30.00 ± 0.20	0.2977	W
34.32 ± 0.20	0.2611	M.

3. The molecular sieve of claim 2, having, on an anhydrous basis, a composition represented by the following empirical formula:

$$pQ:aF:(Co_xAl_yP_z)O_2$$

wherein:

(a) p is the number of moles of 1,3-diisopropylimidazolium cation (Q) per mole of $(Co_xAl_yP_z)O_2$ and p has a value from 0.001 to 1;

(b) *a* is the number of moles of fluoride ion (F) per mole of $(Co_xAl_yP_z)O_2$ and *a* has a value from 0 to 1;

(c) x, y, and z represent the mole fraction of Co, Al, and P, respectively, as tetrahedral units; and

(d) x, y, and z are each a value greater than zero and less than one and the sum of x, y, and z is one.

4. The molecular sieve of claim 3, wherein x is from 0.01 to 0.5, y is from 0.25 to 0.7, and z is from 0.25 to 0.7.

~~5. A method of preparing the crystalline cobalt aluminophosphate molecular sieve of claim 2, comprising contacting under crystallization conditions: (1) at least one source of cobalt; (2) at least one source of aluminum; (3) at least one source of phosphorus; (4) fluoride ions; and (5) a 1,3-diisopropylimidazolium ionic liquid.~~

~~6. The method of claim 5, wherein the molecular sieve is prepared from a reaction mixture comprising, in terms of mole ratios, as follows:~~

CoO/Al_2O_3	0.5 to 5
P_2O_5/Al_2O_3	0.5 to 5
Q/Al_2O_3	≥ 5
F/Al_2O_3	0.5 to 3

wherein Q is a 1,3-diisopropylimidazolium cation.

7. The method of claim 5, wherein the crystallization conditions include a temperature of from 150°C to 200°C.

8. The method of claim 5, wherein the 1,3-diisopropylimidazolium ionic liquid is 1,3-diisopropylimidazolium bromide.

9. The method of claim 5, further comprising recovering the ionic liquid from the reaction mixture.