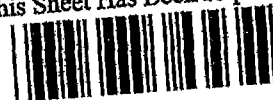


9609391-9
1205 wny

- 1 -

INCREASING VI OF LUBE OIL BY HYDROTREATING USING
BULK Ni Mn Mo OR Ni Cr Mo SULFIDE CATALYSTS
PREPARED FROM LIGATED METAL COMPLEXES

This Sheet Has Been Replaced



Description of the Invention

The Viscosity Index of lubricating and speciality oils can be increased by hydrotreating said oil employing as the hydrotreating catalyst a bulk Ni/Mn/Mo or Cr/Ni/Mo sulfide catalyst prepared from ligated, e.g. ethylene diamine metal complexes wherein the Ni/Mn/Mo sulfide catalyst is prepared from the oxide precursor with subsequent sulfiding while the Cr/Ni/Mo sulfide catalyst is prepared from the sulfide precursor in the presence of H_2S/H_2 .

Background of the Invention

Raffinate hydrotreating and hydrocracking processes to produce lube stocks are well known. In general, these processes employ catalysts which comprise metal oxides or sulfides supported on alumina or silica-alumina. Metals usually include nickel, molybdenum, cobalt, manganese, tungsten, etc. and mixtures thereof in the oxide or sulfide state. When the catalyst is in the oxide state the environment of use being one containing sulfur containing species converts the metal oxides into metal sulfides.

The objectives in the hydrotreater are to increase VI and saturates levels. Catalysts used today in lubes hydrotreating do not perform this function very well insofar as they are employed principally for hydrodesulfurization, hydrodenitrogenation, color improvement and residual extraction solvent removal. High pressure hydrocracking of distillates or deasphalted oil is an acknowledged alternative to solvent extraction and, in the case of distillate hydrocracking all of the VI increase, saturates increase and aromatics decrease are achieved by the catalytic conversion.

The major reactions which contribute to VI increase are of the ring opening type, whether the ring is aromatic or cyclic saturated. Hydrogenation will contribute to a lesser degree to VI

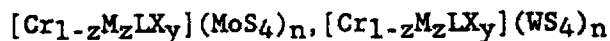
THIS SHEET HAS BEEN
REPLACED



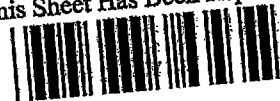
increase. At low severity, hydrosulfurization will give a 3-4 VI unit increase, most likely the result of ring opening that occurs with the hydrosulfurization of thiophenic structures.

US 4,650,563 teaches molybdenum and/or tungsten hydrotreating catalysts prepared from promoter metal thiomolybdate and/or thiotungstate complexes and their use as hydrotreating catalysts. The catalysts are prepared by heating one or more catalyst precursors under oxygen-free conditions in the presence of sulfur at a temperature of at least about 150°C. The catalyst precursors will be one or more compounds of the formula $(ML)(Mo_y W_{1-y} S_4)$ wherein M comprises one or more divalent promoter metals selected from the group consisting of Ni, Co, Zn, Cu and mixtures thereof, wherein y is any value ranging from 0 to 1 and wherein L is one or more neutral, nitrogen containing ligands at least one of which is a chelating polydentate ligand. L may be ethylene diamine.

US 4,716,139 teaches a catalyst composition comprising an amorphous sulfide of trivalent chromium and at least one metal selected from the group consisting of (a) molybdenum, tungsten and mixtures thereof and (b) mixtures of (a) with iron, nickel, cobalt, manganese, copper, zinc and mixtures thereof. The catalyst is prepared by heating one or more precursors at a temperature of at least about 200°C under oxygen-free conditions in the presence of sulfur, wherein said precursor is selected from the group consisting of



and mixtures thereof wherein M is one or more divalent promoter metals selected from the group consisting of Mn, Fe, Co, Ni, Cu, Zn and mixtures thereof, wherein L is one or more neutral nitrogen-containing ligands, wherein ligand X is a singly-charged anionic species such as NO_2^- , OH^- , Cl^- , etc. and wherein $1 > z \geq 0.1 - z \geq y \geq 0$ and wherein n equals $(3-z-y)/2$. The catalysts are good hydroprocessing catalysts, particularly hydrotreating catalysts. These catalysts were demonstrated for their HDS and HDN ability on light catalytic cycle oil (LCCO).



See also U.S. Patent 4,626,339 which teaches the supported version of the catalyst of '139.

U.S. 4,663,023 teaches hydrotreating with self promoted molybdenum and Tungsten sulfide catalysts. Catalysts are made from precursors which are one or more compounds of the formula $ML(Mo_yW_{1-y}O_4)$ wherein M is one or more promoter metals selected from the group consisting essentially of Mn, Fe, Co, Ni, Cu, Zn and mixtures thereof $0 \leq y \leq 1$ and L is a neutral nitrogen containing ligand. This patent would, therefor, embrace Mn/Ni/Mo sulfide and Ni/Mo sulfide catalysts. The patent claims the use of these catalysts in general to hydrotreat hydrocarbon feeds including lube oils.

U.S. 4,698,145 teaches the upgrading of hydrocarbon feeds by use of a supported self-promoted catalyst. The catalyst is prepared by heating a composite of support material and one or more catalyst precursor salts under oxygen-free conditions and in the presence of sulfur. The precursor salt or salts are of the general formula $(ML)(Mo_yW_{1-y}S_4)$ wherein M is one or more promoter metals selected from the group consisting essentially of Mn, Fe, Co, Ni, Zn and mixtures thereof, $0 \leq y \leq 1$, and L is a neutral, nitrogen containing ligand. This patent would, therefore, embrace supported Mn/Ni/Mo sulfide and supported Ni/Mo sulfide catalysts. The claims recite lube oil as one of the hydrocarbon feeds treated using this class of catalysts in general.

U.S. 4,595,672 is directed to a method for making self-promoted sulfide hydrotreating catalysts. The catalysts are made from precursors which are one or more compounds of the formula $ML(Mo_yW_{1-y}O_4)$ where M is one or more promoter metals selected from the group consisting essentially of Mn, Fe, Co, Ni, Cu, Zn and mixtures thereof an $0 \leq y \leq 1$. This patent would, therefore, embrace Ni/Mn/Mo sulfide and Ni/Mo sulfide catalytic materials. L is a neutral nitrogen containing ligand.

U.S. 4,668,376 teaches supported Mn sulfide promoted Mo and W sulfide hydroprocessing catalysts. The catalysts are made from

thiometallate precursor salts of the formula $(ML)(Mo_yW_{1-y}S_4)$ where M comprises one or more divalent promoter metals selected from Mn and mixtures of Mn with one or more of Ni, Co, Cu, Zn or mixtures thereof any $0 \leq y \leq 1$ and L is a neutral nitrogen containing ligand. This patent would, therefore, embrace Ni/Mn/Mo sulfide and Ni/Mo sulfide catalytic materials. The patent claims the use of these catalysts in general for hydrorefining hydrocarbon feeds including lube oils.

U.S. 4,591,429 teaching hydroprocessing of hydrocarbon feeds employing supported sulfided catalyst. The catalyst comprises a sulfide of (1) chromium, (2) Mo, W or mixtures thereof and (3) at least one metal selected from the group consisting of Ni, Co, Mn, Cu, Zn and mixtures thereof, and mixtures thereof with Fe. The catalysts are made by compositing a pre-selected quantity of support material with a precursor comprising a mixture of (1) hydrated oxide of trivalent chromium and (2) a salt containing a thiometallate anion of Mo or W and a cation comprising at least one divalent promoter metal. This patent would, therefore, embrace supported Cr/Ni/Mo sulfide catalyst and claims the use of these catalyst in general for hydroprocessing hydrocarbon feeds, including lube oil.

Description of The Figure

Figure 1 compares the $360^\circ C^+$ product VI as a function of reactor temperature for five bulk metal catalysts and a Ni/W/F/Al₂O₃ reference catalyst.

Detailed Description of the Invention

Lubricating and/or specialty oil stocks such as deasphalted oil stocks, lube oil distillates, and solvent extracted lube oil raffinates can have their viscosity indexes increased by hydrotreating, employing specific bulk metal sulfide hydrotreating catalysts selected from the group consisting of bulk Cr/Ni/Mo sulfide catalyst, bulk Ni/Mo/Mn sulfide catalyst and mixtures thereof wherein the catalysts are prepared from specific metal complexes and wherein the Ni/Mn/Mo sulfide catalyst is prepared from the oxide precursor



decomposed in an inert atmosphere such as N_2 and subsequently sulfided using H_2S/H_2 and the Cr/Ni/Mo sulfide catalyst is prepared from the sulfide precursor and decomposed in a non-oxidizing, sulfur containing atmosphere.

These two catalyst have been found to significantly and, surprisingly, to a higher degree than other catalysts, increase the Viscosity Index of lube oil stocks. The feed stocks are contacted with the catalyst under hydrotreating conditions comprising temperatures in the range of about 200 to $450^\circ C$, preferably about 300 to $380^\circ C$, pressures in the range of about 300 to 3000 psi, preferably at least about 1000 psi, space velocity in the range of about 0.1 to 5.0 v/v/hr., preferably about 0.5 to 2.0 v/v/hr. and a treat gas rate in the range of about 200 to 20,000 SCF H_2 /bbl, preferably about 700 to 5,000 SCF H_2 /bbl.

The catalysts are prepared by using $[Ni(L)_3](Mn(L)_3)[MoO_4]_2$ or $[Ni(L)_3][Cr(L)_3][MoS_4]_{2-2.5}$ as precursor complexes, where L = one or more neutral, nitrogen containing ligands at least one of which is a chelating polydentate ligand such as ethylenediamine. These complexes are prepared by reaction of appropriate metal salts with the ligand solution.

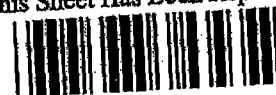
The ligand or ligands L will generally have a denticity of six and will be one or more neutral, nitrogen containing ligands wherein at least one of said ligands is a multi-dentate chelating ligand which chelates the metal cation of the metal salt to form a chelated metal species such as $[ML]^{2+}$, where M is Ni or Cr and Mn.

Those skilled in the art know that the term "ligand" is used to designate functional coordinating groups which have one or more pairs of electrons available for the formation of coordinate bonds. Ligands that can form more than one bond with a metal ion are called polydentate while ligands that can form only one bond with a metal ion are called monodentate. Monodentate ligands are not capable of forming chelates. Hence, if one uses one or more species of monodentate ligands in the precursor molecule, then one must also use



at least one polydentate chelating ligand. Preferably L will be one or more polydentate chelating ligands. The denticity of the ligand L will generally be six, because the promoter metal cations prefer six-fold coordination. Hence, if more than one species of ligand is employed in the precursor molecule, the denticity of the ligand species will usually add up to six. It should be understood that it is possible for ligand L to have a total denticity of less than six, but in most cases L will have a total denticity of six. Thus, L will be three bidentate ligands, two tridentate ligands, a mixture of a bidentate and a quadridentate ligand, a hexadentate ligand or a mixture of a polydentate ligand with monodentate ligands so long as the combination has a total denticity of six. As has heretofore been stated, it is preferred to use chelating bidentate and tridentate ligands. In general, the ligands useful in this invention include alkyl and aryl amines and nitrogen heterocycles. Illustrative but non-limiting examples of ligands useful in the catalyst precursors of this invention are set forth below.

Monodentate ligands will include NH_3 as well as alkyl and aryl amines such as ethyl amine, dimethyl amine, pyridine, etc. Useful chelating bidentate amine ligands are illustrated by ethylenediamine, 2,2'-bipyridine, o-phenylene diamine, tetramethylethylenediamine and propane-1,3 diamine. Similarly, useful chelating tridentate amine ligands are represented by terpyridine and diethylenetriamine while triethylenetetramine is illustrative of a useful chelating quadridentate amine ligands. Useful chelating pentadentate ligands include tetraethylenepentamine while sepulchrate (an octazacryptate) is illustrative of a suitable chelating hexadentate ligand. However, as a practical matter it will be preferred to use chelating, polydentate alkyl amines for L. Illustrative, but not limiting examples of alkyl amines that are useful in the catalyst precursor of this invention include ethylenediamine, diethylenetriamine, and tetraethylenetetramine. It is particularly preferred to use bidentate and tridentate alkyl amines such as ethylenediamine, (en) and diethylenetriamine, (dien).

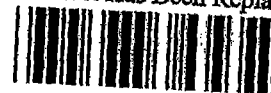


Many of the precursor salts useful in forming the catalysts of this invention and methods for preparing them are known in the art. An article by Diemann and Mueller titled Thio and Seleno Compounds of the Transition Metals With d^0 configuration published in COORD. CHEM. REV. 10:79-122 provides a review of known salts. In general, the precursor salts useful for forming the catalysts useful in this invention may be prepared by mixing an aqueous solution of ammonium thiomolybdate with an aqueous solution of the chelated promoter metal cation $[ML]^{2+}$ which results in the formation of the precursor salt as a precipitate which is readily recovered. The chelating promoter cation is easily formed by, for example, mixing an aqueous solution of one or more water soluble promoter metal salts with the ligand or mixture of ligands. The water soluble salt may be any water soluble salt that is convenient to use such as a halide, sulfate, perchlorate, acetate, nitrate etc. Alternatively, an aqueous solution of ammonium thiomolybdate may be mixed with the ligand with the resulting solution mixed with an aqueous solution of promoter metal salt or the salt can be added to the ligand and dissolved into the solution of thiomolybdate and/or thiotungstate. However, it should be understood that the catalyst precursor preparation is not intended to be limited to aqueous media.

The $[Ni(L)_3](Mn(L)_3)[MoO_4]_2$ precursor is converted into the bulk catalyst by decomposing the precursor complex under inert conditions, such as in the presence of N_2 or an inert gas at 300 to 400°C. The decomposed material is subsequently sulfided using H_2S/H_2 at 200 to 600°C, preferably 250-500°C, most preferably 200-400°C. A preferred ligand is ethylene diamine.

The $[Ni(L)_3](Cr(L)_3)[MoS_4]_{2-2.5}$ precursor is converted into the bulk sulfide catalyst by decomposing the precursor complex in a sulfur containing atmosphere at least 150°C, preferably at least about 200°C for a time sufficient to form the catalyst. A preferred ligand is ethylenediamine.

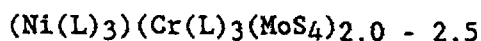
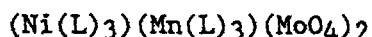
The sulfur required during the formation of the catalyst may be that which is present in the precursor salt. Thus, the Ni/Cr/Mo



catalyst composition used in this invention will be formed if no excess sulfur is present and if the oxygen-free atmosphere is relatively inert, such as nitrogen. It is preferred that the sulfur will be present in an amount in excess of that contained in the precursor salt. It is preferred that the catalyst be formed by heating the precursor in the presence of excess sulfur, it is also preferred that the excess sulfur be present in the form of a sulfur bearing compound which can be one or more solids, liquids, gases or mixtures thereof. Mixtures of hydrogen and H_2S have been found to be particularly suitable. In general the temperature will range between from about 150° - $600^{\circ}C$, preferably 250 - $500^{\circ}C$, and still more preferably from about 300° - $400^{\circ}C$. The non-oxidizing atmosphere may be gaseous, liquid or mixture thereof.

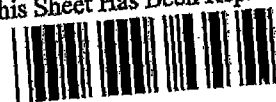
These bulk metal sulfide catalysts are relatively low surface area materials of about $50m^2/gram$.

The catalyst is prepared in a two step operation. Appropriate salts of Ni, Cr, and Mn are reacted in the presence of excess ligand in combination with MoS_4^{2-} or MoO_4^{2-} . Methanol is used as the solvent. One recovers either



which are then decomposed as previously described.

The bulk metal sulfide catalysts are most conveniently used as shaped pellets or extrudates. The bulk metal sulfide can be pelleted by the application of extremely high pressure. Bulk metal sulfide catalyst pellets have been prepared using what is believed to be this high pressure technique but the details are not known. Another technique involves preparing the bulk metal sulfide in the presence of colloidal silica or colloidal alumina, recovering the precipitate, drying, and grinding the recovered cake and pelletizing using polyvinyl alcohol, in accordance with the procedure of U.S. Patent 4,650,563 and U.S. Patent 4,591,429. It is possible that a preferred bulk metal sulfide catalyst pellet or extrudate is produced without the use of any silica or alumina solely by employing standard,



commercial high pressure pelleting, but economics and cost effectiveness will become important issues.

Experimental

Five bulk metal sulfide catalysts were evaluated for raffinate hydrotreating to determine their effect on VI improvement. These catalysts were:

- 1) $\text{Ni}_{0.5} \text{Cr}_{0.5} \text{Mo}$ (from thiomolybdate)
- 2) $\text{Ni}_{0.5} \text{Mn}_{0.5} \text{Mo}$ (from thiomolybdate)
- 3) NiW (from thiotungstate)
- 4) $\text{Ni}_{0.5} \text{Mn}_{0.5} \text{Mo}$ (from molybdate)
- 5) CoMo (from molybdate)

These bulk materials were pelleted using high pressure without any silica or alumina binder material. The pelleted material was then crushed and sieved to secure particles of about 1/16 inch.

The following are representative of typical procedures used to produce materials of the type identified above.

They are not offered as recitations of the exact, specific procedures used to produce the particular materials hereinafter identified.

Typical Preparation of $\text{NiMn}(\text{NH}_2\text{CH}_2\text{CH}_2\text{NH}_2)_6(\text{MoO}_4)_2$ with subsequent sulfiding

88.2g $(\text{NH}_4)_6\text{Mo}_7\text{O}_{24} \cdot 4\text{H}_2\text{O}$ were dissolved in 200 ml H_2O plus 100 ml ethylenediamine. 72.2g $\text{Ni}(\text{NO}_3)_2 \cdot 6\text{H}_2\text{O}$ and 49.5g $\text{MnCl}_2 \cdot 4\text{H}_2\text{O}$ were dissolved in 800 ml deaerated (N_2 bubbler) methanol. The molybdate solution was placed in a 2-liter flask and deaerated with a nitrogen bubbler while stirring with a mechanical stirrer. Under a blanket of N_2 , the Ni- and Mn- containing solution was added dropwise to the flask from a dropping funnel. A precipitate formed. The precipitate was allowed to settle and the supernatant liquid was siphoned off. Two methanol (deaerated) rinsings were performed, each time siphoning off supernatant-liquid. The precipitate was finally recovered by vacuum filtration from methanol in an N_2 atmosphere box. The yield



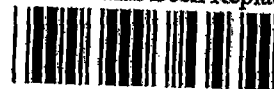
was 87.2%. The catalyst precursor was dried under vacuum at 50°C overnight.

The precursor was then pelletized into a disk and then crushed into 20/40 mesh granules. Sulfiding of the precursor compound to bulk catalysts was done by heating under nitrogen for 50 minutes to 370°C, holding at 370°C for one hour, then heating under a stream of 10% H₂S in hydrogen for ten minutes to 400°C and holding at that temperature under H₂S/H₂ for two hours. Cooling over 1.5 hrs under H₂S/H₂ to <100°C was followed by cooling to room temperature under nitrogen and passivation at room temperature with a stream of 1% O₂/N₂ for half an hour.

Typical Preparation of Cr-Containing Catalysts

Chromium containing catalyst can be prepared via two different chemistries, ololation and chelation in non-rigorous terms. The following reported procedures are representative of the typical techniques employed in producing the material from the thiomolybdate.

Via Ololation: A Ni_{0.5}-Cr_{0.5}-Mo catalyst can be prepared as follows: 40g (NH₄)₂MoS₄ was added to 82 ml of diethylenetriamine (dien) in a one liter flask. Distilled H₂O was used twice to wash off any solution remaining on the sides of the flask. The resulting dark red solution was cooled to 0°C in an ice bath and kept in the bath for the duration of the preparation. In a separate flask, 16.52 grams of CrCl₃·6H₂O and 14.77 gm of NiCl₂·6H₂O were dissolved into a mixture of 250 ml of distilled H₂O and 25 ml of diethylenetriamine to precipitate the hydrated chromium oxide and to chelate the nickel ion. This slurry was allowed to stand for 2-3 hours after which it was slowly added dropwise, to the cold (NH₄)₂MoS₄/dien solution with agitation after each addition keeping the flask as cold as possible, which resulted in a bright orange precipitate being formed. The resulting precipitate-containing mixture was stirred in the ice bath for one half hour after the addition was completed. The precipitate was separated out by vacuum filtration through a Buchner funnel. The resulting cake was washed with distilled water, then with ethanol, and



dried under vacuum for 16-24 hrs. Eighty-three grams of orange colored precipitate were recovered. This material was sulfided in the same way as was the $\text{NiMn(en)}_6(\text{MoO}_4)_2$ material, recited above.

