

Process for the purification of water

The present invention relates to a process for the purification of water and the use of surface-reacted natural calcium carbonate in such a process.

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There is a wide range of water purification techniques that can be used to remove fine solids, micro-organisms and dissolved inorganic and organic materials. The choice of method will depend on the quality of the water being treated, the cost of the treatment process and the quality standards expected of the processed water.

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Flocculation is widely employed in the purification of water such as industrial waste water or drinking water. Flocculation refers to a process where dissolved compounds and/or colloidal particles are removed from the solution in the form of flocs or "flakes." The term is also used to refer to the process by which fine particulates are caused to clump together into flocs. The flocs may then float to the top of the liquid, settle to the bottom of the liquid, or can be readily filtered from the liquid.

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Common flocculants or coagulants used are aluminium sulphate or polyaluminium chloride (PAC). Aluminium sulphate reacts with water to form flocs of aluminium hydroxide. Coagulation with aluminum compounds may leave a residue of aluminium in the finished water. Aluminium can be toxic to humans at high concentrations. In solutions of poly-aluminium chloride (PAC), aluminium ions have formed into polymers consisting of clusters of ions bridged by oxygen atoms. PAC is used e.g. for the treatment of brown drinking water comprising organic materials such as leaves and/or inorganic materials such as iron and manganese compounds which cause the brown discolouration. However, PAC is generally not sufficient to remove all brown discolouration from the water.

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Iron(III) chloride is another common coagulant. Iron(III) coagulants work over a larger pH range than aluminium sulphate but are not effective with many source waters. Coagulation with iron compounds typically leaves a residue of iron in the finished water. This may impart a slight taste to the water, and may cause brownish

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stains on porcelain fixtures. Furthermore, iron(III) chloride presents corrosion risks in the water treatment system.

Further well-known adsorbents for the water treatment based on a high specific
5 surface area such as activated carbon or bentonite have the general drawback that they are very difficult to separate after the adsorption of the substance to be removed from the medium due to their finely divided state

Thus, considering the drawbacks of known agents for water purification mentioned
10 above, it is an object of the present invention to provide an improved water purification process providing for an efficient removal of water impurities.

The object outlined above has been solved by a process for the purification of water, wherein surface-reacted natural calcium carbonate or an aqueous suspension
15 comprising surface-reacted natural calcium carbonate and having a pH of greater than 6.0, measured at 20°C, is added to the water to be purified, the surface-reacted natural calcium carbonate being the reaction product of a natural calcium carbonate with carbon dioxide and one or more acids.

20 The term "purification" is to be interpreted broadly and means any removal of harmful compounds and/or other compounds not tolerated in the water.

The water preferably treated by the process of the present invention includes industrial waste water, drinking water, urban waste water, waste water from
25 breweries or other beverage industries or water in the paper industry, be it white or waste water, as well as agricultural waste-water.

The surface-reacted natural calcium carbonate as defined above and further below is efficiently adsorbing organic as well as inorganic impurities on its surface.

The surface-reacted natural calcium carbonate to be used in the process of the present invention is obtained by reacting a natural calcium carbonate with an acid and with carbon dioxide.

- 5 Preferably, the natural calcium carbonate is selected from the group comprising marble, chalk, calcite, dolomite, limestone and mixtures thereof. In a preferred embodiment, the natural calcium carbonate is ground prior to the treatment with an acid and carbon dioxide. The grinding step can be carried out with any conventional grinding device such as a grinding mill known to the skilled person.

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- The surface-reacted natural calcium carbonate to be used in the water purification process of the present invention is prepared as an aqueous suspension having a pH, measured at 20°C, of greater than 6.0, preferably greater than 6.5, more preferably greater than 7.0, even more preferably greater than 7.5. As will be discussed below,
- 15 the surface-reacted natural calcium carbonate can be brought into contact with the water to be purified by adding said aqueous suspension to the water. It is also possible to modify the pH of the aqueous suspension prior to its addition to the water to be purified, e.g. by dilution with additional water. Alternatively, the aqueous suspension can be dried and the surface-reacted natural calcium carbonate brought
- 20 into contact with the water is in powder form or in the form of granules. In other words, the increase of pH to a value of greater than 6.0 subsequent to treatment with an acid and carbon dioxide is needed to provide the surface-reacted calcium carbonate having the beneficial adsorption properties described herein.

