

## GASEOUS MATTER CAPTURE SYSTEM AND METHOD

### FIELD OF THE INVENTION

The present invention relates to the field of gaseous matter capture, and, more particularly, but not exclusively, to capturing carbon dioxide directly from the atmosphere for the purposes of  
5 climate change mitigation and further use of captured gases.

### BACKGROUND OF THE INVENTION

Climate change has long been a global concern having a potential enormous impact on the global environment and human wellbeing. Human activities such as the combustion of fossilized fuels and deforestation, along with derivative phenomena such as accelerated permafrost thawing,  
10 increase the amount of greenhouse gases in the earth's atmosphere and cause the global climate to change. As a result, many concepts were tested and implemented in order to mitigate the effects of climate change.

Nowadays, the Carbon Dioxide concentration in the earth's atmosphere is 411 parts per million (ppm). This amount increases by over 2 ppm per year, due to the continued emissions in  
15 the multiple and distributed sectors of the world's economy. According to the Paris Agreement led by the UNFCCC and signed by most countries in 2015, mankind has to limit the average temperature increase to 'well below' 2°C compared to pre-industrial levels, in order to avoid catastrophic consequences. In order to try and predict how to avoid said potentially catastrophic 2°C increase limit, models vary between allowing for a remaining carbon quota, but generally aim  
20 at remaining below 430 ppm, whereas 450 ppm indicates an approximate transition to a high probability of irreversible effects as one ppm roughly translates to several Billions of metric tons

of CO<sub>2</sub>, this implies the need to remove green-house gases from the atmosphere in the order of tens of Billions tons per year.

A variety of carbon dioxide capture methods have been described in the art. Among them are scrubbers configured to be implemented in streams of high carbon dioxide concentrations like those found in exhausts of power plants, industrial facilities or vehicles. Although such methods of capturing carbon from flue gas may provide some mitigation to the continuous increase in carbon emissions, they are limited in their applicability, cannot address carbon emissions from distributed sources and also cannot address the high carbon dioxide concentration already circulating as part of the earth's atmosphere.

The capturing of carbon dioxide directly from the earth's atmosphere using chemical processes of different types has been disclosed by several publications, for example, patent applications publications US20170106330A1 and US20170028347A1 disclose the capturing of carbon dioxide conducted by stationary, terrestrial systems using sodium hydroxides that later allows compression of a high purity carbon dioxide stream into liquids or supercritical liquids. Other means of chemical and physical capturing and processing of carbon dioxide have been disclosed, for example, in patent applications publications WO2016185387A1, AU2008239727B2 and US20140061540A1. Patent publications US4963165A and US8702847B2 also disclose the capturing, separating, condensing and recycling of carbon dioxide.

Patent applications publication US20170106330A1 discloses a system for separating and storing molecules, atoms and/or ions from air, comprising at least one collecting tank configured to receive molecules, atoms and/or ions that are separated from ambient air. The system further

comprising at least one storing tank for storing the separated molecules, atoms and/or ions, and at least one outlet, wherein the air collecting means can be a gas tower or a gas balloon.

Patent applications publications US20110146488A discloses capturing and sequestering significant amounts of carbon dioxide molecules from an incoming air stream by directing flow  
5 into an airborne cylindrical carbon composite canister, or “Atmospheric Carbon Dioxide Mitigation” (ATCOM) canister which has the capacity to capture, sequester, and then release the species with negative ionization within a desired high frequency electromagnetic wave field. The initial airflow into the ATCOM canister is slowed to a specific flow velocity as the air stream travels through a volute chamber with resistance added impellers, and then into a free flow  
10 chamber where the incoming flow velocity compresses the air volume, allowing for an osmotic equality distribution of the concentration of CO<sub>2</sub> molecules.

From the state of the art indicated above, one can notice that different trials and development are being conducted, although generally these efforts do not manage to meet the market requirements in terms of price, mitigation (with regard to carbon emissions per ton of  
15 carbon dioxide captured) and applicability.

Neither of the publications indicated above do not teach, alone or in combination, a gaseous matter capture system, comprising an aerial unit and a non-aerial unit and configured to transfer the stored separated gaseous matter from the aerial unit to the non-aerial unit for further processing or storage.

20 There is a need to provide a system and method for capturing gaseous matter directly from the atmosphere in an economical, scalable and applicable manner.

There is a further need to provide a system and method configured to release storage means full of compressed gaseous matter for further processing by a non-aerial unit, thus increasing the gaseous matter collecting efficiency by allowing to capture more gaseous matter such as carbon dioxide mass in a single airborne mission hence reducing regular maintenance and ground time intervals.

#### SUMMARY OF THE INVENTION

The present invention provides a system and method for capturing gaseous matter directly from the atmosphere which is economical and highly scalable with regard to any other available system and method.

Said system and method may further include using the climatic conditions found at high altitude that enable gases' phase transitions at low temperatures and relatively low pressures in order to liquefy or solidify gaseous matter such as carbon dioxide, and thus separate it from the other gases forming the atmospheric mixture.

Said system and method may further include utilizing high altitude platform/vehicle such as a high-altitude balloon configured to capture large amounts of high altitude gaseous matter such as CO<sub>2</sub>, wherein said high altitude CO<sub>2</sub> concentration tends not to be diluted due to the typical strong winds and resulting advection.