Via Chelation: A $\text{Ni}_{0.5}\text{-Cr}_{0.5}\text{-Mo}_{1.25}$ can be prepared as follows: 38 grams of $(\text{NH}_4)_2\text{MoS}_4$ was dissolved in 100 ml degassed ethylene diamine and cooled in an ice bath. Previously prepared $[\text{Cr(en)}_3]\text{Cl}_3(16.4\text{g})$ and $[\text{Ni(en)}_3]\text{Cl}_2(16.7\text{g})$, synthesized in the conventional manner by methanol precipitation from an aqueous solution of $\text{NiCl}_2\cdot\text{H}_2\text{O}$ to which ethylene diamine had been added, were dissolved in a mixture of 150 ml water and 5 ml ethylene diamine. This solution was filtered. The clear filtrate was added dropwise to the $(\text{NH}_4)_2\text{MoS}_4$ solution with vigorous agitation. A red-orange precipitate formed. After addition was complete the red-orange product was filtered. It was washed with a mixture of 450 ml H_2O and 60 ml ethylene diamine in 3 washes, followed by a 300 ml methanol wash and 100 ml diethylether wash. After vacuum drying the yield was about 45 g. - This material was sulfided in the same way as was the $\text{NiMn(en)}_6(\text{MoO}_4)_2$ material, recited above.

Typical Preparation of $\text{Ni(en)}_3\text{WS}_4$

125 grams of $(\text{NH}_4)_2\text{WS}_4$ were placed into a two liter flask and 250 cc of ethylenediamine was slowly added. The flask was then placed in an ice bath and cooled to about 0°C . Then a mixture of 86.2 grams $\text{NiCl}_2\cdot 6\text{H}_2\text{O}$ (4% excess) in 500 ml distilled water was added dropwise, keeping the flask contents as cold as possible. After all the reactants were combined the flask was left in the ice bath for 1/2 hour. The contents were then vacuum filtered and the solids washed with distilled water and ethanol. The catalyst precursor was pressed into 12-20 mesh granules using a 2.25 inch die at 50,000 psi. 25 grams of the 12-20 mesh $\text{Ni(en)}_3\text{WS}_4$ was treated with 15% H_2S (balance H_2) at 350°C for 2 hours then cooled and treated with N_2 bubbling through H_2O to possibly for 1 hour. The procedure was repeated 4 times, then all the products were combined, re-pelleted and sieved to a 20/40 mesh size.



Typical Preparation of $\text{Co(en)}_3\text{MoO}_4$

57.38 grams of $(\text{NH}_4)_6\text{Mo}_7\text{O}_{24} \cdot 4\text{H}_2\text{O}$ was put into a two liter flask and 125 cc of distilled water and 250 cc of ethylenediamine were added. The flask was put into a wet ice bath and cooled to about 0°C with N_2 bubbling through the solution. A solution of 77.38 grams of $\text{CoCl}_2 \cdot 6\text{H}_2\text{O}$ in 500 ml of methanol (methanol had had N_2 sparged through it for 1 hr) was added dropwise. After all the reactants were combined the solution was stirred in the ice bath for 1/2 hour then vacuum filtered under N_2 and washed with four 300 ml portions of methanol then put into a vacuum oven with a N_2 bleed overnight. This material was sulfided in the same way as was the $\text{NiMn(en)}_6(\text{MoO}_4)_2$ material, recited above.

Typical preparation of $\text{Ni}_{0.5}\text{Mn}_{0.5}(\text{en})_3\text{MoS}_4$

100 grams of $(\text{NH}_4)_2\text{MoS}_4$ was placed in a 2 liter flask and 200 cc of ethylenediamine was added slowly after all the ethylenediamine was added the flask was put into a wet ice bath and cooled to about 0°C with stirring. A mixture of 45.67 gm $\text{NiCl}_2 \cdot 6\text{H}_2\text{O}$ (1% excess) and 38.40 grams $\text{MnCl}_2 \cdot 4\text{H}_2\text{O}$ (1% excess) in 500 cc distilled water was then added dropwise. Keeping the mixture as cold as possible after all the mixture was added, the flask was left in the ice bath with stirring for 1/2 hour then vacuum filtered. The solids were washed with distilled water and ethanol and left overnight under N_2 . The solids were then dried in a vacuum oven. This material was sulfided in the same way as was the $\text{NiMn(en)}_6(\text{MoO}_4)_2$ material, recited above.

Test Procedure:

Catalysts were tested on a partially dewaxed ($+8^\circ\text{C}$ pour) Isthmus 600N raffinate (82 VI@ -9°C pour) in a Micro Catalyst Screening Unit (MCSU). In the MCSU, the 1.18 cm ID 347 stainless steel reactor contains a 0.32 cm OD thermowell. The reactor is heated by a Lindberg electric furnace. Typical catalyst bed height is 24 cm (23 ml charge), and the catalyst charge (20-28 mesh size unless stated otherwise) is supported in the hot zone by a $140\text{ }\mu\text{m}$ stainless steel filter.



The unit operated isothermally in a down-flow mode, and cannot run waxy feeds. Inlet gas rate is controlled by a Brooks mass flow meter, and reactor pressure is regulated by a 90W valve. The reactors are equipped with quick fit connections allowing them to be withdrawn and catalysts stored under H_2 or N_2 pressure, and re-installed later if desired. Operation is normally during regular working hours only. At the start of a day's run, the sequence of events is: 1) Start feed and hydrogen flow at desired operating pressure. 2) Heat to target reactor temperature as quickly as possible. Sampling from the gas-liquid separator is carried out hourly, with gas and liquid flows and liquid product yield recorded at that time. At the end of the day's operation, the unit is shutdown as follows: 1) Cool down reactor as quickly as possible under gas and liquid flow, to less than $150^\circ C$. 2) Bottle up reactor under H_2 pressure. Feed characterization data is provided in Table 1. Operating conditions were 8.5 MPa H_2 ; 1.0 LHSV, and treat gas rate of 425 ml/minute (~ 6200 SCF/B), with activity varied by reactor temperature between 325 and $380^\circ C$. Lined out total liquid product from each days run was blended and topped to give $360^\circ C+$ product in a HIVAC model C still. Operating conditions, yields and product characterizations for each of the five catalysts are contained in Tables 2 to 7. Data for fluorided Ni/W/ Al_2O_3 catalyst, a high activity hydrocracking/hydrotreating catalyst, available as Shell 454, is included for comparison purposes and reported in detail in Table 2A. The process of the present invention using bulk Ni/Mn/Mo and Ni/Cr/Mo sulfide catalysts is superior for VI increase and, in the case of the Ni/Mn/Mo sulfide catalyst also superior for aromatics saturation as compared to the other catalysts tested, including fluorided Ni/W/ Al_2O_3 . The process of the present invention, by using the recited bulk sulfide catalysts also avoids the use of any fluoriding agent and, therefore, avoids the corrosion associated with the presence of fluoride in hydrotreating and ancillary equipment.

Activation and Fluoriding of Shell 454 Catalyst

The Ni/W/ Al_2O_3 catalyst was activated by a procedure recommended by the vendor, as follows: The 20-28 mesh sized catalyst was calcined at $485^\circ C$ for 1 hour prior to loading. Twenty-three cc of



catalyst was charged to the MCS unit reactor. Following cold pressure testing with N_2 and H_2 (8.5 MPa), the reactor was purged with N_2 (6 CF/hr) and heated to $420^\circ C$, pressure tested at 8.5 MPa, held at these conditions overnight, then cooled to $107^\circ C$. Treat gas was changed to 2% H_2S in H_2 , and sulphiding feed was introduced (1.49 g dimethyl disulphide per 100 g CWO base) at maximum pump rate, 4.1 MPa pressure, and 136 ml/min (0.28 CF/hr) treat gas rate. When liquid flow through the separator had been established, feed rate was reduced to 1 LHSV, and the reactor quickly heated to $190^\circ C$, held for 3 hours, heated to $330^\circ C$ at $47^\circ C$ per hour, held for 1 hour, then cooled to $50^\circ C$ with gas and liquid flow.

In-situ fluoriding with 2-fluorotoluene is recommended by the catalyst vendor for Shell 454 $Ni/W/Al_2O_3$ catalyst using 400-800 ppm fluoride in feed, but detailed procedures were not provided. Fluoride concentration was 1000 wppm (0.575 g 2-fluorotoluene per 100 g CWO base), and process conditions were $360^\circ C$, 4.0 LHSV, 8.5 MPa, 5000 SCF/B treat gas rate. Catalyst activity was monitored by changes in Refractive Index. Through the first day of fluoriding, activity rapidly increased. After the overnight shutdown, activity was lower than at the end of the previous day, and increased again through the second day of fluoriding. During the third day, steady performance was observed, and fluoriding was terminated at this point. At the start of each days' run, fluoride containing feed was pumped through the reactor as the temperature was increased from ambient to $360^\circ C$, and again as the reactor was being cooled at the end of the days run. The nominal total amount of fluoride added was 1.51 g or 6.8 wt% fluoride on catalyst. It should be noted that product Refractive Index has been found to be an unreliable indicator of catalyst activity for raffinate hydrotreating. In general, the fluoriding procedures have been found to give a large fluoride gradient through the catalyst bed, with a higher concentration at the feed inlet (bottom of bed for upflow operation). The fluoride distribution resulting from the procedures used in these small lab unit (MCS units) experiments is not known at this time.

This Sheet Has Been Replaced

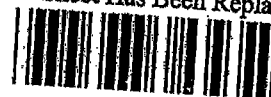


TABLE 1

FEEDSTOCK CHARACTERIZATION

Feed Type:	Isthmus 600N 82VI Raffinate
Dewaxing Conditions:	100% MBK, 1.5/1 solvent/al, +8°C filter temp.

Feed Inspections

Sulfur, wt%	1.07
Total nitrogen, wppm	280
Basic nitrogen, wppm	159
Viscosity @ 40°C, cSt	130.61
Viscosity @ 100°C, cSt	12.8
Viscosity Index	89
Pour Point, °C	+12

HPLC Separation:

Saturates, wt%	55.8
Aromatics, wt%	41.2
Recovery, wt%	96.9

GCD, °C

IBP/5	329/413
20/50	461/495
90/95	568/563
FBP	594

Mass Spectral Composition (LV%)

Paraffins	7.77
Saturates average ring no	1.74
Alkyl benzenes	13.37
2 ring aromatics	11.03
3 ring aromatics	2.18
4 ring aromatics	0.69

THIS SHEET HAS BEEN
REPLACED



360°C+ product VI as a function of reactor temperature is plotted for the five bulk catalysts in Figure 1. VI is corrected to -9°C pour point by subtracting 6 from the measured values in Tables 3-7 (assumes 1 VI unit per 3°C pour point).

Ni/Mn/Mo (from oxide: molybdate)

The bulk Ni/Mn/Mo sulfide catalyst derived from the oxide shows outstanding start of run activity under these conditions exhibiting the production of oil of high VI at lower temperature than the other catalyst tested with the exception of the Ni/Cr/Mo bulk sulfide derived from the sulfide precursor. The catalyst is about 20°C more active than Ni/W/3-Al₂O₃ in achieving the same VI.

The high initial VI (103) obtained at 340°C gradually declined to 95 VI after 30 hours on oil, and returned to about 102 VI when reactor temperature was increased to 350°C. The plotted runs are those performed at a gas rate of 6200 SCF H₂/bbl. The catalyst is also active for aromatics saturation. The catalyst had a BET surface area of 61.7 m²/g.

Ni/Mn/Mo (from sulfide: thiomolybdate)

This material differs nominally from the first catalyst in the source of molybdenum used. This catalyst, however, has very poor activity for both VI increase and aromatics saturation activity as indicated below:

	Ni/Mn/Mo	Ni/Mn/Mo	
<u>Catalyst</u>	<u>(from oxide)</u>	<u>(from sulfide)</u>	<u>Ni/W/F-Al₂O₃</u>
Reactor Temp°C	340	340	340
Product VI	95	85	89
Product Saturates	85	60	72.5
wt%			

This catalyst had a BET surface area of 12.4 m²/g.



Ni/W (from sulphide)

The bulk nickel tungsten sulphide catalyst showed poorer activity for both VI increase and aromatics saturation. This catalyst had a BET surface area of $0.6 \text{ m}^2/\text{g}$.

Co/Mo (from oxide: molybdate)

For the first eight hours on oil, this catalyst showed very high activity, comparable to the Ni/Mn/Mo (oxide precursor) catalyst. However, activity thereafter declined rapidly to a low level, suggestive of rapid poisoning of the active sites by some feed or product component, or possibly catalyst structural arrangement. This catalyst had a BET surface area of $50 \text{ M}^2/\text{g}$.

Ni/Cr/Mo (from sulphide: thiomolybdate)

This catalyst also showed very promising activity, as indicated in Fig. 1. Activity for VI increase was similar to that of Ni Mn Mo (from oxide). Whereas the Ni/Mn/Mo (oxide precursor) catalyst appeared to lose activity somewhat during the course of the evaluation, the Ni/Cr/Mo catalyst activity appeared somewhat steadier. This catalyst showed less aromatics saturation activity than did Ni/Mn/Mo (from oxide: molybdate). This catalyst had a BET surface area of $7.9 \text{ m}^2/\text{g}$.

TABLE 2

MCS UNIT OPERATIONS

UNIT/REACTOR/RUN MCS/116/7 MCS/113/3 MCS/105/8 MCS/111/9 MCS/111/10

Treat Gas

100% H₂

Operating Mode

Isothermal. Once through. downflow

Mass Velocity, kg/hr/m²
(lb/hr/ft²)210
(43)

Length of catalyst bed, cm

23.1

Length of Inert Packing in

Hot Zone, cms:

A) Above catalyst
R) Below catalyst

10.9

Catalyst

Type
Precursor SourceNi/Mo
sulphide

Catalyst charge, ml

23

Catalyst weight, g

30.1

Catalyst size

20 30 mesh

Startup procedure

- 1) Ambient temperature pressure test with nitrogen (9 MPa)
- 2) Reactor heated to 100°C under nitrogen purge (6 CF/hr, 1MPa), held for one hour
- 3) Reactor cooled to ambient temperature under nitrogen purge
- 4) Pressure tested with hydrogen at ambient temperature (9 MPa)
- 5) Cut in feed at ambient temperature, heated to reaction temperature under hydrogen flow (425 ml/min, 8.5 MPa)



TABLE 2A

OPERATING CONDITIONS, YIELDS AND PRODUCT INSPECTIONS

Catalyst Description Shell 454, Crushed and Screened to 20-28 mesh, on-stream fluorided
Run Number MCS-113-4

Operating Conditions and Yields

Feed	Istimus 600N 82 VI Raffinate					
	1.0					
Feed Rate (LHSV)	360	360	370	370	375	380
Gas Rate, SCF/B						340
Reactor Temp., °C						
Unit Pressure, MPa					8.5	
Catalyst Hours	16-20	23-27	30-34	37-41	44-48	51-55

TLP Yield on Feed, wt% 93.7 93.2 95.99 97.11 94.7 94.9

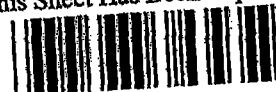
360°Ct Yield on Feed, wt% 79.4 84.0 84.4 75.91 79.5 77.1

360°Ct Product Inspections

Density @ 15°C, kg/dm ³	50	18	17	14	7	5
Sulphur, wppm	<1	<1	<1	1	1	<1
Total Nitrogen, wppm	65.87	74.47	68.44	67.62	60.33	51.80
Viscosity @ 40°C, cSt	8.84	9.50	9.02	8.99	8.48	7.81
Viscosity @ 100°C, cSt	108	105	106	107	112	117
Viscosity Index						95
HPLC Separation:						
Saturates, wt%	88.6	84.5	84.7	86.3	87.0	85.3
Aromatics, wt%	11.0	15.2	14.4	13.4	12.7	13.7
Recovery, wt%	99.6	99.7	99.2	99.7	99.7	98.9

25.7
16
101.24
11.14
95
72.7
26.6
99.3

This Sheet Has Been Replaced



THIS SHEET HAS BEEN REPLACED

This Sheet Has Been Replaced



Table 3

OPERATING CONDITIONS, YIELDS AND PRODUCT INSPECTIONS: MCSU

Unit Run #

MCS 116 1

Catalyst Description

$\text{Ni}_2\text{O}_3, \text{Mn}_2\text{O}_3, \text{Mo}$ (from oxide)

Unit Conditions

Feed

600N 82 VI Raffinate

Feed Rate (LHSV)

1.0

Gas Rate, SCF/B

6200 6200 6200 6200 6200

Reactor Temp., °C

340 340 340 340 340

Unit Pressure, MPa

8.5 8.5 8.5 8.5 8.5

Catalyst Hours

4 6 9 11 16-20 23-21 30 34 37-41 44 48 51 55 58-62

Total Liquid Product

96.1 96.2 96.3 96.4 96.5

TLP Yield, wt%

91.4 91.5 91.5 91.5 91.5

360°C Yield on Feed, wt%

87.2 87.2 87.2 87.2 87.2

Refractive Index @ 75°C

1.4518 1.4584 1.4602 1.4622 1.4600

Hydrogen, wt% by NMR

96.1 96.2 96.3 96.4 96.5

360°C Liquid Product

91.4 91.5 91.5 91.5 91.5

Visc @ 40°C, cSt

64.17 64.20 64.29 64.34 64.38

@ 100°C, cSt

8.82 8.86 8.96 9.01 9.06

Viscosity Index

109 108 104 101 101

Dührman Sulphur, wppm

80 80 74 74 74

X-Ray Sulphur, wt%

2 2 2 2 2

Total Nitrogen, wppm

3 3 3 3 3

Basic Nitrogen, wppm

3 3 3 3 3

HPLC Separation

91.4 91.5 91.5 91.5 91.5

Saturates, wt%

89.1 89.1 89.1 89.1 89.1

Aromatics, wt%

7.1 7.1 7.1 7.1 7.1

Recovery, wt%

97.4 97.4 97.4 97.4 97.4

THIS SHEET HAS BEEN REPLACED



TABLE 3 (CONT'D)

OPERATING CONDITIONS, YIELDS AND PRODUCT INSPECTIONS, MCSU

Unit Run #	MCS 116.7									
Catalyst Description	Ni _{0.5} Mn _{0.5} MoS ₄ (from oxide)									
Unit Conditions										
Feed	600N 82 VI Raffinate									
Feed Rate (LHSV)	6200	500	500	500	500	500	500	500	6200	
Gas Rate, SCF/B	350	340	340	340	355	350	350	350	350	
Reactor Temp., °C	8.5	4.1	4.1	4.1	4.1	4.1	4.1	4.1	8.5	
Unit Pressure, MPa	65-69	75-76	79-83	86-90	93-97	100-103				
Catalyst Hours										
Total Liquid Product										
TLP Yield, wt%	97.7	97.7	99.4	98.4	99.0	97.5				
360°Ct yield on Feed, wt%	87.0	94.6	97.2	94.5	95.2	90.8				
Refractive Index @ 75°C	1.4600	1.4717	1.4712	1.4707	1.4690	1.4632				
Hydrogen, wt% by NMR										
360°Ct Liquid Product										
Visc @ 40°C, cSt	70.67	118.9	109.76	103.54	98.15	86.77				
@ 100°C, cSt	9.21	11.95	11.17	11.24	10.98	10.29				
Viscosity Index	106	88	85	94	96	100				
Dohrman Sulphur, wppm	11	-	-	-	-	-				
X-Ray Sulphur, wt%	<2	0.167	0.131	0.089	0.029	-				
Total Nitrogen, wppm		88	80	-	-	-				
Basic Nitrogen, wppm										
IPLC Separation										
Saturates, wt%	89.8	62.9	62.0	63.6	68.0	83.7				
Aromatics, wt%	9.8	36.2	35.5	36.3	31.5	15.5				
Recovery, wt%	99.5	99.0	97.5	99.9	99.5	99.3				

This Sheet Has Been Replaced



TABLE 4

OPERATING CONDITIONS, YIELDS AND PRODUCT INSPECTIONS, MCSU

Unit Run #	MCS 113-3									
Catalyst Description	Ni/WS (from sulphide)									
Unit Conditions										
Feed	600N 82 VI Raffinate									
Feed Rate (LHSV)	1.0									
Gas Rate, SCF/B	6200									
Reactor Temp., °C	340	375	375	375	375	380				
Unit Pressure, MPa	8.5									
Catalyst Hours	8-12	15-19	22-26	29-33	36-40					
Total Liquid Product										
TLP Yield, wt%	96.1	97.8	96.2	94.9	96.6					
360°Ct Yield on Feed, wt%	94.1	87.4	84.6	82.8	79.0					
Refractive Index @ 75°C	1.4744	1.4662	1.4657	1.4658	1.4626					
Hydrogen, wt% by NMR										
360°Ct Liquid Product										
Visc @ 40°C, cSt	124.2	85.17	81.5	79.75	65.75					
@ 100°C, cSt	12.45	10.21	9.97	9.86	8.93					
Viscosity Index	90	100	102	102	110	160				
Dohrman Sulphur, wppm										
X-Ray Sulphur, wt%	0.677	0.091	0.071	0.089						
Total Nitrogen, wppm	128									
HPLC Separation										
Saturates, wt%	58.9	73.6	74.7	72.4	80.2					
Aromatics, wt%	37.6	25.6	23.9	25.9	19.0					
Recovery, wt%	96.5	99.2	98.6	98.4	99.2					

THIS SHEET HAS BEEN REPLACED

This Sheet Has Been Replaced



TABLE 5

OPERATING CONDITIONS, YIELDS AND PRODUCT INSPECTIONS, MCSU

Unit Run #

MCS 105-8

Catalyst Description

CoMo (from oxide)

