

SILICON-GERMANIUM HYDRIDES AND METHODS FOR MAKING AND USING SAME

5 Cross Reference

This application claims priority to U.S. Provisional Patent Applications SN 60/739,528 filed November 23, 2005 and SN 60/764,128 filed February 1, 2006, both of which are hereby incorporated by reference in its entirety.

10 Statement of Government Interest

This work was supported by National Science Foundation grant DMR-0221993 thus the U.S. government may have certain rights in the invention.

Field of the Invention

15 The present application relates generally to the fields of chemistry, optoelectronics, silicon-germanium alloys, semiconductor structures, and related methods and devices.

Background

20 Heteroepitaxial $\text{Ge}_{1-x}\text{Si}_x$ layers on Si(100) with high Ge contents are of interest due to many applications in key technologies such as solar cells, MEMS, quantum cascade lasers and Si-based photonics^{1,2} including high speed modulators and photodetectors.³ However, these materials are much less developed in spite of this high impact potential in IR optical devices. In addition, they serve as virtual
25 substrates for growth of high mobility, strained Si and Ge device channels, and are also considered as a potential pathway to monolithic integration of III-V based devices with Si technologies.^{4,5} The best current route to these materials is complicated and fraught with difficulties, requiring both high temperature growth of thick ($>10\text{ }\mu\text{m}$) compositionally graded films and a chemical-mechanical
30 planarization step to relieve the misfit strain between the $\text{Ge}_{1-x}\text{Si}_x$ epilayer and Si substrate and produce a flat surface, respectively.⁶

To circumvent these difficulties we have recently developed new chemical vapor deposition (CVD) heteroepitaxy to produce device quality strain-relieved Ge-

rich films on Si without the need for graded compositions.^{7,8} Our strategy involves depositions of single source hydride precursors with direct Si-Ge bonds such as the $(\text{H}_3\text{Ge})_x\text{SiH}_{4-x}$ ($x=1-4$) family of compounds. These are routinely synthesized with semiconductor grade purity via straightforward methodologies utilizing commercially available starting materials.⁷ Their physical and chemical properties such as high volatility and facile reactivity make them particularly useful as reagents in low temperature film growth by CVD and gas source molecular beam epitaxy (GSMBE). Our initial growth experiments have yielded single crystal films with compositions $\text{Si}_{0.50}\text{Ge}_{0.50}$, $\text{Si}_{0.33}\text{Ge}_{0.67}$, $\text{Si}_{0.25}\text{Ge}_{0.75}$ and $\text{Si}_{0.20}\text{Ge}_{0.80}$, precisely matching those of the corresponding precursors SiH_3GeH_3 , $(\text{H}_3\text{Ge})_2\text{SiH}_2$, $(\text{H}_3\text{Ge})_3\text{SiH}$ and $(\text{H}_3\text{Ge})_4\text{Si}$.⁸ Our results demonstrate that exact control of the composition, structure and strain at the atomic level can be achieved via incorporation of the *entire* Si-Ge framework of the precursors into the film. The resulting films are of much higher quality than those previously reported using conventional sources and grow at much lower temperatures that are compatible with CMOS technology.

Within this compositional range the $\text{Si}_{0.50}\text{Ge}_{0.50}$ semiconductor system is of particular importance because it possesses the ideal lattice dimensions to integrate fully strained Si channels exhibiting high electron mobilities into metal oxide silicon field effect transistors (MOSFETS). Thus these materials are ideally suited for immediate technological applications in high-speed devices. However, in our experiments we have observed that stoichiometric $\text{Si}_{0.50}\text{Ge}_{0.50}$, with the required device quality properties (surface planarity, low defect densities), can only be grown near a maximum temperature of $\sim 450^\circ\text{C}$ at a growth rate of only 0.2 nm/min which is too low to be practical for high throughput device fabrication. Negligible film growth was observed at lower temperatures via H_3GeSiH_3 , while films with rough surfaces and high dislocation densities were obtained above $\sim 500^\circ\text{C}$ due to the thermal mismatch with the substrate.

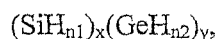
The synthesis of new compounds with increased reactivities and improved properties for device fabrication relative to H_3GeSiH_3 would be of great value in the art.

In this study we synthesize and fully characterize the butane-like $(\text{GeH}_3)_2(\text{SiH}_2)_2$ (1) compound containing highly reactive $-\text{H}_2\text{Si}-\text{SiH}_2-$ functionalities (rather than SiH_3). The $\text{H}_2\text{Si}=\text{SiH}_2$ dimeric unit has been identified as the dominant reactive intermediate in high temperature CVD via decomposition of conventional Si-

hydrides. These species lead to highly efficient growth of homoepitaxial Si.⁹ Precursor (1) is specifically designed to incorporate these necessary -H₂Si-SiH₂- building blocks and provide a new low temperature route for development of a broad range of Si-containing materials. Here we demonstrate the utility of (1) to produce device quality Si_{0.50}Ge_{0.50} alloys at unprecedented low temperatures of 350 °C. A 10-fold improvement in the growth rate (~ 2.0 nm/min) is achieved at 350 °C relative to growth (0.2 nm/minute) at 450 °C via SiH₃GeH₃ leading to uniformly stressed or strain-relaxed thin films directly on Si with planar surfaces and remarkably reduced dislocation densities. En route to (1) we have produced the isobutane analog (GeH₃)₂SiH(SiH₃) (2) and have also isolated and fully characterized the Ge-rich derivative (GeH₃)₂(SiH₂GeH₂) (3). The latter is a moderately volatile liquid, with potential applications in low temperature synthesis of Ge-rich optoelectronic alloys.

Summary of the invention

In one aspect, the present invention provides compounds comprising a silicon-germanium hydride of formula I:



wherein x is 2, 3 or 4;

wherein y is 1, 2, or 3;

wherein x + y is 3, 4, or 5;

wherein n1 is independently 0, 1, 2, or 3 for each Si atom in the compound, to satisfy valency; and

wherein n2 is independently 0, 1, 2, or 3 for each Ge atom in the compound, to satisfy valency;

with the proviso that n2 is not 0 when y is 1;

with the further proviso that when x is 3 and y is 1, then n2 is 2 or 3; and

with the further proviso that when x is 2 and y is 1, then n2 is 3.

In a second aspect, the present invention provides semiconductor structures, comprising a substrate and an Si-Ge layer formed by introducing near the surface of a substrate a gaseous precursor comprising one or more compounds according to formula I.

In further aspect, the invention provides semiconductor structures, comprising a substrate and an Si-Ge layer comprising a backbone of one or more compounds according to formula I.

5 In a third aspect, the present invention provides methods for synthesizing one or more silicon-germanium hydrides of formula I, comprising combining a compound selected from the group consisting of a nonafluorobutane-sulfonic substituted disilane and a triflate substituted disilane, with a compound comprising a GeH_3 ligand under conditions whereby the silicon-germanium hydride is formed.

10 In a fourth aspect, the present invention provides methods for depositing a Si-Ge material on a substrate in a reaction chamber, comprising introducing into the chamber a gaseous precursor comprising one or more compounds according to formula I, under conditions whereby a layer comprising a Si-Ge material is formed on the substrate.

15 In a fifth aspect, the present invention provides methods for depositing an epitaxial Si-Ge layer on a substrate, comprising introducing at or near a surface of the substrate a gaseous precursor comprising one or more compounds of formula I, and dehydrogenating the precursor under conditions whereby epitaxial Si-Ge is formed on the substrate.

20 In a sixth aspect, the present invention provides a compound comprising $(\text{SO}_3\text{C}_4\text{F}_9)_2(\text{SiH}_2)_2$.

In a seventh aspect, the present invention provides methods for synthesizing $(\text{SO}_3\text{C}_4\text{F}_9)_2(\text{SiH}_2)_2$ comprising combining $(p\text{-tolyl})_2(\text{SiH}_2)_2$ and $\text{HSO}_3\text{C}_4\text{F}_9$ under conditions whereby the $(\text{SO}_3\text{C}_4\text{F}_9)_2(\text{SiH}_2)_2$ is formed.

25 Description of the Figures

Figure 1: Molecular structure of the conformational isomer pairs $\{n\text{-GeSiSiGe}, g\text{-GeSiSiGe}\}$, $\{n\text{-GeSiGeGe}, g\text{-GeSiGeGe}\}$ and positional isomer $i\text{-Si}(\text{GeSiSi})$.

30 *Figure 2:* Calculated IR spectra of $n\text{-GeSiSiGe}$, $g\text{-GeSiSiGe}$, and $i\text{-Si}(\text{GeSiSi})$. The low- and high- frequency range is shown in panels $\{(a),(b),(c)\}$ and $\{(d),(e),(f)\}$, respectively. Empirical frequency scaling has not been applied to the spectra shown here. Individual spectral features of the $n\text{-GeSiSiGe}$, $g\text{-GeSiSiGe}$ and $i\text{-Si}(\text{GeSiSi})$ molecules are labeled with “n”, “g” and “i”, respectively. Corresponding molecular structures are drawn as insets in the low-frequency plots.

Figure 3: Calculated IR spectra of *n*-GeSiGeGe and *g*-GeSiGeGe. The low- and high- frequency range is shown in panels {(a),(b)} and {(c),(d)}, respectively. The spectral features of the *n*-GeSiGeGe and *g*-GeSiGeGe isomers are labeled “N” and “G”, respectively. Empirical frequency scaling has not been applied to the spectra shown here. Corresponding molecular structures are drawn as insets in the low-frequency plots.

Figure 4: Comparison of an experimental spectrum of $(\text{GeH}_3)_2(\text{SiH}_2)_2$ (**1**) with a superposition of theoretical spectra for the {*n*-GeSiSiGe, *g*-GeSiSiGe} combination: (a) low-frequency region (500-1100 cm^{-1}) and (b) high-frequency Si-H/Ge-H region (1900-2300 cm^{-1}). Frequency scale factors of 0.984 and 0.975 have been applied to the theoretical spectra in low- and high-frequency regions, respectively. The theoretical spectra consist of an admixture of 39% *n*-GeSiSiGe and 61% *g*-GeSiSiGe.

Figure 5: Comparison of an experimental spectrum of a mixture $(\text{GeH}_3)_2(\text{SiH}_2)_2$ (**1**) with a superposition of theoretical spectra for the {*i*-Si(GeSiSi), *g*-GeSiSiGe} combination: (a) low-frequency region (500-1100 cm^{-1}) and (b) high-frequency Si-H/Ge-H region (1900-2300 cm^{-1}). Frequency scale factors of 0.985 and 0.976 have been applied to the theoretical spectra in low- and high-frequency regions, respectively. The theoretical spectra consist of an admixture of 68% *i*-Si(GeSiSi) and 32% *g*-GeSiSiGe.

Figure 6: Comparison of an experimental spectrum of $(\text{GeH}_3)_2(\text{SiH}_2\text{GeH}_2)$ (**3**) with a superposition of theoretical spectra for the {*n*-GeSiGeGe, *g*-GeSiGeGe} combination: (a) low-frequency region (500-1100 cm^{-1}) and (b) high-frequency Si-H/Ge-H region (1900-2300 cm^{-1}). Frequency scale factors of 0.985 and 0.976 have been applied to the theoretical spectra in low- and high-frequency regions, respectively. The theoretical spectra consist of an admixture of 23 % *n*-GeSiGeGe and 77 % *g*-GeSiGeGe.

Figure 7: Relaxed potential energy surface of as a function of the molecular backbone torsion angle for butane (black), tetrasilane (red) and GeSiSiGe (blue). Solid and dashed lines correspond to CCSD/LANL2DZ and B3LYP/6-311G++(3df,2pd) calculations, respectively.

Figure 8: Temperature dependence of the first layer growth rates for three precursors: H_3GeSiH_3 , Ge_2H_6 and $(\text{H}_3\text{Ge})_2(\text{SiH})_2$ (**1**). The activation energies for these compounds are 1.95, 1.66 and 1.42 eV, respectively.

Figure 9: XTEM micrographs of a $\text{Si}_{0.50}\text{Ge}_{0.50}$ layer grown on Si(100) at 350 °C. No threading defects are observed within the field of view of $\sim 2.5 \mu\text{m}$, indicating a dislocation density of $<10^6/\text{cm}^2$.

Figure 10: Diffraction contrast, high resolution XTEM micrographs of a $\text{Si}_{0.50}\text{Ge}_{0.50}$ layer grown on Si(100) at 400 °C. The film is nearly strain-free ($\sim 85\%$) and displays an atomically smooth surface, free of threading defects (top). The mismatch with Si(100) is accommodated by dislocations located at the interface (bottom left). Defects originating at the interface show the propensity to penetrate downward into the Si(100) substrate (bottom right).

Figure 11: (a) AFM image of $\text{Si}_{0.50}\text{Ge}_{0.50}$ on Si(100) showing the cross-hatch pattern morphology with an RMS roughness of 2.3 nm. (b) AFM image of a strain-relaxed and completely planar film with an RMS roughness of 0.5 nm.

Figure 12: Calculated thermochemical (300K) and structural properties of the conformational isomer pairs $\{n\text{-GeSiSiGe}, g\text{-GeSiSiGe}\}$, $\{n\text{-GeSiGeGe}, g\text{-GeSiGeGe}\}$ and positional isomer $i\text{-Si}(\text{GeSiSi})$. The units for bond lengths, bond angles and energies are Angstroms, degrees and Hartree, respectively.

Figure 13A-B Exemplary sequence of macromolecular, polymeric units based on interconnected $(\text{GeH}_3)_2(\text{SiH}_2)_2$ monomers (a-d), en route to a diamond cubic Si-Ge lattice with concentration Ge_2Si_2 ($\text{Si}_{0.50}\text{Ge}_{0.50}$) (e). This material incorporates the Ge-Si-Si-Ge molecular core of $(\text{GeH}_3)_2(\text{SiH}_2)_2$ as the structural and compositional building block.

Figure 14. Hypothetical growth sequence of a two dimensional island of composition $\text{Si}_{0.50}\text{Ge}_{0.50}$ containing a random distribution of Si and Ge atoms. The schematic process is viewed along the normal to the (111) growth plane, and involves the systematic addition of $n\text{-GeSiSiGe}$ molecules and simultaneous abstraction of H_2 molecules. Legend: dark circles – Ge; light circles – Si; small white circles – H.

