

**OXIDATION PROCESS FOR THE PRODUCTION OF ALKENES AND  
CARBOXYLIC ACIDS**

The present invention relates to a process for the oxidation of a C<sub>2</sub> to C<sub>4</sub> alkane to produce the corresponding alkene and carboxylic acid and to integrated processes in which the alkene and carboxylic acid are further used as reactants.

Carboxylic acids are useful feedstocks for the production of alkenyl  
5 carboxylates. Thus, for example, acetic acid is used to manufacture vinyl acetate which is generally prepared commercially by contacting ethylene and acetic acid with molecular oxygen in the presence of a catalyst active for the production of vinyl acetate. Suitably, the catalyst may comprise palladium, an alkali metal acetate promoter and an optional co-promoter (for example, gold or cadmium) on a catalyst support.

10 Acetic acid may be produced by the catalytic oxidation of ethylene and/or ethane.

Integrated processes for producing acetic acid and/or vinyl acetate are known in the art. EP-A-0 877 727 discloses an integrated process for the production of acetic acid and/or vinyl acetate in any pre-determined and variable proportions from a gaseous feedstock comprising ethylene and/or ethane. The integrated process comprises a first  
15 step wherein ethylene and/or ethane is catalytically oxidised in a first reaction zone to produce a first product stream comprising acetic acid, water and ethylene and optionally ethane, carbon monoxide, carbon dioxide and/or nitrogen. The acetic acid and ethylene produced in this first reaction zone are then contacted in a second reaction zone with a molecular oxygen-containing gas in the presence of a catalyst to produce a second  
20 product stream comprising vinyl acetate, water, acetic acid and optionally ethylene. No mention is made of any control of the production ratio of ethylene to acetic acid from the catalytic oxidation of ethane and/or ethylene.

Research Disclosure 2244 of 1992(June) No. 338 describes a process for the oxidation of ethane and/or ethylene to produce acetic acid in which the by-product carbon monoxide is oxidised to carbon dioxide. According to this document, the acetic acid, unreacted ethane (if present) and ethylene is passed with or without carbon dioxide and water removal, to a reactor having a suitable catalyst for the production of ethyl acetate or, with the addition of oxygen, for the production of vinyl acetate. This document is silent on the control of the ratio of ethylene to acetic acid produced in the oxidation step.

In the manufacture of vinyl acetate from ethylene and acetic acid, the molar ratio of the fresh feed ethylene to acetic acid is desirably unity or approximately unity. Thus, in an integrated process in which ethane is oxidised in an oxidation reaction zone to produce ethylene and acetic acid for use in a second reaction zone for the production of vinyl acetate, to maximise overall integrated process efficiency and also vinyl acetate output, the molar ratio of ethylene to acetic acid produced in the oxidation reaction zone, is desirably unity or approximately unity depending upon the selectivity/yield in the second reaction zone.

The effect of water on the formation of acetic acid during the oxidation of ethane to acetic acid is described in US 4250346, but this does not disclose the effect of ethylene on the ratio of products formed.

Thus, there remains a need for a process for the oxidation of a  $C_2$  to  $C_4$  alkane to produce the corresponding alkene and carboxylic acid in which the molar ratio of alkene to carboxylic acid produced is adjusted or maintained at a pre-determined value.

Typically, in the oxidation of an alkane such as ethane, carbon monoxide and/or carbon dioxide (the carbon oxides) are formed as by-products. The formation of high levels (typically greater than 15 mol%) of these carbon oxides is undesirable as it generally results in increased capital and production costs. A low selectivity to carbon oxides reduces the need for costly reaction and heat removal systems and removal systems for product purification, reduces the cost of carbon oxide removal systems, lowers operating costs and results in increased yields to the desirable carboxylic acid and alkene products.

In view of the above, it would be desirable, if the oxidation of a  $C_2$  to  $C_4$  alkane to produce the corresponding alkene and carboxylic acid and in which the molar ratio of alkene to carboxylic acid produced is adjusted or maintained at a pre-determined value

were to achieve a low selectivity to the carbon oxides, carbon monoxide and/or carbon dioxide.

Accordingly, the present invention provides a process for the oxidation of a  $C_2$  to  $C_4$  alkane to produce the corresponding alkene and carboxylic acid which process comprises contacting in an oxidation reaction zone, said alkane, molecular oxygen-containing gas, and the corresponding alkene and optionally, water, in the presence of at least one catalyst active for the oxidation of the alkane to the corresponding alkene and carboxylic acid, to produce a product stream comprising alkene, carboxylic acid and water, wherein in said process the molar ratio of alkene to carboxylic acid produced in said oxidation reaction zone is adjusted or maintained at a pre-determined value by controlling the concentrations of the alkene and optional water in said oxidation reaction zone and optionally by also controlling one or more of the pressure, temperature and residence time of the oxidation reaction zone.

Each of the alkane, molecular oxygen-containing gas, alkene and water may be introduced into the oxidation reaction zone as fresh feed and/or recycle component

If the at least one catalyst in the oxidation reaction zone de-activates or otherwise changes its selectivity in use, the molar ratio of alkene to carboxylic acid produced may be maintained at a constant, pre-determined value by controlling the concentrations of the alkene and optionally, water, fed to the oxidation reaction zone and optionally by also controlling one or more of the pressure, temperature and residence time of the oxidation reaction zone.

The present invention also provides a method of adjusting the molar ratio of alkene to carboxylic acid produced, for example, in response to changes in demand or requirement in down-stream processes by controlling the concentrations of alkene and optionally, water, fed to the oxidation reaction zone and optionally by also controlling one or more of the pressure, temperature and residence time of the oxidation reaction zone.

The process of the present invention is particularly useful when the alkene and/or carboxylic acid products are used at least in part in integrated downstream processes, for example (a) for the production of ester by reacting the carboxylic acid with the alkene or an alcohol or (b) for the production of alkenyl carboxylate by the reaction of an oxygen-containing gas with the carboxylic acid and alkene. Alkene and/or carboxylic acid may be recovered from the product of the oxidation reaction

zone and/or additional alkene and/or carboxylic acid may be used in the downstream process.

In a further embodiment of the present invention, the alkene and carboxylic acid may be produced in a molar ratio suitable for use in an integrated downstream process, for example (a) for the production of ester by reacting the carboxylic acid with the alkene or (b) for the production of alkenyl carboxylate by the reaction of a molecular oxygen-containing gas with the carboxylic acid and alkene. If alkene and/or carboxylic acid is neither separately recovered from reaction product nor separately added to the downstream process, the molar ratio of alkene to carboxylic acid produced in the oxidation reaction zone is suitably approximately 1 : 1, for example, 0.8 : 1 to 1.4 : 1. A different ratio may be produced if alkene and/or carboxylic acid is separately recovered from the oxidation reaction product or separately added to the down-stream process. The molar ratio of alkene to carboxylic acid may then be adjusted by controlling the concentrations of the alkene and optionally, water, in the total combined feeds to the oxidation reaction zone and optionally, by also controlling one or more of the pressure, temperature and residence time of the oxidation reaction zone, for example to meet changes in market demand or feedstock availability. Suitably, the molar ratio of alkene to carboxylic acid produced in the oxidation reaction zone is in the range 1 : 10 to 10 : 1.