Said system and method may further include transferring the stored separated gaseous matter from the aerial unit to the non-aerial unit for further processing or storage.

Said system and method may further include increasing the gaseous matter collecting efficiency by allowing to capture more gaseous matter such as carbon dioxide mass in a single airborne mission hence reducing regular maintenance and ground time intervals.

The following embodiments and aspects thereof are described and illustrated in  
5 conjunction with systems, devices and methods which are meant to be exemplary and illustrative and not limiting in scope. In various embodiments, one or more of the above-described problems have been reduced or eliminated, while other embodiments are directed to other advantages or improvements.

According to one aspect, there is provided a gaseous matter capture system, comprising at  
10 least one aerial unit configured to be airborne, at least one non-aerial unit, at least one gas separation means configured to be carried by the aerial unit, storage means configured to be carried by the aerial unit, a controller configured to control the system's operation and an energy source configured to enable the system's operation, wherein the at least one gas separation means is configured to separate at least one designated gaseous matter from the air, wherein the at least one  
15 separated gaseous matter is configured to be stored within the storage means and wherein the aerial unit is configured to transfer the stored separated gaseous matter to the non-aerial unit.

According to some embodiments, the at least one gas separation means is operable while the aerial unit is airborne at an altitude range of 5-15 km.

According to some embodiments, the gas separation means comprises at least one pressure  
20 increasing apparatus.

According to some embodiments, the gas separation means comprises chemical catalysts that may be based on sorbents for carbon dioxide and configured to utilize gas separation procedure.

According to some embodiments, the gas separation means comprises biological enzymes  
5 configured to utilize gas separation procedure.

According to some embodiments, the aerial unit is a high-altitude balloon.

According to some embodiments, the aerial unit is configured to be tethered to the non-aerial unit.

According to some embodiments, the aerial unit further comprises self-steering means.

10 According to some embodiments, the aerial unit is configured to be retrofitted/integrated into the propulsion means to an aerial vehicle.

According to some embodiments, the storage means is at least one compressed gas container that may be configured to be released from the aerial unit and reach the non-aerial unit.

According to some embodiments, the non-aerial unit comprises a designated landing area  
15 configured to capture the at least one compressed gas container.

According to some embodiments, the at least one compressed gas container comprises guidance means configured to guide the at least one compressed gas container from the aerial unit to the non-aerial unit.

According to some embodiments, the non-aerial unit is configured to utilize the stored  
20 designated gaseous matter captured by the aerial unit.

According to some embodiments, the non-aerial unit is configured to be located on the ground, on a body of water or on a vessel, wherein a non-aerial unit configured to be located on a body of water may further comprise a docking area.

According to some embodiments, the controller is further configured to generate  
5 navigation commands in order to control the aerial unit.

According to some embodiments, the gas separation means further comprise an air compressor configured to increase air pressure flowing within to 6-10 Bar above the ambient air pressure.

According to some embodiments, the designated gaseous matter is carbon dioxide.

10 According to some embodiments, the airborne aerial unit is configured to exploit the low temperatures at high altitudes in order to liquefy or solidify the designated gaseous matter.

According to some embodiments, the at least one gas separation means carried by the airborne aerial unit is configured to exploit high altitude wind in order to harness an incoming airflow pressure for the purpose of gas separation.

15 According to some embodiments, the potential energy stored within the compressed air may be further utilized by the gaseous matter capture system.

According to some embodiments, the aerial unit is configured to capture carbon dioxide by using a phase transition process at a temperatures range of  $-100^{\circ}$  to  $-10^{\circ}$  and pressures range of 0.2-10 Bar.

According to some embodiments, the energy source is based on solar energy/ wind energy/ prestored power reservoir or configured to power the aerial unit by using a wired connection.

According to a second aspect, there is provided a method for gaseous matter capture using a gaseous matter capture system, comprising the steps of separating at least one designated gaseous matter from the air using at least one gas separation means carried by an airborne aerial unit, storing  
5 the at least one separated gaseous matter within storage means carried by the airborne aerial unit and transferring the stored separated gaseous matter to a non-aerial unit.

#### **BRIEF DESCRIPTION OF THE FIGURES**

Some embodiments of the invention are described herein with reference to the  
10 accompanying figures. The description, together with the figures, makes apparent to a person having ordinary skill in the art how some embodiments may be practiced. The figures are for the purpose of illustrative description and no attempt is made to show structural details of an embodiment in more detail than is necessary for a fundamental understanding of the invention.

In the Figures:

15 **FIG. 1** constitutes a schematic perspective view an aerial unit and non arial unit of a gaseous matter capture system, according to some embodiments of the invention.

**FIG. 2** constitutes a schematic perspective view of the non-aerial unit of a gaseous matter capture system, according to some embodiments of the invention.

**FIG. 3** constitutes a block diagram illustrating possible modules that form an aerial unit of  
20 a gaseous matter capture system, according to some embodiment of the invention.

**FIG. 4A** constitutes a typical phase diagram of CO<sub>2</sub> at various temperatures.

**FIG. 4B** constitutes a combo chart depicting sampled CO<sub>2</sub> concentrations in the ambient air at different altitudes.

**FIG. 5** constitutes a line graph illustrating the various temperatures and pressures having  
5 an effect on liquidation or solidification of CO<sub>2</sub>.