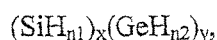
Figure 15. Hypothetical growth sequence of a two dimensional island of composition $\text{Si}_{0.50}\text{Ge}_{0.50}$ containing a random distribution of Si and Ge atoms. The schematic process is viewed along the normal to the (001) growth plane, and involves the systematic addition of $n\text{-GeSiSiGe}$ molecules and simultaneous abstraction of H_2 molecules. Legend: dark circles – Ge; light circles – Si; small white circles – H.

Detailed Description of the Invention

All publications, patents and patent applications cited herein are hereby expressly incorporated by reference for all purposes.

As used herein, the singular forms "a", "an" and "the" include plural referents
5 unless the context clearly dictates otherwise.

In one aspect, the present invention provides compounds comprising or consisting of a silicon-germanium hydride of formula I:



wherein x is 2, 3 or 4;

10 wherein y is 1, 2, or 3;

wherein $x + y$ is 3, 4, or 5;

wherein $n1$ is independently 0, 1, 2, or 3 for each Si atom in the compound, to satisfy valency; and

wherein $n2$ is independently 0, 1, 2, or 3 for each Ge atom in the compound, to
15 satisfy valency;

with the proviso that $n2$ is not 0 when y is 1;

with the further proviso that when x is 3 and y is 1, then $n2$ is 2 or 3; and

with the further proviso that when x is 2 and y is 1, then $n2$ is 3.

20 As used herein, the phrase "to satisfy valency" means to maintain tetra-valency of each Si and Ge atom in the compounds of the invention.

As demonstrated herein, compounds according to formula I can be used, for example, in controlled depositions to produce stoichiometric SiGe films possessing desired properties for semiconductor applications including perfectly crystalline and
25 epitaxial microstructures, smooth morphologies, and uniformly strain-relaxed states. The compounds of formula I are herein demonstrated to be remarkably stable over extended periods of time, and thus represent viable molecular sources for potential industrial applications.

Thus, in one embodiment of this first aspect, x is 2; in an alternative
30 embodiment x is 3; in a further alternative embodiment x is 4. In embodiments where x is 2, y can be 1, 2, or 3 preferably 2 or 3, most preferably 2. In embodiments where x is 3, y can be 1 or 2, preferably 2. In embodiments where x is 4, y is 1.

In a further embodiment of this first aspect, x is 2 or 3, with the proviso that if y is 1, then n_2 is 0, 1, or 2, and with the further proviso that if x is 2 and y is 2, then at least one of the following is true:

- (i) n_1 for at least one Si atom is 0, 1, or 2;
- 5 (ii) n_2 for each Ge atom in the compound is 2; and
- (iii) n_2 for at least one Ge atom in the compound is 0 or 1.

In a still further embodiment, y is 2 or 3 and $x + y$ is 4 or 5, with the proviso that if x is 2 and y is 2, then at least one of the following is true:

- (i) n_1 for at least one Si atom is 0, 1, or 2;
- 10 (ii) n_2 for each Ge atom in the compound is 2; and
- (iii) n_2 for at least one Ge atom in the compound is 0 or 1.

Compounds according to formula I include any conformational form of the compound, including but not limited to *n*-, *g*-, and *iso*-forms of the compounds, and combinations thereof.

15 Exemplary silicon-germanium hydrides according to formula I comprise or consist of those compounds listed in Table 1. All Si and Ge atoms in the compounds are tetravalent. Dashed lines represent bonds between Si and Ge atoms in the linear versions. In the isobutene- and isopentane-like isomers, the Si and Ge atoms inside the brackets are directly bound to the Si or Ge to the left of the brackets; the Si or Ge
20 in parenthesis outside of the brackets at the far right in some of the compounds are directly bound to the last Si or Ge inside of the brackets.

For example, Figure 1 shows compound *i*-Si(SiGeGe). As noted in Table 1, this compound is also referred to as SiH[(SiH₃)(GeH₃)₂].

25 Table 1

3 and 4 metal variants:

(a) Linear

	SiH ₃ -SiH ₂ -GeH ₂ -GeH ₃ ;	Si ₂ Ge ₂ H ₁₀
	SiH ₃ -GeH ₂ -SiH ₂ -GeH ₃ ;	Si ₂ Ge ₂ H ₁₀
30	SiH ₃ -GeH ₂ -GeH ₂ -SiH ₃ ;	Si ₂ Ge ₂ H ₁₀
	GeH ₃ -SiH ₂ -SiH ₂ -GeH ₃ ;	Si ₂ Ge ₂ H ₁₀
	SiH ₃ -SiH ₂ -GeH ₂ -SiH ₃ ;	Si ₃ Ge ₁ H ₁₀
	GeH ₃ -SiH ₂ -SiH ₂ -SiH ₃ ;	Si ₃ Ge ₁ H ₁₀
	SiH ₃ -SiH ₂ -GeH ₃ .	Si ₂ Ge ₁ H ₈

(b) iso-butane-like		
	$\text{SiH}[(\text{SiH}_3)(\text{GeH}_3)_2]$	$\text{Si}_2\text{Ge}_2\text{H}_{10}$
	$\text{SiH}[(\text{SiH}_3)_2(\text{GeH}_3)]$	$\text{Si}_3\text{Ge}_1\text{H}_{10}$
5	$\text{GeH}[(\text{GeH}_3)(\text{SiH}_3)_2]$	$\text{Si}_2\text{Ge}_2\text{H}_{10}$
	$\text{GeH}[(\text{GeH}_3)_2(\text{SiH}_3)]$	$\text{Si}_2\text{Ge}_2\text{H}_{10}$
5 metal atom variants:		
1 Ge:		
10	(a) Linear:	
	$\text{SiH}_3\text{-SiH}_2\text{-SiH}_2\text{-SiH}_2\text{-GeH}_3$	$\text{Si}_4\text{Ge}_1\text{H}_{12}$
	$\text{SiH}_3\text{-SiH}_2\text{-SiH}_2\text{-GeH}_2\text{-SiH}_3$	$\text{Si}_4\text{Ge}_1\text{H}_{12}$
	$\text{SiH}_3\text{-SiH}_2\text{-GeH}_2\text{-SiH}_2\text{-SiH}_3$	$\text{Si}_4\text{Ge}_1\text{H}_{12}$
15	(b) Iso-pentane-like	
	$\text{SiH}[(\text{GeH}_3)(\text{SiH}_3)(\text{SiH}_2)](\text{SiH}_3)$	$\text{Si}_4\text{Ge}_1\text{H}_{12}$
	$\text{SiH}[(\text{SiH}_3)_2(\text{SiH}_2)](\text{GeH}_3)$	$\text{Si}_4\text{Ge}_1\text{H}_{12}$
	$\text{SiH}[(\text{SiH}_3)_2(\text{GeH}_2)](\text{SiH}_3)$	$\text{Si}_4\text{Ge}_1\text{H}_{12}$
	$\text{GeH}[(\text{SiH}_3)_2(\text{SiH}_2)](\text{SiH}_3)$	$\text{Si}_4\text{Ge}_1\text{H}_{12}$
20	$\text{SiH}[(\text{SiH}_3)_2(\text{GeH}_2)](\text{SiH}_3)$	$\text{Si}_4\text{Ge}_1\text{H}_{12}$
Neopentane-like		
	$\text{Si}[(\text{SiH}_3)_3(\text{GeH}_3)]$	$\text{Si}_4\text{Ge}_1\text{H}_{12}$
25	2 Ge:	
	(a) Linear:	
	$\text{SiH}_3\text{-SiH}_2\text{-SiH}_2\text{-GeH}_2\text{-GeH}_3$	$\text{Si}_3\text{Ge}_2\text{H}_{12}$
	$\text{SiH}_3\text{-SiH}_2\text{-GeH}_2\text{-SiH}_2\text{-GeH}_3$	$\text{Si}_3\text{Ge}_2\text{H}_{12}$
	$\text{SiH}_3\text{-GeH}_2\text{-SiH}_2\text{-SiH}_2\text{-GeH}_3$	$\text{Si}_3\text{Ge}_2\text{H}_{12}$
30	$\text{GeH}_3\text{-SiH}_2\text{-SiH}_2\text{-SiH}_2\text{-GeH}_3$	$\text{Si}_3\text{Ge}_2\text{H}_{12}$
	$\text{SiH}_3\text{-SiH}_2\text{-GeH}_2\text{-GeH}_2\text{-SiH}_3$	$\text{Si}_3\text{Ge}_2\text{H}_{12}$
	$\text{SiH}_3\text{-GeH}_2\text{-SiH}_2\text{-GeH}_2\text{-SiH}_3$	$\text{Si}_3\text{Ge}_2\text{H}_{12}$
	(b) Isopentane-like	

	$\text{GeH}[(\text{SiH}_3)_2(\text{GeH}_2)](\text{SiH}_3)$	$\text{Si}_3\text{Ge}_2\text{H}_{12}$
	$\text{GeH}[(\text{SiH}_3)_2(\text{SiH}_2)](\text{GeH}_3)$	$\text{Si}_3\text{Ge}_2\text{H}_{12}$
	$\text{SiH}[(\text{GeH}_3)(\text{SiH}_3)(\text{GeH}_2)](\text{SiH}_3)$	$\text{Si}_3\text{Ge}_2\text{H}_{12}$
	$\text{SiH}[(\text{SiH}_3)(\text{GeH}_3)(\text{SiH}_2)](\text{GeH}_3)$	$\text{Si}_3\text{Ge}_2\text{H}_{12}$
5	$\text{GeH}[(\text{GeH}_3)(\text{SiH}_3)(\text{SiH}_2)]\text{SiH}_3$	$\text{Si}_3\text{Ge}_2\text{H}_{12}$
	$\text{SiH}[(\text{GeH}_3)_2(\text{SiH}_2)]\text{SiH}_3$	$\text{Si}_3\text{Ge}_2\text{H}_{12}$
(c) Neopentane like:		
	$\text{Si}[(\text{SiH}_3)_2(\text{GeH}_3)_2]$	$\text{Si}_3\text{Ge}_2\text{H}_{12}$
10	$\text{Ge}[(\text{SiH}_3)_3(\text{GeH}_3)]$	$\text{Si}_3\text{Ge}_2\text{H}_{12}$
3 Ge:		
(a) Linear		
	$\text{SiH}_3\text{-SiH}_2\text{-GeH}_2\text{-GeH}_2\text{-GeH}_3$	$\text{Si}_2\text{Ge}_3\text{H}_{12}$
15	$\text{SiH}_3\text{-GeH}_2\text{-SiH}_2\text{-GeH}_2\text{-GeH}_3$	$\text{Si}_2\text{Ge}_3\text{H}_{12}$
	$\text{GeH}_3\text{-SiH}_2\text{-SiH}_2\text{-GeH}_2\text{-GeH}_3$	$\text{Si}_2\text{Ge}_3\text{H}_{12}$
	$\text{GeH}_3\text{-SiH}_2\text{-GeH}_2\text{-SiH}_2\text{-GeH}_3$	$\text{Si}_2\text{Ge}_3\text{H}_{12}$
	$\text{SiH}_3\text{-GeH}_2\text{-GeH}_2\text{-GeH}_2\text{-SiH}_3$	$\text{Si}_2\text{Ge}_3\text{H}_{12}$
	$\text{SiH}_3\text{-GeH}_2\text{-GeH}_2\text{-SiH}_2\text{-GeH}_3$	$\text{Si}_2\text{Ge}_3\text{H}_{12}$
20		
(b) Isopentane-like		
	$\text{GeH}[(\text{GeH}_3)(\text{SiH}_3)(\text{GeH}_2)](\text{SiH}_3)$	$\text{Si}_2\text{Ge}_3\text{H}_{12}$
	$\text{GeH}[(\text{SiH}_3)_2(\text{GeH}_2)](\text{GeH}_3)$	$\text{Si}_2\text{Ge}_3\text{H}_{12}$
	$\text{SiH}[(\text{SiH}_3)(\text{GeH}_3)(\text{GeH}_2)](\text{GeH}_3)$	$\text{Si}_2\text{Ge}_3\text{H}_{12}$
25	$\text{GeH}[(\text{GeH}_3)(\text{SiH}_3)(\text{SiH}_2)](\text{GeH}_3)$	$\text{Si}_2\text{Ge}_3\text{H}_{12}$
	$\text{GeH}[(\text{GeH}_3)_2(\text{SiH}_2)](\text{SiH}_3)$	$\text{Si}_2\text{Ge}_3\text{H}_{12}$
	$\text{SiH}[(\text{GeH}_3)_2(\text{GeH}_2)](\text{SiH}_3)$	$\text{Si}_2\text{Ge}_3\text{H}_{12}$
	$\text{SiH}[(\text{GeH}_3)_2(\text{SiH}_2)](\text{GeH}_3)$	$\text{Si}_2\text{Ge}_3\text{H}_{12}$
30	(c) Neopentane like	
	$\text{Si}[(\text{GeH}_3)_3(\text{SiH}_3)]$	$\text{Si}_2\text{Ge}_3\text{H}_{12}$
	$\text{Ge}[(\text{GeH}_3)_2(\text{SiH}_3)_2]$	$\text{Si}_2\text{Ge}_3\text{H}_{12}$

As noted above, these compounds each include the *n* or *g* forms, and stereoisomers thereof.

Figure 1 provides molecular structures of exemplary silicon-germanium hydrides according to this first aspect of the invention.

5 In a preferred embodiment, the silicon germanium hydride is selected from the group consisting of SiH₃-SiH₂-GeH₂-GeH₃; SiH₃-GeH₂-SiH₂-GeH₃; SiH₃-GeH₂-GeH₂-SiH₃; and GeH₃-SiH₂-SiH₂-GeH₃), and SiH[(SiH₃)(GeH₃)₂]

This first aspect also provides compositions comprising combinations of the silicon germanium hydrides. In a preferred embodiment of such combinations, x is 3
10 and y is 1, or x is 2 and y is 2, for each of the two or more compounds. In a most preferred embodiment, the combination comprises compounds selected from the group consisting of SiH₃-SiH₂-GeH₂-GeH₃; SiH₃-GeH₂-SiH₂-GeH₃; SiH₃-GeH₂-GeH₂-SiH₃; and GeH₃-SiH₂-SiH₂-GeH₃), and SiH[(SiH₃)(GeH₃)₂].