- 25 In a preferred process for the preparation of the aqueous suspension, the natural calcium carbonate, either finely divided, such as by grinding, or not, is suspended in water. Preferably, the slurry has a content of natural calcium carbonate within the range of 1 wt% to 80 wt%, more preferably 3 wt% to 60 wt%, and even more preferably 5 wt% to 40 wt%, based on the weight of the slurry.

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In a next step, an acid is added to the aqueous suspension containing the natural calcium carbonate. Preferably, the acid has a pK_a at 25°C of 2.5 or less. If the pK_a at 25°C is 0 or less, the acid is preferably selected from sulphuric acid, hydrochloric acid, or mixtures thereof. If the pK_a at 25°C is from 0 to 2.5, the acid is preferably
5 selected from H_2SO_3 , HSO_4^- , H_3PO_4 , oxalic acid or mixtures thereof. The one or more acids can be added to the suspension as a concentrated solution or a more diluted solution. Preferably, the molar ratio of the acid to the natural calcium carbonate is from 0.05 to 4, more preferably from 0.1 to 2.

- 10 As an alternative, it is also possible to add the acid to the water before the natural calcium carbonate is suspended.

In a next step, the natural calcium carbonate is treated with carbon dioxide. If a strong acid such as sulphuric acid or hydrochloric acid is used for the acid treatment
15 of the natural calcium carbonate, the carbon dioxide is automatically formed. Alternatively or additionally, the carbon dioxide can be supplied from an external source.

Acid treatment and treatment with carbon dioxide can be carried out simultaneously
20 which is the case when a strong acid is used. It is also possible to carry out acid treatment first, e.g. with a medium strong acid having a pK_a in the range of 0 to 2.5, followed by treatment with carbon dioxide supplied from an external source.

Preferably, the concentration of gaseous carbon dioxide in the suspension is, in terms
25 of volume, such that the ratio (volume of suspension):(volume of gaseous CO_2) is from 1:0.05 to 1:20, even more preferably 1:0.05 to 1:5.

In a preferred embodiment, the acid treatment step and/or the carbon dioxide treatment step are repeated at least once, more preferably several times.
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Subsequent to the acid treatment and carbon dioxide treatment, the pH of the aqueous suspension, measured at 20°C, naturally reaches a value of greater than 6.0, preferably greater than 6.5, more preferably greater than 7.0, even more preferably greater than 7.5, thereby preparing the surface-reacted natural calcium carbonate as
5 an aqueous suspension having a pH of greater than 6.0, preferably greater than 6.5, more preferably greater than 7.0, even more preferably greater than 7.5. If the aqueous suspension is allowed to reach equilibrium, the pH is greater than 7. A pH of greater than 6.0 can be adjusted without the addition of a base when stirring of the aqueous suspension is continued for a sufficient time period, preferably 1 hour to 10
10 hours, more preferably 1 to 5 hours.

Alternatively, prior to reaching equilibrium, which occurs at a pH greater than 7, the pH of the aqueous suspension may be increased to a value greater than 6 by adding a base subsequent to carbon dioxide treatment. Any conventional base such as sodium
15 hydroxide or potassium hydroxide can be used.

With the process steps described above, i.e. acid treatment, treatment with carbon dioxide and pH adjustment, a surface-reacted natural calcium carbonate is obtained having good adsorption properties for organic as well as inorganic impurities which
20 might be present in waste water.

Further details about the preparation of the surface-reacted natural calcium carbonate are disclosed in WO 00/39222 and US 2004/0020410 A1, wherein the surface-reacted natural calcium carbonate is described as a filler for paper manufacture, the
25 content of these references herewith being included in the present application.

In a preferred embodiment of the preparation of the surface-reacted natural calcium carbonate, the natural calcium carbonate is reacted with the acid and/or the carbon dioxide in the presence of at least one compound selected from the group consisting
30 of silicate, silica, aluminium hydroxide, earth alkali aluminate such as sodium or

potassium aluminate, magnesium oxide, or mixtures thereof. Preferably, the at least one silicate is selected from an aluminium silicate, a calcium silicate, or an earth alkali metal silicate. These components can be added to an aqueous suspension comprising the natural calcium carbonate before adding the acid and/or carbon
5 dioxide.

