

TITLE OF INVENTION**PROCESS FOR CRYSTALLIZATION OF DICARBOXYLIC ACIDS****FIELD OF THE INVENTION**

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The present invention relates to a process for the crystallization of dicarboxylic acids.

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More specifically, the present invention relates to a novel process for the crystallization of dicarboxylic acids employing the use of additives which help increase the yield and purity of crystals.

The object of the invention is to achieve the desired increased size, high purity and greater yield of crystals.

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The other object of the invention is to achieve a crystallization process in which crystallization is at a faster rate as compared to the conventionally known processes.

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Still another object of the invention is to achieve the desired results by consuming less energy.

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The embodiment of the invention resides in the crystallization of di-carboxylic acids in a novel crystallizer comprising a cylindrical impeller and impeller shaft

Surfactants and /or salts are used as additives to achieve the desired increased size of crystals and high yield.

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Addition of surfactants as additives for the crystallization of dicarboxylic acids results in obtaining bigger crystals with high purity, whereas addition of salts as additives for the crystallization of dicarboxylic acids results in the greater yield of crystals.

The addition of surfactants and salts result in obtaining high yield crystals of having bigger crystal size.

5 BACKGROUND OF THE INVENTION

The modification of the crystal morphology and structure is achieved by introduction of additives which stick on one of the preselected crystal faces thereby inhibiting the growth of the crystal predictably perpendicular to the
10 surface on which they adhere. The size of the crystal structure is drastically varied by carrying out crystallization in electrolyte solutions.

U.S. Pat. No. 5,296,639 claims a process for purification of adipic acid during crystallization by modifying crystal morphology through the introduction of
15 additives which are among caproic acid and surfactants, thereby decreasing the incorporation of impurity on the crystal system.

U.S. Pat. No. 5,827,700 teaches a process of recovery and crystallization of citric acid from impure process stream by introducing very strong salts.
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Juetten, in his thesis entitled "The Enhanced Crystallization of Dicarboxylic Acids in Electrolyte Solutions", Michigan State University, 1992, performed crystallization experiments on aqueous solutions using dicarboxylic acids in order to determine by trial-and-error which electrolytes caused "salting out".
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One of the major drawbacks associated with the use of salts, either buffer salts or acid salts is the instant precipitation, which results in the smaller size of the crystals. In the subject invention to overcome this drawback, the selection of salts with the surfactants has resulted in achieving the desired
30 crystal size along with better yield of crystals.

5 U.S. Patent # 5296639 teaches a process for purification of adipic acid during crystallization, involving the modification of crystal morphology for the purpose of reducing incorporation of impurities, by adding an additive to the crystallizing solution. the additives are caproic acid or surfactant.

10 German Patent # DE 19705329 A (Rauls et al.) teaches a process for manufacture of dicarboxylic acid crystals from a solution containing it by addition of anionic polyelectrolytes having a molecular weight of at least 2000 and crystallizing the dicarboxylic acid. Polyacrylic acid with a molecular weight of 2,50,000 is used as a polyelectrolyte. Its concentration in the crystallizing mixture is preferably between 0.1 and 150 ppm. The dicarboxylic acid crystals have a purity of 99.5%.

15 U.S. Patent # 3102908 A (Raynes M et al.) teaches a process of continuous crystallization of adipic acid employing a vacuum crystalliser. The objective of the invention is to prevent the build-up of adipic acid crystals in the lines and walls of the vacuum crystalliser and to prevent foaming of the crystallizing solution in the
20 crystalliser.

Japanese Patent # 54115314 A (Asahi Chem. Ind. Co. Ltd.) teaches preparation of high purity adipic acid by re-crystallizing and separating adipic acid from a mixture containing adipic, glutaric and succinic acids and feeding the recovered adipic acid to the
25 purification process. The solvent used in re-crystallization is water or alcohol. The separation efficiency is improved by adding 5% w/w inorganic salt, based on the mixture weight, to the mixture from which adipic acid is re-crystallized.

European Patent # 0633240A (Haarmann & Reimer Corporation) teaches an
30 improved process for recovering citric acid from an impure process stream (containing citric acid) by first concentrating the stream, by evaporation of water, to 0-10% of the saturation point of citric acid in solution, followed by addition of strong acid or salts to crystallize out the citric acid. Crystals of citric acid are separated from the mother liquor (containing citric acid, colorants, carbohydrates, strong acid and other
35 impurities). The mother liquor is then passed through an ion exclusion resin column of the type strong acid cation exchange resin. Citric acid is obtained as the final eluate and is combined with the citric acid crystals obtained above, to increase yield, to give a citric acid solution. Final recovery of citric acid is through crystallization or spray granulation. Strong acid or salts added are chosen from among H_2SO_4 , HCl , Na_2SO_4
40 and HNO_3 , preferably H_2SO_4 is used.

