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NOVEL METAL ION-EXCHANGED PHOSPHORUS-VANADIUM COMPOUND AND
SOLID ACID CATALYST USING THE COMPOUND

BACKGROUND OF THE INVENTION

Field of the Invention:

5 This invention relates to a novel metal ion-exchanged
phosphorus-vanadium oxide and to a solid acid catalyst which
contains this phosphorus-vanadium oxide. More particularly,
this novel phosphorus-vanadium oxide is a compound which
exhibits a specific d value (lattice spacing, Å). This
10 compound manifests the qualities of a solid acid and, owing
to the acid qualities, can be used as a catalyst.

Description of the Prior Art:

Various studies have been heretofore made concerning
the phosphorus-vanadium oxides and physical properties and
15 uses of these oxides have been simultaneously explored for the
sake of development. The phosphorus-vanadium compounds are
specific in such qualities as the state of oxidation and the
configuration and consequently have various crystal structures
and find various applications owing to these structures. It
20 is well known that the catalytically effective component of
phosphorus-vanadium oxides is divanadyl pyrophosphate, a
crystalline oxide which has the composition of $(VO)_2P_2O_7$. This
divanadyl pyrophosphate is obtained by synthesizing its
precursor, vanadyl hydrogen orthophosphate hydrate
25 ($VOHPO_4 \cdot nH_2O$), and thermally dehydrating this precursor. It
is also known that the vanadyl hydrogen orthophosphate hydrate
as the precursor is an interlayer compound which is formed by
the hydrogen bondage of opposed $VOHPO_4$ layers through the medium
of water molecules intervening therebetween.

30 It is widely known that phosphorus-vanadium compounds
generally form an effective catalyst for gas-phase oxidation
of such hydrocarbons as butane, butene, and butadiene which
have four carbon atoms (hereinafter referred to briefly as "C4

hydrocarbons"). Various patterns of X-ray diffraction peaks for identifying these phosphorus-vanadium oxides have been published (JP-A-53-61,588, JP-A-56-41,816, JP-A-56-45,815, JP-A-59-132,938, and JP-A-05-15,781). Various inventions
5 concerning the addition of third components to the phosphorus-vanadium oxides have been also disclosed (JP-A-52-135,580, JP-A-54-30,114, and JP-A-57-111,219).

Reports on phosphorus-vanadium oxides are published in various articles of literature besides the patent publications
10 mentioned above. B. K. Hodnett, ed., Catalysts Today, Vol. 1, No. 5 (1987), for example, carries a detail report.

No attempt has ever been made to synthesize a compound having a divalent metal cation exchanged for the H^+ ion which intervenes between the opposed layers of the vanadyl hydrogen
15 orthophosphate hydrate, a precursor of a phosphorus-vanadium oxide.

It is well known that the phosphorus-vanadium compound having such a plane structure as illustrated in the form of a model in Fig. 1, e.g. a vanadyl hydrogen orthophosphate
20 hydrate, $VOHPO_4$, is a layer compound which results from the hydrogen bondage through the medium of constitution water intervening between the $VOHPO_4$ layers as shown in Fig. 2. The interlayer distance is generally 5.70 Å. Absolutely no case of having this interlayer proton exchanged with other metal
25 cation has been known to date.

An object of this invention, therefore, is to provide a novel phosphorus-vanadium compound.

Another object of this invention is to provide a novel phosphorus-vanadium compound which is obtained by the exchange
30 of a proton present between opposed layers in the layer structure of a phosphorus-vanadium compound with a metal ion and a method for the production thereof.

Yet another object of this invention is to provide a

novel solid acid catalyst.

A further object of this invention is to provide a catalyst appropriate for gas-phase partial oxidation of a hydrocarbon.

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SUMMARY OF THE INVENTION

The objects mentioned above are accomplished by the following items, (1) - (8).

- (1) A metal ion-exchanged phosphorus-vanadium compound obtained by treating vanadyl hydrogen orthophosphate
10 hydrate represented by the formula (1)



wherein n fulfills the expression, $0 \leq n \leq 2.0$, with an aqueous divalent metal salt solution thereby effecting the exchange of H^+ present between the layers of the vanadyl
15 hydrogen orthophosphate hydrate and drying the resultant ion exchanger.

(2) A compound according to Item (1) above, wherein the interlayer distance is in the range of 7.0 to 8.2 Å.