Alternatively, the silicate and/or silica and/or aluminium hydroxide and/or earth alkali aluminate and/or magnesium oxide component(s) can be added to the aqueous suspension of natural calcium carbonate while the reaction of natural calcium
10 carbonate with an acid and carbon dioxide has already started. Further details about the preparation of the surface-reacted natural calcium carbonate in the presence of at least one silicate and/or silica and/or aluminium hydroxide and/or earth alkali aluminate component(s) are disclosed in WO 2004/083316, the content of this reference herewith being included in the present application.

15 The surface-reacted natural calcium carbonate can be kept in suspension, optionally further stabilised by a dispersant. Conventional dispersants known to the skilled person can be used. The dispersant can be anionic or cationic. A preferred dispersant is polyacrylic acid.

20 Alternatively, the aqueous suspension described above can be dried, thereby obtaining the surface-reacted natural calcium carbonate in the form of granules or a powder.

25 In a preferred embodiment, the surface-reacted natural calcium carbonate has a specific surface area of from $5 \text{ m}^2/\text{g}$ to $200 \text{ m}^2/\text{g}$, more preferably $20 \text{ m}^2/\text{g}$ to $80 \text{ m}^2/\text{g}$ and even more preferably $30 \text{ m}^2/\text{g}$ to $60 \text{ m}^2/\text{g}$, measured using nitrogen and the BET method according to ISO 9277.

Furthermore, it is preferred that the surface-reacted natural calcium carbonate has a mean grain diameter of from 0.1 to 50 μm , more preferably from 0.5 to 25 μm , even more preferably 1 to 10 μm , measured according to the sedimentation method. The sedimentation method is an analysis of sedimentation behaviour in a gravimetric field. The measurement is made with a SedigraphTM 5100 of Micromeritics Instrument Corporation. The method and the instrument are known to the skilled person and are commonly used to determine grain size of fillers and pigments. The measurement is carried out in an aqueous solution of 0.1 wt% $\text{Na}_4\text{P}_2\text{O}_7$. The samples were dispersed using a high speed stirrer and supersonic.

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In a preferred embodiment, the surface-reacted natural calcium carbonate has a specific surface area within the range of 15 to 200 m^2/g and a mean grain diameter within the range of 0.1 to 50 μm . More preferably, the specific surface area is within the range of 20 to 80 m^2/g and the mean grain diameter is within the range of 0.5 to 25 μm . Even more preferably, the specific surface area is within the range of 30 to 60 m^2/g and the mean grain diameter is within the range of 0.7 to 7 μm .

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In the process of the present invention, the surface-reacted natural calcium carbonate is brought into contact with the water to be purified, e.g. industrial waste water, drinking water, urban waste water, waste water from breweries or other beverage industries or water in the paper industry, be it white or waste water, as well as agricultural waste-water by any conventional means known to the skilled person.

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The surface-reacted natural calcium carbonate can be added as an aqueous suspension, e.g. the suspension described above. Alternatively, it can be added in solid form, e.g. in the form of granules or a powder or in the form of a cake. Within the context of the present invention, it is also possible to provide an immobile phase comprising the surface-reacted natural calcium carbonate, the water to be purified running through said immobile phase, e.g. in the form of a cake or layer, comprising

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the surface-reacted natural calcium carbonate, the water to be purified running through said immobile phase. This will be discussed in further detail below.

5 In a preferred embodiment, the pH of the water to be treated is adjusted to a value of greater than 6.0, more preferably greater than 6.5, and even more preferably greater than 7.0 prior to the addition of the surface-reacted calcium carbonate.

10 Preferably, the surface-reacted natural calcium carbonate is suspended in the water, e.g. by agitation means. The amount of surface-reacted natural calcium carbonate depends on the type of water to be treated as well as on the type and amount of impurities. Preferably an amount of 10 ppm - 1 wt.-%, preferably 100 ppm - 0.5 wt.-%, more preferably 400 ppm - 2000 ppm. surface-reacted natural calcium carbonate, based on the weight of the water to be treated, is added.

15 The water may contain organic impurities, e.g. resulting from human waste, organic materials, soil, surfactants as well as inorganic impurities, in particular heavy metal impurities such as iron- or manganese-containing compounds. Harmful components that can be removed from the water with the purification process of the present invention also include microorganism such as bacteria, fungi, archaea, or protists.

