

APPARATUS AND METHOD FOR ELECTROCHEMICAL DETECTION

Field of the Invention

[0001] The present invention is directed to a sensor with opposing electrodes and test strips using this sensor. The invention can be used to measure blood or plasma coagulation in assays like prothrombin time (PT) and thrombin potential, for example, in point-of-care monitoring of anticoagulants.

Background

[0002] Various patents discuss devices and methods of detecting electrochemical substrates. For example, Jozefonvicz *et al.* (US Patent No. 4,304,853) describes the use of electrochemical substrates to detect proteases, including the enzymes in a coagulation system. Unfortunately, their method is not practical for point-of-care devices. The teaching, as used herein, is useful for the finding that preconditioning of the electrode(s) is necessary to improve reproducibility because the measured currents are very low. The substrate requires DMSO to assist with solubility in aqueous solutions. It also discloses a 3-electrode system with volumes of sample and reagents that are orders of magnitude greater than that required for viable point-of-care tests.

[0003] Ludin *et al.* (US 6,495,336) disclose electrochemical substrates that can be used for detection of thrombin, typically in a coagulation assay. The substrate provided therein is an improvement over that of Jozefonvicz *et al.* because it is apparently water soluble. They do not disclose how to use the substrate in a point-of-care device.

[0004] Frenkel *et al.* (US 6,352,630) discloses use of an electrochemical thrombin substrate, specifically as disclosed by Ludin *et al.* (2002), in a device that appears to be suitable for point-of-care testing. The teaching, as used herein, provides that the substrate is to be immobilized on the working electrode via the end that is opposite the electroactive group that is cleaved off. This requirement adds complexity and limits the electrode materials and fabrication process.

[0005] Immobilization of the substrate on the working electrode is important and/or essential when using the coplanar electrodes that Frenkel *et al.* teach. This is because the

relatively large distance between the working and counter electrodes prohibits any meaningful diffusion of molecules from one electrode to the other.

[0006] Similarly, the teaching provides a Ag/AgCl reference electrode, which may also be a counter electrode, and is desired for coplanar electrode configurations because the Ag/AgCl can act as a sink or source of electrons. A potential issue of this device is that the relatively small amount of electroactive product that is released, say in the reduced form, in one embodiment, can only be detected when it is oxidized. This may result in relatively low electrochemical currents. Furthermore, the preparation of a Ag/AgCl electrode adds to the expense of manufacturing such sensors.

[0007] Unkrig *et al.* (US 2003/0146113A1) also disclose the use of an electrochemical thrombin substrate that appears suitable for point-of-care devices. As used herein, the reference describes a flat 2-electrode system with a Pd working electrode and a Ag/AgCl combined reference and counter electrode. In order to overcome the inherent low signal in this coplanar electrode design in certain embodiments the teaching regenerates the electroactive group by coupling its detection to the oxidation of glucose. An advantage of certain present embodiments is that the opposing electrodes can automatically regenerate the detected species without the need for coupling to another redox reaction.

[0008] Opalsky *et al.* (EP 1,234, 053 B1) also describe a coagulation sensor that uses an electrochemical substrate. In certain embodiments, the teaching discloses a complex cartridge that preferably has a pump to move the sample within the cartridge to mix the sample plus reagent past coplanar electrodes. In contrast, in certain embodiments herein, the opposing electrodes permit a simple strip that does not benefit from a pump as reagent can passively diffuse throughout the smaller volume of sample.

[0009] Opposing electrode biosensors are disclosed by Hodges *et al.* (US 6,284,125 B1) to measure the concentration of a redox species. In particular, a redox species generated by the oxidation of glucose and from this the glucose concentration was calculated. An opposing electrode sensor can consist of:

- (1) Working and counter electrodes (and optionally a reference electrode in any configuration) spaced by a predetermined distance and;

- (2) The electrodes facing each other while lying on different but approximately parallel planes and;
- (3) Selecting the spacing between the working electrode and the counter electrode so that reaction products from the counter electrode can diffuse to the working electrode.

[0010] Opposing electrodes are disclosed by Newman and Chatelier (PCT/1B2007/001990) for a coagulation sensor but the disclosure is limited to using magnetic particle (im)mobility to detect coagulation and makes no teaching of electrochemical substrates.