Accordingly, the present invention provides an integrated process for the production of an alkyl carboxylate which process comprises the steps :

(a) contacting in an oxidation reaction zone a C<sub>2</sub> to C<sub>4</sub> alkane, a molecular oxygen-containing gas, the corresponding alkene and optionally, water, in the presence of at least one catalyst active for the oxidation of the alkane to the corresponding alkene and carboxylic acid, to produce a product stream comprising alkene and carboxylic acid and water; and

(b) contacting in a second reaction zone at least a portion of each of said alkene and carboxylic acid produced in the first reaction zone, in the presence of at least one catalyst active for the production of alkyl carboxylate to produce said alkyl carboxylate, and in which process the molar ratio of alkene to carboxylic acid produced in said oxidation reaction zone is adjusted or maintained at a pre-determined value by controlling the concentrations of the alkene and optional water in said oxidation reaction zone, and optionally by also controlling one or more of the pressure, temperature and

residence time of the oxidation reaction zone.

Also, in another embodiment, the present invention provides an integrated process for the production of an alkenyl carboxylate which process comprises the steps:

- (a) contacting in an oxidation reaction zone a C<sub>2</sub> to C<sub>4</sub> alkane, a molecular oxygen-containing gas, and the corresponding alkene and optionally, water, in the presence of at least one catalyst active for the oxidation of the alkane to the corresponding alkene and carboxylic acid, to produce a product stream comprising alkene and carboxylic acid and water; and
- (b) contacting in a second reaction zone at least a portion of each of said alkene and carboxylic acid produced in the first reaction zone and a molecular oxygen-containing gas, in the presence of at least one catalyst active for the production of alkenyl carboxylate to produce said alkenyl carboxylate, and in which process the molar ratio of alkene to carboxylic acid produced in said oxidation reaction zone is adjusted or maintained at a pre-determined value by controlling the concentrations of the alkene and optional water in said oxidation reaction zone, and optionally, by also controlling one or more of the pressure, temperature and residence time of the oxidation reaction zone.

Preferably, the molar ratio of alkene : carboxylic acid produced in the oxidation reaction zone is maintained at approximately 1 : 1, for example, 0.8 : 1 to 1.4 : 1 for subsequent use in a second reaction zone for the production of alkyl carboxylate or alkenyl carboxylate.

In the present invention, preferably, the C<sub>2</sub> to C<sub>4</sub> alkane is ethane, the corresponding alkene being ethylene and the corresponding carboxylic acid being acetic acid. These products may be reacted in downstream processes to produce ethyl acetate or, with a molecular oxygen-containing gas to produce vinyl acetate.

Typically, the oxidation reaction is performed heterogeneously with solid catalysts and the reactants in the fluid phase. In this case, the concentrations of alkene and optional water may be controlled as partial pressures in the oxidation reaction zone.

Catalysts active for the oxidation of alkane to alkene and carboxylic acid may comprise any suitable catalysts known in the art, for example, for the oxidation of ethane to ethylene and acetic acid as described in US 4596787, EP-A-0407091, DE 19620542, WO 99/20592, DE 19630832, WO 98/47850, WO 99/51339, EP-A-0 1043064, WO 9913980, US 5300682 and US 5300684, the contents of which are hereby

incorporated by reference.

US 4596787 relates to a process for the low temperature oxydehydrogenation of ethane to ethylene using a catalyst having the empirical formula  $Mo_aV_bNb_cSb_dX_e$  as therein defined, the elements being present in combination with oxygen.

5 EP-A-0407091 relates to process and catalyst for the production of ethylene and/or acetic acid by oxidation of ethane and/or ethylene in the presence of an oxidation catalyst comprising molybdenum, rhenium and tungsten.

DE 19620542 relates to molybdenum, palladium, rhenium based oxidation catalysts for the production of acetic acid from ethane and/or ethylene.

10 WO 99/20592 relates to a method of selectively producing acetic acid from ethane, ethylene or mixtures thereof and oxygen at high temperature in the presence of a catalyst having the formula  $Mo_aPd_bX_cY_d$  wherein X represents one or several of Cr, Mn, Nb, Ta, Ti, V, Te and W; Y represents one or several of B, Al, Ga, In, Pt, Zn, Cd, Bi, Ce, Co, Rh, Ir, Cu, Ag, Au, Fe, Ru, Os, K, Rb, Cs, Mg, Ca, Sr, Ba, Nb, Zr, Hf, Ni, P,  
15 Pb, Sb, Si, Sn, Tl and U and  $a=1$ ,  $b=0.0001$  to  $0.01$ ,  $c=0.4$  to  $1$  and  $d=0.005$  to  $1$ .

German patent application DE 196 30 832 A1 relates to a similar catalyst composition in which  $a=1$ ,  $b>0$ ,  $c>0$  and  $d=0$  to  $2$ . Preferably,  $a=1$ ,  $b=0.0001$  to  $0.5$ ,  $c=0.1$  to  $1.0$  and  $d=0$  to  $1.0$ .

WO 98/47850 relates to a process for producing acetic acid from ethane,  
20 ethylene or mixtures thereof and a catalyst having the formula  $W_aX_bY_cZ_d$  in which X represents one or several of Pd, Pt, Ag and Au, Y represents one or several of V, Nb, Cr, Mn, Fe, Sn, Sb, Cu, Zn, U, Ni, and Bi and Z represents one or several of Li, Na, K, Rb, Cs, Be, Mg, Ca, Sr, Ba, Sc, Y, La, Ti, Zr, Hf, Ru, Os, Co, Rh, Ir, B, Al, Ga, In, Tl, Si, Ge, Pb, P, As and Te,  $a=1$ ,  $b>0$ ,  $c>0$  and  $d$  is  $0$  to  $2$ .

25 WO 99/51339 relates to a catalyst composition for the selective oxidation of ethane and/or ethylene to acetic acid which composition comprises in combination with oxygen the elements  $Mo_aW_bAg_cIr_dX_eY_f$  wherein X is the elements Nb and V; Y is one or more elements selected from the group consisting of Cr, Mn, Ta, Ti, B, Al, Ga, In, Pt, Zn, Cd, Bi, Ce, Co, Rh, Cu, Au, Fe, Ru, Os, K, Rb, Cs, Mg, Ca, Sr, Ba, Zr, Hf, Ni, P,  
30 Pb, Sb, Si, Sn, Tl, U, Re and Pd;  $a, b, c, d, e$  and  $f$  represent the gram atom ratios of the elements such that  $0 < a \leq 1$ ,  $0 \leq b < 1$  and  $a + b = 1$ ;  $0 < (c + d) \leq 0.1$ ;  $0 < e \leq 2$ ; and  $0 \leq f \leq 2$ .

EP-A-1043064 relates to a catalyst composition for the oxidation of ethane to

ethylene and/or acetic acid and/or for the oxidation of ethylene to acetic acid which composition comprises in combination with oxygen the elements molybdenum, vanadium, niobium and gold in the absence of palladium according to the empirical formula :  $\text{Mo}_a\text{W}_b\text{Au}_c\text{V}_d\text{Nb}_e\text{Y}_f$  wherein Y is one or more elements selected from the group consisting of : Cr, Mn, Ta, Ti, B, Al, Ga, In, Pt, Zn, Cd, Bi, Ce, Co, Rh, Ir, Cu, Ag, Fe, Ru, Os, K, Rb, Cs, Mg, Ca, Sr, Ba, Zr, Hf, Ni, P, Pb, Sb, Si, Sn, Tl, U, Re, Te, La and Pd; a, b, c, d, e and f represent the gram atom ratios of the elements such that :  $0 < a \leq 1$ ;  $0 \leq b < 1$  and  $a + b = 1$ ;  $10^{-5} < c \leq 0.02$ ;  $0 < d \leq 2$ ;  $0 < e \leq 1$ ; and  $0 \leq f \leq 2$ .

