

## **WASTE INCINERATOR ASH AS AERATING AGENT FOR THE MANUFACTURE OF LIGHTWEIGHT CONSTRUCTION MATERIALS**

### **CROSS-REFERENCE TO RELATED APPLICATION**

5 [0001] This application claims the benefit of priority of US provisional application No. 61/863,308 filed on 7 August 2013 and US provisional application No. 61/879,945 filed on 19 September 2013, the content of which are incorporated herein by reference in their entirety for all purposes.

### **TECHNICAL FIELD**

10 [0002] The invention relates to the manufacture of lightweight construction materials using waste incinerator ash. The ash may be derived from the bottom ash of a municipal solid waste incinerator facility.

### **BACKGROUND**

15 [0003] Incineration greatly reduces volume of municipal solid waste (MSW) and is the dominant waste treatment process in major cities and land scarce countries such as Singapore. Generally, two types of ashes are generated from MSW incineration - about 80 % of the ash is incineration bottom ash (IBA), while the remaining 20 % is incineration fly ash (IFA). Of  
20 the two, IBA contains much lower amounts of leachable heavy metals and highly toxic organic substances such as dioxins. In addition, chloride content in IBA is much lower.

[0004] Studies have been carried out on potential use of IBA in civil engineering applications, such as road construction, embankment, pavement, aggregate, and as fillers for concrete.

25 [0005] Although IBA may potentially be used as replacement material in base course and sub-base for road construction, concerns regarding leaching of heavy metals into soil and underground water, which can lead to serious health and environmental problems, exist. Mechanisms of long-term leaching remain unclear, and remain matters of research. Further, presence of IBA as aggregate in concrete may lead to expansion and cracking of the concrete  
30 due to reaction between metallic aluminum in IBA and cement.

[0006] Unlike ferrous metal which may be easily extracted, recovery efficiency of non-ferrous metals such as aluminum is relatively low. Traditional eddy current separation

method showed an average recovery efficiency of 30 % of the aluminum fed into the furnace of the incineration plant. It was also reported the recovery efficiency is nearly zero for fine IBA particles (size range of 5 mm to 12 mm).

[0007] Even though pre-treatment methods, such as vitrifying, may address some of the issues concerning IBA use in civil engineering applications, high costs and process complexities associated with the methods render them unattractive for practical and industrial application.

[0008] For example, conventional approach of recycling and reutilization of IBA as aggregates requires pre-treatment to improve the safety and usability of IBA. The most economic phosphate-based IBA treatment, such as WES-PHix, costs S\$30 per ton which is about two times the cost of natural aggregates of S\$15/ton. Vitrification involves higher costs of S\$200/ton. Therefore, unless pre-treatment costs may be reduced to less than S\$15/ton, it is not economically attractive to use IBA as a replacement for aggregates.

[0009] In view of the above, there remains a need for an improved method to utilize waste incinerator ash that overcomes or at least alleviates one or more of the above-mentioned problems.

## SUMMARY

[0010] In a first aspect, a method of manufacturing a lightweight concrete is provided. The method comprises

- a) providing ash obtained from incineration of waste, the ash comprising at least about 0.1 wt% metallic aluminum;
- b) mixing the ash with a binder to form a mixture; and
- c) curing the mixture to obtain the lightweight concrete.

[0011] In a second aspect, lightweight concrete manufactured by a method according to the first aspect is provided.

[0012] In a third aspect, lightweight concrete having a density of about 200 kg/m<sup>3</sup> to about 1900 kg/m<sup>3</sup>, a compressive strength of about 0.5 MPa to about 70 MPa, and a thermal conductivity in the range of about 0.05 W/m-k to about 1.3 W/m-k is provided.

[0013] In a fourth aspect, use of ash obtained from incineration of waste, the ash comprising at least about 0.1 wt% metallic aluminum, as an aerating agent is provided.

### BRIEF DESCRIPTION OF THE DRAWINGS

[0014] The invention will be better understood with reference to the detailed description when considered in conjunction with the non-limiting examples and the accompanying drawings, in which:

5 [0015] **FIG. 1A** is an optical microscopy image of pore structure in lightweight concrete using inorganic polymer as binder (pores in black) according to an embodiment. **FIG. 1B** is a binary image of the optical microscopy image in **FIG. 1A**.

[0016] **FIG. 2** is a graph showing effects of mix duration (minutes) and liquid to solid ratio (L/S) on the dry density ( $\text{kg/m}^3$ ) of lightweight concrete using inorganic polymer as binder according to embodiments.

[0017] **FIG. 3** is a graph showing relationship between compressive strength and dry density of lightweight concrete using inorganic polymer as binder according to embodiments.

[0018] **FIG. 4** is a graph showing relationship between dry density ( $\text{kg/m}^3$ ) and macro-porosity (%) of lightweight concrete using inorganic polymer as binder according to

15 [0019] **FIG. 5A** is an optical microscopy image of pore structure in lightweight concrete using hydraulic cement as binder (pores in black) according to an embodiment. **FIG. 5B** is a binary image of the optical microscopy image in **FIG. 5A**.

[0020] **FIG. 6** is a graph showing compressive strength (MPa) versus density ( $\text{kg/m}^3$ ) of lightweight concrete using hydraulic cement as binder according to embodiments.

[0021] **FIG. 7A** is a graph showing effect of alkali molarity (mol/L) on the compressive strength (MPa) and dry density ( $\text{kg/m}^3$ ) of lightweight concrete using hydraulic cement as binder according to embodiments. **FIG. 7B** is a graph showing effect of IBA replacement ratio (%) on the compressive strength (MPa) and dry density ( $\text{kg/m}^3$ ) of lightweight concrete

25 [0022] **FIG. 8** is a graph showing effect of macro-porosity (%) on dry density ( $\text{kg/m}^3$ ) and compressive strength (MPa) of lightweight concrete using hydraulic cement as binder according to embodiments.

[0023] **FIG. 9** shows typical scanning electron microscopy (SEM) images of IBA lightweight concrete using hydraulic cement as binder prepared with (A) tap water, matrix; (B) tap water, wall of macro-pore; (C) sodium hydroxide (NaOH) solution, matrix; and (D) NaOH solution, wall of macro-pore. Scale bar in the figures denote a length of 1  $\mu\text{m}$ .

### DETAILED DESCRIPTION

[0024] Instead of pre-treating waste incinerator ash to remove or to immobilize/stabilize metallic aluminum, embodiments disclosed herein utilize waste incinerator ash containing metallic aluminum as aerating agent for the manufacture of lightweight construction materials, such as aerated concrete.

[0025] By using ash obtained from incineration of waste, addition of metallic aluminum powder is not required for producing the lightweight construction materials. In addition, waste incinerator ash may possess pozzolanic and/or cementitious properties, which render its suitability as a replacement for cement. With rich metallic aluminum and silicate content, waste incinerator ash may be used as a precursor for manufacturing alkaline-activated aluminosilicate, or geopolymer, binders.