[0011] Ohara *et al.* (US 6,620,310 B1) describe detecting a change in viscosity of a coagulating sample by detecting the decrease in steady state current through opposing electrodes. Again the disclosure does not relate to an electrochemical substrate, instead the change in diffusion of a redox couple is detected upon coagulation.

[0012] None of the prior art describes the use of opposing electrodes to detect an electrochemical substrate. Nor do they describe the unexpected discovery that when using opposing electrodes it is preferable not to immobilize the electrochemical substrate on the working electrode.

Summary

[0013] Embodiments of the present invention are directed to an opposing-electrode sensor with a working electrode, a counter electrode, and an electrochemical substrate, wherein the sensor detects cleavage of the electrochemical substrate. The electrochemical substrate can be coated on the counter electrode. These sensors can be used as part of a test strip for testing coagulation times in blood.

[0014] Other embodiments of the invention are directed to methods of using the sensors and strips of the present invention. These methods can be practiced at the point of care.

Brief Description of the Figures

[0015] FIG. 1 illustrates in exploded view an exemplary electrochemical strip in accordance with the present embodiments, including a bottom electrode (1), an insulating separator (2), a top electrode (3), a cover over a fill channel (4), a fill channel (5) and a reaction chamber (6).

[0016] FIG. 2 illustrates an exemplary assembled electrochemical strip in accordance with the present embodiments, including a bottom electrode (1), an insulating separator (2), a top electrode (3), a cover over a fill channel (4), a fill channel (5) and a reaction chamber (6).

[0017] FIG. 3 illustrates an exemplary graph in accordance with the present embodiments; the figure demonstrates improved clot detection when thrombin substrate is coated on the counter electrode; time in seconds is provided as abscissa, and current in microamps is provided as ordinate; the graph labeled “a” represents thrombin substrate coated on counter electrode (the cathode in this example) and thromboplastin coated on working electrode (the anode in this example); the graph labeled “b” represents thrombin substrate coated on working electrode (anode) and thromboplastin coated on counter electrode (cathode).

Detailed Description

[0018] Exemplary embodiments of the invention are discussed in detail below. While specific exemplary embodiments are discussed, it should be understood that this is done for illustration purposes only. A person skilled in the relevant art will recognize that other components and configurations may be used without parting from the spirit and scope of the invention. References to “one embodiment,” “an embodiment,” “example embodiment,” “various embodiments,” etc., may indicate that the embodiment(s) of the invention so described may include a particular feature, structure, or characteristic, but not every embodiment necessarily includes the particular feature, structure, or characteristic. Further, repeated use of the phrase “in one embodiment,” or “in an exemplary embodiment,” do not necessarily refer to the same embodiment, although they may.

[0019] The present invention is directed to a sensor having opposing electrodes (e.g., a working and a counter electrode) and an electrochemical substrate. In some embodiments,

the electrochemical substrate is on the counter electrode. As will be described to follow, these sensors can be used by adding a sample to the sensor (optionally on a test strip) where a protease cleaves an electroactive species off the substrate. This electroactive species can be detected by redox reactions at the electrodes.

[0020] The invention can be used in a wide variety of technologies, for example, in methods of measuring blood or plasma coagulation using assays like prothrombin time (PT) and thrombin potential, particularly in point-of-care monitoring of oral anticoagulants (e.g. coumadins like warfarin and phenprocoumon).

[0021] As discussed more below, one advantage of the present invention is that the opposing electrodes of certain embodiments do not need the complexity of a true reference electrode such as a Ag/AgCl electrode. Instead, in these embodiments, the close proximity of the electrodes permits cycling of an electroactive species. That is, a molecule that has been reduced at the cathode can be reoxidised at the anode, resulting in a higher current. This can make it is possible to use two similar electrodes such as those formed from an inert conducting material such as carbon, Au, Pd, Pt Ir, indium oxide, mixed indium/tin oxide, and/or tin oxide, etc.

[0022] The devices and test strips of the present invention have an electrochemical substrate which can be dried on the counter electrode in some embodiments. The electrochemical substrate can be cleaved by an enzyme, for example, a protease. Proteases are a class of enzymes that cleave peptide bonds. They have an essential role in many biological processes. For example, they are involved with food digestion, HIV replication, the blood coagulation cascade and the complement pathway.

