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DESCRIPTION

"Method for preparing a glucosamine compound, and compound thus obtained"

5 The present invention relates to a method for preparing a glucosamine compound, and to a compound thus obtained.

As is known, glucosamine (2-amino-2-deoxyglucose; chitosamine; GluNH₂) is a component of mucoproteins and mucopolysaccharides. It is usually obtained from chitin and is isolated as the hydrochloride.

10 Glucosamine sulphate is a drug that is well known and widely used in the treatment of rheumatic fever, arthritic and arthrosic complaints, both acute and chronic, and in the treatment of pathological conditions originating from metabolic disorders of osteoarticular tissue. Its synthesis was described at the end of 1898 by Breuer (Chem. Ber. 31, 15 2197). However, its physicochemical properties are such that it is relatively difficult to handle and occasionally unstable. The main problems are derived from its high hygroscopicity.

US-A-4 642 340 teaches how to overcome this drawback by using a presumed mixed salt of formula A:

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$$[\text{GluNH}_3^+]_2 \cdot 2\text{Na}^+ \cdot \text{SO}_4^{2-} \cdot 2\text{Cl}^- \quad \text{A}$$

which is in the form of crystalline powder, melts at more than 300°C, is stable under ambient conditions and has pharmacological properties that are substantially identical to those of glucosamine sulphate.

The said document also describes the preparation of the said presumed mixed salt A, by:

- 25 (a) dissolving, with stirring, anhydrous sodium chloride in 5.5-7.5 parts by weight of distilled water for each part of sodium chloride at a temperature of from 50°C to 70°C;

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- (b) dissolving, with stirring, in the solution obtained in step (a), a stoichiometric amount of glucosamine sulphate at a temperature of from 35°C to 45°C;
- 5 (c) precipitating the presumed mixed salt **A** by adding a precipitating liquid that is water-miscible and in which the presumed mixed salt **A** has a solubility not greater than 0.1% (w/v), this step being performed with stirring at a temperature of from 40°C to 50°C;
- (d) completing the precipitation by reducing the temperature of the mixture; and
- 10 (e) recovering the presumed precipitated mixed salt **A**.

The precipitating liquid used in step (c) is acetone, ethanol, acetonitrile, tetrahydrofuran or dioxane.

A similar method is described in US-A-5 847 107. According to this method, the said presumed mixed salt **A** is obtained by:

- 15 (a) dissolving glucosamine hydrochloride and sodium sulphate, in stoichiometric amounts, in water; and
- (b) precipitating the presumed mixed salt **A** by adding a precipitating liquid that is water-miscible.

In this case also, the precipitating liquid used is acetone, ethanol, 20 acetonitrile, tetrahydrofuran or dioxane.

Another similar method is described in EP-A-0 214 642. According to this method, the said presumed mixed salt **A** and other similar products are obtained by:

- 25 (a) forming glucosamine sulphate from glucosamine and sulphuric acid in water;
- (b) forming the presumed mixed salt **A** by adding an approximately stoichiometric amount of an alkali metal halide or alkaline-earth metal halide; and
- 30 (c) precipitating the presumed mixed salt **A** by adding a water-soluble solvent.

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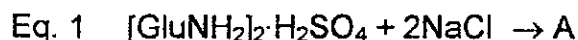
The precipitating solvents mentioned are ethanol, acetone and acetonitrile.

According to US-A-5 902 801, the product present on the market is not thought to be a true mixed salt A, but rather a simple stoichiometric mixture of glucosamine hydrochloride and sodium sulphate (column 1, lines 23-29). This document describes a compound that purports to be novel and to be the true mixed salt A. According to this document, the true mixed salt A is that obtained by (column 3, lines 5-19):

- (a) placing stoichiometric amounts of glucosamine hydrochloride and sodium sulphate in contact in a sufficient amount of water to have a solids concentration of from about 15 to 40% by weight (this step takes from 15 minutes to 2 hours and is carried out at a temperature of from about 20°C to about 40°C); and
- (b) removing the water by freeze-drying at a pressure below 800 milliTorr, preferably from 300 to 500 milliTorr, and at a temperature from -60°C to 0°C and preferably from -40°C to -5°C.

However, the water is not totally removed and the mixed salt A thus obtained contains from 3 to 5% by weight of water (column 2, lines 20-24).

The abovementioned known methods thus involve the use of starting materials in one of the exact stoichiometric ratios given below:



A person skilled in the art will readily appreciate that the abovementioned prior art methods have considerable drawbacks.