20 Further harmful substances such as polycyclic compounds, cholesterol, or endocrine disrupting compounds (EDCs), such as e.g. endogenous hormones such as 17β -estradiol (E2), estrone (E1), estriol (E3), testosterone or dihydro testosterone; phyto and myco hormones such as β -sitosterol, genistein, daidzein or zeraleon; drugs such
25 as 17α -ethinylestradiol (EE2), mestranol (ME), diethylstilbestrol (DES), and industrial chemicals such as 4-nonyl phenol (NP), 4-tert-octyl phenol (OP), bisphenol A (BPA), tributyltin (TBT), methylmercury, phthalates, PAK or PCB.

Furthermore iodised contrast agents for X-ray-analysis e.g. ioxithalamate, fluorochinolones like ciprofoxacin and antibiotics like macrolides and sulfonamides can be removed by the process of the present invention.

- 5 It has also been found that surface-reacted calcium carbonate very efficiently adsorbs activated carbon powder alone as well as activated carbon powder which has already adsorbed substances. As outlined above, activated carbon is a well-known potent adsorbent, having however the drawback that the adsorption very often is not very fast and that its removal from the medium to be treated is very difficult due to its fine
10 division in the medium.

The resulting composites, however, comprising surface-reacted calcium carbonate, activated carbon, and, optionally, further substances like impurities adsorbed to the carbon surface, can be easily separated from the water to be purified.

- 15 The use of surface-reacted calcium carbonate together with activated carbon on the one hand provides for a very efficient adsorption of impurities, such as the substances mentioned above, by the excellent adsorption properties of both of the adsorbents, and on the other hand additionally improves the removability of activated
20 carbon by forming composites therewith, also if it has already substances, e.g. the substances mentioned above as impurities, adsorbed onto its surfaces, resulting in a further improved removal of impurities due to the synergistic interaction between surface-reacted calcium carbonate and activated carbon, the resulting composites being easily separable from the aqueous medium.

- 25 In this respect it is especially preferred that the activated carbon is added to the aqueous medium prior to the addition of the surface-reacted natural calcium carbonate. In this embodiment, the substances are essentially first adsorbed to the activated carbon and the resulting composite is essentially subsequently adsorbed by

the surface-reacted calcium carbonate providing for an improved removability of the activated carbon, and the impurities attached thereto, respectively.

Any activated carbon known in the art can be used in the process of the present invention. Examples of activated carbons useful in the present invention are e.g. available from Fluka, such as article No. 05112 (p.a. for gc; grain size 0.3 to 0.5 mm; bulk density 410 kg/m³), from Aldrich such as article number 484156 (glassy spherical powder, grain size 10 – 40 µm), from Sigma-Aldrich such as article No. 242276 (Darco® G-60, powder, -100 mesh particle size); from Riedel-de Haen (article No. 18002, purum, granulated), or Lurgi Hydrafin CC 8 x 30 (Donau Carbon GmbH & Co. KG, Frankfurt am Main, Germany) or activated carbon available from Fluka (article No. 05112).

For example, the activated carbon particles can have a grain size of 0.1 µm to 5mm, preferably 10 µm to 2 mm, 0.1 mm to 0.5 mm, e.g. 0.3 mm.

Preferably, the weight ratio of the surface-reacted natural calcium carbonate to the activated carbon is from 1:1 to 100:1, more preferably from 5:1 to 80:1, especially 10:1 to 70:1 or 20:1 to 50:1, e.g. 30:1 or 40:1.

Optionally, additives can be added to the water sample to be treated. These might include agents for pH adjustment and conventional flocculants such as polyaluminium chloride, iron chloride or aluminium sulphate.

In a preferred embodiment, a natural calcium carbonate which has not been surface-reacted as described above is added as well.

After the adsorption has been completed, the resulting composite can be separated from the water sample by conventional separation means known to the skilled person such as sedimentation and filtration.

In an alternative approach, the liquid to be purified is preferably passed through a permeable filter comprising the surface-reacted natural calcium carbonate and being capable of retaining, via size exclusion, the impurities on the filter surface as the liquid is passed through by gravity and/or under vacuum and/or under pressure. This process is called "surface filtration".