[0023] Some proteases are specific for a particular peptide sequence while others are less selective. In either case, it can be highly desirable to detect and/or quantify protease activity as this provides an insight into the biological process. As one of skill in the art will appreciate, any naturally occurring or synthetically engineered protease or other cleaving enzyme can be detected or otherwise used in the present invention provided the enzyme is able to recognize and cleave the desired bond or bonds in the electrochemical substrate being used.

[0024] A number of useful peptide substrates are commercially available for a range of proteases. Typically a protease cleaves a dye or fluorescent group from these substrates,

resulting in a change in color or fluorescence. It is also possible to create substrates with electroactive groups that can be detected electrochemically after cleavage. The amino acid sequence of a peptide substrate determines specificity for a protease. Thus, in some embodiments, the specificity of a protease detection method can be changed by selection of a substrate specific for the protease of interest.

[0025] One exemplary protease detection system that embodiments of the present invention can detect is the blood coagulation cascade. The blood coagulation cascade consists of a series of protease reactions and serves as a model system, with practical applications, to investigate the detection of proteases. The time taken for a sample of blood or plasma to coagulate can be indicative of various disease states and useful for monitoring anticoagulation therapy. The specificity of such clotting tests usually comes, not from how the coagulation is detected, but from the nature of the reagent(s) used to initiate coagulation. Thus, the present embodiments are applicable to a wide range of clotting tests.

[0026] For example, one area of use for the present invention is methods of monitoring anticoagulation by oral vitamin K antagonists such as warfarin and phenprocoumon at the point-of-care. These drugs are generally monitored using a prothrombin time test and results are reported as an international normalized ratio (INR). There is a trend towards performing these tests using point-of-care devices that provide a result within a couple of minutes using blood derived from a simple finger-prick. This approach is usually more convenient than traditional venipuncture and plasma preparation but such devices require a method for detecting coagulation in whole blood. The present embodiments improve on techniques that address this issue.

[0027] Plasma coagulation can be detected in a number of ways. For example, the gelling of the sample can be detected by increased resistance to particle (typically magnetic) movement through the sample or by a change in fluidity of the sample. The gel point may also correspond to increasing turbidity of plasma, which can be detected optically. Electrochemical methods can be used to detect changes in the viscosity of a sample by measuring changes in impedance, for example, as found in US Patent Nos. 6,673,622; 6,066,504; and 6,060,323; 6,046,051.

[0028] Coagulation can be indirectly detected using artificial peptide substrates for thrombin in accordance with the embodiments. For example, chromogenic substrates can be

cleaved by thrombin (or other coagulation enzymes) and a dye group may be released. This produces a detectable color change in the solution. Other substrates can be provided where the cleaved group is electrochemically active (can be oxidized and/or reduced at an electrode) and can be detected in an electrochemical sensor. This type of substrate is referred to as an electrochemical substrate and is the subject of certain embodiments herein.

[0029] For the purposes of the present embodiments an electrochemical substrate is defined as a chemical compound, which is cleaved by an enzyme to release an electroactive leaving group. The intact substrate has little, or different, electrochemical activity, under the detection conditions, compared to the free leaving group that is released by cleavage.

[0030] An electroactive species can be detected amperometrically in embodiments of the present invention by, for example, using either a two or three electrode system. A three electrode system can have a working, a counter, and a reference electrode. The reference electrode measures the voltage of the solution near the surface of the working electrode. The potential between the working and counter electrode can be adjusted so that the voltage between the working and reference electrode is set to the desired value. The magnitude of the current through the circuit connecting the working and counter electrode is determined by the rate of reduction or oxidation at the solution interface of the working electrode. Thus, it is at the working electrode where the reaction of interest takes place. The system is designed so that reaction at the working electrode, and not the counter electrode, limits the current.

[0031] In an embodiment, a two electrode system does not have a separate reference electrode. Instead the voltage between the conductors of the two remaining electrodes can be set. The working electrode can again be defined as the electrode at which the reaction of interest takes place. This will be the electrode at which the current-limiting reaction occurs.

