

As originally filed**Closed reversible breathing apparatus with metal-organic framework material**

- 5 The present invention relates to methods for removing carbon dioxide and, if appropriate, water from breathing air in closed or partially closed systems using a porous metal-organic framework material, such systems having at least one breathing apparatus and also their use and methods for regenerating the porous metal-organic framework material.
- 10 In closed or partially closed systems it is necessary that owing to the limited supply of oxygen this must be replenished if, for example, a person wishes to remain in this system for longer than the oxygen supply which is provided by the volume of the system would permit.
- 15 Air or oxygen is generally supplied in this case by corresponding pressure vessels, such as, for example, pressure cylinders.

For instance it is necessary, for example, during diving that the diver, in addition to a diving mask, also carries in conjunction oxygen cylinders if the diver wishes to remain for a
20 relatively long time under water.

Customarily in this case oxygen is supplied to the diver via a mouthpiece from the pressure cylinder, which the diver can breath. The expired air is released to the surrounding water. By this means the diver can remain under water for longer than the air volume of the diving
25 mask would provide.

Nevertheless, the time spent below water is restricted for the diver by the volume of the pressure cylinder. A further possibility for optimization and, in association, a prolongation of the time spent below water is additionally using an adsorbent which is suitable for removing
30 from the air the carbon dioxide present in the expired air in such a manner that the air having the remaining oxygen can be again provided for breathing.

Such systems having adsorbents are known in the prior art. Examples of these are described in EP-A 0 782 953, DE-A 197 167 49 and also DE-A 198 16 373.

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The adsorbents described in the prior art which can be used comprise different materials.

In GB-A 1 438 757, for example, use is made of a soda lime bed for a diving apparatus.

WO-A 01/83294 describes, for example, a breathing apparatus, in which the carbon dioxide absorber is said to be able to be reactivated by heat or reduced carbon dioxide pressure.

5 An example of such an absorber mentioned is calcium hydroxide.

DE-A 33 03 420 describes methods and devices for purifying breathing air from CO₂, molecular sieves acting as adsorbents which can be regenerated by a pressure-swing method.

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Finally, special adsorbents are described in EP-A 1 155 728. These are amino-methylated bead polymers.

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Despite these numerous adsorbents proposed in the prior art, there is still a requirement to provide further optimized adsorbents for removing carbon dioxide and, if appropriate, water from the breathing air.

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An object of the present invention is thus that further improved adsorbents are provided for the abovementioned methods and apparatuses.

The object is achieved by a method for removing carbon dioxide and, if appropriate, water from breathing air in closed or partially closed systems comprising the step

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- contacting the breathing air with a porous metal-organic framework material, the framework material comprising at least one at least bidentate organic compound which is bound by coordination to at least one metal ion.

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The object is further achieved by a closed or partially closed system which comprises at least one breathing apparatus and also a breathing mask, a breathing suit or other life support system, further comprising a porous metal-organic framework material, the framework material comprising at least one at least bidentate organic compound which is bound by coordination to at least one metal ion.

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This is because it has been found that the use of porous metal-organic framework materials in closed or partially closed systems which comprise at least one breathing apparatus and in methods for removing carbon dioxide and, if appropriate, water, from breathing air are particularly efficient and, in addition, can be readily regenerated.

To carry out the inventive method for removing carbon dioxide and, if appropriate, water, particular use can be made of a closed or partially closed system which comprises at least one breathing apparatus and also a breathing mask, a breathing suit, or other life support systems.

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Closed systems are, in particular, those which have no opening to the surroundings through which atmospheric oxygen is to be introduced or removed.

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Partially closed systems are, in particular, those in which no atmospheric oxygen is to be taken up into the system through the surroundings.

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Surroundings of the closed or partially closed system which come into consideration are in principle any surroundings which do not contain surrounding gas or have a surrounding gas, the breathing of which does not ensure the necessary life support or freedom from harm of a human or higher animal.

Surroundings which contain no surrounding gas are situated, for example, under water or in space.

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A surrounding gas, the breathing of which does not ensure the necessary life support or freedom from harm of a human or higher animal is, for example, air whose oxygen fraction or partial pressure is too low for breathing and/or which has other harmful constituents.

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The breathing mask can be, for example, a mask such as is used in diving, therefore a diving mask. However, likewise, it can be a respiratory protection mask, as can be used, for example, in the case of fire, in a chemical accident, during painting or handling hazardous chemical or biological material, in extreme mountain climbing or at a great height (for example in an aircraft). In addition, such systems can also comprise suits. In addition, it is possible that the life support system is a helmet. Typically, such as a helmet can also be integrated into a corresponding suit. Frequently, in this connection, full protective suits can be mentioned. Space suits may also be mentioned in this context. Likewise, it can also be systems for rooms or passages of buildings, for example protective rooms, or of vehicles, for example in submarines, aircraft, in tunnels, mineshafts or the like.

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The closed or partially closed system can in addition have a filter in which the porous metal-organic framework material is present at least as part of an adsorber bed. Other adsorbents such as zeolites can likewise be present.

The filter can be exchangeable or be installed fixed in the system. The filters to be used are known from the prior art. These are typically constituents of the systems which are likewise known in the prior art.

- 5 Preferably, the inventive closed or partially closed system is used for removing carbon dioxide and, if appropriate, water, from breathing air. In this case it is advantageous that, in addition to CO₂, also the water present in the breathing air can be removed. However, this is not a precondition for functioning of the CO₂ adsorption.
- 10 The porous metal-organic framework material is, inter alia, therefore advantageous because ready regeneration is possible.

Therefore, the present invention further relates to a method for regenerating a porous metal-organic framework material from a closed or partially closed system as has been
15 described above comprising the steps

- if appropriate removing the metal-organic framework material; and
- impinging the framework material with a gas.

- 20 The gas can be, for example, air, nitrogen, an inert gas or a mixture thereof. Suitable inert gases are, for example, helium or argon.

The regeneration can be performed, for example, by simply passing the gas through the metal-organic framework material. Preferably, however, the regeneration takes place under
25 pressure-swing and/or temperature-swing adsorption.

Therefore, it is preferred when the inventive method for regeneration is carried out in such a manner that the impingement takes place with the change of at least one parameter selected from pressure and temperature.

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The term "pressure", in the context of the present invention, is to be taken to mean the total pressure and/or the carbon dioxide partial pressure.

The regeneration of the metal-organic framework material can be performed during the use
35 of the inventive closed or partially closed system.

The porous metal-organic framework material to be used is known in the prior art. The

suitability of porous metal-organic framework materials for storage of carbon dioxide has been described, for example, by A.R. Millward et al., J. Am. Chem. Soc. 127 (2005), 17998-17999.

- 5 The porous metal-organic framework material comprises at least one at least bidentate, organic compound which is bound by coordination to a metal ion. This metal-organic framework material (MOF) is described, for example, in US 5,648,508, EP-A-0 790 253, M. O'Keeffe et al., J. Sol. State Chem., 152 (2000), pages 3 to 20, H. Li et al., Nature 402 (1999), page 276, M. Eddaoudi et al., Topics in Catalysis 9 (1999), pages 105 to 111, B.
10 Chen et al., Science 291 (2001), pages 1021 to 1023 and DE-A-101 11 230.

The MOFs according to the present invention comprise pores, in particular micropores and/or mesopores. Micropores are defined as those having a diameter of 2 nm or less and mesopores are defined by a diameter in the range from 2 to 50 nm, in each case in
15 accordance with the definition as reported in Pure & Applied Chem. 57 (1985), 603-619, in particular page 606. The presence of micropores and/or mesopores can be checked with the aid of sorption measurements, these measurements determining the uptake capacity of the metal-organic framework material for nitrogen at 77 Kelvin as specified in DIN 66131 and/or DIN 66134.

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Preferably, the specific surface area, calculated according to the Langmuir model (DIN 66131, 66134) for an MOF in powder form is greater than 5 m²/g, more preferably above 10 m²/g, more preferably greater than 50 m²/g, further more preferably greater than 500 m²/g, further more preferably greater than 1000 m²/g, and particularly preferably
25 greater than 1500 m²/g.

MOF shaped bodies can have a lower specific surface area; but preferably greater than 10 m²/g, more preferably greater than 50 m²/g, further more preferably greater than 500 m²/g.

