

Process for Producing 2-Hydroxy-4-Methylthiobutanoic Acid

BACKGROUND OF THE INVENTION

Field of the Invention

5 [0001]

The present application is filed, claiming the Paris Convention priorities of Japanese Patent Application No. 2008-263734 (filed on October 10, 2008), the entire content of which is incorporated herein by reference.

10 The present invention relates to a process for producing 2-hydroxy-4-methylthiobutanoic acid from 2-hydroxy-4-methylthiobutanenitrile via 2-hydroxy-4-methylthiobutanamide. In particular, the invention relates to a material for an apparatus for use in a solid-liquid
15 separation step which comprises neutralizing a reaction solution which contains 2-hydroxy-4-methylthiobutanoic acid, and separating the same to obtain an organic phase; concentrating the organic phase; and separating from the organic phase a residue containing an inorganic salt, which
20 is left to remain in the organic phase.

Description of the Related Art

[0002]

2-Hydroxy-4-methylthiobutanoic acid useful as an
25 additive to a feed is known to be obtained by a method

which comprises the steps of hydrating 2-hydroxy-4-methylthiobutanenitrile in the presence of sulfuric acid to obtain 2-hydroxy-4-methylthiobutanamide, and hydrolyzing 2-hydroxy-4-methylthiobutanamide to obtain 2-hydroxy-4-methylthiobutanoic acid (cf. JP-A-2007-238555).

[0003]

A nickel-based alloy such as Hastelloy®C-22 is generally used, because of its corrosion resistance, as a material for an apparatus for use in a solid-liquid separation step wherein a reaction solution containing 2-hydroxy-4-methylthiobutanoic acid, obtained by hydration and hydrolysis reactions, is neutralized and phase-separated to obtain an organic phase, which is then concentrated, and a residue containing an inorganic salt, left to remain in the organic phase, is separated. Such alloy contains costly but essential elements of Mo and Ni in higher amounts.

SUMMARY OF THE INVENTION

[0004]

An object of the present invention is to provide a method of using corrosion resistant material containing lesser amounts of such elements for an apparatus for use in a solid-liquid separation step wherein a reaction solution containing 2-hydroxy-4-methylthiobutanoic acid, obtained by

hydration and hydrolysis of 2-hydroxy-4-methylthiobutanenitrile, is neutralized and phase-separated to obtain an organic phase, which is then concentrated, and a residue containing an inorganic salt, left to remain in the organic phase, is separated. Thus, the present method is cost effective.

[0005]

The present invention provides a process for producing 2-hydroxy-4-methylthiobutanoic acid, comprising carrying out

a solid-liquid separation step of separating a residue containing an inorganic salt from an organic phase concentrate containing 2-hydroxy-4-methylthiobutanoic acid, and the residue with an apparatus composed of, as a material,

an alloy which contains 21.0 to 30.0% by weight of a Cr element, 2.5 to 11.0% by weight of a Ni element, 1.0 to 5.0% by weight of a Mo element and a Fe element as the rest, at a temperature of from 60 to 95°C.

[0006]

According to the present invention, there is provided a more inexpensive and corrosion resistant material for an apparatus for use in a solid-liquid separation step wherein a reaction solution containing 2-hydroxy-4-methylthiobutanoic acid, obtained by hydration and

hydrolysis of 2-hydroxy-4-methylthiobutanenitrile, is neutralized and phase-separated to obtain an organic phase, which is then concentrated, and a residue containing an inorganic salt, left to remain in the organic phase, is
5 separated; and thus, 2-hydroxy-4-methylthiobutanoic acid can be economically produced.

Detailed Description of the Invention

[0007]

10 (Production Step)

2-Hydroxy-4-methylthiobutanoic acid (hereinafter optionally referred to as HMTBA) is produced by hydrating 2-hydroxy-4-methylthiobutanenitrile (hereinafter optionally referred to as HMTBN) in the presence of sulfuric acid
15 usually at a temperature of from about 40 to about 70°C to obtain 2-hydroxy-4-methylthiobutanamide (hereinafter optionally referred to as HMTBAA), and adding water to 2-hydroxy-4-methylthiobutanamide to thereby hydrolyze the same usually at a temperature of from about 90 to about
20 130°C.

For example, HMTBN is industrially produced by reacting acrolein with methyl mercaptan to obtain 3-methylthiopropionaldehyde, and reacting this 3-methylthiopropionaldehyde with hydrogen cyanide.

25 [0008]

Generally, HMTBAA is produced by feeding HMTBN, water and sulfuric acid to a hydration tank to hydrate HMTBN.