WO 99/13980 relates to a catalyst for the selective oxidation of ethane to acetic acid of formula:  $\text{Mo}_a\text{V}_b\text{Nb}_c\text{X}_d$  wherein X is at least one promoter element selected from the group consisting of P, B, Hf, Te and As; a is a number ranging from about 1 to about 5; b is 1; c is a number ranging from about 0.01 to about 0.5; and d is a number ranging from greater than 0 to about 0.1.

US 5300682 relates to the use of oxidation catalyst with empirical formula of  $\text{VP}_a\text{MbO}_x$  where M is one or more of Co, Cu, Re, Fe, Ni, Nb, Cr, W, U, Ta, Ti, Zr, Hf, Mn, Pt, Pd, Sn, Sb, Bi, Ce, As, Ag and Au, a is 0.5 to 3, b is 0 1 and x satisfies the valence requirements.

US 5300684 relates to a fluid bed oxidation reaction using for example  $\text{Mo}_{0.37}\text{Re}_{0.25}\text{V}_{0.26}\text{Nb}_{0.07}\text{Sb}_{0.03}\text{Ca}_{0.02}\text{O}_x$ .

Other suitable oxidation catalysts for use in the present invention are described in WO 99/13980 which relates to the use of catalysts with elements in combination with oxygen in the relative gram atom ratios of  $\text{Mo}_a\text{V}_b\text{Nb}_c\text{X}_d$  where  $X = \text{P, B, Hf, Te or As}$ ; US 6030920 which relates to the use of catalysts with elements in combination with oxygen in the relative gram atom ratios of  $\text{Mo}_a\text{V}_b\text{Nb}_c\text{Pd}_d$ ; WO 00/00284 which relates to the use of catalysts with elements in combination with oxygen in the relative gram atom ratios of  $\text{Mo}_a\text{V}_b\text{Nb}_c\text{Pd}_d$  and/or  $\text{Mo}_a\text{V}_b\text{La}_c\text{Pd}_d$ ; US 6087297 which relates to the use of catalysts with elements in combination with oxygen in the relative gram atom ratios of  $\text{Mo}_a\text{V}_b\text{Pd}_c\text{La}_d$ ; WO 00/09260 which relates to the use of catalysts with elements in combination with oxygen in the relative gram atom ratios of  $\text{Mo}_a\text{V}_b\text{La}_c\text{Pd}_d\text{Nb}_e\text{X}_f$  where  $X = \text{Cu or Cr}$  and e and f can be zero; WO 00/29106 and WO 00/29105 which relate to the use of catalysts with elements in combination with oxygen in the relative gram atom ratios of  $\text{Mo}_a\text{V}_b\text{Ga}_c\text{Pd}_d\text{Nb}_e\text{X}_f$  wherein  $X = \text{La, Te, Ge, Zn, Si, In or W}$  and WO 00/38833 which relates to the use of catalysts with elements in combination with

oxygen in the relative gram atom ratios of  $\text{Mo}_a\text{V}_b\text{La}_c\text{Pd}_d\text{Nb}_e\text{X}_f$  wherein  $\text{X} = \text{Al}, \text{Ga}, \text{Ge}$  or  $\text{Si}$ , the contents of which are hereby incorporated by reference.

5 Solid catalysts active for the oxidation of the  $\text{C}_2$  to  $\text{C}_4$  alkane may be supported or unsupported. Examples of suitable supports include silica, diatomaceous earth, montmorillonite, alumina, silica alumina, zirconia, titania, silicon carbide, activated carbon and mixtures thereof.

Solid catalysts active for the oxidation of the  $\text{C}_2$  to  $\text{C}_4$  alkane may be used in the form of a fixed or fluidised bed.

10 The oxidation catalyst would be expected to oxidise at least part of any alkene fed to the oxidation reaction zone, for example to the corresponding carboxylic acid.

The molecular oxygen-containing gas used in the oxidation reaction zone, may be air or a gas richer or poorer in molecular oxygen than air. A suitable gas may be, for example, oxygen diluted with a suitable diluent, for example nitrogen or carbon dioxide. Preferably, the molecular oxygen-containing gas is oxygen. Preferably, at least some of  
15 the molecular oxygen-containing gas is fed to the oxidation reaction zone independently from the alkane and optional alkene feeds, and any recycle streams.

The alkane and alkene fed into the oxidation reaction zone of the process of the present invention may be substantially pure or may be admixed, for example, with one or more of nitrogen, methane, carbon dioxide, carbon monoxide, hydrogen, and low  
20 levels of  $\text{C}_3/\text{C}_4$  alkenes/alkanes.

Suitably, the concentration of alkene (as fresh feed and/or recycle component) is from greater than 0 and up to and including 50 mol % of the total feed, including recycles, to the oxidation reaction zone, preferably from 1 to 20 mol %, more preferably from 1 to 15 mol %.

25 Suitably, the concentration of water (as fresh feed and/or recycle component) is from 0 to 50 mol % inclusive of the total feed, including recycles, to the oxidation reaction zone, preferably from 0 to 25 mol %.

Advantageously, the process of the present invention allows the alkene : carboxylic acid molar ratio to be adjusted or maintained at a pre-determined level whilst  
30 achieving a low selectivity to carbon oxides by controlling the concentration of both ethylene and water fed to the oxidation reaction zone. Thus, for example, where it is desired to obtain an alkene : carboxylic acid molar ratio of approximately 1: 1, the ratio can be achieved by adding alkene only or water only. However, the addition of alkene

alone results in high selectivity to carbon oxides and high quantities of water, for example greater than 15 mol%, are required to achieve low carbon oxide selectivity. By using an alkene, such as ethylene, in combination with water, the 1: 1 molar ratio can be achieved at low selectivity to carbon oxides and using only small amounts of water (such as less than 10 mol%).

By employing a co-feed of alkene and water, the level of carbon oxides produced may be reduced to less than 15 mol%, for example less than 10 mol%.

Thus, in a preferred embodiment of the present invention, the alkene, such as ethylene, and water are co-fed into the oxidation reaction zone.

Suitably, the alkene, for example, ethylene, and water may be used in a ratio of 1 : 0.1-250 by weight, such as 1 : 0.1-100 or 1 : 0.1-50 but preferably in a ratio 1 : 0.1-10 by weight.

Where it is desired to obtain an alkene : carboxylic acid molar ratio in the range 0.8 : 1 to 1.4 : 1, such as in the oxidation of ethane to produce ethylene and acetic acid, the alkene and water are preferably co-fed to the oxidation reaction zone in the ratio 1 : 0.1-10 by weight, such as 1 : 0.1-1 by weight or 1 : 0.1-2 by weight, for example 1 : 1 by weight or 1 : 2 by weight.

When solid catalysts are used in the oxidation reaction zone, the alkane, corresponding alkene, molecular-oxygen containing gas, optional water and any recycle gases are preferably passed through the oxidation reaction zone with a residence time corresponding to a combined gas hourly space velocity (GHSV) of 500-10,000hr<sup>-1</sup>; the GHSV being defined as volume (calculated at STP) of gas passing through the reactor divided by the bulk volume of settled catalyst.

The oxidation reaction of the present invention may suitably be carried out at a temperature in the range from 100 to 400°C, typically in the range 140 to 350°C.

The oxidation reaction of the present invention may suitably be carried out at atmospheric or superatmospheric pressure, for example, in the range from 80 to 400 psig.

