

CLEAN METHOD FOR PREPARING D,L-METHIONINE

Technical Field

The present invention is related to the compound synthesis field, more specifically, a clean
5 method for preparing D,L-methionine.

Background of the Invention

D,L-methionine is a sulfur-containing amino acid essential to human body and is closely
related to metabolism of various sulfur-containing compounds in biological organisms.
10 Methionine belongs to the first restrictive amino acid and cannot be synthesized inside animal
body and needs to be taken from the food. Adding methionine into the feedstuffs can promote
growth of livestock, increase lean meat percentage, and shorten the feeding cycle. Methionine
can also be used for prevention and treatment of liver diseases such as chronic or acute hepatitis
and hepatic cirrhosis because methyl group contained in methionine can be used for methylation
15 of toxicants or medicines and thus have detoxification effect. Methionine can also be used for
alleviation of toxic reaction of hazardous substances such as arsenic, trichloride, carbon
tetrachloride, benzene, pyridine and quinoline. According to relevant statistical information,
market demand for methionine worldwide has reached 1,000,000ton/year. The demand for
methionine in international market is increasing at a year-to-year growth rate of 4% in recent
20 years; whereas the demand in China is increasing at a year-to-year growth rate of 7%.

As introduced in relevant literatures, there are following methods for synthesis of
D,L-methionine:

In the patent publication with the publication number of CN1923807A, Sumitomo
Chemical Co., Ltd. from Japan proposes to prepare D,L-methionine through hydrolysis of
25 5-(β -methylthioethyl)glycolyurea with Alkaline potassium compounds, then acidification with
pressurized CO₂ gas, and then fractionized condensation and crystallization with addition of
polyvinyl alcohol. Although such method can recycle CO₂ gas, due to factitious addition of
polyvinyl alcohol, the crystallization mother solution as discharged cannot be recycled and can
only be wasted, which produces a large quantity of wastewater S and N containing organic
30 compounds. Furthermore, fractionized condensaton method consumes a large quantity of thermal
energy to the extent of resulting in significant increase in product cost for D,L-methionine.

In the patent publication with the publication number of CN85108505A, Wang Jianhua proposes to add a mixture of neutral amino acid and organic acid, containing 10% acrolein, as the catalyst during preparation of 5-(β -methylthioethyl)glycolyurea; it aims to prepare 5-(β -methylthioethyl)glycolyurea with a one-step method by using methanethiol and acrolein, simplify operation process, and improve preparation yield of 5-(β -methylthioethyl)glycolyurea. However, a mixture of neutral amino acid and organic acid that is intentionally added for further saponification with the addition of alkaline substances inevitably participates in the reaction, and produces metal salt. As a result, the metal salts accumulated in the mother solution so that the mother solution cannot be recycled and produces a large quantity of wastewater S and N containing organic compounds.

In the patent publication with the publication number of CN85108531A, Wang Jianhua proposes to produce D,L-methionine through saponification of 5-(β -methylthioethyl)glycolyurea with alkaline sodium compounds, acidification with vitriol and separation with fractionized condensation and crystallization method. Such method may produce a large quantity of by-products such as sodium sulfate, separation and processing of sodium sulfate will become a significant burden for production.

In the patent publication with the publication number of CN1589259A, to obtain a methionine preparation method that can stably produce a granular or tabular methionine crystal of high bulk density and quality, Nippon Soda Co., Ltd. proposes to use one intermetallic compound selected from metal hydroxide, metal carbonate or metal carbonate salt to hydrolyze 5-(β -methylthioethyl)glycolyurea to obtain the methionine metal salt, then neutralize the methionine metal salt by using pressurized CO₂ gas to realize crystallization of methionine. After that, the methionine is separated from the filtrate, and one makes sure that the filtrate can be recycled for hydrolysis of 5-(β -methylthioethyl)glycolyurea during preparation of methionine. According to this method, despite of the fact that partial mother solution for crystallization is again used for hydrolysis, some tailings and by-products are not eliminated during intermediate process and are inevitably accumulated to the extent of affecting application effect and product quality. Meanwhile, because the mother solution is not used for preparation of 5-(β -methylthioethyl)glycolyurea, a large quantity of water is required, causing the water amount cannot be balanced over the whole process. Furthermore, each use will consume a large quantity of water as required by preparation of 5-(β -methylthioethyl)glycolyurea.

In the patent publication with the publication number of CN1103066A, Degussa AG proposes a synthesis technique to obtain an amino nitrile compound through reaction among methylthio propionaldehyde, HCN and ammonia, and hydrolyze amino nitrile compound with a ketone catalyst to obtain the aminoamide, and then finally prepare methionine through high-temperature hydrolysis with an alkaline catalyst. Such methionine synthesis method also has some defects despite of the fact that it is provided with unique features as compared with industrialized production method. The most serious defect lies in a large quantity of wastewater as produced; for instance, aminoamide obtained through addition of ketone catalyst during its preparation shall subject to separation and purification by passing through column, such separation and purification method is inappropriate for such bulk product during industrialized production. Viewed from overall techniques, it fails to solve the problem of a large quantity of waste water.

In the patent publication with the publication number of CN102399177A, Li Kuanyi proposes a green and clean technical method for continuous synthesis of methionine: taking methylthio propionaldehyde as synthesized with acrolein and methanthiol as the material for reaction with hydrogen cyanide to obtain the intermediate 2-hydroxy-4-(methylthio)nitrile; further obtaining hydantoin solution through continuous reaction of the intermediate 2-hydroxy-4-(methylthio)nitrile in the first reaction bed of combined reactor in the presence of excessive ammonia and carbon dioxide. The hydantoin solution will flow out of the first reaction bed, and release excessive carbon dioxide and ammonia in the desorption column. The hydantoin solution subjecting to desorption will flow in the second reaction bed of the combined reactor for hydrolysis under alkaline conditions to obtain the methionine potassium solution. The methionine potassium solution with carbon dioxide are neutralized to obtain methionine and potassium hydrogen carbonate solution. Methionine is separated from the solution by means of crystallization; whereas potassium hydrogen carbonate and mother solution shall be subject to further treatment for recycling. It appears that this technical approach is extremely clean and appropriate for industrialized production. However, according to its patent introduction and analysis from the angle of industrialized production, it still has some disadvantages. First, such technique requires a large quantity of water for reaction during preparation of hydantoin solution; whereas mother solution for crystallization is used for hydrolysis other than preparation of hydantoin. As a result, a large quantity of water is distilled for preparation of hydantoin during

hydrolysis to further apply remaining mother solution for hydrolysis of hydantoin. This process results in increased production cost because it requires a large amount of thermal and electric energies. Second, delayed elimination of by-products as produced during chemical reaction in the whole technological process may seriously affect the quality of final products, and making it more difficult for purification of final products. Meanwhile, it also affects the consumption and batch of mother solution. This results in increased emission of wastewater during industrialized production.