**FIGS. 6A & 6B** constitute a line graph illustrates the absorption capacities of sorbents for carbon capture having high affinity to CO<sub>2</sub>.

#### **DETAILED DESCRIPTION OF SOME EMBODIMENTS**

In the following detailed description, numerous specific details are set forth in order to  
10 provide a thorough understanding of the invention. However, it will be understood by those skilled in the art that the present invention may be practiced without these specific details. In other instances, well-known methods, procedures, and components, modules, units and/or circuits have not been described in detail so as not to obscure the invention. Some features or elements described with respect to one embodiment may be combined with features or elements described with respect  
15 to other embodiments. For the sake of clarity, discussion of same or similar features or elements may not be repeated.

Although embodiments of the invention are not limited in this regard, discussions utilizing terms such as, for example, “controlling” “processing,” “computing,” “calculating,” “determining,” “establishing”, “analyzing”, “checking”, “setting”, “receiving”, or the like, may  
20 refer to operation(s) and/or process(es) of a controller, a computer, a computing platform, a computing system, or other electronic computing device, that manipulates and/or transforms data

represented as physical (e.g., electronic) quantities within the computer's registers and/or memories into other data similarly represented as physical quantities within the computer's registers and/or memories or other information non-transitory storage medium that may store instructions to perform operations and/or processes.

5 Unless explicitly stated, the method embodiments described herein are not constrained to a particular order or sequence. Additionally, some of the described method embodiments or elements thereof can occur or be performed simultaneously, at the same point in time, or concurrently.

The term "Controller", as used herein, refers to any type of computing platform or  
10 component that may be provisioned with a Central Processing Unit (CPU) or microprocessors, and may be provisioned with several input/output (I/O) ports, for example, a general-purpose computer such as a personal computer, laptop, tablet, mobile cellular phone, controller chip, SoC or a cloud computing system.

The term "Sorbents for carbon capture", as used herein, refers to any material with higher  
15 affinity to CO<sub>2</sub> when compared to other atmospheric gases such as nitrogen and oxygen and, more specifically, to diverse range of typically porous, solid-phase materials, that can accommodate a wide variety of cations and may include mesoporous silicas, zeolites, metal-organic frameworks, etc., wherein these materials have the potential to selectively remove CO<sub>2</sub> from large volumes of air.

20 The term "Metal-organic framework (MOFs)" as used herein, refers to a class of porous compounds having low heat capacities and consisting of metal ions or clusters coordinated to organic ligands forming 1D, 2D, or 3D structures. Because of their small pore sizes and high

void fractions, MOFs are considered a promising potential material for use as an adsorbent to capture CO<sub>2</sub> and may provide an efficient alternative comparing to traditional amine solvent-based methods widely used today. CO<sub>2</sub> can bind to a MOF surface through either physisorption, which is caused by Van der Waals interactions, or chemisorption, which is caused by covalent bond formation. Once the MOF is saturated with CO<sub>2</sub>, the CO<sub>2</sub> would be removed from the MOF through either a temperature swing or a pressure swing (a process known as regeneration). In a temperature swing regeneration, the MOF would be heated until CO<sub>2</sub> desorbs. In a pressure swing, the pressure would be decreased until CO<sub>2</sub> desorbs.

The term “ZIFs”, as used herein, refers to a class of metal-organic frameworks (MOFs) that are topologically isomorphic with zeolites, wherein a particular MOF named ZIF-8, has a very high separation factor for hydrogen and carbon dioxide mixtures and a relatively high specificity of carbon dioxide over nitrogen. ZIF-8 is also known to be a relatively stable MOF, thus applicable over a wide range of temperatures and pressures. ZIFs are composed of tetrahedrally-coordinated transition metal ions (e.g. Fe, Co, Cu, Zn) connected by imidazolate linkers and having zeolite-like topologies.

The term “inversion layer” as used herein, refers to a layer in the atmosphere, or as a region in terms of altitude, in which temperatures tend to stop decreasing with any further increase in altitude. While thermal inversion can occur in multiple conditions, it is common to refer to the inversion layer as the altitude at which the vertical temperature gradient reverses, at the top of the troposphere, sometimes referred to as the tropopause.

According to some embodiments, the present invention discloses a gaseous matter capture system comprising a light weight aerial unit configured to be released to the atmosphere and

comprising at least one gas separation means such as a compressor. According to some embodiments, said aerial unit may be a high-altitude balloon and may further comprise controlling, navigation and steering components. According to some embodiments, said aerial unit is configured to capture large amounts of gaseous matter such as carbon dioxide and to throw it  
5 downwards to a designated non-aerial unit where it may be safely caught. According to some embodiments, the aerial unit may be a high-altitude balloon configured for throwing large quantities of captured carbon dioxide in order to allow capturing more carbon dioxide mass in a single airborne mission hence reducing regular maintenance and ground time intervals.

According to some embodiments, the separation of the gaseous matter such as carbon  
10 dioxide from the ambient air can be done using multiple techniques and methods such as: coolers, refrigerators, freezers, heat pumps, pressure pumps, compressors, membranes, separation using chemical means or catalysts, separation by biological enzymes, etc. Such techniques and methods are used to increase the rate at which CO<sub>2</sub> is captured from the atmospheric ambient air.

