate in aqueous solution (Example 16). A blue solid material with low solubility in ethylbenzene was obtained. Infrared analysis (Figure 6) showed a large absorption peak at 1540 cm^{-1} . There was no significant absorption peak at 1600 cm^{-1} (ethylbenzene subtracted spectrum) and no significant free acid peak at 1700 cm^{-1} .

Example 24

[0203] A 10.5 wt% $\text{Cr}(\text{EH})_3$ described in Example 19 was heated, without solvent to determine what changes in the infrared spectrum would be observed. The sample of the 10.5 wt% $\text{Cr}(\text{EH})_3$ was heated to 200-240°C for 2-3 hours to make chromium(III) 2-ethylhexanoate Sample 17-4e. No additional 2-ethylhexanoic acid was added. The infrared spectrum of an ethylbenzene solution of the resulting semisolid (chromium(III) 2-ethylhexanoate Sample 17-4e) showed the absorption peak at 1540 cm^{-1} was larger than the peak at 1600 cm^{-1} (Figure 14).

[0204] Since this heating did not involve additional acid, an increase in chromium oligomers was expected. Comparison of the infrared spectrum of the heated material with that of the starting $\text{Cr}(\text{EH})_3$ (chromium(III) 2-ethylhexanoate Sample 17-4f - Figure 15) shows the absorption peak at 1540 cm^{-1} increases to become larger than the peak at 1600 cm^{-1} . This is the reverse absorption intensity of the unheated $\text{Cr}(\text{EH})_3$. An absorption peak at 1540 cm^{-1} is consistent with the presence of chromium oligomers as was indicated in the extended heating and vacuum stripping samples (Figure 9) discussed earlier. In summary the infrared absorption peak at 1540 cm^{-1} can be attributed to both hydrated

chromium species and chromium oligomers. The methanol solubility test can help in ascertaining the amount of chromium oligomers present, as they are insoluble in methanol. The absorption peak at 1600 cm^{-1} is attributed to the unhydrated $\text{Cr}(\text{EH})_3$ and the absorption peak at 1700 cm^{-1} is due to the free acid carbonyl.

Example 25

[0205] A typical catalyst preparation as used in the Examples is given here. The catalyst preparation was done in a drybox under an inert gas atmosphere. Dry, degassed ethylbenzene (20.01 g) was added to a weighed, dry 100 mL volumetric flask (dried overnight in a glass oven). To this flask was added 12.08 g neat triethylaluminum and 9.26 g neat diethylaluminum chloride. The contents were mixed and allowed to stand for 15 minutes. Then 2,5-dimethylpyrrole (2.76 g) was carefully added down the side of the flask. The contents were mixed (gas evolution was observed) and the weight loss was determined after 30 minutes to be 0.60 g. To this mixture sufficient $\text{Cr}(\text{EH})_3$ diluted in ethylbenzene (6.89 g; 7.27 wt% Cr) was added to supply 500 mg chromium metal. When neat $\text{Cr}(\text{EH})_3$ was used the ethylbenzene dilution was done the day before the catalyst preparation to allow time for the $\text{Cr}(\text{EH})_3$ to dissolve in the ethylbenzene. The chromium solution was added over about 40 minutes. The weight loss was determined 30 minutes after the end of chromium addition and was 1.05 g. The volume was brought to 100 mL by the addition of 39.14 g of ethylbenzene. There was a total of 61.43 g of ethylbenzene added including the ethylbenzene included in the chromium. The net weight of the catalyst was 88.49 g. After standing overnight the solution obtained a reddish orange color. The catalyst had a concentration of 5 mg Cr/mL and was active as prepared.