In a further embodiment, the compounds of the invention are substantially
15 isolated. As used herein, a "substantially isolated" compound of the invention is one compound, or two or more compounds of the invention in combination, that are at least 75% isolated from precursors and other contaminants, and preferably 80%, 85%, 90%, 95%, 98%, or more isolated.

In a second aspect, the present invention provides semiconductor structures,
20 comprising a substrate; and an Si-Ge layer formed by introducing near the surface of a substrate a gaseous precursor comprising or consisting of one or more silicon-germanium hydride compounds of formula I.

As shown herein, the silicon-germanium hydride compounds of formula I are especially useful for fabricating SiGe semiconductor alloys at unprecedented low
25 temperatures that display homogeneous compositional and strain profiles, low threading dislocation densities and atomically planar surfaces. Controlled depositions have produced stoichiometric SiGe films reflecting the Si/Ge content of the precursor, and possessing the desired properties for semiconductor applications including perfectly crystalline and epitaxial microstructures, smooth morphologies, and
30 uniformly strain-relaxed states.

In this aspect, the substrate can be any substrate suitable for semiconductor or flat panel display use, including but not limited to silicon, germanium, silicon on insulator, Ge:Sn alloys, SiO₂, sapphire, quartz, Si:Ge alloys, Si:C alloys, stainless

steel, polyimide or other polymer films such as those used in the fabrication of flexible displays. In a preferred embodiment, the substrate comprises Si(100).

In a preferred embodiment, the semiconductor substrates of this aspect of the invention have an SiGe layer which comprises an SiGe film having a thickness of less than one micron, more preferably a thickness in a range from 50 nm to 500 nm. In a further preferred embodiment, the semiconductor substrates of this second aspect have an SiGe layer comprising an SiGe film having a density of threading defects of $10^6/\text{cm}^2$ or less. In a further preferred embodiment, the semiconductor substrates of this aspect of the invention comprise an SiGe film having a substantially atomically planar surface morphology.

In various embodiments, the gaseous precursor comprises or consists of one or more of the compounds listed in Table 1, each in *n* or *g* forms, and stereoisomers thereof.

In a preferred embodiment of the semiconductor substrates of the second aspect, the gaseous precursor comprises or consists of a compound selected from the group consisting of $\text{SiH}_3\text{-SiH}_2\text{-GeH}_2\text{-GeH}_3$; $\text{SiH}_3\text{-GeH}_2\text{-SiH}_2\text{-GeH}_3$; $\text{SiH}_3\text{-GeH}_2\text{-GeH}_2\text{-SiH}_3$; $\text{GeH}_3\text{-SiH}_2\text{-SiH}_2\text{-GeH}_3$, and $\text{SiH}[(\text{SiH}_3)(\text{GeH}_3)_2]$.

The semiconductor structure may further comprise other features as desired, including but not limited to the inclusion of dopants, such as boron, phosphorous, arsenic, and antimony. These embodiments are especially preferred for semiconductor substrates used as active devices. Inclusion of such dopants into the semiconductor substrates can be carried out by standard methods in the art.

In a further embodiment, the semiconductor structure may further comprise varying quantities of carbon or tin, as desired for a given application. Inclusion of carbon or tin into the semiconductor substrates can be carried out by standard methods in the art. The carbon can be used to reduce the mobility of the dopants in the structure and more specifically boron. Incorporation of Sn can yield materials with novel optical properties such as direct emission and absorption leading to the formation of Si-based lasers and high sensitivity infrared photodetectors.

In further aspect, the invention provides semiconductor structures, comprising a substrate and an Si-Ge layer comprising or consisting of a backbone of one or more compounds according to the formula I. Preferred embodiments are as described above.

5 In a further aspect, the invention provides a composition, comprising one or more compounds according to formula 1 in an inert gas. Such inert gases include, but are not limited to H₂, He, N₂, and argon. Preferred embodiments are as described above.

By using the compounds of formula I as precursors for preparing
10 semiconductor substrates, the semiconductor substrate will comprise an Si-Ge layer comprising or consisting of a backbone of the compounds of the invention. By using the compounds of the invention, as opposed to the separate use of Si and Ge sources for deposition, the resulting Si-Ge layer comprises highly defined and uniform bonding arrangements and strain compensated Si-Ge atomic patterns, as opposed to
15 the random nature of prior Si-Ge films in which the Si and Ge atoms can exist as an intimate mixture of several (or multi) atom Si and Ge clusters leading to inhomogeneous bonding and locally stressed configurations. In mismatched heteroepitaxy applications these materials exhibit copious dislocations (at levels unacceptable for most device applications), strain non-uniformities and high surface
20 roughness. A distinct and important advantage of our compounds, films, and methods over conventional ones is that the incorporation of the entire Si-Ge molecular core promotes the formation of exceptionally uniform bonding arrangements over the entire crystal, leading to relaxed (or uniformly stressed) films with planar surface morphology (no surface ripples). A crucial advantage is the unprecedented low
25 growth temperatures which reduce surface mobility of the Si and Ge atoms and prevent mass segregation thereby resulting in highly uniform compositional and strain profiles at the atomic level.

The intact incorporation of the massive Si/Ge cores into the film also leads to low surface diffusion and high sticking coefficients compared to the conventional low
30 mass compounds. The high reactivity permits promotes high growth rates which increase systematically with the mass of the precursors

Furthermore, the incorporation of the *entire* Si/Ge framework of the gaseous precursor into the lattice is particularly important in the formations of SiGe nanoscale systems (quantum dots and wires) because it allows precise control of morphology,

tunable composition, and structure that cannot be accomplished by vapor deposition of separate silanes and germanes, or by molecular beam epitaxy of solid Si and Ge sources. The precise control of composition and corresponding morphology, both of which ultimately determine the physical properties of these nanoscale materials is a crucial issue. Any compositional variations that exist within and among individual nanostructures will have a significant impact on the optical and electronic properties and on the performance of devices. In the MBE case an average concentration is assigned on the basis of the flux ratio utilized to grow the materials. In the CVD approach the local stoichiometry on the scale of the nanostructure size is even less certain because reactions of various molecular sources cannot be simultaneously controlled at every site throughout the sample. Thus neither technique can guarantee systematic growth of Si-Ge nanoscale structures with controllable composition, strain and size.

Such highly defined patterns can be determined using standard methods in the art, such as neutron diffraction

An example of an Si-Ge layer comprising a backbone of the compounds of the invention can be found in **Figure 13**, which shows an exemplary sequence of macromolecular, polymeric units of the invention, based on interconnected $(\text{GeH}_3)_2(\text{SiH}_2)_2$ monomers (a-d), en route to a diamond cubic Si-Ge lattice with concentration Ge_2Si_2 ($\text{Si}_{0.50}\text{Ge}_{0.50}$) (e). This material incorporates the Ge-Si-Si-Ge molecular core of $(\text{GeH}_3)_2(\text{SiH}_2)_2$ as the structural and compositional building block.

Figures 14 and 15 illustrate schematically, the molecule-by-molecule growth of a random 2D (one layer) clusters of SiGe film. **Figure 14** shows the hypothetical growth sequence of a two dimensional island of composition $\text{Si}_{0.50}\text{Ge}_{0.50}$ containing a random distribution of Si and Ge atoms. The schematic process is viewed along the normal to the (111) growth plane and involves the systematic addition of n -GeSiSiGe molecules and simultaneous abstraction of H_2 molecules. **Figure 15** is similar, but The schematic process is viewed along the normal to the (001) growth plane.

Thus, the compounds of the invention can be used to form a random alloy with intimate mixing of the elements using the GeSiSiGe precursors of formula I. The resulting Si-Ge backbone structure consists of polymerized units of the starting compound of the invention, wherein polymerization requires dehydrogenation.

As will be understood by those of skill in the art, other macromolecular, polymeric units of the invention, based on one or more other polymerized compounds

of the invention can also lead to Si-Ge lattices with a desired concentration of Ge and Si based on the concentration of each in the compound used.

Preferred embodiments of this aspect are as discussed above for the second aspect of the invention.

5

In a third aspect, the present invention provides methods for synthesizing one or more silicon-germanium hydrides of formula I, $(\text{SiH}_{n1})_x(\text{GeH}_{n2})_y$, wherein the method comprises combining one or more perfluoroalkyl-sulfonate-substituted disilane with a compound comprising a GeH_3 ligand under conditions whereby the silicon-germanium hydride is formed. In a preferred embodiment, the perfluoroalkyl-sulfonate-substituted disilane is selected from the group consisting of a nonafluorobutane-sulfonic substituted disilane and a triflate substituted disilane

10

Thus, this aspect provides methods for making the compounds of formula I. The synthetic routes described herein utilize high yield single-step substitution reactions to yield semiconductor grade materials suitable for industrial application and manufacturing processes.

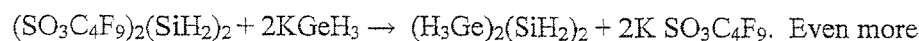
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In one embodiment of this third aspect, the compound comprising a GeH_3 ligand is selected from the group consisting alkali and alkaline earth germanides. In a preferred embodiment, the alkali or alkaline earth germanides are selected from the group consisting of KGeH_3 , NaGeH_3 , and MR_3GeH_3 , wherein M is a group IV element and R is an organic ligand.

20

In a further embodiment of this third aspect, the method comprises combining the compound comprising a GeH_3 ligand with a nonafluorobutane-sulfonic substituted disilane, wherein the nonafluorobutane-sulfonic substituted disilane comprises $(\text{SO}_3\text{C}_4\text{F}_9)_2(\text{SiH}_2)_2$. In other embodiments, entire families of perfluoroalkyl sulfonyl and alkyl sulfonyl substituted silanes (or germanes) can be used. More preferably, the compound comprising a GeH_3 ligand is KGeH_3 , and the silicon germanium hydride is formed according to the equation

25



30

Even more preferably, the method comprises:

- (a) adding $(\text{SO}_3\text{C}_4\text{F}_9)_2(\text{SiH}_2)_2$ to a slurry of KGeH_3 in solvent at between -50°C and -30°C (more preferably about -40°C) to form a mixture;
- (b) warming the mixture to between -20°C and 0°C (more preferably about -10°C);

- (c) stirring the mixture under nitrogen;
- (d) heating the mixture to between 25°C and 45°C (more preferably about 35°C); and
- (e) distilling the mixture to form the $(\text{H}_3\text{Ge})_2(\text{SiH}_2)_2$.

5

In an alternative embodiment of this third aspect, the method comprises combining the compound comprising a GeH_3 ligand with a triflate substituted disilane, and wherein the triflate substituted disilane comprises $(\text{SO}_3\text{CF}_3)_2(\text{SiH}_2)_2$. More preferably, the compound comprising a GeH_3 ligand is KGeH_3 , and the silicon

10 germanium hydride is formed according to the equation

$(\text{SO}_3\text{CF}_3)_2(\text{SiH}_2)_2 + 2\text{KGeH}_3 \rightarrow (\text{H}_3\text{Ge})_2(\text{SiH}_2)_2 + 2\text{KSO}_3\text{CF}_3$. Even more preferably, the method comprises

- (a) adding $(\text{SO}_3\text{CF}_3)_2(\text{SiH}_2)_2$ to a slurry of KGeH_3 in solvent at between -40°C and -10°C (more preferably about -30°C) to form a mixture;
- 15 (b) optionally warming the mixture to between 10°C and 35°C (more preferably about 23°C);
- (c) stirring the mixture under nitrogen; and
- (d) distilling the mixture to form the $(\text{H}_3\text{Ge})_2(\text{SiH}_2)_2$.

In this embodiment, low temperature ($T < -10^\circ\text{C}$) reactions of $(\text{SO}_3\text{CF}_3)_2(\text{SiH}_2)_2$ with KGeH_3 afforded predominately $(\text{H}_3\text{Ge})_2(\text{SiH}_2)_2$ and impurity levels of $(\text{GeH}_3)_2(\text{SiH}_2\text{GeH}_2)$. These species can be separated by distillation. In contrast to the low temperature pathways, the reaction of $(\text{SO}_3\text{CF}_3)_2(\text{SiH}_2)_2$ with KGeH_3 at $T \geq 22^\circ\text{C}$ yields predominately combinations of the positional isomers $(\text{GeH}_3)_2(\text{SiH}_2)_2$ and $(\text{GeH}_3)_2\text{SiH}(\text{SiH}_3)$ in varying proportions (a small amount of $(\text{GeH}_3)_2(\text{SiH}_2\text{GeH}_2)$ is

20 also produced but this is readily separated from the bulk via fractional distillation from the $(\text{SO}_3\text{CF}_3)_2(\text{SiH}_2)_2$ and $(\text{GeH}_3)_2(\text{SiH}_2\text{GeH}_2)$ mixture.

In a fourth aspect, the present invention provides methods for depositing a Si-Ge material on a substrate in a reaction chamber, comprising introducing into the

30 chamber a gaseous precursor comprising or consisting of one or more compounds according to formula I, under conditions whereby a layer comprising a Si-Ge material is formed on the substrate.

In a fifth aspect, the present invention provides methods for depositing an epitaxial Si-Ge layer on a substrate, comprising introducing near a surface of the substrate a gaseous precursor comprising or consisting of one or more compounds according to formula I, and dehydrogenating the precursor under conditions whereby
5 epitaxial Si-Ge is formed on the substrate.

In this aspect, the substrate can be any substrate suitable for semiconductor or flat panel display use, including but not limited to silicon, germanium, silicon on insulator, Ge:Sn alloys, SiO₂, sapphire, quartz, Si:Ge alloys, Si:C alloys, stainless steel, polyimide or other polymer films. In a preferred embodiment, the substrate
10 comprises Si(100).

Embodiments of the gaseous precursors are as described above for previous aspects of the invention. For example, the methods may further comprise adding a dopant on the substrate, including but not limited to the inclusion of dopants, such as boron, phosphorous, arsenic, and antimony. These embodiments are especially
15 preferred for semiconductor substrates used as active devices. Inclusion of such dopants into the semiconductor substrates can be carried out by standard methods in the art.

