

SALT STABLE CATIONIC SILICONE OIL-IN-WATER MICROEMULSION

This invention is directed to cationic oil-in-water microemulsions of polyorganosiloxanes which are stable in the presence of salts. More particularly, cationic oil-in-water microemulsions containing polydimethylsiloxanes and amine functional polysiloxanes do not break upon the addition of greater than twenty percent by weight of salts such as sodium chloride, calcium chloride and ammonium chloride.

Breaking is a term applied to the irreversible coalescence of the individual emulsion or microemulsion droplets to form a continuous phase. To break an emulsion or microemulsion, the stabilizing forces must be weakened or eliminated.

The addition of electrolytes, i.e., salts for example, often causes an emulsion or microemulsion to break, since the electric charge of a boundary layer is very sensitive to the addition of foreign ions. The effectiveness of the added electrolyte increases with its valence. Thus, aluminum sulfate or iron sulfate are more effective than sodium sulfate and sodium sulfate is in turn more effective than sodium chloride, i.e., see Ullmann's Encyclopedia of Industrial Chemistry, Volume A9, Page 311, VCH, Weinheim, Germany (1987).

Accordingly, to provide a microemulsion that was stable in the presence of salts is unexpected and contrary to conventional wisdom.

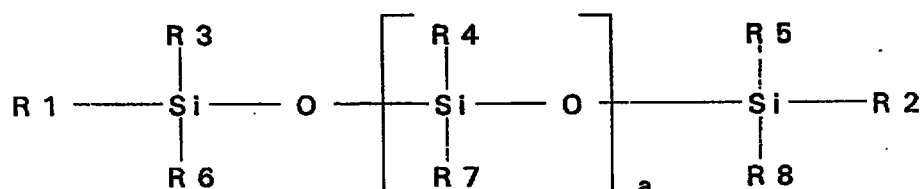
The need and usefulness of such a microemulsion in textile applications, for example, is evidenced by US Patent 5,466,746 (November 14, 1995), in which an alkyl polyglycoside, which is a nonionic surfactant, is used to prepare mechanical microemulsions possessing some degree of stability with respect to the presence of foreign electrolytes such as sodium chloride. This is not totally unexpected, however, as nonionic surfactants carry no charge.

What is unexpected, and what has been discovered according to the present invention, is that microemulsions stable to the presence of foreign electrolytes are conveniently prepared using cationic surfactants, which contain positively charged polar groups, especially if the microemulsion is prepared using an emulsion polymerization technique, rather than prepared mechanically as in the previous patent.

This invention relates to cationic microemulsions that remain stable when incorporated with a variety of ionic inorganic and organic salts. Thus, cationic microemulsions with an average particle size of 140 nanometer (0.14 micrometer) or less, preferably 100 nanometer (0.10 micrometer) or less, and most preferably 40 nanometer (0.04 micrometer) or less, have been found to provide salt

stability. Salt concentrations as high as thirty weight percent of the total solution can be incorporated into such stable systems; while lesser amounts can be easily tolerated, i.e., 0.1 to thirty percent, preferably twenty to thirty percent. In addition, increased solution clarity and reduced tackiness are added benefits provided by such salt stable cationic microemulsions. These microemulsions demonstrate long term system stability, i.e., upwards of six months, when stored at ambient temperatures.

As used herein, polysiloxane is intended to mean a polymer having the general structure:



where a has a value to provide polymers with a viscosity in the range of 1-30,000 centistoke (mm^2/sec). R1-R8 are alkyl radicals of 1-6 carbon atoms or an aryl group such as phenyl. Typically, the value of a is 1-1,250. Illustrative polysiloxanes are polydimethylsiloxane, polydiethylsiloxane, polymethylethylsiloxane, polymethylphenylsiloxane and polydiphenylsiloxane. Preferred are polydimethylsiloxanes having a viscosity of 20-1,000 centistoke (mm^2/sec). Most preferred are polydimethylsiloxanes having a viscosity of 20-350 centistoke (mm^2/sec). By polydimethylsiloxane

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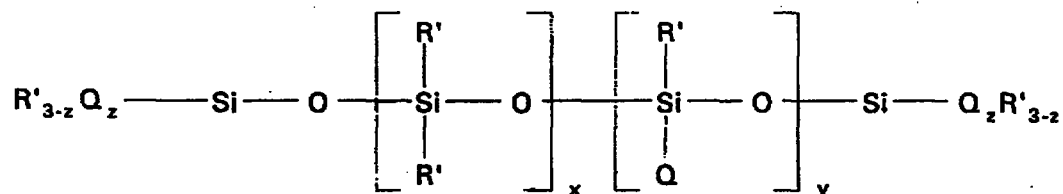


(PDMS), we mean a polysiloxane having the above structure where R4 and R7 are methyl radicals.