Fed to the hydration tank are usually about 20 to about 70 parts by weight, preferably about 25 to about 50 parts by weight of water per 100 parts by weight of HMTBN, and about 0.5 to about 1 mol in total, preferably about 0.6 to about 0.8 mol in total, of sulfuric acid per 1 mol of the HMTBN.

[0009]

10 It is also possible to feed water in the form of a pre-mixture thereof with HMTBN and/or sulfuric acid, in other words, as an aqueous HMTBN solution and/or an aqueous sulfuric acid solution, to the hydration tank. Preferably, water is fed as an aqueous sulfuric acid solution to the
15 hydration tank.

The hydration reaction is carried out usually at a temperature of about 40 to about 70°C for about 1 to about 3 hours. After the reaction, the reaction solution is typically aged. In the reaction solution, HMTBA is
20 produced as a result of hydrolysis of a part of HMTBAA.

[0010]

Then, the reaction solution containing HMTBAA as a main component, obtained in the hydration tank, is typically fed to a hydrolysis tank, and water is added to
25 the reaction solution to thereby hydrolyze HMTBAA to

produce HMTBA.

About 100 to about 200 parts by weight of water is usually fed per 100 parts by weight of the aqueous sulfuric acid solution in the above-described reaction solution.

5 In the hydrolysis tank, HMTBAA is typically hydrolyzed with water and sulfuric acid to produce HMTBA and concurrently to by-produce ammonium bisulfate (NH_4HSO_4) and ammonium sulfate ($(\text{NH}_4)_2\text{SO}_4$). After the addition of water and heating, the hydrolysis reaction is carried out at a
10 temperature of from about 90 to about 130°C for about 2 to about 6 hours.

The hydration and hydrolysis reactions are typically carried out, using a glass-lined, resin-lined, MAT®21-made or AlloyB-2 (Hastelloy® B-2)-made apparatus.

15 [0011]

The reaction solution containing HMTBA, obtained in the hydrolysis tank, is usually distilled to remove low boiling point components from the reaction solution.

The distillation is carried out usually at a
20 temperature of from about 80 to about 120°C under a pressure of from about 50 to about 150 kPa to remove by-produced low boiling point components such as dimethyl sulfide, dimethyl disulfide and formic acid. The low boiling point components are removed as a distillate
25 usually at a rate of about 1 to about 4% by weight relative

to the reaction solution, as required.

The removal of the low boiling point components may be done after neutralization or phase separation as will be described later.

5 [0012]

(Separation Step)

Then, an alkali is added to the reaction solution containing HMTBA, from which the low boiling point components have been removed, to thereby neutralize the
10 reaction solution, and the reaction solution is phase-separated into an organic phase containing HMTBA and an aqueous phase containing water and an inorganic salt (containing ammonium bisulfate and ammonium sulfate). The neutralization and the phase separation are conducted, for
15 example, using a mixer-settler type liquid-liquid extractor in which a stirrer tank and a phase-separation tank are combined as one set.

[0013]

As the alkali, for example, sodium hydroxide, sodium
20 hydrogencarbonate, sodium carbonate or the like is used in the form of an aqueous solution. The alkali is fed typically at a rate of from about 0.5 to about 1.2 mol, preferably from about 0.6 to about 0.8 mol, per 1 mol of ammonium bisulfate in the above-described reaction solution.
25 The addition rate of the alkali may be controlled by a

hydrogen ion concentration (pH) of the reaction solution admixed with the alkali.

The neutralization reaction is carried out usually at a temperature of from about 15 to about 120°C, preferably
5 from about 30 to about 110°C, for about 0.1 to 3 hours, preferably about 0.1 to about 2 hours.

[0014]

After the neutralization, the reaction solution is left to stand still in the phase-separation tank so as to
10 phase-separate the reaction solution into an organic phase as an upper layer and an aqueous phase as a lower layer (liquid separation). The temperature for this separation is usually from about 30 to about 110°C.

[0015]

15 (Concentration Step)

The organic phase separated from the aqueous phase typically contains about 40 to about 60% by weight of HMTBA, about 20 to about 30% by weight of water and about 10 to about 30% by weight of an inorganic salt.

20 This organic phase is concentrated to remove the remaining water. For example, the concentration is done in a concentration tank usually at a temperature of from about 60 to about 150°C under a pressure of from about 1 to about 20 kPa, so that the water in the organic phase is reduced
25 to about 5% by weight or less, preferably about 2% by

weight or less, more preferably about 1% by weight or less. By this concentration, the sulfate ion concentration and kinematic viscosity of a product as will be described later can be decreased.

5 The inorganic salt is precipitated from the organic phase by this concentration, and the resulting organic phase forms a slurry.

