

DESCRIPTION

FINE BUBBLE ELECTROLYZED WATER GENERATING APPARATUS AND METHOD FOR GENERATING FINE BUBBLE ELECTROLYZED WATER

Technical Field

5 [0001]

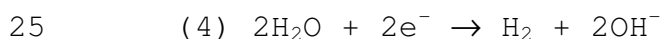
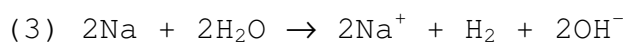
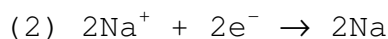
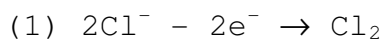
The present invention relates to a fine bubble electrolyzed water generating apparatus and a method for generating fine bubble electrolyzed water, and more particularly to a fine bubble electrolyzed water generating apparatus and a method for generating
10 fine bubble electrolyzed water, wherein electrolyzed water and gas that are generated by electrolysis are used to provide alkaline electrolyzed nanobubble water most suitable for use in the overall cleaning field, particularly, in the cleaning of industrial parts and commercial cleaning, and acidic electrolyzed nanobubble water
15 suitable for use in commercial cleaning.

Background Art

[0002]

Conventional electrolysis of brine includes the following reaction (1) on the anode side (anode chamber) and the following
20 reactions (2), (3), and (4) on the cathode side (cathode chamber).

[0003]



[0004]

Alkaline reducing cathode electrolyzed water is accordingly obtained in the cathode chamber.

[0005]

5 Hydrogen gas is produced along with this cathode electrolyzed water. However, hydrogen gas, which has low solubility, has been released into the air without being dissolved in cathode electrolyzed water. Acidic anode electrolyzed water and chlorine gas are also obtained in the anode chamber.

10 [0006]

Patent Literature 1 has been known as a technique of effectively utilizing chlorine gas produced along with anode electrolyzed water.

[0007]

Patent Literature 1 has disclosed that an aqueous solution
15 of an electrolyte, a mixture of caustic soda and table salt, was electrolyzed and the obtained electrolyzed solution was used as water for washing and cleaning.

Citation List

Patent Literature

20 [0008]

Patent Literature 1: JP2003-251353

Summary of Invention

Technical Problem

[0009]

25 In the method of Patent Literature 1, however, chlorine gas

included in the obtained electrolyzed water easily volatilized, causing low storage stability. When the electrolyzed water was used as water for washing and cleaning, satisfactory cleaning efficiency was not obtained accordingly.

5 [0010]

In order to solve the problem, the present invention provides electrolyzed water that achieves satisfactory cleaning efficiency.

[0011]

The present inventors have earnestly studied and as a result,
10 have found that the above problem can be solved by forming nanobubbles of gas produced by electrolysis of an aqueous solution of sodium chloride, thereby completing the present invention.

Solution to Problem

[0012]

15 The present invention relates to a fine bubble electrolyzed water generating apparatus and a method for generating fine bubble electrolyzed water, the apparatus comprises:

an electrolytic apparatus having a two-chamber structure including an anode chamber with an anode electrode, a cathode chamber
20 with a cathode electrode, and a diaphragm provided between the anode chamber and the cathode chamber, or an electrolytic apparatus having a three-chamber structure including an anode chamber with an anode electrode, a cathode chamber with a cathode electrode, an intermediate chamber provided between these chambers, a diaphragm
25 provided between the anode chamber and the intermediate chamber,

and a diaphragm provided between the cathode chamber and the intermediate chamber;

an acidic electrolyzed water storage tank for storing acidic electrolyzed water and an alkaline electrolyzed water storage tank 5 for storing alkaline electrolyzed water, both provided adjacent to the electrolytic apparatus;

pipes through which these electrolyzed water storage tanks are in communication with the anode chamber and the cathode chamber, respectively; and

10 nanobubble generators for generating nanobubbles with gas and liquid, which are in communication with these electrolyzed water storage tanks, respectively,

wherein these nanobubble generators generate chlorine nanobubble electrolyzed water with chlorine gas and acidic 15 electrolyzed water produced in the anode chamber, and generate hydrogen nanobubble electrolyzed water with hydrogen gas and alkaline electrolyzed water produced in the cathode chamber.

Advantageous Effects of Invention

[0013]

20 According to the present invention, cathode electrolyzed water and excessive hydrogen gas generated at the cathode are supplied to the nanobubble generator in the apparatus. This produces hydrogen nanobubbles in the cathode electrolyzed water to increase the oxidation-reduction potential (ORP) of the cathode electrolyzed 25 water.

[0014]

In addition, anode electrolyzed water and excessive chlorine gas generated at the anode are supplied to the nanobubble generator in the apparatus. This produces chlorine nanobubbles in the anode electrolyzed water to provide significantly advantageous effects such as increase in chlorine concentration in the anode electrolyzed water.

[0015]

Furthermore, chlorine nanobubbles are produced in the anode electrolyzed water and nanobubbles are generated in the obtained anode electrolyzed water. This provides fine bubble electrolyzed water having an excellent bactericidal effect.