In another embodiment, the methods comprise adding varying quantities of carbon or tin to the semiconductor substrate. Inclusion of carbon or tin into the
20 semiconductor substrates can be carried out by standard methods in the art. The carbon can be used to reduce the mobility of the dopants in the structure and more specifically boron. Incorporation of Sn can yield materials with novel optical properties such as direct emission and absorption leading to the formation of Si-based lasers and high sensitivity infrared photodetectors.

25 As demonstrated herein, the silicon-germanium hydrides of the invention were used to deposit device quality layers on substrates that display homogeneous compositional and strain profiles, low threading dislocation densities and atomically planar surfaces.

In a preferred embodiment of the fourth and fifth aspects of the invention, the
30 step of introducing the gaseous precursor comprises introducing the gaseous precursor in substantially pure form. In a further preferred embodiment, the step of introducing the gaseous precursor comprises introducing the gaseous precursor as a single gas source.

In another embodiment of the fourth and fifth aspects, the step of introducing the gaseous precursor comprises introducing the gaseous precursor intermixed with an inert carrier gas. In this embodiment, the inert gas can be, for example, He, argon, H₂ or N₂.

5 In these aspects, the gaseous precursor can be deposited by any suitable technique, including but not limited to gas source molecular beam epitaxy, chemical vapor deposition, plasma enhanced chemical vapor deposition, laser assisted chemical vapor deposition, and atomic layer deposition.

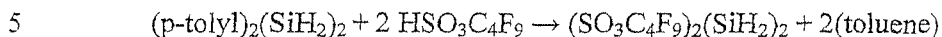
10 In a preferred embodiment of the fourth and fifth aspects, the gaseous precursor is introduced at a temperature of between 300°C and 450°C, more preferably between 350°C and 450°C. Practical advantages associated with this low temperature/rapid growth process include (i) short deposition times compatible with preprocessed Si wafers, (ii) selective growth for application in high frequency devices, and (iii) negligible mass segregation of dopants, which is particularly critical
15 for thin layers.

In various further embodiments of the fourth and fifth aspects, the gaseous precursor is introduced at a partial pressure between 10⁻⁸ Torr and 1000 Torr. In one preferred embodiment, the gaseous precursor is introduced at between 10⁻⁸ Torr and 10⁻⁵ Torr (corresponding to UHV vertical furnace technology). In another preferred
20 embodiment, the gaseous precursor is introduced at between 10⁻⁸ Torr and 100 Torr, corresponding to LPCVD conditions.

In various other preferred embodiments, the Si-Ge material is formed on the substrate as a strain-relaxed layer having a planar surface; the composition of the Si-Ge material is substantially uniform; and/or the entire Si and Ge framework of the
25 gaseous precursor is incorporated into the Si-Ge material or epitaxial Si-Ge.

In a sixth aspect, the present invention provides a compound comprising or consisting of (SO₃C₄F₉)₂(SiH₂)₂, as well as methods for making the compound. As disclosed herein, (SO₃C₄F₉)₂(SiH₂)₂ is useful as an intermediate in the synthesis of the
30 silicon-germanium hydrides of the invention. In one embodiment, methods for synthesizing (SO₃C₄F₉)₂(SiH₂)₂ comprise combining R(SiH₂)₂ and HSO₃C₄F₉ under conditions whereby the (SO₃C₄F₉)₂(SiH₂)₂ is formed, where R is an aryl substituent. Such aryl substituents include, but are not limited to, p-tolyl and benzyl.

In a preferred embodiment, methods for synthesizing $(\text{SO}_3\text{C}_4\text{F}_9)_2(\text{SiH}_2)_2$ comprise combining $(p\text{-tolyl})_2(\text{SiH}_2)_2$ and $\text{HSO}_3\text{C}_4\text{F}_9$ under conditions whereby the $(\text{SO}_3\text{C}_4\text{F}_9)_2(\text{SiH}_2)_2$ is formed. Preferably, the $(\text{SO}_3\text{C}_4\text{F}_9)_2(\text{SiH}_2)_2$ is formed according to the equation



More preferably, the method comprises

- (a) combining $\text{HSO}_3\text{C}_4\text{F}_9$, $(p\text{-tolyl})_2(\text{SiH}_2)_2$, and solvent in a reaction chamber at between -50°C and 0°C to form a mixture;
- (b) warming the mixture to between 20°C and 25°C ;
- 10 (c) stirring the mixture; and
- (d) filtering the solid $(\text{SO}_3\text{C}_4\text{F}_9)_2(\text{SiH}_2)_2$.

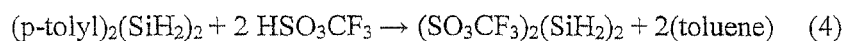
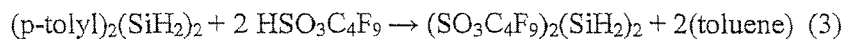
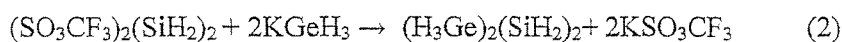
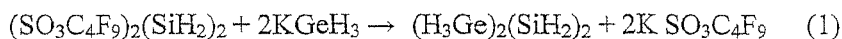
Examples

The synthesis of butane-like $(\text{GeH}_3)_2(\text{SiH}_2)_2$ (Referred to above as GeH3-SiH2-SiH2-GeH3)(1), $(\text{GeH}_3)_2\text{SiH}(\text{SiH}_3)$ (referred to above as SiH $[(\text{GeH}_3)_2(\text{SiH}_3)]$ (2) and $(\text{GeH}_3)_2(\text{SiH}_2\text{GeH}_2)$ (also referred to as GeH3-SiH2-GeH2-GeH3) (3) Si-Ge hydrides with applications in low temperature synthesis of Ge-rich $\text{Si}_{1-x}\text{Ge}_x$ optoelectronic alloys is herein demonstrated. The compositional, vibrational, structural and thermochemical properties of these compounds were studied by FTIR, multinuclear NMR, mass spectrometry, Rutherford backscattering, and density functional theory (DFT) simulations. The analyses indicate that the linear $(\text{GeH}_3)_2(\text{SiH}_2)_2$ (1) and $(\text{GeH}_3)_2(\text{SiH}_2\text{GeH}_2)$ (3) compounds exist as a mixture of the classic normal and gauche conformational isomers which do not seem to interconvert at 22°C . The conformational proportions in the samples were determined using a new fitting procedure, which combines calculated molecular spectra to reproduce those observed by varying the global intensity, frequency scale and admixture coefficients of the individual conformers. The $(\text{GeH}_3)_2(\text{SiH}_2)_2$ (1) species was then utilized to fabricate $\text{Si}_{0.50}\text{Ge}_{0.50}$ semiconductor alloys reflecting exactly the Si/Ge content of the precursor. Device quality layers were grown via gas source MBE directly on Si(100) at unprecedented low temperatures $350\text{--}450^\circ\text{C}$ and display homogeneous compositional and strain profiles, low threading dislocation densities and atomically planar surfaces. Low energy electron microscopy (LEEM) analysis has demonstrated that the precursor is highly reactive on Si(100) surfaces, with H₂

desorption kinetics comparable to those of Ge_2H_6 , despite the presence of strong Si-H bonds in the molecular structure.

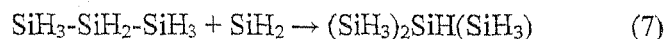
Results and Discussion

5 The synthesis of the main product 1,2-digermyldisilane, $(\text{H}_3\text{Ge})_2(\text{SiH}_2)_2$ (**1**) was accomplished via reactions of KGeH_3 with the nonafluorobutane-sulfonic and triflate substituted disilanes $(\text{C}_4\text{F}_9\text{SO}_3)_2(\text{SiH}_2)_2$ (**4**) and $(\text{SO}_3\text{CF}_3)_2(\text{SiH}_2)_2$ (**5**) as illustrated by Eq 1 and 2 respectively. The starting materials (**4**) and (**5**) were prepared *via* reactions of the corresponding sulfonic acids $\text{HSO}_3\text{C}_4\text{F}_9$ and HSO_3CF_3 with 1,2-(*p*-tolyl) $_2(\text{SiH}_2)_2$ (Eq 3 and 4, respectively).



15 The route described by (Eq 1) consistently produces pure, single-phase $(\text{H}_3\text{Ge})_2(\text{SiH}_2)_2$ (**1**) as a colorless, pyrophoric liquid. Here the replacement of $\text{SO}_3\text{C}_4\text{F}_9$ by GeH_3 occurs exclusively at the 1,2-substitution sites of $(\text{SO}_3\text{C}_4\text{F}_9)_2(\text{SiH}_2)_2$ (**4**), as expected. However, the route described by (Eq 2) yields a mixture of products (Fig. 2) including $(\text{GeH}_3)_2(\text{SiH}_2)_2$ (**1**), the isobutane analog $(\text{GeH}_3)_2\text{SiH}(\text{SiH}_3)$ (**2**), and the Ge-rich derivative $(\text{GeH}_3)_2(\text{SiH}_2\text{GeH}_2)$ (**3**) depending on reaction conditions. Specifically, low temperature ($T < -10^\circ\text{C}$) reactions of $(\text{SO}_3\text{CF}_3)_2(\text{SiH}_2)_2$ (**5**) with KGeH_3 (Eq 2) afforded predominately (**1**) and impurity levels of (**3**). These species are separated by distillation and the minor low volatility fraction (vapor pressure of ~ 1 Torr at 22°C) was characterized by a variety of spectroscopic methods as the Ge-rich derivative (**3**) possessing a heteronuclear $\text{H}_2\text{Si-GeH}_2$ core (Fig. 1). In contrast to the low temperature pathways, the reaction of $(\text{SO}_3\text{CF}_3)_2(\text{SiH}_2)_2$ (**5**) with KGeH_3 at $T \geq 22^\circ\text{C}$ yields predominately combinations of the positional isomers (**1**) and (**2**) in varying proportions (a small amount of (**3**) is also produced but this is readily separated from the bulk via fractional distillation from the (**1**) and (**2**) mixture). In this case the replacement of SO_3CF_3 by GeH_3 initially occurs at the 1-substitution site of $(\text{SO}_3\text{CF}_3)_2(\text{SiH}_2)_2$ (**5**) as expected. The second replacement likely produces germynyl disilane, $\text{SiH}_3\text{SiH}_2\text{GeH}_3$, and a stable germylene intermediate (GeH_2) rather

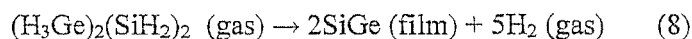
than the 1,2-digermyldisilane, $(\text{H}_3\text{Ge})_2(\text{SiH}_2)_2$ (1). Insertion of GeH_2 into $\text{SiH}_3\text{SiH}_2\text{GeH}_3$ then may yield $(\text{GeH}_3)_2\text{SiH}(\text{SiH}_3)$ (3) as shown in Eq 6. A similar insertion mechanism (discussed in detail in subsequent sections) has been observed in the photochemical generation of gas phase $i\text{-Si}_4\text{H}_{10}$ from SiH_2 insertion into trisilane $\text{SiH}_3\text{-SiH}_2\text{-SiH}_3$ (Eq 7):



The synthesis of (5) is described by (Eq 4) and has been previously demonstrated elsewhere. Its $(p\text{-tolyl})_2(\text{SiH}_2)_2$ starting materials has been produced via coupling reactions of $(p\text{-tolyl})\text{SiH}_2\text{Cl}$ with powdered lithium.¹⁰ In the present work we adopt a more convenient, analogous approach using readily available sodium fragments as the coupling agent. Compound (5) is obtained as a highly reactive liquid, which is difficult to fully handle and purify. Consequently its reactions with KGeH_3 are complex and produce mixtures of products as described above. In contrast the newly prepared compound (4) is readily crystallized as a robust molecular solid that is completely isolated from the reaction byproducts and other contaminants, thereby leading to high purity $(\text{H}_3\text{Ge})_2(\text{SiH}_2)_2$ in subsequent steps according to (Eq 1). The compositional, vibrational and structural properties of the newly synthesized compounds (1), (2), (3) and (4) are presented below.

$(\text{H}_3\text{Ge})_2(\text{SiH}_2)_2$ (1): The $(\text{H}_3\text{Ge})_2(\text{SiH}_2)_2$ (1) species is prepared as described by the reactions depicted in (Eq 1,2) and was isolated and purified by fractional distillation (vapor pressure of ~8-10 Torr at 22 °C). The FTIR, ^{29}Si NMR (proton-coupled) and mass spectra indicate a 1,2-digermyldisilane with a butane-like structure as shown in Figure 1. The IR spectrum shows the Si-H and Ge-H stretching modes at 2147 and 2073 cm^{-1} , respectively, and a set of absorption bands between 910 and 442 cm^{-1} corresponding to the bending modes of the molecule. These assignments are consistent with those made previously for the $(\text{H}_3\text{Ge})_{4-x}\text{SiH}_x$ ($x=1-4$) family of silyl germanes.⁷ A more detailed interpretation of the full IR spectrum is obtained from quantum chemical calculations and the data show fixed combinations between the butane-like normal and gauche conformers of (1) (see section below for details). The

¹H NMR spectra reveals a quartet at 3.29 ppm and a triplet at 3.11 ppm corresponding to the Si-H and Ge-H protons, respectively, which is consistent with a H₃Ge-SiH₂-SiH₂-GeH₃ molecular structure. This structure is further confirmed with 2D ¹H COSY and ¹H-²⁹Si HMQC NMR experiments. The 2D ¹H COSY spectrum showed cross-peaks that correlate to Si-H and Ge-H resonances at 3.29 and 3.11 ppm, respectively, indicating the H-Si-Ge-H connectivity. A proton-decoupled ¹H-²⁹Si HMQC spectrum showed that the H atoms at 3.29 ppm are directly attached to Si atoms at -105 ppm. The mass spectra revealed an isotopic envelope at 210-196 amu as the highest mass peak corresponding to the parent ion, (Si₂Ge₂H_x⁺). The associated fragmentation pattern is consistent with the proposed structure. The existence of (1) was further corroborated by Rutherford backscattering (RBS) analysis of monocrystalline films produced by single source MBE depositions of the compound at 350 °C (Eq 8). The Si/Ge elemental content of the films was consistently 1:1 reflecting precisely the composition of the Si₂Ge₂ core of the molecule.