U.S. Patent # 5149867 (Chen C. T. et al) teaches a process for crystallization and purification of a water-insoluble aromatic dicarboxylic acid such as 2,6-Napthalene dicarboxylic acid (NDA). The process involves forming a soluble di-salt of the crude NDA, by reacting with a base, followed by filtration. Partial acidification is carried out to liberate about 90% of the NDA and some mono salt of NDA as micro crystals. Holding this slurry for a length of time leads to formation of large crystals. Surfactants are added to reduce the holding time. Final acidification is carried out to convert more of the salt to the NDA.

U.S. Patent # 5567842 (Izumisawa Y. et al.) teaches a process for producing terephthalic acid suitable for use as a starting material for the production of polyesters by direct polymerization with a glycol. Important for this polymerization is the particle size distribution of the terephthalic acid crystals, which should be between 50 and 150 um and the absence of impurities. This patent aims at producing terephthalic acid crystals meeting these specifications. The aqueous solution of crude terephthalic acid at 260-320° C is purified and terephthalic acid is crystallized from this solution in a multi-stage crystalliser series. Different temperature and pressure conditions are maintained in each of the crystallizers, ranging from 260 to 80°C.

The present invention relates to a process for crystallization of dicarboxylic acids in a system, which is under continuous shear force that is created in the annular space between the impeller and the crystallizer wall. The system consists of selection and addition of additives in a manner to avoid any effect
5 on the crystal purity.

To achieve the objectives of increased size, high purity and greater yield of crystals, in the present invention crystallization of diacids is carried out in a solution consisting of a solvent or a mixture of solvents, crude diacid and the
10 additives in a crystallizer.

The crude diacids consists of one solute in major amount and impurities in minor amounts. When re-crystallized, the subject process provides a highly purified crystals of the diacids, as the amounts of impurities present, if any,
15 are totally dissolved in the solvent/solution.

The solvents used in the present invention are selected preferably from water, any aliphatic monobasic acid having 2-4 carbons comprising primary and secondary hydrogen, any primary alcohol having 2-4 carbon or any
20 combination thereof. Preferably, the solvent used in the subject process is water or water with small amount of acetic acid, wherein the ratio of water to acetic acid is more than 1:1. More preferably, the solvent with less than 10% of acetic acid is used. The preferred range of the solvent to the crude diacid is between 10 to 1 to 0.1 to 1. Most preferred range of the solvent to the
25 crude diacid is 5 to 0.2

The additives used in the subject invention are selected from surfactants, and /or buffer salts and acid salts.

30 The addition of surfactants only results in obtaining bigger crystals with high purity. The surfactants are selected from non-ionic surfactants, anionic surfactants, cationic surfactants and zwitterionic surfactants. Preferably, the

anionic surfactants are used. The preferred anionic surfactants can be selected from Sodium dodecyl sulfate (SDS), Sodium dodecyl benzene sulfonate (SDBS) and Sodium bis (2-ethyl hexyl) sulfosuccinate (AOT).

- 5 The nonionic surfactant used in the present invention are selected from hexoxy ethylene glycol mono-n-dodecyl ether ($C_{12}E_6$) and Tween 20.

The cationic surfactants can be selected from cetyl trimethyl ammonium bromide (CTAB) and dodecyl dimethyl ammonium bromide (DDAB).

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The zwitterionic surfactants are selected from Phosphatidyl choline (PC) and Phosphatidyl ethanolamine (PE).

- 15 The amount of surfactants that is required for achieving good crystal size and shape is between 0.005% to 2.0% of the weight of the total solution. It has been observed that the concentration of surfactant below the critical micellar concentration (CMC) in the solution has achieved the desired results successfully. Most preferably, the amount of surfactant used is between 0.01% to 1.0% i.e. below the CMC.

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- To increase the yield of crystals the buffer acids and/or buffer salts are used as additives. The addition of salts results in increased rate of crystallization thus resulting in greater yield of crystals. The buffer salts exhibit resistance to the change in pH and acid salts increases the pH. The buffer salts used in the present process are selected from Na_2HPO_4 , NaH_2PO_4 , KH_2PO_4 and K_2HPO_4 . Preferably these salts are selected from Na_2HPO_4 or K_2HPO_4 or a combination thereof. The amount of buffer salts taken is between 0.0025% to 1.0% of the weight of the total solution.

