

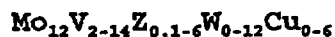
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PROCESS FOR THE MANUFACTURE OF ACRYLIC ACID FROM ACROLEIN
BY A REDOX REACTION AND USE OF A SOLID MIXED OXIDES
COMPOSITION AS REDOX SYSTEM IN THE SAID REACTION ✓

5 The present invention relates to the manufacture
of acrylic acid from acrolein by oxidation according to
a redox reaction. The invention also relates to the use
of a solid mixed oxides composition as redox system for
the said reaction.

10 Industrial production of acrylic acid is at
present carried out by vapour phase catalytic oxidation
of acrolein. All attempts to improve this process have
hitherto related to the development of catalysts giving
the highest possible conversion of acrolein and the
highest possible selectivity for the desired acrylic
15 acid.

Thus, French Patent No. 2 222 349 describes a
catalyst for the preparation of acrylic acid by vapour
phase catalytic oxidation of acrolein using a gas con-
taining molecular oxygen, this catalyst including a
20 catalytic oxide on an inert porous support, this cataly-
tic oxide having the following metal composition:



Z being at least one of Be, Mg, Ca, Ba and Sr, and at
least one of W and Cu being always present.

25 This catalyst can be prepared by adding a support
(powdered material or α -alumina, silicon carbide or
similar beads) to an aqueous solution in which compounds
of the various catalyst elements are dissolved, evaporat-
ing the aqueous solution to dryness to deposit the
30 catalyst elements on the support and calcining the dried
product between 300 and 800°C.

The Applicant Company has now discovered that
acrylic acid can be manufactured by gas phase oxidation
of acrolein in the absence of molecular oxygen, by
35 passing a gaseous mixture of acrolein and water vapour
and, if appropriate, of an inert gas over a specific
solid mixed oxides composition, which acts as a redox
system and supplies the oxygen necessary for the reac-
tion.

The advantages of this new process are the following:

- 5 - the disadvantage of an oxidation with molecular oxygen is the overoxidation promoting the degradation of the products formed; according to the present invention, since the operation is carried out in the absence of molecular oxygen, the formation of CO_x (carbon monoxide and carbon dioxide), degradation products, is reduced, and this allows the selectivity for acrylic acid to be substantially increased;
- 10 - the selectivity for acrylic acid remains good when the degree of reduction of the solid composition increases;
- 15 - once it has undergone a reduction and the progressive loss of its activity, the solid composition can be easily regenerated by heating in the presence of oxygen or of a gas containing oxygen after a certain period of use; after the regeneration the solid recovers its initial activity and can be employed in a new reaction cycle;
- 20 - the separation of the stages of reduction of the solid composition and of its regeneration makes it possible:
- 25 - to increase the selectivity for acrylic acid; and
- to increase the partial pressure of acrolein, such a partial pressure of acrolein feed being no longer restricted by the existence of an explosive region of the acrolein + oxygen mixture.
- 30

The subject of the present invention is therefore firstly the use of a solid mixed oxides composition of formula (I):



35 in which:

- a is between 2 and 14, limits included,
- b is between 0.1 and 6, limits included,
- c is between 0 and 12, limits included,
- d is between 0 and 6, limits included,

- e is between 0 and 15, limits included, and
- x is the quantity of oxygen bonded to the other
elements and depends on their oxidation states,
in the manufacture of acrylic acid by oxidation of
5 acrolein, the said solid composition reacting with
acrolein according to the redox reaction (1):



The oxides of the various metals forming part of
the composition of the mixed oxide of formula (I) can be
10 employed as raw materials in the preparation of this
composition, but the raw materials are not restricted to
the oxides; as other raw materials there may be
mentioned:

- in the case of molybdenum: ammonium molybdate and
15 molybdic acid,
- in the case of vanadium: ammonium metavanadate,
- in the case of strontium: strontium hydroxide,
carbonate or nitrate,
- in the case of tungsten: ammonium tungstate and
20 tungstic acid,
- in the case of copper: copper hydroxide, carbonate
or nitrate,

and, in general, any compounds capable of forming an
oxide on calcination, namely metal salts of organic
25 acids, metal salts of inorganic acids, complex metal
compounds and organic metal compounds, and the like.

The source of silicon generally consists of
colloidal silica.