[0026] Furthermore, by utilizing waste incinerator ash in production of products that have higher value, such as aerated concrete, this may render use of waste incinerator ash economically attractive and viable, as compared to low value applications such as aggregates. By replacing aluminum powder, cement, and other binder materials using waste incinerator ash, it is estimated that raw material costs may be reduced by as much as 30 %.

[0027] With the above in mind, the invention refers in a first aspect to a method of manufacturing a lightweight concrete. As used herein, the term "lightweight concrete" refers to concrete having a density lower than that of normal concrete, where typical density of normal concrete is about  $2,400 \text{ kg/m}^3$ . The term "aerated concrete" as used herein refers to a type of lightweight concrete that has air voids entrapped therein to render its lower density than that of a normal concrete.

[0028] The lightweight concrete disclosed herein may have a density of about  $200 \text{ kg/m}^3$  to about  $1900 \text{ kg/m}^3$ , a compressive strength of about 0.5 MPa to about 70 MPa, and a thermal conductivity of about 0.05 W/m-k to about 1.3 W/m-k, rendering its suitability for thermal insulation and sound-proofing. Applications of such lightweight concrete include, but are not limited to, structural use such as in floors and trench fills, thermal insulation such as roof insulation and other insulating purposes, as well as to make masonry units.

[0029] Various embodiments disclosed herein refer to use of waste incinerator ash as aerating agent to entrain air and to reduce density of construction materials. The method includes providing ash obtained from incineration of waste. Generally, ash obtained from

incineration of waste, or waste incinerator ash, may be classified as bottom ash and fly ash. The term “bottom ash”, otherwise termed herein as “incinerator bottom ash”, IBA, refers to ash that is collected at the bottom of a combustion chamber or incinerator. In some municipal solid waste combustors, the bottom ash is recovered from the furnace grate of the combustor chamber. In other combustors, such as fluidized-bed type combustors, the bottom ash is recovered from the bed, or other solid ash residue access point.

[0030] Fly ash, on the other hand, refers to ash that is lighter than bottom ash and which is collected mainly in the discharge gas processing system connected to the incinerator or in a dust collector provided in the discharge gas processing system. Bottom ash and fly ash have significantly different properties and compositions, with fly ash having a smaller specific weight and containing more volatile components than bottom ash.

[0031] In various embodiments, the ash is derived from bottom ash of a solid waste incinerator facility. Typically, bottom ash contains a heterogeneous mixture of slag, glass, ceramics, ferrous and non-ferrous metals, minerals, non-combustibles; and unburnt organic matter. **TABLE 1** provides an exemplary chemical composition of ash disclosed herein.

[0032] **TABLE 1:** Chemical compositions of incinerator bottom ash

| Oxide                          | Content (%) | Oxide                          | Content (%) |
|--------------------------------|-------------|--------------------------------|-------------|
| SiO <sub>2</sub>               | 32.75       | K <sub>2</sub> O               | 1.24        |
| CaO                            | 29.06       | ZnO                            | 0.81        |
| Fe <sub>2</sub> O <sub>3</sub> | 10.02       | CuO                            | 0.31        |
| Al <sub>2</sub> O <sub>3</sub> | 8.57        | Cr <sub>2</sub> O <sub>3</sub> | 0.22        |
| P <sub>2</sub> O <sub>5</sub>  | 4.77        | MnO                            | 0.15        |
| SO <sub>3</sub>                | 3.01        | PbO                            | 0.12        |
| Na <sub>2</sub> O              | 2.87        | SrO                            | 0.10        |
| MgO                            | 1.75        | NiO                            | 0.06        |
| TiO <sub>2</sub>               | 1.57        | ZrO <sub>2</sub>               | 0.01        |

[0033] The ash obtained from incineration of waste comprises at least about 0.1 wt% metallic aluminum. In various embodiments, the ash obtained from incineration of waste comprises at least about 0.5 wt% metallic aluminum, such as at least about 1 wt%, about 2 wt%, about 3 wt%, about 5 wt%, or at least about 10 wt% metallic aluminum.

[0034] As mentioned above, metallic aluminum present in the ash may act as an aerating agent. It may react in an alkaline environment, such as fresh cement paste disclosed herein, to

generate bubbles of hydrogen gas. With subsequent hardening and curing of the mixture, a lightweight concrete may be obtained.

[0035] Providing ash obtained from incineration of waste may include reducing particle size of the ash to an average particle size of about 500  $\mu\text{m}$  or less. For example, particle size of the ash may be reduced to an average particle size in the range of about 20  $\mu\text{m}$  to about 500  $\mu\text{m}$ , such as about 40  $\mu\text{m}$  to about 500  $\mu\text{m}$ , about 100  $\mu\text{m}$  to about 500  $\mu\text{m}$ , about 150  $\mu\text{m}$  to about 500  $\mu\text{m}$ , about 200  $\mu\text{m}$  to about 500  $\mu\text{m}$ , about 300  $\mu\text{m}$  to about 500  $\mu\text{m}$ , about 400  $\mu\text{m}$  to about 500  $\mu\text{m}$ , about 20  $\mu\text{m}$  to about 400  $\mu\text{m}$ , about 20  $\mu\text{m}$  to about 300  $\mu\text{m}$ , about 20  $\mu\text{m}$  to about 200  $\mu\text{m}$ , about 20  $\mu\text{m}$  to about 100  $\mu\text{m}$ , about 20  $\mu\text{m}$  to about 50  $\mu\text{m}$ , or about 30  $\mu\text{m}$  to about 50  $\mu\text{m}$ . In specific embodiments, providing ash obtained from incineration of waste may include reducing particle size of the ash to an average particle size of about 20  $\mu\text{m}$  to about 40  $\mu\text{m}$ .

[0036] Reducing particle size of the ash may be carried out by any suitable size reduction methods such as, but not limited to, grinding, crushing, cutting, milling, and combinations thereof. In specific embodiments, reducing particle size of the ash is carried out by ball milling.

[0037] In various embodiments, providing ash obtained from incineration of waste comprises drying the ash. Drying the ash may be carried out at ambient temperature, defined herein as a temperature of between about 20  $^{\circ}\text{C}$  to about 40  $^{\circ}\text{C}$ , or at elevated temperatures.

[0038] The method includes mixing the waste incinerator ash with a binder to form a mixture. As used herein, the term “binder” refers to a material that is capable of attaching two or more materials to one another such that the two or more materials are held together.

[0039] Ratio of ash to binder may be in the range of about 0.01 to about 100 by weight. For example, ratio of ash to binder may be in the range of about 1 to about 100, about 10 to about 100, about 25 to about 100, about 50 to about 100, about 75 to about 100, about 0.01 to about 75, about 0.01 to about 50, about 0.01 to about 25, about 1 to about 80, about 10 to about 80, about 25 to about 75, or about 30 to about 60 by weight.

[0040] In various embodiments, the binder is selected from the group consisting of a hydraulic cement binder and an inorganic polymer binder.