Specifically, in the methods of US-A-4 642 340, US-A-5 847 107 and EP-A-0 214 642, the precipitation requires the use of organic solvents and this obliges the application of specific safety standards on account of their flammability and explosiveness. In addition, a specific treatment

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of the wastes is necessary before discharging. This therefore involves factors that, in addition to having an impact on the safety of the industrial production, also involve increases in the production costs. Finally, a further drawback is given by the fact that the organic solvents cannot be completely removed from the presumed mixed salt A either during the drying operation.

As regards the freeze-drying process required in the method of US-A-5 902 801, in addition to being expensive, it does not allow the impurities present in the initial solution to be removed. The mixed salt A thus obtained actually retains all the abovementioned impurities. To overcome this drawback, the glucosamine hydrochloride and the sodium sulphate will have to be very pure and, therefore, very expensive. Finally, small amounts of water also tend to remain in the mixed salt A in the freeze-drying process.

It has now been found, surprisingly, that all the abovementioned drawbacks are overcome by working in water alone.

In a first aspect, the present invention thus comprises a method for preparing a compound comprising protonated glucosamine and Cl^- , Na^+ and SO_4^{2-} ions in the ratio indicated in the following formula:



the method being carried out by:

- (a) placing glucosamine hydrochloride and a stoichiometric excess of sodium sulphate in water,
- (b) heating the mixture obtained in the preceding step (a),
- (c) cooling the mixture, and
- (d) recovering the solid present in the said cooled mixture by filtration.

Preferably, in step (a), from 0.8 to 1 mol of sodium sulphate is used per mole of glucosamine hydrochloride. Even more preferably, 0.9 mol of sodium sulphate is used per mole of glucosamine hydrochloride.

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Advantageously, the amount of water used in step (a) ranges from 180 to 230 ml per mole of glucosamine hydrochloride. Preferably, the amount of water is about 200 ml per mole of glucosamine hydrochloride.

- 5 In step (b), the mixture is preferably heated, with stirring, to a temperature of from 35°C to 55°C and even more preferably to a temperature of from 40°C to 45°C.

The heating in step (b) is preferably continued for a period of from 15 minutes to 5 hours and even more preferably from 2.5 to 3.5 hours.

- 10 In step (c), the mixture is preferably cooled to a temperature of from 0°C to 33°C and even more preferably to about 30°C.

The cooling in step (c) is preferably continued for a period of from 1 to 5 hours and even more preferably from 2.5 to 3.5 hours.

- 15 The solid collected by filtration in step (d) is dried in a vacuum oven or under a stream of air according to standard techniques.

Another advantage of the present invention is that, by always working only in water, the mother liquors obtained from the filtration in step (d) are readily recycled. In this way, the yield for the method is very close to the theoretical yield.

- 20 Since the method of the present invention is based on a filtration of saturated solution, a person skilled in the art will readily appreciate that the concentrations, temperatures and times previously indicated are not limiting. For example, the stoichiometric excess of the sodium sulphate can vary within a wide range depending on the amount of water used and the filtration temperature.

25 The glucosamine compound according to the present invention is of very high purity, does not contain excess sulphates, is virtually anhydrous and can be dried by simple removal of the final traces of moisture.

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In addition, as shown in Fig. 1, it has a specific and characteristic X-ray diffraction diagram.

The X-ray diffraction measurements were carried out using a Selfert XRD3000 powder diffractometer with Cu- α radiations ($\lambda = 1.5406\text{\AA}$).

- 5 The diffractometer was equipped with a secondary graphite monochromator (between the sample and the detector), the incident beam was limited by entry slits of from 3 to 2 mm, and the diffracted beam was limited by detector slits of from 0.3 to 0.2 mm.

- 10 The diffraction diagram was acquired with "scan θ -2 θ " in the range $3 \leq 2\theta \leq 60^\circ$, with a sweep size of 0.04° and a count time of 4 sec/sweep.

- 15 The examples that follow serve to illustrate the invention without, however, limiting it.

EXAMPLE 1

Preparation of the glucosamine compound according to the present Invention

- 20 116.28 g (0.54 mol) of glucosamine hydrochloride and 69.47 g (0.49 mol) of anhydrous sodium sulphate were added to 109 ml of water with stirring. The mixture was heated to 42°C - 45°C and then maintained at this temperature for 3 hours with stirring. The mixture was cooled to 30°C over 30 minutes and maintained at this temperature for a further 3
25 hours with stirring.

The precipitated solid was collected by filtration and dried at 50°C under reduced pressure, to constant weight.

114.7 g (74.1% of the theoretical yield) of the glucosamine compound according to the present invention were thus obtained in the

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form of a white crystalline powder which, on analysis, gave the results shown in Table I.

The mother liquors filtered off (166 g) consisted of water (100.4 ml), glucosamine hydrochloride (27.74 g; 0.13 mol) and sodium sulphate
5 (37.90 g; 0.27 mol).