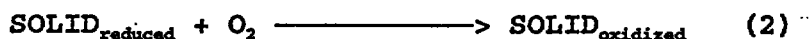
In accordance with specific embodiments, solid
30 compositions of formula (I) can be prepared by mixing,
with stirring, aqueous solutions of ammonium paratung-
state, ammonium metavanadate, ammonium molybdate and
copper nitrate and strontium nitrate, adding colloidal
silica if appropriate, and then drying them and calcining
35 them in air between 300 and 600°C, preferably between 350
and 500°C.

Another subject of the present invention is a
process for the manufacture of acrylic acid from
acrolein, according to which process a gaseous mixture of

acrolein and of water vapour and, if appropriate, of an inert gas such as nitrogen is passed over a solid composition of formula (I) defined above, to conduct the redox reaction (1) as indicated above, by operating at a temperature of 200 to 500°C, especially from 250 to 450°C, at a pressure of 1.01×10^4 to 1.01×10^6 Pa (0.1 to 10 atmospheres), especially from 5.05×10^4 to 5.05×10^5 Pa (0.5 - 5 atmospheres) and with a residence time of 0.01 second to 90 seconds, especially from 0.1 second to 30 seconds, in the absence of molecular oxygen.

The acrolein/water vapour volume ratio in the gaseous phase is not critical and may vary within wide limits.

During the redox reaction (1), the solid composition undergoes a reduction and a gradual loss of its activity. This is why, once the solid composition has changed to the reduced state, regeneration of the said solid composition is conducted according to the reaction (2):



by heating in the presence of an excess of oxygen or of an oxygen-containing gas at a temperature of 250 to 500°C, for the time needed for the reoxidation of the solid composition.

After the regeneration, which can be carried out in temperature and pressure conditions which are identical with or different from those of the redox reaction, the solid composition recovers an initial activity and can be employed in a new reaction cycle.

The redox reaction (1) and the regeneration may be conducted in a two-stage device, namely a reactor and a regenerator which operate simultaneously and in which two charges of solid composition alternate periodically; the redox reaction (1) and the regeneration may also be conducted in the same reactor by alternating the reaction and regeneration periods.

The preparation of acrylic acid according to the invention takes place according to a stoichiometric and

noncatalytic reaction.

The following examples illustrate the present invention without, however, restricting its scope. The conversions, selectivities and yields are defined as follows:

$$\begin{aligned} \text{Conversion (\%)} &= \frac{\text{Number of moles of acrolein which have reacted}}{\text{Number of moles of acrolein introduced}} \times 100 \\ \text{Selectivity (\%)} &= \frac{\text{Number of moles of acrylic acid formed}}{\text{Number of moles of acrolein which have reacted}} \times 100 \\ \text{Selectivity (\%)} &= \frac{\text{Number of moles of acetic acid formed}}{\text{Number of moles of acrolein which have reacted}} \times 100 \end{aligned}$$

EXAMPLE 1(a):

Preparation of a solid of formula

$\text{Mo}_{12}\text{V}_{4.8}\text{Sr}_{0.5}\text{W}_{2.4}\text{Cu}_{2.2}\text{O}_x$, x being the quantity of oxygen bonded to the other elements and depending on their oxidation states.

3.6 g of ammonium paratungstate, 3.0 g of ammonium metavanadate and 12.4 g of ammonium heptamolybdate are introduced into 100 g of water and heated to 100°C. 3.0 g of copper nitrate and 0.62 g of strontium nitrate are introduced into 5 g of water and heated to 100°C. The second solution is added to the first and the resulting solution is then evaporated to dryness and then calcined for 4 hours at 400°C.

EXAMPLE 1(b) (comparative):

Preparation of acrylic acid from acrolein

50 mg of a solid prepared according to Example 1(a) are charged into a tubular reactor and the reactor is then flushed with a continuous flow of 20 ml/minute of air and heated to 300°C. An injection of an aqueous solution of 12 % by weight of acrolein, containing 1.1×10^{-6} mol of acrolein, is introduced onto the solid. 97 % of the acrolein is converted with selectivities for acrylic acid and for acetic acid of 56 % and 4 % respectively.

Example 2:

Preparation of acrylic acid from acrolein by a redox reaction.

After the treatment of Example 1(b) the reactor

is flushed with a continuous flow of 17 ml/minute of nitrogen and heated to 300°C. An injection of an aqueous solution of 12 % by weight of acrolein, containing 1.1×10^{-6} mol of acrolein, is introduced onto the solid.

