

Molecular sieve SCM-15, a process of preparing same and use thereof

Technical Field

The invention relates to a molecular sieve SCM-15, a process of
5 preparing same and use thereof.

Background

In industry, porous inorganic materials are widely used as catalysts
and catalyst supports. A porous material has a relatively high specific
10 surface and an open channel structure, and is therefore a good catalytic
material or catalyst support. The porous material may generally comprise:
amorphous porous materials, crystalline molecular sieves, modified
layered materials, and the like. The subtle differences in the structures of
these materials are indicatives of significant differences in their catalytic
15 and adsorption properties, as well as differences in the various observable
properties used to characterize them, such as the morphology, specific
surface area, porosity and variability of these dimensions.

The basic skeleton structure of crystalline microporous zeolite is
based on a rigid three-dimensional TO_4 (SiO_4 , AlO_4 , etc.) unit structure; in
20 which structure, TO_4 shares oxygen atoms in a tetrahedral structure, and
the charge balance of the skeleton tetrahedron, such as AlO_4 , is maintained
through the presence of surface cations such as Na^+ and H^+ . It can be seen
that the properties of the zeolite can be altered by cation exchange. At the
same time, there exists abundant pores with uniform opening in the
25 structure of a zeolite. These pores are interlaced to form a
three-dimensional network structure, and the skeleton can still be stably
retained after the removal of the occluded water or organic species (US
4439409). Based on the above structure, zeolite not only has good catalytic
activity, excellent shape-selection, but also has good selectivity by
30 modification (US 6162416, US 4954325, and US 5362697) in various
organic reactions.

The specific structure of a molecular sieve is determined by an X-ray diffraction spectrum (XRD), and the X-ray diffraction spectrum (XRD) is measured by an X-ray powder diffractometer using a Cu-K α ray source with a nickel filter. Different zeolite molecular sieves have different XRD
5 spectrums. Known molecular sieves, such as zeolite A (US 2882243), zeolite Y (US 3130007), PSH-3 molecular sieve (US 4439409), ZSM-11 molecular sieve (US 3709979), ZSM-12 molecular sieve (US 3832449), ZSM-23 molecular sieve (US 4076842), ZSM-35 molecular sieve (US 4016245), MCM-22 molecular sieve (US 4954325), etc. each have XRD
10 spectra of their respective characteristics.

At the same time, zeolites with same characteristic XRD spectrum but different types of skeleton atoms will be considered as different molecular sieves. For example, TS-1 molecular sieve (US 4410501) and ZSM-5
15 molecular sieve (US 3702886) have the same characteristic XRD spectrums, but have different skeleton elements. Specifically, the TS-1 molecular sieve comprises skeleton elements of Si and Ti, exhibiting a catalytic oxidation ability, while the ZSM-5 molecular sieve comprises skeleton elements of Si and Al, exhibiting an acidic catalytic ability.

In addition, molecular sieves with the same characteristic XRD
20 spectrum and the same types of skeleton elements but with different relative amounts of the skeleton elements, will be identified as different molecular sieves as well. For example, X zeolite (US 2882244) and Y zeolite (US 3130007), share the same characterizing XRD pattern and the same types of skeleton elements (Si and Al), but with different relative
25 amounts of Si and Al. Specifically, X zeolite has a Si/Al molar ratio of less than 1.5, while Y zeolite has a Si/Al molar ratio of higher than 1.5.

Summary of the invention

The inventors have made deep study on the basis of the prior art and
30 found a novel molecular sieve SCM-15, and further identified beneficial properties thereof.

Specifically, the present invention relates to a molecular sieve SCM-15, characterized in that the molecular sieve has an X-ray diffraction spectrum as substantially shown in the table below.

2θ ($^{\circ}$) ^(a)	d-distance ($\text{\AA}$)	Relative intensity ($I/I_0 \times 100$)
4.811	18.424 ± 1.148	w
6.571	13.468 ± 0.614	m-vs
6.888	12.847 ± 0.559	m-s
9.675	9.143 ± 0.283	m-vs
9.894	8.941 ± 0.270	m-vs
19.888	4.462 ± 0.067	w-m

5 $a = \pm 0.3^{\circ}$.

The present invention also provides a process of preparing the molecular sieve SCM-15.

10 **Technical effect**

According to the present invention, the SCM-15 molecular sieve involved has a skeleton structure which has never been obtained before in the art.

15 **Description of drawings**

Figure 1 is an X-ray diffraction spectrum (XRD) of the molecular sieve (in the prepared form) obtained in Example 1.

Figure 2 is an X-ray diffraction spectrum (XRD) of the molecular sieve (in the prepared form) obtained in Example 8.

20 Figure 3 is an X-ray diffraction spectrum (XRD) of the molecular sieve (in the calcined form) obtained in Example 8.

Embodiments

The embodiments of the present invention will be illustrated in detail below, while it should be understood that the protection scopes of the present invention are not restricted thereto; instead, the protection scopes are defined by the claims attached.

5 All publications, patent applications, patents and other references mentioned in this specification are hereby incorporated by reference. Unless otherwise defined, the scientific and technical terms used in the specification have the meanings conventionally known by those skilled in the art. For conflicting meanings of the terms, they shall be understood by
10 the definitions of the present specification.

When the present specification mentions a material, substance, method, step, device, or component, etc. with the derivative words "known to those skilled in the art", "prior art" or the like, the term derived is intended to cover those conventionally used in the field of the present
15 application, but also cover those that are not currently known, whilst will become known in the art to be useful for the similar purposes.

In the context of this specification, the term "specific surface area" refers to the total area of a unit weight of sample, including the internal surface area and the external surface area. Non-porous samples have only
20 external surface areas, such as silicate cement, some clay mineral particles, etc.; while porous samples have both external and internal surface areas, such as asbestos fibers, diatomaceous earth and molecular sieves. The surface area of pores having pore diameters less than 2 nm in a porous sample is the internal surface area, the surface area after deducting the
25 internal surface area from the total surface area is referred to as the external surface area, and the external surface area per unit weight of a sample is the external specific surface area.

In the context of this specification, the term "pore volume" refers to the volume of pores per unit weight of porous material. The term "total
30 pore volume" refers to the volume of all pores (generally only the pores having a pore diameter of less than 50 nm) per unit weight of molecular sieve. The term "micropore volume" refers to the volume of all micropores

(generally the pores having a pore diameter of less than 2 nm) per unit weight of molecular sieve.

In the context of the present specification, in the XRD data of a molecular sieve, w, m, s, and vs represent the intensities of the diffraction peaks, wherein w represents weak, m represents medium, s represents strong, and vs represents very strong, which are known to those skilled in the art. In general, w is less than 20; m is 20-40; s is 40-70; and vs is greater than 70.

In the context of the present specification, the structure of a molecular sieve is determined by X-ray diffraction (XRD), wherein the X-ray diffraction spectrum (XRD) of the molecular sieve is collected using an X-ray powder diffractometer equipped with a Cu-K α ray source, with K α 1 wavelength $\lambda = 1.5405980$ angstroms ($\text{\AA}$) and a nickel filter.