As proposed in US2004/0039228A1, D,L-methionine can be obtained through reaction among some mother solution for crystallization, certain amount of NH_3 and CO_2 at the temperature of 60°C in the presence of TiO_2 , the catalyst, saponification at the temperature of $180\text{--}300^\circ\text{C}$ and neutralization with CO_2 . Despite of the fact that such method makes use of partial crystallization mother solution indiscriminately to minimize the emission of wastewater, it will result in the following problems due to addition of as TiO_2 catalyst in the technological process: if all mother solution for crystallization is used indiscriminately, it will result in accumulation of TiO_2 ; Furthermore, D,L-methionine subjected to crystallization and filtration will also contain TiO_2 and thus the product will contain metal Ti that will result in reduced product quality, which makes follow-up treatment more complicated and difficult.

In the patent publication with the publication number of EB1761074A1(CN101602700A), Sumitomo Chemical puts forward the following method: using alkaline potassium compound as the catalyst for saponification of 5-(β -methylthioethyl)glycolylurea; using CO_2 for neutralization and crystallization; using crystallization mother solution subjected to aforesaid fractionized concentration and crystallization for saponification of reaction fluid of follow-up 5-(β -methylthioethyl)glycolylurea indiscriminately; adding polyvinyl alcohol into the remaining mother liquor subjected to secondary concentration for crystallization. At this point, mother solution is abandoned other than indiscriminate use. Such technique is used by Sumitomo Chemical for industrialized production of D,L-methionine. However, there are numerous problems with this technique, especially on the aspect of green and clean production. Only partial crystallization mother solution is used. Most of crystallization mother solution is abandoned after one-time use, which produces a large quantity of wastewater.

Summary of the Invention

In view of excessive amount of wastewater with N and S containing organic compounds - produced during industrialized production of D,L-methionine and serious environmental pollution, a major technical problem solved by the present invention is to provide a DL-methionine preparation method that is clean and easy to operate.

5 For this purpose, the present invention provides a clean method for preparing D,L-methionine, comprising the following steps:

(1) preparing 5-(β -methylthioethyl)glycolyurea reaction fluid by using 3-methylthio propionaldehyde, KCN and NH_4HCO_3 solution as raw materials for continuous reaction in a tubular reactor that is gradually heated;

10 (2) proceeding with decompression of aforesaid 5-(β -methylthioethyl)glycolyurea reaction fluid to separate NH_3 and CO_2 from the reaction fluid for delivery to another tubular reactor for decomposition and saponification to obtain a saponification fluid;

(3) eliminating the produced NH_3 and CO_2 during saponification through desorption of the saponification fluid to obtain D, L-methionine potassium solution;

15 (4) obtaining NH_4HCO_3 solution through spraying with water to the absorbed NH_3 and CO_2 as separated in Steps (2) and (3), and directly using the NH_4HCO_3 solution as the raw material for preparation of 5-(β -methylthioethyl)glycolyurea reaction fluid in Step (1);

(5) cooling down D,L-methionine potassium solution obtained in Step (3), and then using an organic solvent for reversed and continuous extraction; directly using the organic solvent
20 recycled through distillation of an organic solvent layer for reversed and continuous extraction; raffinate produced by distillation is treated as a waste fluid;

(6) delivering water as produced through reversed and continuous extraction in Step (5) to the continuous crystallizer, simultaneously delivering CO_2 gas to the continuous crystallizer for acidification to control PH value of reaction fluid in the continuous crystallizer at 6-9, and then
25 obtaining crystallization intermixture;

(7) separating solid and liquid in the crystallization intermixture to obtain the crude D,L-methionine product and crystallization mother solution; washing crude D,L-methionine product with water for solid-liquid separation to obtain filter cake of D,L-methionine and washing filtrate;

30 (8) drying the filter cake of D,L-methionine under the protection of an inert gas to obtain D,L-methionine product;

(9) combining the crystallization mother solution in Step (7) with washing filtrate for delivery to the decomposition column to fully decompose KHCO_3 into K_2CO_3 for release of CO_2 gas; released CO_2 gas subjecting to compression is delivered to a continuous crystallizer for acidification and crystallization;

5 (10) using K_2CO_3 -containing mother solution obtained in Step (9) for absorption of HCN gas to prepare a KCN solution that is used as the material for preparation of 5-(β -methylthioethyl)glycolurea in Step (1).

In the Step (1), the preferred mole ratio of 3-methylthio propionaldehyde, KCN and NH_4HCO_3 is 1:1--1.1:2--3. At this point, an optimal single-pass conversion rate of 3-methylthio
10 propionaldehyde and equipment utilization rate can be obtained.

In the Step (1), the preferred gradual heating range is 50--150°C with a reaction time up to 3-15 minutes.

In the Step (2), the preferred temperature for decomposition and saponification is 140
15 --220°C with a time up to 2--15 minutes.

In the Step (5), the temperature of D,L-methionine potassium solution is reduced to 0--40°C before using an organic solvent equivalent to 0.5-2 times the weight of D,L-methionine potassium solution for reversed and continuous extraction.

