

MAGNESIUM ALLOY, METHOD FOR THE PRODUCTION THEREOF AND USE THEREOF

This patent application relates to a magnesium alloy and to a method for the production thereof and to the use thereof.

It is known that the properties of magnesium alloys are decisively defined by the type and quantity of the alloying elements and impurities as well as the production conditions. The effects of the alloying elements and impurities on the properties of the magnesium alloys have been long known to a person skilled in the art and illustrate the complex nature of determining the properties of binary or ternary magnesium alloys for the use thereof as implant materials.

The alloying element used most frequently for magnesium is aluminum, resulting in increased tensile strength due to solid solution and precipitation hardening and fine grain formation, but also in microporosity. Moreover, in the melt aluminum shifts the iron precipitation boundary toward drastically lower iron contents at which the iron particles precipitate or form intermetallic particles together with other elements.

Calcium exhibits a pronounced grain refining effect and worsens the castability and corrosion resistance.

Undesirable accompanying elements in magnesium alloys include iron, nickel, cobalt and copper, which cause a considerable increase in the corrosion tendency due to the electropositive nature thereof.

Manganese can be found in all magnesium casting alloys and binds iron in the form of AlMnFe precipitations, whereby the formation of local elements is reduced. On the other hand, manganese is not able to bind all the iron, and therefore a remainder of iron and a remainder of manganese always remain in the melt.

Silicon lowers the castability and viscosity, and as the content of Si rises, a worsened corrosion behavior is to be expected. Iron, manganese and silicon have a very high tendency to form an intermetallic phase.

5 The electrochemical potential of this phase is very high and can thus act as a cathode controlling the corrosion of the alloy matrix.

As a result of solid solution hardening, zinc improves the mechanical properties and results in grain refining, however it also leads to microporosity with a tendency toward hot cracking starting at a content of 1.5 to 2% by weight in binary Mg-Zn and ternary
10 Mg-Al-Zn alloys.

Alloying additions made of zirconium increase the tensile strength without lowering the expansion and lead to grain refining, but also to a strong impairment of dynamic recrystallization, which is manifested in an increase of the recrystallization temperature and therefore requires high energy expenditure. Moreover, zirconium cannot be added to
15 melts containing aluminum and silicon because the grain refining effect is lost.

Rare earths such as Lu, Er, Ho, Th, Sc and In all exhibit a similar chemical behavior and form eutectic systems with partial solubility on the magnesium-rich side of the binary phase diagrams such that precipitation hardening is possible.

The addition of further alloying elements, in conjunction with the impurities, is known to
20 cause the formation of different intermetallic phases in binary magnesium alloys. For example, the intermetallic phase $Mg_{17}Al_{12}$ forming at the grain boundaries is brittle and limits the ductility. As compared to the magnesium matrix, this intermetallic phase is more noble and able to form local elements, whereby the corrosion behavior worsens.

25 In addition to these influencing factors, the properties of the magnesium alloys also decisively depend on the metallurgical production conditions. Conventional casting methods automatically introduce impurities when adding, by alloying, the alloying elements. The prior art (US 5,055,254 A) therefore defines tolerance limits for impurities in magnesium casting alloys, which, for example for a magnesium-aluminum-zinc alloy
30 containing approximately 8 to 9.5% Al and 0.45 to 0.9% Zn, mentions tolerance limits of 0.0015 to 0.0024% Fe, 0.0010% Ni, 0.0010 to 0.0024% Cu and no less than 0.15 to 0.5% Mn.

Tolerance limits for impurities in magnesium and the alloys therefore are stated in a variety of technical literature as follows:

Alloy	Production	State	Fe	Fe/Mn	Ni	Cu
Pure Mg	no information		0.017		0.005	0.01
AZ 91	Die casting	F		0.032	0.005	0.040
	High-pressure die casting			0.032	0.005	0.040
	Low-pressure die casting			0.032	0.001	0.040
		T4		0.035	0.001	0.010
		T6		0.046	0.001	0.040
	Gravity die casting	F		0.032	0.001	0.040
AM60	Die casting	F		0.021	0.003	0.010
AM50	Die casting	F		0.015	0.003	0.010
AS41	Die casting	F		0.010	0.004	0.020
AE42	Die casting	F		0.020	0.020	0.100

- 5 It has been found that these tolerance definitions are not sufficient to reliably exclude the formation of corrosion-promoting intermetallic phases, which in terms of electrochemistry have a more noble potential than the magnesium matrix.