[0041] In some embodiments, the binder comprises or consists of a hydraulic cement binder. In specific embodiments, the binder consists of a hydraulic cement binder. As used herein, the term “hydraulic cement” refers to cement that sets and hardens in the presence of

water. In various embodiments, the hydraulic cement binder is selected from the group consisting of Portland cement, blended Portland cement, expansive cement, rapid setting and hardening cement, calcium aluminate cement, magnesium phosphate, and mixtures thereof. In specific embodiments, the hydraulic cement binder comprises or consists of Type I  
5 Portland cement.

[0042] Mixing the ash with a hydraulic cement binder may include adding at least one of a pozzolanic admixture, water, lime, alkali, and a set-control admixture to the mixture. For example, pozzolanic admixtures such as fly ash, silica fume, and/or ground granulated blast-furnace slag may be included in the mixture.

10 [0043] Ratio of water to the hydraulic cement binder in the mixture may be in the range of about 0.2 to about 1 by weight. For example, the ratio of water to the hydraulic cement binder may be in the range of about 0.3 to about 1 by weight, about 0.3 to about 0.9 by weight, or about 0.2 to about 0.8 by weight. In specific embodiments, the ratio of water to the hydraulic cement binder is about 0.35 to about 0.7 by weight.

15 [0044] Lime and/or alkalis may be added to accelerate hydrogen generation from the mixture containing ash and a hydraulic cement binder. In various embodiments, the lime comprises or consists of an alkaline earth metal carbonate. In specific embodiments, the lime comprises or consists of at least one of magnesium carbonate and calcium carbonate.

[0045] The alkali may be selected from the group consisting of an alkali metal hydroxide,  
20 an alkaline earth metal hydroxide, ammonium hydroxide, sodium carbonate, and combinations thereof. Examples of an alkali metal hydroxide include lithium hydroxide, sodium hydroxide, potassium hydroxide, rubidium hydroxide, cesium hydroxide, francium hydroxide, or combinations thereof. Examples of an alkaline earth metal hydroxide include beryllium hydroxide, beryllium hydroxide, magnesium hydroxide, calcium hydroxide,  
25 strontium hydroxide, barium hydroxide, or combinations thereof. In some embodiments, the alkali is selected from the group consisting of sodium hydroxide, potassium hydroxide, calcium hydroxide, magnesium hydroxide, sodium carbonate, and combinations thereof. In specific embodiments, the alkali comprises or consists of sodium hydroxide.

[0046] Concentration of hydroxide in the mixture may be 100 mol/L or less. For example,  
30 the hydroxide may be present in the mixture in the range of about 10 mol/L to about 100 mol/L, such as about 30 mol/L to about 100 mol/L, about 50 mol/L to about 100 mol/L, about 70 mol/L to about 100 mol/L, about 10 mol/L to about 70 mol/L, about 10 mol/L to about 50

mol/L, about 10 mol/L to about 30 mol/L, about 30 mol/L to about 80 mol/L, or about 20 mol/L to about 90 mol/L.

[0047] A set-control admixture may be added to control setting for the hydraulic cement binder. For example, the set-control admixture may be selected from the group consisting of gypsum, anhydrite, and combinations thereof. Gypsum and/or anhydrite may serve to delay hydration and setting so as to provide sufficient time for aeration in the mixture.

[0048] In some embodiments, the binder consists essentially or consists of an inorganic polymer binder. In specific embodiments, the binder consists of an inorganic polymer binder. As used herein, the term "inorganic polymer" refers to a compound having a skeletal structure that does not include carbon atoms. In various embodiments, the inorganic polymer binder is selected from the group consisting of alkali activated aluminosilicates, geopolymer, alkali activated cement, and mixtures thereof.

[0049] The inorganic polymer binder may be obtained by reacting an inorganic polymer precursor with an alkaline activator liquid. The precursor to form inorganic polymer binder may be selected from the group consisting of coal fly ash, metakaoline, waste incinerator ash, ground granulated blast-furnace slag, and combinations thereof.

[0050] The alkaline activator liquid for the inorganic polymer binder may comprise a silicate and/or a hydroxide selected from the group consisting an alkali-metal hydroxide, an alkaline earth metal hydroxide, ammonium hydroxide, and combinations thereof.

[0051] In various embodiments, the silicate comprises or consists of an alkali metal silicate. For example, the alkali metal silicate may be at least one of lithium silicate, sodium silicate, potassium silicate, rubidium silicate, cesium silicate, and francium silicate. In specific embodiments, the silicate comprises or consists of sodium silicate.

[0052] The hydroxide may be selected from the group consisting of an alkali metal hydroxide, an alkaline earth metal hydroxide, ammonium hydroxide, and combinations thereof. For example, the hydroxide may comprise or consist of an alkali metal hydroxide, such as lithium hydroxide, sodium hydroxide, potassium hydroxide, rubidium hydroxide, cesium hydroxide, francium hydroxide, or combinations thereof. The hydroxide may additionally or otherwise contain an alkaline earth metal hydroxide, such as beryllium hydroxide, beryllium hydroxide, magnesium hydroxide, calcium hydroxide, strontium hydroxide, barium hydroxide, or combinations thereof. In various embodiments, the hydroxide comprises or consists of sodium hydroxide.

[0053] In various embodiments, concentration of hydroxide in the alkaline activator liquid is at least 1 mol/L. For example, concentration of hydroxide in the alkaline activator liquid may be at least 2 mol/L, at least 4 mol/L, at least 6 mol/L, at least 8 mol/L, or at least 10 mol/L. In specific embodiments, concentration of hydroxide in the alkaline activator liquid is about 8 mol/L.

[0054] When hydroxide and silicate are both present in the alkaline activator liquid, ratio of hydroxide to silicate in the alkaline activator liquid may be in the range of about 0.2 to about 3 by weight. For example, ratio of hydroxide to silicate by weight in the alkaline activator liquid may be in the range of about 0.5 to about 3, about 1 to about 3 by weight, about 1.5 to about 3, about 2 to about 3, about 2.5 to about 3, about 0.2 to about 2.5, about 0.2 to about 2, about 0.2 to about 1.5, about 0.2 to about 1, about 0.5 to about 2.5, about 1 to about 2, or about 1.5 to about 2.5.

[0055] The silicate and/or the hydroxide may be at least substantially dissolved in an aqueous reagent. In various embodiments, the alkaline activator liquid is an aqueous solution of the silicate and/or the hydroxide.

[0056] Ratio of alkaline activator liquid to inorganic polymer binder may be in the range of about 0.3 to about 2 by weight. In various embodiments, ratio of the alkaline activator liquid to the inorganic polymer binder is in the range of about 0.5 to about 2, about 0.8 to about 2, about 1.5 to about 2, about 0.3 to about 1.5, about 0.3 to about 1, about 0.5 to about 1.5, about 0.5 to about 1, or about 0.8 to about 1.5.

[0057] In addition to the binder, a viscosity control agent and/or water reducing agent may be added to the mixture. The viscosity control agent and/or water reducing agent may be added to the mixture containing the hydraulic cement binder or the inorganic polymer binder to affect rheology of the mixture.