In the Step (5), the preferred organic solvents for reversed and continuous extraction
20 include one or more of toluene, ethylbenzene, dimethylbenzene, n-butyl alcohol, isobutanol, n-pentanol, 2-methyl-1-butanol, isoamylol, sec-amyl alcohol, 3-pentanol, tert-amyl alcohol, n-hexyl alcohol, 4-Methyl-2-pentanol, 2-ethyl alcohol, 2-methyl amyl alcohol, heptanol, 2-heptanol, 3-heptanol, 2-ethylhexyl alcohol, 2-octanol, octanol, 3,5,5-trimethylhexanol, ether, MTBE (methyl tertiary-butyl ether), isopropyl ether, n-propyl ether, n-butyl ether, isoamyl ether,
25 hexyl ether, 2-methyl tetrahydrofuran, anisole, ethoxybenzene, 3-methylanisole, EBE (ethyl benzyl ether), ethylene glycol diethyl ether, diethylene dipropyl ether and ethylene glycol di-butyl ether.

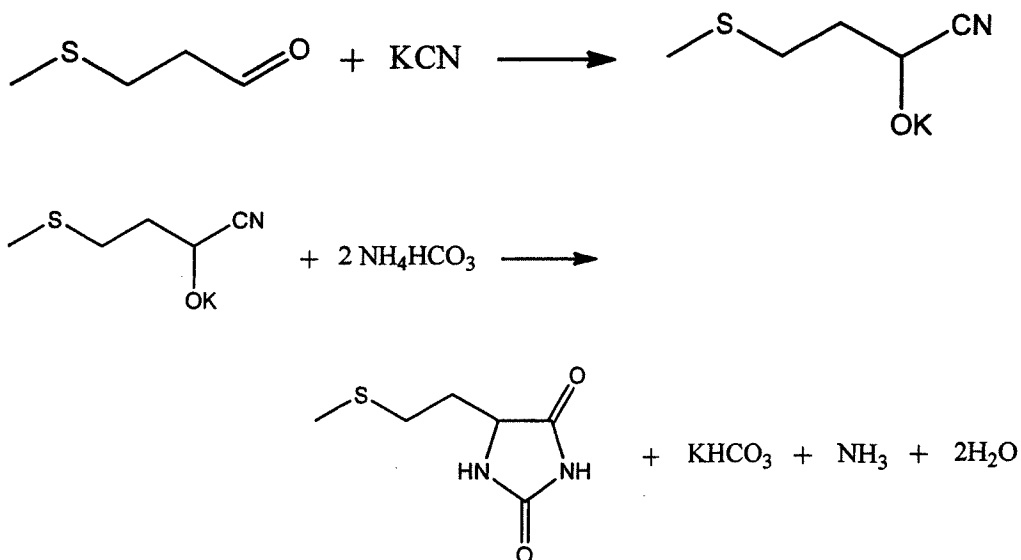
Some oil soluble viscous by-products might be produced during preparation of 5-(β -methylthioethyl)glycolurea and D,L-methionine potassium. These by-products can be
30 extracted by the aforesaid organic solvents; whereas D,L-methionine is a highly water soluble substance, which is unlikely dissolved in the aforesaid organic solvents;. Therefore, by-products

are eliminated in time through extraction by organic solvents.

In the Step (6), the temperature for continuous crystallization is 0--40°C; whereas reaction fluid is to stay inside the continuous crystallizer for 0.5--5 hours.

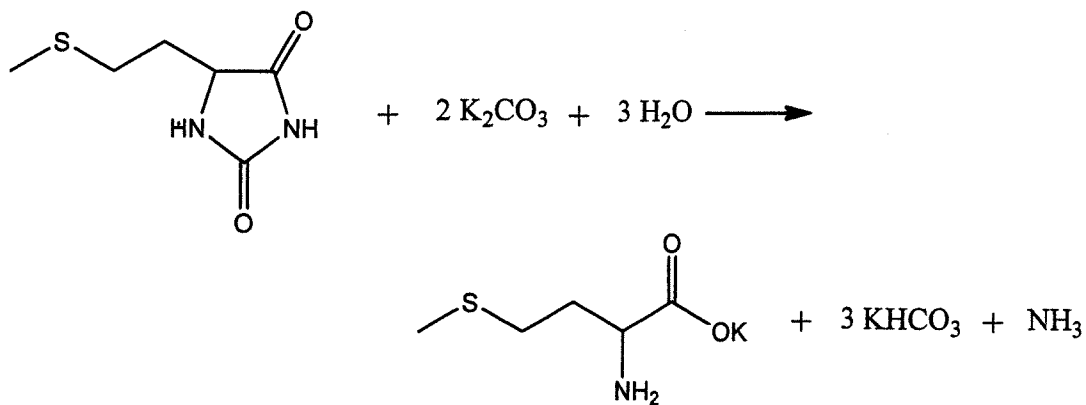
In the Step (9), temperature and pressure of decomposition column are 110--160°C and 0.15--0.8MPa with decomposition reaction time up to 1--4 hours.

The reaction formula in the Step (1) is:

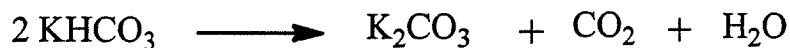


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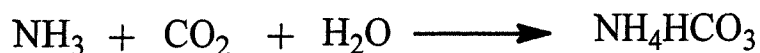
The reaction formula in the Steps (2) and (3) are:



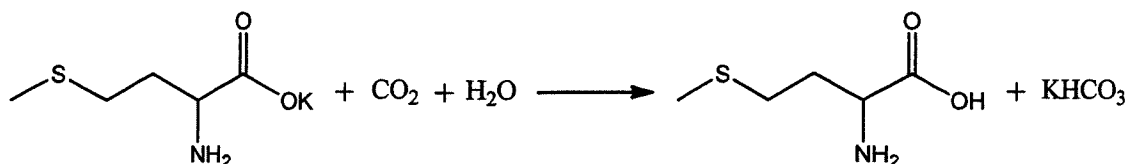
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The reaction formula in the Step (4) is:

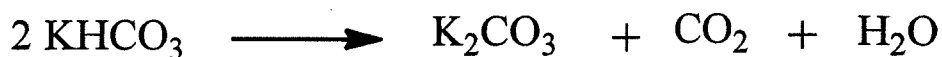


The reaction formula in the Step (6) is:



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The reaction formula in the Step (9) is:



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The reaction formula in the Step (10) is:



In the Step (5), by-products produced during reaction are eliminated by means of extraction so as to ensure repeated and unlimited recycling of process water.