Unit Conditions

Feed

Feed Rate (LHSV)

Gas Rate, SCF/B

Reactor Temp., °C

Unit Pressure, MPa

Catalyst Hours

600N 82 VI Raffinate

1.0

6200

350

350

8.5

35

39

41

44

54

57

64

67

Total Liquid Product

TLP Yield, wt%

360°C Yield on Feed, wt%

Refractive Index @ 75°C

Hydrogen, wt% by NMR

360°C Liquid Product

Visc @ 40°C, cSt

@ 100°C, cSt

Viscosity Index

Dohrman Sulphur, wppm

X-Ray Sulphur, wt%

Total Nitrogen, wppm

Basic Nitrogen, wppm

HPLC Separation

Saturates, wt%

Aromatics, wt%

Recovery, wt%

THIS SHEET HAS BEEN REPLACED

TABLE 6
OPERATING CONDITIONS, YIELDS AND PRODUCT INSPECTIONS, MCSU

Unit Run #	MCS 111-9									
Catalyst Description	Ni _{0.5} Mn _{0.5} Mo (from sulphide)									
Unit Conditions	600N 82 VI Raffinate									
Feed	1.0									
Feed Rate (LHSV)	6200									
Gas Rate, SCF/B	340									
Reactor Temp., °C	360									
Unit Pressure, MPa	360									
Catalyst Hours	5 8	12 15	19 22	26 29	33 36	40 43				
Total Liquid Product										
TLP Yield, wt%	96.8	97.8	98.6	96.6	98.0	96.5				
360°C Yield on Feed, wt%	95.7	97.3	98.1	96.1	97.4	95.3				
Refractive Index @ 75°C	1.4720	1.4717	1.4716	1.4698	1.4695	1.4671				
Hydrogen, wt% by NMR										
360°C Liquid Product										
Visc @ 40°C, cSt	115.72	112.94	110.34	103.17	101.41	89.85				
@ 100°C, cSt	11.98	11.77	11.70	11.32	11.09	10.43				
Viscosity Index	91	91	93	95	94	98				
Dohrman Sulphur, wppm	-	-	-	640	530	260				
X-Ray Sulphur, wt%	0.20	0.16	0.15	-	-	-				
Total Nitrogen, wppm	170	150	150	87	67	37				
Basic Nitrogen, wppm										
HPLC Separation										
Saturates, wt%	62.5	62.8	60.3	66.7	66.9	67.1				
Aromatics, wt%	36.6	36.0	34.9	32.8	30.6	27.7				
Recovery, wt%	99.1	98.7	95.2	99.5	97.5	94.8				

This Sheet Has Been Replaced



THIS SHEET HAS BEEN
REPLACED

This Sheet Has Been Replaced



Table 7

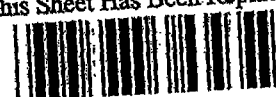
OPERATING CONDITIONS, YIELDS AND PRODUCT INSPECTIONS, MCSU

Unit Run #	MCS 111 10									
Catalyst Description	Ni ₂ , Cr ₂ , Mo (from sulphide)									
Unit Conditions	600N 82 VI Raffinate									
Feed	1.0									
Feed Rate (LHSV)	6200									
Gas Rate, SCF/B	325	340	340	350	350	350	350	355	345	
Reactor Temp., °C										
Unit Pressure, MPa										
Catalyst Hours	11-14	18-21	25-28	32-35	39-42	46-49	53-56	60-63		
Total Liquid Product										
TLP Yield, wt%	97.8	98.5	96.6	93.8	93.2	93.3	93.5	94.3		
360°C Yield on Feed, wt%	97.1	97.6	93.2	82.9	79.1	79.2	79.3	83.7		
Refractive Index @ 75°C	1.4707	1.4079	1.4661	1.4604	1.4555	1.4567	1.4561	1.4594		
Hydrogen, wt% by NMR										
360°C Liquid Product										
Visc @ 40°C, cSt	106.07	88.89	85.42	67.50	56.70	57.60	56.68	64.85		
@ 100°C, cSt	11.38	10.30	10.11	9.05	8.13	8.20	8.16	8.81		
Viscosity Index	92.8	96.6	98.2	109.1	111.9	111.4	112.9	109.3		
Dohrman Sulphur, wppm	58	440	570	160	160	290	185	140		
X-Ray Sulphur, wt%	0.25							47		
Total Nitrogen, wppm	44	13	5							
Basic Nitrogen, wppm										
IPLC Separation										
Saturates, wt%	65.8	63.8	69.0	79.8	88.7	84.5	85.8	83.0		
Aromatics, wt%	35.2	35.7	31.1	18.8	10.1	12.3	10.1	16.0		
Recovery, wt%	101.0	99.4	100.1	98.6	98.8	96.8	95.9	99.0		

THIS SHEET HAS BEEN REPLACED

This Sheet Has Been Replaced

CLAIMS:



1. A method for increasing the Viscosity Index of lubricating and specialty oils by contacting with a catalyst, said method comprising contacting a lube or specialty oil stock under hydrotreating conditions with a specific bulk metal sulfide hydrotreating catalyst selected from the group consisting of bulk Cr/Ni/Mo sulfide catalyst, bulk Ni/Mo/Mn sulfide catalyst and mixtures thereof, wherein the Cr/Ni/Mo sulfide catalyst is prepared by the decomposition in a sulfide containing atmosphere of a $[\text{Ni}(\text{L})_3](\text{Cr}(\text{L})_3)[\text{MoS}_4]_{2-2.5}$ precursor complexes and the Ni/Mo/Mn sulfide catalyst is prepared by the decomposition in inert atmosphere followed by sulfiding using $\text{H}_2\text{S}/\text{H}_2$ of a $[\text{Ni}(\text{L})_3](\text{Mn}(\text{L})_3)[\text{MoO}_4]_2$ precursor complex where L is one or more neutral, nitrogen containing ligand, at least one of which is a chelating polydentate ligand.

2. The method of claim 1 wherein hydrotreating conditions comprise temperature in the range 200 to 450°C, pressure in the range 300 to 3000 psi, space velocity in the range of about 0.1 to 5.0 v/v/hr and a treat gas rate in the range of about 200 to 20,000 SCF H_2 /bbl.

3. The method of claim 1 or 2 wherein the $[\text{Ni}(\text{L})_3](\text{Mn}(\text{L})_3)[\text{MoO}_4]_2$ precursor is converted into the bulk catalyst by decomposing the precursor in the presence of an inert atmosphere at 300 to 400°C and sulfiding using $\text{H}_2\text{S}/\text{H}_2$ at 200 to 600°C and the $[\text{Ni}(\text{L})_3](\text{Cr}(\text{L})_3)[\text{MoS}_4]_{2-2.5}$ precursor is converted into the bulk catalyst by decomposing the precursor in a sulfur containing atmosphere at a temperature of at least 150°C.

4. The method of claim 3 wherein the ligand is ethylene diamine.

5. The method of claim 1 or 2 wherein the bulk metal sulfide catalyst is used in the form of pellets made by compacting bulk metal sulfide under pressure.

THIS SHEET HAS BEEN
REPLACED

INCREASING VI OF LUBE OIL BY HYDROTREATING USING
BULK Ni Mn Mo OR Ni Cr Mo SULFIDE CATALYSTS PREPARED
FROM LIGATED METAL COMPLEXES

ABSTRACT OF THE DISCLOSURE

Lube oils of increased VI are prepared by hydrotreating using bulk Ni/Mn/Mo or Cr/Ni/Mo sulfide catalysts prepared from ligated, e.g., ethylene diamine, metal complexes. In particular the Ni/Mn/Mo sulfide catalyst is prepared from the oxide, precursor with subsequent sulfiding while the Cr/Ni/Mo sulfide catalyst is prepared from the sulfide precursor in the presence of H_2S/H_2 .

1/1

LEGEND

- NI/W/F/Al₂O₃
- NI/Mn/Me(FROM OXIDE)
- NI/W(FROM SULPHIDE)
- Co/Mo(FROM OXIDE)
- △ NI/Mn/Mo(FROM SULPHIDE)
- × NI/Cr/Mo(FROM SULPHIDE)

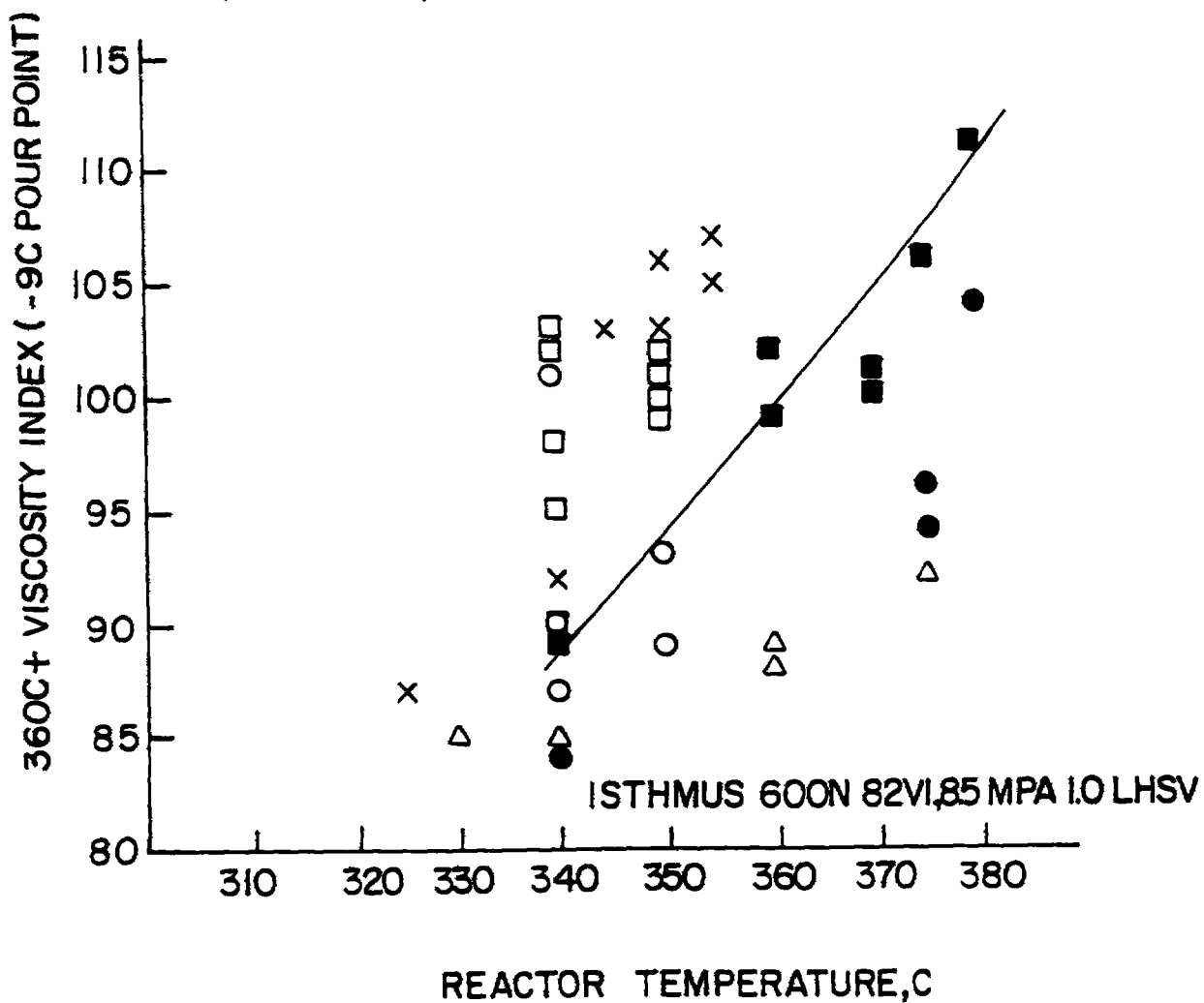
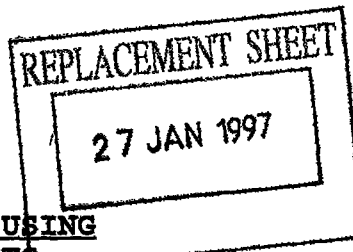


FIG. 1

THIS SHEET HAS BEEN
REPLACED



INCREASING VI OR LUBE OIL BY HYDROTREATING USING
BULK NI MN MO OR NI CR MO SULFIDE CATALYSTS

Description of the Invention

The Viscosity Index of lubricating and speciality oils can be increased by hydrotreating said oil employing as the hydrotreating catalyst a bulk Ni/Mn/Mo or Cr/Ni/Mo sulfide catalyst prepared from ligated, e.g. ethylene diamine metal complexes wherein the Ni/Mn/Mo sulfide catalyst is prepared from the oxide precursor with subsequent sulfiding while the Cr/Ni/Mo sulfide catalyst is prepared from the sulfide precursor in the presence of H_2S/H_2 .

Background of the Invention

Raffinate hydrotreating and hydrocracking processes to produce lube stocks are well known. In general, these processes employ catalysts which comprise metal oxides or sulfides supported on alumina or silica-alumina. Metals usually include nickel, molybdenum, cobalt, manganese, tungsten, etc. and mixtures thereof in the oxide or sulfide state. When the catalyst is in the oxide state the environment of use being one containing sulfur containing species converts the metal oxides into metal sulfides.

The objectives in the hydrotreater are to increase VI and saturates levels. Catalysts used today in lubes hydrotreating do not perform this function very well insofar as they are employed principally for hydrodesulfurization, hydrodenitrogenation, color improvement and residual extraction solvent removal. High pressure hydrocracking of distillates or deasphalted oil is an acknowledged alternative to solvent extraction and, in the case of distillate hydrocracking all of the VI increase, saturates increase and aromatics decrease are achieved by the catalytic conversion.

The major reactions which contribute to VI increase are of the ring opening type, whether the ring is aromatic or cyclic saturated. Hydrogenation will contribute to a lesser degree to VI

This Sheet Has Been Replaced



THIS SHEET HAS BEEN
REPLACED



27 JAN 1997

increase. At low severity, hydrodesulfurization will give a 3-4 VI unit increase, most likely the result of ring opening that occurs with the hydrodesulfurization of thiophenic structures.

US 4,650,563 teaches molybdenum and/or tungsten hydrotreating catalysts prepared from promoter metal thiomolybdate and/or thiotungstate complexes and their use as hydrotreating catalysts. The catalysts are prepared by heating one or more catalyst precursors under oxygen-free conditions in the presence of sulfur at a temperature of at least about 150°C. The catalyst precursors will be one or more compounds of the formula $(ML)(Mo_y W_{1-y} S_4)$ wherein M comprises one or more divalent promoter metals selected from the group consisting of Ni, Co, Zn, Cu and mixtures thereof, wherein y is any value ranging from 0 to 1 and wherein L is one or more neutral, nitrogen containing ligands at least one of which is a chelating polydentate ligand. L may be ethylene diamine.

US 4,716,139 teaches a catalyst composition comprising an amorphous sulfide of trivalent chromium and at least one metal selected from the group consisting of (a) molybdenum, tungsten and mixtures thereof and (b) mixtures of (a) with iron, nickel, cobalt, manganese, copper, zinc and mixtures thereof. The catalyst is prepared by heating one or more precursors at a temperature of at least about 200°C under oxygen-free conditions in the presence of sulfur, wherein said precursor is selected from the group consisting of



and mixtures thereof wherein M is one or more divalent promoter metals selected from the group consisting of Mn, Fe, Co, Ni, Cu, Zn and mixtures thereof, wherein L is one or more neutral nitrogen-containing ligands, wherein ligand X is a singly-charged anionic species such as NO_2^- , OH^- , Cl^- , etc. and wherein $1 \geq z \geq 0.1$ and $z \geq y \geq 0$ and wherein n equals $(3-z-y)/2$. The catalysts are good hydroprocessing catalysts, particularly hydrotreating catalysts. These catalysts were demonstrated for their HDS and HDN ability on light catalytic cycle oil (LCCO).

This Sheet Has Been Replaced



- 3 -

REPLACEMENT SHEET

27 JAN 1991

See also U.S. Patent 4,626,339 which teaches the supported version of the catalyst of '139.

U.S. 4,663,023 teaches hydrotreating with self promoted molybdenum and Tungsten sulfide catalysts. Catalysts are made from precursors which are one or more compounds of the formula $ML(Mo_yW_{1-y}O_4)$ wherein M is one or more promoter metals selected from the group consisting essentially of Mn, Fe, Co, Ni, Cu, Zn and mixtures thereof $0 \leq y \leq 1$ and L is a neutral nitrogen containing ligand. This patent would, therefor, embrace Mn/Ni/Mo sulfide and Ni/Mo sulfide catalysts. The patent claims the use of these catalysts in general to hydrorefine hydrocarbon feeds including lube oils.

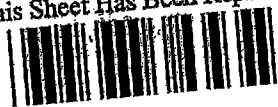
U.S. 4,698,145 teaches the upgrading of hydrocarbon feeds by use of a supported self-promoted catalyst. The catalyst is prepared by heating a composite of support material and one or more catalyst precursor salts under oxygen-free conditions and in the presence of sulfur. The precursor salt or salts are of the general formula $(ML)(Mo_yW_{1-y}S_4)$ wherein M is one or more promoter metals selected from the group consisting essentially of Mn, Fe, Co, Ni, Zn and mixtures thereof, $0 \leq y \leq 1$, and L is a neutral, nitrogen containing ligand. This patent would, therefore, embrace supported Mn/Ni/Mo sulfide and supported Ni/Mo sulfide catalysts. The claims recite lube oil as one of the hydrocarbon feeds treated using this class of catalysts in general.

U.S. 4,595,672 is directed to a method for making self-promoted sulfide hydrotreating catalysts. The catalysts are made from precursors which are one or more compounds of the formula $ML(Mo_yW_{1-y}O_4)$ where M is one or more promoter metals selected from the group consisting essentially of Mn, Fe, Co, Ni, Cu, Zn and mixtures thereof an $0 \leq y \leq 1$. This patent would, therefore, embrace Ni/Mn/Mo sulfide and Ni/Mo sulfide catalytic materials. L is a neutral nitrogen containing ligand.

U.S. 4,668,376 teaches supported Mn sulfide promoted Mo and W sulfide hydroprocessing catalysts. The catalysts are made from

THIS SHEET HAS BEEN
REPLACED

This Sheet Has Been Replaced



- 4 -

REPLACEMENT SHEET

27 JAN 1997

thiometallate precursor salts of the formula $(ML)(Mo_yW_{1-y}S_4)$ where M comprises one or more divalent promoter metals selected from Mn and mixtures of Mn with one or more of Ni, Co, Cu, Zn or mixtures thereof any $0 \leq y \leq 1$ and L is a neutral nitrogen containing ligand. This patent would, therefore, embrace Ni/Mn/Mo sulfide and Ni/Mo sulfide catalytic materials. The patent claims the use of these catalysts in general for hydrorefining hydrocarbon feeds including lube oils.

U.S. 4,591,429 teaching hydroprocessing of hydrocarbon feeds employing supported sulfided catalyst. The catalyst comprises a sulfide of (1) chromium, (2) Mo, W or mixtures thereof and (3) at least one metal selected from the group consisting of Ni, Co, Mn, Cu, Zn and mixtures thereof, and mixtures thereof with Fe. The catalysts are made by compositing a pre-selected quantity of support material with a precursor comprising a mixture of (1) hydrated oxide of trivalent chromium and (2) a salt containing a thiometallate anion of Mo or W and a cation comprising at least one divalent promoter metal. This patent would, therefore, embrace supported Cr/Ni/Mo sulfide catalyst and claims the use of these catalyst in general for hydroprocessing hydrocarbon feeds, including lube oil.

Description of The Figure

Figure 1 compares the $360^\circ C^+$ product VI as a function of reactor temperature for five bulk metal catalysts and a Ni/W/F/Al₂O₃ reference catalyst.

Detailed Description of the Invention

Lubricating and/or specialty oil stocks such as deasphalted oil stocks, lube oil distillates, and solvent extracted lube oil raffinates can have their viscosity indexes increased by hydrotreating, employing specific bulk metal sulfide hydrotreating catalysts selected from the group consisting of bulk Cr/Ni/Mo sulfide catalyst, bulk Ni/Mo/Mn sulfide catalyst and mixtures thereof wherein the catalysts are prepared from specific metal complexes and wherein the Ni/Mn/Mo sulfide catalyst is prepared from the oxide precursor

THIS SHEET HAS BEEN
REPLACED



27 JAN 1997

decomposed in an inert atmosphere such as N_2 and subsequently sulfided using H_2S/H_2 and the Cr/Ni/Mo sulfide catalyst is prepared from the sulfide precursor and decomposed in a non-oxidizing, sulfur containing atmosphere.

These two catalyst have been found to significantly and, surprisingly, to a higher degree than other catalysts, increase the Viscosity Index of lube oil stocks. The feed stocks are contacted with the catalyst under hydrotreating conditions comprising temperatures in the range of about 200 to 450°C, preferably about 300 to 380°C, pressures in the range of about 300 to 3000 psi, preferably at least about 1000 psi, space velocity in the range of about 0.1 to 5.0 v/v/hr., preferably about 0.5 to 2.0 v/v/hr. and a treat gas rate in the range of about 200 to 20,000 SCF H_2 /bbl, preferably about 700 to 5,000 SCF H_2 /bbl.