As mentioned above, low selectivity to carbon oxides may be achieved by feeding both alkene and water to the oxidation reaction zone. Alternatively and/or additionally, the selectivity to carbon oxides may be reduced, for example, to less than 15 mol%, by decreasing the reaction temperature and/or decreasing the reaction pressure whilst maintaining all other reaction parameters constant.

Desirably, an alkene : carboxylic acid molar ratio in the range 0.8 : 1 to 1.4 : 1 may be achieved at low carbon oxide selectivity by reducing the reaction temperature and/or pressure, all other parameters being constant.

Typically, alkane conversions in the range 1 to 99% may be achieved in the oxidation reaction of the present invention.

Typically, oxygen conversions in the range 30 to 100% may be achieved in the oxidation reaction of the present invention.

In the oxidation reaction of the present invention, the catalyst suitably has a productivity in the range 10 to 10000 grams of carboxylic acid, such as acetic acid, per hour per kilogram of catalyst.

Depending upon the nature of any catalyst used in any downstream process, it is desirable that when used for the production of alkenyl carboxylate, such as vinyl acetate, the first product stream should have a low concentration of carbon monoxide by-product, as this may have an adverse effect on some catalysts for the production of alkenyl carboxylates e.g. vinyl acetate. Thus, it is preferred to use a catalyst in the oxidation reaction zone that gives negligible carbon monoxide by-product. An additional catalyst component in the oxidation reaction zone may be used to oxidise carbon monoxide to carbon dioxide. The additional catalyst component may be present in the oxidation catalyst or catalysts or in a secondary reaction zone.

When ethane is used as a reactant for the oxidation process, the product stream comprises acetic acid, ethylene and water, and may also contain ethane and oxygen, inert gas components such as argon and nitrogen and the by-products, acetaldehyde, carbon monoxide and carbon dioxide. Acetaldehyde and carbon monoxide may be converted by the molecular oxygen-containing gas to produce acetic acid and carbon dioxide respectively, either in downstream processes or, after recycling, in the oxidation reaction zone. Ethylene is present in the product stream of the oxidation reaction as unconverted reactant ethylene from the feed and/or as oxidation product of the ethane reactant.

The product from the oxidation process may be fed directly or indirectly after one or more separation stages, and as one or more feed streams, to a second reaction zone together with optional additional molecular oxygen-containing gas, optional additional alkene and optional additional carboxylic acid to produce alkenyl carboxylate, such as vinyl acetate. Carboxylic acid and/or alkene may be optionally

recovered from the product of the oxidation process.

Unconverted alkane and/or alkene may be recycled together or after at least partial separation from the downstream process to the oxidation reaction zone directly or indirectly after one or more separation stages.

5        Catalysts known in the art for the production of alkenyl carboxylates may be used in the process of the present invention. Thus, catalyst active for the production of vinyl acetate which may be used in a second reaction zone of the present invention may comprise, for example, catalysts as described in GB 1 559 540; US 5,185,308 and EP-A-0672453 the contents of which are hereby incorporated by reference.

10        GB 1 559 540 describes a catalyst active for the preparation of vinyl acetate by the reaction of ethylene, acetic acid and oxygen, the catalyst consisting essentially of: (1) a catalyst support having a particle diameter of from 3 to 7 mm and a pore volume of from 0.2 to 1.5 ml/g, a 10% by weight water suspension of the catalyst support having a pH from 3.0 to 9.0, (2) a palladium-gold alloy distributed in a surface layer of the  
15        catalyst support, the surface layer extending less than 0.5 mm from the surface of the support, the palladium in the alloy being present in an amount of from 1.5 to 5.0 grams per litre of catalyst, and the gold being present in an amount of from 0.5 to 2.25 grams per litre of catalyst, and (3) from 5 to 60 grams per litre of catalyst of alkali metal acetate.

20        US 5,185,308 describes a shell impregnated catalyst active for the production of vinyl acetate from ethylene, acetic acid and an oxygen containing gas, the catalyst consisting essentially of: (1) a catalyst support having a particle diameter from about 3 to about 7 mm and a pore volume of 0.2 to 1.5 ml per gram, (2) palladium and gold distributed in the outermost 1.0 mm thick layer of the catalyst support particles, and  
25        (3) from about 3.5 to about 9.5% by weight of potassium acetate wherein the gold to palladium weight ratio in said catalyst is in the range 0.6 to 1.25.

EP-A-0672453 describes palladium containing catalysts and their preparation for fluid bed vinyl acetate processes.

30        An advantage of using a palladium-containing catalyst is that any carbon monoxide produced in the first reaction zone will be consumed in the presence of oxygen and the palladium-containing catalyst in the second reaction zone, thereby eliminating the need for a separate carbon monoxide removal reactor.

Typically, the production of alkenyl carboxylate such as vinyl acetate in the

second reaction zone is carried out heterogeneously with the reactants being present in the gas phase.

Additional alkene reactant may be fed to the second reaction zone for the production of alkenyl carboxylate as well as the alkene from the oxidation reaction zone as oxidation product and/or unconsumed alkene reactant.

Additional alkene introduced into the second reaction zone for the production of alkenyl carboxylate may be substantially pure or may be admixed, for example, with one or more of nitrogen, methane, carbon dioxide, carbon monoxide, hydrogen, and low levels of C<sub>3</sub>/C<sub>4</sub> alkenes/alkanes.

The molecular oxygen-containing gas used in the second reaction zone for the production of alkenyl carboxylate may comprise unreacted molecular oxygen-containing gas from step (a) and/or additional molecular oxygen-containing gas.

The additional molecular oxygen-containing gas, if used, may be air or a gas richer or poorer in molecular oxygen than air. A suitable additional molecular oxygen-containing gas may be, for example, oxygen diluted with a suitable diluent, for example nitrogen or carbon dioxide. Preferably, the additional molecular oxygen-containing gas is oxygen. Preferably, at least some of the molecular oxygen-containing gas is fed independently to the second reaction zone from the alkene and carboxylic acid reactants.

At least part of the carboxylic acid fed to the second reaction zone may be liquid.

When solid catalysts are used in the second reaction zone for the production of alkenyl carboxylate, the product from the oxidation reaction zone, any additional alkene or carboxylic acid reactants, any recycle streams and molecular oxygen-containing gas are preferably passed through the second reaction zone at a combined gas hourly space velocity (GHSV) of 1000-10,000hr<sup>-1</sup>.

The second reaction zone for the production of alkenyl carboxylate may suitably be operated at a temperature in the range from 140 to 200°C.

The second reaction zone for the production of alkenyl carboxylate may suitably be operated at a pressure in the range 50 to 300 psig.

The second reaction zone for the production of alkenyl carboxylate may suitably be operated as either a fixed or a fluidised bed process.

Carboxylic acid conversions in the range 5 to 80% may be achieved in the

second reaction zone for the production of alkenyl carboxylate.

Oxygen conversions in the range 20 to 100% may be achieved in the second reaction zone for the production of alkenyl carboxylate.

Alkene conversions in the range 5 to 100% may be achieved in the second  
5 reaction zone for the production of alkenyl carboxylate.

In the second reaction zone for the production of alkenyl carboxylate, the catalyst suitably has a productivity in the range 10 to 10000 grams of alkenyl carboxylate per hour per kg of catalyst.