[0058] In various embodiments, the viscosity control agent is selected from the group consisting of methylcellulose, polyvinyl alcohol, starch, and combinations thereof.

[0059] Ratio of viscosity control agent to binder may not be more than 0.05 by weight. For example, ratio of viscosity control agent to binder may be in the range of about 0.005 to about 0.05, about 0.01 to about 0.05, about 0.02 to about 0.05, about 0.03 to about 0.05, about 0.005 to about 0.04, about 0.005 to about 0.03, or about 0.005 to about 0.02.

[0060] In various embodiments, the water reducing agent is selected from the group consisting of carboxylated polymer, gypsum, anhydrite, and mixtures thereof. Amount of

water reducing agent used may depend on water content in the mixture as measured by the water-to-binder ratio, for example, type of lightweight filler, and type of water reducing agent used. In various embodiments, ratio of the water reducing agent to the binder is in the range of about 0.001 to about 0.002 by weight.

5 [0061] Apart from adding the viscosity control agent and/or the water reducing agent, mixing the ash with the binder to form a mixture may also include adding aggregates to the mixture. Amount and type of aggregates used may be varied depending on specific physical, mechanical and thermal properties of the lightweight concrete.

10 [0062] The aggregates may include coarse aggregates selected from the group consisting of gravel, crushed rock, slag aggregate, recycled aggregate, crushed quartz, granite, gneiss, basalt, limestone, and combinations thereof. In some embodiments, the aggregates comprise fine aggregates selected from the group consisting of sand, fly ash, silica fume, artificial light weight fine aggregate, and combinations thereof.

15 [0063] The mixing may be carried out in any suitable or conventional type of concrete or mortar mixer.

[0064] The method disclosed herein includes curing the mixture comprising the ash and the binder to obtain the lightweight concrete. Curing the mixture may be carried out at room temperature, at an elevated temperature such as steam curing, or at an elevated temperature and pressure such as in an autoclave.

20 [0065] In various embodiments, curing the mixture to obtain the lightweight concrete is carried out at a temperature in the range of about 20 °C to about 80 °C, such as about 25 °C to about 60 °C, about 25 °C to about 50 °C, about 25 °C to about 40 °C, about 30 °C to about 80 °C, about 45 °C to about 80 °C, about 60 °C to about 80 °C, or about 30 °C to about 60 °C. In specific embodiments, curing the mixture is carried out at a temperature in the range of  
25 about 75 °C to about 80 °C.

[0066] The invention also refers in further aspects, to a lightweight concrete manufactured by a method according to the first aspect, and to a lightweight concrete having a density in the range of about 200 kg/m<sup>3</sup> to about 1900 kg/m<sup>3</sup>, a compressive strength in the range of about 0.5 MPa to about 70 MPa, and a thermal conductivity in the range of about 0.05 W/m-k  
30 to about 1.3 W/m-k.

[0067] The lightweight concrete according to various embodiments may have a density of about 200 kg/m<sup>3</sup> to about 1900 kg/m<sup>3</sup>, such as about 500 kg/m<sup>3</sup> to about 1900 kg/m<sup>3</sup>, about

800 kg/m<sup>3</sup> to about 1900 kg/m<sup>3</sup>, about 1000 kg/m<sup>3</sup> to about 1900 kg/m<sup>3</sup>, about 200 kg/m<sup>3</sup> to about 1000 kg/m<sup>3</sup>, about 200 kg/m<sup>3</sup> to about 800 kg/m<sup>3</sup>, about 200 kg/m<sup>3</sup> to about 500 kg/m<sup>3</sup>, about 500 kg/m<sup>3</sup> to about 800 kg/m<sup>3</sup>, or about 400 kg/m<sup>3</sup> to about 600 kg/m<sup>3</sup>.

[0068] The lightweight concrete may have a compressive strength of about 0.5 MPa to about 70 MPa, such as about 1 MPa to about 70 MPa, about 15 MPa to about 70 MPa, about 30 MPa to about 70 MPa, about 50 MPa to about 70 MPa, about 0.5 MPa to about 50 MPa, about 0.5 MPa to about 30 MPa, or about 10 MPa to about 50 MPa.

[0069] Thermal conductivity of the lightweight concrete may be in the range of about 0.05 W/m-k to about 1.3 W/m-k, such as about 0.1 W/m-k to about 1.3 W/m-k, about 0.5 W/m-k to about 0.3 W/m-k, about 0.7 W/m-k to about 1.3 W/m-k, about 0.05 W/m-k to about 1 W/m-k, about 0.05 W/m-k to about 0.8 W/m-k, about 0.3 W/m-k to about 1.2 W/m-k, or about 0.5 W/m-k to about 1 W/m-k.

[0070] By utilizing ash obtained from incineration of waste which would otherwise occupy space in the landfills, lifespan of landfills may be prolonged. Other benefits include reducing or eliminating disposal costs of ash, preserving land capacity, conserving dwindling supplies of natural raw materials, and mitigating potential environmental impacts.

[0071] The invention refers in a fourth aspect to use of ash obtained from incineration of waste, the ash comprising at least about 0.1 wt% metallic aluminum as an aerating agent. As mentioned above, the aerating agent may be an aerating agent in manufacture of lightweight construction materials such as lightweight concrete or aerated concrete.

[0072] As mentioned above, metallic aluminum and chemical species that are present in the ash obtained from incineration of waste may act as aerating agent without the need for separate aeration agents to be added. This may result in a simplified process as well as reduces costs of manufacturing the lightweight concrete.

[0073] Hereinafter, the present invention will be described more fully with reference to the accompanying drawings, in which exemplary embodiments of the invention are shown. This invention may, however, be embodied in many different forms and should not be construed as limited to the exemplary embodiments set forth herein. Rather, these embodiments are provided so that this disclosure will be thorough and complete, and will fully convey the scope of the invention to those skilled in the art. In the drawings, lengths and sizes of layers and regions may be exaggerated for clarity.

[0074] The terminology used herein is for the purpose of describing particular embodiments only and is not intended to be limiting of the invention. As used herein, the singular forms "a", "an" and "the" are intended to include the plural forms as well, unless the context clearly indicates otherwise. It will be further understood that the terms "comprises" and/or "comprising," when used in this specification, specify the presence of stated features, integers, steps, operations, elements, and/or components, but do not preclude the presence or addition of one or more other features, integers, steps, operations, elements, components, and/or groups thereof.

[0075] The invention illustratively described herein may suitably be practiced in the absence of any element or elements, limitation or limitations, not specifically disclosed herein. Thus, for example, the terms "comprising", "including", "containing", etc. shall be read expansively and without limitation. Additionally, the terms and expressions employed herein have been used as terms of description and not of limitation, and there is no intention in the use of such terms and expressions of excluding any equivalents of the features shown and described or portions thereof, but it is recognized that various modifications are possible within the scope of the invention claimed. Thus, it should be understood that although the present invention has been specifically disclosed by preferred embodiments and optional features, modification and variation of the inventions embodied therein herein disclosed may be resorted to by those skilled in the art, and that such modifications and variations are considered to be within the scope of this invention.