The present invention aims to decompose KHCO_3 into K_2CO_3 through heating, and use K_2CO_3 -containing crystallization mother solution as HCN absorption liquid to prepare KCN solution. This eliminates the need to add other metal salt during the whole technical process so as to lay down a solid foundation for recycling of process water in the whole preparation process.

To ensure a clean process, NH_3 and CO_2 gases produced during preparation of 5-(β-methylthioethyl)glycolurea of the present invention and saponification are to subject to absorption to form NH_4HCO_3 again; whereas NH_4HCO_3 is a material as required for the preparation of 5-(β-methylthioethyl)glycolurea. This aims to ensure maximum utilization of auxiliary materials as produced during the technical process.

According to analysis of industrialized production process, D,L-methionine preparation process using methods as reported in existing literatures may produce a large quantity of

wastewater. Furthermore, N and S-containing organic compounds in the wastewater may result in serious environmental pollution and high process cost. The process of the present invention can prevent production of wastewater from the source. Furthermore, comprehensive utilization of exhaust gas produced during production can basically prevent production of exhaust gas during technical process. On this account, D,L-methionine synthesis technique of the present invention is a clean technical approach in accommodation with industrialized production.

Brief Description of Drawings

Fig. 1 is the flow chart for the present invention.

In the drawing, T1 is a HCN absorption column; T2 is a CO₂ and NH₃ absorption column; R1 is the first tubular reactor; R2 is a pressure relief tank; R3 is the second tubular reactor; T3 is a desorption column; T4 is a continuous extraction column; T5 is an organic solvent recycling and distillation column; T6 is a continuous crystallizer; S1 is the first solid-liquid separator; R4 is a scrubber; S2 is the second solid-liquid separator; T7 is a KHCO₃ decomposition column.

Detailed Description of The Invention

The present invention is described in reference to the following specific embodiments. These embodiments are only for explanation, which shall not be deemed as restriction of scope or implementation method of the present invention.

Embodiment 1

A metering liquid delivery pump was used to deliver treated crystallization mother solution (10% K₂CO₃ solution was used at the initial start-up) to the HCN absorption column T1 at the flow rate of 700Kg/h. Meanwhile, 5.4Kg/h HCN gas was delivered to the HCN absorption column for preparation of a KCN solution in the HCN absorption column T1, and a liquid delivery pump was used to deliver the KCN solution to the first tubular reactor R1. A liquid delivery pump was used to deliver ammonium bicarbonate absorbing ammonia gas and carbon dioxide in the CO₂NH₃ absorption column T2 to the first tubular reactor R1 (16% ammonia bicarbonate solution was used at the initial start-up) at the velocity of 200Kg/h, and a liquid delivery pump was used to deliver 3-methylthio propionaldehyde to the first tubular reactor R1 at the velocity of 20.8Kg/h simultaneously. The reaction fluid was stayed for 7 minutes

respectively at the temperature of 50°C and 100°C, and stayed for another 1 minute at the temperature of 150°C. The reaction fluid discharged from the first tubular reactor R1 was subject to decompression in the decompression tank R2 to discharge surplus ammonia gas and carbon dioxide to CO₂ and NH₃ absorption column T2 to obtain 910Kg/h 5-(β-methylthioethyl)glycolurea solution.

A liquid delivery pump was used to deliver the prepared 5-(β-methylthioethyl)glycolurea solution to the second tubular reactor R3 at the velocity of 910Kg/h. The reaction fluid was stayed in the second tubular reactor for 15 minutes at the temperature of 140°C for saponification prior to decompression in the T3 absorption column. Ammonia gas and carbon dioxide produced during reaction were discharged to the CO₂ and NH₃ absorption column T2 for absorption. The saponification liquid was cooled down to the temperature of 40°C in the heat exchanger, which was further delivered to the continuous extraction column via the pump at the velocity of 906Kg/h. Meanwhile, 453Kg/h ethylene glycol dipropyl ether was delivered to the continuous extraction column T4 for continuous reversed extraction. Continuous differential distillation of ethylene glycol dipropyl ether layer was conducted in the organic solvent recycling distillation column T5 to recycle the solvent for circulated use during extraction process, and incinerate tailings at the same time. A water layer was delivered to the continuous crystallizer T6 at the velocity of 935Kg/h, and simultaneously CO₂ gas was delivered to the continuous crystallizer T6 for acidification until PH value of the reaction system is up to 8. Continuous crystallization was conducted at the temperature of 40°C, and the reaction fluid stayed in continuous crystallizer for 3 hours. The crystallization liquid was subject to continuous centrifugal solid-liquid separation in the first solid-liquid separator S1 to filtrate crude D,L-methionine product, and prepare a crystallization mother solution for further use. A solid delivery equipment was used to deliver crude D,L-methionine product to the scrubber R4 for spiral agitation and washing in the specified quantity at the velocity of 34Kg/h. Meanwhile, 40Kg/h water was continuously delivered to the scrubber R4 to control temperature of materials in the scrubber R4 below 5°C. The solid material was stayed in the scrubber R4 for 15 minutes. The intermixture subjected to continuous agitation and washing in the scrubber R4 was under continuous centrifugal solid-liquid separation via the second solid-liquid separator S2. The washing filtrate was poured into the crystallization mother solution, and D,L-methionine as filtered was airflow dried at the temperature of 110°C under the protection of N₂ to obtain finished 27.5Kg/h D,L-methionine

product. Inspection was conducted according to GB-T17810-2009 Feed Grade DL-Methionine to obtain the result that methionine content was up to 99.5%, and total molar reaction yield based on 3-methylthio propionaldehyde was up to 92.3%.

5 The crystallization mother solution and washing filtrate were delivered to KHCO_3 decomposition column T7 at the velocity of 977 Kg/h, and they stayed for 1 hour at the temperature of 160°C to totally decompose potassium hydrogen carbonate as contained in the mother solution into potassium carbonate. Meanwhile, 210Kg/h water and 45Kg/h ethylene diglycol ether were prepared through distillation at the column top. The distilled water was circulated to the scrubber R4 for agitation prior to further use at 40Kg/h, whereas remaining
10 170Kg/h distilled water was used in the CO_2 and NH_3 absorption column T2. Distilled ethylene diglycol ether was circulated to the extraction column T4 for further use, whereas CO_2 gas discharged through decomposition was used for acidification and crystallization in continuous crystallizer T6 after compression. Potassium carbonate solution containing mother solution discharged from the column bottom was cooled prior to circulation to HCN absorption column
15 T1 for absorption of formonitrile.