The catalysts are prepared by using $[Ni(L)_3](Mn(L)_3)[MoO_4]_2$ or $[Ni(L)_3][Cr(L)_3][MoS_4]_{2-2.5}$ as precursor complexes, where L - one or more neutral, nitrogen containing ligands at least one of which is a chelating polydentate ligand such as ethylenediamine. These complexes are prepared by reaction of appropriate metal salts with the ligand solution.

The ligand or ligands L will generally have a denticity of six and will be one or more neutral, nitrogen containing ligands wherein at least one of said ligands is a multi-dentate chelating ligand which chelates the metal cation of the metal salt to form a chelated metal species such as $[ML]^{2+}$, where M is Ni or Cr and Mn.

Those skilled in the art know that the term "ligand" is used to designate functional coordinating groups which have one or more pairs of electrons available for the formation of coordinate bonds. Ligands that can form more than one bond with a metal ion are called polydentate while ligands that can form only one bond with a metal ion are called monodentate. Monodentate ligands are ~~not~~ capable of forming chelates. Hence, if one uses one or more species of monodentate ligands in the precursor molecule, then one must also use

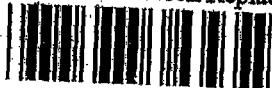
at least one polydentate chelating ligand. Preferably L will be one or more polydentate chelating ligands. The denticity of the ligand L will generally be six, because the promoter metal cations prefer six-fold coordination. Hence, if more than one species of ligand is employed in the precursor molecule, the denticity of the ligand species will usually add up to six. It should be understood that it is possible for ligand L to have a total denticity of less than six, but in most cases L will have a total denticity of six. Thus, L will be three bidentate ligands, two tridentate ligands, a mixture of a bidentate and a quadridentate ligand, a hexadentate ligand or a mixture of a polydentate ligand with monodentate ligands so long as the combination has a total denticity of six. As has heretofore been stated, it is preferred to use chelating bidentate and tridentate ligands. In general, the ligands useful in this invention include alkyl and aryl amines and nitrogen heterocycles. Illustrative but non-limiting examples of ligands useful in the catalyst precursors of this invention are set forth below.

Monodentate ligands will include NH_3 as well as alkyl and aryl amines such as ethyl amine, dimethyl amine, pyridine, etc. Useful chelating bidentate amine ligands are illustrated by ethylenediamine, 2,2'-bipyridine, o-phenylene diamine, tetramethylethylenediamine and propane-1,3 diamine. Similarly, useful chelating tridentate amine ligands are represented by terpyridine and diethylenetriamine while triethylenetetramine is illustrative of a useful chelating quadridentate amine ligands. Useful chelating pentadentate ligands include tetraethylenepentamine while sepulchrates (an octazacryptate) is illustrative of a suitable chelating hexadentate ligand. However, as a practical matter it will be preferred to use chelating, polydentate alkyl amines for L. Illustrative, but not limiting examples of alkyl amines that are useful in the catalyst precursor of this invention include ethylenediamine, diethylenetriamine, and tetraethylenetetramine. It is particularly preferred to use bidentate and tridentate alkyl amines such as ethylenediamine, (en) and diethylenetriamine, (dien).

This Sheet Has Been Replaced



THIS SHEET HAS BEEN
REPLACED



27 JAN 1997

Many of the precursor salts useful in forming the catalysts of this invention and methods for preparing them are known in the art. An article by Diemann and Mueller titled Thio and Seleno Compounds of the Transition Metals With d^0 configuration published in COORD. CHEM. REV. 10:79-122 provides a review of known salts. In general, the precursor salts useful for forming the catalysts useful in this invention may be prepared by mixing an aqueous solution of ammonium thiomolybdate with an aqueous solution of the chelated promoter metal cation $[ML]^{2+}$ which results in the formation of the precursor salt as a precipitate which is readily recovered. The chelating promoter cation is easily formed by, for example, mixing an aqueous solution of one or more water soluble promoter metal salts with the ligand or mixture of ligands. The water soluble salt may be any water soluble salt that is convenient to use such as a halide, sulfate, perchlorate, acetate, nitrate etc. Alternatively, an aqueous solution of ammonium thiomolybdate may be mixed with the ligand with the resulting solution mixed with an aqueous solution of promoter metal salt or the salt can be added to the ligand and dissolved into the solution of thiomolybdate and/or thiotungstate. However, it should be understood that the catalyst precursor preparation is not intended to be limited to aqueous media.

The $[Ni(L)_3](Mn(L)_3)[MoO_4]_2$ precursor is converted into the bulk catalyst by decomposing the precursor complex under inert conditions, such as in the presence of N_2 or an inert gas at 300 to 400°C. The decomposed material is subsequently sulfided using H_2S/H_2 at 200 to 600°C, preferably 250-500°C, most preferably 200-400°C. A preferred ligand is ethylene diamine.

The $[Ni(L)_3](Cr(L)_3)[MoS_4]_{2-2.5}$ precursor is converted into the bulk sulfide catalyst by decomposing the precursor complex in a sulfur containing atmosphere at least 150°C, preferably at least about 200°C for a time sufficient to form the catalyst. A preferred ligand is ethylenediamine.

The sulfur required during the formation of the catalyst may be that which is present in the precursor salt. Thus, the Ni/Cr/Mo

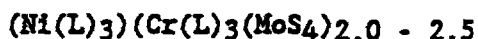


27 JAN 1997

catalyst composition used in this invention will be formed if no excess sulfur is present and if the oxygen-free atmosphere is relatively inert, such as nitrogen. It is preferred that the sulfur will be present in an amount in excess of that contained in the precursor salt. It is preferred that the catalyst be formed by heating the precursor in the presence of excess sulfur, it is also preferred that the excess sulfur be present in the form of a sulfur bearing compound which can be one or more solids, liquids, gases or mixtures thereof. Mixtures of hydrogen and H_2S have been found to be particularly suitable. In general the temperature will range between from about 150° - $600^{\circ}C$, preferably 250 - $500^{\circ}C$, and still more preferably from about 300° - $400^{\circ}C$. The non-oxidizing atmosphere may be gaseous, liquid or mixture thereof.

These bulk metal sulfide catalysts are relatively low surface area materials of about $50m^2/gram$.

The catalyst is prepared in a two step operation. Appropriate salts of Ni, Cr, and Mn are reacted in the presence of excess ligand in combination with MoS_4^{2-} or MoO_4^{2-} . Methanol is used as the solvent. One recovers either



which are then decomposed as previously described.

The bulk metal sulfide catalysts are most conveniently used as shaped pellets or extrudates. The bulk metal sulfide can be pelleted by the application of extremely high pressure. Bulk metal sulfide catalyst pellets have been prepared using what is believed to be this high pressure technique but the details are not known. Another technique involves preparing the bulk metal sulfide in the presence of colloidal silica or colloidal alumina, recovering the precipitate, drying, and grinding the recovered cake and pelletizing using polyvinyl alcohol, in accordance with the procedure of U.S. Patent 4,650,563 and U.S. Patent 4,591,429. It is possible that a preferred bulk metal sulfide catalyst pellet or extrudate is produced without the use of any silica or alumina solely by employing standard,



27 JAN 1997

commercial high pressure pelleting, but economics and cost effectiveness will become important issues.

Experimental

Five bulk metal sulfide catalysts were evaluated for raffinate hydrotreating to determine their effect on VI improvement. These catalysts were:

- 1) $\text{Ni}_{0.5} \text{Cr}_{0.5} \text{Mo}$ (from thiomolybdate)
- 2) $\text{Ni}_{0.5} \text{Mn}_{0.5} \text{Mo}$ (from thiomolybdate)
- 3) NiW (from thiotungstate)
- 4) $\text{Ni}_{0.5} \text{Mn}_{0.5} \text{Mo}$ (from molybdate)
- 5) CoMo (from molybdate)

These bulk materials were pelleted using high pressure without any silica or alumina binder material. The pelleted material was then crushed and sieved to secure particles of about 1/16 inch.

The following are representative of typical procedures used to produce materials of the type identified above.

They are not offered as recitations of the exact, specific procedures used to produce the particular materials hereinafter identified.

Typical Preparation of $\text{NiMn}(\text{NH}_2\text{CH}_2\text{CH}_2\text{NH}_2)_6(\text{MoO}_4)_2$ with subsequent sulfiding

88.2g $(\text{NH}_4)_6\text{Mo}_7\text{O}_{24} \cdot 4\text{H}_2\text{O}$ were dissolved in 200 ml H_2O plus 100 ml ethylenediamine. 72.2g $\text{Ni}(\text{NO}_3)_2 \cdot 6\text{H}_2\text{O}$ and 49.5g $\text{MnCl}_2 \cdot 4\text{H}_2\text{O}$ were dissolved in 800 ml deaerated (N_2 bubbler) methanol. The molybdate solution was placed in a 2-liter flask and deaerated with a nitrogen bubbler while stirring with a mechanical stirrer. Under a blanket of N_2 , the Ni- and Mn- containing solution was added dropwise to the flask from a dropping funnel. A precipitate formed. The precipitate was allowed to settle and the supernatant liquid was siphoned off. Two methanol (deaerated) rinsings were performed, each time siphoning off supernatant-liquid. The precipitate was finally recovered by vacuum filtration from methanol in an N_2 atmosphere box. The yield



27 JAN 1997

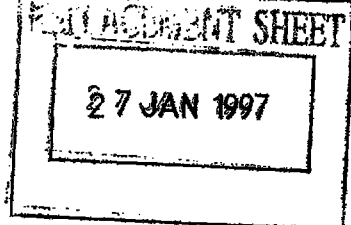
was 87.2%. The catalyst precursor was dried under vacuum at 50°C overnight.

The precursor was then pelletized into a disk and then crushed into 20/40 mesh granules. Sulfiding of the precursor compound to bulk catalysts was done by heating under nitrogen for 50 minutes to 370°C, holding at 370°C for one hour, then heating under a stream of 10% H₂S in hydrogen for ten minutes to 400°C and holding at that temperature under H₂S/H₂ for two hours. Cooling over 1.5 hrs under H₂S/H₂ to <100°C was followed by cooling to room temperature under nitrogen and passivation at room temperature with a stream of 1% O₂/N₂ for half an hour.

Typical Preparation of Cr-Containing Catalysts

Chromium containing catalyst can be prepared via two different chemistries, ololation and chelation in non-rigorous terms. The following reported procedures are representative of the typical techniques employed in producing the material from the thiomolybdate.

Via Ololation: A Ni_{0.5}-Cr_{0.5}-Mo catalyst can be prepared as follows: 40g (NH₄)₂MoS₄ was added to 82 ml of diethylenetriamine (dien) in a one liter flask. Distilled H₂O was used twice to wash off any solution remaining on the sides of the flask. The resulting dark red solution was cooled to 0°C in an ice bath and kept in the bath for the duration of the preparation. In a separate flask, 16.52 grams of CrCl₃·6H₂O and 14.77 gm of NiCl₂·6H₂O were dissolved into a mixture of 250 ml of distilled H₂O and 25 ml of diethylenetriamine to precipitate the hydrated chromium oxide and to chelate the nickel ion. This slurry was allowed to stand for 2-3 hours after which it was slowly added dropwise, to the cold (NH₄)₂MoS₄/dien solution with agitation after each addition keeping the flask as cold as possible, which resulted in a bright orange precipitate being formed. The resulting precipitate-containing mixture was stirred in the ice bath for one half hour after the addition was completed. The precipitate was separated out by vacuum filtration through a Buchner funnel. The resulting cake was washed with distilled water, then with ethanol, and



dried under vacuum for 16-24 hrs. Eighty-three grams of orange colored precipitate were recovered. This material was sulfided in the same way as was the $\text{NiMn(en)}_6(\text{MoO}_4)_2$ material, recited above.

Via Chelation: A $\text{Ni}_{0.5}\text{-Cr}_{0.5}\text{-Mo}_{1.25}$ can be prepared as follows: 38 grams of $(\text{NH}_4)_2\text{MoS}_4$ was dissolved in 100 ml degassed ethylene diamine and cooled in an ice bath. Previously prepared $[\text{Cr(en)}_3]\text{Cl}_3(16.4\text{g})$ and $[\text{Ni(en)}_3]\text{Cl}_2(16.7\text{g})$, synthesized in the conventional manner by methanol precipitation from an aqueous solution of $\text{NiCl}_2\cdot\text{H}_2\text{O}$ to which ethylene diamine had been added, were dissolved in a mixture of 150 ml water and 5 ml ethylene diamine. This solution was filtered. The clear filtrate was added dropwise to the $(\text{NH}_4)_2\text{MoS}_4$ solution with vigorous agitation. A red-orange precipitate formed. After addition was complete the red-orange product was filtered. It was washed with a mixture of 450 ml H_2O and 60 ml ethylene diamine in 3 washes, followed by a 300 ml methanol wash and 100 ml diethylether wash. After vacuum drying the yield was about 45 g. - This material was sulfided in the same way as was the $\text{NiMn(en)}_6(\text{MoO}_4)_2$ material, recited above.

Typical Preparation of $\text{Ni(en)}_3\text{WS}_4$

125 grams of $(\text{NH}_4)_2\text{WS}_4$ were placed into a two liter flask and 250 cc of ethylenediamine was slowly added. The flask was then placed in an ice bath and cooled to about 0°C . Then a mixture of 86.2 grams $\text{NiCl}_2\cdot 6\text{H}_2\text{O}$ (4% excess) in 500 ml distilled water was added dropwise, keeping the flask contents as cold as possible. After all the reactants were combined the flask was left in the ice bath for 1/2 hour. The contents were then vacuum filtered and the solids washed with distilled water and ethanol. The catalyst precursor was pressed into 12-20 mesh granules using a 2.25 inch die at 50,000 psi. 25 grams of the 12-20 mesh $\text{Ni(en)}_3\text{WS}_4$ was treated with 15% H_2S (balance H_2) at 350°C for 2 hours then cooled and treated with N_2 bubbling through H_2O to possibly for 1 hour. The procedure was repeated 4 times, then all the products were combined, re-pelleted and sieved to a 20/40 mesh size.



27 JAN 1997

Typical Preparation of $\text{Co(en)}_3\text{MoO}_4$

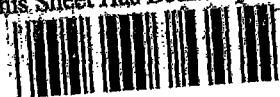
57.38 grams of $(\text{NH}_4)_6\text{Mo}_7\text{O}_{24} \cdot 4\text{H}_2\text{O}$ was put into a two liter flask and 125 cc of distilled water and 250 cc of ethylenediamine were added. The flask was put into a wet ice bath and cooled to about 0°C with N_2 bubbling through the solution. A solution of 77.38 grams of $\text{CoCl}_2 \cdot 6\text{H}_2\text{O}$ in 500 ml of methanol (methanol had had N_2 sparged through it for 1 hr) was added dropwise. After all the reactants were combined the solution was stirred in the ice bath for 1/2 hour then vacuum filtered under N_2 and washed with four 300 ml portions of methanol then put into a vacuum oven with a N_2 bleed overnight. This material was sulfided in the same way as was the $\text{NiMn(en)}_6(\text{MoO}_4)_2$ material, recited above.

Typical preparation of $\text{Ni}_{10} \text{Mn}_5(\text{en})_3\text{MoS}_4$

100 grams of $(\text{NH}_4)_2\text{MoS}_4$ was placed in a 2 liter flask and 200 cc of ethylenediamine was added slowly after all the ethylenediamine was added the flask was put into a wet ice bath and cooled to about 0°C with stirring. A mixture of 45.67 gm $\text{NiCl}_2 \cdot 6\text{H}_2\text{O}$ (1% excess) and 38.40 grams $\text{MnCl}_2 \cdot 4\text{H}_2\text{O}$ (1% excess) in 500 cc distilled water was then added dropwise. Keeping the mixture as cold as possible after all the mixture was added, the flask was left in the ice bath with stirring for 1/2 hour then vacuum filtered. The solids were washed with distilled water and ethanol and left overnight under N_2 . The solids were then dried in a vacuum oven. This material was sulfided in the same way as was the $\text{NiMn(en)}_6(\text{MoO}_4)_2$ material, recited above.

Test Procedure:

Catalysts were tested on a partially dewaxed ($+8^\circ\text{C}$ pour) Isthmus 600N raffinate (82 VI@ -9°C pour) in a Micro Catalyst Screening Unit (MCSU). In the MCSU, the 1.18 cm ID 347 stainless steel reactor contains a 0.32 cm OD thermowell. The reactor is heated by a Lindberg electric furnace. Typical catalyst bed height is 24 cm (23 ml charge), and the catalyst charge (20-28 mesh size unless stated otherwise) is supported in the hot zone by a 140 μm stainless steel filter.



27 JAN 1997

The unit operated isothermally in a down-flow mode, and cannot run waxy feeds. Inlet gas rate is controlled by a Brooks mass flow meter, and reactor pressure is regulated by a 90W valve. The reactors are equipped with quick fit connections allowing them to be withdrawn and catalysts stored under H_2 or N_2 pressure, and re-installed later if desired. Operation is normally during regular working hours only. At the start of a day's run, the sequence of events is: 1) Start feed and hydrogen flow at desired operating pressure. 2) Heat to target reactor temperature as quickly as possible. Sampling from the gas-liquid separator is carried out hourly, with gas and liquid flows and liquid product yield recorded at that time. At the end of the day's operation, the unit is shutdown as follows: 1) Cool down reactor as quickly as possible under gas and liquid flow, to less than $150^\circ C$. 2) Bottle up reactor under H_2 pressure. Feed characterization data is provided in Table 1. Operating conditions were 8.5 MPa H_2 ; 1.0 LHSV, and treat gas rate of 425 ml/minute (~ 6200 SCF/B), with activity varied by reactor temperature between 325 and $380^\circ C$. Lined out total liquid product from each days run was blended and topped to give $360^\circ C+$ product in a HVIAC model C still. Operating conditions, yields and product characterizations for each of the five catalysts are contained in Tables 2 to 7. Data for fluorided $Ni/W/Al_2O_3$ catalyst, a high activity hydrocracking/hydrotreating catalyst, available as Shell 454, is included for comparison purposes and reported in detail in Table 2A. The process of the present invention using bulk $Ni/Mn/Mo$ and $Ni/Cr/Mo$ sulfide catalysts is superior for VI increase and, in the case of the $Ni/Mn/Mo$ sulfide catalyst also superior for aromatics saturation as compared to the other catalysts tested, including fluorided $Ni/W/Al_2O_3$. The process of the present invention, by using the recited bulk sulfide catalysts also avoids the use of any fluoridating agent and, therefore, avoids the corrosion associated with the presence of fluoride in hydrotreating and ancillary equipment.

Activation and Fluoriding of Shell 454 Catalyst

The $Ni/W/Al_2O_3$ catalyst was activated by a procedure recommended by the vendor, as follows: The 20-28 mesh sized catalyst was calcined at $485^\circ C$ for 1 hour prior to loading. Twenty-three cc of

This Sheet Has Been Replaced



REPLACEMENT SHEET

27 JAN 1997

- 14 -

catalyst was charged to the MCS unit reactor. Following cold pressure testing with N_2 and H_2 (8.5 MPa), the reactor was purged with N_2 (6 CF/hr) and heated to 420°C, pressure tested at 8.5 MPa, held at these conditions overnight, then cooled to 107°C. Treat gas was changed to 2% H_2S in H_2 , and sulphiding feed was introduced (1.49 g dimethyl disulphide per 100 g CWO base) at maximum pump rate, 4.1 MPa pressure, and 136 ml/min (0.28 CF/hr) treat gas rate. When liquid flow through the separator had been established, feed rate was reduced to 1 LHSV, and the reactor quickly heated to 190°C, held for 3 hours, heated to 330°C at 47°C per hour, held for 1 hour, then cooled to 50°C with gas and liquid flow.

In-situ fluoriding with 2-fluorotoluene is recommended by the catalyst vendor for Shell 454 Ni/W/ Al_2O_3 catalyst using 400-800 ppm fluoride in feed, but detailed procedures were not provided. Fluoride concentration was 1000 wppm (0.575 g 2-fluorotoluene per 100 g CWO base), and process conditions were 360°C, 4.0 LHSV, 8.5 MPa, 5000 SCF/B treat gas rate. Catalyst activity was monitored by changes in Refractive Index. Through the first day of fluoriding, activity rapidly increased. After the overnight shutdown, activity was lower than at the end of the previous day, and increased again through the second day of fluoriding. During the third day, steady performance was observed, and fluoriding was terminated at this point. At the start of each days' run, fluoride containing feed was pumped through the reactor as the temperature was increased from ambient to 360°C, and again as the reactor was being cooled at the end of the days run. The nominal total amount of fluoride added was 1.51 g or 6.8 wt% fluoride on catalyst. It should be noted that product Refractive Index has been found to be an unreliable indicator of catalyst activity for raffinate hydrotreating. In general, the fluoriding procedures have been found to give a large fluoride gradient through the catalyst bed, with a higher concentration at the feed inlet (bottom of bed for upflow operation). The fluoride distribution resulting from the procedures used in these small lab unit (MCS units) experiments is not known at this time.

