

METHOD FOR FORMING A TITANIUM-CONTAINING LAYER ON A SUBSTRATE USING AN ATOMIC LAYER DEPOSITION (ALD) PROCESS

The invention relates to a method for forming a titanium-containing layer on a substrate using an ALD (Atomic Layer Deposition) process.

The new requirements in terms of dimensions for the future generation semiconductor devices, lead to the development of new materials with a high dielectric constant. High-k dielectrics materials are required in CMOS device to replace SiO_2 which reaches its physical limits with a typical thickness of 1 nm. Similarly, high-k dielectrics are required in Metal-Insulator-Metal architectures for DRAM applications. Various metal compositions have been considered to fulfil both the materials requirements (k, leakage current, crystallization temperature, charge trapping) and the integration requirements (for example, thermal stability on interface or dry etching feasibility). Titanium based oxides are among the most promising candidates, as for instance a Ti containing mixed oxide film, TiO_2 , SrTiO_3 , $(\text{Ba},\text{Sr})\text{TiO}_3$, $\text{Pb}(\text{Ti})\text{O}_3$, $\text{Ti}(\text{OiPr})_4$, TiLnO_x (Ln being selected among lanthanide and rare-earth elements) and more generally TiMO_x , M being an element selected from group II, group IIIa and IIIb, or a transition metal.

Furthermore, hafnium, zirconium and titanium metals composition can also be considered for electrode and/or Cu diffusion barrier applications, such as TiN , HfN , ZrN , HfSi , ZrSi , HfSiN , ZrSiN for MIM electrodes.

Vapor phase deposition is the main industrial way to deposit thin films with reasonable throughput and acceptable purity. Those techniques could be MOCVD (Metal-organic Chemical Vapor Deposition) or ALD (Atomic Layer Deposition). Metal-organic or metal-halide precursors are required for those processes. Various hafnium, zirconium and titanium metal-organic compounds have been considered as precursors to enable such deposition.

Halides such as HfCl_4 , ZrCl_4 are commonly used as Hf or Zr precursors and have been widely described. Triyoso et al. [Journal of Applied Physics 97 (2005) 124107], Kawahara et al. [Japanese Journal of Applied Physics 43 (2004) 4129], Caymax et al. [Material Research Society symp proc Vol 765 (2003) 47], evaluated HfCl_4 for the deposition of HfO_2 by ALD. The main disadvantage is the generation

during the deposition process of some by-products such as HCl or Cl₂. Those compounds can cause surface/interface roughness that can be detrimental to the final properties. Other possible by-products, depending on the O-source, might be hazardous. Moreover, in the case of high-k oxide, Cl or F impurities are highly detrimental to the final electrical properties.

Alkoxides, such as Hf(OtBu)₄, Zr(OtBu)₄, Hf(OtBu)₂(mmp)₂ and Hf(mmp)₄, (mmp = 1-methoxy-2-methyl-2-propanolate), have been evaluated, for instance by Kukli et al. [Chem. Mater. 15 (2003) 1722], and Potter et al. [Chem. Vap. Dep. 11 (2005) 159]. However, as suggested by Potter et al., factors like β-hydride elimination of the alkoxide ligand may prevent the ALD self-limitation. A direct comparison between Hf(OtBu)₄ and Hf(NEtMe)₄ (Senzaki et al. [J. Vac. Sc. Techn. A 22 (2004) 1175]) showed much lower impurity (C, H) content in the oxide film for the latter precursor.