Embodiment 2

A metering liquid delivery pump was used to deliver treated crystallization mother solution (10% K_2CO_3 solution was used at the initial start-up) to the T1 HCN absorption column at the
20 flow rate of 552Kg/h. Meanwhile, HCN gas was delivered to the HCN absorption column at 5.94Kg/h for preparation of the KCN solution in the HCN absorption column T1, and a liquid delivery pump was used to deliver the KCN solution to the first tubular reactor R1. A liquid delivery pump was used to deliver ammonium bicarbonate absorbing ammonia gas and carbon dioxide in the CO_2NH_3 absorption column T2 to the first tubular reactor R1 (16% ammonia
25 bicarbonate solution was used at the initial start-up) at the velocity of 250Kg/h, and a liquid delivery pump was used to deliver 3-methylthio propionaldehyde to the first tubular reactor R1 at the velocity of 20.8Kg/h simultaneously. The reaction fluid stayed for 3 minutes respectively at the temperature of 60°C and 120°C, and stayed for another 1 minute at the temperature of 150°C. The reaction fluid discharged from the first tubular reactor R1 was subject to decompression in
30 the decompression tank R2 to discharge surplus ammonia gas and carbon dioxide to CO_2 and NH_3 absorption column T2 to obtain 805Kg/h 5-(β-methylthioethyl)glycolurea solution.

A liquid delivery pump was used to deliver 5-(β -methylthioethyl)glycolyurea solution prepared to the second tubular reactor R3 at the velocity of 805Kg/h. The reaction fluid stayed in the second tubular reactor for 2 minutes at the temperature of 220°C for saponification prior to decompression in the T3 absorption column. Ammonia gas and carbon dioxide produced during reaction were discharged to CO₂ and NH₃ absorption column T2 for absorption. The saponification liquid was cooled down to the temperature of 0°C in the heat exchanger, which was further delivered to continuous extraction column T4 via the pump at the velocity of 800Kg/h. Meanwhile, 1600Kg/h ethylene glycol dipropyl ether was delivered to the continuous extraction column T4 for continuous reversed extraction. Continuous differential distillation of ethylene glycol dipropyl ether layer was conducted in the organic solvent recycling distillation column T5 to recycle the solvent for circulated use during extraction process, and incinerate tailings at the same time. A water layer was delivered to the continuous crystallizer T6 at the velocity of 790Kg/h, and simultaneously CO₂ gas was delivered to continuous crystallizer T6 for acidification until PH value of the reaction system was up to 6. Continuous crystallization was conducted at the temperature of 0°C, and the reaction fluid stayed in the continuous crystallizer for 0.5 hour. The crystallization liquid was subject to continuous centrifugal solid-liquid separation in the first solid-liquid separator S1 to filtrate crude D,L-methionine product, and the crystallization mother solution was prepared for further use. A solid delivery equipment was used to deliver crude D,L-methionine product to the scrubber R4 for spiral agitation and washed in the specified quantity at the velocity of 34Kg/h. Meanwhile, 102Kg/h water was continuously delivered to the scrubber R4 and the temperature of materials in the scrubber R4 was controlled below 0°C. The solid material stayed in the scrubber R4 for 5 minutes. Continuous centrifugal solid-liquid separation of intermixture was conducted subjecting to continuous agitation and washing in the scrubber R4 via the second solid-liquid separator S2. Washing filtrate was poured into the crystallization mother solution, and the filter cake of D,L-methionine filtered at the temperature of 110°C was airflow dried under the protection of N₂ to obtain 27.2Kg/h finished D,L-methionine product. Inspection was conducted according to GB-T17810-2009 Feed Grade DL-Methionine to obtain the results that methionine content was up to 99.3%, and total molar reaction yield as based on 3-methylthio propionaldehyde was up to 91.3%.

Crystallization mother solution and washing filtrate were delivered to the KHCO₃ decomposition column T7 at the velocity of 890Kg/h, and they stayed for 3 hours at the

temperature of 130°C to decompose potassium hydrogen carbonate as contained in the mother solution into potassium carbonate. Meanwhile, 312Kg/h water and 5Kg/h toluene were obtained through distillation at the column top. 102Kg/h distilled water was circulated to the scrubber R4 for agitation prior to further use; whereas remaining 210Kg/h distilled water was used in the CO₂ and NH₃ absorption column T2. Distilled toluene was circulated to the extraction column T4 for further use; whereas CO₂ gas discharged through decomposition was used for acidification and crystallization in continuous crystallizer T6 after compression. Potassium carbonate solution containing mother solution discharged from the column bottom was cooled prior to circulation to HCN absorption column T1 for absorption of formonitrile.

Embodiment 3

A metering liquid delivery pump was used to deliver treated crystallization mother solution (10% K₂CO₃ solution was used at the initial start-up) to the HCN absorption column T1 at the flow rate of 828Kg/h. Meanwhile, 5.67Kg/h HCN gas was delivered to the HCN absorption column for preparation of the KCN solution in the HCN absorption column T1, and a liquid delivery pump was used to deliver the KCN solution to the first tubular reactor R1. A liquid delivery pump was used to deliver ammonium bicarbonate absorbing ammonia gas and carbon dioxide in CO₂NH₃ absorption column T2 to the first tubular reactor R1 (16% ammonia bicarbonate solution was used at the initial start-up) at the velocity of 300Kg/h, and a liquid delivery pump was used to deliver 3- methylthio propionaldehyde to the first tubular reactor R1 at the velocity of 20.8Kg/h simultaneously. The reaction fluid stayed for 1 minutes and 2 minutes respectively at the temperature of 90°C and 120°C, and stayed for another 2 minutes at the temperature of 150°C. The reaction fluid discharged from the first tubular reactor R1 was subject to decompression in the decompression tank R2 to discharge surplus ammonia gas and carbon dioxide to T2 CO₂ and NH₃ absorption column T2 to obtain 1122Kg/h 5-(β-methylthioethyl)glycolyurea solution.