In the context of the present specification, the so-called prepared state, prepared form or prepared molecular sieve refers to the state of the molecular sieve after the completion of the preparation. As the prepared state, a specific example may be the state directly presented after completion of the preparation (generally called as a precursor of molecular sieve), or a state after organic substances (particularly, an organic template agent) which may be present in the pores of the molecular sieve precursor are further removed by a method other than calcination. Thus, in the prepared state, the molecular sieve may contain water, may contain organic substances, or may be free of water or organic substances.

In the context of this specification, the term "calcined", "calcined form" or "calcined molecular sieve" refers to the state of the molecular sieve after calcination. As the state after calcination, for example, may be a state obtained by further removing organic substances (particularly, organic template agents) and water, etc. that may be present in the pores of the prepared molecular sieve by calcination. Here, the conditions of the calcination include, in particular: calcinating at 550°C for 6 hours in an air atmosphere.

It should be particularly understood that two or more of the aspects (or embodiments) disclosed in the context of this specification can be combined with each other as desired, and that such combined embodiments (e.g., methods or systems) are incorporated herein and constitute a part of this original disclosure, while remaining within the scope of the present invention.

Without otherwise specifically indicated, all percentages, parts, ratios, etc. mentioned in this specification are provided by weight, unless the basis by weight is not in accordance with the conventional knowledge of those skilled in the art.

According to an aspect of the present invention, the invention relates to molecular sieve SCM-15. The molecular sieve, particularly in its prepared form or calcined form, has an X-ray diffraction spectrum substantially as shown in Table A-1 or Table A-2 below.

TABLE A-1

2θ ($^{\circ}$) ^(a)	d-distance ($\text{\AA}$) ^(b)	Relative intensity ($I/I_0 \times 100$)
4.811	18.353	w
6.571	13.441	m-vs
6.888	12.823	m-s
9.675	9.134	m-vs
9.894	8.933	m-vs
19.888	4.461	w-m

(a) = $\pm 0.3^{\circ}$, (b) is a function of 2θ .

TABLE A-2

2θ ($^{\circ}$) ^(a)	d-distance ($\text{\AA}$)	Relative intensity ($I/I_0 \times 100$)
4.811	18.424 ± 1.148	w
6.571	13.468 ± 0.614	m-vs
6.888	12.847 ± 0.559	m-s
9.675	9.143 ± 0.283	m-vs

9.894	8.941±0.270	m-vs
19.888	4.462±0.067	w-m

(a) = ±0.3°.

According to an aspect of the present invention, the X-ray diffraction spectrum may further comprise X-ray diffraction peaks substantially as shown in Table B-1 or Table B-2 below.

TABLE B-1

2θ (°) ^(a)	d-distance (Å) ^(b)	Relative intensity (I/I ₀ X 100)
8.225	10.741	w
15.592	5.679	w
22.031	4.031	w-m
22.435	3.960	w-m

(a) = ±0.3°, (b) is a function of 2θ.

TABLE B-2

2θ (°) ^(a)	d-distance (Å) ^(b)	Relative intensity (I/I ₀ X 100)
8.225	10.755±0.392	w
15.592	5.681±0.109	w
22.031	4.032±0.054	w-m
22.435	3.960±0.052	w-m
25.330	3.514±0.041	w

(a)=±0.3°.

According to an aspect of the present invention, the X-ray diffraction spectrum optionally further comprises X-ray diffraction peaks substantially as shown in the following Table.

2θ (°) ^(a)	d-distance (Å) ^(b)	Relative intensity (I/I ₀ X 100)
14.294	6.194±0.129	w

16.909	5.241±0.0.092	w
19.156	4.630±0.0.072	w-m
19.517	4.546±0.0.069	w-m
21.162	4.196±0.0.059	w-m

(a) = ±0.3°.

According to an aspect of the present invention, the molecular sieve SCM-15 has a schematic chemical composition as shown with the formula "SiO₂·GeO₂". It is known that molecular sieves sometimes contain a certain amount of moisture and organics (especially organic template agents), particularly immediately after preparation, but it is not considered necessary to specify the amount of the moisture and organics in the present invention because the presence or absence of this moisture and organics does not substantially affect the XRD spectrum of the molecular sieve. In view of this, the schematic chemical composition represents in fact the chemical composition of the molecular sieve excluding water and organics. Moreover, it is apparent that the schematic chemical composition represents the framework chemical composition of the SCM-15 molecular sieve, or alternatively, may represent the schematic chemical composition of the calcined SCM-15 molecular sieve.

According to an aspect of the invention, in the molecular sieve SCM-15, the molar ratio of silicon to germanium satisfies SiO₂/GeO₂ ≥ 1, preferably 1 ≤ SiO₂/GeO₂ ≤ 15, preferably 1 ≤ SiO₂/GeO₂ ≤ 10, more preferably 2.5 ≤ SiO₂/GeO₂ ≤ 5.

According to an aspect of the invention, in the molecular sieve SCM-15, germanium on the sekeleton may be partially replaced by a trivalent or tetravalent element other than silicon and germanium, with a replacement ratio not exceeding 10%. Here, the parameter "replacement ratio" is dimensionless. When germanium is replaced by a trivalent element, such as boron or aluminum, the replacement ratio = 2X₂O₃/(2X₂O₃+ GeO₂) × 100%, wherein X is a trivalent element. When germanium is replaced by a tetravalent element, such as tin, zirconium, or

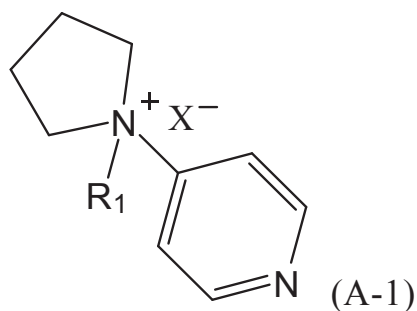
titanium, the replacement ratio = $\text{YO}_2/(\text{YO}_2+\text{GeO}_2) \times 100\%$, wherein Y is a tetravalent element. In calculating the replacement ratio, the moles of the corresponding oxide are used.

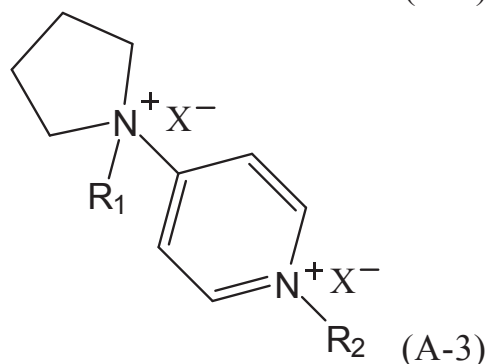
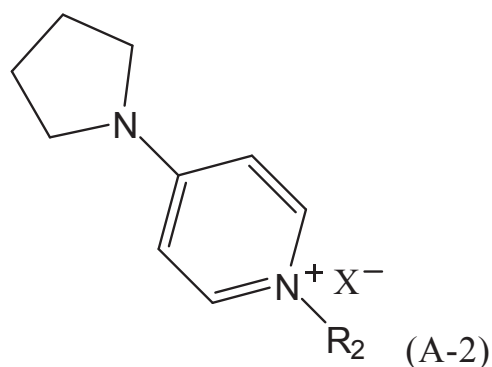
According to an aspect of the invention, the molecular sieve SCM-15 has a specific surface area (according to BET method) of 100-600 m^2/g , preferably 130-500 m^2/g , more preferably 200-400 m^2/g .