Alkylamides precursors (Hf(EMA)₄, Hf(DMA)₄, Hf(DEA)₄) have been widely studied. The deposition of HfO₂ from TEMAH by ALD was studied by Senzaki et al. [J. Vac. Sc. Techn. A 22 (2004) 1175] Kawahara et al. [JJAP 43 (2004) 4129] and others. Those alkylamides are both suitable for ALD and MOCVD processes and present interesting properties for distribution since they are liquid at room temperature (TDEAH and TEMAH) and of sufficient volatility. Furthermore, they allow self-limited ALD at low temperature for limited thermal budget processes. The drawback of the Group IV alkylamide precursors is their limited thermal stability and in particular zirconium compounds and so, they may decompose slightly during the distribution, can generate particles during deposition, may entail un-uniform composition during deep trenches deposition processes and they do not allow a large self-limited ALD temperature window, hence reducing the process window. In particular, Zr(NEtMe)₄ may decompose in the distribution lines and generate particles above 170°C which is a common distribution temperature. Hf(NEtMe)₄ is more thermally stable yet do not afford self-limited atomic layer deposition above 300°C due to thermal decomposition.

In WO2007/055088, Thenappan et al. disclose hafnium and zirconium guanidines complexes and their application for vapor phase deposition. Hf(NEt₂)₂((NiPr-CNEt₂)₂) is given as example. Hafnium and zirconium guanidines are however generally solids with very limited volatility. As exemplified in WO2007/055088 thermal gravimetric analysis, one may not obtain Hf(NEt₂)₂((NiPr-

CNEt₂)₂ in vapour phase without risk of thermal decomposition and subsequent particle generation.

Lehn et al. [CVD 12 (2006) 280] have introduced new tetrakis(trimethylhydrazido)zirconium and hafnium Zr(NMeNMe₂)₄ and their use for low temperature CVD. The exemplified compounds have acceptable volatility (sublimation at 0.06 Torr, 90°C reported) but present the disadvantage of being solid at room temperature.

Carta et al. (Electrochem Soc Proceedings, 260, 2005-09, 2005) have introduced a new family of Zr and Hf compounds as alternative to hafnium and zirconium alkylamides: Bis(cyclopentadienyl)bisdimethyl hafnium, biscyclopentadienylbisdimethylzirconium. They allow efficient ALD with an ALD window up to 400°C and achievement of film with less than 0.2% C in optimized conditions with H₂O as co-reactant. However, HfCp₂Me₂ and ZrCp₂Me₂ are both solid at room temperature. HfCp₂Me₂ melting point is 57.5°C. This prevents IC makers to use those precursors in an industrial manner, that is using delocalized containers filling, and entail both facilitation and process issues. Liquid dicyclopentadienyl derivatives have recently been proposed by Heys et al. in WO2006/131751 A1. However, they still present the disadvantage of limited volatility and also present large steric hindrance that may limit the achieved growth rate.

Heys et al. in WO2006/131751A1 have introduced the bis(cyclopentadienyl)bis(alkoxyde) hafnium and zirconium family. However, HfCp₂(OMe)₂ and HfCp₂(OMe)₂ has the major drawback of being solid. They present the disadvantage to be thermally unstable. A thermogravimetric analysis shows 20 % of residues for an end of vaporization at 310°C. The second drawback for a using in an industrial manner is their solid state.

The main industrial options to enable the deposition of highly uniform and conformal thin films with reasonable throughput in high aspect ratio structures are techniques such as MOCVD (Metal-Organic Chemical Vapor Deposition) or ALD (Atomic Layer Deposition).

However, films deposited by MOCVD need high thermal budget and generally follow a 3D-growth mechanism described by a Volmer-Weber model. Thin films grow by clusters nucleation and such technique also leads to insufficient step coverage.

The typical ALD process (e.g as described in RITALA M., LESKELA M., Atomic Layer Deposition, Handbook of thin films materials) involves gaseous reactants led onto a substrate by pulses, separated by inert gas purging. In MOCVD, gaseous reactants are injected simultaneously and react by thermal self-decomposition while in ALD; the loss of the ligand is thermally induced by reaction with the surface groups on the substrate. In a temperature range, the surface reactions are self-limited, which allow the deposition of highly uniform and conformal films. Precursors must be volatile and stable enough to be easily transferred to the reaction chamber without being decomposed. Moreover, they must be reactive enough with the chemical groups of the surface to ensure reasonable growth rate.