Brief Description of Drawings

[0016]

Fig. 1 is a schematic diagram illustrating an embodiment of the electrolytic apparatus used in the present invention.

Fig. 2 is a schematic diagram illustrating the condition of use of the fine bubble electrolyzed water generating apparatus used in the present invention.

Fig. 3 is a flow chart of the fine bubble electrolyzed water generating apparatus used in the present invention.

Fig. 4 is a graph showing the cleaning effect using the fine bubble electrolyzed water obtained in the present invention.

Fig. 5 is a graph showing the cleaning effect using the fine bubble electrolyzed water obtained in the present invention.

Fig. 6 is a graph showing the cleaning effect using the fine bubble electrolyzed water obtained in the present invention.

Fig. 7 is a graph showing the cleaning effect using the fine bubble electrolyzed water obtained in the present invention.

5 Fig. 8 is a graph showing the cleaning effect using the fine bubble electrolyzed water obtained in the present invention.

Fig. 9 shows photographs substituted for drawings, illustrating the cleaning effect using the fine bubble electrolyzed water obtained in the present invention, wherein Fig. 9(a) shows
10 an Boston lettuce as an analyte; Fig. 9(b) shows the number of bacteria for an untreated analyte in the case of 100 fold dilution; Fig. 9(c) shows the number of bacteria for an analyte treated with tap water in the case of 100 fold dilution; Fig. 9(d) shows the number of bacteria for an analyte treated with acidic electrolyzed water
15 in the case of 10 fold dilution; and Fig. 9(e) shows the number of bacteria for an analyte treated with nanobubble (NB) acidic electrolyzed water in the case of 10 fold dilution.

Description of Embodiments

[0017]

20 The present invention will be described below in detail. The present invention is not limited to the following description and can be appropriately modified without departing from the spirit of the present invention.

[0018]

25 An embodiment of the fine bubble electrolyzed water generating

apparatus of the present invention will be described with reference to Figs. 1, 2, and 3. An electrolytic apparatus (X) constituting part of the apparatus has a three-chamber structure, and the three-chamber structure includes an anode chamber (1) with an anode electrode (6), a cathode chamber (2) with a cathode electrode (7), and an intermediate chamber (3) provided between the anode chamber (1) and the cathode chamber (2). The three-chamber structure further includes a diaphragm (4) provided between the anode chamber (1) and the intermediate chamber (3), and a diaphragm (5) provided between the cathode chamber (2) and the intermediate chamber (3).

[0019]

In the present invention, the electrolytic apparatus (X) is provided with an acidic electrolyzed water storage tank (10) for storing acidic electrolyzed water and a cathode electrolyzed water storage tank (11) for storing alkaline electrolyzed water, both of which are provided close to the anode chamber (1) and the cathode chamber (2) of the apparatus, respectively, and further provided with nanobubble generators (14) and (14) on the anode chamber (1) side and the cathode chamber (2) side, respectively. In addition, pipes (8), (9), (12), (13), (15), and (16) for connecting the chambers (1) and (2), the storage tanks (10) and (11), and the nanobubble generators (14) and (14), and pipes (17) and (18) for discharging from the storage tanks (10) and (11) are provided. Two brine tanks are also provided to supply brine to the anode chamber (1), the cathode chamber (2), and the intermediate chamber (3).

[0020]

Accordingly, the acidic electrolyzed water generated in the anode chamber (1) of the electrolytic apparatus (X) is in communication with the acidic electrolyzed water storage tank (10) 5 through the acidic electrolyzed water connection pipe (8).

[0021]

The acidic electrolyzed water storage tank (10) is in communication with the nanobubble generator (14) through the acidic electrolyzed water conduit (12).

10 [0022]

Moreover, the nanobubble generator (14) is in communication with the acidic electrolyzed water storage tank (10) through the acidic nanobubble electrolyzed water pipe (15).

[0023]

15 The acidic electrolyzed water storage tank (10) is configured so that acidic nanobubble electrolyzed water is discharged through the acidic nanobubble electrolyzed water discharge pipe (17).

[0024]

The alkaline electrolyzed water generated in the cathode 20 chamber (2) of the electrolytic apparatus (X) is in communication with the alkaline electrolyzed water storage tank (11) through the alkaline electrolyzed water connection pipe (9).

[0025]

The alkaline electrolyzed water storage tank (11) is in 25 communication with the nanobubble generator (14') through the

alkaline electrolyzed water conduit (13).

[0026]

In addition, the nanobubble generator (14') is in communication with the alkaline electrolyzed water storage tank (11) through the alkaline nanobubble electrolyzed water pipe (16).

[0027]

The alkaline electrolyzed water storage tank (11) is configured so that alkaline nanobubble electrolyzed water is discharged through the alkaline nanobubble electrolyzed water discharge pipe (18).