[0016]

When the residence time of the organic phase in the
10 concentration tank is usually set to about 0.5 hour or longer, HMTBA is converted into an oligomer (which is mainly a dimer containing small amounts of a trimer and a tetramer). By doing so, the solubility of the inorganic salt in the organic phase is lowered, so that the
15 concentration of the inorganic salt in the organic phase can be lowered. Preferably, the temperature, pressure and residence time are selected so that the weight ratio of the monomer of HMTBA to the oligomer of HMTBA can be about 2 to about 4. Such a concentration is effective to increase the
20 particle size of the inorganic salt precipitated in the organic phase, so that a solid-liquid separation efficiency is improved during a solid-liquid separation as will be described later. Thus, the removal of the inorganic salt is facilitated.

25 [0017]

(Solid-Liquid Separation Step)

Then, the resultant slurry of the organic phase is cooled to usually about 60 to about 95°C, using a heat exchanger or the like. After that, the slurry is separated
5 into a liquid component containing the organic phase and a solid component (or a residue) containing the precipitated inorganic salt.

The apparatus for use in the solid-liquid separation step refers to a solid-liquid separator, a pipe or a heat
10 exchanger with which the slurry of the organic phase is brought into contact.

[0018]

In the present invention, as a material for the apparatus for use in the solid-liquid separation step,
15 employed is an alloy that contains 21.0 to 30.0% by weight of a Cr element, 2.5 to 11.0% by weight of a Ni element, 1.0 to 5.0% by weight of a Mo element and a Fe element as the rest.

SUS329J4L, SUS329J1, SUS329J3L, SCS10, and SCS11
20 exemplify commercially available alloys that contain 21.0 to 30.0% by weight of a Cr element, 2.5 to 11.0% by weight of a Ni element, 1.0 to 5.0% by weight of a Mo element and a Fe element as the rest .

[0019]

25 A conventional nickel-based alloy contains about 13 to

about 15% by weight of a Mo element and a Ni element as the rest (about 61 to about 65% by weight), while the alloy to be used in the present invention contains so small an amount as 1 to 5% by weight of a Mo element and a Fe
5 element as the rest, and thus is inexpensive.

[0020]

The solid-liquid separation is carried out typically at a temperature of from about 60 to about 95°C, preferably from about 60 to 90°C. When this separation is carried out
10 at a temperature higher than about 95°C, the above-described alloy undesirably tends to corrode. When this separation is carried out at a temperature lower than about 60°C, the viscosity of the slurry of the organic phase becomes higher, since the slurry obtained by the
15 concentration in the concentration step has a relatively high viscosity and thus is low in filtration efficiency and deliquoring property. Thus, this is undesirable, since the solid-liquid separation becomes difficult.

[0021]

20 As a solid-liquid separator, a centrifugal separator is usually used. The centrifugal separator may be of cylindrical type, separator plate type, decanter type or the like, among which a decanter type centrifugal separator is preferably used.

25 [0022]

The separated liquid component comprises HMTBA (containing the oligomer produced during the concentration) as a main component. If needed, water is added to the liquid component to provide a product of HMTBA which typically contains about 88 to about 90% by weight of HMTBA, about 10 to about 12.5% by weight of a moisture content and very small amounts of other components.

[0023]

[Recovering Step]

10 The separated solid component contains about 20 to about 60% by weight of HMTBA, and thus is usually admixed with water to dissolve the inorganic salt, so as to phase-separate the solid component into an aqueous phase containing the inorganic salt and an organic phase. This organic phase is recovered. As the water to be added, the aqueous phase separated in the above-described separation step is preferably used. The organic phase to be recovered is mixed into the organic phase obtained during the above-described neutralization and phase-separation in the above-described separation step, so as to be recovered.

Examples

[0024]

An organic phase (HMTBA: 75% by weight, HMTBA dimer: 8.5% by weight, $(\text{NH}_4)_2\text{SO}_4$: 10% by weight, NH_4HSO_4 : 0.5% by

weight, and NaNH_4SO_4 : 6% by weight) to be treated in the solid-liquid separation step was subjected to a corrosion test for the following metallic material.

The organic phase and a test piece of the following
5 metallic material (a flat plate of 25 mm in length X 20 mm in width X 2 mm in thickness was U-like bent, and a stress was applied thereto (a U-bent test piece)) were put in a test container (a glass-made test bottle with a condenser used at 90°C or lower, or a glass-made autoclave used at
10 100°C or higher); the gas phase zone was purged with a nitrogen gas; and the test container was heated at a temperature indicated in Table 1 and was maintained at the same temperature for 168 hours. After that, the test piece was removed, rinsed and dried. Then, the reduced weight of
15 the test piece was measured to determine a corrosion rate. The results are shown in Table 1.