(GeH₃)₂SiH(SiH₃) (2): The isobutane-like structure of (2) was determined by ²⁹Si and ¹H NMR studies. The proton NMR resonances are present as a multiplet centered at 3.049 ppm due to -SiH, and two doublets centered at 3.213 and 3.404 ppm corresponding to the (-GeH₃)₂ and -SiH₃ protons, respectively. The integrated peak intensity ratios of the -SiH, (-GeH₃)₂ and -SiH₃ are 1:6:3, respectively, as expected. The peak intensities, the coupling patterns and the position of the peaks in the spectrum, collectively point to the isobutane-like structure in which an Si center is tetrahedrally bonded to two GeH₃ groups one SiH₃ group and an single H atom. This structure is further confirmed by 2D ¹H COSY and ¹H-²⁹Si HMQC NMR measurements. The 2D ¹H COSY spectrum showed cross-peaks that correlate to the central SiH, and terminal GeH₃ and SiH₃ resonances at 3.049, 3.213 and 3.404 ppm, respectively, indicating a H₃Si-SiH-(GeH₃)₂ connectivity, which is consistent with the proposed isobutane structure. A proton-decoupled ¹H-²⁹Si HMQC spectrum showed that the H atoms at 3.049 and 3.404 ppm are directly bonded to Si atoms at -121 and -93 ppm, respectively. The ²⁹Si chemical shift (-121 ppm, in C₆D₆) due to the central Si atom in this molecule is similar to those found for SiH(GeH₃)₃ (-113 ppm in CDCl₃) and Si(GeH₃)₄ (-128 ppm, in CDCl₃). A close correspondence is also

found between the ^{29}Si chemical shift (-93 ppm) due to the terminal Si atom and those of the terminal SiH_3 groups (-92 ppm) for H_3SiGeH_3 .

Gas-chromatography mass-spectrometry experiments (GCMS) of samples containing (1) $(\text{H}_3\text{Ge})_2(\text{SiH}_2)_2$ and $(\text{GeH}_3)_2\text{SiH}(\text{SiH}_3)$ (2) indicated the presence of two major fractions that displayed slightly different retention times in the GC column, which is consistent with a mixture of (1) and (2) possessing slightly different boiling points. Furthermore, these fractions displayed essentially identical mass spectra. The highest mass peak is observed in the 210–196 amu range corresponding to the $(\text{Si}_2\text{Ge}_2\text{H}_x)^+$ parent ions of (1) and (2). Samples of the mixture were further characterized using IR spectroscopy. The spectra are essentially a simple admixture of the individual (1) and (2) components. The spectrum of (2) was theoretically determined since the compound could not be completely separated in pure form from (1) via our distillation method due to the similarity in their boiling points. The contributions of the each compound in the experimental IR spectra were then determined using a fitting procedure involving both the observed and the calculated spectra.

$(\text{GeH}_3)_2(\text{SiH}_2\text{GeH}_2)$ (3): Compound (3) is a byproduct of the reaction described by Eq 2. Its gas-phase IR spectrum shows the Si–H and Ge–H stretching bands centered at 2145 and 2073 cm^{-1} , respectively. The intensity of the Ge–H peak in the spectrum is significantly stronger than the Si–H due to the greater number of Ge–H bonds relative to Si–H bonds. The ^1H NMR spectra showed a set of two sextets at 3.30 and 3.03 ppm encompassing two triplets at 3.23, 3.13 ppm. The integrated peak intensity ratios of the sextet-triplet-triplet-sextet sequence are 2:3:3:2. The integrated intensities the coupling patterns and the position of the peaks in the NMR spectrum collectively point to a $\text{H}_3\text{Ge-SiH}_2\text{-GeH}_2\text{-GeH}_3$ molecular structure in which a central SiH_2GeH_2 core is bonded with two terminal GeH_3 moieties. In particular, the sextets at 3.30 and 3.03 ppm can be associated, respectively, with the SiH_2 and GeH_2 protons of the central $-\text{SiH}_2\text{GeH}_2-$ while the triplets at 3.23, 3.13 ppm correspond to terminal GeH_3 protons. The SiH_2 and GeH_2 NMR frequencies in this case correlate well with those in the previously reported $(\text{H}_3\text{Ge})_2\text{SiH}_2$ and $\text{GeH}_3\text{GeH}_2\text{SiH}_3$ compounds.⁷ The assignment of the sextet at 3.30 ppm to the SiH_2 protons was further corroborated by a ^1H - ^{29}Si HMQC NMR spectrum, which indicates a direct coupling to a single ^{29}Si resonance found at -98.2 ppm. In addition, a 2D ^1H COSY NMR spectrum was used

to unequivocally establish the specific ordering of the Ge-Si-Ge-Ge backbone constituents. Cross-peaks correlate an Si-H and three Ge-H resonances at 3.30, 3.23, 3.13 and 3.03 ppm. These correspond, respectively, to SiH₂, GeH₃ (terminal) connected to GeH₂, Ge-H₃ (terminal) connected to SiH₂, and GeH₂. The mass spectra showed an isotopic envelope at 256–238 amu as the highest mass peak corresponding to the parent ion, (SiGe₃H_x⁺) and a fragmentation pattern consistent with the molecular structure.

(C₄F₉SO₃)₂(SiH₂)₂ (4): Compound (4) is isolated at 95% yield as a colorless, pyrophoric solid with a melting point of 68 °C. It is readily soluble in ethers, CHCl₃, CH₂Cl₂ and slightly soluble in toluene. The ¹H NMR spectrum showed a singlet centered at 5.08 ppm (δ Si-H) due to SiH₂ and ²⁹Si satellite peaks. These exhibit a triplet of triplets with a 1-bond Si-H coupling of 272 Hz and a H-H 3-bond coupling of 3.6 Hz. This indicates the presence of a second SiH₂ group consistent with the –SiH₂SiH₂– core. Furthermore, a proton decoupled ¹H–²⁹Si HMQC showed that the proton resonance at 5.08 ppm is directly attached to a Si atom at –30.1 ppm. These values are consistent with those observed for the triflate analog (SO₃CF₃)₂(SiH₂)₂ (5).¹⁰ Compound (4) was further characterized by FTIR, mass spectrometry and C, H, F elemental analysis and the data are fully consistent with the proposed structure (see experimental section).

Ab initio simulations of molecular properties

Structural and thermochemical properties: In order to corroborate the identification of the (GeH₃)₂(SiH₂)₂, (GeH₃)₂SiH(SiH₃) and (GeH₃)₂(SiH₂GeH₂) described above we simulated the structural, energetic and vibrational trends of these compounds. In prior studies⁷ we demonstrated that B3LYP hybrid density functional theory (DFT) simulations provide an excellent account of the ground state structural, thermochemical and vibrational properties of the (H₃Ge)_xSiH_{4-x} family of molecular hydrides. In particular use of the 6-311G++(2d,2p) basis set yielded small typical bond length and frequency discrepancies (on the order of 0.4% and 1.4%, respectively). Accordingly, our first attempt at reconciling the observed IR spectra with those calculated was based on the same procedure, and on the assumption that the fundamental structures are *n*-butane like. Although this approach provided a close match to most of the main spectral features observed, several strong low-frequency

features remained unaccounted for in the calculated IR pattern. In view of the prior success of B3LYP approach we concluded that *n*-butane like molecular structures alone could not account for the spectrum of the synthesized compounds. We therefore expanded our calculations to include the gauche conformational isomers in which a terminal GeH₃ group is rotated out of the Ge-Si-Si-Ge (or Ge-Si-Ge-Ge) skeletal plane, as well as the iso-butane like positional isomer *i*-Si(SiGeGe) (Figure 1). These isomers are completely analogous to those previously reported in tetrasilane, and related molecules, and in subsequent discussion we simply refer to them as *n*-GeSiSiGe, *g*-GeSiSiGe, *n*-GeSiGeGe and *g*-GeSiGeGe, and *i*-Si(SiGeGe), as shown in Figure 1. Our results indicate that the observed spectra of both the primary products and the minor impurity compounds are completely accounted for by a linear admixture of the *n*-, *g*-, and *i*-type conformations.

Initial basis set convergence tests showed that a full resolution of differences in the calculated properties of the conformational/positional isomers requires an even more stringent treatment than that used in the prior studies on (H₃Ge)_xSiH_{4-x} hydrides, particularly with regards to calculated vibrational properties. Accordingly all of the calculations described in the present work were carried out at the B3LYP/6-311G++(3df,2pd) level as implemented in the Gaussian03 code.¹¹

The results of the structural optimizations, obtained using “tight” convergence criteria, are presented in **Figure 12**. The figure lists the bond length, bond angle and bond torsion data for *n*-GeSiSiGe, *g*-GeSiSiGe, *n*-GeSiGeGe, *g*-GeSiGeGe and *i*-Si(SiGeGe). In all molecules, the skeletal structure exhibits typical Si-Si and Si-Ge bonds with the gauche conformations exhibiting slightly (0.002 Å) longer/shorter Si-Si/Si-Ge bonds than their linear counterparts. In the iso-butane like isomer *i*-Si(SiGeGe) the Si-Si bond lengths (2.351 Å) are found to be intermediate to those in the other isomers while the Si-Ge bond is dilated to a value 2.400 Å. All the Si-Si values, 2.350-2.352 Å, correspond closely (~0.1 %) to that found in bulk silicon $a\sqrt{3}/4 = 2.352$ Å where *a* (5.431 Å) is the lattice constant of crystalline Si in the diamond structure. The Ge-Ge values found in *n*-GeSiGeGe and *g*-GeSiGeGe, 2.446 Å, is also close to that of bulk germanium (2.449 Å). The heteronuclear skeletal Si-Ge bond lengths exhibit a distribution in the range 2.396-2.400 Å and represent an average of the Si-Si and Ge-Ge values (2.399 Å).

The Si-H and Ge-H bond lengths occur exclusively as central SiH₂ moieties and terminal GeH₃ groups in the *n*- and *g*- GeSiSiGe isomers, with values of 1.486 Å and 1.539 Å, respectively. In the corresponding GeSiGeGe isomers the Si-H bonds associated with the SiH₂ moiety has essentially the same value, 1.485 Å, while the
5 GeH₂ moiety exhibits a slightly longer Ge-H bond (1.542-1.548 Å) than in a GeH₃ terminal group, as expected. Due to its unique structure the *i*-Si(SiGeGe) isomer possesses both terminal SiH₃ and GeH₃ groups and a single SiH moiety. Here the Ge-H bond lengths associated with the terminal GeH₃ groups have the same value (1.539 Å) as in GeH₃ groups of other isomers, while the corresponding terminal group Si-H₃
10 bond lengths are slightly contracted (1.483 Å). The longest Si-H bond length is associated with the central SiH moiety and has a value 1.488 Å. Collectively, these bond length trends are consistent with those found previously in the heaviest members of the (H₃Ge)_NSiH_{4-N} (N=3,4) family of molecules.

The data presented in **Figure 12** also shows that the inter-metallic skeletal
15 bond angles, as well as those involving hydrogen-metal bonds, are essentially unchanged between all isomers studied. The only exception is found in the H-Ge-Si bond distributions, which possess several values in the linear *n*-GeSiSiGe isomer and a unique value in all of the others. The Ge-Si-Si-Ge bond torsion in the linear molecules is 180°, indicating that the metal atoms are confined to a common plane,
20 while the gauche conformations exhibit torsion angles of 66° and 64° in Ge-Si-Si-Ge and Ge-Si-Ge-Ge, respectively. It is also evident from the data that the inter-metallic and metal-hydrogen bond lengths in the *i*-Si(SiGeGe) isomer are, on average, closer to ideal tetrahedral values than those predicted for the other conformational/positional isomers.

Figure 12 also lists the thermochemical energies of the molecules, including
25 the total ground state electronic energy E_0 , and its value corrected for zero point energy correction (E_{ZPE}), thermal energy (E_{CORR}), enthalpy (H_{CORR}) and free energy (G_{CORR}) at 298K. In the case of the Ge₂Si₂H₁₀ molecules our calculations indicate that *i*-Si(SiGeGe) isomer is slightly favored at room temperature, with a free energy
30 difference of 0.0008 Hartree or about 2.0 kJ/mol (note: $k_B T @ 300\text{K} \sim 2.4 \text{ kJ/mol}$) relative to *n*-GeSiSiGe. The latter molecule's free energy is, in turn, more stable than that of *g*-GeSiSiGe by about 2.8 kJ/mol. For the GeSiGeGe impurity compound the *n*-GeSiGeGe isomer free energy is lower than that of its gauche conformation by 0.0020 Hartree (5.3 kJ/mol).

Vibrational spectra: In order to facilitate the interpretation of the observed FTIR spectra we calculated the vibrational frequencies and intensities of the conformational isomers *n*-GeSiSiGe and *g*-GeSiSiGe, and positional isomer *i*-Si(SiGeGe) of the primary product, as well as those of the minor species *n*-GeSiGeGe and *g*-GeSiGeGe, using the B3LYP DFT functional at the B3LYP/6-311G++(3df,2pd) level. No symmetry was imposed in the calculation of the frequency spectra and all molecules studied exhibited a positive definite spectrum of harmonic frequencies indicating that the ground state structures are dynamically stable.