[0076] The invention has been described broadly and generically herein. Each of the narrower species and subgeneric groupings falling within the generic disclosure also form part of the invention. This includes the generic description of the invention with a proviso or negative limitation removing any subject matter from the genus, regardless of whether or not the excised material is specifically recited herein.

[0077] Other embodiments are within the following claims and non-limiting examples. In addition, where features or aspects of the invention are described in terms of Markush groups, those skilled in the art will recognize that the invention is also thereby described in terms of any individual member or subgroup of members of the Markush group.

## EXPERIMENTAL SECTION

[0078] Conventional approach of recycling and reutilization of incinerator bottom ash (IBA) as aggregates requires costly pre-treatment to improve the safety and usability of IBA. Current cost of pre-treatment of IBA as road aggregates is around S\$30 per ton, two times the cost of natural aggregates (S\$15/ton). Other pre-treatment such as vitrification is even more costly (S\$100/ton). It is therefore not attractive to use IBA to replace aggregates as it does not make economic sense, unless costs of the pre-treatment may be reduced to less than S\$15/ton. In embodiments disclosed herein, ash obtained from incineration of waste such as IBA is utilized to produce high value products such as aerated concrete, which are economically attractive and viable, as compared to low value products such as IBA aggregate.

[0079] Typical mix proportions of aerated concrete may be found in **TABLE 2**.

[0080] **TABLE 2:** Typical mix proportion and cost structure of lightweight aerated concrete

| Material        | Mix proportion (kg/m <sup>3</sup> ) | Unit price (S\$/ton) | Cost (S\$/m <sup>3</sup> ) | Cost ratio |
|-----------------|-------------------------------------|----------------------|----------------------------|------------|
| Cement          | 49                                  | 152                  | 7.4                        | 26%        |
| Sand            | 490                                 | 14                   | 6.9                        | 25%        |
| Lime            | 133                                 | 80                   | 10.6                       | 38%        |
| Gypsum          | 28                                  | 20                   | 0.6                        | 2%         |
| Aluminum powder | 0.46                                | 5,200                | 2.4                        | 9%         |
|                 |                                     | Total                | 27.9                       | 100%       |

[0081] Instead of pre-treating ash obtained from incineration of waste such as IBA to remove or to immobilize/stabilize metallic aluminum and heavy metals, embodiments disclosed herein utilizes the ash to replace costly aluminum powder and other binder material such as cement, sand, lime, and gypsum in the production of lightweight aerated concrete. With replacement of cement, lime, gypsum, aluminum powder by ash in production of aerated concrete, it is estimated the resulting lightweight aerated concrete has a cost advantage over the existing autoclaved aerated concrete by at least 20 %.

[0082] **Example 1: Material Preparation For Manufacture of Lightweight Concrete Using Inorganic Polymer as Binder (Embodiment 1)**

[0083] Exemplary mixes for preparing lightweight concrete containing IBA, with mix proportion of IBA to alkaline solution and their mix durations (minutes), are tabulated in **TABLE 3**. Unlike conventional aerated concrete shown in **TABLE 2**, costly aerating agent (e.g. aluminum powder) and cementitious binders (e.g. cement and lime) was exempted in the

current mix design. IBA was used as aerating agent and precursor together with alkaline solution to form inorganic polymer binder.

[0084] **TABLE 3: Mix proportions of examples, parts by weight**

| Mix No. | IBA | Alkaline solution | Mix duration (minutes) |
|---------|-----|-------------------|------------------------|
| 1       | 1   | 0.60              | 15                     |
| 2       | 1   | 0.60              | 30                     |
| 3       | 1   | 0.60              | 60                     |
| 4       | 1   | 0.60              | 120                    |
| 5       | 1   | 0.75              | 15                     |
| 6       | 1   | 0.75              | 30                     |
| 7       | 1   | 0.75              | 60                     |
| 8       | 1   | 0.75              | 120                    |
| 9       | 1   | 0.90              | 15                     |
| 10      | 1   | 0.90              | 30                     |
| 11      | 1   | 0.90              | 60                     |
| 12      | 1   | 0.90              | 120                    |
| 13      | 1   | 1.10              | 60                     |

5 [0085] The IBA was collected from Keppel Seghers Tuas Waste-to-Energy incineration plant. The IBA granular was first dried and followed by ball milling until the average particle size is around 20 microns.

[0086] A mixture of sodium hydroxide and sodium silicate solution was prepared. Sodium hydroxide solution with 8 mol/L concentration was prepared, and sodium silicate solution  
10 (sodium silicate solids : water = 2 : 3) with molar ratio of SiO<sub>2</sub> to Na<sub>2</sub>O of 2.9 to 3.2 was used. Mass ratio of the sodium hydroxide solution to the sodium silicate solution was 1:2.

[0087] The mixture was prepared in a mortar mixer with a planetary rotating blade. The IBA was dry mixed for approximately 1 to 2 minutes, followed by addition of alkaline solution and mixed for pre-determined time durations ranging from 15 to 120 minutes. The  
15 fresh mixture was cast into 50 mm cubic molds and vibrated for 30 seconds. The molds were covered with plastics to prevent water evaporation and the specimens were cured in elevated temperature of 75 °C right after casting. Specimens were demolded after 24 hours and then cured in elevated temperature of 75°C for another 2 days. After which, the specimens were stored in laboratory air.

20 [0088] **Example 2: Characterization (Embodiment 1)**

[0089] Tests on the compressive strength, the dry density, and the pore structure were conducted to evaluate the physical and mechanical properties of the resulting lightweight concrete. For each mix, three specimens were tested and the average was reported for the compressive strength. The loading rate was controlled at 0.25 mm/min.

5 [0090] The dry density was determined in accordance with ASTM C 138-92. The specimens were first dried in oven at 105 °C for 24 hours. The dry density was defined as the ratio between the weight and the volume of the sample.

[0091] To evaluate the macro porosity of lightweight concrete, image of the cross-section of the specimen (12 x 9 mm<sup>2</sup>) was captured by means of an optical microscope. The photo  
10 was post-proceeded into a binary image for quantitative measurement of the macro porosity of the cross-section as shown in **FIG. 1**.

[0092] **TABLE 4** summarizes the measured physical and mechanical properties of the resulting lightweight concrete. The metallic aluminum retained in IBA can react with alkaline solution to liberate hydrogen gas. The formation of porous structure decreases the dry density  
15 of lightweight concrete.