However, common group IV or V (Ti, V, Nb, Ta) based metal-organic precursors are not suitable for the deposition without assisting the thermal ALD process by plasma techniques.

According to the present invention, surprisingly, it appears that in applying Atomic Layer Deposition techniques to a molecule which could have been first proposed for Chemical Vapor Deposition, allows to get very good TiO_2 films and solves the inconvenients cited above.

Moreover, using such metal-organic precursors by MOCVD at 600°C , can not be suitable for the deposition of highly uniform, conformal thin films. Such deposition will result in an excessive carbon contamination and poor step coverage in high aspect ratio structures.

Nevertheless, according to the state of the art, MOCVD process is the best solution in terms of interesting growth rate, so others processes have never been implemented.

According to a first embodiment, the invention relates to a method for forming a titanium-containing layer on a substrate, the method comprising at least the steps of:

a) providing a vapor comprising at least one precursor compound of the formula $\text{Ti}(\text{Me}_5\text{Cp})(\text{OR})_3$ (I), wherein R is selected in the group consisting in methyl, ethyl, isopropyl; or of the formula $\text{Ti}(\text{R}^1\text{Cp})(\text{OR}^2)_3$ (II), wherein R^1 is selected from the group consisting in H, methyl, ethyl, isopropyl and R^2 is independently selected from the group consisting in methyl, ethyl, isopropyl or tert-butyl;

b) reacting the vapor comprising the at least one compound of formula (I) or (II) with the substrate, according to an atomic layer deposition process, to form a layer of a tantalum-containing complex on at least one surface of said substrate.

Furthermore, other embodiments of the invention are:

5 - The method according to the invention further comprising the step:

c) reaction of the complex formed obtained in step b) with a reagent selected from another metal source, reducing reactants and/or nitriding reactants and/or oxidizing reactants.

10 - The method wherein the vapour provided in step a) further comprises one or more metal (M')-organic precursor(s) to produce thin films containing titanium and M'. M' being independently selected from any other element in the group II, III-A, III-B, Sulphur (S), transition metal, lanthanoids, or rare-earth metals.

15 - The method, further comprising providing at least one reaction gas wherein the at least one reaction gas is selected from the group consisting of hydrogen, hydrogen sulfide, hydrogen selenide, hydrogen telluride, carbon monoxide, ammonia, organic amine, silane, disilane, higher silanes, silylamines, diborane, hydrazine, methylhydrazine, chlorosilane and chloropolysilane, metal alkyl, arsine, phosphine, trialkylboron, oxygen, ozone, water, hydrogen peroxide, nitrous oxide, nitrogen monoxide, nitrogen dioxide, alcohols, plasma comprising
20 fragments of those species, and combinations thereof, preferably ozone or water.

- The method, wherein the titanium precursor of formula (I) or (II) is selected in the group consisting of:

25 $\text{Ti}(\text{MeCp})(\text{OEt})_3$; $\text{Ti}(\text{MeCp})(\text{OiPr})_3$; $\text{Ti}(\text{MeCp})(\text{OtBu})_3$; $\text{Ti}(\text{EtCp})(\text{OEt})_3$; $\text{Ti}(\text{EtCp})(\text{OiPr})_3$; $\text{Ti}(\text{EtCp})(\text{OtBu})_3$; $\text{Ti}(\text{Cp})(\text{OEt})_3$; $\text{Ti}(\text{Cp})(\text{OiPr})_3$; $\text{Ti}(\text{Cp})(\text{OtBu})_3$; $\text{Ti}(\text{Me}_5\text{Cp})(\text{OMe})_3$; $\text{Ti}(\text{MeCp})(\text{OMe})_3$; $\text{Ti}(\text{MeCp})(\text{OEt})_3$; $\text{Ti}(\text{MeCp})(\text{OiPr})_3$; $\text{Ti}(\text{MeCp})(\text{OtBu})_3$; $\text{Ti}(\text{EtCp})(\text{OMe})_3$; $\text{Ti}(\text{EtCp})(\text{OEt})_3$; $\text{Ti}(\text{EtCp})(\text{OiPr})_3$; $\text{Ti}(\text{EtCp})(\text{OtBu})_3$; $\text{Ti}(\text{Cp})(\text{OMe})_3$; $\text{Ti}(\text{Cp})(\text{OEt})_3$; $\text{Ti}(\text{Cp})(\text{OiPr})_3$; $\text{Ti}(\text{Cp})(\text{OtBu})_3$; $\text{Ti}(\text{iPrCp})(\text{OMe})_3$; $\text{Ti}(\text{iPrCp})(\text{OEt})_3$; $\text{Ti}(\text{iPrCp})(\text{OiPr})_3$; $\text{Ti}(\text{iPrCp})(\text{OtBu})_3$, preferably $\text{Ti}(\text{Me}_5\text{Cp})(\text{OMe})_3$.