10 Biodegradable implants require a load-bearing function and consequently strength, together with sufficient expandability, during the physiologically necessary support periods thereof. However, especially in this respect, the known magnesium materials cannot come even close to the properties achieved by permanent implants such as titanium, CoCr alloys and titanium alloys. The ultimate tensile strength R_m for implants is approximately 500 MPa to >1000 MPa, while that of magnesium materials is <275 MPa
15 so far, and in most cases <250 MPa.

Another drawback of many technical magnesium materials is that the difference thereof between ultimate tensile strength R_m and proof stress R_p is only small. In the case of implants that allow plastic deformation, such as cardiovascular stents, this means that no
20 further resistance exists against deformation after initial deformation of the material, and

the regions that have already been deformed are deformed further without any load increase, whereby overstretching of parts of the component may be caused and fracture may occur.

5 Many magnesium materials, such as the alloys of the AZ group, for example, additionally exhibit a clearly pronounced mechanical asymmetry, which is manifested in the difference in the mechanical properties, especially the proof stress R_p with tension load and compression load. Such asymmetries are created, for example, during forming processes such as extrusion, rolling and drawing, which are used to produce suitable
10 semi-finished products. A difference between the proof stress R_p during tension and the proof stress R_p during compression that is too large may result in inhomogeneous deformation of a component, such as a cardiovascular stent, which later undergoes multiaxial deformation, and may cause cracking and fracture.

15 Because of the low number of crystallographic slip systems, magnesium alloys can generally also form textures during forming processes such as extrusion, rolling and drawing used to produce suitable semifinished products by orienting the grains during the forming process. Specifically, this means that the semifinished product has different properties in different directions in space. For example, high deformability or elongation
20 at fracture occurs in one direction in space after forming, and reduced deformability or elongation at fracture occurs in another direction in space. The formation of such textures should likewise be avoided, because a stent is subjected to high plastic deformation, and reduced elongation at fracture increases the risk of failure of the implant. One method for substantially avoiding such textures during forming is to adjust as fine a grain as possible
25 prior to forming. Because of the hexagonal lattice structure of magnesium materials, the ability of these materials to deform at room temperature is low, which is characterized by slip in the base plane. If the material additionally has a coarse microstructure, which is to say a coarse grain, so-called twinning is forcibly produced upon further deformation, at which shear strain occurs, which transforms a crystal region into a position that is mirror
30 symmetrical to the starting position.

The resulting twin grain boundaries constitute weak points in the material, where incipient cracking starts, especially with plastic deformation, which ultimately leads to the destruction of the component.

If the grain of the implant materials is sufficiently fine, the risk of such implant failure is drastically reduced. Implant materials should therefore have as fine a grain as possible so as to prevent such undesirable shear strain.

All available technical magnesium materials for implants are subject to high corrosion in physiological media. Attempts have been made in the prior art to curb the corrosion tendency by providing the implants with a corrosion-inhibiting coating, for example made of polymeric materials (EP 2 085 100 A2, EP 2 384 725 A1), an aqueous or alcoholic conversion solution (DE 10 2006 060 501 A1) or an oxide (DE 10 2010 027 532 A1, EP 0 295 397 A1).

The use of polymeric passivation layers is highly contested, because virtually all appropriate polymers also cause strong inflammations in the tissue at times. Thin structures without such protective measures do not reach the required support periods. The corrosion on thin-walled traumatological implants is often times accompanied by an excessively fast loss of tensile strength, which poses an additional burden by forming excessive amounts of hydrogen per unit of time. The consequences are undesirable gas inclusions in the bones and tissue.

In the case of traumatological implants having larger cross-sections, there is a need to be able to deliberately control the hydrogen problem and the corrosion rate of the implant by way of the structure thereof.

Specifically with biodegradable implants, there is a desire for maximum biocompatibility of the elements, because all the chemical elements that are contained are absorbed by the body after decomposition. In any case, highly toxic elements such as Be, Cd, Pb, Cr and the like should be avoided.

Degradable magnesium alloys are especially suitable for implementing implants which have been employed in a wide variety of forms in modern medical technology. Implants are used, for example, to support vessels, hollow organs and vein systems (endovascular

implants, such as stents), for fastening and temporarily fixing tissue implants and tissue transplantations, but also for orthopedic purposes, such as nails, plates or screws. A particularly frequently used form of an implant is the stent.

5 The implantation of stents has become established as one of the most effective therapeutic measures for the treatment of vascular diseases. Stents have the purpose of assuming a supporting function in hollow organs of a patient. For this purpose, stents featuring conventional designs have a filigree supporting structure comprising metal struts, which is initially present in compressed form for introduction into the body and is
10 expanded at the site of the application. One of the main application areas of such stents is to permanently or temporarily widen and hold open vascular constrictions, particularly constrictions (stenosis) of coronary blood vessels. In addition, aneurysm stents are known, which are used primarily to seal the aneurysm. The support function is additionally provided.