Table I

Elemental analysis	Theoretical	Result
	C = 25.14%	C = 25.10%
	H = 4.92%	H = 4.85%
	N = 4.88%	N = 4.70%
Glucosamine		100.2%
Cl ⁻		99.8%
SO ₄ ²⁻		100.5%
Optical rotation		[α] _D ²⁰ = +52.6

The glucosamine content was determined by potentiometric titration with NaOH (0.1 N) and by determining the equivalent point with a glass electrode.

10 The content of Cl⁻ was determined by potentiometric titration with silver nitrate (0.05 N) and the equivalent point was determined with an Ag/AgCl electrode.

The SO₄²⁻ content was determined by ionic chromatography using a DIONEXTM AG-9 analytical column with conductimetric detection and
15 external standardization.

The optical rotation was determined according to Ph. Eur. IV Ed. Par. (2.2.7) 2002, using a 1 dm polarimeter tube after an equilibration time of 3 hours at room temperature (c = 10% in water).

EXAMPLE 2

20 First recycling of the mother liquors

Water (8.45 ml), glucosamine hydrochloride (88.54 g; 0.41 mol) and sodium sulphate (31.59 g; 0.22 mol) were added to the mother liquors

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obtained from Example 1, and the process was performed as described in Example 1.

After drying, 117.7 g (76.1% of the theoretical yield) of the glucosamine compound according to the present invention were
5 obtained in the form of a white crystalline powder which, on analysis, gave results similar to those of Table I.

The combined yield for the two preparations, calculated on the basis of the total number of moles of glucosamine used in the two preparations described in Examples 1 and 2 was 85.4%.

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EXAMPLE 3

Second recycling of the mother liquors

Water (8.45 ml), glucosamine hydrochloride (88.54 g; 0.41 mol) and sodium sulphate (31.59 g; 0.22 mol) were added to the mother liquors obtained from Example 2, and the process was performed as described
15 in Example 2.

After drying, 117.2 g (75.8% of the theoretical yield) of the glucosamine compound according to the present invention were obtained in the form of a white crystalline powder which, on analysis, gave results similar to those of Table I.

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The combined yield for the three preparations, calculated on the basis of the total number of moles of glucosamine used in the three preparations described in Examples 1, 2 and 3, was 89.6%.

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CLAIMS

1. A method for preparing a compound comprising protonated glucosamine and Cl^- , Na^+ and SO_4^{2-} ions in the ratio indicated in the following formula:
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$$[\text{GluNH}_3^+]_2 \cdot 2\text{Na}^+ \cdot \text{SO}_4^{2-} \cdot 2\text{Cl}^- \quad \text{A}$$

the method being carried out by:

 - (a) placing glucosamine hydrochloride and a stoichiometric excess of sodium sulphate in water,
 - (b) heating the mixture obtained in the preceding step (a),
 - 10 (c) cooling the mixture, and
 - (d) recovering the solid present in the said cooled mixture by filtration.
2. A method according to Claim 1, characterized in that, in step (a), from 0.8 to 1 mol of sodium sulphate is used per mole of glucosamine hydrochloride.
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3. A method according to Claim 2, characterized in that, in step (a), 0.9 mol of sodium sulphate is used per mole of glucosamine hydrochloride.
4. A method according to any one of the preceding Claims 1 to 3, characterized in that the amount of water used in step (a) ranges from 180 to 230 ml per mole of glucosamine hydrochloride.
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5. A method according to Claim 4, characterized in that the amount of water is about 200 ml per mole of glucosamine hydrochloride.
6. A method according to any one of Claims 1 to 5, characterized in that, in step (b), the mixture is heated to a temperature of from 35°C to 55°C.
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7. A method according to Claim 6, characterized in that the mixture is heated to a temperature of from 40°C to 45°C.

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8. A method according to any one of the preceding Claims 1 to 7, characterized in that, in step (b), the heating is continued for a period of from 15 minutes to 5 hours.
- 5 9. A method according to Claim 8, characterized in that the heating is continued for a period of from 2.5 to 3.5 hours.
- 10 10. A method according to any one of the preceding Claims 1 to 9, characterized in that, in step (c), the mixture is cooled to a temperature of from 0°C to 33°C.
- 10 11. A method according to Claim 10, characterized in that the mixture is cooled to about 30°C.
12. A method according to any one of the preceding Claims 1-11, characterized in that, in step (c), the cooling is continued for a period of from 1 to 5 hours.
- 15 13. A method according to Claim 12, characterized in that the cooling is continued for a period of from 2.5 to 3.5 hours.

AMENDED SHEET