10 [0028]

According to the present invention, chlorine gas is generated at the anode electrode, and nanobubbles of the chlorine gas are formed and the acidic electrolyzed water including the nanobubbles of the chlorine gas is supplied to the acidic electrolyzed water tank. Furthermore, according to the present invention, hydrogen gas is generated at the cathode electrode, and nanobubbles of the hydrogen gas are formed and the alkaline electrolyzed water including the nanobubbles of the hydrogen gas is supplied to the alkaline electrolyzed water storage tank.

20 [0029]

As described above, the electrolytic apparatus used in the fine bubble electrolyzed water generating apparatus has the three-chamber structure including the anode chamber, the intermediate chamber, and the cathode chamber. This electrolytic apparatus may have a two-chamber structure including an anode chamber

with an anode electrode, a cathode chamber with a cathode electrode, and a diaphragm provided between these chambers. Alternatively, the electrolytic apparatus may have a one-chamber structure including an anode electrode and a cathode electrode without a
5 diaphragm between these electrodes.

[0030]

Electrolyzed water generated in the three-chamber type electrolytic cell includes few undissociated electrolytes (for example, salts). The electrolyzed water accordingly has less
10 influence on the operating environment so that the three-chamber type electrolytic cell is more preferable than the two-chamber type electrolytic cell.

[0031]

Next, an embodiment of the method for generating fine bubble
15 electrolyzed water of the present invention will be described with reference to Figs. 1 and 2.

[0032]

First, an aqueous solution of sodium chloride is electrolyzed. Chlorine gas and acidic electrolyzed water generated at the anode
20 electrode in the electrolytic apparatus are processed by the nanobubble generator. This produces chlorine nanobubble electrolyzed water. Hydrogen gas and alkaline electrolyzed water generated at the cathode electrode in the electrolytic apparatus are processed by the nanobubble generator. This produces hydrogen
25 nanobubble electrolyzed water.

[0033]

In the aqueous solution of sodium chloride used as a raw material of the fine bubble electrolyzed water of the present invention, the concentration of sodium chloride is preferably in the range of 0.05 to 0.2 wt% when the two-chamber type electrolytic cell is used. The aqueous solution of sodium chloride of less than 0.05 wt% is not preferred because the electrolytic efficiency significantly decreases. The aqueous solution of sodium chloride of more than 0.2 wt% is not preferred because undissociated electrolytes remain in the electrolyzed water.

[0034]

In the three-chamber type electrolytic cell, the aqueous solution of sodium chloride in the intermediate chamber is separated from cathode water and anode water through the diaphragms. Accordingly, there is no movement of sodium chloride to both electrodes. The concentration of the aqueous solution of sodium chloride can be thus controlled by circulating saturated brine, and does not require strict control at a level of several tenths of one percent.

[0035]

Conventional electrolyzed water and bactericides have hardly reduced the number of residual bacteria or provided the cleaning effect. The fine bubble electrolyzed water of the present invention, however, can reduce the number of residual bacteria and provide the cleaning effect. The fine bubble electrolyzed water of the

present invention can be used as, for example, rinsing and sterilizing water in food-processing plants for cut vegetables or others and in cleaning business or other businesses.

EXAMPLES

5 [0036]

Next, verification will be made on the cleaning effect of the fine bubble electrolyzed water (hereinafter, referred to as nano-electrolyzed water) generated by this fine bubble electrolyzed water generating apparatus and this method for generating fine bubble
10 electrolyzed water.

[0037]

Example 1:

[Content of Test]

Saturated brine was electrolyzed using an electrolytic cell
15 having a three-chamber structure under the condition where acidic and alkaline electrolyzed waters were generated at 4 L/min, and the following samples were prepared.

[0038]

- Nano-electrolyzed water 1: a general term for electrolyzed
20 water obtained by forming nanobubbles of gas (hydrogen, chlorine) produced during the generation of electrolyzed water and adding the nanobubbles to the electrolyzed water (alkaline electrolyzed water + hydrogen nanobubbles, acidic electrolyzed water + chlorine nanobubbles)

25 - Nano-electrolyzed water 2: a general term for electrolyzed

water obtained by forming nanobubbles of atmosphere and adding the nanobubbles to electrolyzed water (alkaline electrolyzed water + atmosphere nanobubbles, acidic electrolyzed water + atmosphere nanobubbles)

5 - Nano-alkaline electrolyzed water 1: alkaline electrolyzed water obtained by forming nanobubbles of hydrogen gas produced during the generation of alkaline electrolyzed water and adding the nanobubbles to the alkaline electrolyzed water

 - Nano-alkaline electrolyzed water 2: alkaline electrolyzed
10 water obtained by forming nanobubbles of atmosphere and adding the nanobubbles to alkaline electrolyzed water

 - Nano-acidic electrolyzed water 1: acidic electrolyzed water obtained by forming nanobubbles of chlorine gas produced during the generation of acidic electrolyzed water and adding the
15 nanobubbles to the acidic electrolyzed water

 - Nano-acidic electrolyzed water 2: acidic electrolyzed water obtained by forming nanobubbles of atmosphere and adding the nanobubbles to acidic electrolyzed water

 - Electrolyzed water: a general term for water obtained in
20 the anode and cathode chambers by circulating saturated brine in the intermediate chamber of the three-chamber type electrolytic cell for electrolysis

[0039]

Artificially soiled fabrics (produced by EMPA Co.) were washed
25 with the nano-electrolyzed waters to calculate the cleaning

efficiency.