[0032] The inventors have surprisingly found that electrochemical sensors having opposing electrodes can have a number of advantages over those using co-planar electrodes, including, but not limited to the following:

- (1) Small electroactive species can easily diffuse from one opposing electrode to the other. This can result in a recycling of the reduced and oxidized forms, effectively amplifying the signal. This is not possible in co-planar electrodes as the distances are too great for species to

diffuse the required distance to the working electrode in the time available.

- (2) As one of skill in the art will appreciate, any redox reaction must have an oxidation and a reduction component. In opposing electrode sensors the electroactive species of interest can be oxidized at one electrode and reduced at the other in the aforementioned recycling. In coplanar electrode sensors, extra effort is required to ensure that the reaction at the counter electrode can occur. For example the counter electrode may be Ag/AgCl, or a large voltage may be applied, or a complementary electroactive species may be required.
- (3) The close proximity of opposing electrodes to each other means that most of the potential difference between the electrodes is accounted for by the voltage across the electrode interfaces and not across the bulk of the solution. In contrast, co-planar electrodes may have a considerable $I \times R$ voltage drop across the sample, particularly if the solution between the electrodes has a low ionic strength in an embodiment.
- (4) Manufacture of opposing electrodes can be easier than co-planar because to form coplanar electrodes it can be necessary to either deposit or ablate metal (or another conductive material) to create distinct electrodes that are electrically isolated from each other. In contrast, it can be advantageous to have a conductive surface covering the entire material that forms each opposing electrode. This means, for example, that a simple process like sputter coating can be used to coat the electrodes and specific patterns in the conductor are not required. The opposing electrode area that contacts the sample can be defined by the insulating spacer that separates the electrodes.
- (5) Using opposing electrodes can avoid the requirement in certain embodiments of Frenkel *et al.* (US 6,352,630; 2002) requiring immobilization of an electrochemical substrate in a specific orientation. In opposing electrode sensors the electrochemical

substrate can be free to dissolve in the sample and its orientation is irrelevant. This simplifies the manufacturing process.

[0033] One or more embodiments are related to a device, in the configuration of a strip, which measures protease activities. For example, to monitor blood coagulation. The strip connects electrically to a meter that applies a predetermined voltage or current to the strip and measures the resulting current or voltage. As one of skill in the art will appreciate, these methods are known; for example, as amperometry or galvanostatic electrochemical techniques respectively. The meter can also thermally regulate the strip, calculate, record, display and/or transmit the result of the reaction.

[0034] An exemplary strip, shown in FIGS. 1 and 2, can comprise two electrodes, 1 and 3, in an opposing configuration. The electrodes can be electrically insulated from each other, by a plastic separator, 2, or its equivalent. In some embodiments, the electrodes can be less than 0.5mm apart and in certain embodiments between 0.05 and 0.15mm apart.

[0035] The separator, 2, can have a gap in it that forms a cavity between the electrodes and defines a reaction chamber, 6. The electrode surfaces are electrically conductive and in electrical contact with the meter.

[0036] In certain embodiments, the surfaces are sputter coated onto a plastic support. Suitable coatings may be known to those skilled in the art and may include platinum, iridium, carbon, mixed indium/tin oxide, and tin oxide but preferably palladium or gold.

[0037] The surfaces on each of the electrodes can be the same or different. The conductive surfaces face each other. A fill channel, 5, can allow a sample to be introduced into the reaction chamber while the reaction chamber can be maintained at a set temperature within the meter.

[0038] Dry reagent(s), which can be used to detect protease activity, can be coated onto the surface of the reaction chamber. In certain embodiments, the electrochemical substrate is not coated onto the working electrode. After addition of the sample (e.g., blood), cleavage of the substrate releases an electroactive species and a detectable change in the electrochemical activity take place. The reagent(s) may also contain components that activate blood or plasma coagulation enzymes or that facilitate this activation.

[0039] In certain embodiments, the disclosed strip consists of two electrodes in an opposing configuration. The electrodes can comprise palladium sputter coated onto a plastic support. An insulating separator layer, 0.1mm thick, can define the separation of the electrodes. One or more gaps in the separator can define reaction chamber(s). A fill channel can allow sample to be conducted into the reaction chamber(s) by capillary action.