In addition to the aforementioned organic polysiloxanes, the present compositions may also include an amine functional polysiloxane. The amine functional polysiloxane is a silicone fluid with highly polar pendant aminoalkyl modifying groups that enhance the durability of films formed by the polysiloxane and promote adhesion of the films to various substrates.

Particularly preferred amine functional polysiloxanes include reactive and non-reactive hydrolyzable and non-hydrolyzable derivatives which are wholly, or in part, terminally substituted with aminopropyl, aminobutyl or diamino pendant chains.

The amine functional polysiloxane employed in the microemulsions of the present invention has the formula:



wherein R' denotes an alkyl group of 1 to 4 carbons or a phenyl group with the proviso that at least 50 percent of the total number of R' groups are methyl; Q denotes an amine functional substituent of the formula -R''Z wherein R'' is a divalent alkylene radical of 3 to 6 carbon atoms and Z is a monovalent radical selected from the group consisting of

$-\text{NR}_2'''$ and $-\text{NR}'''(\text{CH}_2)_n\text{NR}_2'''$; wherein R''' denotes hydrogen atom or an alkyl group of 1 to 4 carbons, and n is a positive integer having a value of from 2 to 6; z has a value of 0 or 1; x has an average value of 25 to 3000; y has an average value of 0 to 100 when z is 1, y has an average value of 1 to 100 when z is 0; with the proviso that in all cases y has an average value that is not greater than one tenth the average value of x .

Suitable R' groups are represented by and independently selected from methyl, ethyl, propyl, isopropyl, butyl, isobutyl and phenyl, with the proviso that at least fifty percent of the R' groups are methyl.

The alkylene radicals represented by R'' include trimethylene, tetramethylene, pentamethylene, $-\text{CH}_2\text{CHCH}_3\text{CH}_2-$ and $-\text{CH}_2\text{CH}_2\text{CHCH}_3\text{CH}_2-$. Siloxanes where R'' is a trimethylene or an alkyl substituted trimethylene radical such as $-\text{CH}_2\text{CHCH}_3\text{CH}_2-$, are preferred.

Alkyl groups of 1 to 4 carbon atoms as represented by R''' include methyl, ethyl, propyl, isopropyl, butyl and isobutyl.

Useful Z radicals include the unsubstituted amine radical $-\text{NH}_2$, alkyl substituted amine radicals such as $-\text{NHCH}_3$, $-\text{NHCH}_2\text{CH}_2\text{CH}_2\text{CH}_3$ and $-\text{N}(\text{CH}_2\text{CH}_3)_2$; and aminoalkyl substituted amine radicals such as $-\text{NHCH}_2\text{CH}_2\text{NH}_2$, $-\text{NH}(\text{CH}_2)_6\text{NH}_2$ and $-\text{NHCH}_2\text{CH}_2\text{CH}_2\text{N}(\text{CH}_3)_2$.

When z is zero, the amine functional polysiloxane has only pendent amine substituents in the polymer chain.

When z is one, the amine functional polysiloxane has only terminal amine substituents or both terminal and pendent amine substituents in the polymer chain. Preferably, x may vary from a value of 25 to 100, and y may vary from zero to 100 when z is one and from one to 100 when z is zero. Most preferably, the value of $x + y$ is in the range of 50 to 500.

The composition according to the present invention includes as one of its components an aqueous microemulsion of either a polydimethylsiloxane or an amine functional polysiloxane stabilized in the microemulsion by one or more ionic surfactants. These microemulsions are prepared mechanically or by emulsion polymerization.