[0093] **TABLE 4:** Density, compressive strength, and macro porosity of examples

| Mix No. | Density (kg/m <sup>3</sup> ) | Compressive strength (MPa) | Porosity (%) |
|---------|------------------------------|----------------------------|--------------|
| 1       | 1008                         | 2.26                       | 10.3         |
| 2       | 1019                         | 2.54                       | 11.6         |
| 3       | 1008                         | 2.58                       | 11.9         |
| 4       | 1036                         | 2.82                       | 12.5         |
| 5       | 801                          | 1.04                       | 19.8         |
| 6       | 984                          | 1.60                       | 15.7         |
| 7       | 821                          | 0.95                       | 31           |
| 8       | 1017                         | 2.34                       | 13.8         |
| 9       | 612                          | 1.04                       | 27.3         |
| 10      | 929                          | 1.88                       | 15.6         |
| 11      | 869                          | 1.47                       | 25.3         |
| 12      | 939                          | 1.90                       | 11.9         |
| 13      | 676                          | 1.28                       | 37.3         |

[0094] **FIG. 2** shows the effect of mix duration and liquid to solid ratio (L/S) on the dry density of lightweight concrete. L/S values of 0.6, 0.75 and 0.9 were used. As can be seen  
20 from the figure, the dry density of the resulting lightweight concrete ranges from 600 kg/m<sup>3</sup> to around 1,000 kg/m<sup>3</sup>. The dry density increases with prolonged mixing in the first 30

minutes. After that, the dry density decreases when the mix duration is 60 minutes and increases again for the 120-minute mixing. Prolonged mixing may introduce two effects. On one hand, it allows escape of entrained air. On the other hand, mixing promotes reaction between metallic aluminum and alkali and increases the amount of hydrogen generation. The dry density at different mix duration is a result of these two competing mechanisms. It is plausible that the escaping of entrained air dominates the overall density when mixing time is less than 30 minutes. From 30 to 60 minutes, significant increase of hydrogen generation due to prolonged mix results a lower density. With further prolonged mixing, air escaping again dominates as most of the hydrogen reaction has completed.

[0095] As may be seen from the figure, the dry density generally decreases with increase in liquid to solid ratio. Not only does a higher L/S value increase capillary porosity, it also promotes hydrogen generation. An increase in capillary porosity and more hydrogen generation result in a lower density. Therefore, dry density of lightweight concrete may be adjusted through control of mix duration as well as liquid to solid ratio.

[0096] **FIG. 3** plots the compressive strength against the dry density of the sample. As can be seen, there is a strong positive correlation between the compressive strength and the dry density of lightweight concrete. The compressive strength was in vicinity of 1 MPa when the dry density was lower than 800 kg/m<sup>3</sup>. With the dry density exceeding 800 kg/m<sup>3</sup>, the compressive strength showed a rapid increase.

[0097] **FIG. 4** plots the dry density against macro-porosity of each sample. As can be seen, the lightweight concrete contains 10 % to 40 % visible voids. A negative linear correlation between the dry density and the porosity of lightweight concrete is observed as depicted in **FIG. 4**. The increase of porosity is therefore the principal contributor to the decrease in the dry density as well in the compressive strength.

[0098] **Example 3: Material Preparation For Manufacture of Lightweight Concrete Using Hydraulic Cement as Binder (Embodiment 2)**

[0099] Exemplary mixes disclosed herein are for preparing lightweight concrete using hydraulic cement as the binder. IBA is used as the aerating agent to replace aluminum powder and as supplement to partially replace cement.

[00100] The mix proportions are tabulated in **TABLE 5**. Type I Portland cement was used in all exemplary mixes. The IBA was collected from Keppel Seghers Tuas Waste-to-Energy incineration plant. The IBA granular was first dried and followed by ball milling until the

average particle size is around 20 microns. Normal river sand was adopted in the exemplary mixes. The mixes have a constant sand-to-binder ratio of 1 and a water-to-binder ratio between 0.35 and 0.70. The binder consists of IBA and cement in the ratio of 2-to-8 (mixes A and B), 4-to-6 (mix C1), and 6-to-4 (mix C2). Sodium hydroxide was used in mixes A and C to accelerate the hydrogen generation. The molarity of resulting alkaline solution ranges from 0.01 to 1.

[00101] **TABLE 5:** Mix proportions of examples, parts by weight

| Mix No. | Cement (kg/m <sup>3</sup> ) | IBA (kg/m <sup>3</sup> ) | Sand (kg/m <sup>3</sup> ) | Water (kg/m <sup>3</sup> ) | Water-to-binder ratio | NaOH (kg/m <sup>3</sup> ) | Molarity (mol/l) | Curing temp. |
|---------|-----------------------------|--------------------------|---------------------------|----------------------------|-----------------------|---------------------------|------------------|--------------|
| A1      | 460                         | 115                      | 575                       | 201.3                      | 0.35                  | 7.80                      | 1                | 80°C         |
| A2      | 460                         | 115                      | 575                       | 201.3                      | 0.35                  | 0.78                      | 0.1              | 80°C         |
| A3      | 460                         | 115                      | 575                       | 201.3                      | 0.35                  | 0.08                      | 0.01             | 80°C         |
| B1      | 460                         | 115                      | 575                       | 201.3                      | 0.35                  | 0.00                      | -                | 80°C         |
| B2      | 460                         | 115                      | 575                       | 201.3                      | 0.35                  | 0.00                      | -                | 20°C         |
| C1      | 345                         | 230                      | 575                       | 402.5                      | 0.70                  | 15.60                     | 1                | 80°C         |
| C2      | 230                         | 345                      | 575                       | 402.5                      | 0.70                  | 15.60                     | 1                | 80°C         |

[00102] The mixture was prepared in a mortar mixer with a planetary rotating blade. The cement, IBA and sand were dry-mixed first for approximately 1 to 2 minutes, followed by the addition of water or alkaline solution and mixed for another 2 to 3 minutes. The fresh mixture was cast into 50 mm cubic molds and vibrated for 30 seconds. The molds were covered with plastics to prevent water evaporation and the specimens were cured in room temperature of 20 °C (mix B2) or elevated temperature of 80 °C right after casting for 3 days. After that, specimens were demolded and cured in water until testing at the age of 28 days.

[00103] **Example 4: Characterization (Embodiment 2)**

[00104] Dry density was defined as a ratio of the dry weight to the volume of each specimen. The specimens were dried in oven at 105 °C for at least 24 hours. The weight and volume of the dry specimen were recorded to determine the dry density.

[00105] Compression tests were conducted using 50 mm cube specimen at the age of 28 days. The loading rate was 0.50 mm/min and only peak loads were recorded. Each data point was the average of three specimens.

[00106] To evaluate the macro-porosity of lightweight concrete, image of the cross-section of the specimen (12 × 9 mm<sup>2</sup>) was captured by means of the Nikon SMZ745T optical microscope. The photo was post-processed into a binary image for quantitative measurement of the macro-porosity of the cross-section as shown in **FIG. 5A and 5B**.

[00107] Scanning electron microscope (SEM) technique was conducted by using JEOL-7400F field-emission scanning electron microscope to examine the microstructure of lightweight concrete.