According to some embodiments, the gaseous matter capture system is configured to use  
15 the low ambient temperatures and high winds that circulate the surrounding atmosphere. The gaseous matter capture system is further configured to prevent dilution of the carbon dioxide in the incoming air stream since that, although the air density is lower in the high troposphere and the lower stratosphere, the volumetric concentration of carbon dioxide is not significantly lower and is almost similar to the levels found at sea level. It is thus applicable to utilize the above  
20 described carbon capture method at high altitudes above sea level.

According to some embodiments, the gaseous matter capture system is configured to collect a gaseous matter directly from the atmosphere (wherein a preferable gas to be collected

may be carbon dioxide), for purposes of climate change mitigation and gaseous matter reuse. According to some embodiments, the system may be based on carbon dioxide's phase transition at low temperatures, such as in ranges of -100°C to -10°C and increased pressure ranges, such as between 0.2 Bar to 10 Bar.

5           According to some embodiments, in order to remove massive amounts of CO<sub>2</sub> directly from the ambient air, without the need for excessive energy input, without the use of dangerous or scarce resources, and in a fully scalable manner, the use of high altitude vehicles equipped with compressors is suggested.

          According to some embodiments, once carbon dioxide has been separated from the air  
10   flow, it can be stored or utilized further in accordance to various needs or constraints. For example, a separated carbon dioxide may be liquefied/solidified and kept in storage means such as high-pressure containers, wherein said containers can be made of any known material, such as composite carbon fibers, aluminum, polymers etc.

          Reference is now made to **FIG. 1** which schematically illustrates an aerial unit **100** and  
15   non aerial unit **200** of gaseous matter capture system **10**. As shown, gaseous matter capture system **10** may comprise two main units, aerial unit **100** and non-aerial unit **200**. Aerial unit **100** can be, for example, a high-altitude balloon or any other airborne vehicle, configured to be flown at high altitudes, such as altitudes of 5-15km, wherein the standard temperatures at these altitudes are typically around -50°C and the air density is approximately 10-30% of those found at sea level.

20           According to some embodiments, a high-altitude balloon that operates as an aerial unit **100**, may be filled with Helium, Hydrogen gas, hot air or any other known substance used to provide

aerial lift. According to some embodiments, Aerial unit **100** may be tethered or untethered to the non-aerial unit **200**.

According to some embodiments, aerial unit **100** may be any known aerial vehicle or platform, for example, a powered aircraft (either by internal combustion engine, jet propulsion, solar power or electrical power), a gliding aircraft (such as kite, glider etc.) or an aerostat (such as an airship, balloon, etc.) According to some embodiments, aerial unit **100** may be implemented on an existing aerial vehicle, for example, aerial unit **100** may be retrofitted to a commercial aviation plane to be carried upon or implemented with any section of its fuselage, wings or engines. An aerial unit **100** retrofitted upon an aerial vehicle may further rely on already existing systems, for example, it may use an aircraft's engine built-in compressor as a substitute to an integrated gas separation means **102** (disclosed below).

According to some embodiments, gas separation means **102** may comprise a compressor, pump or any known pressure increasing device configured to be carried by the aerial unit **100** and compress the surrounding air at high flow rates to pressures that are approximately 5-10 Bar above the ambient air pressure. According to some embodiments, and as mentioned above, the separation of carbon dioxide from the ambient air by the gas separation means **102** can be achieved by using various techniques and/or methods such as: refrigeration, heat pumping, multiple pumps/compressors, membranal separation, separation using chemical means such as catalysts, separation by biological enzymes, etc. Such techniques and methods may be used to increase the rate at which carbon is captured from the atmospheric ambient air, reduce the required energy for CO2 capturing, etc.

According to some embodiments, and since strong wind is generally abundant at high altitude, the high dynamic pressure caused by said high altitude strong wind in which the aerial unit **100** designed to operate, may be exploited for the purpose of separating CO<sub>2</sub> from an airflow entering separation means **102**. For example, gas separation means **102** located on an aerial unit

5 **100** may use different types of membranes in order to filter the incoming airflow and produce filtered air having an increased CO<sub>2</sub> concentration with regard to Nitrogen or other gases' concentration. Said increased CO<sub>2</sub> concentration may be high over an order of magnitude (x10) with regard to the CO<sub>2</sub> concentration in the ambient airflow, hence, harnessing the wind for the sake of said gaseous matter separation may significantly increase the efficiency of said process by

10 reducing the need to compress the incoming airflow.

According to some embodiments, since the gas separation means **102** of aerial unit **100** are configured to operate at high altitude, much of the energy generally used to compress ambient air at ground altitude is generally unneeded. According to some embodiments, after being separated from its CO<sub>2</sub>, the compressed air may be further utilized by using its potential stored energy. For

15 example, the potential energy stored within the compressed air may be used directly to compress further airflow or indirectly to power various electrical/mechanical systems, thus leading to further energy/weight savings.

According to some embodiments, in order to increase carbon dioxide capture efficiency, the aerial unit **100** may utilize multiple gas separation means **102** in parallel, for example, gas

20 separation means **102** may be multiple compressors aliened in series/in several stages in order to provide an efficient compression and separation of gaseous matter.

According to some embodiments, designated substances may be used for CO<sub>2</sub> capture or CO<sub>2</sub> separation and may be implemented within gas separation means **102** in order to increase the amount of CO<sub>2</sub> available for separation. Said substances may be, for example, MOFs, ZIFs, or any other known sorbents for carbon capture and may be arranged in the form of thin films and  
5 produced by various chemical processes such as nano deposition techniques such as ALD, CVD, PVD, Glancing Angle Deposition, etc.