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The implant, notably the stent, has a base body made of an implant material. An implant material is a non-living material, which is employed for applications in medicine and interacts with biological systems. A basic prerequisite for the use of a material as an implant material, which is in contact with the body environment when used as intended,
20 is the body friendliness thereof (biocompatibility). For the purpose of the present application, biocompatibility shall be understood to mean the ability of a material to induce an appropriate tissue reaction in a specific application. This includes an adaptation of the chemical, physical, biological, and morphological surface properties of an implant to the recipient's tissue with the aim of a clinically desired interaction. The
25 biocompatibility of the implant material is also dependent on the temporal process of the reaction of the biosystem in which it is implanted. For example, irritations and inflammations occur in a relatively short time, which can lead to tissue changes. Depending on the properties of the implant material, biological systems thus react in different ways. According to the reaction of the biosystem, the implant materials can be
30 divided into bioactive, bioinert and degradable or resorbable materials.

Implant materials comprise polymers, metallic materials, and ceramic materials (as coatings, for example). Biocompatible metals and metal alloys for permanent implants comprise, for example, stainless steels (such as 316L), cobalt-based alloys (such as CoCrMo cast alloys, CoCrMo forge alloys, CoCrWNi forge alloys and CoCrNiMo forge alloys), technical pure titanium and titanium alloys (such as cp titanium, TiAl6V4 or TiAl6Nb7) and gold alloys. In the field of biocorrodible stents, the use of magnesium or technical pure iron as well as biocorrodible base alloys of the elements magnesium, iron, zinc, molybdenum, and tungsten are proposed.

The use of biocorrodible magnesium alloys for temporary implants having filigree structures is made difficult in particular in that the degradation of the implant progresses very quickly in vivo. So as to reduce the corrosion rate, this being the degradation speed, different approaches are being discussed. For one, it is attempted to slow the degradation on the part of the implant material by developing appropriate alloys. In addition, coatings are to bring about a temporary inhibition of the degradation. While the existing approaches are promising, none of them has so far led to a commercially available product. Regardless of the efforts made so far, there is rather a continuing need for solutions that make it possible to at least temporarily reduce the corrosion of magnesium alloys in vivo, while optimizing the mechanical properties thereof at the same time.

In light of this prior art, the techniques disclosed herein may allow for provision of a biodegradable magnesium alloy, a method for the production thereof and a use for implants, which may allow the magnesium matrix of the implant to remain in an electrochemically stable state over the required support period with fine grain and high corrosion resistance without protective layers, and to utilize the formation of intermetallic phases, which electrochemically are more noble than the magnesium matrix, while also improving the mechanical properties, such as increasing the tensile strength and proof stress, as well as reducing the mechanical asymmetry so as to adjust the degradation rate of the implants.

These may be achieved by a magnesium alloy, a method, and the use according to the independent claims.

The characteristics listed in the dependent claims allow advantageous refinements of the magnesium alloy, method or use according to the independent claims.

5 The solution may allow for the corrosion resistance and deformability of the magnesium matrix of the implant to be assured over the support period such that the implant is able to absorb multiaxial permanent load without fracture or cracking, and to also utilize the magnesium matrix as a means for the decomposition triggered by the physiological liquids.

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This is achieved by the magnesium alloy comprising:

3 to 7% by weight Zn, 0.001 to 0.5% by weight Ca, the remainder being magnesium containing total impurities, which promote electrochemical potential differences and/or the formation of intermetallic phases, in a total amount of no more than 0.0048% by weight, preferably no more than 0.0045% by weight and still more preferably no more than 0.0036% by weight, wherein the total impurity contains:

- individual impurities selected from the group of Fe, Si, Mn, Co, Ni, Cu, Al, Zr and P in an amount of not more than 0.0038% of weight; and
- alloying elements selected from the group of the rare earths having the ordinal numbers 21, 39, 57 to 71 and 89 to 103 in an amount of no more than 0.001% by weight.

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The magnesium alloy as disclosed herein may have extraordinarily high corrosion resistance, which may be achieved by drastically reducing the content of individual impurities and the combinations thereof in the magnesium matrix, and by also adding precipitation and solid solution hardenable elements, which may be present in completely solid solution. The microstructure that is obtained may have no electrochemical potential differences between the individual matrix phases after the forming and heat treatment processes, and therefore these differences cannot expedite the corrosion in physiological media.