30 - The method, wherein the temperature of the substrate is 25 °C to 600 °C, preferably 380°C to 425°C, and wherein the atomic layer deposition chamber containing the substrate has a pressure of 0.133 Pa to 133 kPa, preferably below 27 kPa.

- The method, further comprising the step of purging excess vapor comprising the at least one compound of formula (I) from the substrate, with an inert gas selected from the group consisting of hydrogen, nitrogen, helium, argon, and mixtures thereof.

- 5 - A method of manufacturing a semiconductor structure, comprising the steps of the method defined in the present invention, wherein the substrate is a semiconductor substrate.

A new technique to deposit titanium containing films using metal-organic precursors described by the following general formula

10 **Ti(Me₅Cp)(OR)₃ as material source.**

The vaporization of the titanium source is realized by introducing a carrier gas into a heated container containing the said metal source. The container is preferably heated at a temperature allowing to get the said metal source at a sufficient vapor pressure. The carrier gas can be selected from Ar, He, H₂, N₂ or mixtures of them. The said titanium source can be mixed to a solvent or to another metal source or to a mixture of them in the container. The container can for instance be heated at temperatures in the range of 25°C-300°C, preferably below 150°C. Those skilled in the art will consider that the temperature of the container can be adjusted to control the amount of precursor vaporized. To control the evaporation level in the container, the pressure in the container can be modified. By reducing the pressure in the container, the level of vaporation of the titanium source can be increased. The pressure in the container can for instance be changed in the range of 0.133 Pa until 133 kPa, preferably below 27 kPa.

The said titanium source can also be fed in liquid state to a vaporizer where it is vaporized. The said metal source can be mixed to a solvent. The said titanium source can be mixed to another metal source. The said mixture of metal sources can be mixed to a solvent or a mixture of solvent. The said titanium source can be mixed to a stabilizer. The said solvent can be selected in the group consisting of alkanes such as hexane, heptane, octane, aromatic solvents such as benzene, toluene, mesitylene, xylene, silicon containing solvent such as hexamethyldisiloxane, hexamethyldisilazane, tetramethylsilane, sulphur containing solvents such as dimethylsulfoxide, oxygen containing solvent such as tetrahydrofuran, dioxane.

The said vaporized titanium source is then introduced into a reaction chamber where it is contacted to the surface of a substrate. The substrate can be heated to sufficient temperature to obtain the desired film at sufficient growth rate and with desired physical state and composition. Typical temperatures range from 150°C to 600°C. Preferably the temperature is lower or equal to 450°C. The process can be assisted by a plasma technique. The use of plasma techniques allows to ionize the precursor vapor and/or the reactant used to form radicals and thus improve the reactivity of the said vaporized metal source and/or the reactivity of other gaseous species, used in the process.