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The metal component in the framework material according to the present invention is preferably selected from the groups Ia, IIa, IIIa, IVa to VIIIa and Ib to VIb. Particular preference is given to Mg, Ca, Sr, Ba, Sc, Y, Ti, Zr, Hf, V, Nb, Ta, Cr, Mo, W, Mn, Re, Fe, Ru, Os, Co, Rh, Ir, Ni, Pd, Pt, Cu, Ag, Au, Zn, Cd, Hg, Al, Ga, In, Tl, Si, Ge, Sn, Pb, As, Sb
35 and Bi. More preference is given to Zn, Cu, Ni, Pd, Pt, Ru, Rh and Co. In particular preference is given to Zn, Al, Ni and Cu. With respect to the ions of these elements, those which may particularly be mentioned are Mg²⁺, Ca²⁺, Sr²⁺, Ba²⁺, Sc³⁺, Y³⁺, Ti⁴⁺, Zr⁴⁺, Hf⁴⁺,

V^{4+} , V^{3+} , V^{2+} , Nb^{3+} , Ta^{3+} , Cr^{3+} , Mo^{3+} , W^{3+} , Mn^{3+} , Mn^{2+} , Re^{3+} , Re^{2+} , Fe^{3+} , Fe^{2+} , Ru^{3+} , Ru^{2+} , Os^{3+} , Os^{2+} , Co^{3+} , Co^{2+} , Rh^{2+} , Rh^{+} , Ir^{2+} , Ir^{+} , Ni^{2+} , Ni^{+} , Pd^{2+} , Pd^{+} , Pt^{2+} , Pt^{+} , Cu^{2+} , Cu^{+} , Ag^{+} , Au^{+} , Zn^{2+} , Cd^{2+} , Hg^{2+} , Al^{3+} , Ga^{3+} , In^{3+} , Tl^{3+} , Si^{4+} , Si^{2+} , Ge^{4+} , Ge^{2+} , Sn^{4+} , Sn^{2+} , Pb^{4+} , Pb^{2+} , As^{5+} , As^{3+} , As^{+} , Sb^{5+} , Sb^{3+} , Sb^{+} , Bi^{5+} , Bi^{3+} and Bi^{+} .

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The term "at least bidentate organic compound" designates an organic compound which comprises at least one functional group which is able to form, to a given metal ion, at least two, preferably two coordinate, bonds and/or to two or more, preferably two, metal atoms in each case one coordinate bond.

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As functional groups via which said coordinate bonds can be developed, in particular the following functional groups may be mentioned by way of example: $-CO_2H$, $-CS_2H$, $-NO_2$, $-B(OH)_2$, $-SO_3H$, $-Si(OH)_3$, $-Ge(OH)_3$, $-Sn(OH)_3$, $-Si(SH)_4$, $-Ge(SH)_4$, $-Sn(SH)_3$, $-PO_3H$, $-AsO_3H$, $-AsO_4H$, $-P(SH)_3$, $-As(SH)_3$, $-CH(RSH)_2$, $-C(RSH)_3$, $-CH(RNH_2)_2$, $-C(RNH_2)_3$, $-CH(ROH)_2$, $-C(ROH)_3$, $-CH(RCN)_2$, $-C(RCN)_3$, in which R, for example, can preferably be an alkylene group having 1, 2, 3, 4 or 5 carbon atoms such as, for example, a methylene, ethylene, n-propylene, isopropylene, n-butylene, isobutylene, tert-butylene or n-pentylene group, or an aryl group comprising one or two aromatic nuclei, such as, for example, 2 C_6 rings which can, if appropriate, be condensed and independently of one another can be suitably substituted with at least in each case one substituent, and/or which, independently of one another, can each comprise at least one heteroatom, such as, for example, N, O and/or S. According to likewise preferred embodiments, functional groups may be mentioned in which the abovementioned radical R is not present. In this respect, inter alia, $-CH(SH)_2$, $-C(SH)_3$, $-CH(NH_2)_2$, $-C(NH_2)_3$, $CH(OH)_2$, $-C(OH)_3$, $-CH(CN)_2$ or $-C(CN)_3$ may be mentioned.

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The at least two functional groups can in principle be bound to any suitable organic compound provided that it is ensured that the organic compound having these functional groups is capable of forming the coordinate bond and for producing the framework material.

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Preferably, the organic compounds which comprise the at least two functional groups are derived from a saturated or unsaturated aliphatic compound or an aromatic compound or a compound which is both aliphatic and aromatic.

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The aliphatic compound or the aliphatic part of the compound which is both aliphatic and aromatic can be linear and/or branched and/or cyclic, a plurality of cycles per compound also being possible. Further preferably, the aliphatic compound or the aliphatic part of the

compound which is both aliphatic and aromatic comprises 1 to 15, further preferably 1 to 14, further preferably 1 to 13, further preferably 1 to 12, further preferably 1 to 11, and in particular preferably 1 to 10, carbon atoms, such as, for example, 1, 2, 3, 4, 5, 6, 7, 8, 9 or 10 carbon atoms. In particular preference is given in this case to, inter alia, methane, adamantane, acetylene, ethylene or butadiene.

The aromatic compound or the aromatic part of the compound which is both aromatic and aliphatic can have one or else a plurality of nuclei, such as, for example, two, three, four or five nuclei, the nuclei being able to be present separately from one another and/or at least two nuclei in condensed form. Particularly preferably, the aromatic compound or the aromatic part of the compound which is both aliphatic and aromatic has one, two or three nuclei, one or two nuclei being particularly preferred. Independently of one another, in addition, each nucleus of said compound can comprise at least one heteroatom, such as, for example, N, O, S, B, P, Si, Al, preferably N, O and/or S. Further preferably, the aromatic compound or the aromatic part of the compound which is both aromatic and aliphatic comprises one or two C₆ nuclei, the two either being present separately of one another or in condensed form. In particular, as aromatic compounds, mention may be made of benzene, naphthalene and/or biphenyl and/or bipyridyl and/or pyridyl.

Particularly preferably, the at least bidentate, organic compound is derived from a di-, tri-, or tetracarboxylic acid, or their sulfur analogs. Sulfur analogs are the functional groups -C(=O)SH and also their tautomers and C(=S)SH which can be used instead of one or more carboxylic acid groups.

The term "derive" in the context of the present invention means that the at least bidentate, organic compound in the framework material can be present in partly deprotonated or completely deprotonated form. In addition, the at least bidentate, organic compound can comprise further substituents such as, for example, -OH, -NH₂, -OCH₃, -NH(CH₃), -N(CH₃)₂, -CN and also halides.

For example, in the context of the present invention, mention may be made of dicarboxylic acids, such as

oxalic acid, succinic acid, tartaric acid, 1,4-butanedicarboxylic acid, 4-oxopyran-2,6-dicarboxylic acid, 1,6-hexanedicarboxylic acid, decanedicarboxylic acid, 1,8-heptadecanedicarboxylic acid, 1,9-heptadecanedicarboxylic acid, hepta-decanedicarboxylic acid, acetylenedicarboxylic acid, 1,2-benzenedicarboxylic acid,

2,3-pyridinedicarboxylic acid, pyridine-2,3-dicarboxylic acid, 1,3-butadiene-1,4-dicarboxylic acid, 1,4-benzenedicarboxylic acid, p-benzenedicarboxylic acid, imidazole-2,4-dicarboxylic acid, 2-methylquinoline-3,4-dicarboxylic acid, quinoline-2,4-dicarboxylic acid, quinoxaline-2,3-dicarboxylic acid, 6-chloroquinoxaline-2,3-dicarboxylic acid,

5 4,4'-diaminophenylmethane-3,3'-dicarboxylic acid, quinoline-3,4-dicarboxylic acid, 7-chloro-4-hydroxyquinoline-2,8-dicarboxylic acid, diimidodicarboxylic acid, pyridine-2,6-dicarboxylic acid, 2-methylimidazole-4,5-dicarboxylic acid, thiophene-3,4-dicarboxylic acid, 2-isopropylimidazole-4,5-dicarboxylic acid, tetrahydropyran-4,4-dicarboxylic acid, perylene-3,9-dicarboxylic acid, perylenedicarboxylic acid, Pluriol E 200 dicarboxylic acid, 3,6-dioxaoctanedicarboxylic acid, 3,5-cyclohexadiene-1,2-dicarboxylic acid, octadecarboxylic acid,

10 pentane-3,3-carboxylic acid, 4,4'-diamino-1,1'-biphenyl-3,3'-dicarboxylic acid, 4,4'-diaminobiphenyl-3,3'-dicarboxylic acid, benzidine-3,3'-dicarboxylic acid, 1,4-bis(phenylamino)benzene-2,5-dicarboxylic acid, 1,1'-binaphthyl-5,5'-dicarboxylic acid, 7-chloro-8-methylquinoline-2,3-dicarboxylic acid, 1-anilinoanthraquinone-2,4'-dicarboxylic acid, polytetrahydrofuran-250-dicarboxylic acid, 1,4-bis(carboxymethyl)piperazine-2,3-

15 dicarboxylic acid, 7-chloroquinoline-3,8-dicarboxylic acid, 1-(4-carboxy)phenyl-3-(4-chloro)phenylpyrazoline-4,5-dicarboxylic acid, 1, 4, 5, 6, 7, 7-hexachloro-5-norbornene-2,3-dicarboxylic acid, phenylindanedicarboxylic acid, 1,3-dibenzyl-2-oxoimidazolidine-4,5-dicarboxylic acid, 1,4-cyclohexanedicarboxylic acid, naphthalene-1,8-dicarboxylic acid,