[0025]

(Test Piece)

SUS329J4L (Cr: 24.9% by weight, Ni: 7.1% by weight,
20 Mo: 3.1% by weight, N: 0.16 part by weight, and Fe: the rest)

AlloyC-22 (Hastelloy® C-22) (Cr: 21.4% by weight, Mo: 13.3% by weight, and Ni: the rest)

AlloyC-276 (Hastelloy® C-276) (Cr: 15.1% by weight, Mo:
25 15.7% by weight, W: 3.6% by weight, and Ni: the rest)

Alloy59 (Cr: 22.9% by weight, Mo: 15.7% by weight, and Ni: the rest)

[0026]

[Table 1]

No.	Test piece	Temp. (°C)	Corrosion rate (mm/year)	Corroded state
1	SUS329J4L	90	0.00	Passivated surface
2	SUS329J4L	100	0.10	Active corroded surface
3	SUS329J4L	120	0.06	Active corroded surface
4	AlloyC-22	120	0.00	Passivated surface
5	AlloyC-276	120	0.00	Passivated surface
6	Alloy59	120	0.00	Passivated surface

5

[0027]

AlloyC-22, AlloyC-276 and Alloy59 which were nickel-based alloys still had corrosion resistance even at 120°C.

Although SUS329J4L corroded at 100°C or higher, it did not corrode at 90°C.

Therefore, the use of an alloy which contains 21.0 to 30.0% by weight of a Cr element, 2.5 to 11.0% by weight of a Ni element, 1.0 to 5.0% by weight of a Mo element and a Fe element as the rest, as a material for the apparatus for use in the solid-liquid separation step at a temperature of from 60 to 95°C, is effective to realize economical production of 2-hydroxy-4-methylthiobutanoic acid without any corrosion of the apparatus.

WHAT IS CLAIMED IS:

1. A process for producing 2-hydroxy-4-methylthiobutanoic acid, comprising carrying out

a solid-liquid separation step of separating a residue
5 containing an inorganic salt, from an organic phase concentrate containing 2-hydroxy-4-methylthiobutanoic acid and the residue, with an apparatus composed of, as a material,

an alloy which contains 21.0 to 30.0% by weight of a
10 Cr element, 2.5 to 11.0% by weight of a Ni element, 1.0 to 5.0% by weight of a Mo element and a Fe element as the rest, at a temperature of from 60 to 95°C.

2. A process for producing 2-hydroxy-4-methylthiobutanoic acid, which comprises
15

a production step of producing 2-hydroxy-4-methylthiobutanoic acid by adding water and an acid to 2-hydroxy-4-methylthiobutanenitrile and/or 2-hydroxy-4-methylthiobutanamide;

20 a separation step of neutralizing a reaction solution which contains 2-hydroxy-4-methylthiobutanoic acid, obtained in the production step, by adding an alkali thereto, and phase-separating said reaction solution into

an organic phase containing 2-hydroxy-4-methylthiobutanoic acid and an aqueous phase containing water and an inorganic salt;

a concentration step of concentrating 2-hydroxy-4-methylthiobutanoic acid by removing the remaining water from the organic phase separated in the separation step; and

a solid-liquid separation step of separating a residue containing an inorganic salt, from the organic phase concentrate containing 2-hydroxy-4-methylthiobutanoic acid and the residue, with an apparatus composed of, as a material, an alloy which contains 21.0 to 30.0% by weight of a Cr element, 2.5 to 11.0% by weight of a Ni element, 1.0 to 5.0% by weight of a Mo element and a Fe element as the rest, at a temperature of from 60 to 95°C.

3. The process of Claim 1 or 2, which comprises a recovering step of mixing the residue removed in the solid-liquid separation step, with the aqueous phase separated in the separation step; phase-separating the mixture into an organic phase containing 2-hydroxy-4-methylthiobutanoic acid and an aqueous phase containing water and an inorganic salt; and recovering the separated organic phase.

4. The process of Claim 1 or 2, wherein the alloy which contains 21.0 to 30.0% by weight of a Cr element, 2.5 to 11.0% by weight of a Ni element, 1.0 to 5.0% by weight of a Mo element and a Fe element as the rest is SUS329J4L,
5 SUS329J1, SUS329J3L, SCS10 or SCS11.

5. Use of an alloy as a material of an apparatus in a process for producing 2-hydroxy-4-methylthiobutanoic acid comprising a solid-liquid separation step of separating a
10 residue containing an inorganic salt, from an organic phase concentrate containing 2-hydroxy-4-methylthiobutanoic acid and the residue, wherein said alloy is

an alloy which contains 21.0 to 30.0% by weight of a Cr element, 2.5 to 11.0% by weight of a Ni element, 1.0 to
15 5.0% by weight of a Mo element and a Fe element as the rest.