Suitable mechanical microemulsions include compositions containing a polydimethylsiloxane or amine functional polysiloxane stabilized by a cationic surfactant. The siloxane is present in the microemulsion in the form of particles having a diameter less than about 140 nanometer (0.14 micrometer). Such microemulsions can be prepared by mechanical processes described in US Patent 5,017,221 (May 21, 1991) and EP 463 431 (January 2, 1992), for example.

According to such processes, water, one or more ionic surfactants and the siloxane, are simply mixed together and homogenized using a laboratory homogenizer or other device for applying vigorous agitation. These mechanical microemulsions typically have a pH of 7-9.5 and contain 10-80 percent by weight of the siloxane, preferably

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20-60 percent, 0.01-15 percent by weight of the ionic surfactant(s), the balance being water.

Emulsion polymerization is also used to prepare suitable aqueous microemulsions containing polydimethylsiloxanes or amine functional polysiloxanes, stabilized in the microemulsions by an ionic surfactant. The siloxane is present in the microemulsion as particles having a diameter of less than 140 nanometer, preferably less than 40 nanometer. By way of comparison, fine emulsions contain a siloxane as particles with a diameter of 140-300 nanometer (0.14-0.30 micrometer). Standard emulsions, on the other hand, contain a siloxane as particles with a diameter greater than 300 nanometer (0.30 micrometer).

In any event, the microemulsion is readily prepared by the emulsion polymerization process described in EP 459 500 (December 4, 1992), for example. According to that process, stable, oil-free polysiloxane microemulsions are prepared by mixing a cyclic siloxane, an ionic surfactant, water and a condensation polymerization catalyst. The mixture is heated and agitated at a polymerization reaction temperature until essentially all of the cyclic siloxane is reacted and a stable, oil-free microemulsion is formed. The microemulsions typically have a pH of 6.5-8.0 and contain 10-70 percent by weight of the siloxane polymer, preferably 25-60 percent, 0.05-30 percent by weight of the ionic surfactant, preferably 0.5-20

percent, with the balance being water. The microemulsion can also be prepared by the emulsion polymerization process described in US Patent 2,891,920 (June 23, 1959) or in US Patent 4,733,677 (March 29, 1988).

The ionic surfactant according to the present invention is any conventional cationic emulsifier used in emulsion polymerization. Such cationic surfactants include fatty acid amines, amides and their salts. Suitable cationic surfactants are exemplified by aliphatic fatty amines and their derivatives such as dodecylamine acetate, octadecylamine acetate and acetates of the amines of tallow fatty acids; homologues of aromatic amines having fatty chains such as dodecylaniline; fatty amides derived from aliphatic diamines such as undecyl imidazoline; fatty amides derived from di-substituted amines such as oleyl aminodiethylamine; derivatives of ethylene diamine; quaternary ammonium compounds such as tallow trimethyl ammonium chloride, dioctadecyl dimethyl ammonium chloride, didodecyl dimethyl ammonium chloride and dihexadecyl dimethyl ammonium chloride; amide derivatives of amino alcohols such as beta-hydroxyethyl stearylamine; amine salts of long chain fatty acids; quaternary ammonium bases derived from fatty amides of di-substituted diamines such as oleylbenzyl aminoethylene diethylamine hydrochloride; quaternary ammonium bases of benzimidazolines such as methylheptadecyl benzimidazol hydrobromide; basic compounds of pyridinium and its derivatives such as cetyl pyridinium

chloride; sulfonium compounds such as octadecyl sulfonium methyl sulfate; quaternary ammonium compounds of betaine such as betaine compounds of diethylamino acetic acid and octadecyl chloromethyl ether; urethanes of ethylene diamine such as condensation products of stearic acid and diethylene triamine; polyethylene diamines; and polypropanol polyethanol amines.

Examples of commercial cationic surfactants are those products sold under the tradenames ARQUAD T-27 W, ARQUAD 16-29 W, ARQUAD C-33 W, ARQUAD T-50 and ETHOQUAD T-13-27 W ACETATE by Akzo Chemicals, Inc., Chicago, Illinois.