In another preferred technique known as depth filtration, a filtering aid comprising of a number of tortuous passages of varying diameter and configuration retains impurities by molecular and/or electrical forces adsorbing the impurities onto the surface-reacted natural calcium carbonate which is present within said passages, and/or by size exclusion, retaining the impurity particles if they are too large to pass through the entire filter layer thickness.

The techniques of depth filtration and surface filtration may additionally be combined by locating the depth filtration layer on the surface filter; this configuration presents the advantage that those particles that might otherwise block the surface filter pores are retained in the depth filtration layer.

One option to introduce a depth filtration layer onto the surface filter is to suspend a flocculating aid in the liquid to be filtered, allowing this aid to subsequently decant such that it flocculates all or part of the impurities as it is deposited on a surface filter, thereby forming the depth filtration layer. This is known as an alluvium filtration system. Optionally, an initial layer of the depth filtration material may be pre-coated on the surface filter prior to commencing alluvium filtration.

The following figures, examples and tests will illustrate the present invention, but are not intended to limit the invention in any way.

Description of the figures:

Figure 1 shows the adsorption behaviour of 17 α -ethinylestradiol on SRCC for different equilibration times.

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Figure 2 shows the adsorption isotherms for the adsorption of 17 α -ethinylestradiol on SRCC, activated carbon and manganese dioxide in terms of the adsorbed amount per amount of adsorbent.

10 Figure 3 shows the adsorption isotherms for the adsorption of 17 α -ethinylestradiol on SRCC, activated carbon and manganese dioxide in terms of the adsorbed amount per surface area of adsorbent.

15 Figure 4 shows the turbidity reduction by the adsorption of activated carbon by SRCC

Figure 5 shows a SEM image of activated carbon.

20 Figure 6 shows a SEM image of activated carbon and the SRCC.

Figure 7 shows an enlargement of Figure 6.

EXAMPLES:

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The type and amount of impurities which are present in the water samples before and after treatment according to the process of the present invention were determined using Optima 3200 XL ICP-OES instrumentation from Perkin-Elmer. The samples were directly analysed following treatment with SRCC.

Example 1

In example 1, the adsorption capacity of the surface-reacted natural calcium carbonate with regard to different heavy metals is determined.

500 g of a heavy metal solution, containing 5 ppm each of cadmium, chromium, copper, mercury, nickel, and lead was stirred for 15 minutes with 3% (w/w%) surface-reacted natural calcium carbonate. The solution was left for 24 hours at pH 11.5 and the upper liquid phase was analysed with ion chromatography (Dionex DX 120 Ion-Chromatograph).

The surface-reacted natural calcium carbonate was prepared as follows:

A finely divided natural calcium carbonate originating from Omev, France, was suspended to achieve a suspension of approximately 16 % by weight of dry matter. The slurry thus formed is then treated by slow addition of phosphoric acid at a temperature of approximately 55°C, sufficient to provide a product featuring a BET specific surface area of 35 m²/g according to ISO standard 92777, and an approximate number average diameter of 10 micrometers estimated from scanning electron microscope images obtained using a LEO 435 VPe scanning electron microscope.

Table 1: Treatment of heavy metal ion solution with surface-reacted calcium carbonate

	Amount of cadmium (ppm)	Amount of chromium (ppm)	Amount of copper (ppm)	Amount of mercury (ppm)	Amount of nickel (ppm)	Amount of lead (ppm)
Untreated solution	5.0	5.0	5.0	5.0	5.0	5.0
Solution after treatment	< 0.1	< 0.1	< 0.1	< 0.1	0.5	< 0.1
Sediment obtained after treatment	2.6	2.5	3.2	2.1	3.7	2.6

It is obvious that the surface-reacted natural calcium carbonate was able to reduce cadmium, chromium, copper, mercury, and lead very efficiently. From the originally dissolved 5 ppm, less than 0.1 ppm was recovered in the treated sample. For nickel, 90% of the starting amount of 5 ppm was adsorbed and thus removed from the solution.

10 **Example 2**

In Example 2, the adsorption capacity of the surface-reacted natural calcium carbonate with regard to microorganisms is determined.

- 15 A barm suspension of 100 ppm was used with a germ count of $6 * 10^5$ cfu/cm³. In a first experiment, the barm suspension was filtered through a blue band paper filter. In a second experiment, the suspension was filtered through a layer made of the surface-reacted natural calcium carbonate, the layer having a diameter of 90 mm and a thickness of 30 mm.