[0040] Dry reagent(s), which can be useful and/or necessary for the initiation or detection of coagulation, can be coated onto the surface of the reaction chamber(s). These reagents can be, for example, contact activators, tissue factors, animal venoms, phospholipids, buffers, preservatives, stabilizers, dyes, surfactants, and combinations thereof. These dry reagents can react when a blood or plasma sample is introduced into the reaction chamber. The activated coagulation enzymes cleave the electroactive group from the substrate that is present in the reagent(s).

[0041] In a specific embodiment the reagents can contain compounds, including thromboplastin, necessary for a prothrombin time test.

[0042] In a further embodiment, the reagents can contain compounds, such as dilute thromboplastin or tissue factor, necessary for a thrombin generation test from which parameters like lag-time, thrombin peak and endogenous thrombin potential can be derived.

[0043] In another embodiment the reagents can contain compounds, including activators from venoms, necessary for tests such as a dilute Russell viper venom time test or an ecarin clotting time test.

[0044] In another embodiment the reagents can contain factors Xa and/or thrombin and/or antithrombin to measure anticoagulants that inhibit factor Xa or thrombin activity.

[0045] In another embodiment the reagents may also contain plasma components (for example, from bovine plasma) to compensate for variability in coagulation factors like factor V and fibrinogen. This can be analogous to an Owren PT test.

[0046] In some embodiments it is advantageous to coat one or more electrodes with an electroactive compound that ensures the counter electrode does not limit the current. For example, if the cathode is the counter electrode then coating it with iron (III) EDTA provides a reducible compound to support the detection (in this embodiment, by oxidation) of the

species of interest. This can be particularly advantageous if the species of interest is not completely electrochemically reversible.

[0047] In some embodiments, it has been found to be advantageous to dry thromboplastin and the electrochemical thrombin substrate on opposite electrodes to prevent adverse reactions between the two reagents. In one exemplary embodiment, the electrochemical thrombin substrate (Toluolsulfonyl-glycyl-prolinyl-arginin-4-amido-2-chlorophenol) can be coated onto the counter electrode and thromboplastin can be coated onto the working electrode. In this example, the leaving group of the thrombin substrate is cleaved in the reduced form. Here, it is the species that is being detected and limits the current. Since the leaving group is in the reduced form it can be oxidized on the anode to pass current, and thus the anode is the working electrode.

[0048] In general, it would be expected that coating the thrombin substrate on the working electrode would produce faster clot detection because the leaving group is released at the electrode where it is detected. However, in some embodiments, the opposite is true. For example, FIG. 3 shows that the reaction kinetics was faster when the thrombin substrate was dried on the counter electrode (cathode) rather than the working electrode (anode). The faster reaction rate of these embodiments can make determination of the clot time more consistent.

[0049] Many biological assays (e.g., thrombin generation assays) involve calculating reaction rates. In certain embodiments, it is observed that it can be easier to measure the reaction kinetics of a chemically generated electroactive species when the species is diffusing to the working electrode rather than being generated at the working electrode. See, e.g., US Patent Nos. 5,942,102 and 6,444,115. This is another reason why in certain embodiments it can be preferable to coat electrochemical substrates away from the working electrode.

[0050] In another embodiment, the electrochemical substrate and other reagents (e.g., thromboplastin) are formulated so as not to adversely interact, which can allow all or the reagent components to be coated together on the counter electrode. This has an advantage of leaving the working electrode free of any reagent that can foul it. Fouling is less of a concern on the counter electrode because that electrode is not limiting the current. Fouling of the working electrode may not be immediately apparent and often only becomes apparent after storage where it can be a cause of shortened shelf life for a product embodying the present

invention. Thus, coating on the counter electrode can avoid the risk of fouling the working electrode.

[0051] The test strips of the present invention can also include means for verifying strip and/or sample integrity. In one embodiment, the strip has a reaction chamber containing reagents that function to verify strip and/or sample integrity. The reagents can include coagulation factors that are otherwise deficient in the sample. This reaction chamber can produce a coagulation time that is approximately normal, regardless of the degree of deficiency in the sample, and be used to demonstrate strip integrity when the coagulation time falls within predetermined limits.