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The applicant surprisingly found that the alloy matrix, having a content of 3 to 5% by weight Zn and 0.2 to 0.4% by weight Ca, has a more noble electrochemical potential than the intermetallic phase $\text{Ca}_2\text{Mg}_6\text{Zn}_3$, while the MgZn phase, which can be precipitated by heat treatment, is considerably more noble as compared to the alloy matrix, whereby the
5 MgZn phase can be precipitated from the lattice of the alloy matrix by means of targeted heat treatment and can act as cathodes for the alloy matrix. This provides the option of deliberately influencing the degradation rate of the alloy matrix.

Another surprising result is that, despite freedom of Zr, or Zr contents that are
10 considerably lower than those stated in the prior art, a grain refining effect can be achieved which is attributed to the intermetallic phase $\text{Ca}_2\text{Mg}_6\text{Zn}_3$, which blocks the movements of the grain boundaries, limits the grain size during recrystallization and thus prevents undesirable grain growth, while also increasing the proof stress and tensile strength values.

15 Within the meaning of the aforementioned mechanical properties, a Zr content of <0.0003% by weight, and preferably <0.0001% by weight, is advantageous for the magnesium alloy disclosed herein.

The previously known tolerance limits for the individual impurities within the total
20 impurity do not take into account that wrought magnesium alloys often times are subjected to a thermomechanical treatment, and more particularly to an extended annealing process, which creates the near-equilibrium structures. The metallic elements bond by way of diffusion and form what are known as intermetallic phases, which have a different electrochemical potential, notably a considerably higher potential, than the
25 magnesium matrix, and therefore these intermetallic phases act as cathodes and can trigger galvanic corrosion processes.

The applicant found that a formation of such intermetallic phases can be reliably prevented when complying with the following tolerance limits of the individual
30 impurities within the total impurity:

Fe, Si, Mn, Co, Ni, Cu each with <0.0005% by weight; Zr, Al each with <0.0003% by weight; and P <0.0002% by weight.

It is noted that in this embodiment the total impurity additionally may contain individual impurities selected from the group of rare earths having the ordinal numbers 21, 57 to 71 and 89 to 103 in total <0.001% by weight, preferably no more than 0.0005% by weight and still more preferably no more than 0.0003% by weight..

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Preferably the corrosion-staple alloy matrix contains total impurities in a total amount of no more than 0.0036 % by weight, containing individual impurities selected from the group of Fe, Si, Mn, Co, Ni, Al, Zr, and P of in total no more than 0.0026% by weight, wherein the total amount can be achieved when complying with the following tolerance

10 limits of individual impurities in % by weight:

Fe, Si, Mn each with <0.0005; Co, Ni, Cu each with <0.0002; Zr, <0.0003; and P, Al each with <0.0001.

Additionally the total impurities may contain the preferred amount of rare earths having the ordinal numbers 21, 57 to 71 and 89 to 103 in total <0.001.

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In particular the preferred corrosion-staple alloy matrix may contain total impurities in a total amount of no more than 0.00215 % by weight, containing individual impurities selected from the group of Fe, Si, Mn, Co, Ni, Al, Zr, and P of in total no more than 0.00115% by weight, wherein the total amount can be achieved when complying with the

20 following tolerance limits of individual impurities in % by weight:

Fe, Si, Mn each with <0.0002; Co, Ni, Cu, Zr, P each with <0.0001; and Al <0.00005.

Additionally the total impurities may contain the preferred amount of rare earths having the ordinal numbers 21, 57 to 71 and 89 to 103 in total <0.001.

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When the individual impurities are combined, the formation of intermetallic phases, which are more noble than the alloy matrix, may be suppressed if the sum of individual impurities consisting of Fe, Si, Mn, Co, Ni, Cu and Al is no more than 0.0033% by weight, preferably no more than 0.0022% by weight, still more preferably no more than

30 0.00095% by weight, the content of Al is no more than 0.0003% by weight, preferably no more than 0.0001% by weight, particularly preferably no more than 0.00005% by weight

and the content of Zr is preferably no more than 0.0003% by weight, preferably no more than 0.0001% by weight.

There are different mechanisms of action by which the aforementioned individual impurities impair the corrosion resistance of the material.

5 If small Fe particles form in the alloy because of a high Fe content, these particles will act as cathodes for a corrosive attack; the same applies to Ni and Cu.

Moreover, especially Fe and Ni with Zr, but also Fe, Ni and Cu with Zr, can precipitate as intermetallic particles in the melt; these will also act as very effective cathodes for corrosion of the matrix.

10 Intermetallic particles having a very high potential difference as compared to the matrix and a very high formation tendency may include the phases of Fe and Si and Fe, Mn and Si, which is why contaminations containing these elements must be minimized.