The method of the present invention is a method assisted or not by plasma techniques, for depositing a titanium-containing film on the surface of a substrate comprising at least: exposition of the surface of a substrate to a titanium precursor described by the following general formula (I) or (II) and a reagent selected from reducing reactants, oxidizing reactants and/or nitriding reactants, to produce a titanium containing thin film on the surface. Example of reagent can be selected in the list: H_2 , N_2H_2 , methylhydrazine, NH_3 , SiH_4 , Si_2H_6 , Si_3H_8 , TSA, Si_2Cl_6 or any chlorosilane or chloropolysilane, trimethylaluminium, $ZnEt_2$ or any metal alkyl, BH_3 , B_2H_6 , PH_3 , AsH_3 , trimethylboron, triethylboron, CO, monoamines, diamines, mixtures of them or plasma comprising fragment of those species. Example of oxidizing reagents can be selected in the list: O_2 , O_3 , H_2O , H_2O_2 , NO, NO_2 , N_2O , CH_3OH or any alcohol, mixture of them or plasma comprising fragments of those species. Oxidizing species may alternatively be metal-organic compounds containing a metal-oxygen bond.

The deposition method, of this invention improves upon known methods by allowing the use of lower temperatures and producing higher quality films

Furthermore, the thermal stability is improved: the ALD process will be provided with a higher temperature.

Another advantage is that the growth rate reaches a good level with the process of the invention: $0.6 \text{ Angstrom.cycle}^{-1}$.

In one embodiment, the method of the invention consists in introducing alternatively a titanium-organic precursor described by the general formula (I) or (II) into a reaction chamber with a reagent. In a temperature range which depends on the precursor thermal stability and physical properties, the said titanium-organic precursor reacts in a self-limited manner with the chemical bonds present onto the

surface of a substrate, chosen without limitation. Preferably, un-deposited titanium-organic precursors molecules are removed from the reaction chamber. The reagent introduced, reacts also in a self-limited manner.

Once all the complexes present on the surface of the substrate have
5 reacted with the reagent, species are removed from the reaction chamber by a
purge gas. The purge gas can for instance be selected within N₂, Ar, He, H₂
mixtures of them. The purge gas may additionally contain other gas species that
do not modify the chemical reactivity of the surface. Alternatively, the purge can be
realized by vacuum. This process can be repeated as many times as necessary to
10 reach the desired film thickness. The reagent is selected from reducing reactants,
nitriding reactants, oxidizing reactants, or a mixture of them. Example of reagent
can be selected in the list: H₂, N₂H₂, methylhydrazine, NH₃, SiH₄, Si₂H₆, Si₃H₈,
TSA, Si₂Cl₆ or any chlorosilane or chloropolysilane, trimethylaluminium, ZnEt₂ or
any metal alkyl, BH₃, B₂H₆, PH₃, AsH₃, trimethylboron, triethylboron, CO,
15 monoamines, diamines, mixtures of them or plasma comprising fragment of those
species. Example of oxidizing reagents can be selected in the list: O₂, O₃, H₂O,
H₂O₂, NO, NO₂, N₂O, CH₃OH or any alcohol, mixture of them or plasma
comprising fragments of those species, preferably plasma comprising H₂, NH₃ or
O₂. Oxidizing species may alternatively be metal-organic compounds containing a
20 metal-oxygen bond.

In one embodiment, the method of the invention consists in introducing
alternatively first a titanium-organic precursor described by the general formula (I)
or (II) into a reaction chamber and second a reagent or another metal source. This
another metal source being independently selected from any other element in the
25 group II, III-A, III-B, Sulphur (S), transition metal, lanthanoids, or rare-earth metals.
In a temperature range which depends on the precursor thermal stability and
physical properties, the said metal-organic precursors react in a self-limited
manner with the chemical bonds present onto the surface of a substrate.
Preferably, un-deposited metal-organic precursors molecules are removed from
30 the reaction chamber. The reagent introduced, reacts also in a self-limited manner.

In another embodiment of the invention, the said metal source is fed in
liquid state to a vaporizer where it is vaporized. The said metal source can be
mixed to another metal source. The said mixture of metal sources can be mixed to
a solvent or a mixture of solvent. The said metal source can be mixed to a

stabilizer. The said solvent can be selected for example from the group consisting of octane, hexane, pentane, tetramethylsilane.