When the alkane used in the process of the present invention is ethane, the  
10 product stream from the second reaction zone for the production of alkenyl carboxylate may comprise vinyl acetate, water and acetic acid and optionally also unreacted ethylene, ethane, acetaldehyde, nitrogen, argon, carbon monoxide and carbon dioxide. Such a product stream may be separated by azeotropic distillation into an overhead fraction comprising vinyl acetate and water and a base fraction comprising acetic acid  
15 and water. The base fraction may be removed from the distillation column as liquid from the bottom of the column, or as a vapour one or more stages above the bottom of the column. Prior to such a distillation step, ethylene, ethane, acetaldehyde, carbon monoxide and carbon dioxide, if any, may be removed from the second product stream, suitably as an overhead gaseous fraction from a scrubbing column, in which a liquid  
20 fraction comprising vinyl acetate, water and acetic acid is removed from the base. The ethylene and/or ethane may be recycled to step (a) and/or step (b).

Vinyl acetate is recovered from the overhead fraction, suitably for example by decantation. The recovered vinyl acetate may, if desired, be further purified in known manner.

25 The base fraction comprising acetic acid and water may be recycled, with or preferably without further purification, to step (b) of the process. Alternatively, acetic acid is recovered from the base fraction and may be further purified if desired, in known manner, for example by distillation.

A suitable process for the production of esters by reaction of the carboxylic acid  
30 with the alkene is described in EP-A-0926126, the contents of which are hereby incorporated by reference and which relates to an esterification process comprising reacting in an addition reaction a lower olefin with a saturated lower aliphatic mono-carboxylic acid in the vapour phase in the presence of a heteropolyacid catalyst

characterised in that the reaction is carried out in a plurality of reactors set up in series such that the gases comprising the unreacted gases and products exiting from a first reactor are fed as the feed gas to a second reactor and those exiting from the second reactor are fed as feed gas to the third reactor and so on for the subsequent reactors, and  
5 an aliquot of the reactant monocarboxylic acid is introduced into the feed gas to each of the second and subsequent reactors so as to maintain the olefin to monocarboxylic acid ratio in the feed gas to each of the second and subsequent reactors within a pre-determined range.

The invention will now be illustrated by way of example only and with reference  
10 to the Figure and to the following examples.

The Figure represents in schematic block-diagram, apparatus suitable for use in the process of the present invention.

The apparatus comprises an oxidation reaction zone (1) provided with a supply of ethane and optionally ethylene (3), a supply of a molecular oxygen-containing gas  
15 (4), a supply of recycle gas comprising ethane and ethylene (5) and an outlet (18) for a first product stream. Depending on the scale of the process, the oxidation reaction zone (1) may comprise either a single reactor or several reactors in parallel or series.

The apparatus also comprises a second reaction zone (2) for acetoxylation of ethylene to vinyl acetate which is provided with means (17) for conveying at least a  
20 portion of the product from the first reaction zone into the second reaction zone, a supply of molecular oxygen-containing gas (9), a supply of recycle acetic acid (10) and an optional supply or supplies of ethylene and/or acetic acid (8). Depending on the scale of the process, the second reaction zone (2) may comprise either a single reactor or several reactors in parallel or in series.

The apparatus further comprises an optional scrubber (6) for the first reaction product; a scrubber (12) for the product from the second reaction zone; means (13) for separating acetic acid from the product of the second reaction zone; vinyl acetate purification means (14); optional acetic acid purification means (15) and one or more  
25 separation means (16) for separating carbon dioxide from recycle gases from the second  
30 reaction zone and optionally for recovery of ethylene product.

In use, the oxidation reaction zone (1) is provided with at least one catalyst each active for the oxidation of the ethane to form acetic acid and ethylene. Suitably the oxidation catalysts are solid catalysts. Molecular oxygen-containing gas is fed to the

oxidation reaction zone (1) from supply (4) through one or more inlets. A gaseous feedstock comprising ethane, and optionally ethylene is fed to the oxidation reaction zone (1) from supply (3). Recycle gas comprising ethane and ethylene is also fed to the oxidation reactor from supply (5). The molecular oxygen-containing gas, ethane and recycle gas are introduced into the oxidation reaction zone through one or more inlets separately or in partial or complete combination. Optionally at least one of the streams fed to the oxidation reactor also comprises water.

In the oxidation reactor a first product stream is produced which comprises ethylene (as product and/or unreacted feed), acetic acid, water, optionally unconsumed molecular oxygen-containing gas and by-products such as carbon monoxide, carbon dioxide, inerts and acetaldehyde. This may optionally be passed to a scrubber (16) from which gas and liquid are removed. The gas may be recycled after separating by-products such as carbon dioxide and optionally recovering ethylene product by methods known in the art. Acetic acid may be recovered from the liquid, for example by distillation.

At least a portion of the first product stream is fed by means (17) into the second reaction zone which is provided with an acetoxylation catalyst, suitably a solid catalyst. The at least a portion of the first product stream may be fed directly into the second reaction zone. Alternatively, prior to being fed into the second reaction zone, the at least a portion of the first product stream may be separated by suitable separation means (not shown) into a number of component streams, for example, an ethylene stream and an acetic acid stream, which are then fed into the second reaction zone.

A molecular oxygen-containing gas is fed to the second reaction zone from supply (9). Acetic acid is fed to the second reaction zone from recycle supply (10). Optionally, additional ethylene and/or acetic acid may be fed to the second reaction zone from supply or supplies (8). The first product stream, molecular oxygen-containing gas, recycle acetic acid and optional additional supplies of ethylene and/or acetic acid are fed into the second reaction zone through one or more inlets separately or in partial or complete combination.

In the second reaction zone the ethylene, acetic acid and molecular oxygen react to produce a second product stream comprising vinyl acetate.

The second reaction product is passed to scrubber (12) from which gas and liquid are separated. Carbon dioxide is separated from the gas and optionally ethylene

product recovered, in one or more separation stages (16) by methods known in the art. The remaining ethylene and ethane may be recycled to the first and/or second reactors. Acetic acid is separated from the scrubber liquid and is recycled to the second reaction zone. Optionally, acetic acid product may be recovered from the recycle stream by means (15), for example by distillation. Vinyl acetate product is recovered from the scrubber liquid by means (14), for example by distillation.

In use, if selectivity of the at least one catalyst in the oxidation reaction zone changes, for example by deactivation, the molar ratio of alkene to carboxylic acid produced in the oxidation reaction zone is adjusted or maintained at a pre-determined value by controlling the amounts of alkene and water fed to the oxidation reaction zone and optionally by also controlling one or more of the pressure, temperature and residence time of the oxidation reaction zone.

Preferably, the molar ratio of ethylene: acetic acid produced in the oxidation reaction zone is maintained at approximately 1 : 1 , for example, 0.8 : 1 to 1.4 : 1 for subsequent use in the second reaction zone for the production of vinyl acetate. A different ratio may be maintained if ethylene and/or acetic acid is separately recovered from the oxidation reaction product or separately added to the second reaction zone for the production of vinyl acetate. The molar ratio of ethylene to acetic acid may then be adjusted by controlling the amounts of the alkene and water fed to the oxidation reaction zone and optionally by also controlling one or more of the pressure, temperature and residence time of the oxidation reaction zone, for example to meet changes in market demand or feedstock availability.

#### **Preparation of catalyst active for ethane oxidation (Catalyst A)**

A solution was prepared by dissolving 45.88 g ammonium molybdate, 12.86 g ammonium vanadate, 8.10 g niobium chloride, 0.0713 g ammonium gold tetrachloride and 6.75 g oxalic acid in 400 ml water heated to 70°C with stirring. After 15 minutes, the solution water was heated to boiling point followed by evaporation to dryness over 2 hours. The resulting catalyst cake was ground and then calcined in static air in an oven at 400°C for 5 hours. The nominal empirical formula of the catalyst was :

$\text{Mo}_{1.00}\text{V}_{0.423}\text{Nb}_{0.115}\text{Au}_{0.0008}\text{O}_x$ .