Any catalyst that is capable of polymerizing cyclic siloxanes in the presence of water is useful in the preparation of the microemulsion. Catalysts include condensation polymerization catalysts capable of cleaving siloxane bonds. Some examples of condensation polymerization catalysts are strong acids such as substituted benzene sulfonic acids, aliphatic sulfonic acids, hydrochloric acid and sulfuric acid; and strong bases such as quaternary ammonium hydroxides and metal hydroxides. Ionic surfactants such as dodecylbenzene sulfonic acid, i.e., BIOSOFT S-100, additionally function as catalyst. Other useful catalytic systems include phase transfer catalysts such as tetrabutyl ammonium hydroxide or ion exchange resins where a catalyst is formed in situ.

The reaction is stopped by neutralization of the catalyst by the addition of an equal or slightly greater

stoichiometric amount of an acid or a base, depending upon the type of catalyst used in the reaction. A strong or weak acid or base may be used; however, care must be taken when using a strong acid or base not to over neutralize, as it may be possible to re-catalyze the reaction. It is preferred to neutralize with a sufficient quantity of an acid or a base such that the resulting emulsion has a pH of less than 7.

As used herein, the term salt is intended to mean an inorganic salt or an organic salt, including compounds commonly referred to as electrolytes.

Some examples of suitable inorganic salts include calcium chloride, magnesium sulfate, magnesium chloride, sodium sulfate, sodium thiosulfate, sodium chloride, sodium phosphate, ammonium chloride, ammonium carbonate, iron sulfate, aluminum sulfate, aluminum chloride, aluminum chlorohydrate, aluminum sesquichlorohydrate, aluminum dichlorohydrate, aluminum zirconium tetrachlorohydrate, aluminum zirconium trichlorohydrate, aluminum zirconium tetrachlorohydrate, aluminum zirconium pentachlorohydrate and aluminum zirconium octachlorohydrate.

Some examples of suitable organic salts include sodium aluminum lactate, sodium acetate, sodium dehydroacetate, sodium butoxy ethoxy acetate, sodium caprylate, sodium citrate, sodium lactate, sodium dihydroxy glycinate, sodium gluconate, sodium glutamate, sodium hydroxymethane sulfonate, sodium oxalate, sodium phenate,

sodium propionate, sodium saccharin, sodium salicylate, sodium sarcosinate, sodium toluene sulfonate, magnesium aspartate, calcium propionate, calcium saccharin, calcium d-saccharate, calcium thioglycolate, aluminum caprylate, aluminum citrate, aluminum diacetate, aluminum glycinate, aluminum lactate, aluminum methionate, aluminum phenosulfonate, potassium aspartate, potassium biphthalate, potassium bitartrate, potassium glycosulfate, potassium sorbate, potassium thioglycolate, potassium toluene sulfonate and magnesium lactate.

The microemulsion may optionally contain a small amount of a preservative, generally in an amount of 10 to 1,000 ppm, to prevent microbial contamination and multiplication. Some representative examples are compounds such as formaldehyde; 1,3-dimethylol-5,5-dimethyl (DMDM) hydantoin; 5-bromo-5-nitro-1,3-dioxane; methyl paraben (methyl para-hydroxy benzoate); propyl paraben (propyl para-hydroxy benzoate); sorbic acid; imidazolidinyl urea; 5-chloro-2-methyl-4-isothiazolin-3-one or KATHON® LX by the Rohm and Haas Company, Philadelphia, Pennsylvania; Quaternium-15 (chloroallyl methenamine chloride); and GERMABEN® II-E by Sutton Laboratories, Chatham, New Jersey.

EXAMPLES

In order to determine stability limits, several salts and microemulsions were included in the following test protocols. The salts that were used included a 50 percent by weight aqueous solution of aluminum chlorohydrate (ACH-

303); powdered aluminum chlorohydrate (ACH-323 Powder); a 35 percent by weight aqueous solution of aluminum zirconium tetrachlorohydrate glycine (AZG-368); powdered aluminum zirconium tetrachlorohydrate glycine (AZG-370 Powder); a 50 percent by weight aqueous solution of aluminum zirconium tetrachlorohydrate glycine (AZG-370); sodium chloride; calcium chloride; and ammonium chloride.