Ge₂Si₂H₁₀ molecules: Figures 2 shows plots of the calculated spectra of *n*-GeSiSiGe, *g*-GeSiSiGe and *i*-Si(SiGeGe), convoluted with a Gaussian of width $\sim 20 \text{ cm}^{-1}$ to simulate experimental broadening. Both the high- and low-frequency regions are shown. It should also be noted that frequency scaling typically employed to reconcile the slight frequency overestimates obtained using this level of theory, has not been applied to the calculated spectra shown in these plots. The interpretation of the high-frequency spectral features (Figure 2) is straightforward: the lower frequency pairs of peaks designated n_6 and n_7 in *n*-GeSiSiGe (Figure 2(d)), g_{11} and g_{12} in *g*-GeSiSiGe (Figure 2(e)), and i_{11} and i_{12} in *i*-Si(SiGeGe) (Figure 2(f)), correspond to symmetric and anti-symmetric Ge-H stretching vibrations, respectively. Similarly, the symmetric and anti-symmetric Si-H stretching vibrations are designated n_8 and n_9 in *n*-GeSiSiGe (Figure 2(d)), g_{13} and g_{14} in *g*-GeSiSiGe (Figure 2(e)), and i_{14} and i_{15} in *i*-Si(SiGeGe) (Figure 2(f)), respectively. The calculated splitting ($13\text{--}14 \text{ cm}^{-1}$) between the symmetric and anti-symmetric bands (in either Si-H or Ge-H) is approximately the same in all isomers, however, all of these bands are systematically lower ($\sim 4\text{--}5 \text{ cm}^{-1}$) in frequency in the gauche isomer. In addition to these features the high-frequency spectrum of the positional isomer *i*-Si(SiGeGe), shown in Figure 2(f), also exhibits an isolated vibrational feature near 2077 cm^{-1} (designated by i_{13} in Figure 2(f)), which corresponds to the central Si-H stretching vibration.

The low-frequency non-skeletal vibrational structure, also shown in Figure 2, is considerably more complex and involves in- and out-of-plane, symmetric and antisymmetric Si-H/Ge-H wagging vibrations. The most striking difference between the spectra of *n*-GeSiSiGe and *g*-GeSiSiGe (parts (a) and (b) of Figure 2, respectively) is that four distinct features ($n_2\text{--}n_5$) of the linear isomer are effectively split in its gauche counterpart. For example, the strongest band in *n*-GeSiSiGe, n_2 , corresponding

to highly symmetric Si-H wagging vibrations parallel with the molecular backbone, is split into three bands denoted by g_3 , g_5 and g_6 in the gauche isomer exhibiting very strong Si-H wagging parallel to the Si-Si bond. Among these g_5 involves the least “crowded” protons in the molecule and accordingly exhibits the greatest intensity.

5 The weak shoulder indicated by g_4 in the *g*-GeSiSiGe spectrum is the anti-symmetric counterpart to mode g_5 . Finally, the distinct feature near 925 cm^{-1} in the *n*-GeSiSiGe, n_5 , corresponds to Si-H wagging vibrations perpendicular to the Si-Si axis involving symmetric H-Si-H bending motion. In the gauche isomer this band is split into symmetric and anti-symmetric counterparts denoted by g_9 and g_{10} . Only a few bands
10 appear to be common to both isomers. In particular features ($n_1:g_1$) near 430 cm^{-1} (asymmetrical wagging of the terminal Ge-H protons), ($n_3:g_7$) near 800 cm^{-1} (symmetrical terminal Ge-H wagging parallel to the Si-Si bond) and ($n_4:g_8$) near 876 cm^{-1} (terminal Ge-H wagging perpendicular to the Si-Ge bonds). The weak band designated g_2 in the Figure 2(b) involves symmetrical terminal Ge-H wagging, and
15 has a silent counterpart in the *n*-GeSiSiGe molecule.

The low-frequency spectrum of *i*-Si(SiGeGe), shown in Figure 2(c), has a slightly simpler structure due the molecule’s higher symmetry. The lowest frequency wagging modes found in the *g*-GeSiSiGe $\{g_1, g_2\}$ are effectively split into asymmetrical pairs $\{i_1, i_2\}$, and $\{i_3, i_4\}$, respectively, in *i*-Si(SiGeGe). The dominant
20 features near 800 cm^{-1} (i_7 and i_8) correspond to in-phase and anti-phase symmetrical Ge-H wags of the terminal GeH_3 groups while feature i_9 is the asymmetrical analog to these vibrations. In the SiH_3 terminal group the Si-H wagging vibrations (i_{10}) occur at $\sim 951\text{ cm}^{-1}$. Finally, features i_5 and i_6 are unique to the *i*-Si(SiGeGe) molecule, and involve perpendicular Si-H wags of the SiH moiety.

25 **SiGe₃H₁₀ molecules:** For the *n*- and *g*- (GeH_3)₂(SiH₂GeH₂) isomers we corroborate the NMR characterization by directly comparing the observed IR spectrum with that calculated for a $\sim 1:3$ mixture of *n*-GeSiGeGe and *g*-GeSiGeGe conformations (a detailed explanation is given in the following section). The calculated IR spectra for both the *n*- and *g*- isomer, convoluted with a Gaussian of width $\sim 20\text{ cm}^{-1}$ to simulate
30 experimental broadening, are given in Figure 3. Both the high- and low-frequency regions are shown (no frequency scaling has been applied). Peaks in the simulated spectrum originating from *n*-GeSiGeGe or *g*-GeSiGeGe conformations are denoted by N_i and G_i , respectively.

The high-frequency Si-H stretching bands are similar to those found in the GeSiSiGe analogs, exhibiting calculated splitting of 13-14 cm^{-1} between the symmetric and anti-symmetric bands and a systematic overall $\sim 4\text{-}5 \text{ cm}^{-1}$ shift to lower frequency in the gauche isomer. Here (N_{11} : G_{17}) and (N_{10} : G_{16}) correspond to asymmetric and symmetric Si-H stretches, respectively. However, a simple comparison for corresponding Ge-H stretching vibrations for these isomers is vitiated by the lower symmetry of GeSiGeGe isomers. For example the asymmetrical Ge-H stretch (N_9) in *n*-GeSiGeGe is split into contributions G_{12} , G_{14} and G_{15} in the gauche isomer while the symmetric Ge-H stretch, N_8 in the *n*-GeSiGeGe, is split into bands G_{11} and G_{13} .

In the low-frequency range of the spectrum we find a strong correspondence between mode assignments of the *n*- and *g*- isomers above 700 cm^{-1} . Specifically, the asymmetrical H-Si-H wagging vibrations N_3 in *n*-GeSiGeGe appears as G_6 in *g*-GeSiGeGe, with a frequency near 731 cm^{-1} . As in our earlier discussion we designate such corresponding pairs between isomer conformations as (N_3 : G_6). Intense Ge-H wagging vibrations (N_4 : G_7) associated with terminal GeH_3 groups appear near 808 cm^{-1} . Similarly, H-Ge-H "scissor" modes at 866 cm^{-1} are designated as (N_5 : G_8), while their H-Si-H counterparts (N_7 : G_{10}) appear at 936 cm^{-1} (the frequency of the corresponding mode in the GeSiSiGe isomers was $\sim 925 \text{ cm}^{-1}$). The features denoted (N_6 : G_9) consist of asymmetrical "scissor"-like vibrations associated with the terminal GeH_3 groups.

The most significant differences between the IR spectra of the *n*- and *g*- isomers are found below 700 cm^{-1} . The most intense spectral feature in *n*-GeSiGeGe (N_2) is associated with a highly symmetric wagging of the protons bound to the GeH_2 moiety parallel to the molecular backbone. In the gauche isomer this mode is split into three features denoted G_3 , G_4 and G_5 , corresponding to asymmetrical Ge-H wagging of combinations of individual protons in the GeH_2 moiety. Finally, the mode N_1 in the *n*- isomer, consisting of concerted symmetrical wagging of terminal GeH_3 protons at both ends of the molecule, is split into two modes G_1 and G_2 in the gauche isomer in which symmetrical wagging occurs independently in the two GeH_3 terminal groups.

Isomer Admixture Analysis: To estimate the molar concentration of the conformational and positional isomers in the observed spectra, we adopt an approach commonly employed in "multi-phase" fitting of x-ray spectra, in which a normalized

admixture of known pure-phase candidate spectra is fitted to an unknown observed spectrum of the combined molecules. Using the same procedures as in our prior studies on Si-Ge hydrides we have calculated the spectra of 10 digermyl-disilyl positional/conformational isomers (to be described elsewhere). These calculated spectra serve as a proxy for the unknown pure isomer candidate spectra normally supplied by experiment. A comparison of our calculated spectra with that of the synthesized compounds suggests the presence of simple admixtures of conformational $\{n\text{-GeSiSiGe}, g\text{-GeSiSiGe}\}$ and positional $\{i\text{-Si}(\text{SiGeGe}), g\text{-GeSiSiGe}\}$ isomers for the $\text{Si}_2\text{Ge}_2\text{H}_{10}$ family. To proceed with the analysis all of the required theoretical spectra (n -, g - and iso - isomers, see Figure 2) are first calculated on a fine frequency grid (1 cm^{-1} spacing), and then partitioned into low-frequency ($0\text{-}1600 \text{ cm}^{-1}$) and high-frequency ($1600\text{-}2600 \text{ cm}^{-1}$) regions. We denote the low-frequency spectra of the $\{n\text{-GeSiSiGe}, g\text{-GeSiSiGe}\}$ pair by $\{I_n(\omega), I_g(\omega)\}$ and those of $\{i\text{-Si}(\text{SiGeGe}), g\text{-GeSiSiGe}\}$ by $\{I_i(\omega), I_g(\omega)\}$. The corresponding observed spectrum of the synthesized compound, $I_{\text{exp}}(\omega)$, produced on a discrete frequency grid $\{\omega_k\}$ by the spectrometer, is then used to construct an objective function (Eq 9). For example, in the case of the $\{n\text{-GeSiSiGe}, g\text{-GeSiSiGe}\}$ admixture:

$$O(\beta, \gamma, \eta) = \sum_k \left\{ I_{\text{exp}}(\omega_k) - \beta \left[\gamma I_n(\eta\omega_k) + (1-\gamma) I_g(\eta\omega_k) \right] \right\}^2 \quad (9)$$

where γ is the normalized admixture of n - and g - spectra, β is the global amplitude of the theoretical spectrum and η is the common frequency scale factor applied to the theoretical spectra. An identical procedure is used to treat the $\{i\text{-Si}(\text{SiGeGe}), g\text{-GeSiSiGe}\}$ combination. In all cases the theoretical spectra were evaluated at arbitrary frequencies using cubic splines, and a robust simplex method was used to minimize the objective function. For the $\{n\text{-GeSiSiGe}, g\text{-GeSiSiGe}\}$ pair this yields a scale factor of $\eta = 0.984$ and an admixture of $\gamma = 0.392$ (i.e., $\sim 39\%$ $n\text{-GeSiSiGe}$). This optimized scale factor is very similar to the value 0.989 used in previous studies of the $(\text{H}_3\text{Ge})_x\text{SiH}_{4-x}$ hydrides⁷ and Si_2H_6 , Ge_2H_6 and H_3GeSiH_3 molecules¹² at this level of theory. As shown in Figure 4(a) the resulting low-frequency composite theoretical spectrum ($\omega < 1600 \text{ cm}^{-1}$), generated using this scale factor, agrees remarkably well with the observed one. To simulate the composite high-frequency spectrum for $\{n\text{-GeSiSiGe}, g\text{-GeSiSiGe}\}$ we employ the admixture parameter derived from low-frequency behavior, $\gamma = 0.392$, and reduce the scale factor to $\eta' = (0.9909\eta) = 0.975$.

This low- to high-frequency scale factor ratio was previously found to yield an excellent account of the scaling behavior in related molecules¹². Figure 4(b) demonstrates that the optimized parameters obtained using fits to the low-frequency spectra also faithfully and independently reproduce the high-frequency behavior, confirming the integrity of our procedure. Application of the foregoing low-frequency analysis to the positional isomeric combination $\{i\text{-Si}(\text{SiGeGe}), g\text{-GeSiSiGe}\}$ yields $\eta = 0.985$ and $\gamma = 0.678$ (i.e., $\sim 68\%$ $i\text{-Si}(\text{SiGeGe})$). The sample spectrum is compared with the calculated admixture in Figure 5, which indicates that the vibrational structure is largely accounted for using the admixture procedure.

The corresponding results for the GeSiGeGe impurity compound are shown in Figure 6. A completely analogous fitting procedure to that above yields, in this case, an admixture parameter $\gamma = 0.235$ (i.e., 23.5 % $n\text{-GeSiGeGe}$) indicating that the gauche conformation ($g\text{-GeSiGeGe}$) is dominant in the gas phase. As in the GeSiSiGe analysis, we find that fitting parameters obtained from the low-frequency spectrum ($\omega < 1600 \text{ cm}^{-1}$) yields an excellent account of the high-frequency stretching bands using the reduced frequency scale factor 0.975 (Figure 6(b)).

We have observed that the proportions derived for the conformational isomers are essentially constant between different samples. To elucidate this observation we calculated the potential energy surface (PES) of the butane-like GeSiSiGe as a function of the Ge-Si-Si-Ge backbone torsion angle. The linear conformation corresponds a torsion angle of 180° while the gauche modification occurs at $\sim 66^\circ$ (see Figure 12). For each fixed value of the torsion angle all remaining structural degrees of the freedom are optimized resulting in a “relaxed” PES. Figure 7 compares the torsional PES for butane, tetrasilane and GeSiSiGe at the DFT B3LYP/6-311G++(3df,2pd) and *ab initio* CCSD/LANL2DZ levels of theory. In all cases the value of the linear or *trans*- configuration is chosen as the reference energy. The plots reveal that all three molecules exhibit very similar energy-torsion profiles with a global minimum at 180° , a $n\text{-g}$ barrier ($E_{n\text{-g}}$) at $\sim 120^\circ$, a local minimum at the gauche configuration (E_g) near 66° and a maximum value corresponding to the eclipse saddle point (E_e). Using the DFT B3LYP/6-311G++(3df,2pd) description we obtain $\Delta E_{n\text{-g}} = 0.942$, $\Delta E_g = 0.362$ and $\Delta E_e = 1.628$ kcal/mol for GeSiSiGe, while the corresponding tetrasilane values are 0.895, 0.360 and 1.591 kcal/mol, respectively. At this level of theory the conformational energy structure of GeSiSiGe and butane-like

tetrasilane is essentially identical. This is not unexpected in view of their very similar bond character. At the same level of description we find that the analogous PES for butane shares the same qualitative form as that of tetrasilane and GeSiSiGe, but with significantly larger rotational barriers $\Delta E_{n-g} = 3.235$, $\Delta E_g = 0.884$ and $\Delta E_e = 5.713$ kcal/mol.