THIS SHEET HAS BEEN
REPLACED

This Sheet Has Been Replaced



- 15 -

REPLACEMENT SHEET

27 JAN 1997

TABLE 1

FEEDSTOCK CHARACTERIZATION

Feed Type: Isthmus 600N 82VI Raffinate
Dewaxing Conditions: 100% MBK, 1.5/l solvent/al,
+8°C filter temp.

Feed Inspections

Sulfur, wt% 1.07
Total nitrogen, wppm 280
Basic nitrogen, wppm 159
Viscosity @ 40°C, cSt 130.61
Viscosity @ 100°C, cSt 12.8
Viscosity Index 89
Pour Point, °C +12

HPLC Separation:

Saturates, wt% 55.8
Aromatics, wt% 41.2
Recovery, wt% 96.9

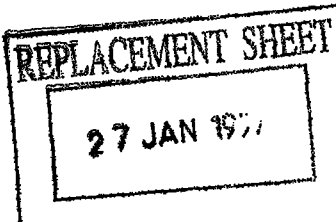
GCD, °C

IBP/5 329/413
20/50 461/495
90/95 568/563
FBP 594

Mass Spectral Composition (LVZ)

Paraffins 7.77
Saturates average ring no 1.74
Alkyl benzenes 13.37
2 ring aromatics 11.03
3 ring aromatics 2.18
4 ring aromatics 0.69

THIS SHEET HAS BEEN
REPLACED



360°C+ product VI as a function of reactor temperature is plotted for the five bulk catalysts in Figure 1. VI is corrected to -9°C pour point by subtracting 6 from the measured values in Tables 3-7 (assumes 1 VI unit per 3°C pour point).

Ni/Mn/Mo (from oxide: molybdate)

The bulk Ni/Mn/Mo sulfide catalyst derived from the oxide shows outstanding start of run activity under these conditions exhibiting the production of oil of high VI at lower temperature than the other catalyst tested with the exception of the Ni/Cr/Mo bulk sulfide derived from the sulfide precursor. The catalyst is about 20°C more active than Ni/W/3-Al₂O₃ in achieving the same VI.

The high initial VI (103) obtained at 340°C gradually declined to 95 VI after 30 hours on oil, and returned to about 102 VI when reactor temperature was increased to 350°C. The plotted runs are those performed at a gas rate of 6200 SCF H₂/bbl. The catalyst is also active for aromatics saturation. The catalyst had a BET surface area of 61.7 m²/g.

Ni/Mn/Mo (from sulfide: thiomolybdate)

This material differs nominally from the first catalyst in the source of molybdenum used. This catalyst, however, has very poor activity for both VI increase and aromatics saturation activity as indicated below:

	Ni/Mn/Mo	Ni/Mn/Mo	
Catalyst	(from oxide)	(from sulfide)	Ni/W/F-Al ₂ O ₃
Reactor Temp°C	340	340	340
Product VI	95	85	89
Product Saturates	85	60	72.5
wt%			

This catalyst had a BET surface area of 12.4 m²/g.

This Sheet Has Been Replaced



- 17 -

REPLACEMENT SHEET

27 JAN 1997

Ni/W (from sulphide)

The bulk nickel tungsten sulphide catalyst showed poorer activity for both VI increase and aromatics saturation. This catalyst had a BET surface area of 0.6 m²/g.

Co/Mo (from oxide: molybdate)

For the first eight hours on oil, this catalyst showed very high activity, comparable to the Ni/Mn/Mo (oxide precursor) catalyst. However, activity thereafter declined rapidly to a low level, suggestive of rapid poisoning of the active sites by some feed or product component, or possibly catalyst structural arrangement. This catalyst had a BET surface area of 50 M²/g.

Ni/Cr/Mo (from sulphide: thiomolybdate)

This catalyst also showed very promising activity, as indicated in Fig. 1. Activity for VI increase was similar to that of Ni Mn Mo (from oxide). Whereas the Ni/Mn/Mo (oxide precursor) catalyst appeared to lose activity somewhat during the course of the evaluation, the Ni/Cr/Mo catalyst activity appeared somewhat steadier. This catalyst showed less aromatics saturation activity than did Ni/Mn/Mo (from oxide: molybdate). This catalyst had a BET surface area of 7.9 m²/g.

THIS SHEET HAS BEEN
REPLACED

27 JAN 1997

This Sheet Has Been Replaced



- 18 -

TABLE 2

MCS UNIT OPERATIONS

UNIT/REACTOR/RUN	MCS/116/1	MCS/113/2	MCS/102/0	MCS/111/9	MCS/111/10
Treat Gas			100% H ₂		
Operating Mode			isothermal, Once through, downflow		
Mass Velocity, kg/hr/m ² (lb/hr/ft ²)			210 (43)		
Length of catalyst bed, cm	23.3	25	23.9	25	23.1
Length of Inert Packing in Hot Zone, cms:					
A) Above catalyst	12.5	11	11.3	10.3	10.9
B) Below catalyst					
Catalyst					
Type	Ni/Mn/Mo oxide	NiW sulphide	CoMo oxide	NiMoMo sulphide	Ni/Cr/Mo sulphide
Precursor Source					
Catalyst charge, ml	22	23	23	23	21
Catalyst weight, g	34.5	34.0	37.7	31.5	30.1
Catalyst size	20 30 mesh	20 30 mesh	20 30 mesh	20 30 mesh	20 30 mesh

Startup procedure

- 1) Ambient temperature pressure test with nitrogen (9 MPa)
- 2) Reactor heated to 100°C under nitrogen purge (6 CF/hr, 1MPa), held for one hour
- 3) Reactor cooled to ambient temperature under nitrogen purge
- 4) Pressure tested with hydrogen at ambient temperature (9 MPa)
- 5) Cut in feed at ambient temperature, heated to reaction temperature under hydrogen flow (425 ml/min, 8.5 MPa)

This Sheet Has Been Replaced



- 19 -

REPLACEMENT SHEET

27 JAN 1997

TABLE 2A

OPERATING CONDITIONS, YIELDS AND PRODUCT INSPECTIONS

Catalyst Description Shell 454, Crushed and Screened to 20-28 mesh, on-stream fluorided
Run Number -----MCS-113-4-----

Operating Conditions and Yields

Feed	Istimus 600N 82 VI Raffinate				
Feed Rate (LHSV)	1.0				
Gas Rate, SCF/B	360	360	370	375	380
Reactor Temp., °C	8.5				
Unit Pressure, MPa	8.5				
Catalyst Hours	16 20	23-27	30 34	37 41	44 48
					51 55
					58 61
TLP Yield on Feed, wt%	93.7	93.2	95.99	97.11	94.7
360°C Yield on Feed, wt%	79.4	84.0	84.4	75.91	79.5
					77.1
					94.3

360°C Product Inspections

Density @ 15°C, kg/dm ³	50	18	17	14	7	5	257
Sulphur, wppm	<1	<1	<1	1	1	<1	16
Total Nitrogen, wppm	65.87	74.47	68.44	67.62	60.33	51.80	101.24
Viscosity @ 40°C, cSt	8.84	9.50	9.02	8.99	8.48	7.81	11.14
Viscosity @ 100°C, cSt	108	105	106	107	112	117	95
Viscosity Index							
HPLC Separation:							
Saturates, wt%	88.6	84.5	84.7	86.3	87.0	85.3	72.7
Aromatics, wt%	11.0	15.2	14.4	13.4	12.7	13.7	26.6
Recovery, wt%	99.6	99.7	99.2	99.7	99.7	98.9	99.3

THIS SHEET HAS BEEN REPLACED



27 JAN 1997

Table 3

OPERATING CONDITIONS, YIELDS AND PRODUCT IDENTIFICATION, MSU

Unit Run #	MS 116 /									
Catalyst Description	Ni _{0.5} Mn _{0.5} Mo (from oxide)									
Unit Conditions	600N 82 VI Raffinate									
Feed										
Feed Rate (liters/h)	6200	6200	6200	6200	6200	6200	6200	6200	6200	6200
Gas Rate, SCFH	340	340	340	340	340	340	340	350	350	350
Reactor Temp., °C	8.5	8.5	8.5	8.5	8.5	8.5	8.5	8.5	8.5	8.5
Unit Pressure, MPa	4.6	9.1	16.20	23.21	30.14	31.41	44.48	51.55	58.62	58.62
Catalyst Hours										
Total Liquid Product										
TIP Yield, wt%	98.1	97.5	96.2	96.5	97.4	97.4	97.4	97.4	96.1	96.1
360°C Yield on Feed, wt%	83.5	85.4	84.1	87.1	89.7	91.9	85.5	85.2	84.3	84.3
Refractive Index @ 15°C	1.4518	1.4585	1.4584	1.4602	1.4622	1.4600	1.4592	1.4580	1.4580	1.4580
Hydrogen, wt% by H ₂ O										
Product Liquid Product										
Visc @ 40°C, cSt	64.17	68.56	68.20	74.29	80.94	73.24	69.12	67.08	67.10	67.10
Visc @ 100°C, cSt	8.82	9.10	9.06	9.46	9.89	9.42	9.11	8.97	8.98	8.98
Viscosity Index	109	108	108	104	101	105	107	108	108	108
Dohman Sulphur, wt%	80	26	23	22	44	21	11	7	10	10
X-Ray Sulphur, wt%										
Total Nitrogen, wt%	3	2	<2	<2	<2	<2	<2	<2	<2	<2
Basic Nitrogen, wt%										
TPIC Separation										
Saturated wt%	95.2	93.0	92.1	90.2	84.4	90.1	91.3	91.5	89.4	89.4
Aromatics, wt%	4.8	6.7	6.5	9.4	12.8	7.9	7.5	6.3	7.1	7.1
Recovery, wt%	100	99.1	99.2	99.6	97.2	91.6	98.8	91.8	91.4	91.4

This Sheet Has Been Replaced



REPLACEMENT SHEET

27 JAN 1997

- 21 -

TABLE 2 (CONT'D)

OPERATING CONDITIONS, YIELDS AND PRODUCT INSPECTIONS, MCSU

Unit No. #	MCS 116 7									
Catalyst Description	MoO ₃ MoS ₂ (from oxide)									
Unit Conditions	600N 82 VI Raffinate									
Feed	1.0									
Feed Rate (LHSV)	6200	500	500	500	500	500	500	500	6200	
Gas Rate, SCF/B	350	340	340	340	355	350	350	350	350	
Reactor Temp., °C	8.5	4.1	4.1	4.1	4.1	4.1	4.1	4.1	8.5	
Unit Pressure, MPa	65-69	75-76	79-83	86-90	93	97	100	103		
Catalyst Hours										
Total Liquid Product										
TLV Yield, wt%	97.7	97.7	99.4	98.4	99.0	97.5				
360°C Yield on Feed, wt%	87.0	94.6	97.2	94.5	95.2	90.8				
Refractive Index @ 75°C	1.4600	1.4717	1.4712	1.4707	1.4690	1.4632				
Hydrogen, wt% by NMR										
260°C Liquid Product										
Visc @ 40°C, cSt	70.67	118.9	109.76	103.54	98.15	86.77				
@ 100°C, cSt	9.21	11.95	11.17	11.24	10.98	10.29				
Viscosity Index	106	88	85	94	96	100				
Dohrman Sulphur, wppm	11									
X-Ray Sulphur, wt%		0.167	0.131	0.089	0.029					
Total Nitrogen, wppm	<2	88	80							
Basic Nitrogen, wppm										
IPLC Separation										
Saturates, wt%	89.8	62.9	62.0	63.6	68.0	83.7				
Aromatics, wt%	9.8	36.2	35.5	36.3	31.5	15.5				
Recovery, wt%	99.5	99.0	97.5	99.9	99.5	99.3				

THIS SHEET HAS BEEN REPLACED

This Sheet Has Been Replaced



- 22 -

REPLACEMENT SHEET

27 JAN 1997

TABLE 4

OPERATING CONDITIONS, YIELDS AND PRODUCT INSPECTIONS, MCSU

Unit Run #	MCS 113 3			
Catalyst Description	NI/VS (from sulphide)			
Unit Conditions				
Feed	600N 82 VI Raffinate			
Feed Rate (LHSV)	1.0			
Gas Rate, SCF/B	6200			
Reactor Temp., °C	340	375	375	380
Unit Pressure, MPa	8.5			
Catalyst Hours	8 12	15-19	22-26	29-33 36-40
Total Liquid Product				
TLP Yield, wt%	96.1	97.8	96.2	94.9 96.6
360°C Yield on Feed, wt%	94.1	87.4	84.6	82.8 79.0
Refractive Index @ 75°C	1.4744	1.4662	1.4657	1.4658 1.4626
Hydrogen, wt% by NMR				
360°C Liquid Product				
Visc @ 40°C, cSt	124.2	85.17	81.5	79.75 65.75
@ 100°C, cSt	12.45	10.21	9.97	9.86 8.91
Viscosity Index	90	100	102	102 110 160
Dohrman Sulphur, wppm				
X-Ray Sulphur, wt%	0.677	0.091	0.071	0.089
Total Nitrogen, wppm	128			
HPLC Separation				
Saturates, wt%	58.9	73.6	74.7	72.4 80.2
Aromatics, wt%	37.6	25.6	23.9	25.9 19.0
Recovery, wt%	96.5	99.2	98.6	98.4 99.2

THIS SHEET HAS BEEN REPLACED

This Sheet Has Been Replaced



REPLACEMENT SHEET

27 JAN 1997

TABLE 5
OPERATING CONDITIONS, YIELDS AND PRODUCT INSPECTIONS, MCSU

Unit Run #	MCS 105 B									
Catalyst Description	CoMo (from oxide)									
Unit Conditions	600N 82 VI Raffinate									
Feed	1.0									
Feed Rate (LHSV)	6200									
Gas Rate, SCF/B	340	340	340	350	350	350	350	350	350	350
Reactor Temp., °C	8.5									
Unit Pressure, MPa	8.5									
Catalyst Hours	7 11	14 18	21 25	28 32	35 39	41 44	54 57	64 61		
Total Liquid Product										
TLP Yield, wt%	95.7	98.8	96.6	98.3	98.3	98.4	98.8	98.8	98.8	98.8
360°C Yield on Feed, wt%	83.4	94.8	94.3	94.0	94.0	95.2	95.4	95.4	97.5	97.5
Refractive Index @ 75°C	1.4622	1.4680	1.4695	1.4672	1.4672	1.4672	1.4679	1.4679		
Hydrogen, wt% by NMR										
360°C Liquid Product										
Visc @ 40°C, cSt	68.74	102.4	107.59	96.05	96.05	96.31	98.14	98.31	86.31	86.31
@ 100°C, cSt	9.08	11.28	11.50	10.81	10.81	10.82	10.95	10.21	10.21	10.21
Viscosity Index	107	96	93	95	95	95	95	99	99	99
Dohrmag Sulphur, wppm	58	440	570	160	160	290	185	140	140	140
X-Ray sulphur, wt%										
Total Nitrogen, wppm										
Basic Nitrogen, wppm										
HPLC Separation										
Saturates, wt%	80.0	62.7	64.6	64.7	64.7	67.0	66.5	69.3	69.3	69.3
Aromatics, wt%	18.9	32.2	34.0	28.4	28.4	30.2	28.3	29.6	29.6	29.6
Recovery, wt%	98.9	94.9	98.6	93.0	93.0	97.1	94.8	98.9	98.9	98.9

THIS SHEET HAS BEEN REPLACED

This Sheet Has Been Replaced



- 24 -

REPLACEMENT SHEET

27 JAN 1977

TABLE 6

OPERATING CONDITIONS, YIELDS AND PRODUCT INSPECTIONS, PCSU

Unit Run	MCS 111-9	
Catalyst Description	Ni _{0.5} Mo _{0.5} Mo (from sulphide)	
Unit Conditions	600N 82 VI Raffinate	
Feed	1.0	
Feed Rate (LHSV)	6200	
Gas Rate, SCF/B	330	340
Reactor Temp., °C	340	360
Unit Pressure, MPa	8.5	8.5
Catalyst Hours	5 8	12 15
		19 22
		26 29
		33 36
		40 43
Total Liquid Product		
T.P. Yield, wt%	96.8	97.8
360°C Yield on Feed, wt%	95.7	97.3
Refractive Index @ 75°C	1.4720	1.4717
Hydrogen, wt% by NMR		
360°C Liquid Product		
Visc @ 40°C, cSt	115.72	112.94
@ 100°C, cSt	11.98	11.77
Viscosity Index	91	91
Dohrman Sulphur, wppm		
X-Ray Sulphur, wt%	0.20	0.16
Total Nitrogen, wppm	110	150
Basic Nitrogen, wppm		
HPLC Separation		
Saturates, wt%	62.5	62.8
Aromatics, wt%	36.6	36.0
		60.3
		34.9
		66.7
		32.8
		66.9
		30.6
		67.1
		27.7
		89.85
		10.43
		98
		260
		11
		101.41
		11.09
		94
		530
		67
		103.17
		11.32
		95
		640
		87
		96.6
		96.1
		1.4698
		98.0
		97.4
		1.4695
		96.5
		95.3
		1.4671

THIS SHEET HAS BEEN REPLACED

This Sheet Has Been Replaced



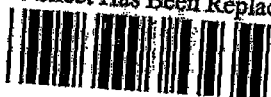
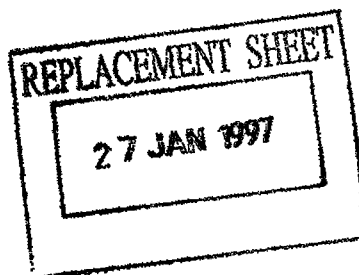
- 25 -

REPLACE THIS SHEET
27 JAN 1971

Table 7
OPERATING CONDITIONS, YIELDS AND PRODUCT INSPECTIONS, MCSU

Unit Run #	MCS 111 10									
Catalyst Description	Ni ₂ , Cr ₂ , Mo (from sulphide)									
Unit Conditions										
Feed	600N 82 VI Raffinate									
Feed Rate (LHSV)	1.0									
Gas Rate, SCF/hr	6200									
Reactor Temp., °C	325	340	340	350	350	350	355	345		
Unit Pressure, MPa	8.5									
Catalyst Hours	11-14	18-21	25-28	32-35	39-42	46-49	53-56	60-63		
Total Liquid Product										
TLP Yield, wt%	97.8	98.5	96.6	93.8	93.2	93.3	93.5	94.3		
360°C Yield on Feed, wt%	97.1	97.6	93.2	82.9	79.1	79.2	79.3	83.7		
Refractive Index @ 75°C	1.4707	1.4879	1.4661	1.4604	1.4555	1.4561	1.4561	1.4594		
Hydrogen, wt% by gas										
360°C Liquid Product										
Visc @ 40°C, cSt	106.07	88.89	85.42	67.50	56.70	51.60	56.68	64.85		
@ 100°C, cSt	11.38	10.30	10.11	9.05	8.13	8.20	8.16	8.81		
Viscosity Index	92.8	96.6	98.2	109.1	111.9	111.4	112.9	109.3		
Dohrmann Sulphur, wppm	58	440	570	160	160	290	185	140		
X-Ray Sulphur, wt%	0.25									
Total Nitrogen, wppm	44	13	5					47		
Basic Nitrogen, wppm										
HPLC Separation										
Saturates, wt%	65.8	63.8	69.0	79.8	88.7	84.5	85.8	83.0		
Aromatics, wt%	35.2	35.7	31.1	18.8	10.1	12.3	10.1	16.0		
Recovery, wt%	101.0	99.4	100.1	98.6	98.8	96.8	95.9	99.0		

THIS SHEET HAS BEEN REPLACED



CLAIMS:

1. A method for increasing the Viscosity Index of lubricating and specialty oils by contacting with a catalyst, said method comprising contacting a lube or specialty oil stock under hydrotreating conditions with a specific bulk metal sulfide hydrotreating catalyst selected from the group consisting of bulk Cr/Ni/Mo sulfide catalyst, bulk Ni/Mo/Mn sulfide catalyst and mixtures thereof, wherein the Cr/Ni/Mo sulfide catalyst is prepared by the decomposition in a sulfide containing atmosphere of a $[\text{Ni}(\text{L})_3](\text{Cr}(\text{L})_3)[\text{MoS}_4]_{2-2.5}$ precursor complexes and the Ni/Mo/Mn sulfide catalyst is prepared by the decomposition in inert atmosphere followed by sulfiding using $\text{H}_2\text{S}/\text{H}_2$ of a $[\text{Ni}(\text{L})_3](\text{Mn}(\text{L})_3)[\text{MoO}_4]_2$ precursor complex where L is one or more neutral, nitrogen containing ligand, at least one of which is a chelating polydentate ligand.

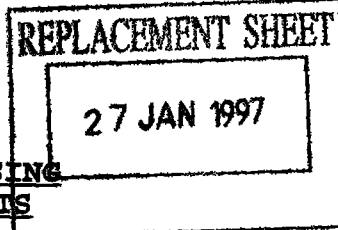
2. The method of claim 1 wherein hydrotreating conditions comprise temperature in the range 200 to 450°C, pressure in the range 300 to 3000 psi, space velocity in the range of about 0.1 to 5.0 v/v/hr and a treat gas rate in the range of about 200 to 20,000 SCF H_2 /bbl.