[0040]

As comparative objects, the same tests were performed using tap water, a detergent, and electrolyzed water.

5 The results were shown in the graphs of Figs. 3 to 6.

[0041]

The nano-electrolyzed water 1 (in each table and drawing, expressed by an encircled numeral 1) was prepared by a method in which hydrogen gas produced during the generation of alkaline
10 electrolyzed water was returned to the alkaline electrolyzed water again (at 0.2 L/min for 15 minutes). After washing the fabric with the nano-electrolyzed water 1, the used alkaline electrolyzed water was removed and the fabric was further washed with acidic electrolyzed water (produced chlorine gas was returned thereto (at 0.2 L/min)
15 for 15 minutes).

[0042]

Subsequently, the fabric was rinsed with tap water. The nano-electrolyzed water 2 (in each table and drawing, expressed by an encircled numeral 2) was alkaline electrolyzed water or acidic
20 electrolyzed water prepared by a method of blowing air (at 0.2 L/min for 15 minutes) instead of gases produced in respective electrolyzed waters. Using this, the fabric was washed with the alkaline electrolyzed water and the acidic electrolyzed water in this order and finally rinsed with tap water.

25 [0043]

[Test Method]

Artificially soiled fabrics (five kinds including an unsoiled fabric, carbon black/mineral oil, blood, cocoa, and red wine) having a size of 15 cm square were sewed on towels (64 cm wide × 27 cm long) 5 respectively and washed according to the following steps.

[0044]

In order to reproduce friction between laundry pieces in a tub, the fabric was washed together with nine bath towels (128 cm wide × 60 cm long).

10 [0045]

A. Order of Washing Process with Tap Water and Detergent (1. → 4.)

(Detergent Used) synthetic laundry detergent, Attack bio EX (produced by Kao Corporation)

15 1. Washing (water temperature: 8 °C) for 15 minutes

[Condition] water level: low, add 60 g of detergent during washing with detergent

Temporary dehydration for 1 minute

2. Rinsing (water temperature: 8 °C) for 10 minutes

20 [Condition] water level: low

Temporary dehydration for 1 minute

3. Rinsing 2 (water temperature: 8 °C) for 15 minutes

[Condition] water level: low

4. Dehydration for 15 minutes

25 [0046]

B. Order of Washing Process with Electrolyzed Water (Setting for Whites) (1. → 4.)

1. Washing with alkaline electrolyzed water (water temperature: 7 °C) for 15 minutes

5 [Condition] water level: low, pH: 10.80, ORP: -192, chlorine concentration: 19 ppm

Temporary dehydration for 1 minute

2. Washing with acidic electrolyzed water (water temperature: 7 °C) for 10 minutes

10 [Condition] water level: low, pH: 4.05, chlorine concentration: 19 ppm

Temporary dehydration for 1 minute

3. Rinsing 2 (water temperature: 6 °C) for 3 minutes

[Condition] water level: low

15 4. Dehydration for 5 minutes

[0047]

C. Washing Process with Nano-electrolyzed Water [common in (1) and (2), order (1. → 4.)]

20 1. Washing with nano-alkaline electrolyzed water (water temperature: 7 °C) for 15 minutes

[Condition] water level: low, pH (1): 12.17, ORP (1): -596

pH (2): 11.78, ORP (2): -202

Temporary dehydration for 1 minute

25 2. Washing with nano-acidic electrolyzed water (water

temperature: 6 °C) for 10 minutes

[Condition] water level: low, pH (1): 4.41, chlorine
concentration (1): 18 ppm

pH (2): 4.40, chlorine concentration (2): 18 ppm

5 Temporary dehydration for 1 minute

3. Rinsing (water temperature: 7 °C) for 3 minutes

[Condition] water level: low

4. Dehydration for 5 minutes

[0048]

10 [Instrument Used]

- Spectrophotometer: CM-600d (produced by Konica Minolta
Sensing, Inc.)

- Portable electrical conductivity/pH meter: WM-32EP
(produced by DKK-TOA CORPORATION)

15 - ORP combination electrode: PST-2739C (produced by DKK-TOA
CORPORATION)

- Chlorine meter: RC-2Z (produced by Kasahara Chemical
Instruments Co., Ltd.)

- Commercial washing machine 22 kg type: WN220 (produced by
20 Yamamoto Manufacturing Co., Ltd.)

[0049]

[Test Results]

The pH, ORP, and chlorine concentration of the
nano-electrolyzed water in the test are shown in Tables 1 and 2.