According to some embodiments, at least one storage means **106**, for example, a compressed gas tank/s, is/are configured to store the separated gaseous matter such as CO<sub>2</sub> in a liquid, solid or gas form after it has been extracted from the compressed air flow.

10 According to some embodiments, in order to maximize the CO<sub>2</sub> extraction efficiency by aerial unit **100** and to allow maximum collection of gaseous matter given limited resources and constrains and in order to allow for more airborne time before the need of aerial unit **100** to land, the gaseous matter capture system **10** may be configured to release full gas tanks designed to safely land on a designated non-aerial unit/s or on predefined ground/watery locations for further  
15 utilization.

According to some embodiments, the gaseous matter capture system **10** further comprises a designated mechanism (not shown) configured to enable a controlled release of at least one storage means **106** that may be a compressed gas tank. For example, the controlled release means may be configured to disconnect storage means **106** after it has been filled with gaseous matter in  
20 order to eliminate excess weight from the airborne vehicle. According to some embodiments, storage means **106** is configured to be released and fall downwards to a pre-designated non-aerial unit, where it may be safely caught and collected.

According to some embodiments, storage means **106** may be a free-falling tank/s, configured to be thrown by/released from the aerial unit **100**, using the aforementioned designated release mechanism, or alternatively, fall in a guided manner. Storage means **106** may include utilization of parachutes, gliding wings, propellers, gas injections or jet thrusters in order to provide  
5 trajectory correction ability or any other known steering/navigation means.

According to some embodiments, controller **104** is further configured to provide general operational control of the gaseous matter capture system **10**. According to some embodiments, controller **104** may be positioned upon the aerial unit **100**, upon the non-aerial unit **200**, or may be located elsewhere, for example, on a remote server or as part of cloud computing platform.  
10 According to some embodiments, controller **104** is configured to provide navigation control to aerial unit **100**, wherein said navigation control may be conducted automatically or manually by a user monitoring the operation of the gaseous matter capture system **10**.

According to some embodiments, the aerial unit **100** may further comprise propulsive/steering means (not shown) that can be any known propulsive component configured  
15 to provide a controlled aerial deployment of the aerial unit **100**. According to some embodiments, controller **104** may control the propulsive/steering means that may be jet thrusters, rocket propulsion, flaps, propeller of any sort or any other known means of propulsion.

According to some embodiments, the gaseous matter capture system **10** further comprises communications means (not shown) configured to provide a reliable and fast communication track  
20 between the aerial unit **100** and the non-aerial unit **200**. For example, a communication system that may be controlled by the controller **104** may provide navigation commands to the aerial unit in

accordance with various needs or restrains and may be operated either automatically or manually by a user monitoring the operation of the gaseous matter capture system **10**.

According to some embodiments, the gaseous matter capture system **10** further comprises an energy source **108** that may be a power reservoir/battery, a hydrogen reservoir (that may simultaneously be used for lift purposes), solar panels/paints/sheets, wind turbines (in order to take advantage of the surrounding strong wind), nuclear power generators, thermal-nuclear power sources in conjunction with thermoelectric elements, etc. According to some embodiments, a tethered wire connected to the ground, the non-aerial unit **200** or to another airborne vehicle may provide the energy needed for the operation of aerial unit **100**. According to some embodiments, the energy sources used to provide power to the gaseous matter capture system **10** are configured to be carbon neutral or close to it, in order not to contradict the main purpose of carbon dioxide extraction.

According to some embodiments, aerial unit **100** may be configured to be deployed in a relative position that has the ability to provide constant or near constant energetical availability, or has the ability to provide aerial unit **100** with improved gaseous matter capturing conditions. this can be achieved by adaptively changing the altitude/position of aerial unit **100** in order to utilize different wind directions or solar radiation conditions. According to some embodiments, in order to change the relative deployment of aerial unit **100**, propulsion and/or navigation and steering means may be used as previously disclosed.

According to some embodiments, the final product of the gaseous matter capture system **10** may be high purity carbon dioxide intended for either storage or reuse in applications such as agriculture, food industry, research, synthetic near emission neutral fuels manufacturing, etc.

Reference is now made to **FIG. 2** which schematically illustrates a non-aerial unit **200** of gaseous matter capture system **10** (previously disclosed). As shown, a designated landing area **202** that may be any kind of capturing platform (such as a trampoline), is configured to provide a safe area for high velocity, large mass falls. The fast falling storage means **106** may land upon the  
5 landing area **202** to be further collected by any mechanical, robotic or manual means (not shown).

According to some embodiments, the non-aerial unit **200** may include a designated facility **204** configured to provide either maintenance requirements for aerial unit **100** and/or processing of the at least one storage means **106** after it has been loaded with gaseous matter and captured by the landing area **202**.

10 According to some embodiments, the non-aerial unit **200** may be located on the ground or on a body of water floating platform, alternatively, the non-aerial unit **200** may be located upon a moving platform such as any kind of marine vessel or terrestrial vehicle. According to some embodiments, After the capturing of storage means **106** by the landing area **202** and the delivery of storage means **106** to the designated facility **204**, industrial procedures that may be either  
15 chemical and/or mechanical may utilize the collected gaseous matter for further storage or use.