3. The method of claim 1 or 2 wherein the $[\text{Ni}(\text{L})_3](\text{Mn}(\text{L})_3)[\text{MoO}_4]_2$ precursor is converted into the bulk catalyst by decomposing the precursor in the presence of an inert atmosphere at 300 to 400°C and sulfiding using $\text{H}_2\text{S}/\text{H}_2$ at 200 to 600°C and the $[\text{Ni}(\text{L})_3](\text{Cr}(\text{L})_3)[\text{MoS}_4]_{2-2.5}$ precursor is converted into the bulk catalyst by decomposing the precursor in a sulfur containing atmosphere at a temperature of at least 150°C.

4. The method of claim 3 wherein the ligand is ethylene diamine.

5. The method of claim 1 or 2 wherein the bulk metal sulfide catalyst is used in the form of pellets made by compacting bulk metal sulfide under pressure.





INCREASING VI OR LUBE OIL HYDROTREATING USING
BULK NI MN MO OR NI CR MO SULFIDE CATALYSTS

ABSTRACT

Lube oils of increased VI are prepared by hydrotreating using bulk Ni/Mn/Mo or Cr/Ni/Mo sulfide catalysts prepared from ligated, eg., ethylene diamine, metal complexes. In particular the Ni/Mn/Mo sulfide catalyst is prepared from the oxide precursor with subsequent sulfiding while the Cr/Ni/Mo sulfide catalyst is prepared from the sulfide precursor in the presence of H_2S/H_2 .

FIGURE 1.



THIS SHEET HAS BEEN
REPLACED

LEGEND

- NI/W/F/Al2O3
- NI/Mn/Me(FROM OXIDE)
- NI/W(FROM SULPHIDE)
- Co/Mo(FROM OXIDE)
- △ NI/Mn/Mo(FROM SULPHIDE)
- × NI/Cr/Mo(FROM SULPHIDE)

This Sheet Has Been Replaced

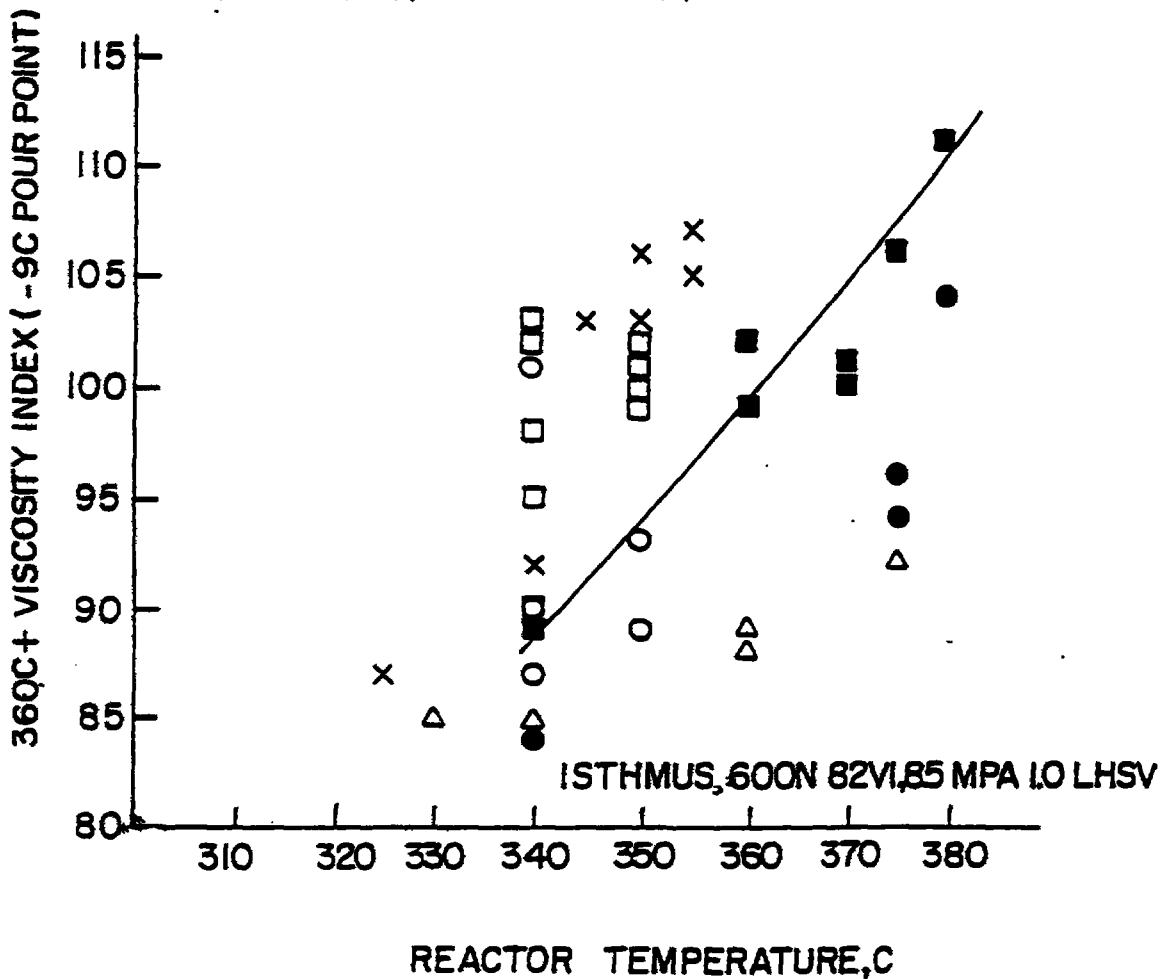


FIG. 1

INCREASING VI OF LUBE OIL BY HYDROTREATING

Description

Description of the invention

The Viscosity Index of lubricating and speciality oils can be increased by hydrotreating said oil employing as the hydrotreating catalyst a bulk Ni/Mn/Mo or Cr/Ni/Mo sulfide catalyst prepared from ligated, e.g. ethylene diamine metal complexes wherein the Ni/Mn/Mo sulfide catalyst is prepared from the oxide precursor with subsequent sulfiding while the Cr/Ni/Mo sulfide catalyst is prepared from the sulfide precursor in the presence of H_2S/H_2 .

Background of the invention

Raffinate hydrotreating and hydrocracking processes to produce lube stocks are well known. In general, these processes employ catalysts which comprise metal oxides or sulfides supported on alumina or silica-alumina. Metals usually include nickel, molybdenum, cobalt, manganese, tungsten, etc. and mixtures thereof in the oxide or sulfide state. When the catalyst is in the oxide state the environment of use being one containing sulfur containing species converts the metal oxides into metal sulfides.

The objectives in the hydrotreater are to increase VI and saturates levels. Catalysts used today in lubes hydrotreating do not perform this function very well insofar as they are employed principally for hydrodesulfurization, hydrodenitrogenation, color improvement and residual extraction solvent removal. High pressure hydrocracking of distillates or deasphalted oil is an acknowledged alternative to solvent extraction and, in the case of distillate hydrocracking all of the VI increase, saturates increase and aromatics decrease are achieved by the catalytic conversion.

The major reactions which contribute to VI increase are of the ring opening type, whether the ring is aromatic or cyclic saturated. Hydrogenation will contribute to a lesser degree to VI increase. At low severity, hydrodesulfurization will give a 3-4 VI unit increase, most likely the result of ring opening that occurs with the hydrodesulfurization of thiophenic structures.

US 4,650,563 teaches molybdenum and/or tungsten hydrotreating catalysts prepared from promoter metal thio-molybdate and/or thio-tungstate complexes and their use as hydrotreating catalysts. The catalysts are prepared by heating one or more catalyst precursors under oxygen-free conditions in the presence of sulfur at a temperature of at least about 150°C. The catalyst precursors will be one or more compounds of the formula $(ML)(Mo_y W_{1-y} S_4)$ wherein M comprises one or more divalent promoter metals selected from the group consisting of Ni, Co, Zn, Cu and mixtures thereof, wherein y is any value ranging from 0 to 1 and wherein L is one or more neutral, nitrogen containing ligands at least one of which is a chelating polydentate ligand. L may be ethylene diamine.

US 4,716,139 teaches a catalyst composition comprising an amorphous sulfide of trivalent chromium and at least one metal selected from the group consisting of (a) molybdenum, tungsten and mixtures thereof and (b) mixtures of (a) with iron, nickel, cobalt, manganese, copper, zinc and mixtures thereof. The catalyst is prepared by heating one or more precursors at a temperature of at least about 200°C under oxygen-free conditions in the presence of sulfur, wherein said precursor is selected from the group consisting of



and mixtures thereof wherein M is one or more divalent promoter metals selected from the group consisting of Mn, Fe, Co, Ni, Cu, Zn and mixtures thereof, wherein L is one or more neutral nitrogen-containing ligands, wherein ligand X is a singly-charged anionic species such as NO_2^- , OH^- , Cl^- , etc. and wherein $1 \geq z \geq 0.1$, $z \geq y \geq 0$ and wherein n equals $(3-z-y)/2$. The catalysts are good hydroprocessing catalysts, particularly hydrotreating catalysts. These catalysts were demonstrated for their HDS and HDN ability on light catalytic cycle oil (LCCO).

See also U.S. Patent 4,626,339 which teaches the supported version of the catalyst of '139.

U.S. 4,663,023 teaches hydrotreating with self promoted molybdenum and Tungsten sulfide catalysts. Catalysts are made from precursors which are one or more compounds of the formula $ML(Mo_y W_{1-y} O_4)$ wherein M is one or more promoter metals selected from the group consisting essentially of Mn, Fe, Co, Ni, Cu, Zn and mixtures thereof $0 \leq y \leq 1$ and L is a neutral nitrogen containing ligand. This patent would, therefore, embrace Mn/Ni/Mo sulfide and Ni/Mo sulfide catalysts. The patent claims the use of these catalysts in general to hydrotreat hydrocarbon feeds including lube oils.

U.S. 4,698,145 teaches the upgrading of hydrocarbon feeds by use of a supported self-promoted catalyst. The catalyst is prepared by heating a composite of support material and one or more catalyst precursor salts under oxygen-free conditions and in the presence of sulfur. The precursor salt or salts are of the general formula $(HL)(Mo_y W_{1-y} S_4)$ wherein M is one or more promoter metals selected from the group consisting essentially of Mn, Fe, Co, Ni, Zn and mixtures thereof, $0 \leq y \leq 1$, and L is a neutral, nitrogen containing ligand. This patent would, therefore, embrace supported Mn/Ni/Mo sulfide and supported Ni/Mo sulfide catalysts. The claims recite lube oil as one of the hydrocarbon feeds

03 DEC 1997

treated using this class of catalysts in general.

U.S. 4,595,672 is directed to a method for making self-promoted sulfide hydrotreating catalysts. The catalysts are made from precursors which are one or more compounds of the formula $ML(Mo_yW_{1-y}O_4)$ where M is one or more promoter metals selected from the group consisting essentially of Mn, Fe, Co, Ni, Cu, Zn and mixtures thereof and $0 \leq y \leq 1$. This patent would, therefore, embrace Ni/Mn/Mo sulfide and Ni/Mo sulfide catalytic materials. L is a neutral nitrogen containing ligand.

U.S. 4,668,376 teaches supported Mn sulfide promoted Mo and W sulfide hydroprocessing catalysts. The catalysts are made from thiometallate precursor salts of the formula $(ML)(Mo_yW_{1-y}S_4)$ where M comprises one or more divalent promoter metals selected from Mn and mixtures of Mn with one or more of Ni, Co, Cu, Zn or mixtures thereof and $0 \leq y \leq 1$ and L is a neutral nitrogen containing ligand. This patent would, therefore, embrace Ni/Mn/Mo sulfide and Ni/Mo sulfide catalytic materials. The patent claims the use of these catalysts in general for hydrotreating hydrocarbon feeds including lube oils.

U.S. 4,591,429 teaches hydroprocessing of hydrocarbon feeds employing supported sulfided catalyst. The catalyst comprises a sulfide of (1) chromium, (2) Mo, W or mixtures thereof and (3) at least one metal selected from the group consisting of Ni, Co, Mn, Cu, Zn and mixtures thereof, and mixtures thereof with Fe. The catalysts are made by compositing a pre-selected quantity of support material with a precursor comprising a mixture of (1) hydrated oxide of trivalent chromium and (2) a salt containing a thiometallate anion of Mo or W and a cation comprising at least one divalent promoter metal. This patent would, therefore, embrace supported Cr/Ni/Mo sulfide catalyst and claims the use of these catalyst in general for hydroprocessing hydrocarbon feeds, including lube oil.

Description of The Figure

Figure 1 compares the 360°C+ product VI as a function of reactor temperature for five bulk metal catalysts and a Ni/W/F/Al₂O₃ reference catalyst.

Detailed Description of the Invention

Lubricating and/or specialty oil stocks such as deasphalted oil stocks, lube oil distillates, and solvent extracted lube oil raffinates can have their viscosity indexes increased by hydrotreating, employing specific bulk metal sulfide hydrotreating catalysts selected from the group consisting of bulk Cr/Ni/Mo sulfide catalyst, bulk Ni/Mo/Mn sulfide catalyst and mixtures thereof wherein the catalysts are prepared from specific metal complexes and wherein the Ni/Mn/Mo sulfide catalyst is prepared from the oxide precursor decomposed in an inert atmosphere such as N₂ and subsequently sulfided using H₂S/H₂ and the Cr/Ni/Mo sulfide catalyst is prepared from the sulfide precursor and decomposed in a non-oxidizing, sulfur containing atmosphere.

These two catalyst have been found to significantly and, surprisingly, to a higher degree than other catalysts, increase the Viscosity Index of lube oil stocks. The feed stocks are contacted with the catalyst under hydrotreating conditions comprising temperatures in the range of about 200 to 450°C, preferably about 300 to 380°C, pressures in the range of about 300 to 3000 psi (20.7 to 206.9 bar), preferably at least about 1000 psi (69 bar), space velocity in the range of about 0.1 to 5.0 v/v/hr., preferably about 0.5 to 2.0 v/v/hr. and a treat gas rate in the range of about 200 to 20,000 SCF H₂/bbl (35.6 to 3557.9 m³H₂/m³ feed), preferably about 700 to 5,000 SCF H₂/bbl (124.5 to 889.5 m³H₂/m³ feed).

The catalysts are prepared by using $[Ni(L)_3](Mn(L)_3)[MoO_4]_2$ or $[Ni(L)_3][Cr(L)_3][MoO_4]_{2-2.5}$ as precursor complexes, where L = one or more neutral, nitrogen containing ligands at least one of which is a chelating polydentate ligand such as ethylenediamine. These complexes are prepared by reaction of appropriate metal salts with the ligand solution.

The ligand or ligands L will generally have a denticity of six and will be one or more neutral, nitrogen containing ligands wherein at least one of said ligands is a multi-dentate chelating ligand which chelates the metal cation of the metal salt to form a chelated metal species such as $[ML]^{2+}$, where M is Ni or Cr and Mn.

Those skilled in the art know that the term "ligand" is used to designate functional coordinating groups which have one or more pairs of electrons available for the formation of coordinate bonds. Ligands that can form more than one bond with a metal ion are called polydentate while ligands that can form only one bond with a metal ion are called monodentate. Monodentate ligands are not capable of forming chelates. Hence, if one uses one or more species of monodentate ligands in the precursor molecule, then one must also use at least one polydentate chelating ligand. Preferably L will be one or more polydentate chelating ligands. The denticity of the ligand L will generally be six, because the promoter metal cations prefer six-fold coordination. Hence, if more than one species of ligand is employed in the precursor molecule, the denticity of the ligand species will usually add up to six. It should be understood that it is possible for ligand L to have a total denticity of less than six, but in most cases L will have a total denticity of six. Thus, L will be three bidentate ligands, two tridentate ligands, a mixture of a bidentate and a quadridentate ligand, a hexadentate ligand or a mixture of a polydentate ligand with monodentate ligands so long as the combination has a total

denticity of six. As has heretofore been stated, it is preferred to use chelating bidentate and tridentate ligands. In general, the ligands useful in this invention include alkyl and aryl amines and nitrogen heterocycles. Illustrative but non-limiting examples of ligands useful in the catalyst precursors of this invention are set forth below.

Monodentate ligands will include NH_3 as well as alkyl and aryl amines such as ethyl amine, dimethyl amine, pyridine, etc. Useful chelating bidentate amine ligands are illustrated by ethylenediamine, 2,2'-bipyridine, o-phenylene diamine, tetramethylethylenediamine and propane-1,3 diamine. Similarly, useful chelating tridentate amine ligands are represented by terpyridine and diethylenetriamine while triethylenetetramine is illustrative of a useful chelating quadridentate amine ligands. Useful chelating pentadentate ligands include tetraethylenepentamine while sepuichrate (an octazacryptate) is illustrative of a suitable chelating hexadentate ligand. However, as a practical matter it will be preferred to use chelating, polydentate alkyl amines for L. Illustrative, but not limiting examples of alkyl amines that are useful in the catalyst precursor of this invention include ethylenediamine, diethylenetriamine, and tetraethylenetetramine. It is particularly preferred to use bidentate and tridentate alkyl amines such as ethylenediamine, (en) and diethylenetriamine, (dien).

Many of the precursor salts useful in forming the catalysts of this invention and methods for preparing them are known in the art. An article by Diemann and Mueller titled Thio and Seleno Compounds of the Transition Metals With d^0 configuration published in COORD. CHEM. REV. 10:79-122 provides a review of known salts. In general, the precursor salts useful for forming the catalysts useful in this invention may be prepared by mixing an aqueous solution of ammonium thiomolybdate with an aqueous solution of the chelated promoter metal cation $[\text{ML}]^{2+}$ which results in the formation of the precursor salt as a precipitate which is readily recovered. The chelating promoter cation is easily formed by, for example, mixing an aqueous solution of one or more water soluble promoter metal salts with the ligand or mixture of ligands. The water soluble salt may be any water soluble salt that is convenient to use such as a halide, sulfate, perchlorate, acetate, nitrate etc. Alternatively, an aqueous solution of ammonium thiomolybdate may be mixed with the ligand with the resulting solution mixed with an aqueous solution of promoter metal salt or the salt can be added to the ligand and dissolved into the solution of thiomolybdate and/or thiotungstate. However, it should be understood that the catalyst precursor preparation is not intended to be limited to aqueous media.

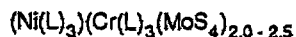
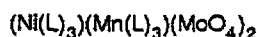
The $[\text{Ni}(\text{L})_3](\text{Mn}(\text{L})_3)(\text{MoO}_4)_2$ precursor is converted into the bulk catalyst by decomposing the precursor complex under inert conditions, such as in the presence of N_2 or an inert gas at 300 to 400°C. The decomposed material is subsequently sulfided using $\text{H}_2\text{S}/\text{H}_2$ at 200 to 600°C, preferably 250-500°C, most preferably 200-400°C. A preferred ligand is ethylene diamine.

The $[\text{Ni}(\text{L})_3](\text{Cr}(\text{L})_3)(\text{MoS}_4)_{2.5}$ precursor is converted into the bulk sulfide catalyst by decomposing the precursor complex in a sulfur-containing non-oxidizing atmosphere, preferably at a temperature of at least 150°C, more preferably at least about 200°C, for a time sufficient to form the catalyst. A preferred ligand is ethylenediamine.

The sulfur required during the formation of the catalyst may be that which is present in the precursor salt. Thus, the Ni/Cr/Mo catalyst composition used in this invention will be formed if no excess sulfur is present and if the oxygen-free atmosphere is relatively inert, such as nitrogen. It is preferred that the sulfur will be present in an amount in excess of that contained in the precursor salt. It is preferred that the catalyst be formed by heating the precursor in the presence of excess sulfur, it is also preferred that the excess sulfur be present in the form of a sulfur bearing compound which can be one or more solids, liquids, gases or mixtures thereof. Mixtures of hydrogen and H_2S have been found to be particularly suitable. In general the temperature will range between from about 150°-600°C, preferably 250-500°C, and still more preferably from about 300°-400°C. The non-oxidizing atmosphere may be gaseous, liquid or mixture thereof.

These bulk metal sulfide catalysts are relatively low surface area materials of about 50m²/gram.

The catalyst is prepared in a two step operation. Appropriate salts of Ni, Cr, and Mn are reacted in the presence of excess ligand in combination with MoS_4^{2-} or MoO_4^{2-} . Methanol is used as the solvent. One recovers either



which are then decomposed as previously described.

The bulk metal sulfide catalysts are most conveniently used as shaped pellets or extrudates. The bulk metal sulfide can be pelleted by the application of extremely high pressure. Bulk metal sulfide catalyst pellets have been prepared using what is believed to be this high pressure technique but the details are not known. Another technique involves preparing the bulk metal sulfide in the presence of colloidal silica or colloidal alumina, recovering the precipitate, drying, and grinding the recovered cake and pelletizing using polyvinyl alcohol, in accordance with the procedure of U.S. Patent 4,650,563 and U.S. Patent 4,591,429. It is possible that a preferred bulk metal sulfide catalyst pellet or extrudate

03 DEC 1997

is produced without the use of any silica or alumina solely by employing standard, commercial high pressure pelleting, but economics and cost effectiveness will become important issues.