25 [0050]

[Table 1]

TABLE 1. VARIOUS VALUES WHEN GENERATING NANO-ALKALINE ELECTROLYZED WATER

TEST WATER	NANO-ALKALINE ELECTROLYZED WATER①		NANO-ALKALINE ELECTROLYZED WATER②	
	RAW WATER	AFTER PRODUCING NANOBUBBLES	RAW WATER	AFTER PRODUCING NANOBUBBLES
pH	12.04	12.17	11.89	11.78
ORP(mV)	-346	-596	-324	-202

[0051]

Various values in generating the nano-alkaline electrolyzed water are listed in Table 1.

[0052]

[Table 2]

TABLE 2. VARIOUS VALUES WHEN GENERATING NANO-ACIDIC ELECTROLYZED WATER

TEST WATER	NANO-ACIDIC ELECTROLYZED WATER①		NANO-ACIDIC ELECTROLYZED WATER②	
	RAW WATER	AFTER PRODUCING NANOBUBBLES	RAW WATER	AFTER PRODUCING NANOBUBBLES
pH	3.73	3.77	3.73	3.75
CHLORINE CON- CENTRATION (ppm)	34	51	51	44

* For use in the process with the nano-acidic electrolyzed water, the nano-acidic electrolyzed water was diluted with tap water to have the same chlorine concentration as that in the washing with electrolyzed water.

[0053]

Various values in generating the nano-acidic electrolyzed water are listed in Table 2.

[0054]

In Table 2, for use in the process with the nano-acidic electrolyzed water the nano-acidic electrolyzed water was diluted with tap water to have the same chlorine concentration as that in the washing with the electrolyzed water.

[0055]

After washing the fabric, the reflectance (R) at 520 nm was

measured with the spectrophotometer. The K/S value and cleaning efficiency (%) were calculated according to the following formulas.

[0056]

[Formula]

5 1. The formula is $K/S \text{ Value} = [1 - \text{Reflectance}(R)]^2 / 2 / \text{Reflectance}(R)$

2. Cleaning Efficiency (%) = $[(K/S \text{ of Fabric to be Washed}) - (K/S \text{ of Washed Soiled Fabric})] / [(K/S \text{ of Soiled Fabric}) - (K/S \text{ of Unsoiled Fabric})] \times 100$

10 [0057]

[Table 3]

TABLE 3. NON-SOILED FABRIC			
TEST WATER		REFLECTANCE (R)	K/S
NON-SOILED FABRIC (RAW FABRIC)	TAP WATER	0.7913	0.0275
	DETERGENT	0.7932	0.0270
	ELECTROLYZED WATER	0.7429	0.0445
	NANO-ELECTROLYZED WATER ^①	0.7920	0.0273
	NANO-ELECTROLYZED WATER ^②	0.7913	0.0275

* The reflectance (R) was the average of the measurements for ten times in total on both sides of each soiled fabric.

[0058]

Table 3 shows the results of the non-soiled fabric. The
15 reflectance (R) is the average of the measurements for ten times
in total on both sides of each soiled fabric.

[0059]

[Table 4]

TABLE 4. CARBON BLACK/MINERAL OIL (BEFORE WASHING)

TEST WATER (AFTER WASHING)		REFLECTANCE (R)	K/S	CLEANING EFFICIENCY (%)
CARBON BLACK / MINERAL OIL	TAP WATER	0.3320	0.6720	37.32
	DETERGENT	0.3810	0.5028	54.15
	ELECTROLYZED WATER	0.3525	0.5947	47.87
	NANO-ELECTROLYZED WATER①	0.3690	0.5395	50.70
	NANO-ELECTROLYZED WATER②	0.3519	0.5968	45.97

[0060]

Table 4 shows the results of carbon black/mineral oil (before 5 washing).

[0061]

[Table 5]

TABLE 5. CARBON BLACK/MINERAL OIL (AFTER WASHING)

TEST WATER (BEFORE WASHING)		K/S	
CARBON BLACK / MINERAL OIL	TAP WATER	0.2596	1.0558
	DETERGENT	0.2583	1.0649
	ELECTROLYZED WATER	0.2534	1.0999
	NANO-ELECTROLYZED WATER①	0.2581	1.0663
	NANO-ELECTROLYZED WATER②	0.2560	1.0811

* The cleaning efficiency (%) was indicated in red, orange, and green in order of precedence.

[0062]

10 Table 5 shows the results of carbon black/mineral oil (after washing).

[0063]

[Table 6]

TABLE 6. BLOOD (BEFORE WASHING)

TEST WATER (BEFORE WASHING)		REFLECTANCE (R)	K/S
BLOOD	TAP WATER	0.1634	2.1417
	DETERGENT	0.1644	2.1236
	ELECTROLYZED WATER	0.1651	2.1110
	NANO-ELECTROLYZED WATER①	0.1643	2.1254
	NANO-ELECTROLYZED WATER②	0.1634	2.1417

* The reflectance (R) was the average of the measurements for ten times in total on both sides of each soiled fabric.

