including the cellulose components of tobacco. Any suitable ionic liquid may be used as a solvent to dissolve tobacco. As used herein, an "ionic liquid" is an ionic compound in a liquid state. An ionic compound may be a compound having positively and negatively charged moieties, such as *N*-methylmorphione-*N*-oxide (NMMO), or a salt. Suitable ionic liquids typically have melting points of about 150°C or less, such as about 100°C or less. In embodiments, the ionic liquid has a melting temperature of about 40°C or less, such as about 25°C or less, about 23°C or less, about 20°C or less, about 15°C or less, about 10°C or less, about 5°C or less, about 0°C or less, about -10°C or less, about -20°C or less, or about -30°C or less. Ionic liquids are typically liquid over a wide temperature range from the melting point to the decomposition temperature. Preferably, the ionic liquids are liquid at room temperature; i.e., have a melting temperature of less than about room temperature, which is typically considered to be between about 20°C and about 25°C. More preferably, the ionic liquids are liquid at temperatures 10°C or more below room temperature, such as below about 15°C, or below about 10°C. Most preferably, the ionic liquids are liquid at a temperature of below about 0°C, such as below about -10°C or below about -20°C. By way of example, one suitable ionic liquid, 1-ethyl-3-methylimidazolium acetate, has a melting temperature of about -20°C.

In many embodiments, ionic liquids are salts. Examples of cation moieties of ionic liquid salts include cyclic and acyclic cations. Cyclic cations include pyridinium, imidazolium, and imidazole. Acyclic cations include alkyl quaternary ammomnium and alkyl quaternary phosphorous cations. Substituent groups, (i.e. R groups), on the cations can be C_1, C_2, C_3, and C_4, which may be saturated or unsaturated. The ionic liquid salt may have any suitable counter anion. In embodiments, counter anions of the cation moiety are selected from the group consisting of halogen, pseudohalogen and carboxylate. Carboxylates include acetate, citrate, malate, maleate, formate, and oxylate. Halogens include chloride, bromide, zinc chloride/choline chloride, 3-methyl-N-butyl-pyridinium chloride and benzyldimethyl (tetradecyl) ammonium chloride.

Examples of compounds which are ionic liquids and which may be used to dissolve tobacco include, but are not limited to, NMMO, 1-ethyl-3-methylimidazolium chloride, 1-allyl-3-methylimidazolium chloride, 1-butyl-3-methylimidazolium chloride, 1-ethyl-3-methylimidazolium methanesulfonate, 1-ethyl-3-methylimidazolium diethylphosphate, 1,3-dimethylimidazolium dimethylphosphate, 1-ethyl-3-methylimidazolium acetate, ~~1-butyl-3-methylimidazolium chloride,~~ 1-butyl-3-methylimidazolium methylcarbonate, and tris-(2-hydroxyethyl)-methylammonium methylsulfate.

Tobacco may be dissolved, fully or partially, in an ionic liquid at any suitable concentration to obtain a liquid tobacco composition. As used herein, a "liquid tobacco