[0052] In another embodiment, strip integrity may be verified by including an electroactive compound in the reagent. Such a compound may be different from the electrochemical substrate, which may perhaps only release an electroactive group on enzymatic cleavage. The compound that monitors strip integrity may change (increase or decrease) electrochemical activity under storage conditions (e.g. elevated temperature or humidity) that may damage other components of the strip. Examples of such compounds may include 4-amino-2-chlorophenol, ferricyanide, ferrocyanide, and the organic N-oxide or nitroso compounds disclosed in U.S. Patent Application No. US 2005/0123441. The electrochemical activity of such compounds can be compared to predetermined limits for the batch of strips. If the value falls outside the limits then the meter may give an error message instead of a result.

[0053] Reagents to monitor strip integrity may be in the same reaction chamber that detects substrate cleavage or they may be in different chambers. In the embodiment where chambers are separate then the chambers may be electrically isolated by discontinuity of the electrodes or, in an alternative embodiment, the chambers may use electrically connected electrodes. Where the chambers share a continuous electrode, the signal detected by the meter may be a combination of the chambers. Interference between the integrity and cleavage reaction signals can be minimized by detecting the integrity reaction before or after the cleavage reaction and/or by reducing the cross-reacting electrode area.

[0054] Some embodiments of the present are directed to an on-board control suitable for monitoring the viability of electrochemical biosensor strips, such as, but not limited to those disclosed herein. Electrode materials, and the coating on them, are chosen such that a

potential difference is established between the electrodes. This voltage can be measured directly or the current that it induces is measured. In some embodiments, the sensor briefly behaves like a battery, generating charge. The potential difference breaks down after a relatively short time and does not interfere with the subsequent test reaction.

[0055] Such controls are useful in, for example, point-of-care tests, based on a meter that measures an electrochemical reaction in a disposable strip, which are becoming more common. There is a need to ensure that test results are accurate so controls to ensure strip integrity are desirable. For end users, e.g., a patient, the most convenient control is one that is “on-board” the strip and verifies the integrity of the same strip that is being tested. Examples of on-board controls can be found in the area of point-of-care coagulation devices already on the market (e.g., Coaguchek XS, INRatio). Also, US2005123441A1 provides a concise introduction to the advantages of an on-board control.

[0056] Electrochemical biosensors typically become more complicated the more isolated the control reaction is from the test reaction. For example, in order of increasing complexity:

- a. A strip with a single reaction chamber that includes both control reaction and test reaction is comparatively easy to manufacture.
- b. A strip with control and test chambers that are separate but electrically connected is more complicated as it requires deposition of multiple reagents in different areas.
- c. A strip with control and test chambers that are not only separate but also electrically isolated can be quite challenging to manufacture economically at scale. It usually requires deposition or ablating electrode material in a pattern.

[0057] However, the difficulty in finding suitable control reaction chemistry is in the reverse order to that above.

- a. A strip with separate, and electrically isolated, control and test chambers will not have any problems with one reaction (e.g. control) interfering with the other (e.g. test).
- b. A strip with separate, but electrically connected, control and test chambers may suffer from electrical interference between the chambers as electrons from one reaction are

indistinguishable from electrons from the other but the control and test will not interfere with each other chemically.

- c. A strip with a single reaction chamber requires components that will not interfere with each other chemically or cause interfering electrical signals.

[0058] The present invention, in some embodiments, is directed to a novel way of assessing ~~a control reaction in an electrochemical~~ sensor as well as a way to prevent electrical signals from the control reaction and test reaction from interfering with each other.

[0059] US2005123441A1 discloses an on-board control reaction that utilizes N-oxide or nitroso compounds that become reduced upon exposure to conditions that may damage strip performance. The change in on-board control can be detected optically or electrochemically. The electrochemical detection method disclosed in US2005123441A1 involves applying -700mV, relative to Ag/AgCl, across the working electrode and then applying -100mV.

[0060] Embodiments of the present invention can be advantageous because the inventive controls only require 1-2 sec instead of 4.5sec for some previous devices. Our control does not risk applied voltage interfering with the test.

[0061] One specific embodiment involves an on-board control for monitoring the viability of electrochemical sensors that detect blood coagulation. In this specific embodiment the control reaction reagent, which measures strip viability, is contained within the same chamber as the test reaction reagents, which measure clot time. The signal from the control reaction is differentiated from that of the test by the time at which they occur. Specifically, the control reaction is assessed in the first few seconds after adding the sample and then the test reaction is assessed.