[0064]

Table 6 shows the results of blood (before washing), and the reflectance (R) is the average of the measurements for ten times 5 in total on both sides of each soiled fabric.

[0065]

[Table 7]

TABLE 7. BLOOD (AFTER WASHING)

TEST WATER (AFTER WASHING)		REFLECTANCE (R)	K/S	CLEANING EFFICIENCY (%)
BLOOD	TAP WATER	0.2764	0.9472	56.50
	DETERGENT	0.5093	0.2364	90.01
	ELECTROLYZED WATER	0.3136	0.7512	65.80
	NANO-ELECTROLYZED WATER①	0.6643	0.0848	97.26
	NANO-ELECTROLYZED WATER②	0.6280	0.1102	96.09

* The reflectance (R) was the average of the measurements for ten times in total on both sides of each soiled fabric.

* The cleaning efficiency (%) was indicated in red, orange, and green in order of precedence.

[0066]

10 Table 7 shows the results of blood (after washing).

[0067]

[Table 8]

TABLE 8. COCOA (BEFORE WASHING)

TEST WATER (BEFORE WASHING)		REFLECTANCE (R)	K/S
COCOA	TAP WATER	0.3213	0.7168
	DETERGENT	0.3224	0.7121
	ELECTROLYZED WATER	0.3162	0.7394
	NANO-ELECTROLYZED WATER①	0.3226	0.7112
	NANO-ELECTROLYZED WATER②	0.3176	0.7331

[0068]

Table 8 shows the results of cocoa (before washing).

[0069]

5 [Table 9]

TABLE 9. COCOA (AFTER WASHING)

TEST WATER (AFTER WASHING)		REFLECTANCE (R)	K/S	CLEANING EFFICIENCY (%)
COCOA	TAP WATER	0.4347	0.3676	50.67
	DETERGENT	0.4658	0.3063	59.22
	ELECTROLYZED WATER	0.4020	0.4448	42.39
	NANO-ELECTROLYZED WATER①	0.5426	0.1928	75.80
	NANO-ELECTROLYZED WATER②	0.5110	0.2340	70.74

[0070]

Table 9 shows the results of cocoa (after washing).

[0071]

10 [Table 10]

TABLE 10. RED WINE (BEFORE WASHING)

TEST WATER (BEFORE WASHING)		REFLECTANCE (R)	K/S
RED WINE	TAP WATER	0.4633	0.3109
	DETERGENT	0.4629	0.3116
	ELECTROLYZED WATER	0.4586	0.3196
	NANO-ELECTROLYZED WATER①	0.4620	0.3133
	NANO-ELECTROLYZED WATER②	0.4615	0.3142

[0072]

Table 10 shows the results of red wine (before washing).

[0073]

[Table 11]

TABLE 11. RED WINE (AFTER WASHING)

TEST WATER (AFTER WASHING)		REFLECTANCE (R)	K/S	CLEANING EFFICIENCY (%)
RED WINE	TAP WATER	0.5553	0.1781	46.87
	DETERGENT	0.5569	0.1763	47.54
	ELECTROLYZED WATER	0.5923	0.1403	65.16
	NANO-ELECTROLYZED WATER①	0.5590	0.1740	48.72
	NANO-ELECTROLYZED WATER②	0.5693	0.1629	52.77

5

[0074]

Table 11 shows the results of red wine (after washing).

[0075]

For the fabric to which carbon black dispersed in mineral oil was attached as an artificially soiled fabric, a large difference in detergency was not observed from that of an ordinary detergent. The reason is considered as follows. The mineral oil is emulsified in the washing process by the reaction with a surfactant in the detergent and removed from the fabric. During that time, foams are produced and carbon black is incorporated thereinto as particles.

[0076]

The nano-electrolyzed water, however, does not function as a surfactant. Free carbon black was accordingly incorporated into a woven fiber along with a part of oil, causing a small reduction in the cleaning effect. This is supposed to be the reason.

[0077]

Next, for the fabric having blood attached, the entire fabric in which erythrocytes contained in proteins in the blood remain may discolor. When the fabric is washed with the nano-electrolyzed 5 waters prepared by the method of the nano-electrolyzed water 1 and the nano-electrolyzed water 2, the blood incorporated into the fiber may be removed from the fabric by the nano effect. Consequently, the cleaning effect as expected was obtained.

[0078]

10 Moreover, cocoa is one of luxury grocery items and includes cacao polyphenols which attract attention from health aspects. These are called antioxidants and have a function of self-protection against the attack of active oxygen.

[0079]

15 Cocoa also includes lipid, so that it may be firmly attached to the fiber. Accordingly, it is difficult for tap water, detergents, or others to efficiently remove this kind of lipid incorporated into the woven fiber.

[0080]

20 It was found out that the lipid incorporated into the fiber was efficiently removed when the fabric was washed with the nano-electrolyzed water prepared by the method (1) among the nano-electrolyzed waters.