The said vaporized metal source is introduced into a reaction chamber where it is contacted to a substrate. The substrate can be selected from the group consisting of Si, SiO₂, SiN, SiON, W containing film and other metal containing
5 films. The substrate can be heated to sufficient temperature to obtain the desired film at sufficient growth rate and with desired physical state and composition. Typical temperature ranges from 150°C to 600°C. Preferably the temperature is lower or equal to 450°C. The pressure in the reaction chamber is controlled to
10 obtain the desired metal containing film at sufficient growth rate. Typical pressure ranges from 1m Torr (133.10^{-3} Pa) level to 100 Torr (133.10^2 Pa) or higher.

In one embodiment of the invention, the said metal source is mixed to a reactant species prior to the reaction chamber.

In one embodiment of the invention where the targeted metal based film
15 contains oxygen, such as for example metal oxide or metal oxy-nitride, the said reactant species include an oxygen source which is selected from, but not limited to, oxygen (O₂), oxygen radicals (for instance O or OH), for instance generated by a remote plasma, ozone (O₃), NO, N₂O, NO₂, moisture (H₂O) and H₂O₂.

In one embodiment of the invention where the targeted metal based film
20 contains nitrogen, such as for example metal nitride or metal carbo-nitride, the said reactant species include a nitrogen source which is selected from, but not limited to nitrogen (N₂), ammonia, hydrazine and alkyl derivatives, N-containing radicals (for instance N, NH, NH₂), NO, N₂O, NO₂, amines.

In one embodiment of the invention where the targeted metal based film
25 contains carbon, such as for example metal carbide or metal carbo-nitride, the said reactant species include a carbon source which is selected from, but not limited to, methane, ethane, propane, butane, ethylene, propylene, t-butylene, isobutylene, CCl₄.

In one embodiment of the invention where the targeted metal based film
30 contains silicon, such as for example metal silicide, silico-nitride, silicate, silico-carbo-nitride, the said reactant species include a silicon source which is selected from, but not limited to, SiH₄, Si₂H₆, Si₃H₈, TriDMAS, BDMAS, BDEAS, TDEAS, TDMAS, TEMAS, (SiH₃)₃N, (SiH₃)₂O, trisilylamine, disiloxane, trisilylamine, disilane, trisilane, a alkoxysilane SiH_x(OR³)_{4-x}, a silanol Si(OH)_x(OR³)_{4-x}; preferably

Si(OH)(OR³)₃ ; more preferably Si(OH)(OtBu)₃, an aminosilane SiH_x(NR³R⁴)_{4-x} (where x is comprised between 0 and 4; R³ and R⁴ are independently H or a C1-C6 carbon chain, either linear, branched or cyclic); preferably TriDMAS SiH(NMe₂)₃, BTBAS SiH₂(NHtBu)₂, BDEAS SiH₂(NEt₂)₂ and mixtures thereof. The
 5 targeted film can alternatively contain Germanium. The above-mentioned Si containing sources could be replaced by Ge containing sources.

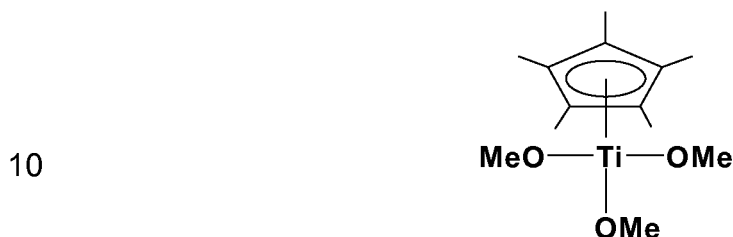
The said first metal source, the possible said second metal source and the reactant species are introduced sequentially in the reaction chamber (atomic layer deposition). The reactor pressure is selected in the range from 1mTorr (133.10⁻³
 10 Pa) to 100 Torr (133.10² Pa). Preferably, the reactor pressure is comprised between 1 and 10 Torr (1330 Pa). A purge gas is introduced between the metal source pulse and the reactant species pulse. The purge gas can be selected from the group consisting of N₂, Ar, He. The metal source, purge gas and reactant species pulse duration is comprised between 0.1 and 100s. Preferably the pulse
 15 duration is comprised between 0.5 and 10s.