Aluminum zirconium tetrachlorohydrate glycine is a coordination complex of aluminum zirconium tetrachlorohydrate $\text{Al}_4\text{Zr}(\text{OH})_{12}\text{Cl}_4 \cdot n\text{H}_2\text{O}$ and the amino acid glycine $\text{H}_2\text{NCH}_2\text{COOH}$, in which some of the water molecules normally coordinated to the metal have been displaced by the glycine.

Anionic, nonionic and cationic emulsions and microemulsions were selected for the test protocols. Emulsions and microemulsions with particle size characteristics of the polysiloxane ranging from less than 40 nanometer (nm) to 450 nanometer (nm) were considered in the course of our tests.

As used herein, the term microemulsion is intended to include compositions in which a siloxane is present as particles having an average particle diameter of less than 140 nanometer. The term emulsion is intended to include compositions in which a siloxane is present as particles having an average particle diameter of more than 140 nanometer.

The internal phase, i.e., the polysiloxane, viscosity for the eleven (11) emulsions and microemulsions

considered ranged from 2000 centipoise (cP)/mPa·s to 1,000,000 cP/mPa·s. Table I shows each of these emulsions and microemulsions and their physical properties. Table I also identifies the type of polysiloxane in each emulsion/microemulsion, as PDMS for those containing a polydimethylsiloxane polymer and as AMINO for those containing an amine functional polysiloxane polymer.

TABLE I - Emulsions/Microemulsions Tested for Salt Stability

Silicone Emulsion or Microemulsion	Surfactant Type	Polymer Viscosity centipoise*	Particle Size nanometer	pH	Percent Silicone	Salt Stability
Microemulsion 1000	Cationic	15,000 PDMS	42-72	5-6.5	30	Stable
Microemulsion 1100	Cationic	3,000 AMINO	< 40	7.5 to 8.5	29	Stable
Microemulsion 1200	Cationic	80,000 AMINO	~ 33	6.5-8	25	Stable
Microemulsion 1300	Cationic	3,000 AMINO	~ 26	6.5-8	25	Stable
Emulsion 1400	Cationic	2,000 AMINO	120-150	6.5-9	36	Unstable
Emulsion 1500	Cationic	1,000,000 PDMS	100-170	5-7	36	Unstable
Microemulsion 1600	Anionic	> 15,000 PDMS	< 40	6.5-8	25	Unstable
Microemulsion 1700	Anionic	60,000 to 600,000 PDMS	45-50	6.5-8	25	Unstable
Microemulsion 1800	Anionic	4,500 to 7,000 PDMS	< 40	6.5-8	25	Unstable
Emulsion 1900	Anionic	1,000,000-1,500,000 PDMS	< 1,000	6-8	50	Unstable
Emulsion 2000	Nonionic	110-180 AMINO	> 300	9-11.5	36	Unstable

*Centipoise = mPa·s

Silicone 1000 was a cationic microemulsion prepared by emulsion polymerization. Silicone cationic microemulsion 1000 contained as the cationic surfactant ETHOQUAD T-13-27 W, a tallow alkyltris (2-hydroxyethyl) ammonium acetate.

Silicone 1100 was a cationic microemulsion prepared by emulsion polymerization. Silicone cationic microemulsion 1100 contained as the cationic surfactant ARQUAD 16-29 W, i.e., hexadecyl (cetyl) trimethyl ammonium chloride.

Silicone 1200 was a cationic microemulsion prepared by emulsion polymerization. Silicone cationic microemulsion 1200 contained as the cationic surfactant ARQUAD 16-29 W.

Silicone 1300 was a cationic microemulsion prepared by emulsion polymerization. Silicone cationic microemulsion 1300 contained as the cationic surfactant ARQUAD 16-29 W.

Silicone 1400 was a cationic emulsion prepared by emulsion polymerization. Silicone cationic emulsion 1400 contained as the cationic surfactant ARQUAD 16-29 W.

Silicone 1500 was a cationic emulsion prepared by emulsion polymerization. Silicone cationic emulsion 1500 contained as the cationic surfactant ARQUAD T-27 W, a tallow trimethyl ammonium chloride.