20 2-benzoylbenzene-1,3-dicarboxylic acid, 1,3-dibenzyl-2-oxoimidazolidine-4,5-cis-dicarboxylic acid, 2,2'-biquinoline-4,4'-dicarboxylic acid, pyridine-3,4-dicarboxylic acid, 3,6,9-trioxaundecanedicarboxylic acid, O-hydroxybenzophenonedicarboxylic acid, Pluriol E 300 dicarboxylic acid, Pluriol E 400 dicarboxylic acid, Pluriol E 600 dicarboxylic acid, pyrazole-3,4-dicarboxylic acid, 2,3-pyrazinedicarboxylic acid, 5,6-dimethyl-2,3-

25 pyrazinedicarboxylic acid, 4,4'-diaminodiphenyl ether diimidodicarboxylic acid, 4,4'-diaminodiphenylmethane diimidodicarboxylic acid, 4,4'-diaminodiphenyl sulfone diimidodicarboxylic acid, 2,6-naphthalenedicarboxylic acid, 1,3-adamantanedicarboxylic acid, 1,8-naphthalenedicarboxylic acid, 2,3-naphthalenedicarboxylic acid, 8-methoxy-2,3-naphthalenedicarboxylic acid, 8-nitro-2,3-naphthalenecarboxylic acid, 8-sulfo-2,3-

30 naphthalenedicarboxylic acid, anthracene-2,3-dicarboxylic acid, 2',3'-diphenyl-p-terphenyl-4,4''-dicarboxylic acid, diphenyl ether-4,4'-dicarboxylic acid, imidazole-4,5-dicarboxylic acid, 4(1H)-oxothiochromene-2,8-dicarboxylic acid, 5-tert-butyl-1,3-benzenedicarboxylic acid, 7,8-quinolinedicarboxylic acid, 4,5-imidazoledicarboxylic acid, 4-cyclohexene-1,2-dicarboxylic acid, hexatriacontanedicarboxylic acid, tetradecanedicarboxylic acid,

35 1,7-heptadecarboxylic acid, 5-hydroxy-1,3-benzenedicarboxylic acid, pyrazine-2,3-dicarboxylic acid, furan-2,5-dicarboxylic acid, 1-nonene-6,9-dicarboxylic acid, eicosenedicarboxylic acid, 4,4'-dihydroxydiphenylmethane-3,3'-dicarboxylic acid, 1-amino-

4-methyl-9,10-dioxo-9,10-dihydroanthracene-2,3-dicarboxylic acid, 2,5-pyridinedicarboxylic acid, cyclohexene-2,3-dicarboxylic acid, 2,9-dichlorofluororubine-4,11-dicarboxylic acid, 7-chloro-3-methylquinoline-6,8-dicarboxylic acid, 2,4-dichlorobenzophenone-2',5'-dicarboxylic acid, 1,3-benzenedicarboxylic acid, 2,6-pyridinedicarboxylic acid, 1-methylpyrrole-3,4-dicarboxylic acid, 1-benzyl-1H-pyrrole-3,4-dicarboxylic acid, anthraquinone-1,5-dicarboxylic acid, 3,5-pyrazoledicarboxylic acid, 2-nitrobenzene-1,4-dicarboxylic acid, heptane-1,7-dicarboxylic acid, cyclobutane-1,1-dicarboxylic acid, 1,14-tetradecanedicarboxylic acid, 5,6-dehydronorbornane-2,3-dicarboxylic acids, or 5-ethyl-2,3-pyridinedicarboxylic acid,

tricarboxylic acid, such as

2-hydroxy-1,2,3-propanetricarboxylic acid, 7-chloro-2,3,8-quinolinetricarboxylic acid, 1,2,4-benzenetricarboxylic acid, 1,2,4-butanetricarboxylic acid, 2-phosphono-1,2,4-butanetricarboxylic acid, 1,3,5-benzenetricarboxylic acid, 1-hydroxy-1,2,3-propanetricarboxylic acid, 4,5-dihydro-4,5-dioxo-1H-pyrrolo[2,3-F]quinoline-2,7,9-tricarboxylic acid, 5-acetyl-3-amino-6-methylbenzene-1,2,4-tricarboxylic acid, 3-amino-5-benzoyl-6-methylbenzene-1,2,4-tricarboxylic acid, 1,2,3-propanetricarboxylic acid or aurintricarboxylic acid,

or tetracarboxylic acids, such as

perylene-1,10,12,13-tetracarboxylic acid, perylene-3,4,9,10-tetracarboxylic acid, perylene-1,12,13,14-tetracarboxylic acid, perylene-1,12,13,14-bis(sulfone)-3,4,9,10-tetracarboxylic acid, butanetetracarboxylic acid such as 1,2,3,4-butanetetracarboxylic acid or meso-1,2,3,4-butanetetracarboxylic acid, decane-2,4,6,8-tetracarboxylic acid, 1,4,7,10,13,16-hexaoxacyclooctadecane-2,3,11,12-tetracarboxylic acid, 1,2,4,5-benzenetetracarboxylic acid, 1,2,11,12-dodecanetetracarboxylic acid, 1,2,5,6-hexanetetracarboxylic acid, 1,2,7,8-octanetetracarboxylic acid, 1,4,5,8-naphthalenetetracarboxylic acid, 1,2,9,10-decanetetracarboxylic acid, benzophenonetetracarboxylic acid, 3,3',4,4'-benzophenonetetracarboxylic acid, tetrahydrofuran-tetracarboxylic acid or cyclopentanetetracarboxylic acids such as cyclopentane-1,2,3,4-tetracarboxylic acid.

Very particularly preferably, use is made of, if appropriate at least monosubstituted, mono-, di-, tri-, tetranuclear or higher nuclear aromatic di-, tri- or tetracarboxylic acids, with each of the nuclei being able to comprise at least one heteroatom, with two or more nuclei being

able to comprise identical or different heteroatoms. For example, preference is given to mononuclear dicarboxylic acids, mononuclear tricarboxylic acids, mononuclear tetracarboxylic acids, dinuclear dicarboxylic acids, dinuclear tricarboxylic acids, dinuclear tetracarboxylic acids, trinuclear dicarboxylic acids, trinuclear tricarboxylic acids, trinuclear tetracarboxylic acids, tetranuclear dicarboxylic acids, tetranuclear tricarboxylic acids and/or tetranuclear tetracarboxylic acids. Suitable heteroatoms are, for example, N, O, S, B, P, Si, Al, preferred heteroatoms here are N, S, and/or O. A suitable substituent which may be mentioned in this respect, is, inter alia, -OH, a nitro group, an amino group or an alkyl or alkoxy group.

In particular preferably, as at least bidentate, organic compounds, use is made of acetylenedicarboxylic acid (ADC), benzenedicarboxylic acids, naphthalenedicarboxylic acids, biphenyldicarboxylic acids, such as, for example, 4,4'-biphenyldicarboxylic acid (BPDC), bipyridinedicarboxylic acids, such as, for example, 2,2'-bipyridinedicarboxylic acids, such as, for example, 2,2'-bipyridine-5,5'-dicarboxylic acid, benzenetricarboxylic acids, such as, for example, 1,2,3-benzenetricarboxylic acid or 1,3,5-benzenetricarboxylic acid (BTC), adamantane tetracarboxylic acid (ATC), adamantane dibenzoate (ADB), benzene tribenzoate (BTB), methane tetrabenzoate (MTB), adamantane tetrabenzoate or dihydroxyterephthalic acids, such as, for example, 2,5-dihydroxyterephthalic acid (DHBDC).

Very particularly preferably, use is made of, inter alia, isophthalic acid, terephthalic acid, 2,5-dihydroxyterephthalic acid, 1,2,3-benzenetricarboxylic acid, 1,3,5-benzenetricarboxylic acid or 2,2'-bipyridine-5,5'-dicarboxylic acid.

In addition to these at least bidentate, organic compounds, the MOF can also comprise one or more unidentate ligands.

Suitable solvents for producing the MOF are, inter alia, ethanol, dimethylformamide, toluene, methanol, chlorobenzene, diethylformamide, dimethyl sulfoxide, water, hydrogen peroxide, methylamine, sodium hydroxide solution, N-methylpyrrolidone ether, acetonitrile, benzyl chloride, triethylamine, ethylene glycol and mixtures thereof. Further metal ions, at least bidentate, organic compounds, and solvents for the production of MOF are described, inter alia, in US-A 5,648,508 or DE-A 101 11 230.