The content of P should be minimized to the extent possible, because Mg phosphides form even when minute amounts of P are present and dramatically impair the mechanical
15 properties of the structure.

Such low concentrations thus ensure that the magnesium matrix no longer contains any intermetallic phases, which have a more positive electrochemical potential as compared to the matrix.

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In the magnesium alloy disclosed herein, the sum of the contents of elements of the rare earths and scandium (ordinal numbers 21, 57 to 71 and 89 to 103) may be less than 0.001% by weight, preferably less than 0.0005% by weight and still more preferably less than 0.0003.

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These additions make it possible to increase the tensile strength of the magnesium matrix and raise the electrochemical potential of the matrix, whereby a corrosion-reducing action, notably with respect to physiological media, may develop. The precipitations preferably have a size of no more than 5 μm , and preferably of no more than 1 μm , and
30 are located on the grain boundaries and in the grain interior, whereby the movement of grain boundaries during thermal treatment as well as dislocations during deformation may be impaired and the strength of the magnesium alloy may be increased.

The magnesium alloy disclosed herein may achieve a tensile strength of >275 MPa, and preferably >300 MPa, a yield point of >200 MPa, and preferably >225 MPa, and a yield ratio of <0.8, and preferably <0.75, wherein the difference between the tensile strength and yield point may be >50 MPa, and preferably >100 MPa, and the mechanical asymmetry may be <1.25.

These significantly improved mechanical properties of the magnesium alloys may allow the implants, for example cardiovascular stents, to be able to withstand the multiaxial permanent load in the implanted state over the entire support period, despite onsetting degradation of the magnesium matrix due to corrosion.

So as to minimize the mechanical asymmetry, it is particularly important for the magnesium alloy to have a particularly fine microstructure having a grain size of no more than 5 μm .

In another exemplary technique, there is provided a method for producing a magnesium alloy having improved mechanical and electrochemical properties. The method comprises the following steps:

- a) generating a high-purity magnesium by way of vacuum distillation;
- b) generating a billet of the alloy by synthesis of the magnesium according to step a) with high-purity Zn and Ca in a composition of 3 to 7% by weight Zn, 0.001 to 0.5% by weight Ca, the remainder being magnesium containing total impurities, which promote electrochemical potential differences and/or the formation of intermetallic phases, in an amount of no more than 0.0048 by weight, preferably no more than 0.0045 % by weight and still more preferably no more than 0.0036% by weight, wherein the total impurity contains individual impurities selected from the group of Fe, Si, Mn, Co, Ni, Cu, Al, Zr and P in an amount of not more than 0.0038% of weight; and wherein the alloying elements are selected from the group of the rare earths having the ordinal numbers 21, 39, 57 to 71 and 89 to 103 in a total amount of no more than 0.001% by weight;
- c) homogenizing the alloy by annealing at a temperature between 350°C and 450°C with a holding period of 4 to 40 hours;

- c) at least single forming of the homogenized alloy in the temperature range between 250°C and 350°C; and
- d) optionally heat treating the formed alloy in the temperature range between 100°C and 300°C with a holding period of 1 minute to 3 hours.

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A content of preferably 3 to 5% by weight Zn and a content of 0.2 to 0.4% by weight Ca assure that volume contents of up to 2% of the intermetallic phase and the precipitable MgZn phase, respectively, are created in the matrix lattice. The electrochemical potentials of the two phases differ considerably, the MgZn phase being more noble than the matrix and the $\text{Ca}_2\text{Mg}_6\text{Zn}_3$ phase being less noble than the matrix. By means of the heat treatment, the MgZn phase can be prompted to precipitate to the desired degree in a regime preselected by the temperature and holding period, whereby the degradation rate of the alloy matrix can be adjusted.

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In addition to the corrosion-inhibiting action, the intermetallic $\text{Ca}_2\text{Mg}_6\text{Zn}_3$ phase also has the surprising effect of grain refining created by the forming process, which results in a considerable increase in tensile strength and proof stress. This allows Zr to be dispensed with as an alloying element and the temperatures for the recrystallization to be lowered.

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Vacuum distillation is preferably used to produce a starting material for a high-purity magnesium-zinc-calcium alloy having the required threshold values.

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The sum of impurities and the content of addition elements triggering the precipitation and solid solution hardening and raising the matrix potential can be selectively adjusted and in % by weight are:

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- a) for the individual impurities:

Fe, Si, Mn, Co, Ni, Cu each with <0.0005% by weight;

Al, Zr each with <0.0003% by weight;

and P <0.0002% by weight.