In one embodiment of the invention, the second metal source is a lanthanide and rare-earth metal source (Sc, Y, La, Ce, Pr, Nd, Gd...) source and is selected, but not limited to, from the group consisting of rare earth diketonates Ln(-O-C(R¹)-C(R²)-C(R³)-O-)(-O-C(R⁴)-C(R⁵)-C(R⁶)-O-)(-O-C(R⁷)-C(R⁸)-C(R⁹)-O-)
 20 (where each Rⁱ is independently H or a C1-C6 carbon chain, either linear, branched or cyclic), a cyclopentadienyl Ln(R¹Cp)(R²Cp)(R³Cp) (where each Rⁱ is independently H or a C1-C6 carbon chain, either linear, branched or cyclic), Ln(NR¹R²)(NR³R⁴)(NR⁵R⁶) and mixtures thereof. The other metal source can alternatively be an aluminum source and is selected from, but not limited to, the
 25 group consisting of trimethylaluminum, dimethylaluminum hydride, an alkoxyalane AlRⁱ_x(OR^j)_{3-x} (where x is comprised between 0 and 2; R¹ and R² are independently H or a C1-C6 carbon chain, either linear, branched or cyclic); preferably AlR¹R²OR^j, most preferably AlMe₂(OiPr), an amidoalane AlRⁱ_x(NR^jR^k)_{3-x} (where x is comprised between 0 and 2; R¹ and R² are independently H or a C1-C6 carbon
 30 chain, either linear, branched or cyclic) and mixtures thereof. The other metal source can alternatively be a tungsten or molybdenum source. The other metal source can be a titanium source such as Ti(OR¹)₄ or other alkoxide-containing metal sources, M(NR¹R²)₄, or adducts containing these species. The second metal source can alternatively be a divalent metal source (preferably Sr, Ba, Mg, Ca, and

Zn) selected from, but not limited to metal β -diketonates, cyclopentadienyl or adducts containing these species.

EXAMPLES

5 **Synthesis of pentamethylcyclopentadienyl trimethoxy titanium**
Ti(Me₅Cp)(OMe)₃:



The synthesis is carried out. A yellow liquid is obtained. The purified compound is analyzed by NMR. A portion is analyzed by TGA. 50% of the mass is evaporated at 205°C and the molecule is fully evaporated at 230°C, with a negligible amount of residues (cf figure 1). It shows the excellent stability of the molecule until the full evaporation temperature and the much higher volatility of this molecule compared to Ti(OMe)₄.

20 **Synthesis of methylcyclopentadienyl trimethoxy titanium**
Ti(MeCp)(OMe)₃:



The synthesis is carried out. A light yellow liquid is obtained. The purified compound is analyzed by NMR. A portion is analyzed by TGA. 50% of the mass is evaporated at 196°C and the molecule is fully evaporated at 218°C, with a negligible amount of residues (cf figure 2). It shows the excellent stability of the molecule until the full evaporation temperature and the much higher volatility of this molecule compared to Ti(OMe)₄.

Depositions of TiO_2 using this titanium precursor are carried out at temperatures above 400°C , showing the higher stability of $\text{Ti}(\text{MeCp})(\text{OMe})_3$ compared to $\text{Ti}(\text{OMe})_4$.

The deposition of SrTiO_3 using this titanium precursor and $\text{Sr}(\text{iPr}_3\text{Cp})_2, \text{THF}$ is carried out above 400°C , leading to a stoichiometric film.