- (3) A metal ion-exchanged phosphorus-vanadium
20 compound obtained by treating vanadyl hydrogen orthophosphate hydrate represented by the formula (1)



wherein n fulfills the expression, $0 \leq n \leq 2.0$, with an aqueous divalent metal salt solution thereby effecting the
25 exchange of H^+ present between the layers of the vanadyl hydrogen orthophosphate hydrate, drying the resultant ion exchanger, and calcining the dried ion exchanger.

(4) A compound according to Item (1) or Item (3) above, wherein the divalent metal is at least one member selected from
30 the group consisting of cobalt, nickel, copper, zinc, manganese, iron, magnesium, palladium, platinum, and germanium.

(5) A method according to Item (1) or Item (3) above, wherein the drying is carried out at a temperature in the range

of 150° to 250 °C.

(6) A method according to Item (3), wherein the calcination is carried out in the presence of a hydrocarbon under an oxidizing atmosphere at a temperature in the range
5 of 350° to 600 °C.

(7) A solid catalyst formed of a metal ion-exchanged phosphorus-vanadium compound set forth in any of Items (1) - (6).

(8) A catalyst for gas-phase partial oxidation of a
10 hydrocarbon, formed of a metal ion-exchanged phosphorus-vanadium compound set forth in any of Items (1) - (6).

BRIEF DESCRIPTION OF THE DRAWINGS

Fig. 1 is a model plan view illustrating the structure of vanadyl hydrogen orthophosphate hydrate.

15 Fig. 2 is a model diagram illustrating a cross section of the structural diagram shown in Fig. 1.

Fig. 3 is a model diagram illustrating the structure of vanadyl hydrogen orthophosphate hydrate and the cross-sectional structure of an ion exchanger thereof.

20 Fig. 4 is an X-ray diffraction chart illustrating one example of a metal ion-exchanged phosphorus-vanadium compound according to this invention.

Fig. 5 is an X-ray diffraction chart illustrating another example of the metal ion-exchanged phosphorus-vanadium
25 compound according to this invention.

DESCRIPTION OF THE PREFERRED EMBODIMENT

The metal ion-exchanged phosphorus-vanadium compound in accordance with this invention is obtained by causing a vanadium compound to react with a phosphorus compound thereby
30 forming vanadyl hydrogen orthophosphate hydrate represented by the formula (1)



wherein n fulfills the expression, $0 \leq n \leq 2.0$, treating the

vanadyl hydrogen orthophosphate hydrate with an aqueous divalent metal salt solution thereby effecting the exchange of H^+ present between the layers of the vanadyl hydrogen orthophosphate hydrate and drying the resultant ion exchanger.

5 The dried ion exchanger, when necessary, is further calcined.

The interlayer distance of the metal ion-exchanged phosphorous-vanadium compound in accordance with the present invention is 7.0- 8.2 Å, preferably 7.1- 8.1 Å

The method for the production of the metal ion-exchanged
10 phosphorus-vanadium compound according to this invention will be described in detail below.

For the purpose of this production, this invention permits adoption of any method which is capable of obtaining a product exhibiting a specific X-ray diffraction spectrum
15 which will be described afterward. The typical methods available for the production are known in the following type kinds, depending on the water content of the vanadyl hydrogen orthophosphate hydrate ($VOHPO_4 \cdot nH_2O$). First, the raw materials to be used for the preparation of the catalyst will
20 be explained.

The vanadium source may be any of the vanadium compounds which are generally adopted for the production of phosphorus-vanadium oxides. As typical examples of the vanadium compounds, pentavalent, tetravalent, or trivalent
25 vanadium-containing compounds such as vanadium pentoxide, metavanadates, and vanadium oxyhalogenides may be cited. Among other vanadium compounds mentioned above, vanadium pentoxide proves particularly advantageous.

The phosphorus source, similarly to the vanadium source,
30 may be any of the phosphorus compounds which are generally adopted for the production of phosphorus-vanadium oxides. As typical examples of the phosphorus compounds, orthophosphoric acid, pyrophosphoric acid, phosphorous acid, polyphosphoric

acid, and phosphorus pentoxide may be cited. Among other phosphorus compounds mentioned above, orthophosphoric acid proves particularly advantageous.

Method 1 for preparation

5 Method 1 for preparation will be explained. The content of H_2O in a compound is represented by a specific amount for the sake of convenience.