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It is noted that in this embodiment the total impurity additionally contains individual impurities selected from the group of rare earths

having the ordinal numbers 21, 57 to 71 and 89 to 103 in total <0.001% by weight.

5 aa) for the individual impurities in a preferred amount of impurities of no more than 0.0026% by weight,:

Fe, Si, Mn each with <0.0005;

Co, Ni, Cu each with <0.0002;

Zr, <0.0003; and

Al, P each with <0.0001.

10 In this preferred embodiment the amount of total impurities which is the sum of the individual impurities and the impurities selected from the group of rare earths having the ordinal numbers 21, 57 to 71 and 89 to 103 is in total <0.0036% by weight.

15 ab) for the individual impurities in a particularly preferred amount of impurities of no more than 0.00115% by weight:

Fe, Si, Mn each with <0.0002;

Co, Ni, Cu, Zr, P each with <0.0001; and

Al <0.0001.

20 In this preferred embodiment the amount of total impurities which is the sum of the individual impurities and the impurities selected from the group of rare earths having the ordinal numbers 21, 57 to 71 and 89 to 103 is in total <0.00215% by weight.

25 b) for the combination of individual impurities in total:

Fe, Si, Mn, Co, Ni, Cu and Al no more than 0.0033% by weight, preferably no more than 0.0022% by weight, still more preferably 0.00095% by weight, the content of Al no more than 0.0003 preferably no more than 0.0001% by weight, particularly preferably no more than 0.00005% by weight, and the content of Zr is no more than 0.0003% by weight, preferably no more than 0.0001% by weight;

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- c) for the addition elements which are also contained in the total impurities:
rare earths in a total amount of no more than 0.001% by weight, preferably no more than 0.0005% by weight.

5 It is particularly advantageous that the method disclosed herein may only require a small number of forming steps. Extrusion, equal channel angular extrusion and/or multiple forging can thus preferably be employed, which assure that a substantially homogeneous fine grain of <15 μm is achieved.

Because of the heat treatment, MgZn precipitations having a grain size of 1 nm to 50 nm
10 in a fine-grained structure having a grain size of preferably <7.5 μm dispersely distributed at the grain boundaries and in the interior of the grains, whereby the tensile strength of the alloy reaches values which at >275 MPa, and preferably >300 MPa, may be considerably higher than the prior art.

15 In another exemplary technique, there is provided a use of the magnesium alloy produced according to the method disclosed herein, which may have the aforescribed advantageous composition and structure, in medical technology, notably for the production of implants, for example endovascular implants such as stents, for fastening and temporarily fixing tissue implants and tissue transplantations, orthopedic and dental
20 implants, and neuroimplants.

All implants in the sense of this patent application are in the Cardiovascular field, osteosynthesis field or other areas.

25 Cardiovascular field in the sense of this application means

- the field of diagnostic, prevention and treatment of all diseases of the cardiovascular system, i.e. heart and blood vessel system,
- by mean of active and non-active implants used to support vessels, and vein systems
- including coronary, cerebral and peripheral vascular implants like stents, valves,
30 closure devices, occluders, clips, coils, staples, implantable regional drug delivery devices,
- implantable electrostimulators (like pacemakers and defibrillators), implantable monitoring devices, implantable electrodes,

- system for fastening and temporarily fixing tissue implants and tissue transplantations
- field also includes any type of stent as mechanical fix or temporary scaffold to support hollow organs (or bodies?) including bones, intervertebral disks

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Osteosynthesis in the sense of this application means

- the field of treatment of fractured bones for internal fixation and stabilization by mechanical devices such as metal plates, pins, rods, wires, screws, clips, nails, staples excluding stent technology

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Examples of areas out of the osteosynthesis field or the cardiovascular field are:

- Devices for the treatment of diseases of the sinews, joints, muscles, cartilages,
- oral (including dental) and maxillo facial implants (excl. osteosynthesis means),
- esthetic implants,
- supporting tools out of the body, (examples?)
- tissue engineering,
- soft tissue implants,
- devices for wound care,
- suture material and clamps,
- neurosurgery
- local drug delivery (excl. cardiovascular, i.e. lever)
- renal applications

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Example

A magnesium alloy is to be generated which is composed of 5% by weight Zn and 0.15% by weight Ca, the remainder being Mg, and contains the following individual impurities in % by weight:

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Fe: <0.0005; Si: <0.0005; Mn: <0.0005; Co: <0.0002; Ni: <0.0001; Cu <0.0002, wherein the sum of individual impurities consisting of Fe, Si, Mn, Co, Ni, Cu and Al should be no more than 0.0015% by weight, the content of Al < 0.001% by weight and that of Zr

<0,0003% by weight, and the content of rare earths having the ordinal numbers 21, 39, 57 to 71 and 89 to 103 in total should be less than 0.001% by weight, which leads to an alloy with total impurities of in total no more than 0.0035% by weight.