CLAIMS

1. A method for forming a titanium-containing layer on a substrate, the
5 method comprising at least the steps of:

a) providing a vapor comprising at least one precursor compound of the
formula $\text{Ti}(\text{Me}_5\text{Cp})(\text{OR})_3$ (I), wherein R is selected in the group consisting in
methyl, ethyl, isopropyl; or of the formula $\text{Ti}(\text{R}^1\text{Cp})(\text{OR}^2)_3$ (II), wherein R^1 is
selected from the group consisting in H, methyl, ethyl, isopropyl and R^2 is
10 independently selected from the group consisting in methyl, ethyl, isopropyl or tert-
butyl;

b) reacting the vapor comprising the at least one compound of formula (I)
or (II) with the substrate, according to an atomic layer deposition process, to form
a layer of a tantalum-containing complex on at least one surface of said substrate.
15

2. The method of claim 1 further comprising the step:

c) reaction of the complex formed obtained in step b) with a reagent
selected from another metal source, reducing reactants and/or nitriding reactants
and/or oxidizing reactants.
20

3. The method of one of claims 1 to 2 wherein each R is different from one
another.

4. The method of one of claims 1 to 3 wherein the vapour provided in step
25 a) further comprises one or more metal (M')-organic precursor(s) to produce thin
films containing titanium and M' .

5. The method of one of claims 1 to 4, further comprising providing at least
one reaction gas wherein the at least one reaction gas is selected from the group
30 consisting of hydrogen, hydrogen sulfide, hydrogen selenide, hydrogen telluride,
carbon monoxide, ammonia, organic amine, silane, disilane, higher silanes,
silylamines, diborane, hydrazine, methylhydrazine, chlorosilane and
chloropolysilane, metal alkyl, arsine, phosphine, trialkylboron, oxygen, ozone,
water, hydrogen peroxide, nitrous oxide, nitrogen monoxide, nitrogen dioxide,

alcohols, plasma comprising fragments of those species, and combinations thereof, preferably ozone or water.

6. The method of one of claims 1 to 5, wherein the titanium precursor is
5 selected in the group consisting of:

Ti(MeCp)(OEt)₃; Ti(MeCp)(OiPr)₃; Ti(MeCp)(OtBu)₃; Ti(EtCp)(OEt)₃; Ti(EtCp)(
OiPr)₃; Ti(EtCp)(OtBu)₃; Ti(Cp)(OEt)₃; Ti(Cp)(OiPr)₃; Ti(Cp)(OtBu)₃,
Ti(Me₅Cp)(OMe)₃, Ti(MeCp)(OMe)₃, Ti(MeCp)(OEt)₃, Ti(MeCp)(OiPr)₃,
Ti(MeCp)(OtBu)₃, Ti(EtCp)(OMe)₃, Ti(EtCp)(OEt)₃, Ti(EtCp)(OiPr)₃,
10 Ti(EtCp)(OtBu)₃, Ti(Cp)(OMe)₃, Ti(Cp)(OEt)₃, Ti(Cp)(OiPr)₃, Ti(Cp)(OtBu)₃,
Ti(iPrCp)(OMe)₃, Ti(iPrCp)(OEt)₃, Ti(iPrCp)(OiPr)₃, Ti(iPrCp)(OtBu)₃.

7. The method of one of claims 1 to 6, wherein the temperature of the
substrate is 25°C to 600°C, preferably 380°C to 425°C, and wherein the atomic
15 layer deposition chamber containing the substrate has a pressure of 0.133 Pa to
133 kPa, preferably below 27 kPa.

8. The method of one of claims 1 to 7 further comprising the step of
purging excess vapor comprising the at least one compound of formula (I) from the
20 substrate, with an inert gas selected from the group consisting of hydrogen,
nitrogen, helium, argon, and mixtures thereof.

9. A method of manufacturing a semiconductor structure, comprising the
steps of the method defined in any one of claims 1 to 8, wherein the substrate is a
25 semiconductor substrate.