A liquid delivery pump was used to deliver the prepared 5-(β-methylthioethyl)glycolyurea solution to the second tubular reactor R3 at the velocity of 1122Kg/h. The reaction fluid stayed in the second tubular reactor R3 for 4 minutes at the temperature of 200°C for saponification prior to decompression in the T3 absorption column. The produced ammonia gas and carbon dioxide during reaction were discharged to CO₂ and NH₃ absorption column T2 for absorption.

The saponification liquid was cooled down to the temperature of 20°C in the heat exchanger, which was further delivered to the continuous extraction column T4 via the pump at the velocity of 1117Kg/h. Meanwhile, 1000Kg/h ethylene glycol dipropyl ether was delivered to the continuous extraction column T4 for continuous reversed extraction. Continuous differential distillation of ethylene glycol dipropyl ether layer in the organic solvent recycling distillation column T5 was conducted to recycle the solvent for circulated use during extraction process, and incinerate tailings at the same time. A water layer was delivered to continuous crystallizer T6 at the velocity of 1130Kg/h, and simultaneously CO₂ gas was delivered to the continuous crystallizer T6 for acidification until PH value of the reaction system is up to 8. Continuous crystallization was conducted at the temperature of 20°C, and the reaction fluid stayed in continuous crystallizer for 5 hours. The crystallization liquid was subject to continuous centrifugal solid-liquid separation in the first solid-liquid separator to filtrate crude D,L-methionine product, and crystallization mother solution was prepared for further use. A solid delivery equipment was used to deliver crude D,L-methionine product to the scrubber R4 for spiral agitation and washing in the specified quantity at the velocity of 35Kg/h. Meanwhile, 70Kg/h water was continuously delivered to the scrubber R4 to control temperature of materials in the scrubber R4 below 20°C; solid material is to be stay in the scrubber R4 for 10 minutes; proceed with continuous centrifugal solid-liquid separation of intermixture subjecting to continuous agitation and washing in the scrubber R4 via the second solid-liquid separator S2; pour washing filtrate into the crystallization mother solution, and proceed with airflow drying of filter cake of D,L-methionine as filtered at the temperature of 140°C under the protection of N₂ to obtain 27.9Kg/h finished D,L-methionine product. Proceed with inspection as per GB-T17810-2009: Feed Grade DL-methionine to make sure that methionine content is up to 99.4%, and total molar reaction yield as based on 3-methylthio propionaldehyde is up to 93.6%.

Deliver crystallization mother solution and washing filtrate to KHCO₃ decomposition column T7 at the velocity of 1207Kg/h, and let them to stay for 4 hours at the temperature of 110°C to decompose potassium hydrogen carbonate as contained in the mother solution into potassium carbonate. Meanwhile, 328Kg/h water and 25Kg/h amyl carbinol were prepared through distillation at the column top. 70Kg/h distilled water was circulated to the scrubber R4 for agitation prior to further use; whereas remaining 258Kg/h distilled water was used to CO₂ and NH₃ absorption column T2. Distilled amyl carbinol was circulated to the extraction column

T4 for further use; whereas CO₂ gas discharged through decomposition was used for acidification and crystallization in the continuous crystallizer T6 after compression. Potassium carbonate solution contained mother solution as discharged from the column bottom was cooled prior to circulation to the HCN absorption column T1 for absorption of formonitrile.

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Embodiment 4

A metering liquid delivery pump was used to deliver treated crystallization mother solution (10% K₂CO₃ solution was used at the initial start-up) to the HCN absorption column T1 at the flow rate of 700Kg/h. Meanwhile, 5.4Kg/h HCN gas was delivered to the HCN absorption column for preparation of the KCN solution in HCN absorption column T1, and a liquid delivery pump was used to deliver the KCN solution to the first tubular reactor R1. A liquid delivery pump was used to deliver ammonium bicarbonate absorbing ammonia gas and carbon dioxide in the CO₂NH₃ absorption column T2 to the first tubular reactor R1 (16% ammonia bicarbonate solution was used at the initial start-up) at the velocity of 200Kg/h, and a liquid delivery pump was used to deliver 3-methylthio propionaldehyde to the first tubular reactor R1 at the velocity of 20.8Kg/h simultaneously. The reaction fluid stayed for reaction for 3 minutes at the temperature of 150°C. The reaction fluid discharged from the first tubular reactor R1 was subject to decompression in the decompression tank R2 to discharge surplus ammonia gas and carbon dioxide to the CO₂ and NH₃ absorption column T2 to obtain 910Kg/h 5-(β-methylthioethyl)glycolurea solution.

A liquid delivery pump was used to deliver the prepared 5-(β-methylthioethyl)glycolurea solution to the second tubular reactor R3 at the velocity of 910Kg/h. The reaction fluid stayed in the second tubular reactor R3 for 10 minutes at the temperature of 170°C for saponification prior to decompression in the T3 absorption column. The produced ammonia gas and carbon dioxide during reaction were discharged to the CO₂ and NH₃ absorption column T2 for absorption. The saponification liquid was cooled down to the temperature of 10°C in the heat exchanger, which was further delivered to continuous extraction column T4 via the pump at the velocity of 905Kg/h. Meanwhile, 905Kg/h ethylene glycol dipropyl ether was delivered to the continuous extraction column T4 for continuous reversed extraction. Continuous differential distillation of the ethylene glycol dipropyl ether layer in the organic solvent recycling distillation column T5 was conducted to recycle the solvent for circulated use during extraction process, and incinerate