It should be noted that our B3LYP/6-311G++(3df,2pd) tetrasilane values for ΔE_{n-g} and ΔE_e are qualitatively different from those calculated by Albinsson *et al*¹³ at the MP2/6-31* level for the same molecule. In particular the latter authors predict a symmetric *n-g* barrier and a near identity of gauche and linear conformation energy ($\Delta E_{n-g} \sim 0.588$, $\Delta E_g \sim -0.04$ and $\Delta E_e \sim 1.236$ kcal/mol). Unfortunately, because the basis set levels and model chemistries in our two approaches both differ, the origin of this discrepancy is ambiguous. We therefore carried out a series of coupled cluster calculations at the CCSD/LANL2DZ level for the sequence butane, tetrasilane and GeSiSiGe. The results, shown as solid lines in Figure 7, are remarkably consistent with the B3LYP-based DFT calculations (dashed lines in the figure), and indicate that the linear butane-like conformation is the ground state structure, with a gauche value of $\Delta E_g \sim 0.2$ - 0.4 kcal/mol. Furthermore, the magnitude of the *barrier* connecting the gauche and linear conformations in GeSiSiGe is reduced to 0.36 kcal/mol (~ 188 K) in the CCSD/LANL2DZ calculations. This small value of the torsional barriers in tetrasilane and GeSiSiGe is consistent with the facile low-temperature conformational interconversions observed experimentally.

It should be noted that the 68% value obtained for the positional isomeric spectrum, shown in Figure 5, is characteristic of *one* particular sample. Unlike the conformational isomers, which exhibited a constant proportion between samples, the proportions of the positional isomers sensitively depend on the specific reaction conditions. This suggests that once the $\{i\text{-Si}(\text{SiGeGe}), g\text{-GeSiSiGe}\}$ mixture has been isolated no inter-conversion between positional isomers occurs. Chain cleavage and spontaneous isomerization are observed in the closely related tetrasilane Si_4H_{10} system during photochemical matrix isolation experiments producing trisilane (Si_3H_8) and silylene (SiH_2).¹³ The formation *i*- Si_4H_{10} observed as a mixture with *n*- Si_4H_{10} is then believed to follow from SiH_2 insertion into the Si-H bonds of trisilane. As discussed in detail by Jasinski *et al.* this insertion reaction is one of the characteristic mechanisms in gas phase polysilane chemistry, possessing a slight *negative* activation

energy and a very high collisional efficiency.¹⁴ We surmise that an analogous low energy germylene insertion reaction channel is present in our synthesis route (see Eq.2) involving the triflate-substituted disilanes at $T > 22^\circ\text{C}$, leading to the condensation of *i*-Si(SiGeGe). One plausible mechanism could involve the replacement of $-\text{SO}_3\text{CF}_3$ by GeH_3 at the 1-substitution site of $(\text{SO}_3\text{CF}_3)_2(\text{SiH}_2)_2$ (**5**), followed by insertion of GeH_2 in the Si-H bonds of the central SiH_2 site in the $-\text{SiH}_2-\text{SiH}_2-\text{GeH}_3$ intermediate.

Deposition studies

The deposition experiments were conducted on a clean Si(100) substrates in ultra-high vacuum system equipped with a low energy electron microscope (LEEM) that allowed *in situ* real time observation of the growth process. The $(\text{H}_3\text{Ge})_2(\text{SiH}_2)_2$ gaseous precursor reacted on the Si surface via complete H_2 elimination at partial pressures in the range of 10^{-5} - 10^{-8} Torr. Initially, the reaction growth kinetics was investigated using LEEM and the activation energy (E_{act}) of the compound with respect to H_2 desorption from the Si surface was measured. It is known that for conventional silanes and germanes the E_{act} is much lower on a Ge containing surface (including $\text{Si}_{1-x}\text{Ge}_x$) than on pristine Si. Thus, an accurate determination of E_{act} can only be obtained from measuring the growth rate of the first monolayer, (which grows directly on pure Si) and not that of subsequent layers, which grow on Ge containing surfaces and have higher growth rates. In this regard LEEM is particularly unique because its dynamic imaging differentiates the $\text{Si}(001)-(2\times 1)$ surface from the $\text{Si}_{1-x}\text{Ge}_x$ epilayer and thus provides unambiguous means for the correct determination of E_{act} on Si by measuring the growth rate of the first $\text{Si}_{1-x}\text{Ge}_x$ monolayer vs. temperature.

For comparison the growth rates of Ge_2H_6 and H_3GeSiH_3 were also determined using the same method. For these and for $(\text{H}_3\text{Ge})_2(\text{SiH}_2)_2$ the first monolayer took longer to complete on pristine Si than subsequent monolayers. Figure 8 is a graph showing plots of the temperature dependence of the first layer growth rates. The activation energies were determined using the relation $R \sim \exp(-E_a/kT)$. The data are consistent with the first-order H_2 desorption kinetics and yield E_{act} of 1.95, 1.66 and 1.42 eV, for H_3GeSiH_3 , Ge_2H_6 and $(\text{H}_3\text{Ge})_2(\text{SiH}_2)_2$, respectively. This indicates that the latter it is not only more reactive than H_3GeSiH_3 , but is also more reactive than pure Ge_2H_6 despite the presence of strong Si-H bonds in the molecular structure. Accordingly, it represents a unique high-reactivity species suitable for low

temperature (300-400 °C) rapid growth of thick layers (100-500 nm) with stoichiometry $\text{Si}_{0.50}\text{Ge}_{0.50}$ reflecting that of the corresponding precursor. Below we describe the deposition procedure and key results regarding morphology, microstructure, composition and strain for the films.

5 The synthesis of $\text{Si}_{0.5}\text{Ge}_{0.5}$ films was investigated in the 350°C - 500°C range via dehydrogenation of $(\text{H}_3\text{Ge})_2(\text{SiH}_2)_2$. The highest temperature that yields planar layers is ~ 450°C. The growth rate is ~ 4 nm/min and decreases to 2 nm/min at 350°C. No appreciable growth was observed below 300°C (1×10^{-5} Torr) due to the reduced reactivity of the compound. Thick films with rough surfaces and defective
10 microstructures were produced at 500 °C with growth rates as high as 15 nm/min. The composition and heteroepitaxial character of the films were investigated by random and channeling Rutherford backscattering (RBS), which in all cases indicate a highly stoichiometric material aligned with the substrate. The ratio of the aligned versus the random peak heights (χ_{min}), which measures the degree of crystallinity,
15 decreases from 15 % at the interface to 6% at the surface indicating a dramatic reduction in dislocation density across the thickness of the film. This is confirmed by high resolution XTEM, which revealed commensurate SiGe/Si interfaces and perfectly monocrystalline layer microstructures. A survey of TEM samples showed no defects penetrating through the layers within a field of view of ~ 2.5 μm in the
20 micrographs indicating an upper limit in the threading dislocation density of ~ $10^6/\text{cm}^2$ (Figure 9). Most of the defects due to the lattice mismatch are concentrated near the interface region. In fact, a number of these appear to originate at the SiGe/Si interface and show a propensity to propagate downward into the substrate (Figure 9). Occasional edge dislocations spaced by ~150 nm are also observed in the vicinity of
25 steps on the Si surface (Figure 10). . These are parallel to the interface plane and may provide stress relief between the film and the substrate. The presence of edge dislocations is unusual since SiGe-based materials typically display (111)-threading dislocations and stacking faults. X-ray diffraction (XRD) revealed highly aligned layers with mosaics spreads as low as 0.1 degrees and relatively strain free
30 microstructures. Raman scattering experiments conducted with several Ar+ laser lines revealed a high degree of vertical uniformity in concentration and strain.⁸ The Raman analysis indicated that the composition agrees well with the RBS value and that the

strain relaxation is 75% and 95 % for films with thickness of 200 and 500 nm, respectively which similar to those obtained by XRD.

Atomic force microscopy (AFM) studies show that films with thickness of 50 and 750 nm display RMS values of 0.5 and 1.5 nm, respectively, for $10 \times 10 \mu\text{m}^2$ areas, indicating highly planar surfaces for samples grown at $T < 450^\circ\text{C}$ as shown Figure 11(a). For layers grown at $T > 450^\circ\text{C}$ the AFM images [Figure 11(b)] show a classic “cross hatched” surface pattern caused by misfit dislocations produced during strain relaxation. It has been shown that both defect induced interface and bulk alloy scattering considerably reduces the electron mobilities in these samples and degrades the device performance.¹⁵ Nevertheless conventional CVD at ultra low pressure is known to produce $\text{Si}_{1-x}\text{Ge}_x$ materials with significantly lower cross hatching patterns leading to higher quality electrical properties. In contrast to these methodologies, which typically involve low-order silanes and germanes (SiH_4 , Si_2H_6 , GeH_4 , Ge_2H_6), the present route utilizes the highly reactive and more massive $(\text{H}_3\text{Ge})_2(\text{SiH}_2)_2$ compound. The direct incorporation of the Si_2Ge_2 molecular unit into the film leads to low surface diffusion compared to the conventional low mass compounds. Furthermore, the high reactivity permits low growth temperature and promotes high growth rates at pressures of $\sim 10^{-5}$ Torr leading to device quality films devoid of cross hatched surface morphologies.

The thermal properties of the films were investigated between 450°C and 750°C , a range well within device processing temperatures. The XRD lattice constant, the RBS χ_{min} values, and the AFM roughness were measured for the annealed samples and compared with the corresponding values of the as-grown materials. Samples with a thickness as low as 200 nm did not show any increase in surface roughness even after annealing at 750°C indicating that the surface planarity of our samples is thermally robust. The XRD and RBS aligned spectra were essentially identical before and after annealing. This result also confirms full relaxation in the as-grown layers at $T < 450^\circ\text{C}$.

30 Conclusion

We have demonstrated the synthesis of the butane-like compounds with molecular formulas $(\text{GeH}_3)_2(\text{SiH}_2)_2$ (1) $(\text{GeH}_3)_2\text{SiH}(\text{SiH}_3)$ (2) and $(\text{GeH}_3)_2(\text{SiH}_2\text{GeH}_2)$ (3) and their conformational isomers. Detailed characterization of the products was

conducted via a range of spectroscopic and analytical methods including quantum chemical simulations. A new fitting procedure was developed, based on the admixture of normal and gauche calculated spectra, to obtain the proportions of the individual conformers in the product sample at 22°C. Controlled depositions of (1) have produced stoichiometric SiGe films possessing the desired properties for semiconductor applications including perfectly crystalline and epitaxial microstructures, smooth morphologies, and uniformly strain-relaxed states. Unique practical advantages associated with this low temperature (~ 350-450°C) rapid growth include: (i) short deposition times compatible with preprocessed Si wafers, (ii) selective growth for application in high frequency devices, and (iii) negligible mass segregation of dopants, which is particularly critical for thin layers. Finally we have found that these compounds are remarkably stable over extended periods of time, with evidence of decomposition, and thus represent viable molecular sources for potential industrial applications.

Experimental Section

All manipulations were carried out under inert conditions using standard Schlenk and drybox techniques. Dry, air-free solvents were distilled from either anhydrous CaCl₂ or sodium benzophenone ketyl prior to use. All NMR spectra were collected either on a Varian INOVA 400 spectrometer operating at 400 MHz or a Gemini 300 spectrometer. Samples were dissolved in CDCl₃ or C₆D₆ and all nuclei were referenced either directly or indirectly to the proton signal of TMS or the residual solvent peak as indicated. IR spectra were recorded using a 10 cm gas cell with KBr windows. Elemental analyses were performed by Desert Analytics (Tucson, Arizona). Gas chromatography mass spectrometry (GCMS) data were obtained using a JEOL JMS-GC Mate II spectrometer. Powdered lithium (Aldrich), lithium tetrahydroaluminate (Aldrich), trichloro(*p*-tolyl)silane (Gelest), trifluoromethane sulfonic acid (Alfa Aesar), nonafluorobutane-1-sulfonic acid (Aldrich) and electronic grade germane gas (Voltaix, Inc.) were used as received. The starting materials *p*-tolylsilane, chloro(*p*-tolyl)silane, 1,2-*bis*(*p*-tolyl)disilane and 1,2-*bis*(trifluoromethylsulfonyloxy)disilane were prepared according to literature procedures and their purity checked by NMR spectroscopy. Potassium germyl was synthesized in monoglyme using sodium-potassium (80% K) alloy. The coupling

reactions using powdered lithium were performed in a helium environment rather than in an N₂ atmosphere to prevent formation of lithium nitride.

5 (p-tolyl)₂(SiH₂)₂. A solution of monochloro(p-tolyl)silane (10.0 g, 63.8 mmol) in di-n-butyl ether (120 mL) and toluene (40 mL) is stirred with sodium pieces (1.68 g, 146.2 mmol) for 5 h at 125 °C. The mixture is filtered and the filtrate is distilled to remove the solvents and other impurities. The remaining colorless liquid product (4.40 g, ~ 60 %) was examined by NMR to determine its identity and purity and was subsequently used without purification.

10

(C₄F₉SO₃)₂(SiH₂)₂: A sample of C₄F₉SO₃H (7.9 g, 26.3 mmol) was added to a 15 mL solution of 1,2-(p-tolyl)disilane (3.18 g, 13.1 mmol) in toluene at -40 °C. A colorless precipitate was immediately formed at low temperatures. The reaction flask was warmed to 23 °C and stirred for 1-2 h. The colorless solid was filtered and dried *in vacuo*. Repeated concentration and cooling of the toluene filtrate at -20 °C yielded additional product. Yield = 7.9 g (95 %). Mp = 68 °C. ¹H NMR (300 MHz, CDCl₃): δ 5.078 (s, ¹J_{SiH} = 272 Hz, Si-H₂, ³J_{HH} = 3.6 Hz). ²⁹Si NMR (79.5 MHz, CDCl₃): δ -30.1. Anal. Calcd for C₈F₁₈S₂O₆Si₂H₄: C, 14.59; H, 0.61; F, 51.95. Found: C, 14.27; H, 0.87; F, 51.74.