Silicone 1600 was an anionic microemulsion prepared by emulsion polymerization. Silicone anionic microemulsion 1600 contained as the anionic surfactant

dodecylbenzene sulfonic acid (DBSA) BIOSOFT S-100 by the Stepan Company, Northfield, Illinois.

Silicone 1700 was an anionic microemulsion prepared by emulsion polymerization. Silicone anionic microemulsion 1700 contained as the anionic surfactant DBSA.

Silicone 1800 was an anionic microemulsion prepared by emulsion polymerization. Silicone anionic microemulsion 1800 contained as the anionic surfactant DBSA.

Silicone 1900 was an anionic emulsion prepared by emulsion polymerization. Silicone anionic emulsion 1900 contained as the anionic surfactant DBSA.

Silicone 2000 was a nonionic emulsion prepared mechanically. Silicone nonionic emulsion 2000 contained as nonionic surfactants TERGITOL® TMN-6, i.e., ethoxylated trimethylnonanol and TRITON® X-405, i.e., octylphenoxy polyethoxy ethanol; products of Union Carbide Corporation, Danbury, Connecticut.

Silicone Emulsions and Microemulsions 1000 to 1900 were generally prepared according to the method described in EP 459 500, as noted above, in which a cyclic siloxane, ionic surfactant, water and catalyst, are mixed, heated and agitated.

Silicone Cationic Emulsions and Microemulsions 1000 to 1500 were also prepared generally according to the method described in US Patent 2,891,920, as noted above.

Silicone Anionic Emulsions and Microemulsions 1600 to 1900 were also prepared generally according to the method

described in US Patent 3,294,725 (December 27, 1966); or generally according to the method described in the Journal of Polymer Science, Part C, entitled "Anionic Emulsion Polymerization of Siloxanes", Pages 27-34, (1969).

Silicone Mechanical Emulsion 2000 was prepared generally according to the method described in US Patent 5,017,221 and EP 463 431, as noted above, in which a functional or non-functional polysiloxane, surfactant and water, were mixed and emulsified using a high shear device.

EXAMPLE I - Solutions

Solutions were prepared containing a salt and one of the eleven emulsions or microemulsions shown in Table I. These solutions were monitored for stability after aging under ambient temperatures, i.e., at 20 to 25°C. and at 40°C. The solutions were also subjected to freeze/thaw cycles to challenge their stability. In addition, the solutions containing the emulsions and microemulsions were monitored for changes in clarity using a Bausch & Lomb™ Spectronic 21 Spectrophotometer. The results are shown below in Tables II-VII.

EXAMPLE II - Anionic Microemulsions

In this example, it was determined that all of the anionic emulsions and microemulsions shown in Table I were not salt stable. Thus, when salts were added to anionic emulsions and microemulsions seen in Table I, strong signs of instability were observed, i.e., thickening and separation. As an example, several solutions were prepared

using several different salts and one of the anionic microemulsions and the results provide evidence of the instability which results from a combination of an anionic emulsion or microemulsion and a salt. The results are shown in Table II.

TABLE II - Examples of Anionic Microemulsion Instability

Micro emulsion	Salt	Wt % Salt Powder Solution	Initial Observation	Observation after 24 Hours
1800	ACH-323 Powder	20	Hard to mix solid/opaque	Opaque/Thick
1800	AZG-370 Powder	20	Hard to mix solid/opaque	Opaque/Thick
1800	Calcium Chloride	20	Exotherm/Opaque	Opaque/Thick
1800	Sodium Chloride	20	Opaque/Thick	Opaque/Thick Separated
1800	Ammonium Chloride	20	Endotherm/Opaque	Opaque

EXAMPLE III - Nonionic Emulsion

In this example, a nonionic emulsion containing an amine functional polysiloxane was tested and found unstable when combined with several different types of salts. The results are shown in Table III.