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30 The acid salts used in the present invention are selected from $NaHSO_4$ and $KHSO_4$, more preferably $KHSO_4$. The amount of acid salts taken is between 0.0025% to 1.0% of the weight of the total solution.

Hence, in the subject crystallization process, both the acid and buffer salts along with the surfactants are used to achieve the desired average crystal size as surfactants makes the crystal grow and the salts precipitate out of the di-carboxylic acid.

The presence of surfactants results in growth of crystals, thereby results in obtaining crystals of higher purity. Salts present in the system results in the increased yield i.e. the weight of the crystals formed is more than the weight of crystals formed in a conventional process. The same is shown in the comparative table :-

	Conven- tional	with surfactant only(0.5% wt of solution)	with surfactant +salts (0.25+0.2 5% wt of solution)	with salts only(0.5 % wt of solution)
Yield (%)	60.95	62.75	76.67	81.23
Purity(%)	99.8	100	99.92	99.8
Time for crystallization (hrs)	1.00	1.00	1.00	1.00
Crystal size (microns)	655	1814	1041	310

TABLE-1

The amount of surfactants and buffer or acid salts is between 0.005% to 4.0 % of the weight of the total solution.

As can be clearly seen from the Table-1, the presence of surfactants helps only in the growth of the crystals and the size of the crystal is increased by three times approximately. Increased size results in increasing the purity of crystals. Here, the yield increases but the increase is not considerable.

In case of the use of salts only as additives, the size of the crystals decreases, but the yield increases.

However, in the presence of both the salts and surfactants, there is increase
5 in the yield, purity and size of the crystals.

In accordance with the present invention the diacid is crystallized in medium consisting of a suitable solvent, which is more preferably a mixture of water and acetic acid in different proportions and additives which are selected from
10 surfactants, buffer salts or/and acid salts in a novel crystallizer having a impeller with impeller shaft.

In the process for the crystallization of di-carboxylic acid, the crude diacid at an acid value of 15-35, is put in the annular space, between the impeller and
15 crystallizer wall. The crude diacid is dissolved in solvent at a suitable temperature at which the crude diacid is dissolved in the solvent to form a solution. The said solution is evaporated till a final acid value of 18-64 is achieved at a temperature of between 60-75° C. The resultant solution is then treated with additives and is cooled to 25- 35 ° C and crystallized for one
20 hour, which results in the crystals of the di-acid to purge out of the solution which are later recovered by filtration/centrifugation.

SUMMARY

25 Accordingly, the subject invention relates to a process for the crystallization of di- carboxylic acids comprising :-

dissolving the crude dicarboxylic acid at a temperature between 30° C to 75°C to dissolve the crude diacid in the solvent in an evaporator whereby
portion of the said solvent evaporates increasing the concentration of the
30 dicarboxylic acid in the said solution, transferring the said solution from the evaporator to a crystallizer, adding the additives selected from surfactant, buffer salts or/and acid salts or mixture thereof in the said evaporated solution

in the said crystallizer, cooling the said mixture resulting in the formation of crystals of dicarboxylic acids, wherein the said crystallization takes place in the annular space between the impeller and the crystallizer wall at 10-80° C.

- 5 The crystal forms are needle shaped, monoclinic crystals. The additives modifies the shape of crystal by increasing the leniency and decreasing the surface area.

- 10 Dissolution of the crude dicarboxylic acid in a solvent and the addition of additives in the said solvent can be carried in a separate evaporator out side the said crystallizer.

BRIEF DESCRIPTION OF THE ACCOMPANYING DRAWINGS

- 15 Figure 1 depicts the diagram of the novel crystallizer

DETAILED DESCRIPTION OF THE INVENTION

- 20 The novel crystallizer used in the subject invention has a cylinder coaxial impeller (1) or series of impellers having different diameters separated from each other, moving at different speeds with the innermost impeller having the diameter between 5 cm to 3 meters, moving at highest speed preferably at the rate of 10-50 RPM and the outermost impeller having the diameter of between 5 cm to 10 meters, moving at lowest speed preferably at the rate of
- 25 1-10 RPM, rotated by a common shaft (3) which is attached to the motor. For these corresponding diameters of the innermost and the outermost impellers the corresponding annular space (2) between the two cylindrical impellers is between 1 mm to 50 cms

- 30 The process of crystallization of dicarboxylic acid takes place in the annular space between the walls of cylindrical impellers. Since, the flow is unidirectional in the subject crystallizer, crystals are aligned in one direction.

The surfactant molecules present in the said solution comes over the crystals, which results in increasing the size of the crystals. The crystals thus formed settle down due to the force of gravity.