The pore size of the MOF can be controlled by selection of suitable ligands and/or the at least bidentate, organic compound. In general it is true that the larger the organic compound, the larger the pore size. Preferably, the pore size is from 0.2 nm to 30 nm,

particularly preferably the pore size is in the range from 0.3 nm to 3 nm, based on the crystalline material.

In an MOF shaped body, however, larger pores also occur, the size distribution of which can vary. Preferably, however, more than 50% of the total pore volume, in particular more than 75%, is formed by pores having a pore diameter of up to 1000 nm. Preferably, however, a majority of the pore volume is formed by pores of two diameter ranges. It is therefore further preferred if more than 25% of the total pore volume, in particular more than 50% total pore volume, is formed by pores which are in a diameter range from 100 nm to 800 nm and if more than 15% of the total pore volume, in particular more than 25% of the total pore volume, is formed by pores which are in a diameter range of up to 10 nm. The pore distribution can be determined by means of mercury porosimetry.

Examples of MOFs are given below. In addition to the designation of the MOF, the metal and also the at least bidentate ligand, furthermore the solvent and also the cell parameters (angles α , β and γ , and also the distances A, B and C in Å) are reported. The latter were determined by X-ray diffraction.

MOF-n	Constituents molar ratio M+L	Solvent s	α	β	γ	a	b	c	Space group
MOF-0	Zn(NO ₃) ₂ ·6H ₂ O H ₃ (BTC)	Ethanol	90	90	120	16.711	16.711	14.189	P6(3)/ Mcm
MOF-2	Zn(NO ₃) ₂ ·6H ₂ O (0.246 mmol) H ₂ (BDC) 0.241 mmol)	DMF Toluene	90	102.8	90	6.718	15.49	12.43	P2(1)/n
MOF-3	Zn(NO ₃) ₂ ·6H ₂ O (1.89 mmol) H ₂ (BDC) (1.93mmol)	DMF MeOH	99.72	111.11	108.4	9.726	9.911	10.45	P-1
MOF-4	Zn(NO ₃) ₂ ·6H ₂ O (1.00 mmol) H ₃ (BTC) (0.5 mmol)	Ethanol	90	90	90	14.728	14.728	14.728	P2(1)3
MOF-5	Zn(NO ₃) ₂ ·6H ₂ O	DMF	90	90	90	25.669	25.669	25.669	Fm-3m

	(2.22 mmol) H ₂ (BDC) (2.17 mmol)	Chloro- benzene							
MOF-38	Zn(NO ₃) ₂ ·6H ₂ O (0.27 mmol) H ₃ (BTC) (0.15 mmol)	DMF Chloro- benzene	90	90	90	20.657	20.657	17.84	l4cm
MOF-31 Zn(ADC) ₂	Zn(NO ₃) ₂ ·6H ₂ O 0.4 mmol H ₂ (ADC) 0.8 mmol	Ethanol	90	90	90	10.821	10.821	10.821	Pn(-3)m
MOF-12 Zn ₂ (ATC)	Zn(NO ₃) ₂ ·6H ₂ O 0.3 mmol H ₄ (ATC) 0.15 mmol	Ethanol	90	90	90	15.745	16.907	18.167	Pbca
MOF-20 ZnNDC	Zn(NO ₃) ₂ ·6H ₂ O 0.37 mmol H ₂ NDC 0.36 mmol	DMF Chloro- benzene	90	92.13	90	8.13	16.444	12.807	P2(1)/c
MOF-37	Zn(NO ₃) ₂ ·6H ₂ O 0.2 mmol H ₂ NDC 0.2 mmol	DEF Chloro- benzene	72.38	83.16	84.33	9.952	11.576	15.556	P-1
MOF-8 Tb ₂ (ADC)	Tb(NO ₃) ₃ ·5H ₂ O 0.10 mmol H ₂ ADC 0.20 mmol	DMSO MeOH	90	115.7	90	19.83	9.822	19.183	C2/c
MOF-9 Tb ₂ (ADC)	Tb(NO ₃) ₃ ·5H ₂ O 0.08 mmol H ₂ ADB 0.12 mmol	DMSO	90	102.09	90	27.056	16.795	28.139	C2/c
MOF-6	Tb(NO ₃) ₃ ·5H ₂ O 0.30 mmol H ₂ (BDC) 0.30 mmol	DMF MeOH	90	91.28	90	17.599	19.996	10.545	P21/c

MOF-7	Tb(NO ₃) ₃ ·5H ₂ O 0.15 mmol H ₂ (BDC) 0.15 mmol	H ₂ O	102.3	91.12	101.5	6.142	10.069	10.096	P-1
MOF-69A	Zn(NO ₃) ₂ ·6H ₂ O 0.083 mmol 4,4'-BPDC 0.041 mmol	DEF H ₂ O ₂ MeNH ₂	90	111.6	90	23.12	20.92	12	C2/c
MOF-69B	Zn(NO ₃) ₂ ·6H ₂ O 0.083 mmol 2,6-NCD 0.041 mmol	DEF H ₂ O ₂ MeNH ₂	90	95.3	90	20.17	18.55	12.16	C2/c
MOF-11 Cu ₂ (ATC)	Cu(NO ₃) ₂ ·2.5H ₂ O 0.47 mmol H ₂ ATC 0.22 mmol	H ₂ O	90	93.86	90	12.987	11.22	11.336	C2/c
MOF-11 Cu ₂ (ATC) dehydr.			90	90	90	8.4671	8.4671	14.44	P42/ mmc
MOF-14 Cu ₃ (BTB)	Cu(NO ₃) ₂ ·2.5H ₂ O 0.28 mmol H ₃ BTB 0.052 mmol	H ₂ O DMF EtOH	90	90	90	26.946	26.946	26.946	Im-3
MOF-32 Cd(ATC)	Cd(NO ₃) ₂ ·4H ₂ O 0.24 mmol H ₄ ATC 0.10 mmol	H ₂ O NaOH	90	90	90	13.468	13.468	13.468	P(-4)3m
MOF-33 Zn ₂ (ATB)	ZnCl ₂ 0.15 mmol H ₄ ATB 0.02 mmol	H ₂ O DMF EtOH	90	90	90	19.561	15.255	23.404	Imma
MOF-34 Ni(ATC)	Ni(NO ₃) ₂ ·6H ₂ O 0.24 mmol H ₄ ATC 0.10 mmol	H ₂ O NaOH	90	90	90	10.066	11.163	19.201	P2 ₁ 2 ₁ 2 ₁

MOF-36 Zn ₂ (MTB)	Zn(NO ₃) ₂ ·4H ₂ O 0.20 mmol H ₄ MTB 0.04 mmol	H ₂ O DMF	90	90	90	15.745	16.907	18.167	Pbca
MOF-39 Zn ₃ O(HBTB)	Zn(NO ₃) ₂ ·4H ₂ O 0.27 mmol H ₃ BTB 0.07 mmol	H ₂ O DMF EtOH	90	90	90	17.158	21.591	25.308	Pnma
NO305	FeCl ₂ ·4H ₂ O 5.03 mmol formic acid 86.90 mmol	DMF	90	90	120	8.2692	8.2692	63.566	R-3c
NO306A	FeCl ₂ ·4H ₂ O 5.03 mmol formic acid 86.90 mmol	DEF	90	90	90	9.9364	18.374	18.374	Pbcn

NO29 MOF-0 similar	Mn(Ac) ₂ ·4H ₂ O 0.46 mmol H ₃ BTC 0.69 mmol	DMF	120	90	90	14.16	33.521	33.521	P-1
BPR48 A2	Zn(NO ₃) ₂ ·6H ₂ O 0.012 mmol H ₂ BDC 0.012 mmol	DMSO Toluene	90	90	90	14.5	17.04	18.02	Pbca
BPR69 B1	Cd(NO ₃) ₂ ·4H ₂ O 0.0212 mmol H ₂ BDC 0.0428 mmol	DMSO	90	98.76	90	14.16	15.72	17.66	Cc
BPR92 A2	Co(NO ₃) ₂ ·6H ₂ O 0.018 mmol H ₂ BDC 0.018 mmol	NMP	106.3	107.63	107.2	7.5308	10.942	11.025	P1