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The surface-reacted natural calcium carbonates was prepared as follows:

Two samples were tested and provided the same results within 0.5%.

First surface-reacted natural calcium carbonate:

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A finely divided natural calcium carbonate originating from Omev, France, was suspended to achieve a suspension of approximately 16 % by weight of dry matter. The slurry thus formed is then treated by slow addition of phosphoric acid at a temperature of approximately 55°C, sufficient to provide a product featuring a BET
10 specific surface area of 35 m²/g according to ISO standard 92777, and an approximate number average diameter of 10 micrometers estimated from scanning electron microscope images obtained using a LEO 435 VPe scanning electron microscope.

15 Second surface-reacted natural calcium carbonate:

A finely divided natural calcium carbonate originating from Molde, Norway, was suspended to achieve a suspension of approximately 16 % by weight of dry matter. The slurry thus formed is then treated by slow addition of phosphoric acid at a
20 temperature of approximately 55°C, sufficient to provide a product featuring a BET specific surface area of 50 m²/g according to ISO standard 92777, and an approximate number average diameter of 20 micrometers estimated from scanning electron microscope images obtained using a LEO 435 VPe scanning electron microscope.

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The results are shown in Table 2.

Table 2: Filtration of barm suspension through different filter materials

	Germ count of suspension (cfu/ml)
Untreated suspension	$6 * 10^5$
Suspension after filtering through paper filter	$6 * 10^5$
Suspension after filtering through filter layer made of surface-reacted calcium carbonate	$7 * 10^3$

The results indicate that the use of the surface-reacted natural calcium carbonate reduced the germ count by a factor of 100.

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Example 3

In Example 3, river water was subjected to the purification method of the present invention. 100 ppm surface-reacted natural calcium carbonate and 4 ppm
10 polyaluminium chloride were suspended in a river water sample. After two minutes, the flocculated solids were filtered off and the filtrate was analysed for its content of iron and manganese.

The surface-reacted natural calcium carbonate used in Example 3 was prepared as
15 follows:

A finely divided natural calcium carbonate originating from Vermont, U.S.A, containing 800 ppm of magnesium oxide and 2500 ppm anionic polyacrylate dispersant per equivalent dry gram of calcium carbonate, was suspended to achieve a
20 suspension of approximately 16 % by weight of dry calcium carbonate. The slurry thus formed is then treated by slow addition of phosphoric acid at a temperature of approximately 55°C, sufficient to provide a product featuring a BET specific surface area of 68 m²/g according to ISO standard 92777, and a d50 of 10 micrometers measured by means of the Sedigraph™ 5100 from Micromeritics™ .

The results are summarized in Table 3.

Table 3: Treatment of river water with surface-reacted calcium carbonate

	Amount of iron (mg/l)	Amount of manganese (mg/l)
River water sample before treatment	2.03	0.88
River water sample after treatment	0.023	0.104

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The results clearly indicate that treatment with the surface-reacted natural calcium carbonate significantly reduces the amounts of heavy metal ions such as iron and manganese.

10 **Example 4**

A suspension of approximately 25 wt.-% based on the dry weight of finely divided natural calcium carbonate originating from Omev, France, was prepared. The slurry thus formed was then treated by slow addition of phosphoric acid at a temperature of approximately 55°C.

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The resulting slurry had a BET specific surface area of 60 m²/g according to ISO standard 92777, and a d₅₀ of about 7 µm measured by means of the Sedigraph™ 5100 from Micromeritics™.

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1. Adsorption of 17α-ethinylestradiol

17α-ethinylestradiol (EE2) (> 98% pure) was used, which is supplied by Sigma Aldrich (Bornem, Belgium) (article No. E4876).

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a) Preparation of the test mixture

A standard experiment was performed using the following procedure:

- 5 0.4 g SRCC suspension having a solids content of 25 Gew.-% was added to 7 ml of
17 α -ethinylestradiol solution having different concentrations of 50, 100, 200, 500
and 1000 μ g/l and shaken for 30 min., 1 h, 2 h and 24 h at 25 °C.