According to some embodiments, non-aerial unit **200** may further comprise docking area **206** configured for either marine vessels or terrestrial vehicles, in order to enable the transfer of the captured gaseous matter to another location.

Reference is now made to **FIG. 3** which illustrates a block diagram disclosing possible  
20 modules that form aerial unit **100**. As shown, energy module **300** may be a power reservoir such as a battery, hydrogen-based fuel cell, solar panel/paint/sheet, wind turbine, nuclear generator or

any other power source that does not emit greenhouse gasses or generate reduced levels of greenhouse gasses.

According to some embodiments, control module **302** is further configured to provide general operational control of the gaseous matter capture system **10** and may comprise a controller  
5 positioned upon the aerial unit **100**, upon the non-aerial unit **200** (not shown), or may be located elsewhere, for example, on a remote server or as part of cloud computing platform. According to some embodiments, control module **302** is configured to generate navigation/steering commands in order to control aerial unit **100**, wherein said navigation/steering control may be conducted automatically or manually by a user monitoring the operation of the gaseous matter capture system  
10 **10**. According to some embodiments, control module **302** is configured to monitor the various parameters and operations that are part of the gaseous matter capture system **10**'s activity.

According to some embodiments, gas separation module **304** is configured to enable the separation of gaseous matter such as carbon dioxide from the ambient air which may be done using multiple techniques and methods such as: single/multiple pumps or compressors (or any known  
15 pressure increasing device), membranes, separation using chemical means or catalysts, separation by biological enzymes, etc.

According to some embodiments, gaseous matter storage module **306** is configured to store a liquefied/solid or high pressurized gaseous matter such as carbon dioxide in at least one high-pressure container (storage means **106** previously disclosed), wherein said container may be  
20 a free-falling container configured to be thrown by/released from the aerial unit **100**, using the aforementioned designated release mechanism, or, alternatively, fall in a guided manner. According to some embodiments, said high-pressure container may include utilization of

parachutes, gliding wings, propellers, gas injections, jet thrusters or any other known steering/navigation means in order to provide trajectory corrections ability.

According to some embodiments, navigation module **308** is configured to provide navigation abilities to the aerial unit **100** and may comprise designated navigation components  
5 such as GPS, altitude/velocity sensors, etc. in order to determine the exact location, height and relative position of the aerial unit **100**. The navigation module **308** may further utilize a database regarding the wind regime at a certain location and altitude in order to adapt aerial unit **100**'s operation to changing weather conditions. According to some embodiments, navigation module **308** may be a separated module of may be integrated within control module **302**.

10 According to some embodiments, propulsion module **310** is configured to propel the aerial unit **100** to a desired location/altitude, wherein the propulsion of aerial unit **100** may be conducted using jet thrusters, rocket propulsion, flaps, propellers of any sort or any other known means of propulsion.

Reference is now made to **FIG. 4A** which illustrates a phase diagram depicting a carbon  
15 dioxide's phase transition at relatively low temperatures and relatively high pressures.

According to some embodiments, in order to liquefy the gas at a temperature of approximately -55°C, the required pressure should be approximately 6 atmospheres. Due to the said relatively high liquefaction pressure, gas liquification performed at high altitude may represent a challenge for a carbon capture procedure. As shown, the separation of carbon dioxide  
20 can be done by liquefaction or solidification. According to some embodiments, reaching any point below the triple point temperature of approximately -56°C and having sufficiently high pressures, will result in carbon dioxide solidification. Conversely, increasing the pressure at temperatures

above the triple point, for example by applying pressures of 6-10 Bar, will result in the liquefaction of carbon dioxide.

According to some embodiments, any combination of temperature and pressure within the limits that enable carbon dioxide liquefaction/solidification may be used during the operation of  
5 gaseous matter capture system **10**.

According to some embodiments, since in high altitude the air density is approximately one third of the ambient air at sea-level, at least one stage of compression may be used in order to increase the incoming air pressure from said typical pressure at high altitude (0.2-0.8 Bar) in order to reach the range enabling carbon dioxide solidification/liquefaction. The required increase in  
10 pressure may be multiplied by a ratio of 3 to 50 at the end of all compression stages in order to reach said typical transition pressure of 6-10 Bar.

According to some embodiments, when said pressure levels are reached, and given an appropriate temperature (approximately -55°C), a liquefaction of the CO<sub>2</sub> contained within the processed incoming air is likely to follow. Alternatively, further reducing the temperature by 40°C  
15 -70°C while exposing the incoming air to sufficiently high pressure will mainly result in solidification of the CO<sub>2</sub> contained within the incoming air.

According to some embodiments, pressure or temperature changes may be done during a single compression/separation stage or during multiple stages. According to some embodiments, when using multiple compression/separation stages as part of the process of carbon dioxide  
20 capturing disclosed above, the coefficient of performance in cooling that represents some of the energetic efficiency, may be higher.

Reference is now made to **FIG. 4B** which illustrates a combo chart depicting sampled CO<sub>2</sub> concentrations in the ambient air at different altitudes (9-10km) and various temperatures. As shown, the volumetric concentration of CO<sub>2</sub> remains almost the same at high altitudes, namely slightly under 400 ppm when compared to 411 at sea level. This may be due to the fact that the strong winds provide constant airflow and prevent dilution of CO<sub>2</sub>.