Example 26

[0206] Batch reactor trimerization runs were conducted as indicated in various Examples. Trimerization runs were carried out in a 1 Liter stainless steel autoclave reactor equipped with a magna drive stirrer and a bottom valve. The temperature was controlled at the desired set point by an internal steam/water cooling coil. A typical run was carried out using the following procedure. The reactor was purged with nitrogen at $100\text{ }^{\circ}\text{C}$ for at least 15 minutes to remove moisture, oxygen and other volatile impurities. During this time 2-3 mL of 2-ethyl-1-hexanol was added to the line between the reactor bottom valve and the product cylinder. The catalyst (1 mL) was diluted to 25 mL with dry, nitrogen purged cyclohexane and the amount of diluted catalyst was weighed and added to a 50 mL stainless steel cylinder in the drybox. A measured amount of cyclohexane (225 mL) was added to the reactor and the catalyst charged into the line to the reactor. The catalyst was then washed into the reactor with another measured amount of cyclohexane (225 mL). The weight of the measured amount of cyclohexane had been previously determined by weighing several charges. The reactor was then allowed to come to the

desired operating temperature (115 °C) within about 5 minutes. Care was taken during the addition of the components to add minimum nitrogen pressure to the reactor. Hydrogen (a delta of 50 psig) was then added. The reactor was then pressurized with ethylene to the desired pressure (650-800 psig) and ethylene was fed on demand at the desired pressure for 30 minutes. The reactor stirrer was stopped and the contents of the reactor were then charged to the product cylinder through the bottom valve. The product cylinder was weighed to determine the weight gain from the reaction. The consumption of ethylene was determined from the ethylene flow meter and verified by the increase in the weight of the reactor contents. A sample of the contents of the product cylinder was sent to the GC through a liquid sampling valve.

Example 27

[0207] Continuous reactor trimerization runs were conducted as indicated in various Examples. The continuous ethylene trimerization runs were carried out in a 1 Liter stirred autoclave reactor. In a typical run, cyclohexane solvent was first charged to the reactor. A 3 mL aliquot of 1.9 M TEA was then added to the reactor to remove any residual moisture and the reactor was heated to 115 °C (solvent was pumped through the system while the reactor was being heated). The catalyst feed pump was then started and the catalyst was added to the reactor for 30 minutes at double its normal rate. The catalyst feed rate was then reduced to normal and ethylene and hydrogen were added to the solvent feed at the desired feed rates. Separate control systems were provided for each individual feed to the reactor. Standard catalyst preparations (Cr/DMP/TEA/DEAC mole ratios of 1/3/11/8) were used for these tests. The run was carried out continuously for six hours.

[0208] The product stream was monitored by on-line GC samples on an hourly basis and readings of pressure and temperature were recorded every 30 minutes. The catalyst feed rate was also checked every 30 minutes. The reactor was run liquid full on back pressure control. 2-Ethyl-1-hexanol was added to the reactor product stream immediately after the reactor and the reactor product then passed through a cooled filter (1 Liter autoclave packed with stainless steel sponge) before being sent to a product tank. At the end of the run the ethylene, hydrogen and catalyst feeds were shut off. The reactor was flushed with solvent for 30 minutes and allowed to cool. The contents of the reactor were then blown into the product tank, the reactor depressurized and disassembled for cleaning. Any polymer formed in the reactor was collected, dried and weighed. The polymer filter material was also collected, dried and weighed. The filter system was cleaned and new, weighed filter material put in place for the next run.

[0209] While preferred embodiments of the invention have been shown and described, modifications thereof can be made by one skilled in the art without departing from the spirit and teachings of the invention. The embodiments described herein are exemplary only, and are not intended to be limiting.

Many variations and modifications of the invention disclosed herein are possible and are within the scope of the invention. Where numerical ranges or limitations are expressly stated, such express ranges or limitations should be understood to include iterative ranges or limitations of like magnitude falling within the expressly stated ranges or limitations (e.g., from about 1 to about 10 includes, 2, 3, 4, etc.; greater than 0.10 includes 0.11, 0.12, 0.13, etc.). Use of the term "optionally" with respect to any element of a claim is intended to mean that the subject element is required, or alternatively, is not required. Both alternatives are intended to be within the scope of the claim. Use of broader terms such as comprises, includes, having, etc. should be understood to provide support for narrower terms such as consisting of, consisting essentially of, comprised substantially of, etc.

