

REACTIVE CRYSTALLIZATION METHOD TO IMPROVE PARTICLE SIZE

Background of the Invention

This invention relates to a method for simultaneously synthesizing a
5 compound and isolating the product in a crystallization procedure, which controls
particle size.

U. S. Patent number 5,314,506, incorporated herein by reference, provides a
method of impinging fluid jet streams in a continuous crystallization process to
achieve high intensity micromixing of fluids so as to form a homogenous composition
10 prior to the start of nucleation. This process permits direct crystallization of the high
surface area particles of high purity and stability. Particle size may be controlled by
varying the concentration of solutions, temperature and velocity of the solution
through the jets. With this method it is possible to prepare compounds of 5 to 1000
microns.

15 WO 00/44468, incorporated herein by reference, describes an apparatus and
process for crystallizing submicron-sized particles with the introduction of a sonication
probe with impinging jets.

The use of an impinging jet device optionally with a sonication probe to
achieve high intensity mixing coupled with a chemical reaction to provide a new
20 chemical compound of controlled particle size of crystalline materials is novel.

Summary of the Invention

The present invention provides a method for preparation and crystallization of
pharmaceutical compounds or their intermediates which directly produces high
surface area end product crystals with greatly improved stability and purity and
25 thereby eliminates the need for subsequent high intensity milling to meet
bioavailability requirements. By removing the need for milling, the novel jet process
avoids associated problems of noise and dusting, cuts yield loss, and saves the time
and extra expense incurred during milling. It also removes an extra opportunity for
personnel contact with a highly potent pharmaceutical agent, or for adverse effects
30 on labile compounds. The small particle size attained with the jet process is
consistent within a single run and is reproducible between runs. Reproducibility is an
attribute of this process that is not common to "reverse addition" methods typically
used to produce small crystals.

The pure, high surface area particles that result from the jet process also display superior crystal structure when compared to particles formed via standard slow crystallization plus milling methods using the same quality and kind of feed compound. Improvements in crystal structure result in decreases in decomposition rate and therefore longer shelf life for the crystallized product or a pharmaceutical composition containing the crystallized material.

This invention provides a process for preparation and crystallization of a chemical compound in a continuous process.

More particularly, this invention relates to the use of impinging jets to achieve high intensity micromixing of solvents containing chemical reactants so as to produce a chemical reaction which forms a reaction product under high supersaturation conditions leading to a rapid nucleation in a continuous reaction and crystallization process.

Nucleation and precipitation can be initiated by utilizing the effect of temperature reduction on the solubility of the compound to be crystallized in a particular solvent (thermoregulation), or by taking advantage of the solubility characteristics of the compound in solvent mixtures, or by some combination of the two techniques. Further, the product of the reaction will usually be highly insoluble in the final solvent or combination of solvents.

The novel process of this invention provides for the direct crystallization of high surface area particles of high purity and stability.

In a first aspect of the invention, there is provided a process for synthesis and crystallization of a chemical compound comprising contacting one or more jet streams of a solution in a first solvent of a first reactive intermediate and one or more jet streams of a solution in a second solvent of a second reactive intermediate, said jet streams impinging to create high turbulence at their point of impact under conditions of temperatures and pressure which permit reaction of said first and second reactive intermediates to produce a product; and

selecting said first and second solvents so that said product is of limited solubility in a mixture of said first and second solvents; and

said jet streams impinging to create high turbulence at their point of impact and each jet stream having sufficient linear velocity to achieve high intensity

micromixing of said solutions followed by reaction of first and second reactive intermediates, followed by nucleation of said product and production of small crystals of controlled particle size.

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A first embodiment may relate to a process as defined in the first aspect wherein one of said first and second reactive intermediates may be a basic intermediate and the other intermediate may be an acidic intermediate.

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A second embodiment may relate to a process as defined in the first aspect wherein one of said first and second reactive intermediates may be a zwitterion and the other intermediate may be an acid. A third embodiment may relate to a process as defined in the first aspect wherein one of said first and second reactive intermediates may be a zwitterion and the other intermediate may be a base.

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A fourth embodiment may relate to a process as defined in the first aspect wherein one of said first and second reactive intermediates may be an organic salt form and the other intermediate may be a neutralizing acid compound.

A fifth embodiment may relate to a process as defined above wherein one of said first and second reactive intermediates may be an organic salt form and the other intermediate may be a neutralizing basic compound.