According to some embodiments, the high altitude that aerial unit **100** is configured to be located at, represents a tradeoff between low temperatures that allow the use of less energy consumption to reach CO<sub>2</sub> phase transition, along with low overall pressures, which increase the energy consumption needed to reach a desired pressure in order to achieve CO<sub>2</sub> phase transition by either liquefaction or solidification. As noted above, the CO<sub>2</sub> concentration within the ambient air at different altitudes remains approximately similar, and does not significantly change the efficiency of the CO<sub>2</sub> separation process.

According to some embodiments, and as previously shown, compressing the incoming air entering the separation means **102** located on aerial unit **100** at high altitude in order to reach pressures of 6-10 bar results in CO<sub>2</sub> separating from the incoming airflow and stored in a liquid/solid form inside storage means **106**. Considering the gas molar volumetric concentration (22.4 liters per mole) in standard conditions, in high altitude where the temperatures range is approximately -50°C and the pressure is about 0.25 bar, the gas molar volumetric concentration is ~70 liters per mole (22.4 L/mole, multiplied by (223°C / 273°C) for the temperature correction and by (101.3KPa/26.5KPa) for the pressure correction results in 70 L/mole).

According to some embodiments, at a concentration of 400 ppm CO<sub>2</sub>, and given a molar mass of 44 g/mole, there is a need to compress  $70 \times 2500$  (2500 being the reciprocal of 400 ppm)

liters to get to one mole, (or 44 grams). According to the above, in order to produce one ton of CO<sub>2</sub>, there is a need to compress approximately 4 million cubic meters of ambient air.

Nowadays, simple and inexpensive compressors reach flow rates measured at several cubic feet per minute, or several liters per second. According to that, in order to reach a level of a metric  
5 ton per day would require several compressors in parallel.

According to some embodiments, it is possible to compress the air travelling at a typical high-altitude wind of 100km/h through an orifice having a certain diameter, in order to reach high capacity, for example, 4 million m<sup>3</sup> of flow rate within less than a day.

According to some embodiments, in order to introduce an overall solution to climate  
10 change, and assuming that each aerial unit **100** is able handling around one metric ton of CO<sub>2</sub> captured per day, a hypothetical gaseous matter capture system **10** will need to be comprised of 54 million aerial units **100** in order to capture all excess CO<sub>2</sub> introduced into the atmosphere in 2018. Considering that each aerial unit **100** has an annual price tag of \$100K, the complete solution of an annual global CO<sub>2</sub> emissions capture will cost around \$5,000B. this estimation is not only far  
15 lower than any known alternative, but significantly lower than the expected economic damage associated with climate change. For reference, the International Panel on Climate Change (IPCC) stated that it needs \$13,000B to reverse the increasing trend of carbon emissions and lower it by 10 Billion Tons of CO<sub>2</sub>. Meaning over twice the cost for less than half the result.

Reference is now made to **FIG. 5** which illustrates a line graph depicting temperature  
20 differences affecting phase transition of CO<sub>2</sub> based on standard US atmospheric data. More particularly, the line graph represents the difference in temperatures affecting the initiation of

phase transition of captured CO<sub>2</sub> in accordance with the decrease in temperatures associated with increased altitude and lower/higher pressures.

According to some embodiments, at the inversion layer, (not shown, typically at altitudes of 11-13 kilometers above sea level, but potentially varying beyond these numbers in equatorial or polar regions), the required temperature difference decreases along all 3 lines, and hence the inversion layer represents the ideal altitude for performing high altitude gas separation. According to some embodiments, the ambient temperature at said inversion layer is approximately -50°C, and the ambient pressure is approximately 0.3 Bar. This means that to separate CO<sub>2</sub> by solidifying it, a reduction of approximately 40°C is required as seen in **FIG. 4** showing the freezing point of CO<sub>2</sub> is approximately -90°C at 0.3 Bar.

According to some embodiments, and as part of the assumption that the ideal height for a gas separation process is at the inversion layer or its vicinity, there are some additional parameters that may be taken into account. Due to the increased pressure of the ambient air found in said inversion layer, various methods of gaseous matter separation using physical and chemical separation procedures may be enabled. For example, the use of molecular sieves such as Metal Organic Frameworks MOF, Zeolites or other designated substances may benefit in terms of gas separation efficiency from higher pressures occurring at the inversion layer or its vicinity.

According to some embodiments, the temperature differences required at a given altitude in order to perform gas separation procedure according to the standard US atmosphere data are disclosed:

According to some embodiments, an estimated calculation of either cooling the air to CO<sub>2</sub> freezing temperature or compressing it to liquid/solid form is provided as follows: Cooling

incoming air as part of the operation of the gaseous matter capture system **10** requires low air temperature with sufficient CO<sub>2</sub> mass. For example, To cool  $m_{CO_2} = 100Kg$ , by approximately 40°K lower than the ambient air temperature, will require  $M_{tot} \cdot C_{air} \cdot (40^\circ K)$ , where  $C_{air} \sim 0.71KJ/KG$ , and thus the energy cost is  $7 \cdot 10^6KJ$ , or in a period or over a 12 hours day, 5 160KW. In addition, the enthalpy of sublimation (latent heat) is  $\sim 590KJ/Kg$ , meaning that an additional 1.5KW are required for the phase transition, though this may be considered negligible compared to the total required cooling energy.