Experimental

Five bulk metal sulfide catalysts were evaluated for raffinate hydrotreating to determine their effect on VI improvement. These catalysts were:

- 1) $\text{Ni}_{0.5}\text{Cr}_{0.5}\text{Mo}$ (from thiomolybdate)
- 2) $\text{Ni}_{0.5}\text{Mn}_{0.5}\text{Mo}$ (from thiomolybdate)
- 3) NiW (from thiotungstate)
- 4) $\text{Ni}_{0.5}\text{Mn}_{0.5}\text{Mo}$ (from molybdate)
- 5) CoMo (from molybdate)

These bulk materials were pelleted using high pressure without any silica or alumina binder material. The pelleted material was then crushed and sieved to secure particles of about 1/16 inch (1.5875 mm).

The following are representative of typical procedures used to produce materials of the type identified above. They are not offered as recitations of the exact, specific procedures used to produce the particular materials hereinafter identified.

Typical Preparation of $\text{NiMn}(\text{NH}_2\text{CH}_2\text{CH}_2\text{NH}_2)_6(\text{MoO}_4)_2$ with subsequent sulfiding

88.2g $(\text{NH}_4)_6\text{Mo}_7\text{O}_{24}\cdot 4\text{H}_2\text{O}$ were dissolved in 200 ml H_2O plus 100 ml ethylenediamine. 72.2g $\text{Ni}(\text{NO}_3)_2\cdot 6\text{H}_2\text{O}$ and 49.5g $\text{MnCl}_2\cdot 4\text{H}_2\text{O}$ were dissolved in 800 ml deaerated (N_2 bubbler) methanol. The molybdate solution was placed in a 2-liter flask and deaerated with a nitrogen bubbler while stirring with a mechanical stirrer. Under a blanket of N_2 , the Ni- and Mn- containing solution was added dropwise to the flask from a dropping funnel. A precipitate formed. The precipitate was allowed to settle and the supernatant liquid was siphoned off. Two methanol (deaerated) rinsings were performed, each time siphoning off supernatant-liquid. The precipitate was finally recovered by vacuum filtration from methanol in an N_2 atmosphere box. The yield was 87.2%. The catalyst precursor was dried under vacuum at 50°C overnight.

The precursor was then pelletized into a disk and then crushed into 20/40 (Tyler) mesh (840-420 μm) granules. Sulfiding of the precursor compound to bulk catalysts was done by heating under nitrogen for 50 minutes to 370°C , holding at 370°C for one hour, then heating under a stream of 10% H_2S in hydrogen for ten minutes to 400°C and holding at that temperature under $\text{H}_2\text{S}/\text{H}_2$ for two hours. Cooling over 1.5 hrs under $\text{H}_2\text{S}/\text{H}_2$ to $<100^\circ\text{C}$ was followed by cooling to room temperature under nitrogen and passivation at room temperature with a stream of 1% O_2/N_2 for half an hour.

Typical Preparation of Cr-Containing Catalysts

Chromium containing catalyst can be prepared via two different chemistries, ololation and chelation in non-rigorous terms. The following reported procedures are representative of the typical techniques employed in producing the material from the thiomolybdate.

Via Ololation: A $\text{Ni}_{0.5}\text{Cr}_{0.5}\text{Mo}$ catalyst can be prepared as follows: 40g $(\text{NH}_4)_2\text{MoS}_4$ was added to 82 ml of diethylenetriamine (dien) in a one liter flask. Distilled H_2O was used twice to wash off any solution remaining on the sides of the flask. The resulting dark red solution was cooled to 0°C in an ice bath and kept in the bath for the duration of the preparation. In a separate flask, 16.52 grams of $\text{CrCl}_3\cdot 6\text{H}_2\text{O}$ and 14.77 gm of $\text{NiCl}_2\cdot 6\text{H}_2\text{O}$ were dissolved into a mixture of 250 ml of distilled H_2O and 25 ml of diethylenetriamine to precipitate the hydrated chromium oxide and to chelate the nickel ion. This slurry was allowed to stand for 2-3 hours after which it was slowly added dropwise, to the cold $(\text{NH}_4)_2\text{MoS}_4/\text{dien}$ solution with agitation after each addition keeping the flask as cold as possible, which resulted in a bright orange precipitate being formed. The resulting precipitate-containing mixture was stirred in the ice bath for one half hour after the addition was completed. The precipitate was separated out by vacuum filtration through a Buchner funnel. The resulting cake was washed with distilled water, then with ethanol, and dried under vacuum for 16-24 hrs. Eighty-three grams of orange colored precipitate were recovered. This material was sulfided in the same way as was the $\text{NiMn}(\text{en})_6(\text{MoO}_4)_2$ material, recited above.

Via chelation: A $\text{Ni}_{0.5}\text{Cr}_{0.5}\text{Mo}_{1.25}$ can be prepared as follows: 38 grams of $(\text{NH}_4)_2\text{MoS}_4$ was dissolved in 100 ml degassed ethylene diamine and cooled in an ice bath. Previously prepared $[\text{Cr}(\text{en})_3]\text{Cl}_3$ (16.4g) and $[\text{Ni}(\text{en})_3]\text{Cl}_2$ (16.7g), synthesized in the conventional manner by methanol precipitation from an aqueous solution of $\text{NiCl}_2\cdot \text{H}_2\text{O}$ to which ethylene diamine had been added, were dissolved in a mixture of 150 ml water and 5 ml ethylene diamine. This solution

03 DEC 1997

was filtered. The clear filtrate was added dropwise to the $(\text{NH}_4)_2\text{MoS}_4$ solution with vigorous agitation. A red-orange precipitate formed. After addition was complete the red-orange product was filtered. It was washed with a mixture of 450 ml H_2O and 60 ml ethylene diamine in 3 washes, followed by a 300 ml methanol wash and 100 ml diethylether wash. After vacuum drying the yield was about 45 g. - This material was sulfided in the same way as was the $\text{NiMn}(\text{en})_6(\text{MoO}_4)_2$ material, recited above.

Typical Preparation of $\text{Ni}(\text{en})_3\text{WS}_4$

125 grams of $(\text{NH}_4)_2\text{WS}_4$ were placed into a two liter flask and 250 cm^3 of ethylenediamine was slowly added. The flask was then placed in an ice bath and cooled to about 0°C . Then a mixture of 86.2 grams $\text{NiCl}_2 \cdot 6\text{H}_2\text{O}$ (4% excess) in 500 ml distilled water was added dropwise, keeping the flask contents as cold as possible. After all the reactants were combined the flask was left in the ice bath for 1/2 hour. The contents were then vacuum filtered and the solids washed with distilled water and ethanol. The catalyst precursor was pressed into 12-20 (Tyler) mesh (1410-840 μm) granules using a 2.25 inch (5.715 cm) die at 50,000 psi (3448.3 bar). 25 grams of the 12-20 (Tyler) mesh (1410-840 μm) $\text{Ni}(\text{en})_3\text{WS}_4$ was treated with 15% H_2S (balance H_2) at 350°C for 2 hours then cooled and treated with N_2 bubbling through H_2O to possibly for 1 hour. The procedure was repeated 4 times, then all the products were combined, re-pelleted and sieved to a 20/40 (Tyler) mesh (840-420 μm) size.

Typical Preparation of $\text{Co}(\text{en})_3\text{MoO}_4$

57.38 grams of $(\text{NH}_4)_6\text{Mo}_7\text{O}_{24} \cdot 4\text{H}_2\text{O}$ was put into a two liter flask and 125 cm^3 of distilled water and 250 cm^3 of ethylenediamine were added. The flask was put into a wet ice bath and cooled to about 0°C with N_2 bubbling through the solution. A solution of 77.38 grams of $\text{CoCl}_2 \cdot 6\text{H}_2\text{O}$ in 500 ml of methanol (methanol had had N_2 sparged through it for 1 hr) was added dropwise. After all the reactants were combined the solution was stirred in the ice bath for 1/2 hour then vacuum filtered under N_2 and washed with four 300 ml portions of methanol then put into a vacuum oven with a N_2 bleed overnight. This material was sulfided in the same way as was the $\text{NiMn}(\text{en})_6(\text{MoO}_4)_2$ material, recited above.

Typical preparation of $\text{Ni}_{0.5}\text{Mn}_{0.5}(\text{en})_3\text{MoS}_4$

100 grams of $(\text{NH}_4)_2\text{MoS}_4$ was placed in a 2 liter flask and 200 cm^3 of ethylenediamine was added slowly after all the ethylenediamine was added the flask was put into a wet ice bath and cooled to about 0°C with stirring. A mixture of 45.67 gm $\text{NiCl}_2 \cdot 6\text{H}_2\text{O}$ (1% excess) and 38.40 grams $\text{MnCl}_2 \cdot 4\text{H}_2\text{O}$ (1% excess) in 500 cc distilled water was then added dropwise. Keeping the mixture as cold as possible after all the mixture was added, the flask was left in the ice bath with stirring for 1/2 hour then vacuum filtered. The solids were washed with distilled water and ethanol and left overnight under N_2 . The solids were then dried in a vacuum oven. This material was sulfided in the same way as was the $\text{NiMn}(\text{en})_6(\text{MoO}_4)_2$ material, recited above.

Test Procedure:

Catalysts were tested on a partially dewaxed (+8°C pour) Isthmus 600N raffinate (82 VI @ -9°C pour) in a Micro Catalyst Screening Unit (MCSU). In the MCSU, the 1.18 cm ID 347 stainless steel reactor contains a 0.32 cm OD thermowell. The reactor is heated by a Lindberg electric furnace. Typical catalyst bed height is 24 cm (23 ml charge), and the catalyst charge (20-28 (Tyler) mesh (840-590 μm) size unless stated otherwise) is supported in the hot zone by a 140 μm stainless steel filter. The unit operated isothermally in a down-flow mode, and cannot run waxy feeds. Inlet gas rate is controlled by a Brooks mass flow meter, and reactor pressure is regulated by a 90W valve. The reactors are equipped with quick fit connections allowing them to be withdrawn and catalysts stored under H_2 or N_2 pressure, and re-installed later if desired. Operation is normally during regular working hours only. At the start of a day's run, the sequence of events is: 1) Start feed and hydrogen flow at desired operating pressure. 2) Heat to target reactor temperature as quickly as possible. Sampling from the gas-liquid separator is carried out hourly, with gas and liquid flows and liquid product yield recorded at that time. At the end of the day's operation, the unit is shutdown as follows: 1) Cool down reactor as quickly as possible under gas and liquid flow, to less than 150°C . 2) Bottle up reactor under H_2 pressure. Feed characterization data is provided in Table 1. Operating conditions were 8.5 MPa H_2 ; 1.0 LHSV, and treat gas rate of 425 ml/minute (~ 6200 SCF/B ; 1103 m^3 gas/ m^3 feed), with activity varied by reactor temperature between 325 and 380°C . Lined out total liquid product from each days run was blended and topped to give $360^\circ\text{C}+$ product in a HIVAC model C still. Operating conditions, yields and product characterizations for each of the five catalysts are contained in Tables 2 to 7. Data for fluorided $\text{Ni/W/Al}_2\text{O}_3$ catalyst, a high activity hydrocracking/hydrotreating catalyst, available as Shell 454, is included for comparison purposes and reported in detail in Table 2A. The process of the present invention

03 DEC 1991

using bulk Ni/Mn/Mo and Ni/Cr/Mo sulfide catalysts is superior for VI increase and, in the case of the Ni/Mn/Mo sulfide catalyst also superior for aromatics saturation as compared to the other catalysts tested, including fluorided Ni/W/Al₂O₃. The process of the present invention, by using the recited bulk sulfide catalysts also avoids the use of any fluoriding agent and, therefore, avoids the corrosion associated with the presence of fluoride in hydrotreating and ancillary equipment.

Activation and Fluoriding of Shell 454 Catalyst

The Ni/W/Al₂O₃ catalyst was activated by a procedure recommended by the vendor, as follows: The 20-28 (Tyler) mesh (840-590 μ m) sized catalyst was calcined at 485°C for 1 hour prior to loading. Twenty-three cm³ of catalyst was charged to the MCS unit reactor. Following cold pressure testing with N₂ and H₂ (8.5 MPa), the reactor was purged with N₂ (6 CF/hr; 0.17 m³/h) and heated to 420°C, pressure tested at 8.5 MPa, held at these conditions overnight, then cooled to 107°C. Treat gas was changed to 2% H₂S in H₂, and sulphiding feed was introduced (1.49 g dimethyl disulphide per 100 g CWO base) at maximum pump rate, 4.1 MPa pressure, and 136 ml/min (0.28 CF/hr; 0.05 m³/h) treat gas rate. When liquid flow through the separator had been established, feed rate was reduced to 1 LHSV, and the reactor quickly heated to 190°C, held for 3 hours, heated to 330°C at 47°C per hour, held for 1 hour, then cooled to 50°C with gas and liquid flow.

In-situ fluoriding with 2-fluorotoluene is recommended by the catalyst vendor for Shell 454 Ni/W/Al₂O₃ catalyst using 400-800 ppm fluoride in feed, but detailed procedures were not provided. Fluoride concentration was 1000 wppm (0.575 g 2-fluorotoluene per 100 g CWO base), and process conditions were 360°C, 4.0 LHSV, 8.5 MPa, 5000 SCF/B (889.5 m³ gas/m³ feed) treat gas rate. Catalyst activity was monitored by changes in Refractive Index. Through the first day of fluoriding, activity rapidly increased. After the overnight shutdown, activity was lower than at the end of the previous day, and increased again through the second day of fluoriding. During the third day, steady performance was observed, and fluoriding was terminated at this point. At the start of each days' run, fluoride containing feed was pumped through the reactor as the temperature was increased from ambient to 360°C, and again as the reactor was being cooled at the end of the days run. The nominal total amount of fluoride added was 1.51 g or 6.8 wt% fluoride on catalyst. It should be noted that product Refractive Index has been found to be an unreliable indicator of catalyst activity for raffinate hydrotreating. In general, the fluoriding procedures have been found to give a large fluoride gradient through the catalyst bed, with a higher concentration at the feed inlet (bottom of bed for upflow operation). The fluoride distribution resulting from the procedures used in these small lab unit (MCS units) experiments is not known at this time.

TABLE 1

FEEDSTOCK CHARACTERIZATION	
Feed Type:	Isthmus 600N 82VI Raffinate
Dewaxing Conditions:	100% MBK, 1.5/l solvent/al, +8°C filter temp.
Feed Inspections	
Sulfur, wt%	1.07
Total nitrogen, wppm	280
Basic nitrogen, wppm	159
Viscosity @ 40°C, cSt (or 10 ⁶ m ² /s)	130.61
Viscosity @ 100°C, cSt (or 10 ⁶ m ² /s)	12.8
Viscosity Index	89
Pour Point, °C	+12

03 DEC 1997

TABLE 1 (continued)

FEEDSTOCK CHARACTERIZATION	
HPLC Separation:	
Saturates, wt%	55.8
Aromatics, wt%	41.2
Recovery, wt%	96.9
GCD, °C	
IBP/5	329/413
20/50	461/495
90/95	568/563
FBP	594
Mass Spectral Composition (LV%)	
Paraffins	7.77
Saturates average ring no	1.74
Alkyl benzenes	13.37
2 ring aromatics	11.03
3 ring aromatics	2.18
4 ring aromatics	0.69

360°C+ product VI as a function of reactor temperature is plotted for the five bulk catalysts in Figure 1. VI is corrected to -9°C pour point by subtracting 6 from the measured values in Tables 3-7 (assumes 1 VI unit per 3°C pour point).

Ni/Mn/Mo (from oxide: molybdate)

The bulk Ni/Mn/Mo sulfide catalyst derived from the oxide shows outstanding start of run activity under these conditions exhibiting the production of oil of high VI at lower temperature than the other catalyst tested with the exception of the Ni/Cr/Mo bulk sulfide derived from the sulfide precursor. The catalyst is about 20°C more active than Ni/W/3-Al₂O₃ in achieving the same VI.

The high initial VI (103) obtained at 340°C gradually declined to 95 VI after 30 hours on oil, and returned to about 102 VI when reactor temperature was increased to 350°C. The plotted runs are those performed at a gas rate of 6200 SCF H₂/bbl (1102.9 m³H₂/m³ feed). The catalyst is also active for aromatics saturation. The catalyst had a BET surface area of 61.7 m²/g.

Ni/Mn/Mo (from sulfide: thiomolybdate)

This material differs nominally from the first catalyst in the source of molybdenum used. This catalyst, however, has very poor activity for both VI increase and aromatics saturation activity as indicated below:

Catalyst	Ni/Mn/Mo (from oxide)	Ni/Mn/Mo (from sulfide)	Ni/W/F-Al ₂ O ₃
Reactor Temp°C	340	340	340
Product VI	95	85	89
Product Saturates wt%	85	60	72.5

This catalyst had a BET surface area of 12.4 m²/g.

NiW (from sulphide)

The bulk nickel tungsten sulphide catalyst showed poorer activity for both VI increase and aromatics saturation. This catalyst had a BET surface area of 0.6 m²/g.

Co/Mo (from oxide: molybdate)

For the first eight hours on oil, this catalyst showed very high activity, comparable to the Ni/Mn/Mo (oxide precursor) catalyst. However, activity thereafter declined rapidly to a low level, suggestive of rapid poisoning of the active sites by some feed or product component, or possibly catalyst structural arrangement. This catalyst had a BET surface area of 50 m²/g.

Ni/Cr/Mo (from sulphide: thiomolybdate)

This catalyst also showed very promising activity, as indicated in Fig. 1. Activity for VI increase was similar to that of Ni Mn Mo (from oxide). Whereas the Ni/Mn/Mo (oxide precursor) catalyst appeared to lose activity somewhat during the course of the evaluation, the Ni/Cr/Mo catalyst activity appeared somewhat steadier. This catalyst showed less aromatics saturation activity than did Ni/Mn/Mo (from oxide: molybdate). This catalyst had a BET surface area of 7.9 m²/g.

TABLE 2

PCB UNIT OPERATIONS

UNIT/REACTOR/RUN	PCB/11677	PCB/11773	PCB/10570	PCB/11172	PCB/11170
Treat Gas			100% N ₂		
Operating Mode			isothermal. Once through, downflow		
Mass Velocity, kg/hr/m ² (lb/hr/ft ²)			210 (43)		
Length of catalyst bed, cm	23.3	25	23.9	25	23.1
Length of Inert Packing in Hot Zone, cm:					
A) Above catalyst	12.5	11	11.3	10.3	10.9
B) Below catalyst					

9

Catalyst

Type	PCB/11677	PCB/11773	PCB/10570	PCB/11172	PCB/11170
Precursor Source					
Catalyst charge, ml	22	23	23	23	23
Catalyst weight, g	34.5	34.6	31.7	31.5	30.1
Catalyst size (Tyler) µm	20-30 mesh 840-545	20-30 mesh 840-545	20-30 mesh 840-545	20-30 mesh 840-545	20-30 mesh 840-545

STARTUP PROCEDURE

- 1) Ambient temperature pressure test with nitrogen (9 MPa)
- 2) Reactor heated to 100°C under nitrogen purge (6 CF/hr (0.17 m³/h), 1MPa), held for one hour
- 3) Reactor cooled to ambient temperature under nitrogen purge
- 4) Pressure tested with hydrogen at ambient temperature (9 MPa)
- 5) Cut in feed at ambient temperature, heated to reaction temperature under hydrogen flow (425 ml/min, 8.5 MPa)

TABLE 2A

Catalyst Description	Run Number
Shell 454, Crushed and Screened to 20-28 mesh, on-stream fluorided	-----MCS-113-4-----

Operating Conditions

Feed	Feed Rate (LHSV)	Gas Rate	Reactor Temp., °C	Unit Pressure, MPa	Catalyst Hours
Isotamus 600N B2 VI Raffinate	360	360	370	370	380
6200 SCF/B (1102.9m ³ gas/m ³ feed)	16.20	23.27	30.34	37.41	44.48
				8.5	51.55
					58.61

Run	Temp. °C.	Time, hr.	Yield on Feed, wt%	Yield on Feed, wt%
1	350	1	100	100
2	350	2	100	100
3	350	3	100	100
4	350	4	100	100
5	350	5	100	100
6	350	6	100	100
7	350	7	100	100
8	350	8	100	100
9	350	9	100	100
10	350	10	100	100
11	350	11	100	100
12	350	12	100	100
13	350	13	100	100
14	350	14	100	100
15	350	15	100	100
16	350	16	100	100
17	350	17	100	100
18	350	18	100	100
19	350	19	100	100
20	350	20	100	100
21	350	21	100	100
22	350	22	100	100
23	350	23	100	100
24	350	24	100	100
25	350	25	100	100
26	350	26	100	100
27	350	27	100	100
28	350	28	100	100
29	350	29	100	100
30	350	30	100	100
31	350	31	100	100
32	350	32	100	100
33	350	33	100	100
34	350	34	100	100
35	350	35	100	100
36	350	36	100	100
37	350	37	100	100
38	350	38	100	100
39	350	39	100	100
40	350	40	100	100
41	350	41	100	100
42	350	42	100	100
43	350	43	100	100
44	350	44	100	100
45	350	45	100	100
46	350	46	100	100
47	350	47	100	100
48	350	48	100	100
49	350	49	100	100
50	350	50	100	100
51	350	51	100	100
52	350	52	100	100
53	350	53	100	100
54	350	54	100	100
55	350	55	100	100
56	350	56	100	100
57	350	57	100	100
58	350	58	100	100
59	350	59	100	100
60	350	60	100	100
61	350	61	100	100
62	350	62	100	100
63	350	63	100	100
64	350	64	100	100
65	350	65	100	100
66	350	66	100	100
67	350	67	100	100
68	350	68	100	100
69	350	69	100	100
70	350	70	100	100
71	350	71	100	100
72	350	72	100	100
73	350	73	100	100
74	350	74	100	100
75	350	75	100	100
76	350	76	100	100
77	350	77	100	100
78	350	78	100	100
79	350	79	100	100
80	350	80	100	100
81	350	81	100	10