#### **General Ethane Oxidation Reaction Method.**

Typically 5 ml of powdered catalyst A was mixed with 15 ml of glass beads of diameter 0.4 mm to form a diluted catalyst bed of 20 ml in volume. The diluted catalyst

was then loaded into a fixed bed reactor made of Hastelloy of dimensions 12 mm internal diameter and length 40 cm. The catalyst was maintained in position in the centre of the reactor using quartz wall plugs together with inert packing material above and below the catalyst bed. The apparatus was then pressure-tested at 20 bar with helium to check for leaks. Catalyst A was then activated by heating to 220°C at 5°C/min in helium at 21 bar for 4 hours, to ensure full decomposition of catalyst precursors.

The required flows of ethane, ethylene, 20 % oxygen in helium and water were then introduced to the reactor, to ensure the required inlet composition. This composition was typically in the range 42-57 % v/v ethane, 6.6 % oxygen, 0-20 % v/v ethylene, 0-25 % v/v water and balance helium. The total feed flow rate was maintained at a level to ensure a feed GHSV of 2000-9000/h. After equilibrating for 60 minutes, gas samples were taken from the outlet stream to a GC system (model Unicam 4400) to quantify ethane, ethylene, oxygen and helium.

The setpoint temperature of the reactor was increased until 50-75 % oxygen conversion was achieved, as indicated by the calculated level of oxygen in the outlet stream. Following a further equilibration period of 60 minutes, catalyst A was then evaluated under steady state conditions for a period of typically 4-5 hours. Exit gas volume was measured over the run period by a water-gas meter. Liquid products were collected and weighed after the run period. Composition of gas and liquid products was measured using GC analysis (Unicam 4400 and 4200 fitted with TCD and FID detectors respectively).

From analysis of the feed and product flow rates and compositions the following parameters were calculated :

$$\text{ethane conversion} = (\text{inlet mol ethane} - \text{outlet mol ethane}) / \text{inlet mol ethane} * 100$$

$$\text{oxygen conversion} = (\text{inlet mol oxygen} - \text{outlet mol oxygen}) / \text{inlet mol oxygen} * 100$$

$$\text{selectivity to acetic acid (C-mol \%)} = \frac{(\text{outlet mol acetic acid} * 2)}{(\text{mol ethane converted} * 2)} * 100$$

$$\text{selectivity to ethylene (C-mol \%)} = \frac{(\text{outlet mol ethylene} - \text{inlet mol ethylene}) * 2}{(\text{mol ethane converted} * 2)} * 100$$

selectivity to CO (C-mol %) = (outlet mol CO) / (mol ethane converted \* 2) \* 100

selectivity to CO<sub>2</sub> (C-mol %) = (outlet mol CO<sub>2</sub>) / (mol ethane converted \* 2) \* 100

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ethylene/acetic acid ratio = (outlet mol ethylene – inlet ethylene mol) / (mol acetic acid)  
\* 100

STY (space time yield) % = (g acetic acid) / kg catalyst bed / hour

10

Typically, mass balance and carbon balance for a reaction was found to be 100 +/- 5 %.

In the following experiments the catalyst used was prepared according to the procedure described above, but the catalysts were not necessarily all from the same batch.

15 **Example 1.**

In this example varying concentrations of ethylene were utilised as feedstock in the general reaction method above. No water was used. The results are shown in Table I below.

**Table I**

| Ethylene (mol %) | Ethane conversion % | Oxygen conversion % | Ethylene sel. % | COx sel. % | Acetic acid sel. % | Ethylene/ Acetic Acid molar ratio | STY Acetic acid g/kg-cat/h |
|------------------|---------------------|---------------------|-----------------|------------|--------------------|-----------------------------------|----------------------------|
| 0                | 11.4                | 76.8                | 76.6            | 10.8       | 12.6               | 6.08                              | 60.2                       |
| 5                | 8.6                 | 57.2                | 75.9            | 5.8        | 18.3               | 4.15                              | 60.5                       |
| 10               | 7.5                 | 56.7                | 64.6            | 12.3       | 23.1               | 2.80                              | 58.0                       |
| 15               | 5.9                 | 38.7                | 29.3            | 57.7       | 28.4               | 1.03                              | 47.2                       |

20

Process conditions : 300°C, 16 bar, 3200/h, 42-57 % ethane, 6.6 % oxygen, 0-15 % ethylene, balance helium.

From the above Table I it can be seen that at a level of 15 mol% ethylene in the reaction mixture, the molar ratio of ethylene to acetic acid is approximately unity.

25 **Example 2 (comparison).**

In this experiment, water in varying concentrations was utilised in the above general reaction method. This is not an example according to the present invention because no ethylene was used in the feed. The results are given in Table 2 below.

5

**Table 2**

Process conditions : 300°C, 16 bar, 3200/h, 42-57 % ethane, 6.6 % oxygen, 0-25 % water, balance helium.

| Water (mol %) | Ethane conversion % | Oxygen conversion % | Ethylene sel. % | COx sel. % | Acetic acid sel. % | Ethylene/ Acetic acid molar ratio | STY Acetic acid g/kg-cat/h |
|---------------|---------------------|---------------------|-----------------|------------|--------------------|-----------------------------------|----------------------------|
| 0             | 11.4                | 76.8                | 76.6            | 10.8       | 12.6               | 6.08                              | 60.2                       |
| 5             | 9.9                 | 64.5                | 70.6            | 5.9        | 23.5               | 3.00                              | 86.6                       |
| 15            | 14.4                | 97.3                | 65.7            | 7.4        | 27.0               | 2.43                              | 131.7                      |
| 25            | 10.5                | 66.8                | 50.1            | 5.6        | 44.3               | 1.13                              | 145.7                      |

From Table 2 it can be seen that 25 mol % water has to be used in the feed stream, to achieve a molar ratio of ethylene to acetic acid of approximately unity. It can also be seen that as the amount of water fed to the reactor increases, the selectivity to carbon oxides decreases.

10

**Examples 3-5.**

In these examples ethylene and water were co-fed in the general reaction method. The process conditions employed are shown in Table 3. The results are given in Table 4.

15

**Table 3**

| Example | Temp (deg.C) | Pressure (barg) | GHSV (1/h) | Contact time (secs) | Ethane (mol %) | Oxygen (mol %) | Ethylene (mol %) | Water (mol %) |
|---------|--------------|-----------------|------------|---------------------|----------------|----------------|------------------|---------------|
| 3       | 281          | 16              | 3200       | 8.9                 | 52             | 6.6            | 5                | 10            |
| 4       | 299          | 22              | 3200       | 11.8                | 52             | 6.6            | 10               | 5             |
| 5       | 300          | 16              | 3200       | 8.6                 | 52             | 6.6            | 10               | 5             |

20

25

**Table 4**

| Example | Ethane conversion % | Oxygen conversion % | Ethylene sel. % | COx sel. % | Acetic acid sel. % | Ethylene/ Acetic acid molar ratio | STY Acetic acid g/kg-cat/h |
|---------|---------------------|---------------------|-----------------|------------|--------------------|-----------------------------------|----------------------------|
| 3       | 8.9                 | 54.4                | 45.1            | 6.0        | 48.8               | 0.92                              | 162.9                      |
| 4       | 11.3                | 99.6                | 42.4            | 12.8       | 44.8               | 0.95                              | 197.3                      |
| 5       | 12.2                | 99.6                | 44.9            | 14.3       | 40.8               | 1.10                              | 175.1                      |

Table 4 shows that the molar ratio of ethylene to acetic acid is approximately 1 to 1 in each of Examples 3-5.