[0210] Accordingly, the scope of protection is not limited by the description set out above but is only limited by the claims which follow, that scope including all equivalents of the subject matter of the claims. Each and every claim is incorporated into the specification as an embodiment of the present invention. Thus, the claims are a further description and are an addition to the preferred embodiments of the present invention. The discussion of a reference in the Description of Related Art is not an admission that it is prior art to the present invention, especially any reference that may have a publication date after the priority date of this application. The disclosures of all patents, patent applications, and publications cited herein are hereby incorporated by reference, to the extent that they provide exemplary, procedural or other details supplementary to those set forth herein.

CLAIMS

What is claimed is:

1. A method of making an olefin oligomerization catalyst, comprising contacting a chromium-containing compound, a heteroatomic ligand, and a metal alkyl, wherein the chromium-containing compound comprises less than about 5 weight percent chromium oligomers.
2. The method of claim 1 wherein the olefin oligomerization catalyst further comprises a halide containing compound.
3. The method of claim 1 wherein the chromium-containing compound comprises less than about 5 weight percent hydrated chromium species.
4. The method of claim 1 wherein the chromium-containing compound comprises less than about 50 weight percent free acid.
5. The method of 1 wherein the chromium-containing compound comprises a chromium carboxylate.
6. The method of claim 5 wherein the chromium carboxylate comprises chromium(III) 2-ethylhexanoate.
7. The method of claim 1 wherein the heteroatomic ligand comprises a nitrogen-containing compound.
8. The method of claim 1 wherein the heteroatomic ligand comprises a pyrrole-containing compound.
9. The method of claim 1 wherein a composition comprising the chromium-containing compound is added to a composition comprising the metal alkyl.
10. The method of claim 1 further comprising abating all or a portion of water, acidic protons, or both from a composition comprising the chromium-containing compound, a composition comprising the heteroatomic ligand, or combinations thereof.
11. The method of claim 10 wherein the olefin oligomerization catalyst further comprises a non metal halide containing compound and the method further comprises abating all or a portion of water, acidic protons, or both from the non metal halide containing compound.

12. A method of oligomerizing olefins comprising contacting the catalyst of claim 1 with an alpha olefin.
13. The method of claim 12 wherein the method of oligomerizing olefins comprises trimerizing or tetramerizing ethylene.
14. A method of making an olefin oligomerization catalyst comprising a chromium-containing compound, a nitrogen-containing compound, and a metal alkyl, the method comprising adding a composition comprising the chromium-containing compound to a composition comprising the metal alkyl.
15. The method of claim 14 wherein the chromium-containing compound comprises less than about 5 weight percent chromium oligomers.
16. A method of making an olefin oligomerization catalyst comprising a chromium-containing compound, a nitrogen-containing compound, and a metal alkyl, the method comprising abating all or a portion of water, acidic protons, or both from a composition comprising the chromium-containing compound, a composition comprising the nitrogen-containing compound, or combinations thereof prior to or during the preparation of the catalyst.
17. The method of claim 16 wherein the olefin oligomerization catalyst further comprises a non metal halide containing compound and the method further comprises abating all or a portion of water, acidic protons, or both from the non metal halide containing compound prior to or during the preparation of the catalyst.
18. The method of claim 16 wherein the chromium-containing compound comprises less than about 5 weight percent chromium oligomers.
19. A method of making an olefin oligomerization catalyst comprising a chromium-containing compound, a heteroatomic ligand, a metal alkyl, the method comprising abating all or a portion of water, acidic protons, or both from a composition comprising the chromium-containing compound, a composition comprising the heteroatomic ligand, or combinations thereof, and wherein the heteroatomic ligand is described by the general formula $(R)_nA-B-C(R)_m$ wherein A and C are independently selected from a group consisting of phosphorus arsenic, antimony, oxygen, bismuth, sulfur, selenium, and nitrogen; B is a linking group between A and C; each R is independently selected from any homo or hetero hydrocarbyl group; and n and m are determined by the respective valence and oxidation state of A and C.

20. The method of claim 19 wherein the olefin oligomerization catalyst further comprises a non metal halide containing compound and the method further comprises abating all or a portion of water, acidic protons, or both from the non metal halide containing compound prior to or during the preparation of the catalyst.

21. The method of claim 19 wherein the chromium-containing compound comprises less than about 5 weight percent chromium oligomers.