- 5 The use of novel crystallizer of the subject invention as compared to the conventionally available crystallizer has found to give better yield, high purity, increased crystal size and reduced time used for crystallization, as shown in the comparative Table 2.

	CONVENTIONAL	NOVEL CRYSTALLIZER OF THE SUBJECT INVENTION
YIELD (%)	60.95	61.45
PURITY(%)	99.8	99.8
TIME FOR CRYSTALLIZATION IN MINUTES	60	40
CRYSTAL SIZE IN MICRONS	655	710

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TABLE-2

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The newly designed impeller used in the present invention helped in relatively increasing the rate of crystallization and growth of crystals over conventional crystallization processes. The crystals that are formed in the top part of the crystallizer fall down rapidly due to the force of gravity. Crystals are continuously removed from time to time if the process is continuous and are allowed to stay at the bottom of the crystallizer if the process is in batch mode.

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To achieve the faster crystallization rate the specially designed impeller is used. The criticality of the inventive part of the crystallizer resides in the gap i.e. annular space between the impeller and the wall of the crystallizer. The

annular space between the agitator and the crystallizer is maintained between 2 cms to 100 cms.

The subject invention can be best illustrated by way of following examples.

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EXAMPLES 1-5

Example 1

Crystallization of crude succinic acid is carried out in an annular space
10 between the impeller and crystallizer wall. 10 gms of crude succinic acid
(around 98% pure) is initially taken with an acid value of 18.37. The solvent is
water. The additives are introduced. The amount of surfactant, sodium lauryl
sulfate is 0.5% of the crude acid weight. The percentage of acid salts is 0.5%
of the crude acid weight. The salts taken is KHSO_4 . The solution is
15 evaporated till a final acid value of 20.132 is reached. The temperature of
evaporation is 65°C . The system is then cooled to around 32°C and
crystallized in the novel crystallizer for 1 hr. The solution is then filtered and
dried. A total yield of 72.014% of highly pure succinic acid is obtained. The
amount of impurities are less than 100 ppm. The average size of the crystals
20 is 996 microns. The conventional process gave a yield of 52.36% of lesser
purity. The average size of the crystals in conventional process was 514
microns.

Example 2

25 Crystallization of crude succinic acid is carried out in an annular space
between the impeller and crystallizer wall. 10 gms of crude succinic acid
(around 98% pure) is initially taken with an acid value of 20. The solvent is
water. The additives are introduced. The amount of surfactant, sodium lauryl
sulfate is 1.0% of the total diacid weight. The percentage of acid salts is 1.0%
30 of the total diacid weight. The salts taken is KHSO_4 . The solution is
evaporated till a final acid value of 30 is reached. The temperature of
evaporation is 65°C . The system is then cooled to around 28°C and

crystallized in the novel crystallizer for 1 hr. The solution is then filtered and dried. A total yield of 83.66% of highly pure succinic acid is obtained. The amount of impurities are less than 150 ppm. The average size of the crystals is 1151 microns. The conventional process gave a yield of 58.36% of lesser purity (99.12%). The average size of the crystals in conventional process was 506 microns.

Example 3

Crystallization of crude oxalic acid is carried out in an annular space between the impeller and crystallizer wall. 10 gms of crude oxalic acid (around 98% pure) is initially taken with an acid value of 30.04, in a mixture of water (90% by volume) and acetic acid (10% by volume). Sodium lauryl sulfate in an amount of 1.0% of the total solution weight is added along with acid salts in an amount of 1.0% of the total weight of the solution. The salts taken are a mixture of KHSO_4 and K_2HPO_4 . The solution is evaporated till a final acid value of 52.62 is reached. The temperature of evaporation is 65°C . The system is then cooled to around 30°C and crystallized in the novel crystallizer for 1 hr. The solution is then filtered and dried. A total yield of 78.86% of highly pure oxalic acid is obtained. The amount of impurities are less than 50 ppm and the average crystal size obtained is 998 microns.

Example 4

Crystallization of crude Adipic acid is carried out in an annular space between the impeller and crystallizer wall. 6.034 gms of crude Adipic acid is initially taken with an acid value of 30.04 and the purity of 66.7% in water. Sodium lauryl sulfate in an amount of 1.0% of the total solution weight is added along with acid salts in an amount of 1.0% of the total weight of the solution. The salt taken is Na_2HPO_4 . The solution is evaporated till a final acid value of 60.62 is reached. The temperature of evaporation is 73°C . The system is then cooled to around 30°C and crystallized in the novel crystallizer for 1 hr. The solution is then filtered and dried. A total yield of 89.22% of highly pure adipic

acid is obtained. The amount impurities are less than 50 ppm. The average
5 crystal size obtained is 880 microns.