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A sixth embodiment may relate to a process as defined above, wherein said synthesis results in the forming or breaking of a covalent bond and wherein the product or its corresponding salt form crystallizes.

This invention also discloses a process wherein said first solution may be a pharmaceutical salt form solution and said second solution may be a solution containing

either an acid or a base. For example, a solution of Voriconazole camphorsulfonate may be reacted with a solution of sodium acetate to produce Voriconazole free base with controlled particle size. Such a reaction is known as free-basing or conversion to the free base. The process may also be applicable to the conversion to free acids by treating a pharmaceutical salt form with acid to form the free acid. Salt forms include, but are not limited to, those compounds containing anions such as hydrochloride, acetate, besylate, citrate, hydrobromide, D or L lactate, mesylate, succinate, camphorsulfonate, sulfate, D or L tartrate, stearates, tosylates, and cations such as calcium, potassium, sodium and ethylenediamine.

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A seventh embodiment may relate to a process as defined above wherein said chemical compound to be formed may be a pharmaceutically acceptable salt

selected from the group consisting of: hydrochloride, acetate, besylate, citrate, hydrobromide, D or L lactate, mesylate, succinate, sulfate, D or L tartrate, stearate, tosylate; calcium, potassium and sodium; and ethylenediamine.

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An eighth embodiment may relate to a process as defined above wherein said chemical compound may be Ziprasidone hydrochloride monohydrate.

A ninth embodiment may relate to a process wherein one reactive intermediate may be Ziprasidone free base and the other reactive intermediate may be an aqueous hydrochloric acid solution.

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A tenth embodiment may relate to a process as defined above wherein said product to be formed may be Voriconazole free base.

An eleventh embodiment may relate to a process as defined above wherein the basic intermediate may be selected from, but not limited to, sodium acetate, potassium acetate, sodium hydroxide, potassium hydroxide, sodium carbonate, sodium bicarbonate, and said acidic intermediate is Voriconazole R-(-)- camphorsulfonic acid salt.

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A twelfth embodiment may relate to a process as defined above wherein the compound to be crystallized may be 5,8,14-triazatetracyclo[10.3.1.0^{2,11}.0^{4,9}]-hexadeca-2(11),3,5,7,9-pentaene L-tartrate.

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A thirteenth embodiment may relate to a process wherein the first reactive intermediate may be the 5,8,14-triazatetracyclo[10.3.1.0^{2,11}.0^{4,9}]-hexadeca-2(11),3,5,7,9-pentaene free base and the second reactive intermediate may be L-tartaric acid.

A fourteenth embodiment may relate to a process as defined above wherein the desired product may be in the form of crystalline aggregates which may be of controlled size.

A fifteenth embodiment may relate to a process wherein the linear velocity of jet streams may be at least 5 m/sec.

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A sixteenth embodiment may relate to a process wherein said product may be solvated.

A seventeenth embodiment may relate to a sonication probe along with impinging jets to achieve high intensity micromixing.

In a second aspect of the present invention, there is provided a compound when made by the process as defined in the first aspect of the invention and its embodiments.

Brief Description of the Drawings

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An embodiment of the invention has been chosen for purposes of illustration and description, and is shown in the accompanying drawings forming a part of the specification wherein:

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Figure 1 is a schematic diagram showing a crystal production system depicting the jet chamber 3, the transfer line 4, the stirred vessel 5, the agitation device 6 and the entry point of two fluids 1 and 2 into the system.

After micromixing in a jet chamber, the material leaves the jet chamber as depicted in Figure 1, travels into a stirred vessel 5 either directly or via a transfer line 4, and after an appropriate age time, the product suspension flows out of the vessel as indicated by arrow A.

In another embodiment, shown in Figure 2, two fluids 11 and 12 enter directly into the stirred vessel 13 containing a liquid which is a solvent or antisolvent where the jets 16 emit fluid jet streams that impinge and micromix near the effluent stream of the impeller 15.

Detailed Description of the Invention

The novel process of this invention involves the use of jets to create impinging fluid jet streams and thereby achieve high intensity micromixing and reaction of the fluids prior to nucleation in a crystallization process. Two or more jets may be used to micromix two or more fluids. Preferably, two jets are used to micromix two fluids. When using two jets, preferably the two impinging jet streams should be substantially diametrically opposed to each other, i.e., they should be at or close to a 180 degree angle to each other from an overhead view. Figure 1 shows one embodiment of this invention wherein two jets are employed; fluids 1 and 2 enter the jet chamber 3 where micromixing takes place.