5 99.1 % of the acrolein is converted with selectivities for acrylic acid and for acetic acid of 65 % and 6.5 % respectively.

EXAMPLE 3:

After having conducted the reaction of Example 2

10 the same solid is again subjected to eleven successive injections of acrolein in the same test conditions as Example 2. The performance values obtained are reported in Table 1. (Injection No. 1 corresponds to Example 2).

Table 1

15	Injection No.	Acrolein conversion (%)	Selectivity for acrylic acid (%)	Selectivity for acetic acid (%)
	1	99.1	65	6.5
	2	98.2	79	6.5
	3	96.5	81	5.9
20	4	95.7	86	5.3
	5	92.8	89	3.9
	6	89.8	90	3.6
	7	85.5	90	3.3
	8	83.4	92	3.1
25	9	83.9	91	3.4
	10	84.0	89	3.4
	11	73.6	90	3.6
	12	70.8	91	3.6

EXAMPLE 4:

30 After the reducing treatment of Example 3 the solid is regenerated for 2 hours at 300°C under a flow of air and then placed back under a flow of nitrogen. Twenty-four new successive injections of aqueous solution of 12 % by weight of acrolein, containing 1.1×10^{-6} mol

35 of acrolein, are introduced onto the solid.

The performance values obtained are reported in

Table 2.

Table 2

	Injection No.	Acrolein conversion (%)	Selectivity for acrylic acid (%)	Selectivity for acetic acid (%)
5	1	99.5	61	5.3
	2	90.8	79	3.8
	3	97.7	84	5.3
	4	94.2	90	4.0
	5	93.0	88	3.8
10	6	91.4	89	3.5
	7	83.2	91	2.9
	8	82.3	90	3.0
	9	82.4	90	3.3
	10	78.6	90	3.4
15	11	71.4	91	3.6
	12	68.5	89	3.8
	13	67.8	90	3.7
	14	62.5	90	3.7
	15	61.5	89	3.7
20	16	58.1	90	3.8
	17	56.2	92	3.8
	18	54.7	91	3.7
	19	54.1	91	3.7
	20	49.1	92	3.9
25	21	47.5	92	3.9
	22	49.1	92	3.7
	23	47.8	93	3.7
	24	42.6	94	3.9



CLAIMS

1. Use of a solid mixed oxides composition of formula (I):



5 in which:

- a is between 2 and 14, limits included,
- b is between 0.1 and 6, limits included,
- c is between 0 and 12, limits included,
- d is between 0 and 6, limits included,
- 10 - e is between 0 and 15, limits included, and
- x is the quantity of oxygen bonded to the other elements and depends on their oxidation state,

in the manufacture of acrylic acid by oxidation of acrolein, the said solid composition reacting with the acrolein according to the redox reaction (1):



2. Process for the manufacture of acrylic acid from acrolein, characterized in that a gaseous mixture of acrolein and of water vapour and, if appropriate, of an inert gas is passed over a solid composition of formula (I) as defined in Claim 1, to conduct the redox reaction (1) as indicated in Claim 1, by operating at a temperature of 200 to 500°C, at a pressure of 1.01×10^4 to 1.01×10^6 Pa (0.1 to 10 atmospheres) and with a residence time of 0.01 second to 90 seconds, in the absence of molecular oxygen.

3. Process according to Claim 2, characterized in that the redox reaction (1) is conducted at a temperature of 250 to 450°C.

30 4. Process according to either of Claims 2 and 3, characterized in that the redox reaction (1) is conducted at a pressure of 5.05×10^4 - 5.05×10^5 Pa (0.5 - 5 atmospheres).

5. Process according to one of Claims 2 to 4, characterized in that the redox reaction (1) is conducted with a residence time of 0.1 second to 30 seconds.

6. Process according to one of Claims 2 to 5, characterized in that, once the solid composition has



changed to the reduced state, regeneration of the said solid composition is conducted according to the reaction (2):



5 by heating in the presence of an excess of oxygen or of an oxygen-containing gas at a temperature of 250 to 500°C, for the time needed for the reoxidation of the solid composition.

7. Process according to Claim 6, characterized in
10 that the redox reaction (1) and the regeneration are conducted in a two-stage device, namely a reactor and a regenerator which operate simultaneously and in which two charges of solid composition alternate periodically.

8. Process according to Claim 6, characterized in
15 that the redox reaction (1) and the regeneration are conducted in the same reactor by alternating the reaction and regeneration periods.