*b) Determination of the equilibrium concentration in the adsorption of 17 α -
10 ethinylestradiol onto the surface of surface-reacted calcium carbonate*

- For the determination of the equilibrium concentration of 17 α -ethinylestradiol with
respect to the adsorption onto the surface of SRCC, the initial and final concentration
after the addition of SRCC was measured at certain concentrations and treatment
15 (shaking) times. It can be taken from figure 1 that the adsorption of 17 α -
ethinylestradiol onto the surface of SRCC was very fast and no change was observed
after 30 minutes for any concentration, which indicates that the adsorption
equilibrium is reached after 30 minutes, independently from the concentration.

- 20 The concentration was determined by the well-known method of HPLC (Column:
Gemini 5 μ m C18 / ODS C18; Eluent A: Water (45 %); Eluent B: Acetonitrile 55 %;
Eluent Type: Isocratic; Flow 1ml/min; Temp. 25 °C; Detection: UV-Vis Absorption
@ 205 nm)

*c) Determination of adsorption isotherms of 17 α -ethinylestradiol for several
25 adsorbents*

For determining the efficiency of 17 α -ethinylestradiol adsorption the adsorption
isotherms were determined.

After the equilibrium time, the SRCC is separated from the liquid phase by settling, filtration or centrifugation. The concentration in the upper liquid phase is determined by the above described HPLC method, resulting in the equilibrium concentration. The lower solid phase is isolated by decantation of the supernatant phase. The SRCC
5 is dissolved in a certain amount of hydrochloric acid and the concentration measured with the above mentioned method, resulting in the adsorbed amount per amount SRCC.

Based on the findings above, an equilibrium time of 1 h was used for the
10 determination of the adsorption isotherm of 17α -ethinylestradiol with respect to SRCC.

Furthermore, the adsorption isotherm for activated carbon acting as the adsorbent was determined. The data for the activated carbon were measured after 16 hours. The
15 activated carbon used was the commercially available Lurgi Hydrafin CC 8 x 30 (Donau Carbon GmbH & Co. KG, Frankfurt am Main, Germany). It has a grain size of 0.5 to 2.5 mm and a density of $480 \pm 50 \text{ kg/m}^3$.

From figure 2, it can be taken that in terms of the adsorbed amount of 17α -
20 ethinylestradiol per amount of adsorbent in ng/g, the activated carbon is superior in 17α -ethinylestradiol adsorption.

However, in terms of adsorbed amount of 17α -ethinylestradiol per surface area of adsorbent as given in figure 3, it can be seen that SRCC shows the same level of
25 efficiency in 17α -ethinylestradiol adsorption, which is markedly higher than that of activated carbon.

These findings show that less activated carbon in terms of weight amount might be needed for adsorbing 17α -ethinylestradiol, but that SRCC is more effective in terms

of its surface area, i.e. although it has a lower specific surface area, more estrogen per surface area is adsorbed.

Furthermore, it was found that the use of SRCC together with activated carbon
5 shows an unexpected synergy.

2. Adsorption of activated carbon

As mentioned in the description, activated carbon cannot be separated easily from
10 water. However, surface-reacted calcium carbonate is able to clarify an activated carbon suspension nicely simplifying the separation of the active carbon.

Activated carbon, which is commercially available from Fluka (article No. 05112 (p.a. for gc; grain size 0.3 to 0.5 mm; bulk density 410 kg/m³) and the SRCC were
15 subjected to the following treatment:

0.02 g activated carbon was added to 30 g water. Then, 0.2g SRCC was added, and the resulting mixture was shaken for 2 min. The mixture was allowed to settle. Subsequently, the upper phase was separated by decanting.

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From Figure 4, it can be taken that turbidity of the upper liquid phase could be significantly reduced. The turbidity of pure activated carbon was constantly above 1000 NTU which is the maximal value which can be measured with this device (a Hach 2100P Iso Turbidimeter).

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The resulting composites can be easily separated, e.g. by filtration.

Figure 5 shows a SEM image of activated carbon.

Figure 6 shows a SEM image of activated carbon and the SRCC. Figure 7 shows an enlargement of Figure 6.