[00108] **TABLE 6:** Density and compressive strength of examples

| Mix No. | Dry density (kg/m <sup>3</sup> ) | Compressive strength (MPa) |
|---------|----------------------------------|----------------------------|
| A1      | 1324.0                           | 5.7                        |
| A2      | 1443.7                           | 12.5                       |
| A3      | 1512.0                           | 16.0                       |
| B1      | 1553.6                           | 19.5                       |
| B2      | 1604.8                           | 21.5                       |
| C1      | 1056.4                           | 2.7                        |
| C2      | 931.6                            | 1.5                        |

[00109] The compressive strength and dry density are summarized in **TABLE 6**. As can be seen, dry density of the resulting lightweight concrete ranges from 900 kg/m<sup>3</sup> to 1,600 kg/m<sup>3</sup> with compressive strength between 1.5 MPa and 21.5 MPa. The general trend shows a positive correlation between the compressive strength and the density of lightweight concretes. As shown in **FIG. 6**, the compressive strength of lightweight concrete increased with the increase of dry density, which satisfied an exponential correlation in the whole range.

[00110] **FIG. 7** shows the effect of alkali molarity (**FIG. 7A**) and IBA content (**FIG. 7B**) on the properties of lightweight concrete. As can be seen, compressive strength as well as the dry density decreases with the increase of alkali molarity or IBA content. Higher alkali molarity greatly promotes hydrogen generation from IBA resulting in more aeration and lower density and compressive strength. Similarly, higher IBA content generates more hydrogen which reduces the density of samples C1 and C2. In addition, IBA itself has lower specific gravity than cement. The inclusion of higher dosage of IBA naturally reduces the dry density of the resulting material.

[00111] The macro-porosity of aerated lightweight concrete in this research ranges from 15 % to 40 %. **FIG. 8** shows the effect of macro-porosity on the dry density and compressive strength of each mix of lightweight concrete. As can be seen from the figure, both dry density and compressive strength decrease with the increase of macro-porosity. Decrease of dry density as well as compressive strength mainly results from the increase of macro-porosity. Both suitable alkali molarity and proper dosage of IBA are the critical factors for obtaining

better pore structures and further for improving the mechanical and physical properties of lightweight concrete.

[00112] **FIG. 9** shows the microstructure of lightweight concrete prepared with tap water (denoted by TW) and sodium hydroxide solution (denoted by NH) curing in oven at 80 °C.

5 As can be seen, matrix of the TW sample (**FIG. 9A**) has a denser microstructure as compared to that of the NH sample (**FIG. 9C**). This is mainly due to accelerated hydration when sodium hydroxide is used in cement. Coarse calcium silicate hydrates and large amount of ettringite are formed at early age resulting in a loose microstructure. In addition, crystal structure is found in the wall of macro-pores of the NH sample (**FIG. 9D**). Addition of  
10 sodium hydroxide increases the concentration of hydroxyl ions in the pore solution which depresses the solubility of  $\text{Ca}^{2+}$  ions. As a result, calcium carbonate and/or calcium hydroxide may precipitate on the wall of macro-pores as shown in **FIG. 9D**.

[00113] **Example 5: Commercial applications**

[00114] According to the use and physical properties, lightweight concretes may be  
15 classified into the following three categories as shown in **TABLE 7**.

[00115] **TABLE 7:** Classification of lightweight concretes

| Classification                   | Density (kg/m <sup>3</sup> ) | Typical strength range (MPa) | Thermal conductivity (W/m-K) | Use                    |
|----------------------------------|------------------------------|------------------------------|------------------------------|------------------------|
| Structural lightweight concretes | 1,350-1,900                  | 17-63                        | 0.50-1.30                    | Structural             |
| Moderate-strength concretes      | 800-1,350                    | 7-14                         | 0.15-0.50                    | Structural /Insulating |
| Low-density concretes            | 300-800                      | 0.7-2                        | 0.05-0.15                    | Insulating             |

[00116] Lightweight concrete may be formed of the constituents as shown in **TABLE 8**.

[00117] **TABLE 8:** Typical mix proportion and cost structure of lightweight concrete

| Material        | Mix proportion (kg/m <sup>3</sup> ) | Unit price (S\$/ton) | Cost (S\$/m <sup>3</sup> ) | Cost ratio |
|-----------------|-------------------------------------|----------------------|----------------------------|------------|
| Cement          | 49                                  | 152                  | 7.4                        | 26%        |
| Sand            | 490                                 | 14                   | 6.9                        | 25%        |
| Lime            | 133                                 | 80                   | 10.6                       | 38%        |
| Gypsum          | 28                                  | 20                   | 0.6                        | 2%         |
| Aluminum powder | 0.46                                | 5,200                | 2.4                        | 9%         |
|                 |                                     | Total                | 27.9                       | 100%       |

[00118] As can be seen from **TABLE 8**, aluminum powder may make up 10 % of total material cost of lightweight concrete. Advantageously, IBA contains a significant amount of metallic aluminum due to increasing amounts of household waste, and which are being incinerated. The metallic aluminum in IBA allows its reutilization as an aerating agent in production of lightweight concrete.

[00119] Furthermore, IBA may possess high levels of calcium oxide, silicon oxide, and aluminum oxide content to allow its use as a supplementary cementitious material to partially replace cement and other binders used in lightweight concrete. The calcium oxide, silicon oxide, and aluminum oxide may possess pozzolanic properties that may lead to remarkable improvement of concrete strength.

[00120] By using municipal solid waste incineration bottom ash (IBA) as aerating agent to replace costly aluminum powder, as well as supplementary cementitious material to replace cement and other binders for the production of lightweight concrete, the proposed method provides a practical and cost effective solution to recycle and to reutilize incineration bottom ash (IBA) for civil engineering applications. The proposed method contributes not only to prolong the lifespan of landfills but also to create alternative resources from wastes. Other tangible and intangible benefits of the proposed method include alleviating the disposal costs of the ash, preserving land capacity, conserving dwindling supplies of natural raw materials, and mitigating potential environmental impacts. With lower density ( $200 \text{ kg/m}^3$  to  $1,900 \text{ kg/m}^3$ ) and thermal conductivity ( $0.05 \text{ W/m-K}$  to  $1.30 \text{ W/m-K}$ ), IBA lightweight concrete is an idea construction material to reduce self-weight of structure and to provide thermal insulation and sound-proofing.

[00121] Potential applications include structural concretes, floors, trench fills, roof insulation and other insulating purposes, as well as to make masonry units. Preliminary market analysis shows that global market demand for aerated lightweight concrete is around 110 million cubic meter per year, which represents an annual market size of S\$18 billion.

[00122] While the present invention has been particularly shown and described with reference to exemplary embodiments thereof, it will be understood by those of ordinary skill in the art that various changes in form and details may be made therein without departing from the spirit and scope of the present invention as defined by the following claims.