Example 5

Crystallization of crude succinic acid is carried out in an annular space
between the impeller and crystallizer wall. 5.006 gms of crude succinic acid
10 (around 98% pure) is initially taken with an acid value of 16.583. The solvent is
water. The additives are introduced. The amount of surfactant, sodium
dodecyl sulfate is 0.5% of the crude succinic acid weight. The weight
percentage of salts is 0.5% of the crude succinic acid weight. The salts taken
is KHSO_4 . The solution is evaporated till a final acid value of 20 is reached.
15 The temperature of evaporation is 65°C . The system is then cooled to around
 28°C and crystallized in the novel crystallizer for 1 hr. The solution is then
filtered and dried. A total yield of 72.78% of highly pure succinic acid is
obtained. The amount of impurities are less than 100 ppm. The average size
of the crystals is 1008 microns. The conventional process gave a yield of
20 52.36% of lesser purity. The average size of the crystals in conventional
process was 514 microns.

The subject process is a mere statement of invention and should not in any
way construed to restrict the scope of the invention.

CLAIMS:

5 1. A process for the crystallization of dicarboxylic acids comprising the steps of:-

recrystallization a solution of crude dicarboxylic acid in a crystallizer in the presence of at least one additive selected from the group consisting of surfactants, buffer salts, acid salts or a mixture thereof, wherein the
10 crystallizer comprises a cylindrical impeller, a crystallizer wall and a rotatable impeller shaft; and
cooling the mixture to form crystals of dicarboxylic acids between the cylindrical impeller and the crystallizer wall at about 10°C to about 80°C, wherein the impeller and the crystallizer wall creates a space which is
15 between about 2 to about 100 cm.

2. A process as claimed in claim 1, wherein the said crystallizer comprises a cylinder coaxial impeller or series of such impellers having
20 different diameters separated from each other conforming to the shape of the crystallizer, and each of these impellers moving at different speeds with the innermost impeller having a diameter between about 0.05m to about 3m, moving at a speed at a rate of about 10 to about 50 RPM and the outermost impeller having a diameter between about 0.05m to about 10m, moving at a
25 lower speed than the inner impeller at a rate of about 1 to about 10 RPM, the impeller or series of impeller being rotated by a common gear shaft driven by a motor.

30 3. A process as claimed in claim 2, wherein the space between the innermost impeller and the outermost impeller is between about 0.01cm to about 50cm.

5 4. A process as claimed in claim 1, wherein the dicarboxylic acid solution
comprises said dicarboxylic acid dissolved in a solvent, wherein the solvent is
selected from the group comprising of water, any aliphatic monobasic acid
having 2-4 carbons, wherein the monobasic acid has a primary or secondary
hydrogen such as acetic acid, any primary alcohol having 2-4 carbons or any
10 combination thereof.

15 5. A process as claimed in claim 4, wherein the solvent is a mixture of
water and acetic acid in the ratio greater than 1:1.

20 6. A process as claimed in claim 4, wherein, the solvent comprises less
than 10% acetic acid.

25 7. A process as claimed in claim 1, wherein the surfactants are selected
from a group comprising anionic surfactants, non-ionic surfactants, cationic
surfactants and zwitterionic surfactants, most preferably anionic surfactants in
the ratio of 0.005% to 2.0% of the weight of the solution.

30 8. A process as claimed in claim 7, wherein the anionic surfactants are
selected from sodium dodecyl sulfate (SDS), sodium dodecyl benzene
sulfonate (SDBS) and sodium bis (2-ethyl hexyl) sulfosuccinate (AOT).

5 9. A process as claimed in claim 7, wherein the nonionic surfactants are selected from hexoxy ethylene glycol mono-n-dodecyl ether ($C_{12}E_6$) and polysorbates.

10 10. A process as claimed in claim 7, wherein, the cationic surfactants are selected from cetyl trimethyl ammonium bromide (CTAB) and dodecyl dimethyl ammonium bromide (DDAB).

15 11. A process as claimed in claim 7, wherein the zwitterionic surfactants are selected from phosphatidyl choline (PC) and phosphatidyl ethanolamine (PE).

12. A process as claimed in claim 1, wherein the solution after evaporation is cooled to a temperature of between about 28°C to about 32°C.

20 13. A process as claimed in claim 1, wherein the solution of crude diacid is evaporated to have a final acid value of about 18 to about 64 at a temperature between about 60° C to about 75°C.

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