- 5 This alloy, produced using magnesium vacuum distillation, is subjected to homogenizing annealing at a temperature of 350°C for a duration of 12 hours, and subsequently to multiple extrusion processes at a temperature of 300°C, so as to produce a precision tube for a cardiovascular stent.

The subsequent heat treatment was carried out at a temperature of 250°C with a holding
10 period of 0.1 hour. The grain size was <7.5 µm.

The magnesium alloy reached a tensile strength of 310 to 320 MPa and proof stress of <250 MPa [sic]. The yield ratio was 0.7 and the mechanical asymmetry was 1.2.

In artificial body fluid, the $\text{Ca}_2\text{Mg}_6\text{Zn}_3$ phase degraded more quickly than the matrix and is therefore less noble than the matrix. This means that these intermetallic particles
15 cannot act as cathodes for the alloy matrix. The MgZn phase expedited the degradation in artificial body fluid and was therefore more noble electrochemically than the alloy matrix, whereby it is able to induce corrosion.

Because of the subsequent heat treatment, it is thus possible to precipitate the MgZn phase from the alloy matrix, rendering the alloy matrix less noble.

- 20 The subsequent degradation rate under physiological usage conditions can thus be adjusted by means of the heat treatment.

PATENT CLAIMS

1. A magnesium alloy having improved mechanical and electrochemical properties, comprising: 3 to 7% by weight Zn, 0.001 to 0.5% by weight Ca, the remainder
5 being magnesium containing total impurities, which promote electrochemical potential differences and/or the formation of intermetallic phases, in a total amount of no more than 0.0048% by weight wherein the total impurity contains:
 - individual impurities selected from the group of Fe, Si, Mn, Co, Ni, Cu, Al, Zr and P in an amount of not more than 0.0038% of weight; and
 - 10 - alloying elements selected from the group of the rare earths having the ordinal numbers 21, 39, 57 to 71 and 89 to 103 in an amount of no more than 0.001% by weight.
2. The magnesium alloy according to claim 1, characterized in that the total impurities
15 have an amount in total of no more than 0.0045% by weight, preferably no more than 0.0036% by weight.
3. The magnesium alloy according to claim 1, characterized in that the content of Zn
20 is preferably 3 to 5% by weight and that of Ca is preferably at least 0.2 to 0.4% by weight, wherein the alloy contains an intermetallic $\text{Ca}_2\text{Mg}_6\text{Zn}_3$ phase and optionally a precipitable MgZn phase, each in a volume content of close to 0 to 2%.
4. The magnesium alloy according to claim 1 or claim 2, characterized in that the
25 individual impurities in the sum of total impurities amount to the following in % by weight: Fe <0.0005; Si <0.0005; Mn <0.0005; Co <0.0005; Ni <0.0005; Cu <0.0005; Al <0.0003; Zr <0.0003; P <0.0002.
5. The magnesium alloy according to claim 1 or claim 2, characterized in that when
30 the individual impurity elements Fe, Si, Mn, Co, Ni, Cu and Al are combined, the sum of these individual impurities is no more than 0.0033% by weight, preferably no more than 0.0022% by weight particularly preferably no more than 0.00095% by weight, the content of Al is no more than 0.0003% by weight, preferably no more

than 0.0001% by weight particularly preferably no more than 0.00005% by weight and the content of Zr is preferably no more than 0.0001% by weight

- 5 6. The magnesium alloy according to claim 1 or claim 2, characterized in that the individual elements from the group of the rare earths amount to no more than 0.001% by weight, preferably 0.0005% by weight, and still more preferably 0.0003% by weight in the total sum.
- 10 7. The magnesium alloy according to any of the preceding claims 1 to 6, characterized in that the alloy has a fine-grained microstructure having a grain size of less than 7.5 μm , and preferably $<5 \mu\text{m}$, and particularly preferably $<2.5 \mu\text{m}$, without electrochemical potential differences between the individual matrix phases.
- 15 8. The magnesium alloy according to any of claims 1 to 7, characterized in that the intermetallic $\text{Ca}_2\text{Mg}_6\text{Zn}_3$ phase is equally as noble as, or less noble, and the precipitable MgZn phase is more noble, than the matrix phases.
- 20 9. The magnesium alloy according to any of the preceding claims 1 to 8, characterized in that the precipitations have a size of $<5 \mu\text{m}$, and preferably $<1 \mu\text{m}$, and are dispersely distributed at the grain boundaries or in the grain interior.
- 25 10. The magnesium alloy according to any of the preceding claims 1 to 9, characterized in that the alloy has a tensile strength of $>275 \text{ MPa}$, and preferably $>300 \text{ MPa}$, a yield point of $>200 \text{ MPa}$, and preferably $>225 \text{ MPa}$, and a yield ratio of <0.8 , and preferably <0.75 , wherein the difference between the tensile strength and yield point is $>50 \text{ MPa}$, and preferably $>100 \text{ MPa}$, and the mechanical asymmetry is <1.25 .
- 30 11. A method for producing a magnesium alloy having improved mechanical and electrochemical properties, comprising the following steps:
 - a) generating a high-purity magnesium by way of vacuum distillation;