Regardless of the number of jets used, the jet nozzles should be placed so that the fluid streams they emit will impinge, either inside the jet chamber or directly in the stirred vessel. The fluid jets must impinge to create an immediate high turbulence impact; concentric or converging jets generally create insufficient turbulence to achieve the required micromixing. When two jets are used with a jet chamber, as show in Figure 1, the two jet nozzles are preferably arranged so that they are substantially diametrically opposed to each other with their outlet tips directed to face each other; i.e., the two jet nozzles are at or close to a 180 degree angle to each other from an overhead view. Preferably, each jet outlet nozzle can have a slight

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downward angle from the horizontal of about 10 degrees to help the flowing material move down and out of the chamber.

Likewise, two jet nozzles placed directly inside the stirred vessel are preferably arranged so that they are substantially diametrically opposed to each other with their outlet tips directed to face each other. When the jet nozzles are so placed, each nozzle can have a slight upward or downward angle from the horizontal of from 0 degrees up to about 15 degrees, but preferably the two nozzles have just enough downward angle from the horizontal (ca. 13 degrees) to ensure that the fluid stream of one will not enter the outlet hole of the opposite nozzle, possibly causing crystallization in the jet. The use of a sonication probe along with impinging jets aids in achieving high intensity micromixing in a continuous process.

One jet nozzle is used to transport one of the two fluids from an external source into the chamber and the other jet is used to similarly transport the other fluid. The distance between the nozzle tips inside the jet chamber or stirred vessel should be such that the hydrodynamic form of each fluid jet stream remains essentially intact up to the point of impingement. Therefore, the maximum distance between the nozzle tips will vary depending on the linear velocity of the fluids inside the jet nozzles. To obtain good results for generally nonviscous fluids, linear velocity in the jet nozzles should be at least about 5 meters/sec., more preferably above 10 meters/sec., and most preferably between about 20 to 25 meters/sec., although the upper limit of linear velocity is only limited by the practical difficulties involved in achieving it. Linear velocity and flow rate can both be controlled by various known methods, such as altering the diameter of the entry tube and/or that of the nozzle outlet, and/or varying the strength of the external force that moves the fluid into and through the nozzle. Each jet apparatus can be manipulated independently to attain a desired final fluid composition ratio. When the desired flow ratio of one jet to the other differs from unity, preferably the difference is compensated for by appropriate sizing of the entry tubes. For example, if a 4:1 volumetric ratio of feed solution to anti-solvent is desired, the entry tube delivering feed solution should be twice the diameter of the entry tube delivering anti-solvent. When the jet streams impinge inside a jet chamber, residence time for the fluid inside the jet chamber is typically very short, i.e., less than ten seconds.

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A transfer line 4 as shown in Figure 1 may or may not be used to deliver the fluid mixture into a stirred vessel 5 from the jet chamber. Solvent, anti-solvent or mixtures thereof optionally containing seed and optionally heated to attain optimum results can be put inside the stirred vessel at the start of the process before the
5 micromixed fluids enter the stirred vessel; this technique is especially preferred when the jet streams impinge directly in the stirred vessel. Crystal digestion (Ostwald ripening, improvement of surface structure) takes place inside the stirred vessel.

Solvents used in this invention are preferred to be miscible. The first solvent and second solvent may be identical provided that the ultimate product has limited
10 solubility in the common solvent.

In the usual case, the first and second solvents are not identical, the first solvent being adapted to dissolve one reactive intermediate and the second solvent the second reactive intermediate. At least one of the solvents should be an "anti-solvent" which is chosen for its relatively low solvation properties with respect to the
15 product. Although this invention has been exemplified for preparing and crystallizing pharmaceutical salts, it will be obvious to those skilled in this art that the process described herein is applicable to prepare many types of small molecules with controlled particle size. The process is particularly useful for single step reactions, which proceed at a high rate under moderate conditions, for example, salt-formations,
20 free-basing and nucleophilic reactions.

In the present invention, in certain cases, first and second solvents can be identical. This has been demonstrated and applies when both, first and second reactive intermediates are soluble in said solvent but whose reaction product or resulting salt form is highly insoluble in said solvent. In this case, the solvent acts as
25 both the solvent and antisolvent for the reactive crystallization process.

The following examples are given for the purpose of illustrating the present invention and should not be construed as limitations on the scope or spirit of the present invention.