## CLAIMS

1. Method of manufacturing a lightweight concrete, the method comprising
  - a) providing ash obtained from incineration of waste, the ash comprising at least about 0.1 wt% metallic aluminum;
  - b) mixing the ash with a binder to form a mixture; and
  - c) curing the mixture to obtain the lightweight concrete.
2. Method according to claim 1, wherein the ash is derived from bottom ash of a solid waste incinerator facility.
3. Method according to claim 1 or 2, wherein providing ash obtained from incineration of waste comprises reducing particle size of the ash to an average particle size of about 500  $\mu\text{m}$  or less.
4. Method according to any one of claims 1 to 3, wherein providing ash obtained from incineration of waste comprises reducing particle size of the ash to an average particle size of about 20  $\mu\text{m}$  to about 40  $\mu\text{m}$ .
5. Method according to claim 3 or 4, wherein reducing particle size of the ash is carried out by ball milling.
6. Method according to any one of claims 1 to 5, wherein providing ash obtained from incineration of waste comprises drying the ash.
7. Method according to any one of claims 1 to 6, wherein the binder is selected from the group consisting of a hydraulic cement binder and an inorganic polymer binder.
8. Method according to any one of claims 1 to 7, wherein the binder comprises or consists of a hydraulic cement binder.

9. Method according to claim 8, wherein the hydraulic cement binder is selected from the group consisting of Portland cement, blended Portland cement, expansive cement, rapid setting and hardening cement, calcium aluminate cement, magnesium phosphate, and mixtures thereof.
- 5
10. Method according to claim 8 or 9, wherein the hydraulic cement binder comprises or consists of Type I Portland cement.
- 10
11. Method according to any one of claims 8 to 10, wherein mixing the ash with a binder comprises adding at least one of a pozzolanic admixture, water, lime, alkali, and a set-control admixture to the mixture.
- 15
12. Method according to claim 11, wherein the pozzolanic admixture is selected from the group consisting of fly ash, silica fume, ground granulated blast-furnace slag, and combinations thereof.
13. Method according to claim 11 or 12, wherein ratio of water to the binder is in the range of about 0.2 to about 1 by weight.
- 20
14. Method according to any one of claims 11 to 13, wherein ratio of water to the binder is in the range of about 0.35 to about 0.7 by weight.
15. Method according to any one of claims 11 to 14, wherein the lime comprises or consists of an alkaline earth metal carbonate.
- 25
16. Method according to any one of claims 11 to 15, wherein the lime comprises or consists of at least one of magnesium carbonate and calcium carbonate.
- 30
17. Method according to any one of claims 11 to 16, wherein the alkali is selected from the group consisting of an alkali metal hydroxide, an alkaline earth metal hydroxide, ammonium hydroxide, and combinations thereof.

18. Method according to any one of claims 11 to 17, wherein the alkali is selected from the group consisting of sodium hydroxide, potassium hydroxide, calcium hydroxide, magnesium hydroxide, sodium carbonate, and combinations thereof.
- 5 19. Method according to any one of claims 11 to 18, wherein the alkali comprises or consists of sodium hydroxide.
20. Method according to any one of claims 11 to 19, wherein concentration of hydroxide in the mixture is 100 mol/L or less.
- 10 21. Method according to any one of claims 11 to 20, wherein the set-control admixture is selected from the group consisting of gypsum, anhydrite, and combinations thereof.
22. Method according to any one of claims 1 to 7, wherein the binder comprises or  
15 consists of an inorganic polymer binder.
23. Method according to claim 22, wherein the inorganic polymer binder is selected from the group consisting of alkali activated aluminosilicates, geopolymers, alkali activated cement, and mixtures thereof.
- 20 24. Method according to claim 22 or 23, wherein the inorganic polymer binder is obtained by reacting an inorganic polymer precursor with an alkaline activator liquid.
- 25 25. Method according to claim 24, wherein the inorganic polymer precursor is selected from the group consisting of coal fly ash, metakaoline, waste incinerator ash, ground granulated blast-furnace slag, and combinations thereof.
- 30 26. Method according to claim 24 or 25, wherein the alkaline activator liquid comprises a silicate and/or a hydroxide selected from the group consisting of an alkali metal hydroxide, an alkaline earth metal hydroxide, ammonium hydroxide, and combinations thereof.

27. Method according to claim 26, wherein the silicate comprises or consists of an alkali metal silicate.
28. Method according to claim 26 or 27, wherein the silicate comprises or consists of sodium silicate.
29. Method according to any one of claims 26 to 28, wherein the hydroxide comprises or consists of an alkali metal hydroxide.
30. Method according to any one of claims 26 to 29, wherein the hydroxide comprises or consists of sodium hydroxide.
31. Method according to any one of claims 26 to 30, wherein concentration of hydroxide in the alkaline activator liquid is at least 1 mol/L.
32. Method according to any one of claims 26 to 31, wherein ratio of hydroxide to silicate is in the range of about 0.2 to about 3 by weight.
33. Method according to any one of claims 26 to 32, wherein the alkaline activator liquid is an aqueous solution of the silicate and/or the hydroxide.
34. Method according to any one of claims 24 to 33, wherein ratio of alkaline activator liquid to the inorganic polymer binder is in the range of about 0.3 to about 2 by weight.
35. Method according to any one of claims 1 to 34, wherein ratio of ash to binder in the mixture is in the range of about 0.01 to about 100 by weight.
36. Method according to any one of claims 1 to 35, wherein mixing the ash with a binder comprises adding at least one of a viscosity control agent, a water reducing agent, and aggregates to the mixture.

37. Method according to claim 36, wherein the viscosity control agent is selected from the group consisting of methylcellulose, polyvinyl alcohol, starch, and combinations thereof.
- 5 38. Method according to claim 36 or 37, wherein ratio of the viscosity control agent to the binder is not more than 0.05 by weight.
39. Method according to any one of claims 36 to 38, wherein the water reducing agent is selected from the group consisting of carboxylated polymer, gypsum, anhydrite, and mixtures thereof.
- 10 40. Method according to any one of claims 36 to 39, wherein ratio of the water reducing agent to the binder is in the range of about 0.001 to about 0.002 by weight.
- 15 41. Method according to any one of claims 36 to 40, wherein the aggregates comprise coarse aggregates selected from the group consisting of gravel, crushed rock, slag aggregate, recycled aggregate, crushed quartz, granite, gneiss, basalt, limestone, and combinations thereof.
- 20 42. Method according to any one of claims 36 to 41, wherein the aggregates comprise fine aggregates selected from the group consisting of sand, fly ash, silica fume, artificial light weight fine aggregate, and combinations thereof.
43. Method according to any one of claims 1 to 42, wherein curing the mixture to obtain the lightweight concrete is carried out at a temperature in the range of about 20 °C to about 80 °C.
- 25 44. Method according to any one of claims 1 to 43, wherein curing the mixture to obtain the lightweight concrete is carried out at a temperature in the range of about 75 °C to about 80 °C.
- 30

45. Lightweight concrete manufactured by a method according to any one of claims 1 to 44.
46. Lightweight concrete having a density in the range of about 200 kg/m<sup>3</sup> to about 1900 kg/m<sup>3</sup>, a compressive strength in the range of about 0.5 MPa to about 70 MPa, and a thermal conductivity in the range of about 0.05 W/m-k to about 1.3 W/m-k.
47. Use of ash obtained from incineration of waste, the ash comprising at least about 0.1 wt% metallic aluminum, as an aerating agent.
48. Use according to claim 47, wherein the aerating agent is an aerating agent in manufacture of lightweight construction materials.