Claims

1. A process of producing tin(II) alkoxides comprising:

reacting anhydrous tin (II) chloride with n-alcohol and diethylamine in the presence of one or more non-polar aprotic solvents at a temperature of 25^0C to 38^0C to form a reaction mixture containing tin(II) alkoxide, wherein the diethylamine is added to react with a pre-mixture of anhydrous tin (II) chloride and non-polar aprotic solvent at a temperature of 15-20^0C for 3 to 6 hours prior to adding the n-alcohol to form the mixture;

filtering the reaction mixture to acquire a filtrate thereof; and

drying the filtrate to obtain tin(II) alkoxides in a liquid preparation,

wherein the obtained tin (II) alkoxide is represented by a formula $Sn(OnR)_2$ where R is selected from the group consisting of CH_3, C_2H_5, nC_3H_7, nC_4H_9, nC_5H_{11}, nC_6H_{13}, nC_7H_{15}, or nC_8H_{17},

wherein the n-alcohol is in 2.0-2.1 mole equivalence of the anhydrous tin(II) chloride;

wherein the diethylamine is in 2.0-2.1 mole equivalence of the anhydrous tin(II) chloride.

2. The process of claim 1, wherein the n-alcohol is butanol, hexanol or octanol.

3. The process of claim 1, wherein the non-polar aprotic solvent is one of *n*-heptane, *n*-hexane, cyclohexane, benzene, toluene, xylene, and tetrahydrofuran.

4. The process of claim 1, wherein the drying step is performed under vacuum.

5. The process of claim 1, wherein the tin alkoxides in the liquid form are substantially free from self-aggregation resulting in solid precipitation thereof.

6. The process of claim 1, wherein the reacting step is performed under inert atmosphere and the reaction mixture is subjected to continuous stirring.

7. The process of claim 1, wherein the tin(II) alkoxides prepared in the liquid form are fully solubilized in chloroform, toluene or n-heptane at room temperature.

8. The process of claim 1, wherein the non-polar aprotic solvent impedes self-aggregation of the formed tin (II) alkoxides.