According to an aspect of the invention, the molecular sieve SCM-15 has a micropore volume (according to t-plot method) of 0.04-0.25 cm^3/g , preferably 0.05-0.20 cm^3/g , more preferably 0.09-0.18 cm^3/g .

According to an aspect of the present invention, the molecular sieve SCM-15 can be prepared by the following processes. In view of this, the present invention also relates to a process of preparing a molecular sieve SCM-15, comprising the step of: crystallizing a mixture (hereinafter called as mixture) comprising or formed from a silicon source, a germanium source, a fluorine source, an organic template agent and water, to obtain said molecular sieve.

According to an aspect of the present invention, in the process of preparing the molecular sieve, the organic template agent is selected from 4-pyrrolidiny pyridine, or a quaternary ammonium form represented by formula (A-1), formula (A-2) or formula (A-3), preferably 4-pyrrolidiny pyridine. These organic Template agents may be used alone or as a combination in a desired ratio.





In each formula, R_1 and R_2 are each independently H or C_{1-8} alkyl, preferably C_{1-4} alkyl, more preferably C_{1-2} alkyl, and X^- are each independently a halogen ion (such as Cl^- , Br^- , and I^-) and a hydroxide ion (OH^-), preferably hydroxide ion (OH^-).

According to an aspect of the present invention, in the process of preparing the molecular sieve, the crystallization step may be performed in any manner conventionally known in the art, such as a method of mixing the silicon source, the germanium source, the fluorine source, the organic template agent and water in a given ratio, and hydrothermally crystallizing the obtained mixture under the crystallization conditions. Stirring may be applied as required.

According to an aspect of the present invention, in the process of preparing the molecular sieve, any silicon source conventionally used in the art for this purpose may be used as the silicon source. Examples thereof include silicic acid, silica gel, silica sol, tetraalkyl orthosilicate, and water glass. These silicon sources may be used alone or as a combination in a desired ratio.

According to an aspect of the present invention, in the process of preparing the molecular sieve, any germanium source conventionally used

in the art for this purpose may be used as the germanium source, including but not limited to germanium oxide, germanium nitrate, and tetraalkoxy germanium.

5 According to an aspect of the present invention, in the process of preparing the molecular sieve, any fluorine source conventionally used for this purpose in the art may be used as the fluorine source, and examples thereof include fluoride or an aqueous solution thereof, particularly hydrofluoric acid and the like.

10 According to an aspect of the present invention, in the process of preparing the molecular sieve, the molar ratio of the silicon source (calculated as SiO_2), the germanium source (calculated as GeO_2), the fluorine source (calculated as F), the organic template agent and water is generally from 1:(0-1):(0.1-2.0):(0.1-2.0):(3-30); preferably 1:(1/15-1.5):(0.2-1.5):(0.2-1.5):(4-25); more preferably 1:(0.1-0.5):(0.4-1.2):
15 (0.4-1.2):(5-20); more preferably 1:(0.2-0.4):(0.6-1.0):(0.6-1.0):(5-15).

According to an aspect of the invention, in the process of preparing the molecular sieve, the crystallization conditions include: a crystallization temperature of 131-210°C, preferably 150-190°C, more preferably 160-180 °C; and a crystallization duration of 1-20 days, preferably 2-10 days, more
20 preferably 2-7 days.

According to an aspect of the present invention, in the process of preparing the molecular sieve, an ageing step before crystallization is included, and the ageing conditions include: an ageing temperature of 50-90°C, and an ageing duration of 2 hours to 2 days.

25 According to an aspect of the present invention, in the process of preparing the molecular sieve, when germanium atoms are replaced with trivalent or tetravalent elements other than silicon and germanium, a source of the trivalent or tetravalent elements other than silicon and germanium, preferably a source of oxide of the trivalent or tetravalent elements other
30 than silicon and germanium, is added to the mixture. As the source of oxide, at least one selected from the group consisting of a boron oxide source, an aluminum oxide source, a tin oxide source, a zirconium oxide

source, and a titanium oxide source is preferable. Specific examples of the aluminum oxide source include at least one selected from the group consisting of aluminum hydroxide, sodium aluminate, aluminum salt, kaolin and montmorillonite. Specific examples of the boron oxide source include at least one selected from the group consisting of boron oxide, borax, sodium metaborate, and boric acid. Specific examples of the tin oxide source include at least one selected from the group consisting of tin tetrachloride, stannous chloride, alkyl tin, alkoxy tin, and organic stannates. Specific examples of the zirconia source include at least one selected from the group consisting of zirconium salts (e.g., zirconium nitrate or zirconium sulfate), alkyl zirconium, alkoxy zirconium, and organic zirconates. Specific examples of the titanium oxide source include one or more selected from tetraalkyl titanates (e.g., tetramethyl titanate, tetraethyl titanate, tetrapropyl titanate, tetra-n-butyl titanate), TiCl_4 , hexafluorotitanic acid, $\text{Ti}(\text{SO}_4)_2$, and hydrolysis products thereof.

According to an aspect of the present invention, in the process of preparing the molecular sieve, the molar ratio of the oxide source (calculated as the corresponding oxide) to the germanium source (calculated as GeO_2) when used is generally (0.01-0.1): 1, preferably (0.02-0.08): 1.

According to an aspect of the present invention, in the process of preparing the molecular sieve, after the crystallization is completed, the molecular sieve can be separated as a product from the obtained reaction mixture by any separation methods conventionally known, thereby obtaining the molecular sieve SCM-15, which is also called as prepared form of molecular sieve SCM-15. The separation method includes, for example, a method of filtering, washing and drying the obtained reaction mixture.

According to an aspect of the present invention, in the process of preparing the molecular sieve, the filtering, washing and drying may be performed by any method conventionally known in the art. Specifically, for example, the reaction mixture obtained may be simply filtered by

suction. Examples of the washing include washing with deionized water. The drying temperature is, for example, 40 to 250°C, preferably 60 to 150°C, and the drying duration is, for example, 8 to 30 hours, preferably 10 to 20 hours. The drying may be carried out under normal pressure or under
5 reduced pressure.

According to an aspect of the invention, in the process of preparing the molecular sieve, the molecular sieve SCM-15 may not be calcined. Nevertheless, if needed, organics (particularly the organic template agent) possibly existing in the molecular sieve can be removed by a UV/ozone
10 method, thereby obtaining an organic-free molecular sieve. Such molecular sieve also belongs to the molecular sieve SCM-15 according to the present invention, which is also called as the prepared form of molecular sieve SCM-15. The UV/ozone method is known in the art, and a specific example thereof comprises placing a molecular sieve at a distance of 2 to 3
15 mm from an ultraviolet lamp, irradiating the molecular sieve for 12 to 48 hours with ultraviolet light having a wavelength of 184 to 257 nm. The ultraviolet light can be generated by a low-pressure mercury lamp or a medium-pressure mercury lamp (10-20 mW • cm⁻²) sealed in a box.