[0081]

25 Finally, for washing the fabric to which red wine containing

many polyphenols such as flavonoid, anthocyanin, catechin, and tannin is attached, an oxygen bleach is typically used to remove red wine stains. However, an oxygen bleach cannot be used for all textiles.

5 [0082]

Red wine includes about 10% to 15% of alcohol. In order to remove polyphenols dissolved in alcohol absorbed into the fiber of the fabric, the oxidation-reduction potential (ORP, unit: mV) of the electrolyzed water must be taken into consideration.

10 [0083]

Specifically, a larger positive value of ORP results in a larger oxidizing power (aging) and a larger negative value results in a larger reducing power (anti-aging), increasing its function.

[0084]

15 Based on the value of ORP, it was considered that the washing with alkaline electrolyzed water having an ORP value of -192 exerted a high detergency effect. This would be because nanobubbles of the alkaline electrolyzed water were produced by bubbling to increase the ORP value and enhance the reducing power, causing an effect
20 on decrease in cleaning effect.

[0085]

In this test, verification was made using two kinds of nano-electrolyzed waters. Electrolyzed water (2) is nano-electrolyzed water having air sucked. The ORP value of this
25 nano-electrolyzed water shifted to the positive side as compared

with raw water ($-596 \rightarrow -202$ mV). As a result, the detergency of the nano-electrolyzed water 2 was slightly better than that of the nano-electrolyzed water 1.

[0086]

5 According to the above results, two kinds of nano-electrolyzed waters prepared are properly used depending on the types of foods that soiled the fabric. This method efficiently confirmed that the washing with the nano-electrolyzed waters had a higher cleaning effect than that of the washing with a conventional detergent.

10 [0087]

Next, verification will be made on the bactericidal effect of the nano-electrolyzed water generated by this fine bubble electrolyzed water generating apparatus and this method for generating fine bubble electrolyzed water.

15 [0088]

[Content of Test]

Using nano-electrolyzed water, three leaves were randomly taken from a Boston lettuce (commercially available) and one leaf was divided into four parts, each of which was an analyte. The analyte
20 weighted about 8 g. The analyte was lightly washed with water and test water was then sucked up with a bath pump (BP-101K, produced by KOSHIN LTD.) from a sink containing about 50 L test water so that the analyte was washed with running water for two minutes. After the washing, the analyte was lightly washed with water.

25 [0089]

The analyte was placed in a sterilization bag. A 10-fold diluted solution is added to the sterilization bag, followed by homogenization for 1 minute. One ml of the obtained analyte solution was dispensed on a simple medium (petrifilm, plate for general viable bacteria, produced by Sumitomo 3M Limited).

[0090]

Serial dilution was carried out according to the pour plate method. Incubation was carried out under the environment at 35 °C for 48 hours in an incubator. The number of colonies on the medium was counted and the average value of three leaves was obtained as the number of bacteria per g or ml.

[0091]

The followings were used as test water:

- 1) tap water;
- 2) acidic electrolyzed water, pH 3.01, concentration of free residual chlorine: 40 ppm; and
- 3) nanobubble acidic electrolyzed water, pH 3.05, concentration of free residual chlorine: 40 ppm.

Regarding the generation conditions, an electrolytic water generating apparatus (Mamorumizu: registered trademark) (product number: ESS-ZERO) and a fine bubble generator (NANOAQUA: registered trademark) (product number: MN-20) according to production and distribution by the applicant of the present application are manually connected to each other and continuously operated for 15 minutes. The flow rate of acidic electrolyzed water was about 3 L/min.

[0092]

[Number of General Viable Bacteria]

The number of bacteria before and after washing a lettuce was counted. A petrifilm for general viable bacteria was used as a medium. 5 Incubation was carried out in an incubator for about 48 hours and the number of colonies was counted. When the average value of three analytes ($n = 3$) was taken and the average value was smaller than that of untreated analytes by two orders, the bactericidal effect would be obtained.

10 [0093]

[Results]

Compared with no treatment, the treatments with the nanobubble acidic electrolyzed water and the acidic electrolyzed water demonstrated a reduction in the number of bacteria by two or more 15 orders, showing a significant difference. In particular, the treatment with the nanobubble acidic electrolyzed water showed a reduction in the number of bacteria by four or more orders, which was the most excellent result.

[0094]

20 [Table 12]

	NO TREATMENT	TAP WATER	ACIDIC ELECTROLYZED WATER	NANOBUBBLE ACIDIC ELECTROLYZED WATER
NUMBER OF GENERAL VIABLE BACTERIA (cfu/ml)	5.5×10^5	1.5×10^4	8.0×10^3	5.0×10^2

$n = 3$

[0095]

[Summary]

As shown in the above (Table 12) and Fig. 8, the treatment with the nanobubble acidic electrolyzed water demonstrated the largest reduction in the number of general bacteria. This indicated that the nanobubbles were effective for the characteristics (loose 5 leaves are flexible and their thickness is thin.) of the vegetable itself. Furthermore, reproduction and validation or the like have to be succeedingly made on loose leaves. The term "no treatment" means that an analyte of Boston lettuce was purchased but not treated by washing and others.