tailings at the same time. A water layer was delivered to the continuous crystallizer T6 at the velocity of 915Kg/h, and simultaneously CO₂ gas was delivered to the continuous crystallizer T6 for acidification until PH value of the reaction system was up to 7.5. Continuous crystallization was conducted at the temperature of 10°C, and the reaction fluid stayed in the continuous crystallizer for 2 hours. The crystallization liquid was subject to continuous centrifugal solid-liquid separation in the first solid-liquid separator S1 to filtrate crude D,L-methionine product, and the crystallization mother solution was prepared for further use. A solid delivery equipment was used to deliver crude D,L-methionine product to the scrubber R4 for spiral agitation and washing in the specified quantity at the velocity of 35Kg/h. Meanwhile, 35Kg/h water was continuously delivered to the scrubber R4 to control temperature of materials in the scrubber R4 below 10°C. The solid material stayed in the scrubber R4 for 10 minutes. The intermixture subjected to continuous agitation and washing in the scrubber R4 was under continuous centrifugal solid-liquid separation via the second solid-liquid separator S2. The washing filtrate was poured into the crystallization mother solution, and the filter cake of D,L-methionine as filtered was airflow dried at the temperature of 130°C under the protection of N₂ to obtain 28.0Kg/h finished D, L-methionine product. Inspection was conducted according to GB-T17810-2009 Feed Grade DL-methionine to obtain the results that methionine content was up to 99.2%, and total molar reaction yield as based on 3-methylthio propionaldehyde was up to 94.0%.

The crystallization mother solution and washing filtrate were delivered to the KHCO₃ decomposition column T7 at the velocity of 947Kg/h, and they stayed for 1.5 hours at the temperature of 150°C to decompose potassium hydrogen carbonate as contained in the mother solution into potassium carbonate. Meanwhile, 207Kg/h water and 18Kg/h diethyl ether were prepared through distillation at the column top. The distilled water was circulated to the scrubber R4 at 35Kg/h for agitation prior to further use; whereas remaining 172Kg/h distilled water was used to the CO₂ and NH₃ absorption column T2. The distilled diethyl ether was circulated to the extraction column T4 for further use; whereas CO₂ gas discharged through decomposition was used for acidification and crystallization in continuous crystallizer T6 after compression. The potassium carbonate solution containing mother solution discharged from the column bottom was cooled prior to circulation to HCN absorption column T1 for absorption of formonitrile.

Embodiment 5

A metering liquid delivery pump was used to deliver treated crystallization mother solution (10% K_2CO_3 solution was used at the initial start-up) to the HCN absorption column T1 at the flow rate of 838Kg/h. Meanwhile, 5.5Kg/h HCN gas was delivered to the HCN absorption column for preparation of the KCN solution in the HCN absorption column T1, and a liquid delivery pump was used to deliver the KCN solution to the first tubular reactor R1. A liquid delivery pump was used to deliver ammonium bicarbonate absorbing ammonia gas and carbon dioxide in the CO_2NH_3 absorption column T2 to the first tubular reactor R1 (16% ammonia bicarbonate solution was used at the initial start-up) at the velocity of 220Kg/h, and a liquid delivery pump was used to deliver 3-methylthio propionaldehyde to the first tubular reactor R1 at the velocity of 20.8Kg/h simultaneously. A reaction fluid stayed for 3 minutes respectively at the temperature of 60°C and 100°C, and stayed for another 4 minutes at the temperature of 140°C. The as discharged reaction fluid from the first tubular reactor R1 was subject to decompression in the decompression tank R2 to discharge surplus ammonia gas and carbon dioxide to the CO_2 and NH_3 absorption column T2 to obtain 1062Kg/h 5-(β -methylthioethyl)glycolurea solution.

A liquid delivery pump was used to deliver the prepared 5-(β -methylthioethyl)glycolurea solution to the second tubular reactor R3 at the velocity of 1062Kg/h. A reaction fluid stayed in the second tubular reactor for 9 minutes at the temperature of 190°C for saponification prior to decompression in the T3 absorption column. Ammonia gas and carbon dioxide produced during reaction were discharged to the CO_2 and NH_3 absorption column T2 for absorption. The saponification liquid was cooled down to the temperature of 30°C in the heat exchanger, which was further delivered to the continuous extraction column T4 via the pump at the velocity of 1056Kg/h. Meanwhile, 1200Kg/h ethylene glycol dipropyl ether was delivered to the continuous extraction column T4 for continuous reversed extraction. Continuous differential distillation of the ethylene glycol dipropyl ether layer was conducted in the organic solvent recycling distillation column T5 to recycle the solvent for circulated use during extraction process, and incinerate tailings at the same time. A water layer was delivered to continuous crystallizer T6 at the velocity of 1050Kg/h, and simultaneously CO_2 gas was delivered to continuous crystallizer T6 for acidification until PH value of the reaction system was up to 9. Continuous crystallization was conducted at the temperature of 30°C, and the reaction fluid stayed in continuous crystallizer for 3 hours. The crystallization liquid was subject to continuous centrifugal solid-liquid

separation in the first solid-liquid separator S1 to filtrate crude D,L-methionine product, and the crystallization mother solution was prepared for further use. A solid delivery equipment was used to deliver crude D, L-methionine product to the scrubber R4 for spiral agitation and washing in the specified quantity at the velocity of 35Kg/h. Meanwhile, 60Kg/h water was continuously delivered to the scrubber R4 to control temperature of materials in the scrubber R4 below 30°C. The solid material stayed in the scrubber R4 for 10 minutes. The intermixture subjected to continuous agitation and washing in the scrubber R4 was under continuous centrifugal solid-liquid separation via the second solid-liquid separator S2. A washing filtrate was poured into the crystallization mother solution, and the filter cake of D,L-methionine as filtered was airflow dried at the temperature of 130°C under the protection of N₂ to obtain 27.8Kg/h finished D, L-methionine product. Inspection according to GB-T17810-2009 Feed Grade DL-Methionine was conducted to obtain the result that methionine content was up to 99.3%, and total molar reaction yield as based on 3- methylthio propionaldehyde was up to 93.3%.

The crystallization mother solution and washing filtrate were delivered to the KHCO₃ decomposition column T7 at the velocity of 1106Kg/h, and they stayed for 3 hours at the temperature of 120°C to decompose potassium hydrogen carbonate contained in the mother solution into potassium carbonate. Meanwhile, 245Kg/h water and 7Kg/h 2-ethylhexyl alcohol were prepared through distillation at the column top. The distilled water was circulated to the scrubber R4 at 60Kg/h for agitation prior to further use; whereas remaining 185Kg/h distilled water was used in the CO₂ and NH₃ absorption column T2. Distilled 2-ethylhexyl alcohol was circulated to the extraction column T4 for further use; whereas CO₂ gas discharged through decomposition was used for acidification and crystallization in continuous crystallizer T6 after compression. The potassium carbonate solution containing mother solution discharged from the column bottom was cooled prior to circulation to the HCN absorption column T1 for absorption of formonitrile.