BPR95 C5	Cd(NO ₃) ₂ ·4H ₂ O 0.012 mmol H ₂ BDC 0.36 mmol	NMP	90	112.8	90	14.460	11.085	15.829	P2(1)/n
Cu C ₆ H ₄ O ₆	Cu(NO ₃) ₂ ·2.5H ₂ O 0.370 mmol H ₂ BDC(OH) ₂ 0.37 mmol	DMF Chloro- benzene	90	105.29	90	15.259	14.816	14.13	P2(1)/c
M(BTC) MOF-0 similar	Co(SO ₄)·H ₂ O 0.055 mmol H ₃ BTC 0.037 mmol	DMF	as MOF-0						
Tb(C ₆ H ₄ O ₆)	Tb(NO ₃) ₃ ·5H ₂ O 0.370 mmol H ₂ (C ₆ H ₄ O ₆) 0.56 mmol	DMF Chloro- benzene	104.6	107.9	97.147	10.491	10.981	12.541	P-1
Zn (C ₂ O ₄)	ZnCl ₂ 0.370 mmol oxalic acid 0.37 mmol	DMF Chloro- benzene	90	120	90	9.4168	9.4168	8.464	P(-3)1m
Co(CHO)	Co(NO ₃) ₂ ·5H ₂ O 0.043 mmol formic acid 1.60 mmol	DMF	90	91.32	90	11.328	10.049	14.854	P2(1)/n
Cd(CHO)	Cd(NO ₃) ₂ ·4H ₂ O 0.185 mmol formic acid 0.185 mmol	DMF	90	120	90	8.5168	8.5168	22.674	R-3c
Cu(C ₃ H ₂ O ₄)	Cu(NO ₃) ₂ ·2.5H ₂ O 0.043 mmol malonic acid 0.192 mmol	DMF	90	90	90	8.366	8.366	11.919	P43

Zn ₆ (NDC) ₅ MOF-48	Zn(NO ₃) ₂ ·6H ₂ O 0.097 mmol 14 NDC 0.069 mmol	DMF Chloro- benzene H ₂ O ₂	90	95.902	90	19.504	16.482	14.64	C2/m
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MOF-47	Zn(NO ₃) ₂ ·6H ₂ O 0.185 mmol H ₂ (BDC[CH ₃) ₄) 0.185 mmol	DMF Chloro- benzene H ₂ O ₂	90	92.55	90	11.303	16.029	17.535	P2(1)/c
MO25	Cu(NO ₃) ₂ ·2.5H ₂ O 0.084 mmol BPhDC 0.085 mmol	DMF	90	112.0	90	23.880	16.834	18.389	P2(1)/c
Cu-Thio	Cu(NO ₃) ₂ ·2.5H ₂ O 0.084 mmol thiophene dicarboxylic acid 0.085 mmol	DEF	90	113.6	90	15.4747	14.514	14.032	P2(1)/c
CIBDC1	Cu(NO ₃) ₂ ·2.5H ₂ O 0.084 mmol H ₂ (BDCCl ₂) 0.085 mmol	DMF	90	105.6	90	14.911	15.622	18.413	C2/c
MOF-101	Cu(NO ₃) ₂ ·2.5H ₂ O 0.084 mmol BrBDC 0.085 mmol	DMF	90	90	90	21.607	20.607	20.073	Fm3m
Zn ₃ (BTC) ₂	ZnCl ₂ 0.033 mmol H ₃ BTC 0.033 mmol	DMF EtOH base added	90	90	90	26.572	26.572	26.572	Fm-3m
MOF-j	Co(CH ₃ CO ₂) ₂ ·4H ₂ O O (1.65 mmol) H ₃ (BZC) (0.95 mmol)	H ₂ O	90	112.0	90	17.482	12.963	6.559	C2

MOF-n	Zn(NO ₃) ₂ ·6H ₂ O H ₃ (BTC)	Ethanol	90	90	120	16.711	16.711	14.189	P6(3)/mcm
PbBDC	Pb(NO ₃) ₂ (0.181 mmol) H ₂ (BDC) (0.181 mmol)	DMF Ethanol	90	102.7	90	8.3639	17.991	9.9617	P2(1)/n
Znhex	Zn(NO ₃) ₂ ·6H ₂ O (0.171 mmol) H ₃ BTB (0.114 mmol)	DMF p-Xylene Ethanol	90	90	120	37.1165	37.117	30.019	P3(1)c
AS16	FeBr ₂ 0.927 mmol H ₂ (BDC) 0.927 mmol	DMF anhydr.	90	90.13	90	7.2595	8.7894	19.484	P2(1)c
AS27-2	FeBr ₂ 0.927 mmol H ₃ (BDC) 0.464 mmol	DMF anhydr.	90	90	90	26.735	26.735	26.735	Fm3m
AS32	FeCl ₃ 1.23 mmol H ₂ (BDC) 1.23 mmol	DMF anhydr. Ethanol	90	90	120	12.535	12.535	18.479	P6(2)c
AS54-3	FeBr ₂ 0.927 BPDC 0.927 mmol	DMF anhydr. n-Propa- nol	90	109.98	90	12.019	15.286	14.399	C2
AS61-4	FeBr ₂ 0.927 mmol m-BDC 0.927 mmol	Pyridine anhydr.	90	90	120	13.017	13.017	14.896	P6(2)c
AS68-7	FeBr ₂ 0.927 mmol m-BDC 1.204 mmol	DMF anhydr. Pyridine	90	90	90	18.3407	10.036	18.039	Pca2 ₁
Zn(ADC)	Zn(NO ₃) ₂ ·6H ₂ O	DMF	90	99.85	90	16.764	9.349	9.635	C2/c

	0.37 mmol H ₂ (ADC) 0.36 mmol	Chloro- benzene							
MOF-12 Zn ₂ (ATC)	Zn(NO ₃) ₂ ·6H ₂ O 0.30 mmol H ₄ (ATC) 0.15 mmol	Ethanol	90	90	90	15.745	16.907	18.167	Pbca
MOF-20 ZnNDC	Zn(NO ₃) ₂ ·6H ₂ O 0.37 mmol H ₂ NDC 0.36 mmol	DMF Chloro- benzene	90	92.13	90	8.13	16.444	12.807	P2(1)/c
MOF-37	Zn(NO ₃) ₂ ·6H ₂ O 0.20 mmol H ₂ NDC 0.20 mmol	DEF Chloro- benzene	72.38	83.16	84.33	9.952	11.576	15.556	P-1
Zn(NDC) (DMSO)	Zn(NO ₃) ₂ ·6H ₂ O H ₂ NDC	DMSO	68.08	75.33	88.31	8.631	10.207	13.114	P-1
Zn(NDC)	Zn(NO ₃) ₂ ·6H ₂ O H ₂ NDC		90	99.2	90	19.289	17.628	15.052	C2/c
Zn(HPDC)	Zn(NO ₃) ₂ ·4H ₂ O 0.23 mmol H ₂ (HPDC) 0.05 mmol	DMF H ₂ O	107.9	105.06	94.4	8.326	12.085	13.767	P-1
Co(HPDC)	Co(NO ₃) ₂ ·6H ₂ O 0.21 mmol H ₂ (HPDC) 0.06 mmol	DMF H ₂ O/ Ethanol	90	97.69	90	29.677	9.63	7.981	C2/c
Zn ₃ (PDC)2 .5	Zn(NO ₃) ₂ ·4H ₂ O 0.17 mmol H ₂ (HPDC) 0.05 mmol	DMF/ ClBz H ₂ O/ TEA	79.34	80.8	85.83	8.564	14.046	26.428	P-1

Cd ₂ (TPDC) ₂	Cd(NO ₃) ₂ ·4H ₂ O 0.06 mmol H ₂ (HPDC) 0.06 mmol	Methanol/CHP H ₂ O	70.59	72.75	87.14	10.102	14.412	14.964	P-1
Tb(PDC)1. 5	Tb(NO ₃) ₃ ·5H ₂ O 0.21 mmol H ₂ (PDC) 0.034 mmol	DMF H ₂ O/ Ethanol	109.8	103.61	100.14	9.829	12.11	14.628	P-1

ZnDBP	Zn(NO ₃) ₂ ·6H ₂ O 0.05 mmol dibenzyl phosphate 0.10 mmol	MeOH	90	93.67	90	9.254	10.762	27.93	P2/n
Zn ₃ (BPDC)	ZnBr ₂ 0.021 mmol 4,4'-BPDC 0.005 mmol	DMF	90	102.76	90	11.49	14.79	19.18	P21/n
CdBDC	Cd(NO ₃) ₂ ·4H ₂ O 0.100 mmol H ₂ (BDC) 0.401 mmol	DMF Na ₂ SiO ₃ (aq)	90	95.85	90	11.2	11.11	16.71	P21/n
Cd-mBDC	Cd(NO ₃) ₂ ·4H ₂ O 0.009 mmol H ₂ (mBDC) 0.018 mmol	DMF MeNH ₂	90	101.1	90	13.69	18.25	14.91	C2/c
Zn ₄ OBND C	Zn(NO ₃) ₂ ·6H ₂ O 0.041 mmol BNDC	DEF MeNH ₂ H ₂ O ₂	90	90	90	22.35	26.05	59.56	Fmmm
Eu(TCA)	Eu(NO ₃) ₃ ·6H ₂ O 0.14 mmol TCA 0.026 mmol	DMF Chloro- benzene	90	90	90	23.325	23.325	23.325	Pm-3n