10 [0096]

[Discussion]

As described above, the Boston lettuce was washed with running nanobubble acidic electrolyzed water for 2 minutes. This successfully eliminated detectable general bacteria. For other 15 leaf vegetables, red leaf lettuce, cabbage and other vegetables were used for preliminary tests but they showed variations in the number of bacteria after the treatments. However, a longer treatment time solved such a problem.

[0097]

20 In addition, frilly lettuce provided good results on actual site. This indicated that there was a difference in the effect between head leaf vegetables and loose leaf vegetables among leaf vegetables.

[0098]

25 Loose leaves have a more flexible surface, a thinner thickness,

and more bacteria attached to the surface than head leaves do. Accordingly, being stirred by washing with running water may allow nanobubbles to easily reach the depths of the vegetables.

[0099]

5 The nanobubble generator (14) generates a gas-liquid mixture containing micro-nanobubbles. For example, the generator (14) may be composed utilizing the invention described in Japanese Patent No. 4563496.

Industrial Applicability

10 [0100]

The present invention can be industrially applied by establishing the technique of the apparatus and the method for generating the fine bubble electrolyzed water, and manufacturing and distributing the apparatus based on the established technique.

15 Reference Signs List

[0101]

- | | | |
|----|-----|--|
| | 1 | Anode chamber |
| | 2 | Cathode chamber |
| | 3 | Intermediate chamber |
| 20 | 4,5 | Diaphragm |
| | 6 | Anode electrode |
| | 7 | Cathode electrode |
| | 8 | Acidic electrolyzed water pipe |
| | 9 | Alkaline electrolyzed water pipe |
| 25 | 10 | Acidic electrolyzed water storage tank |

	11	Alkaline electrolyzed water storage tank
	12	Acidic electrolyzed water conduit
	13	Alkaline electrolyzed water conduit
	14	Nanobubble generator
5	15	Acidic nanobubble electrolyzed water pipe
	16	Alkaline nanobubble electrolyzed water pipe
	17	Acidic nanobubble electrolyzed water discharge pipe
	18	Alkaline nanobubble electrolyzed water discharge pipe
	X	Electrolytic apparatus

CLAIMS

1. A fine bubble electrolyzed water generating apparatus,
5 comprising:

an electrolytic apparatus having a two-chamber structure including an anode chamber with an anode electrode, a cathode chamber with a cathode electrode, and a diaphragm provided between the anode chamber and the cathode chamber, or an electrolytic apparatus having
10 a three-chamber structure including an anode chamber with an anode electrode, a cathode chamber with a cathode electrode, an intermediate chamber provided between these chambers, a diaphragm provided between the anode chamber and the intermediate chamber, and a diaphragm provided between the cathode chamber and the
15 intermediate chamber;

an acidic electrolyzed water storage tank for storing acidic electrolyzed water and an alkaline electrolyzed water storage tank for storing alkaline electrolyzed water, both provided adjacent to the electrolytic apparatus;

20 pipes through which these electrolyzed water storage tanks are in communication with the anode chamber and the cathode chamber, respectively; and

nanobubble generators for generating nanobubbles with gas and liquid, which are in communication with these electrolyzed water
25 storage tanks, respectively,

wherein these nanobubble generators generate chlorine nanobubble electrolyzed water with chlorine gas and acidic electrolyzed water produced in the anode chamber, and generate hydrogen nanobubble electrolyzed water with hydrogen gas and 5 alkaline electrolyzed water produced in the cathode chamber.

2. The fine bubble electrolyzed water generating apparatus according to claim 1, wherein the apparatus generates nanobubbles of chlorine gas in anode electrolyzed water with the chlorine gas 10 and the anode electrolyzed water produced at the anode electrode, and supplies the anode electrolyzed water including the nanobubbles to the acidic electrolyzed water storage tank.

3. The fine bubble electrolyzed water generating apparatus 15 according to claim 1, wherein the apparatus generates nanobubbles of hydrogen gas in cathode electrolyzed water with the hydrogen gas and the cathode electrolyzed water produced at the cathode electrode, and supplies the cathode electrolyzed water including the nanobubbles to the alkaline electrolyzed water storage tank.

20

4. The fine bubble electrolyzed water generating apparatus according to any one of claims 1 to 3, wherein the fine bubble acidic electrolyzed water containing the nanobubbles generated in the anode electrolyzed water is used for sterilization.

25

5. A method for generating fine bubble electrolyzed water, comprising: electrolyzing an aqueous solution of sodium chloride; processing chlorine gas and acidic electrolyzed water generated at an anode electrode by a nanobubble generator to generate chlorine 5 nanobubble electrolyzed water; and processing hydrogen gas and alkaline electrolyzed water generated at a cathode electrode by a nanobubble generator to generate hydrogen nanobubble electrolyzed water.