According to an aspect of the present invention, in the process of
20 preparing the molecular sieve, the molecular sieve obtained by crystallization may also be calcined, as needed, to remove the organic template agent and if any, the water, etc., and thus to obtain the calcined molecular sieve, which is also called as calcined form of molecular sieve SCM-15. The calcination may be carried out in any manner conventionally
25 known in the art, for example, the calcination temperature is generally 300 to 750°C, preferably 400 to 600°C, and the calcination duration is generally 1 to 10 hours, preferably 3 to 6 hours. In addition, the calcination is generally carried out in an oxygen-containing atmosphere, such as air or oxygen atmosphere.

30 According to an aspect of the present invention, the molecular sieve SCM-15 may be in any physical form, such as a powder, granule, or molded article (e.g., a bar, clover, etc.). These physical forms can be

obtained in any manner conventionally known in the art and are not particularly limited.

According to an aspect of the present invention, the molecular sieve SCM-15 may be used in combination with other materials, thereby obtaining a molecular sieve composition. Examples of the other materials include active materials and inactive materials. Examples of the active material include synthetic zeolite and natural zeolite, and examples of the inactive material (generally called as a binder) include clay, carclazyte, and alumina. These other materials may be used alone or as a combination in any ratio. The amounts of the other materials can refer to those conventionally used in the art, without particular limitation.

According to an aspect of the present invention, the molecular sieve SCM-15 or the molecular sieve composition may be used as an adsorbent, for example to separate at least one component from a mixture of a plurality of components in the gas or liquid phase. Thus, at least one component may be partially or substantially completely separated from the mixture of the plurality of components by contacting the mixture with said molecular sieve SCM-15 or said molecular sieve composition, so as to selectively adsorb such a component.

According to an aspect of the present invention, the molecular sieve SCM-15 or the molecular sieve composition may also be used as a catalyst (or as a catalytically active component thereof) either directly or after having been subjected to necessary treatments or conversions (such as ion exchange, etc.) conventionally performed in the art for molecular sieves. To this end, according to an aspect of the present invention, it is possible, for example, to subject a reactant (such as a hydrocarbon) to a given reaction in the presence of the catalyst, and thereby obtain a target product.

Examples

The present invention will be described in further detail with reference to examples, whilst the present invention is not limited to these examples.

Example 1

43.2 g of deionized water, 42.63 g of organic template agent of 4-pyrrolidinylpyridine (98 wt%), 8.37g of germanium oxide (99 wt%), 14.0g of hydrofluoric acid (40 wt%) and 60.0 g of silica sol (SiO₂ 40 wt%) were uniformly mixed to obtain a reaction mixture, wherein the material ratios (molar ratios) of the reaction mixture were as follows:

$$\text{SiO}_2/\text{GeO}_2=5$$

$$\text{Template agent}/\text{SiO}_2=0.70$$

$$\text{F}/\text{SiO}_2=0.70$$

$$\text{H}_2\text{O}/\text{SiO}_2=12$$

After being mixed uniformly, the mixture was loaded into a stainless steel reactor for aging in a 80 °C water bath for 2 hours, and then crystallizing at 170 °C for 5 days under stirring. After crystallization, the solution was filtered, washed and dried at 120 °C for 12 hours to obtain a molecular sieve.

The XRD spectrum data of the molecular sieve (in the prepared form) of the product was shown in Table 1, and the XRD spectrum was shown in FIG.1.

The product molecular sieve was measured by inductively coupled plasma atomic emission spectroscopy (ICP) to have a SiO₂/GeO₂=5.1.

TABLE 1

2 θ/°	d/Å	I/I ₀ 100
4.829	18.2824	2.4
6.594	13.393	65.4
6.889	12.8198	32
8.234	10.7289	13.6
9.691	9.1187	77.1
9.892	8.9341	100
10.296	8.5844	1
11.89	7.4371	1.2

13.147	6.7288	4.4
13.5	6.5536	2.4
14.215	6.2255	2.3
15.613	5.671	8.4
16.084	5.5059	1.4
16.541	5.355	6.2
16.931	5.2324	9.7
17.821	4.9731	2.8
19.199	4.6191	5.8
19.46	4.5577	11
19.911	4.4554	18.9
20.779	4.2713	9.7
21.115	4.2041	16.8
21.589	4.1129	11
22.042	4.0293	26.2
22.477	3.9523	21.8
23.066	3.8528	4
23.487	3.7846	9
23.957	3.7115	6.6
24.511	3.6288	3.8
24.98	3.5617	2.2
25.279	3.5202	6.8
25.633	3.4724	5.3
26.243	3.393	9.3
26.58	3.3507	10.1
27.903	3.1949	5.2
28.435	3.1362	3.8
28.692	3.1088	5
28.948	3.0818	6.6
29.385	3.037	3.7
30.11	2.9655	4.2

30.976	2.8845	1
31.374	2.8489	1.4
31.727	2.8179	2.7
32.007	2.7939	2.2
32.419	2.7594	2.1

Example 2

Example 1 was repeated except that the reaction mixture was prepared in the following ratios (molar ratios):

- 5 $\text{SiO}_2/\text{GeO}_2=4$
- Template agent/ $\text{SiO}_2=0.60$
- $\text{F}/\text{SiO}_2=0.60$
- $\text{H}_2\text{O}/\text{SiO}_2=10$

After being mixed uniformly, the mixture was loaded into a stainless
10 steel reactor for aging in a 80°C water bath for 3 hours, and then crystallizing at 165°C for 4 days under stirring.

The XRD spectrum data for the product (in the prepared form) was shown in Table 2, and the XRD spectrum was similar to that of FIG.1.

The product molecular sieve was measured by inductively coupled
15 plasma atomic emission spectroscopy (ICP) to have a $\text{SiO}_2/\text{GeO}_2= 4.2$.

TABLE 2

$2\theta/^\circ$	$d/\text{\AA}$	$I/I_0 \times 100$
4.817	18.33	2.3
6.578	13.4255	60.6
6.873	12.8496	32.1
8.292	10.6542	12.3
9.674	9.1353	76.2
9.889	8.9367	100
10.295	8.5851	1.3
11.898	7.4319	1.1

13.128	6.7384	4.4
13.52	6.544	2
14.224	6.2217	2.2
15.612	5.6713	9.9
16.083	5.5064	1.2
16.521	5.3614	6
16.912	5.2383	9.8
17.88	4.9567	2.2
19.163	4.6277	5.1
19.44	4.5624	10.2
19.873	4.4639	17.6
20.76	4.2751	10
21.096	4.2078	18.9
21.569	4.1166	9.4
22.023	4.0327	28.2
22.495	3.9491	22.7
23.011	3.8618	3.4
23.485	3.7849	9.5
23.937	3.7144	6.5
24.505	3.6296	4.2
25.276	3.5207	6.9
25.631	3.4726	5.5
26.245	3.3929	7.2
26.543	3.3554	10.2
27.918	3.1931	5.1
28.418	3.1381	3
28.671	3.111	5.1
28.91	3.0858	5.2
29.381	3.0374	3.3
30.09	2.9674	4.3
30.979	2.8842	1.5

31.71	2.8194	2.5
31.984	2.7959	1.9
32.576	2.7464	2
33.326	2.6863	2.1
34.786	2.5768	2.3
35.219	2.5461	3
35.834	2.5038	2.7

Example 3

Example 1 was repeated except that the reaction mixture was prepared in the following ratios (molar ratios):

- 5 $\text{SiO}_2/\text{GeO}_2=4$
Template agent/ $\text{SiO}_2=0.66$
 $\text{F}/\text{SiO}_2=0.66$
 $\text{H}_2\text{O}/\text{SiO}_2=8$

10 After being mixed uniformly, the mixture was loaded into a stainless steel reactor for crystallizing at 170°C for 5 days under stirring.