Tb(TCA)	Tb(NO ₃) ₃ ·6H ₂ O 0.069 mmol TCA 0.026 mmol	DMF Chloro- benzene	90	90	90	23.272	23.272	23.372	Pm-3n
Formates	Ce(NO ₃) ₃ ·6H ₂ O 0.138 mmol formic acid 0.43 mmol	H ₂ O Ethanol	90	90	120	10.668	10.667	4.107	R-3m
	FeCl ₂ ·4H ₂ O 5.03 mmol formic acid 86.90 mmol	DMF	90	90	120	8.2692	8.2692	63.566	R-3c
	FeCl ₂ ·4H ₂ O 5.03 mmol formic acid 86.90 mmol	DEF	90	90	90	9.9364	18.374	18.374	Pbcn
	FeCl ₂ ·4H ₂ O 5.03 mmol formic acid 86.90 mmol	DEF	90	90	90	8.335	8.335	13.34	P-31c
NO330	FeCl ₂ ·4H ₂ O 0.50 mmol formic acid 8.69 mmol	Form- amide	90	90	90	8.7749	11.655	8.3297	Pnna
NO332	FeCl ₂ ·4H ₂ O 0.50 mmol formic acid 8.69 mmol	DIP	90	90	90	10.031 3	18.808	18.355	Pbcn

NO333	FeCl ₂ ·4H ₂ O 0.50 mmol formic acid 8.69 mmol	DBF	90	90	90	45.275 4	23.861	12.441	Cmcm
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NO335	FeCl ₂ ·4H ₂ O 0.50 mmol formic acid 8.69 mmol	CHF	90	91.372	90	11.596 4	10.187	14.945	P21/n
NO336	FeCl ₂ ·4H ₂ O 0.50 mmol formic acid 8.69 mmol	MFA	90	90	90	11.794 5	48.843	8.4136	Pbcm
NO13	Mn(Ac) ₂ ·4H ₂ O 0.46 mmol benzoic acid 0.92 mmol bipyridine 0.46 mmol	Ethanol	90	90	90	18.66	11.762	9.418	Pbcn
NO29 MOF-0 similar	Mn(Ac) ₂ ·4H ₂ O 0.46 mmol H ₃ BTC 0.69 mmol	DMF	120	90	90	14.16	33.521	33.521	P-1
Mn(hfac) ₂ (O ₂ CC ₆ H ₅)	Mn(Ac) ₂ ·4H ₂ O 0.46 mmol Hfac 0.92 mmol bipyridine 0.46 mmol	Ether	90	95.32	90	9.572	17.162	14.041	C2/c
BPR43G2	Zn(NO ₃) ₂ ·6H ₂ O 0.0288 mmol H ₂ BDC 0.0072 mmol	DMF CH ₃ CN	90	91.37	90	17.96	6.38	7.19	C2/c
BPR48A2	Zn(NO ₃) ₂ ·6H ₂ O 0.012 mmol H ₂ BDC 0.012 mmol	DMSO Toluene	90	90	90	14.5	17.04	18.02	Pbca

BPR49B1	Zn(NO ₃) ₂ 6H ₂ O 0.024 mmol H ₂ BDC 0.048 mmol	DMSO Metha- nol	90	91.172	90	33.181	9.824	17.884	C2/c
BPR56E1	Zn(NO ₃) ₂ 6H ₂ O 0.012 mmol H ₂ BDC 0.024 mmol	DMSO n- Propanol	90	90.096	90	14.587 3	14.153	17.183	P2(1)/n
BPR68D10	Zn(NO ₃) ₂ 6H ₂ O 0.0016 mmol H ₃ BTC 0.0064 mmol	DMSO Benzene	90	95.316	90	10.062 7	10.17	16.413	P2(1)/c
BPR69B1	Cd(NO ₃) ₂ 4H ₂ O 0.0212 mmol H ₂ BDC 0.0428 mmol	DMSO	90	98.76	90	14.16	15.72	17.66	Cc

BPR73E4	Cd(NO ₃) ₂ 4H ₂ O 0.006 mmol H ₂ BDC 0.003 mmol	DMSO Toluene	90	92.324	90	8.7231	7.0568	18.438	P2(1)/n
BPR76D5	Zn(NO ₃) ₂ 6H ₂ O 0.0009 mmol H ₂ BzPDC 0.0036 mmol	DMSO	90	104.17	90	14.419 1	6.2599	7.0611	Pc
BPR80B5	Cd(NO ₃) ₂ 4H ₂ O 0.018 mmol H ₂ BDC 0.036 mmol	DMF	90	115.11	90	28.049	9.184	17.837	C2/c
BPR80H5	Cd(NO ₃) ₂ 4H ₂ O 0.027 mmol H ₂ BDC 0.027 mmol	DMF	90	119.06	90	11.474 6	6.2151	17.268	P2/c

BPR82C6	Cd(NO ₃) ₂ ·4H ₂ O 0.0068 mmol H ₂ BDC 0.202 mmol	DMF	90	90	90	9.7721	21.142	27.77	Fdd2
BPR86C3	Co(NO ₃) ₂ ·6H ₂ O 0.0025 mmol H ₂ BDC 0.075 mmol	DMF	90	90	90	18.344 9	10.031	17.983	Pca2(1)
BPR86H6	Cd(NO ₃) ₂ ·6H ₂ O 0.010 mmol H ₂ BDC 0.010 mmol	DMF	80.98	89.69	83.41 2	9.8752	10.263	15.362	P-1
	Co(NO ₃) ₂ ·6H ₂ O	NMP	106.3	107.63	107.2	7.5308	10.942	11.025	P1
BPR95A2	Zn(NO ₃) ₂ ·6H ₂ O 0.012 mmol H ₂ BDC 0.012 mmol	NMP	90	102.9	90	7.4502	13.767	12.713	P2(1)/c
CuC ₆ F ₄ O ₄	Cu(NO ₃) ₂ ·2.5H ₂ O 0.370 mmol H ₂ BDC(OH) ₂ 0.37 mmol	DMF Chloro- benzene	90	98.834	90	10.967 5	24.43	22.553	P2(1)/n
Fe Formic	FeCl ₂ ·4H ₂ O 0.370 mmol formic acid 0.37 mmol	DMF	90	91.543	90	11.495	9.963	14.48	P2(1)/n
Mg Formic	Mg(NO ₃) ₂ ·6H ₂ O 0.370 mmol formic acid 0.37 mmol	DMF	90	91.359	90	11.383	9.932	14.656	P2(1)/n
MgC ₆ H ₄ O ₆	Mg(NO ₃) ₂ ·6H ₂ O 0.370 mmol H ₂ BDC(OH) ₂ 0.37 mmol	DMF	90	96.624	90	17.245	9.943	9.273	C2/c

Zn C ₂ H ₄ BDC MOF-38	ZnCl ₂ 0.44 mmol CBBDC 0.261 mmol	DMF	90	94.714	90	7.3386	16.834	12.52	P2(1)/n
MOF-49	ZnCl ₂ 0.44 mmol m-BDC 0.261 mmol	DMF CH ₃ CN	90	93.459	90	13.509	11.984	27.039	P2/c
MOF-26	Cu(NO ₃) ₂ ·5H ₂ O 0.084 mmol DCPE 0.085 mmol	DMF	90	95.607	90	20.879 7	16.017	26.176	P2(1)/n
MOF-112	Cu(NO ₃) ₂ ·2.5H ₂ O 0.084 mmol o-Br-m-BDC 0.085 mmol	DMF Ethanol	90	107.49	90	29.324 1	21.297	18.069	C2/c
MOF-109	Cu(NO ₃) ₂ ·2.5H ₂ O 0.084 mmol KDB 0.085 mmol	DMF	90	111.98	90	23.880 1	16.834	18.389	P2(1)/c
MOF-111	Cu(NO ₃) ₂ ·2.5H ₂ O 0.084 mmol o-BrBDC 0.085 mmol	DMF Ethanol	90	102.16	90	10.676 7	18.781	21.052	C2/c
MOF-110	Cu(NO ₃) ₂ ·2.5H ₂ O 0.084 mmol thiophene dicarboxylic acid 0.085 mmol	DMF	90	90	120	20.065 2	20.065	20.747	R-3/m
MOF-107	Cu(NO ₃) ₂ ·2.5H ₂ O 0.084 mmol thiophene dicarboxylic acid 0.085 mmol	DEF	104.8	97.075	95.20 6	11.032	18.067	18.452	P-1