Run	Temp., °C	Time, hr	Yield on Feed, wt%	Yield on Feed, wt%	Yield on Feed, wt%	Yield on Feed, wt%
1	360	1	93.7	97.2	95.99	97.11
2	360	1	79.4	84.0	84.4	75.91
3	360	1				94.7
4	360	1				94.9
5	360	1				77.1
6	360	1				94.3

360°C Product Inspections

Density @ 15°C. kg/dm ³									
Sulphur, wppm	50	18	17	14	7	5	25.7		
Total Nitrogen, wppm	<1	<1	<1	1	1	<1	16		
Viscosity @ 40°C. cSt or 10 ⁶ m ² /s	65.87	74.47	68.44	67.62	60.33	51.80	101.24		
Viscosity @ 100°C. cSt or 10 ⁶ m ² /s	8.84	9.50	9.02	8.99	8.48	7.81	11.14		
Viscosity Index	108	105	106	107	112	117	95		
HPLC Separation:									
Saturates, wt%	88.6	84.5	84.7	86.3	87.0	85.3	72.7		
Aromatics, wt%	11.0	15.2	14.4	13.4	12.7	13.7	26.6		
Recovery, wt%	99.6	99.7	99.2	99.7	99.7	98.9	99.3		

[illegible]

Feed		600N B2 VI Raffinate									
Feed Rate (lit/hr)		1.0									
Gas Rate, SCF/hr ($m^3/gas/m^3$ feed)		6200 (1102.9)	6200 (1102.9)	6200 (1102.9)	6200 (1102.9)	6200 (1102.9)	6200 (1102.9)	6200 (1102.9)	6200 (1102.9)	6200 (1102.9)	6200 (1102.9)
Reactor Temp., °C		340	340	340	340	340	340	340	340	340	340
Unit Pressure, MPa		8.5	8.5	8.5	8.5	8.5	8.5	8.5	8.5	8.5	8.5
Catalyst Hours		4.6	9.11	16.20	23.21	30.14	37.41	44.48	51.55	58.62	65.69
Total Liquid Product											
TIP Yield, wt%		98.1	97.5	96.7	96.5	97.2	97.4	97.5	97.4	97.1	96.7
100°C Yield on Feed, wt%		43.5	55.4	64.1	67.1	69.1	71.9	75.5	78.2	81.1	84.1
Refractive Index @ 15°C		1.4510	1.4585	1.4584	1.4602	1.4622	1.4640	1.4652	1.4668	1.4680	1.4690
Hydrogen, wt% by HMT											
360°C Liquid Product											
Visc @ 60°C, cSt (or $10^6 m^2/s$)		64.17	68.56	68.20	74.29	80.94	73.24	69.12	67.00	67.10	67.10
Viscosity Index @ 100°C, cSt (or $10^6 m^2/s$)		8.02	9.10	9.06	9.46	9.89	9.62	9.11	8.97	8.98	8.98
Dobson Sulphur, wt%		109	108	108	104	101	105	103	108	108	108
N-Ray Sulphur, wt%		80	26	23	22	44	21	11	1	1	1
Total Nitrogen, wt%		3	2	4.2	4.2	4.2	4.2	4.2	4.2	4.2	4.2
Basic Nitrogen, wt%											
100°C Liquid Product											
Saturates, wt%		95.2	93.0	94.1	90.2	88.4	90.7	91.3	91.5	89.7	89.7
Aromatics, wt%		4.8	6.7	6.5	9.4	12.4	7.9	7.5	6.1	7.7	7.7
Recovery, wt%		100	99.1	99.2	99.6	97.2	97.6	98.0	97.0	97.4	97.4

03 DEC 1997

12

TABLE 2 (CONT'D)
OPERATING CONDITIONS, YIELDS AND PRODUCT INSPECTIONS, MCSU

Unit Run #	MCS 116.7									
Catalyst Description	NiO., MnO., MoS, (from oxide)									
Unit Conditions										
Feed	600N 82 VI Raffinate									
Feed Rate (LHSV)	1.0									
Gas Rate, SCF/h (m ³ gas/m ³ feed)	6200 (1102.9)	500 (88.9)	500 (88.9)	500 (88.9)	500 (88.9)	500 (88.9)	500 (88.9)	500 (88.9)	500 (88.9)	6200 (1102.9)
Reactor Temp., °C	350	340	340	340	340	340	340	340	340	350
Unit Pressure, MPa	8.5	4.1	4.1	4.1	4.1	4.1	4.1	4.1	4.1	8.5
Catalyst Hours	65-69	75-76	79-83	79-83	86-90	86-90	93-97	93-97	93-97	100-103
Total Liquid Product										
TLF Yield, wt%	97.7	97.7	99.4	99.4	98.4	98.4	99.0	99.0	99.0	97.5
360°C Yield on Feed, wt%	87.0	94.6	97.2	97.2	94.5	94.5	95.2	95.2	95.2	90.8
Refractive Index @ 75°C	1.4600	1.4717	1.4712	1.4712	1.4707	1.4707	1.4690	1.4690	1.4690	1.4632
Hydrogen, wt% by Norm										
360°C Liquid Product										
Visc @ 40°C, cSt (or 10 ⁶ m ² /s)	70.67	110.9	109.76	109.76	103.54	103.54	98.15	98.15	98.15	86.77
@ 100°C, cSt (or 10 ⁶ m ² /s)	9.21	11.95	11.17	11.17	11.24	11.24	10.98	10.98	10.98	10.29
Viscosity Index	106	88	85	85	94	94	96	96	96	100
Dohrman Sulphur, wppm	11	-	-	-	-	-	-	-	-	-
X-Ray Sulphur, wt%	<2	0.167	0.131	0.131	0.089	0.089	0.029	0.029	0.029	-
Total Nitrogen, wppm		88	80	80						
Basic Nitrogen, wppm										
ILPC Separation										
Saturates, wt%	89.8	62.9	62.0	62.0	63.6	63.6	68.0	68.0	68.0	83.7
Aromatics, wt%	9.8	36.2	35.5	35.5	36.3	36.3	31.5	31.5	31.5	15.5
Recovery, wt%	99.5	99.0	97.5	97.5	99.9	99.9	99.5	99.5	99.5	99.3

TABLE 4
OPERATING CONDITIONS, YIELDS AND PRODUCT INSPECTIONS, MCSU

Unit Run #	MCS 113-3			
Catalyst Description	Ni/MS (from sulphide)			
Unit Conditions				
Feed	600N 82 VI Raffinate			
Feed Rate (LHSV)	1.0			
Gas Rate, SCF/a (m ³ gas/m ³ feed)	6200 (1102.9)			
Reactor Temp., °C	340	375	375	380
Unit Pressure, MPa	0.5			
Catalyst Hours	8-12	15-19	22-26	29-33
Total Liquid Product	36-40			
TLP Yield, wt%	96.1	97.8	96.2	94.9
360°C Yield on Feed, wt%	94.1	87.4	84.6	82.8
Refractive Index @ 75°C	1.4744	1.4662	1.4657	1.4658
Hydrogen, wt% by NMR				
360°C Liquid Product				
Visc @ 40°C, cSt (or 10 ⁶ m ² /s)	124.2	85.17	81.5	79.75
@ 100°C, cSt (or 10 ⁶ m ² /s)	12.45	10.21	9.97	9.86
Viscosity Index	90	100	102	102
Dohrman Sulphur, wppm				
X-Ray Sulphur, wt%	0.677	0.091	0.071	0.089
Total Nitrogen, wppm	128			
HPLC Separation				
Saturates, wt%	58.9	73.6	74.7	72.4
Aromatics, wt%	37.6	25.6	21.9	25.9
Recovery, wt%	96.5	99.2	98.6	98.4
				80.2
				19.0
				99.2

TABLE 2
OPERATING CONDITIONS, YIELD AND PRODUCT INJECTIONS, WASH

Unit Run #

MCH 105.8

Catalyst Description

CofMo (from oxide)

Unit Conditions

600N 82 v1 Raffinate

Feed

Feed Rate (LHSV)

1.0

Gas Rate, SCF/h (m³gas/m³ feed)

6200 (1102.9)

Reactor Temp., °C

340

340

340

350

350

350

Unit Pressure, MPa

8.5

Catalyst Hours

7 11

14 18

21 25

28 32

35 39

41 44

54 57

64 61

Total Liquid Product

TLP Yield, wt%

95.7

90.8

96.6

98.3

98.3

98.4

98.8

98.8

360°C Yield on Feed, wt%

83.4

94.8

94.3

94.0

94.0

95.2

95.4

95.4

Refractive Index @ 75°C

1.4622

1.4680

1.4695

1.4672

1.4672

1.4672

1.4679

1.4679

Hydrogen, wt% by NMR

360°C Liquid Product

Visc @ 40°C, cSt (or 10⁶ m²/s)

60.74

102.4

107.59

96.05

96.05

96.31

98.34

86.31

@ 100°C, cSt (or 10⁶ m²/s)

9.08

11.28

11.50

10.81

10.81

10.82

10.95

10.21

Viscosity Index

107

96

93

95

95

95

95

99

Dohman Sulphur, wppm

50

440

570

160

160

290

185

140

X-Ray Sulphur, wt%

Total Nitrogen, wppm

Basic Nitrogen, wppm

HELC Separation

Saturates, wt%

80.0

62.7

64.6

64.7

67.7

67.0

66.5

69.3

Aromatics, wt%

18.9

32.2

34.0

28.4

30.5

30.2

28.3

29.6

Recovery, wt%

98.9

94.9

98.6

93.0

98.2

97.1

94.8

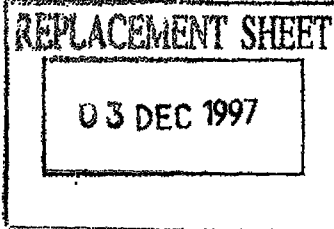
98.9

TABLE 6
OPERATING CONDITIONS, YIELD AND PRODUCT INSTRUCTIONS, P220

Unit Run #	MCS-111-9									
Catalyst Description	Ni ₂ S Mo ₂ S Mo (from sulphide)									
Unit Conditions										
Feed	600N 82 VI Raffinate									
Feed Rate (LHSV)	1.0									
Gas Rate, SCFH (m ³ gas/m ³ feed)	6200 (1102.9)									
Reactor Temp., °C	330	340	340	340	360	360	360	375		
Unit Pressure, MPa	0.5									
Catalyst Hours	5.8	12.15	19.22	26.29	33.36	40.43				
Total Liquid Product										
TiP Yield, wt%	96.8	97.8	98.6	96.6	98.0	96.5				
160°C Yield on Feed, wt%	95.7	97.3	98.1	95.1	97.4	95.3				
Refractive Index @ 75°C	1.4720	1.4717	1.4715	1.4698	1.4695	1.4677				
Hydrogen, wt% by NMR										
260°C Liquid Product										
Visc @ 40°C, cSt (or 10 ⁶ m ² /s)	115.72	112.94	110.34	103.17	101.41	89.85				
@ 100°C, cSt (or 10 ⁶ m ² /s)	11.98	11.77	11.70	11.32	11.09	10.43				
Viscosity Index	91	91	93	95	94	98				
Dohrmann Sulphur, wppm	-	0.16	0.15	640	530	260				
X-Ray Sulphur, wt%	0.20	150	150	87	61	11				
Total Nitrogen, wppm	170									
Basic Nitrogen, wppm										
HPLC Separation										
Saturates, wt%	62.5	62.8	60.3	66.7	66.9	67.1				
Aromatics, wt%	36.6	36.0	34.9	32.6	30.6	27.7				
Recovery, wt%	99.1	98.7	95.2	99.5	97.5	94.8				

Table 7
OPERATING CONDITIONS, YIELDS AND PRODUCT INSPECTIONS, MCSII

Unit Run #	MCS III 10									
Catalyst Description	N ₂ , Cr ₂ O ₃ , Mo (from sulphide)									
Unit Conditions										
Feed	600N 82 VI Raffinate									
Feed Rate (L/min)	1.0									
Gas Rate, SCF/S (m ³ gas/m ³ feed)	6200 (1102.9)									
Reactor Temp., °C	325	340	340	350	350	350	350	355	345	
Unit Pressure, MPa	8.5									
Catalyst Hours	11-14	18-21	25-28	32-35	39-42	46-49	53-56	60-63		
Total Liquid Product										
TLP Yield, wt%	97.0	98.5	96.6	93.0	93.2	93.3	93.5	94.3		
360°C Yield on Feed, wt%	91.1	97.6	93.2	82.9	79.1	79.2	79.3	83.7		
Refractive Index @ 75°C	1.4707	1.4079	1.4661	1.4604	1.4555	1.4567	1.4561	1.4594		
Hydrogen, wt% by H ₂ O										
360°C Liquid Product										
Visc @ 40°C, cSt (or 10 ⁶ m ² /s)	106.07	88.89	85.42	67.50	56.70	57.60	56.68	64.85		
@ 100°C, cSt (or 10 ⁶ m ² /s)	11.38	10.30	10.11	9.05	8.13	8.20	8.16	8.81		
Viscosity Index	92.8	96.6	98.2	109.1	111.9	111.4	112.9	109.3		
Schram Sulphur, wppm	58	440	570	160	160	250	185	140		
X-Ray Sulphur, wt%	0.25							47		
Total Nitrogen, wppm	44	13	5							
Basic Nitrogen, wppm										
HPLC Separation										
Saturates, wt%	65.8	63.8	69.0	79.8	88.7	84.5	85.8	83.0		
Aromatics, wt%	35.2	35.7	31.1	18.8	10.1	12.3	10.1	16.0		
Recovery, wt%	101.0	99.4	100.1	98.6	98.0	96.6	95.9	99.0		

**Claims**

1. A method for increasing the Viscosity Index of lubricating and specialty oils by contacting with a catalyst, said method comprising contacting a lube or specialty oil stock under hydrotreating conditions with a specific bulk metal sulfide hydrotreating catalyst selected from the group consisting of bulk Cr/Ni/Mo sulfide catalyst, bulk Ni/Mo/Mn sulfide catalyst and mixtures thereof, wherein the Cr/Ni/Mo sulfide catalyst is prepared by decomposition in a non-oxidizing sulfur-containing atmosphere of a $[\text{Ni}(\text{L})_3](\text{Cr}(\text{L})_3)(\text{MoS}_4)_{2-2.5}$ precursor complexes and the Ni/Mo/Mn sulfide catalyst is prepared by decomposition in inert atmosphere followed by sulfiding using $\text{H}_2\text{S}/\text{H}_2$ of a $[\text{Ni}(\text{L})_3](\text{Mn}(\text{L})_3)(\text{MoO}_4)_2$ precursor complex where L is one or more neutral, nitrogen containing ligand, at least one of which is a chelating polydentate ligand.
2. The method of claim 1 wherein hydrotreating conditions comprise temperature in the range of from 200 to 450°C, pressure in the range of from 300 to 3000 psi (20.69 to 206.89 bar), space velocity in the range of from about 0.1 to 5.0 v/v/hr and a treat gas rate in the range of from about 200 to 20,000 SCF H_2 /bbl (35.58 to 3557.86 $\text{m}^3/\text{H}_2/\text{m}^3$ feed).
3. The method of claim 1 or claim 2 wherein the $[\text{Ni}(\text{L})_3](\text{Mn}(\text{L})_3)(\text{MoO}_4)_2$ precursor is converted into the bulk catalyst by decomposing the precursor in the presence of an inert atmosphere at 300 to 400°C and sulfiding using $\text{H}_2\text{S}/\text{H}_2$ at 200 to 600°C and the $[\text{Ni}(\text{L})_3](\text{Cr}(\text{L})_3)(\text{MoS}_4)_{2-2.5}$ precursor is converted into the bulk catalyst by decomposing the precursor in a sulfur containing atmosphere at a temperature of at least 150°C.
4. The method of claims 1 to 3 wherein the ligand is ethylene diamine.
5. The method of claims 1 to 4 wherein the bulk metal sulfide catalyst is used in the form of pellets made by compacting bulk metal sulfide under pressure.

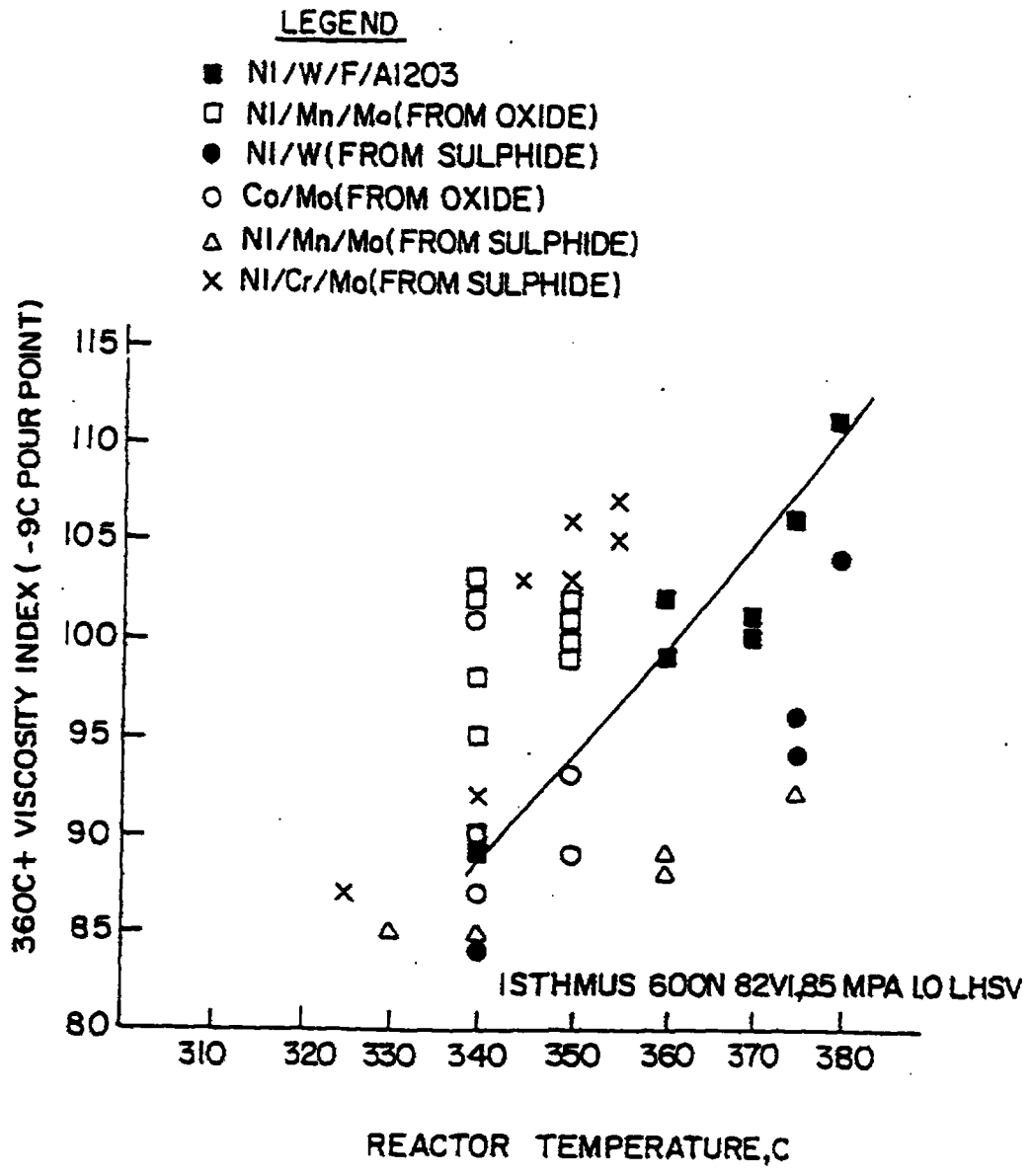
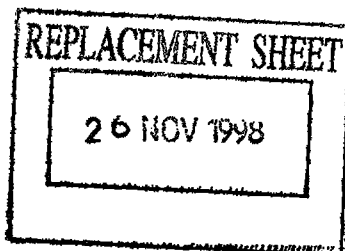


FIG. 1

INCREASING VI OR LUBE OIL HYDROTREATINGABSTRACT

Lube oils of increased VI are prepared by hydrotreating using bulk Ni/Mn/No or Cr/Ni/Mo sulfide catalysts prepared from ligated, eg., ethylene diamine, metal complexes. In particular the Ni/Mn/Mo sulfide catalyst is prepared from the oxide precursor with subsequent sulfiding while the Cr/Ni/Mo sulfide catalyst is prepared from the sulfide precursor in the presence of H_2S/H_2 .

FIGURE 1.