The XRD spectrum data for the product (in the prepared form) was shown in Table 3, and the XRD spectrum was similar to that of FIG.1.

The product molecular sieve was measured by inductively coupled plasma atomic emission spectroscopy (ICP) to have a $\text{SiO}_2/\text{GeO}_2= 4.4$.

TABLE 3

$2\theta/^\circ$	$d/\text{\AA}$	I/I_0 100
4.739	18.6329	2.1
6.516	13.5541	85.9
6.81	12.9693	38
8.23	10.7341	20
9.579	9.225	74.3
9.811	9.0075	100
10.228	8.6416	5

11.618	7.6103	0.9
13.167	6.7185	5.7
14.147	6.2553	2.3
15.533	5.7	26.8
16.003	5.5338	2
16.441	5.3871	6.5
16.832	5.2629	9.8
17.703	5.0059	4.8
19.084	4.6466	5.2
19.358	4.5814	10.5
19.794	4.4816	21.5
20.664	4.2948	16.1
21.017	4.2235	21.5
21.492	4.1312	16.6
21.944	4.0471	32.3
22.34	3.9762	23.4
22.932	3.875	5.2
23.404	3.7978	10.5
23.838	3.7296	10.7
24.41	3.6435	7.5
24.908	3.5719	3.1
25.178	3.5341	7.2
25.674	3.4669	10.3
26.106	3.4105	18.9
26.462	3.3655	14.2
26.913	3.3101	1.3
27.823	3.2039	7.3
28.554	3.1234	6.7
28.867	3.0903	10.7
29.28	3.0476	3.3
29.973	2.9787	4.8

30.935	2.8883	1.3
31.606	2.8285	3.4
31.905	2.8027	2
32.32	2.7676	2
33.209	2.6955	2.4
34.076	2.6289	1.5
34.707	2.5825	2.5
35.16	2.5503	3.4
35.655	2.516	6.4
37.056	2.424	0.5
37.906	2.3716	1.4
38.238	2.3518	1.1
38.853	2.3159	1.7
39.442	2.2827	3.9
40.407	2.2304	2.5
41.791	2.1597	1.7
42.38	2.131	1.6
44.236	2.0458	1.3
44.669	2.027	1.2
45.106	2.0084	2.1
46.464	1.9527	2.6
46.76	1.9411	2.7
48.103	1.89	0.8
48.57	1.8729	0.5
49.149	1.8522	1.7
49.759	1.8309	3.1
51.144	1.7845	1.4
51.814	1.763	0.8
53.489	1.7117	0.8

Example 4

Example 1 was repeated except that the reaction mixture was prepared in the following ratios (molar ratios):

$\text{SiO}_2/\text{GeO}_2=5$

Template agent/ $\text{SiO}_2=0.60$

5 $\text{F}/\text{SiO}_2=0.60$

$\text{H}_2\text{O}/\text{SiO}_2=12$

After being mixed uniformly, the mixture was loaded into a stainless steel reactor for crystallizing at 150°C for 5 days under stirring.

The XRD spectrum data for the product (in the prepared form) was shown in Table 4, and the XRD spectrum was similar to that of FIG.1.

The product molecular sieve was measured by inductively coupled plasma atomic emission spectroscopy (ICP) to have a $\text{SiO}_2/\text{GeO}_2=5.3$.

TABLE 4

$2\theta/^\circ$	$d/\text{\AA}$	$I/I_0\ 100$
4.901	18.0158	1.2
6.712	13.1586	100
7.007	12.6053	24.3
8.372	10.552	37.9
9.797	9.0205	42.8
10.026	8.8154	63.7
10.421	8.4815	5.4
11.761	7.5182	1.3
13.399	6.6027	7.7
13.657	6.4786	6.1
14.382	6.1533	1.5
15.729	5.6293	20.1
16.208	5.464	1.8
16.698	5.3049	6
17.055	5.1947	9.8
18.038	4.9137	6.1

19.324	4.5894	6.5
19.634	4.5178	9.5
20.068	4.4211	24.8
20.859	4.255	11.6
21.391	4.1504	26.8
21.726	4.0872	33.5
22.182	4.0042	26.6
22.597	3.9317	17.8
23.191	3.8323	6
23.68	3.7542	9.6
24.015	3.7025	7.1
24.668	3.606	4.9
25.039	3.5535	4.9
25.417	3.5014	8.9
25.791	3.4514	10.6
26.323	3.383	3.1
26.835	3.3195	9.4
27.21	3.2746	2.3
28.097	3.1732	6.1
28.77	3.1005	6.7
29.047	3.0716	12.5
29.612	3.0143	3.6
30.249	2.9522	4.2
31.49	2.8386	3.6
31.892	2.8038	2.2
32.167	2.7804	2.1
32.619	2.7429	3.3
33.522	2.671	2.6
35.075	2.5562	1.5
35.733	2.5107	7.5
36.047	2.4895	3.9

36.566	2.4554	1.5
37.908	2.3715	1.1
38.276	2.3495	1.8
38.538	2.3341	1.1
39.224	2.2949	1.5
39.797	2.2631	3.1
40.701	2.2149	3.4
42.047	2.1471	1.5
42.759	2.113	1.3
43.965	2.0578	1.4
44.474	2.0354	2.1
45	2.0128	1.3
45.462	1.9935	1.5
46.78	1.9403	3.8
47.037	1.9303	4
48.342	1.8812	1.3
49.547	1.8383	1

Example 5

Example 1 was repeated except that the reaction mixture was prepared in the following ratios (molar ratios):

- 5 $\text{SiO}_2/\text{GeO}_2=3.5$
Template agent/ $\text{SiO}_2=0.72$
 $\text{F}/\text{SiO}_2=0.72$
 $\text{H}_2\text{O}/\text{SiO}_2=9$

After being mixed uniformly, the mixture was loaded into a stainless
10 steel reactor for crystallizing at 150°C for 5 days under stirring.

The XRD spectrum data for the product (in the prepared form) was shown in Table 5, and the XRD spectrum was similar to that of FIG.1.

The product molecular sieve was measured by inductively coupled plasma atomic emission spectroscopy (ICP) to have a $\text{SiO}_2/\text{GeO}_2= 3.9$.