According to the second law of thermodynamics, the cooling efficiency is limited such that in order to remove  $\sim 160KW$  of heat from the certain mass of air, and giving an ideal coefficient of performance of  $CoP = \frac{T_c}{T_h - T_c} \cong 4.5$  (given a freezing temperature of -90°C, or 183°K, and a 10 starting temperature of -50°C, or 223°K, put into the equation as  $T_c$  for the lower temperature and  $T_h$  for the higher temperature), the suggested calculation may, for every watt used for heat removal, remove approximately 4.5W from the cooled air by a single stage cooling at a maximal temperature difference, wherein multiple cooling stages will increase the efficiency of said cooling process in accordance with a decrease in temperature differences. Assuming that the average 15 multiple stage cooling requires approximately 2W of power, the result is a required energy of  $\sim 80KW$ .

According to some embodiments, the above-mentioned power requirements may be achieved by harnessing the surface area of areal unit **100**. For example, and using a technology of 20 commercially available plastic based solar energy films, a power generation density of 100Wp/m<sup>2</sup> may be achieved by an aerial unit **100** having approximately 10-meter radius and approximately 200m<sup>2</sup> of surface area available at a given moment for power generation. Said surface area covered

with solar energy films may produce approximately 40KW of power that represents marginally sufficient energy at very high COPs (coefficient of performance).

According to some embodiments, the air processed by the areal unit **100** may be cooled down to the freezing temperature of CO<sub>2</sub> and the captured CO<sub>2</sub> may be stored in storage means

5 **106**. CO<sub>2</sub> may be frozen by lowering its temperature to a range of -80°C to -100°C depending on the surrounding air pressure. Since the ambient temperature at high altitudes where the areal unit **100** is configured to operate is around -50°C, the potential cooling component of areal unit **100** can work typically at a Carnot efficiency of ~3.5 to ~6.4 [ $COP < T_c / (T_h - T_c)$ ] when the air pressure is lower and resulting phase transition is at -100°C (thus putting the numbers in units of °K into

10 the equation for COP, as 173°K for the lower temperature and 223°K for the higher temperature), the resulting COP is  $-173 / (223 - 173) = 3.46$ . and when the air pressure is higher and resulting required temperature is -80°C or 193°K  $-193 / (223 - 193) = 6.43$ ].

According to some embodiments, in order to overcome the latent sublimation heat that may be produced as part of the air-cooling procedure, there is a need to invest at least 200J per

15 each gram of CO<sub>2</sub>, meaning that for a rate of one gram per second, there is a need to remove 200W of heat in order to overcome the produced latent heat. In accordance to the above, the upper limit of CO<sub>2</sub> freezing rate (given a 1KW power input) would be 17.5g per second (756Kg over a 12 hours' period), neglecting all other power needs. When dismissing 20% of this approximate efficiency as a reasonable assumption, we get that a 1KW of power input is sufficient to allow

20 phase transition of approximately 3 grams CO<sub>2</sub> per second.

According to some embodiments, further capturing techniques using substances having high affinity to CO<sub>2</sub>, that include, among others, Metal Organic Frameworks (MOFs) may be used

as part of the operation of areal unit **100**. Using such techniques will increase the tendency of CO<sub>2</sub> to nucleate on them and require less air volume to be cooled.

Reference is now made to **FIGS. 6A & 6B** which illustrates a line graph depicting absorption capacity of sorbents for carbon capture with high affinity to CO<sub>2</sub> according to some embodiments of the invention. As shown, **FIG. 6A** depicts the adsorption capacity and absorption dynamics of CO<sub>2</sub> in Zeolites with regard to various pressures. **FIG. 6B** depicts a particular MOF compound (ZIF-8) absorption property of CO<sub>2</sub> at different temperatures and pressures. The results depicted on **FIGS. 6A & 6B** suggest that when using the aforementioned materials, the effect of decreasing temperatures is more prominent than variable pressure values with regard to the  
10 aforementioned substances' absorbing efficiency of CO<sub>2</sub>. In other words, the aforementioned substances ability to absorb CO<sub>2</sub> increases dramatically when the temperature drops down to the inversion layer average temperature, as previously disclosed.

According to some embodiments, by using other substances and in other circumstances, various pressure swings required to adsorb and desorb a certain separated gas may be more prominent. In such cases and in others, deploying areal unit **100** at lower altitude with increased  
15 ambient temperature and higher air pressure, or alternatively, deploying areal unit **100** at higher altitude, typically with similar temperature yet with lower air pressure, may have a critical effect on the absorption efficiency of the gas separation process conducted by the areal unit **100**. Hence, for some applications and embodiments, the gas separation process may be conducted at ambient  
20 temperature as high as approximately -10°C and at a typical altitude of approximately 5 km above sea level. Alternatively, and according to some embodiments, in higher altitude of approximately 15 km, the pressure may drop to approximately 0.2 Bar while still allow a gas separation process

with a benefit of decreasing the required energy for pressure reduction as part of desorbing processes.

Although the present invention has been described with reference to specific embodiments, this description is not meant to be construed in a limited sense. Various modifications of the disclosed embodiments, as well as alternative embodiments of the invention will become apparent  
5 to persons skilled in the art upon reference to the description of the invention. It is, therefore, contemplated that the appended claims will cover such modifications that fall within the scope of the invention.