[0062] In this embodiment the control reaction can be created by coating, on one electrode, reagent that contains a low (e.g. 0.5mM) concentration of ferricyanide and on the other electrode reagent without ferricyanide. The difference in electrode chemistry creates a voltage across the electrodes when the sample is added. This voltage can be measured directly with a voltmeter or, in a preferred embodiment, the current through in an external circuit can be recorded by an ammeter.

[0063] The electrochemical voltage can dissipate by at least two methods: Current flowing through an external circuit effectively flattens the electrochemical battery and diffusion of ferricyanide from one electrode to the other decreases the ferricyanide concentration gradient and thus the voltage. Once the signal from the control reaction has dissipated then the analysis of the test reaction can begin. This may, for example, consist of applying 0.3-0.4V across the electrodes and measuring the resulting amperometric current.

[0064] In some embodiments, the concentration of ferricyanide is sufficiently low so that it does not substantially interfere with the detection of the test reaction. As one of skill in the art will appreciate, however, other embodiments using alternatives to ferricyanide can be produced. For example, iodine, ascorbate, ferrocyanide, 4-amino-2-chlorophenol can produce a battery effect.

[0065] In an embodiment, the battery effect is measured by measuring the current without any external voltage being applied across the sensor electrodes. However, it is possible to measure the current with a small voltage applied provided that the voltage does not cause substantial interference with the test reaction.

[0066] Embodiments of the present invention can include those with sensors that contain two chambers. One chamber can contain the test-reaction reagents and the other the control-reaction reagents. This approach is advantageous if the components of one reaction (e.g. the control reaction) may chemically inhibit the other (e.g. the test reaction). The chambers do not need to be electrically isolated provided the test and control reaction do not interfere with each other electrically. This embodiment is particularly advantageous if a neutralising agent is included in the control reaction chamber (e.g. on an opposing electrode).

[0067] In another embodiment, concentrations of control-reaction reagent that would normally interfere with the test-reaction can be used. In this embodiment the control chamber contains separate areas of control reagent and a neutralising agent. When the sample is added the control reagent produces a battery effect or some other electrochemical signal but is then quickly neutralised by the neutralising agent. The neutralising effect can be by a chemical reaction (e.g. iodine and ascorbate neutralise each other) or more of a physical effect such as precipitation (e.g. Co^{2+} or Mn^{2+} ions precipitate ferricyanide). Precipitation of an electroactive compound can render it incapable of interacting with an electrode.

[0068] In an embodiment that measures blood or plasma coagulation, a meter can maintain the sensor at 37°C. The electrical signal detected by the meter can be interpreted to calculate the clot time. For example, the clot time can be defined as the point where the signal, or the rate of change thereof, reaches a threshold. The clot time can be reported, displayed, or otherwise outputted with the units of time or it can be converted into different units like INR by using calibration data assigned to the sensor and/or meter.

[0069] Other results may also be calculated and reported. For example, with the right reagents, the area under the curve can indicate the endogenous thrombin potential, which can similarly be reported. The assay result may be reported by displaying on the meter's screen but it can also be stored in memory and/or transmitted to another device.

[0070] Embodiments of the present invention are also directed to kits having a sensor or test strip described herein. Such kits may be sterile packaged. The test strips or sensors can come in individual packages as well as multipacks. Optionally, the kits include a meter for interacting with the test strips or sensor and/or instructions for using the kit, test strip, or sensor.