TABLE 5

$2\theta/^\circ$	$d/\text{\AA}$	I/I_0 100
4.741	18.6215	1.6
6.555	13.4725	100
6.833	12.9252	25.3
8.234	10.7291	29.7
9.638	9.169	49.7
9.87	8.954	80.7
10.262	8.6126	6
13.186	6.7089	4.9
13.501	6.5531	4.1
14.207	6.2287	1.6
15.572	5.6859	12
16.578	5.3431	6.4
16.894	5.2436	6.9
17.822	4.9729	3
18.826	4.7098	2.4
19.221	4.6138	4.2
19.494	4.5497	6.4
19.894	4.4592	21.2
20.72	4.2833	10.9
21.45	4.1391	22.9
22.023	4.0328	19.3
22.44	3.9587	14.4
23.089	3.849	4.3
23.524	3.7788	12.9
23.914	3.718	3.3
24.454	3.6371	2.2
24.894	3.5737	1.6
25.242	3.5253	5

25.618	3.4744	6.9
26.208	3.3975	2.6
26.679	3.3386	9.8
27.96	3.1885	3
28.592	3.1194	6.8
28.907	3.0861	6.4
29.48	3.0274	2
30.346	2.9429	3
31.331	2.8526	1.7
31.734	2.8174	1.7
32.004	2.7942	1.3
32.46	2.756	2
33.404	2.6803	2.1
34.876	2.5704	0.9
35.319	2.5392	2.5
35.635	2.5174	3.6
35.968	2.4948	2.3
37.329	2.4069	1
38.334	2.3461	1.5
39.093	2.3023	1.2
39.596	2.2742	2.3
40.565	2.2221	2.4
42.559	2.1224	1.2
43.865	2.0623	0.9
44.312	2.0425	1.3
44.948	2.015	1
45.357	1.9978	0.9

Example 6

Example 1 was repeated, except that boric acid was added into the system as a boron source to replace a part of the germanium source, and the replacement ratio was 1%.

The XRD spectrum data for the product (in the prepared form) was shown in Table 6, and the XRD spectrum was similar to that of FIG.1.

The product molecular sieve was measured by inductively coupled plasma atomic emission spectroscopy (ICP) to have a $\text{SiO}_2/\text{GeO}_2=5.4$ and $\text{SiO}_2/\text{B}_2\text{O}_3=991.3$.

10 TABLE 6

$2\theta/^\circ$	$d/\text{\AA}$	I/I_0 100
4.822	18.3094	2.3
6.613	13.3545	100
6.892	12.8156	25.3
8.312	10.6289	29.4
9.678	9.1317	50
9.911	8.9171	80.2
10.322	8.5629	8.7
11.193	7.8986	0.7
11.629	7.6033	0.9
13.283	6.6601	6.8
13.52	6.544	4.3
14.233	6.2178	1.6
15.63	5.6647	21.3
16.085	5.5056	2.1
16.598	5.3366	5.1
16.916	5.2371	6.9
17.939	4.9405	4.8
19.239	4.6097	3
19.497	4.5491	7.2
19.933	4.4506	20.4

20.777	4.2716	14.4
21.214	4.1846	17.2
21.626	4.1058	24.8
22.043	4.0292	21.4
22.459	3.9554	16.5
23.167	3.8361	4.6
23.561	3.7729	11
23.915	3.7178	7.2
24.527	3.6264	3.8
25.005	3.5581	3.9
25.297	3.5178	6.8
25.789	3.4517	8.7
26.243	3.3931	4.8
26.64	3.3433	5.6
27.921	3.1929	3.3
28.613	3.1172	6.3
28.948	3.0819	8.8
29.46	3.0295	2.1
30.074	2.969	3
31.374	2.8488	2.9
31.783	2.8132	2.6
32.119	2.7845	1.7
32.537	2.7496	2
33.386	2.6816	1.8
35.278	2.542	2.7
35.715	2.5119	4.7
38.158	2.3565	1.3
39.105	2.3016	1.1
39.661	2.2706	2.8
40.605	2.22	2.2
42.638	2.1187	1.2

43.787	2.0657	0.9
44.294	2.0433	1.5
44.874	2.0182	1.2
45.419	1.9952	1.5
46.9	1.9356	2.9
48.306	1.8825	1
48.7	1.8682	0.8

Example 7

Example 1 was repeated, except that tetrabutyl titanate was added into the system as a titanium source to replace a part of the germanium source,
5 and the replacement ratio was 2%.

The XRD spectrum data for the product (in the prepared form) was shown in Table 7, and the XRD spectrum was similar to that of FIG.1.

The product molecular sieve was measured by inductively coupled plasma atomic emission spectroscopy (ICP) to have a $\text{SiO}_2/\text{GeO}_2=5.2$ and
10 $\text{SiO}_2/\text{TiO}_2=197.2$.

TABLE 7

$2\theta/^\circ$	$d/\text{\AA}$	I/I_0 100
4.831	18.27649	4
6.5962	13.38897	63
6.8914	12.81609	32.05
8.2366	10.72574	12.95
9.6938	9.11644	76.65
9.895	8.931518	100
10.2992	8.581884	1.15
11.8934	7.435085	1.15
13.1506	6.726979	4.4
13.5038	6.551815	2.2
14.219	6.223833	2.25

15.6172	5.669621	9.15
16.0884	5.504613	1.3
16.5456	5.353521	6.1
16.9358	5.231045	9.75
17.826	4.971771	2.5
19.2042	4.617964	5.45
19.4654	4.556583	10.6
19.9166	4.454362	18.25
20.7848	4.270113	9.85
21.121	4.202895	17.85
21.5952	4.111665	10.2
22.0484	4.028164	27.2
22.4836	3.951168	22.25
23.068	3.852375	3.7
23.4892	3.78424	9.25
23.9594	3.711028	6.55
24.5136	3.62837	4
24.9828	3.561279	4.55
25.282	3.519807	6.15
25.6362	3.471973	5.3
26.245	3.392799	8.25
26.5822	3.350521	10.15
27.9054	3.194581	5.15
28.4376	3.135994	3.4
28.6948	3.108467	5.05
28.951	3.081539	5.9
29.3882	3.036683	3.5
30.1134	2.965184	4.25
30.9796	2.884222	1.25
31.3778	2.848522	1.95
31.731	2.817616	2.3

32.0112	2.793589	2.1
32.4234	2.759009	2.1
33.3326	2.685801	2.2
34.1988	2.61973	1.3
34.83	2.573686	1.6
35.2462	2.544242	3.15
35.8744	2.501117	2.9
38.0476	2.363106	1.2
39.0508	2.304677	1.2
39.528	2.277946	2.8
40.5332	2.223744	2.15
42.4654	2.126925	1.05
43.683	2.070421	0.95
44.3132	2.042428	0.8
44.7354	2.024128	1.2
45.3056	1.999969	1.7
46.5658	1.948737	2
46.922	1.93477	1.25
49.2602	1.848265	0.7

Example 8

Example 1 was repeated except that the reaction mixture was prepared in the following ratios (molar ratios):

- 5 $\text{SiO}_2/\text{GeO}_2=4$
- Template agent/ $\text{SiO}_2=0.9$
- F/ $\text{SiO}_2=0.9$
- $\text{H}_2\text{O}/\text{SiO}_2=11$

10 After being mixed uniformly, the mixture was loaded into a stainless steel reactor for crystallizing at 170°C for 6 days under stirring.