Examples

Example 1

Reaction and Crystallization of Ziprasidone to achieve desired salt form

Ziprasidone (CAS-138982-67-9) free base (3.7 grams) was dissolved into 200
5 ml (55 volumes) of tetrahydrofuran (THF) in a 250 ml 3-neck round bottom flask and
heated to reflux at 65 °C for 20 minutes. The solution was hot filtered through a glass
sintered Buchner funnel. The saturated solution was added to a 2L jacketed glass
vessel preheated to a jacket temperature of 80 °C and equipped with reflux
condenser. The bottom outlet of this vessel is connected to a pump where the liquid
10 is pumped to a 1/16 OD jet nozzle (0.02 inch ID). 200 ml of an aqueous 0.8 wt %
HCl solution is added to a second 2L jacketed vessel held at 25°C. The bottom outlet
of the second vessel is plumbed to a second 1/16 OD jet nozzle (0.02 inch ID). The
jets were secured inside a 1 liter jacketed receiver vessel and were diametrically
opposed to each other with approximately 1 cm of distance between the two jets such
15 that the liquid streams meet at a point of impingement to create a vertical
impingement film. The 1-liter jacketed receiving vessel was held at 20 °C and was
empty at the start of the run. The impinging jets were started with flow rates of 120
ml/min for the THF solution and 115 ml/min for the HCl solution. Linear velocities
were 9.9 m/sec and 9.5 m/sec for the THF and HCl solutions respectively. The
20 solution slurry was collected in the receiver vessel with the volume level eventually
rising above the jets. A Rushton turbine impeller at 200 rpm was used in the receiver
to provide mixing. The slurry was allowed to age for 36 minutes. The crystals were
filtered and washed with 150 ml of DI water and dried in a crystallizing dish over
night.

25 X-ray diffraction (XRD) and differential scanning calorimetry (DSC) indicated
the correct form of Ziprasidone hydrochloride monohydrate was produced by the
impinging jet reactive crystallization process. The dried powder crystals were sieved
through a 600 micron sieve as a means to delump. The particles were analyzed by
Malvern using hexane as a dispersant. The particle size results indicated that 90% of
30 the particles, D_{90} , by volume were less than 83 microns while the volume mean
diameter, $D[4,3]$, was 43 microns. No particles with diameters of less than 0.9 micron
were detected indicating that fines were not produced.

Example 2

Reaction and crystallization of Ziprasidone to achieve desired salt form to obtain preferred particle size

The reaction crystallization of example 1 was conducted with the following modifications: 60 volumes of THF were used relative to Ziprasidone free base. The temperature of the Ziprasidone free base in THF solution was maintained at 60°C. No hot filtration was performed prior to preparing the Ziprasidone solution in the 2L vessel. The aqueous HCl solution was maintained at 15 °C while the receiving vessel was held at 30 °C. The flow rates for the Ziprasidone free base solution in THF and the aqueous HCl solution were 120 ml/min and 100 ml/min respectively and were run for 3 minutes. The age time for the crystals was 2 minutes. The crystals were vacuum filtered, washed with 160 ml DI water, and air dried.

XRD and DSC indicated that the correct form of Ziprasidone HCl salt was produced by the reactive crystallization process. The crystals were sieved through a 600 micron sieve to delump and the particles were analyzed by Malvern using hexane as a dispersant. The particle size results indicated that 90% of the particles by volume were less than 41 microns while the volume mean diameter, D[4,3], was 22.5 microns. No particles of less than 1.0 micron were detected indicating that fines were not produced. According to patent EU 0 965 343 A2 (08.06.1999), the most preferred mean particle size of Ziprasidone hydrochloride monohydrate is in the 5-30 micron range obtained through milling operations which the present invention achieves without the need for high intensity milling. The particle sizes stipulated herein and in the claims refer to particle sizes determined with Malvern light scattering.

In a comparative 15.6 kg sample produced by the typical batch crystallization process method followed by milling using a pin mill at 10,000 rpm, the volume mean diameter, D[4,3], was 23.6 microns with 90% of the particles by volume less than 51 microns, while 5.4% of the batch by the volume distribution had particle size less than 1 micron indicating a much greater number of fines. The impinging jet produced reaction crystallization material, in both examples 1 and 2, showed a narrower particle size distribution with fewer fines produced while eliminating the milling operation.