Claims

1. A process for the purification of water, wherein surface-reacted natural calcium carbonate or an aqueous suspension comprising surface-reacted natural calcium carbonate and having a pH of greater than 6.0, measured at 20°C, is added to the water to be purified, wherein the surface-reacted natural calcium carbonate is the reaction product of a natural calcium carbonate, carbon dioxide and one or more acids.
2. A process according to claim 1, wherein the surface-reacted natural calcium carbonate is added in an amount of 10 ppm - 1 wt.-%, preferably 100 ppm - 0.5 wt.-%, more preferably 400 ppm - 2000 ppm, based on the weight of the water to be treated.
3. A process according to claim 1, wherein the pH of the water to be purified is adjusted to a value of > 6.5, preferably > 6.5, more preferably > 7 before being brought into contact with the surface-reacted natural calcium carbonate.
4. A process according to claim 1, wherein the water to be purified is brought into contact with the surface-reacted natural calcium carbonate by surface filtration, depth filtration and/or alluvium filtration.
5. A process according to claim 1, wherein the water to be purified is selected from the group comprising industrial waste water, drinking water, urban waste water, waste water from breweries or other beverage industries, water in the paper industry and agricultural waste-water.
6. A process according to any of the preceding claims, wherein the water to be purified contains heavy metal impurities.
7. A process according to any of the preceding claims, wherein the water to be purified contains microorganisms as impurities.
8. A process according to any of the preceding claims, wherein the water to be purified contains EDC impurities being selected from the group comprising endogenous hormones such as 17 β -estradiol (E2), estrone (E1), estriol (E3), testosterone or dihydro testosterone; phyto and myco hormones such as β -

sitosterol, genistein, daidzein or zeraleon; drugs such as 17 α - ethinylestradiol (EE2), mestranol (ME), diethylstilbestrol (DES), and industrial chemicals such as 4-nonyl phenol (NP), 4-tert-octyl phenol (OP), bisphenol A (BPA), tributyltin (TBT), methylmercury, phthalates, PAK or PCB.

9. A process according to any of the preceding claims, wherein the water to be purified contains activated carbon impurities.

10. A process according to any of the preceding claims, wherein the activated carbon impurities have further impurities adsorbed onto their surface.

11. A composite of surface-reacted calcium carbonate for removing impurities from water, wherein the surface-reacted natural calcium carbonate is a reaction product of a natural calcium carbonate carbon dioxide and one or more acids and the impurities removed from the water.

12. A composite of surface-reacted calcium carbonate according to claim 11, wherein the surface-reacted natural calcium carbonate is prepared as an aqueous suspension having a pH of greater than 6.5, preferably greater than 7.0 and most preferably 7.5, measured at 20°C.

13. A composite of surface-reacted calcium carbonate according to any of the preceding claims, wherein the natural calcium carbonate is selected from the group comprising marble, calcite, chalk, dolomite, limestone and mixtures thereof.

14. A composite of surface-reacted calcium carbonate according to any of the preceding claims, wherein the acid has a pKa at 25°C of 2.5 or less.

15. A composite of surface-reacted calcium carbonate according to claim 11, wherein the acids are selected from the group comprising hydrochloric acid, sulphuric acid, sulphurous acid, hydrosulphate, phosphoric acid, oxalic acid and mixtures thereof.

16. A composite of surface-reacted calcium carbonate according to claim 11, wherein the natural calcium carbonate is reacted with the acid and/or the carbon dioxide in the presence of at least one silicate and/or silica, aluminium hydroxide, earth alkali metal aluminate, magnesium oxide, or mixtures thereof.

17. A composite of surface-reacted calcium carbonate according to claim 16, wherein at least one silicate is selected from an aluminium silicate, a calcium silicate, or an alkali metal silicate.

18. A composite of surface-reacted calcium carbonate according to any of the preceding claims, wherein the surface-reacted natural calcium carbonate has a specific surface area of from 5 m²/g to 200 m²/g, preferably from 20 m²/g to 80 m²/g, measured using nitrogen and the BET method according to ISO 9277.

19. A composite of surface-reacted calcium carbonate according to any of the preceding claims, wherein the surface-reacted natural calcium carbonate has a mean grain diameter d₅₀ of from 0.1 to 50 µm, preferably from 1 to 10 µm, measured according to the sedimentation method.

20. A composite of surface-reacted calcium carbonate according to any of the preceding claims, wherein the aqueous suspension comprising surface-reacted natural calcium carbonate is stabilised with one or more dispersant.

21. A composite of surface-reacted calcium carbonate according to any of the preceding claims, wherein the surface-reacted natural calcium carbonate is used in powder form and/or in the form of granules.

22. Use of the surface-reacted natural calcium carbonate as defined in one of the claims 1 to 20 for water purification.