20

(H₃Ge)₂(SiH₂)₂, **method A**: A 20 mL solution of (C₄F₉SO₃)₂(SiH₂)₂ (4.67 g, 9.5 mmol) was added to a slurry of KGeH₃ (2.00 g, 20.9 mmol) in diethyl ether (15 mL), at -40 °C. The reaction was then warmed to -10 °C and stirred under nitrogen for 12h after which it was heated at 35 °C for an additional 1h. The volatiles were fractionally distilled several times through U-traps held at -20, -50 and -196 °C. The -50 °C trap retained pure (H₃Ge)₂(SiH₂)₂ (260 mg, 26 % Yield) while the -196 °C trap contained solvent, and traces GeH₄ and SiH₄. Vapor pressure: 7.0 Torr (23 °C). IR (gas, cm⁻¹): 2147 (s), 2073 (vs), 910 (w), 879 (w), 799 (s), 740 (vw), 682 (s), 651 (m), 442 (vw). ¹H NMR (400 MHz, C₆D₆, 7.15 ppm): δ 3.106 (t, J = 4.1 Hz, 6H, Ge-H₃), δ 3.290 (q, J = 3.9 Hz, 4H, Si-H₂). ²⁹Si NMR (79.5 MHz, C₆D₆): δ -105.0 (this value is close to that observed for (H₃Ge)₂SiH₂ at -102.45) GCMS: *m/z* 210-196 (Si₂Ge₂H_x⁺), 180-170 (SiGe₂H_{8-x}⁺), 152-143 (Ge₂H_{6-x}⁺), 140-126 (Si₂GeH_{8-x}⁺), 107-98 (H₃SiGeH₃⁺), 77-72 (GeH₄⁺), 62-58 (Si₂H_{6-x}⁺), 32-28 (SiH₄⁺).

30

(H₃Ge)₂(SiH₂)₂, method B: A sample of (CF₃SO₃)₂(SiH₂)₂ (3.39 g, 9.5 mmol) was added to a slurry of 2.40 g (20.9 mmol) of KGeH₃ in diethyl ether (30 mL) at -30 °C. The mixture was then warmed to 23 °C and stirred under nitrogen for 5h. The volatiles were fractionally distilled and the (H₃Ge)₂(SiH₂)₂ product in its pure form was collected at a -50 °C trap (0.250 g, ~13% yield).

(SiH₂GeH₂)(GeH₃)₂: The compound is obtained as a byproduct in the synthesis of (H₃Ge)₂(SiH₂)₂, and was isolated by repeated distillation through U-traps under static vacuum maintained at -20 and -50 °C to collect SiH₂GeH₂(GeH₃)₂ and (H₃Ge)₂(SiH₂)₂, respectively. Vapor pressure: 1.0 Torr (22 °C). IR (gas, cm⁻¹): 2145 (m), 2073 (vs), 910 (w), 878 (w), 793 (vs), 723 (w), 679 (w), 615 (s), 443 (vw). ¹H NMR (400 MHz, C₆D₆, 7.15 ppm): δ 3.30 (sept, *J* = 3.9 Hz, 2H, Si-H₂), δ 3.23 (t, *J* = 4.6 Hz, 3H, Ge-H₃), δ 3.13 (t, *J* = 3.90 Hz, 3H, Ge-H₃), δ 3.03 (sept, *J* = 4.0 Hz, 2H, Ge-H₂). ²⁹Si NMR (79.5 MHz, C₆D₆): δ -98.2. GCMS: *m/z* 256-240 (SiGe₃H_{10-x}⁺), 230-212 (SiGe₂H_{8-x}⁺), 184-169 (SiGe₂H_{8-x}⁺), 154-141 (Ge₂H_{6-x}⁺), 136-128 (Si₂GeH_{8-x}⁺), 108-98 (SiGeH_{6-x}⁺), 77-70 (GeH_{4-x}⁺), 32-28 (SiH₄⁺).

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We claim:

1. A compound comprising a silicon-germanium hydride selected from the group consisting of
 $(\text{H}_3\text{Ge})_2(\text{SiH}_2)_2$; and
5 $(\text{GeH}_3)_2\text{SiH}(\text{SiH}_3)$.
2. The compound of claim 1 wherein the silicon germanium hydride is substantially isolated $(\text{H}_3\text{Ge})_2(\text{SiH}_2)_2$.
3. The compound of claim 2, further comprising $(\text{GeH}_3)_2(\text{SiH}_2\text{GeH}_2)$.
4. The compound of claim 1 wherein the silicon germanium hydride comprises
10 $(\text{H}_3\text{Ge})_2(\text{SiH}_2)_2$ and $(\text{GeH}_3)_2\text{SiH}(\text{SiH}_3)$.
5. A composition, comprising one or more compounds according to claim 1 in an inert gas.
6. The composition of claim 5, wherein the inert gas is selected from the group consisting of H_2 , He, N_2 , and argon.
- 15 7. A method of synthesizing one or more of the silicon-germanium hydrides of claim 1, comprising combining a perfluoroalkyl-sulfonate-substituted disilane with a compound comprising a GeH_3 ligand under conditions whereby the silicon-germanium hydride is formed.
8. The method of claim 7 wherein the compound comprising a GeH_3 ligand is
20 selected from the group consisting of alkali and alkaline earth germanides.
9. The method of claim 7 wherein the method comprises combining the compound comprising a GeH_3 ligand with a nonafluorobutane-sulfonic substituted disilane, and wherein the nonafluorobutane-sulfonic substituted disilane comprises
 $(\text{SO}_3\text{C}_4\text{F}_9)_2(\text{SiH}_2)_2$.

10. The method of claim 7 wherein the method comprises combining the compound comprising a GeH_3 ligand with a triflate substituted disilane, and wherein the triflate substituted disilane comprises $(\text{SO}_3\text{CF}_3)_2(\text{SiH}_2)_2$.
11. The method of claim 10, wherein the compound comprising a GeH_3 ligand is KGeH_3 , and wherein the silicon germanium hydride is formed according to the equation
- $$(\text{SO}_3\text{CF}_3)_2(\text{SiH}_2)_2 + 2\text{KGeH}_3 \rightarrow (\text{H}_3\text{Ge})_2(\text{SiH}_2)_2 + 2\text{KSO}_3\text{CF}_3.$$
12. The method of claim 11, wherein the method comprises
- (a) adding $(\text{CF}_3\text{SO}_3)_2(\text{SiH}_2)_2$ to a slurry of KGeH_3 in solvent at between -40°C . and -20°C . to form a mixture;
 - (b) warming the mixture to between 0°C . and 35°C .;
 - (c) stirring the mixture under nitrogen; and
 - (d) distilling the mixture to form the $(\text{H}_3\text{Ge})_2(\text{SiH}_2)_2$.
13. The method of claim 12 wherein the reaction is carried out at a temperature of -10°C . or less.
14. The method of claim 12 wherein the reaction is carried out at a temperature of 22°C . or higher.
15. The method of claim 7, wherein the compound comprising a GeH_3 ligand is KGeH_3 , and wherein the silicon germanium hydride is formed according to the equation
- $$(\text{SO}_3\text{C}_4\text{F}_9)_2(\text{SiH}_2)_2 + 2\text{KGeH}_3 \rightarrow (\text{H}_3\text{Ge})_2(\text{SiH}_2)_2 + 2\text{KSO}_3\text{C}_4\text{F}_9.$$
16. The method of claim 15, wherein the method comprises
- (a) adding $(\text{C}_4\text{F}_9\text{SO}_3)_2(\text{SiH}_2)_2$ to a slurry of KGeH_3 in solvent at between -50°C . and -30°C . to form a mixture;
 - (b) warming the mixture to between -20°C . and 0°C .;
 - (c) stirring the mixture under nitrogen;
 - (d) heating the mixture to between 25°C . and 45°C .; and
 - (e) distilling the mixture to form the $(\text{H}_3\text{Ge})_2(\text{SiH}_2)_2$.

17. A method for depositing a Si-Ge material on a substrate in a reaction chamber, comprising introducing into the chamber a gaseous precursor comprising the compound of claim 1, under conditions whereby a layer comprising a Si-Ge material is formed on the substrate.
- 5 18. The method of claim 17 wherein the step of introducing the gaseous precursor comprises introducing the gaseous precursor in substantially pure form.
19. The method of claim 17 wherein the step of introducing the gaseous precursor comprises introducing the gaseous precursor as a single gas source.
20. The method of claim 17 wherein the step of introducing the gaseous precursor
10 comprises introducing the gaseous precursor intermixed with an inert carrier gas.
21. The method of claim 20 wherein the inert carrier gas comprises H₂.
22. The method of claim 20 wherein the inert carrier gas comprises N₂.
23. The method of claim 17 wherein the gaseous precursor is deposited by gas source molecular beam epitaxy.
- 15 24. The method of claim 17 wherein the gaseous precursor is deposited by chemical vapor deposition.
25. The method of claim 17 wherein the gaseous precursor is introduced at a temperature of between 300° C. and 450° C.
26. The method of claim 17 wherein the gaseous precursor is introduced at a partial
20 pressure between 10⁻⁸ Torr and 1000 Torr.
27. The method of claim 17 wherein the Si-Ge material is formed on the substrate as a strain-relaxed or uniformly strained layer having a planar surface.
28. The method of claim 17 wherein the composition of the Si-Ge material is substantially uniform.
- 25 29. The method of claim 17 wherein the entire Si and Ge framework of the gaseous precursor is incorporated into the Si-Ge material or epitaxial Si-Ge.

30. The method of claim 17 wherein the gaseous precursor is deposited by plasma enhanced chemical vapor deposition.
31. The method of claim 17 wherein the gaseous precursor is deposited by laser assisted chemical vapor deposition.
- 5 32. The method of claim 17 wherein the gaseous precursor is deposited by atomic layer deposition.
33. The method of claim 17 wherein the gaseous precursor comprises $(\text{H}_3\text{Ge})_2(\text{SiH}_2)_2$.
34. The method of claim 17, wherein the substrate comprises germanium, silicon on
10 insulator, Ge:Sn alloys, SiO_2 , sapphire, quartz, Si:Ge alloys, Si:C alloys, stainless steel, and polyimide films.
35. The method of claim 34, wherein the substrate comprises silicon.
36. A method for depositing an epitaxial Si—Ge layer on a substrate, comprising introducing near a surface of the substrate a gaseous precursor comprising the
15 compound of claim 1; and
dehydrogenating the precursor under conditions whereby epitaxial Si—Ge is formed on the substrate.
37. The method of claim 36 wherein the step of introducing the gaseous precursor comprises introducing the gaseous precursor in substantially pure form.
- 20 38. The method of claim 36 wherein the step of introducing the gaseous precursor comprises introducing the gaseous precursor as a single gas source.
39. The method of claim 36 wherein the step of introducing the gaseous precursor comprises introducing the gaseous precursor intermixed with an inert carrier gas.
40. The method of claim 39 wherein the inert carrier gas comprises H_2 .
- 25 41. The method of claim 39 wherein the inert carrier gas comprises N_2 .
42. The method of claim 36 wherein the gaseous precursor is deposited by gas

source molecular beam epitaxy.

43. The method of claim 36 wherein the gaseous precursor is deposited by chemical vapor deposition.

44. The method of claim 36 wherein the gaseous precursor is introduced at a
5 temperature of between 300° C. and 450° C.

45. The method of claim 36 wherein the gaseous precursor is introduced at a partial pressure between 10^{-8} Torr and 1000 Torr.

46. The method of claim 36 wherein the Si—Ge material is formed on the substrate as a strain-relaxed or uniformly strained layer having a planar surface.

10 47. The method of claim 36 wherein the composition of the Si—Ge material is substantially uniform.

48. The method of claim 36 wherein the entire Si and Ge framework of the gaseous precursor is incorporated into the Si—Ge material or epitaxial Si—Ge.

49. The method of claim 36 wherein the gaseous precursor is deposited by plasma
15 enhanced chemical vapor deposition.

50. The method of claim 36 wherein the gaseous precursor is deposited by laser assisted chemical vapor deposition.

51. The method of claim 36 wherein the gaseous precursor is deposited by atomic layer deposition.

20 52. The method of claim 36 wherein the gaseous precursor comprises $(\text{H}_3\text{Ge})_2(\text{SiH}_2)_2$.

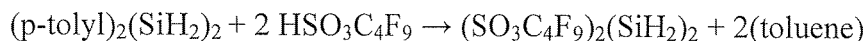
53. The method of claim 36, wherein the substrate comprises germanium, silicon on insulator, Ge:Sn alloys, SiO_2 , sapphire, quartz, Si:Ge alloys, Si:C alloys, stainless steel, and polyimide films.

25 54. A compound comprising $(\text{SO}_3\text{C}_4\text{F}_9)_2(\text{SiH}_2)_2$.

55. A method for synthesizing $(\text{SO}_3\text{C}_4\text{F}_9)_2(\text{SiH}_2)_2$ comprising combining $\text{R}(\text{SiH}_2)_2$ and $\text{HSO}_3\text{C}_4\text{F}_9$ under conditions whereby the $(\text{SO}_3\text{C}_4\text{F}_9)_2(\text{SiH}_2)_2$ is formed, where R is an aryl substituent such as p-tolyl or benzyl.

56. A method for synthesizing $(\text{SO}_3\text{C}_4\text{F}_9)_2(\text{SiH}_2)_2$ comprising combining (p-
5 tolyl) $_2(\text{SiH}_2)_2$ and $\text{HSO}_3\text{C}_4\text{F}_9$ under conditions whereby the $(\text{SO}_3\text{C}_4\text{F}_9)_2(\text{SiH}_2)_2$ is formed.

57. The method of claim 56, wherein the $(\text{SO}_3\text{C}_4\text{F}_9)_2(\text{SiH}_2)_2$ is formed according to the equation



10 58. The method of claim 56 wherein the method comprises

(a) combining $\text{HSO}_3\text{C}_4\text{F}_9$, (p-tolyl) $_2(\text{SiH}_2)_2$, and solvent in a reaction chamber at between -50°C . and 0°C . to form a mixture;

(b) warming the mixture to between 20°C . and 25°C .;

(c) stirring the mixture; and

15 (d) filtering the solid $(\text{SO}_3\text{C}_4\text{F}_9)_2(\text{SiH}_2)_2$.