From inspection of Tables 3 and 4 it can be seen that the use of ethylene in conjunction with low amounts of water results in the production of low levels of carbon oxides. It can also be seen from Examples 3 and 5, that the selectivity to carbon oxides is much lower at the lower reaction temperature of Example 3.

The results of the above Examples 1-5 show that the molar ratio of ethylene to acetic acid produced in the first reaction zone can be adjusted or maintained at a pre-determined value by controlling the concentrations of ethylene and water fed to the oxidation reaction zone under different reaction conditions and, in particular, a molar ratio of about unity may be achieved. Furthermore, the molar ratio of approximately unity may be achieved at low carbon oxide selectivity.

#### **Examples 6 – 14**

Further reactions were performed at different temperatures and pressures to show the effect of these parameters on the molar ratio of ethylene : acetic acid. The results are shown in Table 5 below. The results obtained show the effect of varying pressure and temperature on the performance of catalyst A. Thus, a comparison for the ethylene : acetic acid molar ratio for Examples 6, 9 and 12 shows that reducing pressure increases the molar ratio of ethylene : acetic acid. A similar trend is seen in general at different temperatures for Examples 7, 10 and 13 and for Examples 8, 11 and 14. A comparison of Examples 6, 7 and 8 shows that at 22 bara reducing the temperature increases the molar ratio ethylene : acetic acid but decreases the ethylene : acetic acid molar ratio at 9 bara. The results show that the molar ratio of ethylene to acetic acid can be adjusted by controlling the temperature and/or pressure of the reaction. A ratio of approximately 1 : 1 is achieved in Examples 6 to 11.

It can also be seen from Table 5 that decreasing the pressure from 22 bara to 9 bara at a given temperature, results in a reduction in the amount of carbon oxides produced. In addition, at a given pressure, a decrease in temperature, generally, also results in lower carbon oxide formation.

**Table 5**

Conditions : 3200/h Total feed : 52 mol. % ethane, 6.6 mol. % oxygen, 10 % mol. ethylene, 5 mol. % water, balance helium.

5

| Ex. | Temp<br>°C | Pressure<br>bara | Ethane<br>conversion<br>% | Oxygen<br>conversion<br>% | Ethylene<br>selectivity<br>C-mol % | COx<br>selectivity<br>C-mol % | Acetic<br>acid<br>selectivity<br>C-mol % | Ethylene/<br>acetic acid<br>molar ratio | STY<br>Acetic acid<br>g/kg-cat/h |
|-----|------------|------------------|---------------------------|---------------------------|------------------------------------|-------------------------------|------------------------------------------|-----------------------------------------|----------------------------------|
| 6   | 299        | 22               | 11.3                      | 99.6                      | 42.4                               | 12.8                          | 44.8                                     | 0.95                                    | 197.3                            |
| 7   | 291        | 22               | 7.4                       | 86.8                      | 44.3                               | 18.1                          | 37.6                                     | 1.18                                    | 103.1                            |
| 8   | 281        | 22               | 6.6                       | 71.3                      | 47.4                               | 13.7                          | 38.9                                     | 1.22                                    | 92.4                             |
| 9   | 300        | 16               | 12.2                      | 99.6                      | 44.9                               | 14.3                          | 40.8                                     | 1.10                                    | 175.1                            |
| 10  | 291        | 16               | 9.3                       | 87.8                      | 44.2                               | 13.5                          | 42.3                                     | 1.04                                    | 148.9                            |
| 11  | 281        | 16               | 6.9                       | 61                        | 48.0                               | 11.1                          | 40.9                                     | 1.17                                    | 104.6                            |
| 12  | 299        | 9                | 6.6                       | 64.2                      | 63.3                               | 9.7                           | 27.0                                     | 2.35                                    | 66.7                             |
| 13  | 290        | 9                | 5.3                       | 46.4                      | 60.6                               | 9.4                           | 30.0                                     | 2.02                                    | 61.3                             |
| 14  | 280        | 9                | 4.6                       | 34.6                      | 58.5                               | 4.8                           | 36.7                                     | 1.59                                    | 62.0                             |

22

### Preparation of catalyst active for ethane oxidation (Catalyst B)

A solution 'A' was prepared by dissolving 43.2 g of ammonium molybdate in 100 ml of distilled water heated to 70°C with stirring. A solution 'B' was prepared by dissolving 11.4 g of ammonium vanadate in 120 ml of distilled water heated to 70°C with stirring. A solution 'C' was prepared by dissolving 16.18 g of ammonium niobium oxalate and 2.5 g of oxalic acid in 100 ml of distilled water heated to 70°C with stirring. Each of the solutions A, B and C was allowed to stand for 15 minutes to allow maximum solubilisation of the reaction components. Solution C was then added to solution B rapidly with stirring at 70°C. The solution B/C was stirred for 15 minutes at 70°C then solution A was rapidly added to it. After 15 minutes a solution 'D' (2.57 g of ammonium phosphate dissolved in 20 ml water) was added with stirring. The A/B/C/D solution was then heated to boiling point followed by evaporation to dryness over 1.5 hours. The resulting catalyst paste was oven dried at 120°C for 16 hours. After drying, the resulting catalyst cake was ground, sieved through a 0.2 mm mesh sieve and then calcined in static air in an oven at 350°C for 4 hours. The nominal formula of the catalyst was :



### Examples 15-19

In these Examples Catalyst B as prepared above was used in the general reaction method. The reaction was performed at different temperatures to show the effect of temperature on the molar ratio of ethylene to acetic acid. The results are shown in Table 6.

From an inspection of Table 6 it can be seen that by controlling the temperature the molar ratio of ethylene to acetic acid may be adjusted. Table 6 also illustrates that by reducing the operating temperature, not only can the desirable ethylene : acetic acid molar ratio of approximately unity be achieved, but the molar ratio can be achieved at low carbon oxide selectivity.

**Table 6**

Process conditions : 3200/h, 16 bar, 52 mol % ethane, 6.6 mol % oxygen, 10 mol % ethylene, 5 mol % water, balance helium.

| Ex | Temp<br>°C | Ethane<br>conversion<br>% | Ethylene<br>sel. % | Acetic<br>acid<br>sel. % | COx<br>sel. % | Ethylene/<br>acetic acid<br>molar ratio | STY<br>acetic acid<br>g/kg-cat/h |
|----|------------|---------------------------|--------------------|--------------------------|---------------|-----------------------------------------|----------------------------------|
| 15 | 290        | 4.1                       | 45.9               | 43.2                     | 10.9          | 1.06                                    | 65.7                             |
| 16 | 300        | 4.6                       | 37.4               | 47.5                     | 15.1          | 0.79                                    | 80.7                             |
| 17 | 310        | 4.9                       | 32.7               | 50.9                     | 16.4          | 0.64                                    | 93.4                             |
| 18 | 321        | 5.8                       | 31.2               | 49.6                     | 19.2          | 0.63                                    | 106.7                            |
| 19 | 331        | 6.7                       | 27.5               | 50.4                     | 22.1          | 0.55                                    | 125.7                            |

5

**Examples 20-22**

In these Examples catalyst B was employed in the general reaction method.

Varying concentrations of ethylene were co-fed into the reaction.

The process conditions employed and the results achieved are shown in Table 7.