CLAIMS

1. A clean method for preparing D, L-methionine, comprising the following steps:

(1) preparing a 5-(β -methylthioethyl)glycolyurea reaction fluid by using 3-methylthio propionaldehyde, KCN and NH_4HCO_3 solutions as raw materials for continuous reaction in the tubular reactor that is gradually heated;

(2) decompressing the 5-(β -methylthioethyl)glycolyurea reaction fluid to separate NH_3 and CO_2 for delivery to another tubular reactor for decomposition and saponification to obtain a saponification reaction fluid;

(3) eliminating NH_3 and CO_2 produced during saponification through desorption of the saponification fluid to obtain a D,L-methionine potassium solution;

(4) obtaining a NH_4HCO_3 solution through water spray and absorption of NH_3 and CO_2 as separated in steps (2) and (3), and directly using the NH_4HCO_3 solution as a material for preparation of the 5-(β -methylthioethyl)glycolyurea reaction fluid in step (1);

(5) cooling down the D,L-methionine potassium solution in step (3), and then using an organic solvent for reversed and continuous extraction; directly using an organic solvent recycled through distillation of an organic solvent layer for reversed and continuous extraction; raffinate produced by distillation being treated as waste fluid;

(6) delivering water that is produced through counter-current and continuous extraction in step (5) to a continuous crystallizer, simultaneously delivering CO_2 gas to the continuous crystallizer for acidification to control PH value of reaction fluid in the continuous crystallizer at 6-9, and then obtaining a crystallization intermixture;

(7) separating solid and liquid in the crystallization intermixture to obtain crude D,L-methionine product and a crystallization mother solution; washing crude D,L-methionine product with water for solid-liquid separation to obtain a filter cake of D,L-methionine and washing filtrate;

(8) drying the filter cake of D,L-methionine under the protection of inert gas to obtain D,L-methionine product;

(9) combining the crystallization mother solution and the washing filtrate from step (7) for delivery to a decomposition column to fully decompose KHCO_3 into K_2CO_3 for release of CO_2 gas; released CO_2 gas that is subjected to compression is delivered to continuous crystallizer for

acidification and crystallization;

(10) using the K_2CO_3 containing mother solution as obtained in step (9) for absorption of HCN gas to prepare a KCN solution that is the material for preparation of 5-(β -methylthioethyl)glycolyurea in step (1).

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2. The clean method for preparing D,L-methionine according to Claim 1, characterized in that, in the step (1), a mole ratio of 3-methylthio propionaldehyde, KCN and NH_4HCO_3 is 1:1-1.1:2-3;

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3. The clean method for preparing D,L-methionine according to Claim 1, characterized in that, in the step (1), a gradual heating range is 50-150°C with reaction time up to 3-15 minutes.

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4. The clean method for preparing D,L-methionine according to Claim 1, characterized in that, in the step (2), the temperature for decomposition and saponification is 140-220°C with time up to 2-15 minutes.

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5. The clean method for preparing D,L-methionine according to Claim 1, characterized in that, in the step (5), the temperature of the D, L-methionine potassium solution is reduced to 0-40°C before using the organic solvent equivalent to 0.5-2 times the weight of the D, L-methionine potassium solution for reversed and continuous extraction.

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6. The clean method for preparing D,L-methionine according to Claim 1, characterized in that, in the step (5), the organic solvent for reversed and continuous extraction includes one or more of toluene, ethylbenzene, dimethylbenzene, n-butyl alcohol, isobutanol, n-pentanol, 2-methyl-1-butanol, isoamylol, sec.-amyl alcohol, 3-pentanol, tert-amyl alcohol, n-hexyl alcohol, 4-methyl-2-pentanol, 2-ethyl alcohol, 2-methyl amyl alcohol, heptanol, 2-heptanol, 3-heptanol, 2-ethylhexyl alcohol, 2-octanol, octanol, 3,5,5-trimethylhexanol, ether, methyl tertiary-butyl ether, isopropyl ether, n-propyl ether, n-butyl ether, isoamyl ether, hexyl ether, 2-methyl tetrahydrofuran, anisole, ethoxybenzene, 3-methylanisole, ethyl benzyl ether, ethylene glycol diethyl ether, diethylene dipropyl ether and ethylene glycol di-butyl ether.

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7. The clean method for preparing D,L-methionine according to Claim 1, characterized in that, in the step (6), the temperature for continuous crystallization is 0-40°C, whereas the reaction fluid stays inside the continuous crystallizer for 0.5-5 hours.

5 8. The clean method for preparing D,L-methionine according to Claim 1, characterized in that, in the step (9), the temperature and pressure of decomposition column are 110-160°C and 0.15-0.8MPa with decomposition reaction time up to 1-4 hours.

10 9. The clean method for preparing D,L-methionine according to Claim 5, characterized in that, in the step (5), the organic solvent for reversed and continuous extraction includes one or more of toluene, ethylbenzene, dimethylbenzene, n-butyl alcohol, isobutanol, n-pentanol, 2-methyl-1-butanol, isoamylol, sec.-amyl alcohol, 3-pentanol, tert-amyl alcohol, n-hexyl alcohol, 4-methyl-2-pentanol, 2-ethyl alcohol, 2-methyl amyl alcohol, heptanol, 2-heptanol, 3-heptanol, 2-ethylhexyl alcohol, 2-octanol, octanol, 3,5,5-trimethylhexanol, ether, methyl tertiary-butyl
15 ether, isopropyl ether, n-propyl ether, n-butyl ether, isoamyl ether, hexyl ether, 2-methyl tetrahydrofuran, anisole, ethoxybenzene, 3-methylanisole, ethyl benzyl ether, ethylene glycol diethyl ether, diethylene dipropyl ether and ethylene glycol di-butyl ether.