MOF-108	Cu(NO ₃) ₂ ·2.5H ₂ O 0.084 mmol thiophene dicarboxylic acid 0.085 mmol	DBF/ Methanol	90	113.63	90	15.474 7	14.514	14.032	C2/c
MOF-102	Cu(NO ₃) ₂ ·2.5H ₂ O 0.084 mmol H ₂ (BDCCl ₂) 0.085 mmol	DMF	91.63	106.24	112.0 1	9.3845	10.794	10.831	P-1
Clbdc1	Cu(NO ₃) ₂ ·2.5H ₂ O 0.084 mmol H ₂ (BDCCl ₂) 0.085 mmol	DEF	90	105.56	90	14.911	15.622	18.413	P-1
Cu(NMOP)	Cu(NO ₃) ₂ ·2.5H ₂ O 0.084 mmol NBDC 0.085 mmol	DMF	90	102.37	90	14.923 8	18.727	15.529	P2(1)/m
Tb(BTC)	Tb(NO ₃) ₃ ·5H ₂ O 0.033 mmol H ₃ BTC 0.033 mmol	DMF	90	106.02	90	18.698 6	11.368	19.721	
Zn ₃ (BTC) ₂	ZnCl ₂ 0.033 mmol H ₃ BTC 0.033 mmol	DMF Ethanol	90	90	90	26.572	26.572	26.572	Fm-3m
Zn ₄ O(NDC)	Zn(NO ₃) ₂ ·4H ₂ O 0.066 mmol 14NDC 0.066 mmol	DMF Ethanol	90	90	90	41.559 4	18.818	17.574	aba2
CdTDC	Cd(NO ₃) ₂ ·4H ₂ O 0.014 mmol thiophene 0.040 mmol DABCO 0.020 mmol	DMF H ₂ O	90	90	90	12.173	10.485	7.33	Pmma

IRMOF-2	Zn(NO ₃) ₂ ·4H ₂ O 0.160 mmol o-Br-BDC 0.60 mmol	DEF	90	90	90	25.772	25.772	25.772	Fm-3m
IRMOF-3	Zn(NO ₃) ₂ ·4H ₂ O 0.20 mmol H ₂ N-BDC 0.60 mmol	DEF Ethanol	90	90	90	25.747	25.747	25.747	Fm-3m

IRMOF-4	Zn(NO ₃) ₂ ·4H ₂ O 0.11 mmol [C ₃ H ₇ O] ₂ -BDC 0.48 mmol	DEF	90	90	90	25.849	25.849	25.849	Fm-3m
IRMOF-5	Zn(NO ₃) ₂ ·4H ₂ O 0.13 mmol [C ₅ H ₁₁ O] ₂ -BDC 0.50 mmol	DEF	90	90	90	12.882	12.882	12.882	Pm-3m
IRMOF-6	Zn(NO ₃) ₂ ·4H ₂ O 0.20 mmol [C ₂ H ₄]-BDC 0.60 mmol	DEF	90	90	90	25.842	25.842	25.842	Fm-3m
IRMOF-7	Zn(NO ₃) ₂ ·4H ₂ O 0.07 mmol 1,4NDC 0.20 mmol	DEF	90	90	90	12.914	12.914	12.914	Pm-3m
IRMOF-8	Zn(NO ₃) ₂ ·4H ₂ O 0.55 mmol 2,6NDC 0.42 mmol	DEF	90	90	90	30.092	30.092	30.092	Fm-3m
IRMOF-9	Zn(NO ₃) ₂ ·4H ₂ O 0.05 mmol BPDC 0.42 mmol	DEF	90	90	90	17.147	23.322	25.255	Pnnm

IRMOF-10	Zn(NO ₃) ₂ ·4H ₂ O 0.02 mmol BPDC 0.012 mmol	DEF	90	90	90	34.281	34.281	34.281	Fm-3m
IRMOF-11	Zn(NO ₃) ₂ ·4H ₂ O 0.05 mmol HPDC 0.20 mmol	DEF	90	90	90	24.822	24.822	56.734	R-3m
IRMOF-12	Zn(NO ₃) ₂ ·4H ₂ O 0.017 mmol HPDC 0.12 mmol	DEF	90	90	90	34.281	34.281	34.281	Fm-3m

IRMOF-13	Zn(NO ₃) ₂ ·4H ₂ O 0.048 mmol PDC 0.31 mmol	DEF	90	90	90	24.822	24.822	56.734	R-3m
IRMOF-14	Zn(NO ₃) ₂ ·4H ₂ O 0.17 mmol PDC 0.12 mmol	DEF	90	90	90	34.381	34.381	34.381	Fm-3m
IRMOF-15	Zn(NO ₃) ₂ ·4H ₂ O 0.063 mmol TPDC 0.025 mmol	DEF	90	90	90	21.459	21.459	21.459	Im-3m
IRMOF-16	Zn(NO ₃) ₂ ·4H ₂ O 0.0126 mmol TPDC 0.05 mmol	DEF NMP	90	90	90	21.49	21.49	21.49	Pm-3m

ADC Acetylenedicarboxylic acid

NDC Naphthalenedicarboxylic acid

5 BDC Benzenedicarboxylic acid

ATC Adamantanetetracarboxylic acid

BTC Benzenetricarboxylic acid

BTB	Benzenetribenzoic acid
MTB	Methanetetra benzoic acid
ATB	Adamantanetetra benzoic acid
ADB	Adamantanedibenzoic acid

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Further metal-organic framework materials are MOF-2 to 4, MOF-9, MOF-31 to 36, MOF-39, MOF-69 to 80, MOF103 to 106, MOF-122, MOF-125, MOF-150, MOF-177, MOF-178, MOF-235, MOF-236, MOF-500, MOF-501, MOF-502, MOF-505, IRMOF-1, IRMOF-61, IRMOP-13, IRMOP-51, MIL-17, MIL-45, MIL-47, MIL-53, MIL-59, MIL-60, MIL-61, MIL-63, MIL-68, MIL-79, MIL-80, MIL-83, MIL-85, CPL-1 to 2, SZL-1 which are described in the literature.

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In particular preference is given to a porous metal-organic framework material in which Zn, Al or Cu is present as metal ion and the at least bidentate, organic compound is terephthalic acid, isophthalic acid, 2,6-naphthalenedicarboxylic acid or 1,3,5-benzenetricarboxylic acid.

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In addition to the conventional method for production of MOFs, as described, for example, in US 5,648,508, they can also be produced by the electrochemical route. In this respect, reference is made to DE-A 103 55 087 and also WO-A 2005/049892. The MOFs produced in this way exhibit particularly good properties in relation to adsorption and desorption of chemical substances, in particular gases. They thus differ from those which are produced conventionally, even when these are formed from the same organic and metal ion constituents and are therefore to be considered novel framework materials. In the context of the present invention, electrochemically produced MOFs are particularly preferred.

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Consequently, the electrochemical production relates to a crystalline porous metal-organic framework material comprising at least one at least bidentate, organic compound which is bound by coordination to at least one metal ion and which is obtained in a reaction medium comprising the at least one bidentate organic compound by at least one metal ion being generated by oxidation of at least one anode comprising the corresponding metal.

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The term "electrochemical production" designates a production method in which the formation of at least one reaction product is associated with the migration of electric charges or the occurrence of electric potentials.

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The term "at least one metal ion", as used in connection with the electrochemical

production, designates embodiments according to which at least one ion of a metal or at least one ion of a first metal and at least one ion of at least one second metal different from the first metal are provided by anodic oxidation.

- 5 Consequently, the electrochemical production comprises embodiments in which at least one ion of at least one metal is provided by anodic oxidation and at least one ion of at least one metal is provided via a metal salt, the at least one metal in the metal salt and the at least one metal which is provided as metal ion via anodic oxidation can be identical or different from one another. Therefore the present invention, with respect to
- 10 electrochemically produced MOFs, comprises, for example, an embodiment according to which the reaction medium comprises one or more different salts of a metal and the metal ion present in this salt or in these salts is additionally provided by anodic oxidation of at least one anode comprising this metal. Likewise, the reaction medium can comprise one or more different salts of at least one metal and at least one metal different from these metals
- 15 can be provided by anodic oxidation as metal ion in the reaction medium.

According to a preferred embodiment of the present invention in connection with the electrochemical production, the at least one metal ion is provided by anodic oxidation of at least one anode comprising this at least one metal, though no further metal being provided

20 via a metal salt.

The term "metal", as used in the context of the present invention in connection with the electrochemical production of MOFs, comprises all elements of the Periodic Table of the Elements which can be provided via anodic oxidation via the electrochemical route in a

25 reaction medium and together with at least one at least bidentate, organic compound are able to form at least one metal-organic porous framework material.

Independently of its production, the resultant MOF occurs in pulverulent form or as agglomerate. This can be used as such as sorbent in the inventive method alone or

30 together with other sorbents or further materials. Preferably this occurs as bulk material, in particular in a fixed bed. In addition, the MOF can be converted into a shaped body. Preferred methods in this case are rod extrusion or tableting. In the shaped body production, further materials, such as, for example, binders, lubricants or other additives, can be added to the MOF. Likewise, it is conceivable that mixtures of MOF and other

35 adsorbents, for example activated carbon, are produced as shaped bodies or separately result in shaped bodies which are then used as shaped body mixtures.