 The XRD spectrum data of the product (in the prepared form) was shown in Table 8, and the XRD spectrum was shown in FIG.2.

The product molecular sieve was measured by inductively coupled plasma atomic emission spectroscopy (ICP) to have a SiO₂/GeO₂= 4.2.

TABLE 8

2 θ /°	d/Å	I/I ₀ 100
4.741	18.6215	1.6
6.555	13.4725	100
6.833	12.9252	25.3
8.234	10.7291	29.7
9.638	9.169	49.7
9.87	8.954	80.7
13.186	6.7089	4.9
13.501	6.5531	4.1
14.207	6.2287	1.6
15.572	5.6859	12
16.578	5.3431	6.4
16.894	5.2436	6.9
17.822	4.9729	3
18.826	4.7098	2.4
19.221	4.6138	4.2
19.494	4.5497	6.4
19.894	4.4592	21.2
20.72	4.2833	10.9
21.45	4.1391	22.9
22.023	4.0328	19.3
22.44	3.9587	14.4
23.089	3.849	4.3
23.524	3.7788	12.9
23.914	3.718	3.3
24.454	3.6371	2.2
24.894	3.5737	1.6

25.242	3.5253	5
25.618	3.4744	6.9
26.208	3.3975	2.6
26.679	3.3386	9.8
27.96	3.1885	3
28.592	3.1194	6.8
28.907	3.0861	6.4
29.48	3.0274	2
30.346	2.9429	3
31.331	2.8526	1.7
31.734	2.8174	1.7
32.004	2.7942	1.3
32.46	2.756	2
33.404	2.6803	2.1
34.876	2.5704	0.9
35.319	2.5392	2.5
35.635	2.5174	3.6
35.968	2.4948	2.3

The product was calcined at 550°C for 6 hours in the air atmosphere, and the XRD spectrum of the obtained sample (in the calcined form) was shown in FIG.3, and the spectrum data was shown in Table 9. The calcined sample had a specific surface area (BET method) of 337.4 m²/g and a micropore volume (t-plot method) of 0.14 cm³/g, as measured by nitrogen desorption.

TABLE 9

2 θ /°	d/Å	I/I ₀ 100
4.816	18.3323	2.2
6.604	13.3741	100
6.965	12.6807	34.7

8.39	10.5297	30.9
9.682	9.1272	29.7
9.976	8.859	50.8
13.249	6.6769	6.6
14.22	6.2232	7.5
15.124	5.8531	6.1
15.554	5.6923	18.5
17.26	5.1334	2.3
17.908	4.9492	3.8
19.487	4.5515	3.7
19.966	4.4433	9.2
20.203	4.3917	8
20.574	4.3135	6.5
21.174	4.1924	7.1
21.735	4.0855	21.5
22.24	3.9939	11.8
22.408	3.9643	11.4
23.748	3.7435	6.9
25.221	3.5282	6.4
25.681	3.466	11.5
26.224	3.3955	3.8
26.683	3.3381	4.7
27.12	3.2852	3.2
28.786	3.0988	7.3
30.163	2.9604	2.9
31.189	2.8654	2.6
32.237	2.7746	3.5
35.543	2.5237	4.9
36.185	2.4804	2
38.257	2.3506	1.8
46.871	1.9367	1.7

Example 9

The calcined molecular sieve obtained in example 8 and 0.7 wt% $\text{Al}(\text{NO}_3)_3$ solution were loaded into a three-neck flask at a weight ratio of molecular sieve: $\text{Al}(\text{NO}_3)_3$ solution = 1:50. Reaction was conducted for 6 hours in an oil bath at 80 °C under stirring. The solid sample was centrifuged and washed after reaction, and was put into an oven at 100 °C for overnight drying. The dried sample was then reacted with an 0.01 mol/L of HCl solution at a weight ratio of molecular sieve: HCl solution = 1:50 at room temperature for 6 hours under stirring. The solid sample was centrifuged and washed after the reaction, and was dried in an oven at 100 °C overnight to obtain a powder. The product was measured by inductively coupled plasma atomic emission spectroscopy (ICP) to have a $\text{SiO}_2/\text{GeO}_2=5.8$ and a $\text{SiO}_2/\text{Al}_2\text{O}_3=112.5$.

15

Example 10

3 g of the powder sample prepared in example 9 was mixed with 2 g of alumina and 0.3 g of sesbania powder, kneaded with 5 wt% of nitric acid, extruded into a rod of ϕ 1.6*2 mm, and then dried at 110 °C and calcined at 550 °C for 6 hours in an air atmosphere to remove organic substances, so as to prepare a desired molecular sieve composition. The molecular sieve composition could be used as an adsorbent or a catalyst.

Example 11

The molecular sieve composition prepared in the example 10 was crushed and sieved. 30 mg of particles having a particle size of 20-40 meshes was loaded into a pulse fixed bed reactor, activated for 1h in a nitrogen atmosphere at 300 °C, and cooled to the reaction temperature of 250 °C. A pulse sample injection mode was adopted to inject 0.4 microliter of cumene into the reactor instantly at one time. The cumene was subjected to a cracking reaction through a molecular sieve composition bed layer. The mixture after the reaction was directly fed into a gas chromatography

30

for analysis. The conversion rate of the cumene was 33.7%, and main products were propylene and benzene.

CLAIMS

1. A molecular sieve SCM-15, characterized in that the molecular sieve, particularly in its prepared form or calcined form, has an X-ray diffraction spectrum substantially as shown in Table A-1 or Table A-2 below,

TABLE A-1

2θ ($^{\circ}$) ^(a)	d-distance ($\text{\AA}$) ^(b)	Relative intensity ($I/I_0 \times 100$)
4.811	18.353	w
6.571	13.441	m-vs
6.888	12.823	m-s
9.675	9.134	m-vs
9.894	8.933	m-vs
19.888	4.461	w-m

(a)= $\pm 0.3^{\circ}$, (b) is a function of 2θ ,

TABLE A-2

2θ ($^{\circ}$) ^(a)	d-distance ($\text{\AA}$)	Relative intensity ($I/I_0 \times 100$)
4.811	18.424 ± 1.148	w
6.571	13.468 ± 0.614	m-vs
6.888	12.847 ± 0.559	m-s
9.675	9.143 ± 0.283	m-vs
9.894	8.941 ± 0.270	m-vs
19.888	4.462 ± 0.067	w-m

10 (a) = $\pm 0.3^{\circ}$.