- 10 The results given in Table 7 demonstrate that the molar ratio of ethylene : acetic acid may be adjusted or maintained at a predetermined level by controlling the concentration of ethylene co-fed into the oxidation reaction zone. As can be seen from Example 22, a molar ratio of about unity may be achieved.

15

**Table 7**

Process conditions : 300°C, 16 bar, 3200/h, 52 mol % ethane, 6.6 mol % oxygen, 0-10 mol % ethylene, 5 mol % water, balance helium.

| Ex.             | Ethylene<br>mol % | Ethane<br>conversion<br>% | Ethylene<br>sel. % | Acetic acid<br>sel. % | COx<br>sel. % | Ethylene/<br>acetic acid<br>molar ratio | STY<br>acetic acid<br>g/kg-cat/h |
|-----------------|-------------------|---------------------------|--------------------|-----------------------|---------------|-----------------------------------------|----------------------------------|
| 20 <sup>a</sup> | 0.0               | 2.8                       | 67.3               | 27.5                  | 5.2           | 2.45                                    | 28.2                             |
| 21              | 5.0               | 4.5                       | 55.4               | 35.1                  | 9.5           | 1.58                                    | 58.2                             |
| 22              | 10.0              | 4.6                       | 46.7               | 40.8                  | 12.5          | 1.14                                    | 70.0                             |

<sup>a</sup> this Example is not an example of the invention as no ethylene is fed into the oxidation reaction zone

20



**Claims :**

1. A process for the oxidation of a C<sub>2</sub> to C<sub>4</sub> alkane to produce the corresponding alkene and carboxylic acid which process comprises contacting in an oxidation reaction zone, said alkane, molecular oxygen-containing gas, and the corresponding alkene and optionally, water, in the presence of at least one catalyst active for the oxidation of the  
5 alkane to the corresponding alkene and carboxylic acid, to produce a product stream comprising alkene, carboxylic acid and water, wherein in said process the molar ratio of alkene to carboxylic acid produced in said oxidation reaction zone is adjusted or maintained at a pre-determined value by controlling the concentrations of the alkene and optional water in said oxidation reaction zone and optionally by also controlling one or  
10 more of the pressure, temperature and residence time of the oxidation reaction zone.
2. An integrated process for the production of an alkyl carboxylate which process comprises the steps :
  - (a) contacting in an oxidation reaction zone a C<sub>2</sub> to C<sub>4</sub> alkane, a molecular oxygen-containing gas, the corresponding alkene and optionally, water in the presence of at  
15 least one catalyst active for the oxidation of the alkane to the corresponding alkene and carboxylic acid, to produce a product stream comprising alkene and carboxylic acid and water; and
  - (b) contacting in a second reaction zone at least a portion of each of said alkene and carboxylic acid produced in the first reaction zone, in the presence of at least one  
20 catalyst active for the production of alkyl carboxylate to produce said alkyl carboxylate, and in which process the molar ratio of alkene to carboxylic acid produced in said oxidation reaction zone is adjusted or maintained at a pre-determined value by controlling the concentrations of the alkene and optional water in said oxidation reaction



zone, and optionally by also controlling one or more of the pressure, temperature and residence time of the oxidation reaction zone.

3. An integrated process for the production of an alkenyl carboxylate which process comprises the steps:

- 5 (a) contacting in an oxidation reaction zone a C<sub>2</sub> to C<sub>4</sub> alkane, a molecular oxygen-containing gas, and the corresponding alkene and optionally, water in the presence of at least one catalyst active for the oxidation of the alkane to the corresponding alkene and carboxylic acid, to produce a product stream comprising alkene and carboxylic acid and water; and
- 10 (b) contacting in a second reaction zone at least a portion of each of said alkene and carboxylic acid produced in the first reaction zone and a molecular oxygen-containing gas, in the presence of at least one catalyst active for the production of alkenyl carboxylate to produce said alkenyl carboxylate, and in which process the molar ratio of alkene to carboxylic acid produced in said
- 15 oxidation reaction zone is adjusted or maintained at a pre-determined value by controlling the concentrations of the alkene and optional water in said oxidation reaction zone, and optionally by also controlling one or more of the pressure, temperature and residence time of the oxidation reaction zone.
4. A process as claimed in any one of the preceding claims which the molar ratio of
- 20 alkene to carboxylic acid produced in the oxidation reaction zone is in the range 10 : 1 to 1 : 10.
5. A process as claimed in claim 4 in which the molar ratio of alkene to carboxylic acid produced in the oxidation reaction zone is in the range 0.8 : 1 to 1.4 : 1.
6. A process as claimed in any one of claims 2 to 4 in which alkene and/or
- 25 carboxylic acid is separately recovered from the oxidation reaction product or separately added to the second reaction zone.
7. A process as claimed in any one of the preceding claims in which the alkane is ethane, the corresponding alkene being ethylene and the corresponding carboxylic acid being acetic acid.
- 30 8. A process according to claim 2 wherein the alkyl carboxylate is ethyl acetate
9. A process according to claim 3 wherein the alkenyl carboxylate is vinyl acetate
10. A process according to claim 8 or claim 9 and wherein the molar ratio of alkene to carboxylic acid produced in the oxidation reaction zone is in the range 0.8 : 1 to

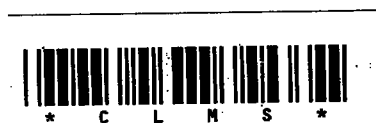


1.4 : 1.

11. A process according to any one of the preceding claims wherein the concentration of alkene fed to the oxidation reaction zone is from greater than 0 and up to and including 50 mol% of the total feed, including recycles.
- 5 12. A process according to any one of the preceding claims wherein the concentration of water fed to the oxidation reaction zone is from 0 to 50 mol% inclusive of the total feed, including recycles.
13. A process according to any one of the preceding claims wherein alkene and water are fed into the oxidation reaction zone.
- 10 14. A process according to any one of the preceding claims wherein the alkene and water are fed into the oxidation reaction zone in an alkene : water ratio of 1 to 0.1-250 by weight.
15. A process according to claim 14 wherein the ratio of alkene : water is 1 to 0.1-10 by weight.
- 15 16. A process according to claim 15 wherein the molar ratio of alkene : carboxylic acid is in the range 0.8 : 1 to 1.4 : 1.
17. A process according to any one of the preceding claims in which the at least one catalyst in the oxidation reaction zone comprises molybdenum.
18. A process according to any one of claims 3 to 17 in which the at least one
- 20 catalyst in the second reaction zone comprises palladium.
19. A process according to any one of the preceding claims in which the oxidation reaction is carried out at a temperature in the range 100 to 400 °C
20. A process according to any one of the preceding claims in which the oxidation reaction is carried out at atmospheric or superatmospheric pressure
- 25 21. A process according to any one of the preceding claims in which the oxidation reaction is carried out at a GHSV of 500-10,000 hr<sup>-1</sup>.
22. A process according to any one of the preceding claims wherein the product stream from the oxidation reaction zone also comprises carbon oxides in an amount of less than 15 mol%.
- 30 23. A process according to any one of claims 1 to 3 in which the alkane is ethane, the corresponding alkene is ethylene, the corresponding carboxylic acid is acetic acid and wherein ethylene and water are fed into the oxidation reaction zone in a ratio of 1 to 0.1 -10 by weight, the molar ratio of ethylene to acetic acid produced is in the range 0.8

: 1 to 1.4 : 1 and the product stream from the oxidation reaction zone also comprises carbon oxides in an amount of less than 15 mol%.

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