[0071] All documents cited herein, including websites, journal articles or abstracts, published or corresponding U.S. or foreign patent applications, issued or foreign patents, or any other documents, are each entirely incorporated by reference herein, including all data, tables, figures, and text presented in the cited document. For example, the following items are specifically incorporated in their entirety, including their related commonly invented and/or assigned references, if any, are herein incorporated by reference in their entireties: U.S. Patent No. 4,304,853 to Jozefonvicz *et al.*; U.S. Patent No. 5,942,102 to Hodges *et al.*; U.S. Patent No. 6,060,323 to Jina; U.S. Patent No. 6,066,504 to Jina; U.S. Patent No. 6,046,051 to Jina; U.S. Patent No. 6,248,125 to Hodges *et al.*; U.S. Patent No. 6,444,115 to Hodges *et al.*; U.S. Patent No. 6,352,630 to Frankel *et al.*; U.S. Patent No. 6,495,336 to Ludin *et al.*; U.S. Patent No. 6,620,310 to Ohara *et al.*; U.S. Patent No. 6,673,622 to Jina; U.S. Patent Application No. 2005/0123441 to Unkrig *et al.*; U.S. Patent Application No. 2003/0146113 to Unkrig *et al.*; European Patent No. 1,234,053 to Opalsky *et al.*; International Application No. PCT/IB2007/01990 to Newman *et al.*; Neurath, H. and Walsh, K.A. (1976). Role of proteolytic enzymes in biological regulation (a review). *Proc. Natl. Acad. Sci. USA.* 73: 3825-3832; Spronk, H.M.H. *et al.* (2003). The blood coagulation system

as a molecular machine. *BioEssays*. 25:1220-1228; and Bates, S.M. and Weitz, J.I. (2005). Coagulation assays. *Circulation*. 112:53-60.

[0072] While the invention has been particularly shown and described with reference to some embodiments thereof, it will be understood by those skilled in the art that they have been presented by way of example only, and not limitation, and various changes in form and details can be made therein without departing from the spirit and scope of the invention. Thus, the breadth and scope of the present invention should not be limited by any of the above-described exemplary embodiments, but should be defined only in accordance with the following claims and their equivalents. Any headings used herein are provided solely for organizational purposes and are not intended to impart any division or meaning to this document, unless specifically indicated.

What Is Claimed Is:

1. An opposing-electrode sensor comprising a working electrode, a counter electrode, and an electrochemical substrate, wherein the sensor detects cleavage of the electrochemical substrate.
2. The sensor of claim 1, wherein the electrochemical substrate is not immobilized on the working electrode.
3. The sensor of claim 1, wherein the electrochemical substrate is a substrate for thrombin comprising a peptide bound to a leaving group that becomes electroactive once cleaved by thrombin.
4. The sensor of claim 1, further comprising a meter connected to the sensor, wherein the meter determines a coagulation time for a sample being tested.
5. The sensor of claim 4, wherein the meter displays or outputs a result comprising the coagulation time or a result or value dependent upon the coagulation time.
6. The sensor of claim 1, further comprising a meter connected to the sensor, wherein the meter displays or outputs a result comprising a parameter of thrombin generation in the sample tested.
7. A test strip for measuring blood or plasma coagulation time in a sample comprising an opposing-electrode sensor comprising a working electrode, a counter electrode, and an electrochemical substrate, wherein the sensor detects cleavage of the electrochemical substrate.
8. The test strip of claim 7, wherein the electrochemical substrate in the sensor is not immobilized on the working electrode.
9. The test strip of claim 7, wherein the substrate in the sensor is a substrate for thrombin comprising a peptide bound to a leaving group that becomes electroactive once cleaved by thrombin.

10. The test strip of claim 7, further comprising a meter connected to the sensor, wherein the meter determines a coagulation time for a sample being tested.
11. The test strip of claim 10, wherein the meter displays or outputs a result comprising the coagulation time or a result or value dependent upon the coagulation time.
12. The test strip of claim 7, further comprising a meter connected to the sensor, wherein the meter displays or outputs a result comprising a parameter of thrombin generation in the sample tested.
13. The test strip of claim 7, further comprises means for verifying strip integrity.
14. A method of measuring blood or plasma coagulation comprising:
 - obtaining a blood sample;
 - adding the blood sample to an opposing-electrode sensor comprising a working electrode, a counter electrode, and an electrochemical substrate, wherein the sensor detects cleavage of the electrochemical substrate; and
 - analyzing the sample using the sensor to determine at least one blood or plasma coagulation property of the sample.
15. The method of claim 14, wherein the method is performed at the point of care.
16. The method of claim 14, wherein the method further comprises a meter connected to the sensor, wherein the meter displays or outputs a result comprising a parameter of thrombin generation in the sample tested.
17. The method of claim 14, wherein the patient is taking an anticoagulant medication.
18. The method of claim 17, wherein the anticoagulant medication is a vitamin K antagonist.
19. The method of claim 14, wherein the sample is obtained from a patient by a finger prick of the patient.