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Example 3

Reaction and crystallization of Ziprasidone to achieve desired salt form but with larger particle size

5 The reaction crystallization of example 1 was modified to demonstrate ability to selectively control and achieve a larger particle size distribution. The modifications to example 1 were: a) 3.1 g of Ziprasidone free base dissolved in 65 volumes of THF was used, b) no hot filtration was performed in preparing the Ziprasidone free base THF solution, and c) the age time for the resulting crystals, after the impingement
10 period, was 55 minutes prior to filtration. The crystals were vacuum filtered, washed with 150 ml of deionized water, and air dried. The dry Ziprasidone hydrochloride monohydrate particles were free flowing. Microscopy revealed the formation of a uniform distribution of crystalline aggregates that give superior filtration and powder properties such as flowability. Thus this method can influence degree of aggregation
15 and provide benefits to bulk handling characteristics. The powder was not sieved prior to Malvern light scattering as it required no delumping. The volume mean particle size, $D[4,3]$, was 104 micron with 90% by volume, D_{90} , less than 210 microns. No particles below 1 micron were measured, indicating no fines were produced.

Example 4

Reaction-Crystallization of Voriconazole Camphorsulfonic Acid Salt to Voriconazole Free base (U.S. Patent No. 5,567,817; CAS-137234-62-9)

28.75 g of Voriconazole R-(-)-camphorsulfonic acid salt was dissolved in 120 ml of a 50:50 volume ratio of ethanol and deionized water and placed in a 2L jacketed
25 vessel equilibrated at 30 °C. 4 liters of an aqueous solution of sodium acetate (3.4 wt%) was placed in a stainless steel pressure can at room temperature and pressurized to 26 psig with nitrogen. Each vessel was plumbed to individual jets assembled in a chamber similar to the configuration depicted in Figure 1. The jets were 1/16 inch OD and 0.02 inch ID. The flow of sodium acetate solution was started
30 through the jet at 115 ml/min, controlled via the nitrogen pressure. The Voriconazole camphorsulfonic acid salt solution was then turned on with flow rate of 115 ml/min through the jet. The flow of the two solutions created an impingement zone between

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the two jets and white crystalline material was observed to be forming in the impingement chamber which flowed down to the receiving vessel that was equilibrated to 5 °C.

5 The jets were run until the Voriconazole salt solution was consumed. The slurry was aged for 10 minutes then vacuum filtered, and washed with 70 ml DI water.

The crystals were dried in a vacuum oven at 50 °C and 10 inches of Hg with an air bleed for 24 hours. The dried material was sieved through a 30 mesh (600 µm) screen as a means to delump. Differential scanning calorimetry and x-ray diffraction
10 confirmed the complete conversion to the Voriconazole free-base. No unreacted Voriconazole camphorsulfonate was detected.

Dry particle sizing analysis (SympaTec) revealed a volume mean diameter of 22 microns with 90% of particles (D[v,0.9]) less than 41 microns and 50% of particles (D[v,0.50]) less than 18 microns. Specification of product conventionally obtained by
15 jet milling is 90% less than 130 microns and 50% less than 50 microns.

Example 5

Reactive crystallization of 5,8,14-triazatetracyclo[10.3.1.0^{2,11}.0^{4,9}]-hexadeca-
2(11),3,5,7,9-pentaene to achieve desired salt form and particle size using
20 submerged impinging jets

1.006 grams of 5,8,14-triazatetracyclo[10.3.1.0^{2,11}.0^{4,9}]-hexadeca-
2(11),3,5,7,9-pentaene free base was dissolved in 30 ml of ethyl acetate and 5 mL of methanol in a conical shaped centrifuge tube and held at room temperature. In a separate conical shaped centrifuge tube 0.96 g of (L)-tartaric acid was dissolved in 40
25 mL methanol at room temperature.

A receiving vessel contained 250 ml of methanol and 250 ml of ethyl acetate to which 20 milligrams of jet milled 5,8,14-triazatetracyclo[10.3.1.0^{2,11}.0^{4,9}]-hexadeca-
2(11),3,5,7,9-pentaene (L)-tartrate were added as seed. The seed had a volume mean diameter of 2.7 microns. The receiver was equipped with magnetic stirrer to
30 suspend the seed in the methanol-ethyl acetate solvent system. The impinging jet apparatus was submerged in the receiver vessel with both jets having inside diameters of 0.007 inches-(ID).