With respect to the possible geometries of these MOF shaped bodies, there are essentially no restrictions. For example, mention may be made of, inter alia, pellets, such as, for example, disk-shaped pellets, pills, spheres, granules, extrudates, for example rod extrudates, honeycombs, meshes or hollow bodies.

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For the production of these shaped bodies, in principle all suitable methods are possible. In particular, the following procedures are preferred:

- 10 - kneading the framework material alone or together with at least one binder and/or at least one pasting agent and/or at least one template compound to obtain a mixture; shaping the resultant mixture by means of at least one suitable method, such as, for example, extrusion; optionally washing and/or drying and/or calcining the extrudate; optionally final processing.
- 15 - Applying the framework material to at least one, if appropriate, porous support material. The resultant material can then be further processed to give a shaped body in accordance with the above described method.
- 20 - Applying the framework material to at least one, if appropriate, porous substrate.
- Foaming into porous plastics, such as, e.g. polyurethane.

25 Kneading and shaping can proceed according to any suitable method, such as, for example, as described in Ullmanns Enzyklopädie der Technischen Chemie [Ullmann's Encyclopedia of Industrial Chemistry], 4th edition, volume 2, pp. 313 ff. (1972), the contents of which in this respect are hereby incorporated in entirety by reference into the context of the present application.

30 For example, preferably, the kneading and/or shaping can proceed by means of a piston press, roller press in the presence or absence of at least one binder, compounding, pelleting, tableting, extrusion, co-extrusion, foaming, spinning, coating, granulating, preferably spray-granulating, spraying, spray-drying or a combination of two or more of these methods.

35 Very particularly, pellets and/or tablets are produced.

The kneading and/or shaping can proceed at elevated temperatures, such as, for example,

in the range from room temperature to 300°C and/or at elevated pressure, such as, for example, in the range from atmospheric pressure up to a few 100 bar and/or in a protective gas atmosphere such as, for example, in the presence of at least one noble gas, nitrogen, or a mixture of two or more thereof.

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The kneading and/or shaping is carried out according to a further embodiment with addition of at least one binder, as binder, use being able to be made in principle of any chemical compound which ensures the viscosity of the mix to be kneaded and/or shaped desired for kneading and/or shaping. Consequently, binders in the context of the present invention can be not only viscosity-increasing compounds, but also viscosity-reducing compounds.

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As binders which are preferred, inter alia, mention may be made of, for example, aluminum oxide or aluminum oxide-containing binders, that are described, for example, in WO 94/29408, silicon dioxide as described, for example, in EP 0 592 050 A1, mixtures of silicon dioxide and aluminum oxide as are described, for example, in WO 94/13584, clay minerals; as are described, for example, in JP 03-037156 A, for example montmorillonite, kaolin, bentonite, hallosite, dickite, nacrite and anauxite, alkoxysilanes, as are described, for example, in EP 0 102 544 B1, for example tetraalkoxysilanes such as, for example, tetramethoxysilane, tetraethoxysilane, tetrapropoxysilane, tetrabutoxysilane, or, for example, trialkoxysilanes such as, for example, trimethoxysilane, triethoxysilane, tripropoxysilane, tributoxysilane, alkoxytitanates, for example tetraalkoxytitanates, such as, for example, tetramethoxytitanate, tetraethoxytitanate, tetrapropoxytitanate, tetrabutoxytitanate, or, for example, trialkoxytitanates, such as, for example, trimethoxytitanate, triethoxytitanate, tripropoxytitanate, tributoxytitanate, alkoxyzirconates, for example tetraalkoxyzirconates, such as, for example, tetramethoxyzirconate, tetraethoxyzirconate, tetrapropoxyzirconate, tetrabutoxyzirconate, or, for example, trialkoxyzirconates such as, for example, trimethoxyzirconate, triethoxyzirconate, tripropoxyzirconate, tributoxyzirconate, silica sols, amphiphilic substances and/or graphites. In particular preference is given to graphite.

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As viscosity-increasing compound, use can also be made of, for example, if appropriate in addition to the abovementioned compounds, an organic compound and/or a hydrophilic polymer such as, for example, cellulose or a cellulose derivative such as, for example, methylcellulose and/or a polyacrylate and/or a polymethacrylate and/or a polyvinyl alcohol and/or a polyvinylpyrrolidone and/or a polyisobutene and/or a polytetrahydrofuran.

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As pasting agent, use can be made of, inter alia, preferably water or at least one alcohol

such as, for example, a monohydric alcohol having 1 to 4 carbon atoms such as, for example, methanol, ethanol, n-propanol, isopropanol, 1-butanol, 2-butanol, 2-methyl-1-propanol or 2-methyl-2-propanol or a mixture of water and at least one of said alcohols or a polyhydric alcohol such as, for example, a glycol, preferably a water-miscible polyhydric alcohol, alone or a mixture with water and/or at least one of said monohydric alcohols.

Further additives which can be used for the kneading and/or shaping are, inter alia, amines or amine derivatives such as, for example, tetraalkylammonium compounds or aminoalcohols and carbonate-containing compounds such as calcium carbonate. Such further additives are described, for instance in EP 0 389 041 A1, EP 0 200 260 A1 or WO 95/19222.

The sequence of the additives such as template compound, binder, pasting agent, viscosity-increasing substance in the shaping and kneading is in principle not critical.

According to a further preferred embodiment, the shaped body obtained according to kneading and/or shaping is subjected to at least one drying, which is generally carried out at a temperature in the range from 25 to 300°C, preferably in the range from 50 to 300°C, and particularly preferably in the range from 100 to 300°C. Likewise, it is possible to dry in a vacuum or under a protective gas atmosphere or by spray drying.

According to a particularly preferred embodiment, in the context of this drying operation, at least one of the compounds added as additives is at least in part removed from the shaped body.

Example

Figure 1 shows the adsorption isotherms of CO₂ at 20°C to the metal-organic framework material aluminum terephthalate in the form of 3×3 mm tablets, P giving the absolute pressure in mbar and A giving the amount of adsorbed gas in mg per g of adsorbent.

At an exemplary diving depth of 20 m (3 bar), the partial pressure of CO₂ (4%×3 bar) is 120 mbar (see point 1 in Fig. 1). This corresponds to a loading of 24 to 27 mg/g (see point 2 in Fig. 1).

At the surface, for example in the dive center, fresh air is blown through the adsorber bed in countercurrent flow using the conventional compressor. The CO₂ partial pressure is

(0.03%×1 bar) 0.3 mbar (see point 3 in Fig. 1).

This means that 1 mol of CO₂ can be adsorbed using about 2 kg of framework material.

Patent claims

1. A method for removing carbon dioxide and, if appropriate, water from breathing air in closed or partially closed systems comprising the step
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 - contacting the breathing air with a porous metal-organic framework material, the framework material comprising at least one at least bidentate organic compound which is bound by coordination to at least one metal ion.
- 10 2. The method according to claim 1, wherein the closed or partially closed system comprises at least one breathing apparatus and also a breathing mask, a breathing suit or other life support system.
- 15 3. The method according to claim 1 or 2, wherein the porous metal-organic framework material is present as a shaped body.
4. The method according to one of claims 1 to 3, wherein the porous metal-organic framework material is at least part of an adsorber bed in a filter.
- 20 5. The method according to one of claims 1 to 4, wherein the at least one metal ion is an ion selected from the group consisting of aluminum, zinc and copper.
6. The method according to one of claims 1 to 5, wherein the at least bidentate organic compound is derived from a di-, tri-, tetracarboxylic acid, or their sulfur analogs.
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7. A closed or partially closed system which comprises at least one breathing apparatus and also a breathing mask, a breathing suit or other life support system, in addition comprising a porous metal-organic framework material, the framework material comprising at least one, at least bidentate, organic compound which is bound by coordination to at least one metal ion.
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8. The closed or partially closed system according to claim 7, wherein this in addition has a filter, in which the porous metal-organic framework material is present at least as part of an adsorber bed.
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9. The closed or partially closed system according to claim 8, wherein the filter is exchangeable or is installed fixed in the system.

10. The use of a closed or partially closed system according to one of claims 7 to 9 for removing carbon dioxide and, if appropriate, water from breathing air.
- 5 11. A method for regenerating a porous metal-organic framework material from a closed or partially closed system according to one of claims 7 to 9 comprising the steps
- if appropriate removing the metal-organic framework material; and
 - impinging the framework material with a gas.
- 10 12. The method according to claim 11, wherein the gas is air, nitrogen, an inert gas, or a mixture thereof.
13. The method according to claim 11 or 12, wherein the impingement is together with the change of at least one parameter selected from pressure and temperature.
- 15 14. The method according to one of claims 11 to 13, wherein the regeneration of the metal-organic framework material is performed during the use of the closed or partially closed system according to one of claims 7 to 9.