TABLE III - Non-ionic Emulsion Instability

Emulsion	Salt	Wt Percent Salt/Powder Solution	Initial Observation	Observation after 24 Hours
2000	Sodium Chloride	20	No Change	Phases Separated
2000	ACH-323 Powder	20	Thick	Slight ring on top/thick
2000	AZG-370 Powder	20	Thick	Phases Separated

EXAMPLE IV - Cationic Microemulsions

In this example, cationic microemulsions of varying particle size and internal phase viscosity as shown in Table I were tested. Thus, three cationic microemulsions containing an amine functional polysiloxane with a particle size of less than 40 nanometer were found to be stable upon salt addition. However, a cationic microemulsion containing an amine functional polysiloxane with a particle size greater than 120 nanometer was found to be unstable upon addition of the salt. A cationic microemulsion containing a polydimethylsiloxane with a particle size less than 40 nanometer was found to be stable upon addition of the salt. However, a cationic microemulsion containing a polydimethylsiloxane with a particle size greater than 100 nanometer was found to be unstable upon addition of the salt. Differences in internal phase viscosity did not appear to negatively or positively affect the stability results for the cationic microemulsions containing an amine functional polysiloxane; however, it did impact stability results for the two cationic microemulsions containing polydimethylsiloxanes as seen in Table I.

An improvement in clarity was noted for several of the salt containing solutions that were tested. In general, stability was achieved with a salt content as high as 30 percent by weight of the salt, as shown in Tables IV-VII which show results obtained with one particular cationic

microemulsion containing an amine functional polysiloxane, that was examined using various salts in various concentrations in several solutions.

TABLE IV - Salt Stable Microemulsion

Wt % 1300 % T = 96.5	% Si	Wt % AZG-370	Wt % Water & Total Water	Observation	% T	Ambient Stability (24 Hours)	Ambient Stability (6 Months)
40	10	20	40/62.8	Clear	98	stable	stable
40	10	25	35/57.8	Clear	98.5	stable	stable
40	10	30	30/52.8	Clear	98	stable	stable
40	10	35	25/47.8	Clear	99	stable	separated
40	10	40	20/42.8	Hard to mix	96.5	stable	thick
40	10	45	15/37.8	Hard to mix cloudy	67	viscous	cloudy thick
40	10	50	10/32.8	Hard to mix cloudy	3	viscous	cloudy thick
50	13	20	30/58.5	Clear	97	stable	stable
80	20	20	0/45.6	Clear	99	stable	stable thick

TABLE V - Salt Stable Microemulsion

Wt % 1300 % T = 96.5	% Si	Wt % Sodium Chloride	Wt % Water & Total Water	Observation	% T	Ambient Stability (24 Hours)	Ambient Stability (6 Months)
40	10	20	40/62.8	Clear	94	stable	stable
40	10	25	35/57.8	Salt hard to dissolve	96	stable	separated

TABLE VI - Salt Stable Microemulsion

Wt % 1300 % T = 96.5	% Si	Wt % Calcium Chloride	Wt % Water & Total Water	Observation	% T	Ambient Stability (24 Hours)	Ambient Stability (6 Months)
40	10	20	40/62.8	Clear	99	stable	stable
40	10	25	35/57.8	Clear	100	stable	stable
40	10	30	30/52.8	Clear	98	stable	separated

TABLE VII - Salt Stable Microemulsion

Wt % 1300 % T = 96.5	% Si	Wt % Ammonium Chloride	Wt % Water & Total Water	Observation	% T	Ambient Stability (24 Hours)	Ambient Stability (6 Months)
40	10	20	40/62.8	Clear	98	stable	stable
40	10	25	35/57.8	Salt hard to dissolve	98.5	stable	separated

Salt stable cationic microemulsions according to this invention are useful generally in textile applications and in personal care products, where the presence and tolerance of large amounts of one or more salts is beneficial.

Claims



1. A composition comprising a stable mixture of (A) a silicone component which is an aqueous oil-in-water microemulsion containing a polysiloxane stabilized in the microemulsion by a cationic surfactant, the polysiloxane in the microemulsion having a particle diameter of less than 100 nanometer; and (B) an inorganic salt or an organic salt.

2. A composition according to claim 1 in which the composition contains 0.1 to thirty percent by weight of the salt, based on the total weight of the composition, the remainder of the composition being the microemulsion; the microemulsion containing 10 to 80 percent by weight of the polysiloxane, 0.05 to 30 percent by weight of the cationic surfactant, with the balance being water, based on the total weight of the microemulsion.

3. A composition according to any of the preceding claims in which the polysiloxane in the microemulsion has a particle diameter less than 40 nanometer.

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