2. The molecular sieve SCM-15 according to claim 1, characterized in that the X-ray diffraction spectrum further comprises X-ray diffraction peaks substantially as shown in Table B-1 or Table B-2 below,

TABLE B-1

2θ ($^{\circ}$) ^(a)	d-distance ($\text{\AA}$) ^(b)	Relative intensity ($I/I_0 \times 100$)
8.225	10.741	w
15.592	5.679	w

22.031	4.031	w-m
22.435	3.960	w-m
25.330	3.513	w

(a)= $\pm 0.3^\circ$, (b) is a function of 2θ ,

TABLE B-2

2θ ($^\circ$) ^(a)	d-distance ($\text{\AA}$)	Relative intensity ($I/I_0 \times 100$)
8.225	10.755 ± 0.392	w
15.592	5.681 ± 0.109	w
22.031	4.032 ± 0.054	w-m
22.435	3.960 ± 0.052	w-m
25.330	3.514 ± 0.041	w

(a)= $\pm 0.3^\circ$,

said X-ray diffraction spectrum optionally further comprising X-ray
5 diffraction peaks substantially as shown in the table below,

2θ ($^\circ$) ^(a)	d-distance ($\text{\AA}$)	Relative intensity ($I/I_0 \times 100$)
14.294	6.194 ± 0.129	w
16.909	5.241 ± 0.092	w
19.156	4.630 ± 0.072	w-m
19.517	4.546 ± 0.069	w-m
21.162	4.196 ± 0.059	w-m

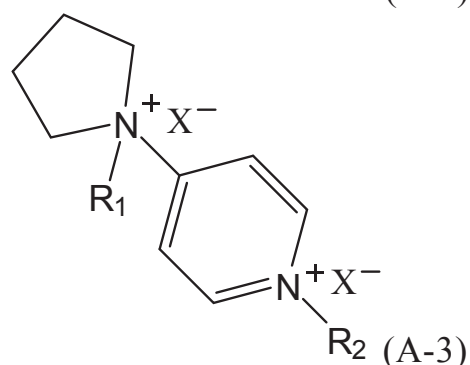
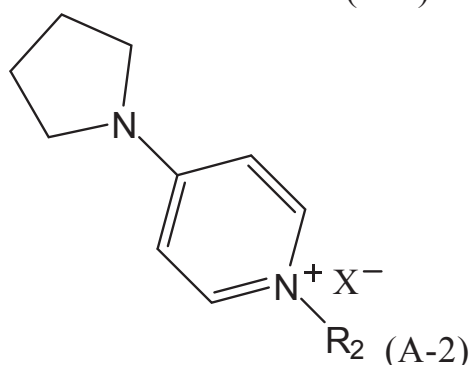
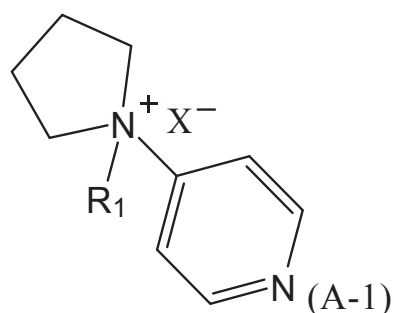
(a) = $\pm 0.3^\circ$.

3. The molecular sieve SCM-15 according to claim 1, characterized in that the molecular sieve has a schematic chemical composition of formula "SiO₂·GeO₂", wherein the molar ratio of silicon to germanium satisfies
10 SiO₂/GeO₂ ≥ 1 , preferably $1 \leq \text{SiO}_2/\text{GeO}_2 \leq 15$, more preferably $2 \leq \text{SiO}_2/\text{GeO}_2 \leq 10$, or more preferably $2.5 \leq \text{SiO}_2/\text{GeO}_2 \leq 5$.

4. The molecular sieve SCM-15 according to claim 3, characterized in that not more than 10% of the Ge atoms in the molecular sieve are substituted by atoms of at least one element other than silicon and
15 germanium.

5. The molecular sieve SCM-15 according to claim 4, characterized in that the element other than silicon and germanium is at least one selected from the group consisting of boron, aluminum, tin, zirconium and titanium.

6. A process of preparing molecular sieve SCM-15, comprising a step
5 of crystallizing a mixture comprising a silicon source, a germanium source, a fluorine source, an organic template and water or consisting of a silicon source, a germanium source, a fluorine source, an organic template and water, to obtain the molecular sieve, wherein the organic template is selected from 4-pyrrolidinyropyridine, or a quaternary ammonium form
10 represented by formula (A-1), formula (A-2) or formula (A-3), preferably 4-pyrrolidinyropyridine,



15 in each formula, R₁ and R₂ are each independently H or C₁₋₈ alkyl, preferably C₁₋₄ alkyl, more preferably C₁₋₂ alkyl, and X⁻ are each

independently a halogen ion (such as Cl^- , Br^- , and I^-) and a hydroxide ion (OH^-), preferably hydroxide ion (OH^-).

7. The process of preparing molecular sieve SCM-15 according to claim 6, characterized in that the silicon source is at least one selected from the group consisting of silicic acid, silica gel, silica sol, tetraalkyl orthosilicate and water glass; the germanium source is at least one selected from the group consisting of germanium oxide, germanium nitrate and tetraalkoxygermanium;

and the molar ratio of the silicon source (calculated by SiO_2), the germanium source (calculated by GeO_2), the fluorine source (calculated by F), the organic template agent and water is 1:(0-1):(0.1-2.0):(0.1-2.0):(3-30); preferably 1:(1/15-1.5):(0.2-1.5):(0.2-1.5):(4-25); more preferably 1:(0.1-0.5):(0.4-1.2):(0.4-1.2):(5-20); more preferably 1:(0.2-0.4):(0.6-1.0):(0.6-1.0):(5-15).

8. The process of preparing molecular sieve SCM-15 according to claim 6, characterized in that the crystallization conditions comprise: a crystallization temperature of 131 to 210°C, preferably 150 to 190°C, more preferably 160 to 180°C; and a crystallization duration of 1 to 20 days, preferably 2 to 10 days, more preferably 2 to 7 days.

9. The process of preparing molecular sieve SCM-15 according to claim 6, characterized in that the process comprises an aging step prior to crystallization; and the aging conditions include: an ageing temperature of 50-90°C, and an ageing duration of 2 hours to 2 days.

10. The process of preparing molecular sieve SCM-15 according to claim 6, characterized in that the mixture further comprises a source of an element other than silicon and germanium, preferably at least one selected from the group consisting of boron source, aluminum source, tin source, zirconium source and titanium source; more preferably at least one oxide source selected from the group consisting of a boron oxide source, an alumina source, a tin oxide source, a zirconium oxide source and a titanium oxide source;

and the molar ratio of the oxide source (calculated as the corresponding oxide) to the germanium source (calculated as GeO_2) is (0.01-0.1):1, preferably (0.02-0.08): 1.

11. A molecular sieve composition, comprising the molecular sieve
5 SCM-15 according to any one of claims 1 to 5 or the molecular sieve
SCM-15 prepared according to the process of preparing the molecular
sieve SCM-15 according to any of claims 6 to 10, and a binder.

12. Use of the molecular sieve SCM-15 according to any one of
claims 1 to 5, the molecular sieve SCM-15 prepared according to the
10 process of preparing the molecular sieve SCM-15 according to any one of
claims 6 to 10, or the molecular sieve composition according to claim 11,
as an adsorbent or catalyst.