- b) generating a billet of the alloy by synthesis of the magnesium according to step a) with high-purity Zn and Ca in a composition of 3 to 7% by weight Zn, 0.001 to 0.5% by weight Ca, the remainder being magnesium containing total impurities, which promote electrochemical potential differences and/or the formation of intermetallic phases, in a amount of no more than 0.0048% by weight wherein the total impurity contains individual impurities selected from the group of Fe, Si, Mn, Co, Ni, Cu, Al, Zr and P in an amount of not more than 0.0038% of weight; and wherein the alloying elements are selected from the group of the rare earths having the ordinal numbers 21, 39, 57 to 71 and 89 to 103 in a total amount of no more than 0.001% by weight;
- c) homogenizing the alloy by annealing at a temperature between 350°C and 450°C with a holding period of 4 to 40 hours;
- c) at least single forming of the homogenized alloy in the temperature range between 250°C and 350°C; and
- d) optionally heat treating the formed alloy in the temperature range between 100°C and 300°C with a holding period of 1 minute to 3 hours.

12. The method according to claim 11, characterized in that the content of Zn is preferably at least 3 to 5% by weight, and the content of Ca is 0.2 to 0.4% by weight.

13. The method according to claim 11 or claim 12, characterized in that the MgZn phase, which is more noble than the alloy matrix, is precipitated from the alloy matrix by the heat treatment and the potential difference that exists between the alloy matrix and the MgZn precipitations is used to adjust the degradation rate of the alloy matrix.

14. The method according to claim 11 or claim 12, characterized in that the grain refining during forming is generated by the intermetallic $\text{Ca}_2\text{Mg}_6\text{Zn}_3$ instead of Zr.

15. The method according to claim 11, characterized in that total impurities have an amount in total of no more than 0.0045% by weight, preferably no more than 0.0036% by weight.

5 16. The method according to claim 11 or claim 15, characterized in that the the individual impurities in the sum of total impurities amount to the following in % by weight:

Fe <0.0005; Si <0.0005; Mn <0.0005; Co <0.0005; Ni <0.0005; Cu <0.0005; Al <0.0003; Zr <0.0003; P <0.0002.

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17. The method according to claim 11, claim 15 or claim 16, characterized in that when the individual impurity elements Fe, Si, Mn, Co, Ni, Cu and Al are combined, the sum of these impurities is no more than 0.0033, and preferably no more than 0.0022% by weight, , still more preferably 0.00095% by weight, the content of Al is no more than 0.0003% by weight, and preferably no more than 0.0001% by weight, still more preferably 0.00005% by weight, and the content of Zr is preferably no more than 0.0001% by weight

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18. The method according to claim 11 or claim 15, characterized in that the individual elements from the rare earths amount to no more than 0.001, and preferably no more than 0.0005% by weight in the total sum.

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19. The method according to claim 11, characterized in that the forming process is an extrusion, equal channel angular extrusion (EACE) and/or multiple forging process.

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20. The method according to claim 11 or claim 13, characterized in that MgZn precipitations after the heat treatment are present in a size of 1 nm to 50 nm in a fine-grained structure having a grain size of preferably <7.5 μm dispersely distributed at the grain boundaries and in the grain.

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21. Use of a magnesium alloy according to any one of the claims 1 to 10 for producing a biodegradable implant.

22. Use of a magnesium alloy according to claim 21, characterized in that the implant is elected from the group of endovascular implants such as stents, implants for fastening and temporarily fixing tissue implants and tissue transplantations, orthopedic and dental implants, and neuroimplants.
23. Use of a magnesium alloy produced by the method according to any one of the claims 11 to claim 20 for producing a biodegradable implant.
24. Use of a magnesium alloy according to claim 23, characterized in that the implant is elected from the group of endovascular implants such as stents, implants for fastening and temporarily fixing tissue implants and tissue transplantations, orthopedic and dental implants, and neuroimplants.