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Flow of neat methanol and ethyl acetate was initiated through their respective jets at 20 ml/min for 5 seconds from separate reservoirs to establish steady state flow. By way of two 3-way valves the solutions of 5,8,14-triazatetracyclo[10.3.1.0^{2,11}.0^{4,9}]-hexadeca-2(11),3,5,7,9-pentaene free base solution and (L)-tartaric acid were switched on so they would impinge each other at the same 20 ml/min flow rate. The velocities through the jets were 13.4 meters/sec. After 1 minute of impingement the flows were stopped and the resulting slurry was allowed to age in the agitated receiver for 5 minutes. The slurry was vacuum filtered and dried at 40 °C overnight and delumped through a 500 micron sieve. Polymorph
10 characterization by x-ray diffraction revealed it to be of the desired polymorphic form. Particle size analysis (Aerosizer) resulted in a volume mean diameter of 10 microns with narrow particle size distribution: 5% of particles were less than 5 microns and 95% of the particle were less than 15 microns.

CLAIMS

1. A process for synthesis and crystallization of a chemical compound comprising contacting one or more jet streams of a solution in a first solvent of a first reactive intermediate and one or more jet streams of a solution in a second solvent of a second reactive intermediate, said jet streams impinging to create high turbulence at their point of impact under conditions of temperature and pressure which permit reaction of said first and second reactive intermediates to produce a product; and

selecting said first and second solvents so that said product is of limited solubility in a mixture of said first and second solvents; and

said jet streams impinging to create high turbulence at their point of impact and each jet stream having sufficient linear velocity to achieve high intensity micromixing of said solutions followed by reaction of first and second reactive intermediates, followed by nucleation of said product and production of small crystals of controlled particle size.

2. The process of claim 1, wherein one of said first and second reactive intermediates is a basic intermediate and the other intermediate is an acidic intermediate.

3. The process of claim 1, wherein one of said first and second reactive intermediates is a zwitterion and the other intermediate is an acid.

4. The process of claim 1, 2 or 3, wherein one reactive intermediate is Ziprasidone free base and the other reactive intermediate is aqueous HCl solution.

5. The process of claim 2 or 4, wherein the basic intermediate is selected from, but not limited to, sodium acetate, potassium acetate, sodium hydroxide, potassium hydroxide, sodium carbonate, sodium bicarbonate, and said acidic intermediate is Voriconazole R-(-)-camphorsulfonic acid salt.

6. The process of claim 1 or 2, wherein one reactive intermediate is 5,8,14-triazatetracyclo[10.3.1.0^{2,11}.0^{4,9}]-hexadeca-2(11),3,5,7,9-pentaene free base and the other reactive intermediate is a L-tartaric acid solution.

7. The process of claim 1, wherein one of said first and second reactive intermediates is a zwitterion and the other intermediate is a base.

8. The process of claim 1, wherein one of said first and second reactive intermediates is in an organic salt form and the other intermediate is a neutralising acid compound.
9. The process of claim 1, wherein one of said first and second reactive intermediates is in an organic salt form and the other intermediate is a neutralising basic compound.
10. The process of any one of the preceding claims, wherein said synthesis results in the forming or breaking of a covalent bond and where said product or its corresponding salt form crystallises.
11. The process of any one of the preceding claims, wherein said chemical compound is a pharmaceutically acceptable salt selected from the group consisting of: hydrochloride, acetate, besylate, citrate, hydrobromide, D or L lactate, mesylate, succinate, sulfate, D or L tartrate, stearate, tosylate, calcium, potassium, sodium and ethylenediamine.
12. The process of any one of the preceding claims, wherein the desired product is in the form of crystalline aggregates which are of controlled size.
13. The process of any one of the preceding claims, wherein said product is Voriconazole free base.
14. The process of any one of the preceding claims, wherein said product produced upon crystallisation is solvated.
15. The process of any one of the preceding claims, wherein at least one of said jet streams is directed about 10° downward to avoid crystallisation of said product in said jets.
16. The process of any one of the preceding claims, wherein the linear velocity of the jet streams is at least 5 m/sec.

17. The process of any one of the preceding claims, wherein a sonication probe is included with said impinging jet streams.
18. The process of any one of the preceding claims, wherein said chemical compound is Ziprasidone hydrochloride monohydrate.
19. The process of any one of the claims 1 to 17, wherein said chemical compound is 5,8,14-triazatetracyclo[10.3.1.0^{2,11}.0^{4,9}]-hexadeca-2(11),3,5,7,9-pentaene(L)-tartrate.
20. A compound when made by the process as defined in any one of claims 1 to 19.