 The method 1 for preparation comprises causing a vanadium compound to react with a phosphorus compound in an
10 aqueous organic solvent thereby forming vanadyl hydrogen orthophosphate hydrate such as, for example, $VOHPO_4 \cdot 0.5H_2O$, combining the vanadyl hydrogen orthophosphate hydrate with distilled water and further with an alkali salt, then adjusting the pH value of the resultant mixture to 5 - 7, filtering and
15 cleaning the mixture, and treating the cleaned mixture with an aqueous metal salt solution thereby effecting required ion exchange. Then, the resultant ion exchanger is dried and the dried ion exchanger such as, for example, $VO_{0.5}PO_4$ is calcined when necessary. More specifically, the preparation is carried
20 out by the following procedure.

 (1) Step of reaction of vanadium compound with phosphorus compound

 This step obtains vanadyl hydrogen orthophosphate hydrate such as, for example, $VOHPO_4 \cdot 0.5H_2O$ by the reaction of
25 a vanadium compound with a phosphorus compound in an aqueous solvent or organic solvent.

 The aqueous solvent or organic solvent for the reaction of the vanadium compound with the phosphorus compound is required to combine the functions of a reducing agent for
30 reducing the pentavalent vanadium compound and of a reaction solvent. Any of the aqueous organic solvents combining these functions can be used. As typical examples of these organic solvents, aliphatic alcohols such as isopropanol, 2-butanol,

isobutanol, 2-pentanol, and isopentanol; aromatic alcohols such as benzyl alcohol, methylbenzyl alcohol, dimethylbenzyl alcohol, and ethylbenzyl alcohol; and aromatic aldehydes such as benzaldehyde, tolualdehyde, dimethyl benzaldehyde, and anisaldehyde may be cited. As the organic solvent, aqueous organic solvent is preferable. It is allowable to use aliphatic alcohols, aromatic alcohols, and aromatic aldehydes as suitably mixed. Preferably, isobutanol and benzyl alcohol are used either singly or together in a mixed state. Further, when the reaction of the vanadium compound with the phosphorous compound is carried out in the aqueous solvent, concentrated hydrochloric acid, hydrazine or hydroxylamine may be used as a reducing agent which can reduce the pentavalent vanadium to tetravalent vanadium compound.

The ratio of the amounts of the phosphorus source and the vanadium source to be used is properly in the range of 0.9/1 - 1.2/1, preferably 0.95/1 - 1.1/1 as phosphorus/vanadium (atomic ratio).

The reaction of the vanadium compound with the phosphorus compound carried out in the solvent mentioned above at a temperature in the range of 60° - 150°C, preferably 80° - 140 °C, for a period in the range of 2 - 24 hours, preferably 4 - 12 hours produces the vanadyl hydrogen orthophosphate hydrate.

(2) Step for pH adjustment

This step comprises combining the produced vanadyl hydrogen orthophosphate hydrate with distilled water and further with an alkali salt, then adjusting the pH value of the resultant mixture to 5 - 7, preferably 6 - 7, and filtering and cleaning the mixture.

The pH adjustment is effected generally with an alkali solution. The vanadyl hydrogen orthophosphate hydrate and distilled water added thereto are together stirred thoroughly.

The resultant mixture, by gradual addition of the aqueous solution of such alkali as sodium hydroxide or potassium hydroxide, has the pH value thereof adjusted to 5 - 7, preferably 6 - 7. After the pH adjustment, the precipitate is immediately
5 separated by filtration and thoroughly cleaned.

(3) Step for metal ion exchange

This step comprises combining the washed precipitate with distilled water and subjecting the resultant mixture to a treatment with an aqueous metal salt solution to effect the
10 ion exchange.

The resultant precipitate is added to the aqueous solution of a metal salt and then heated at a temperature in the range of 60° - 100°C, preferably 80° - 100 °C, for a period in the range of 5 - 60 hours, preferably 15 - 60 hours, to induce
15 an ion- exchange reaction as required. The metal salts available for this treatment include nitrates, carbonates, and organic acid salts. Preferably, divalent metal salts of such organic acids as acetic acid are used. The metal is appropriately used in an amount in the range of 0.5 - 5 mols
20 per mol of the $\text{VOHPO}_4 \cdot 0.5\text{H}_2\text{O}$. As typical examples of the divalent metal, cobalt, nickel, copper, zinc, manganese, iron, magnesium, palladium, platinum, and germanium may be cited. Among other divalent metals mentioned above, cobalt, nickel, copper, and zinc prove particularly advantageous. These
25 divalent metals may be used either singly or in the form of a mixture of two or more members. The product of the ion-exchange reaction is filtered and washed to remove the excess metal salt.

(4) Step for drying

This step comprises drying the precipitate consequently
30 obtained. The precipitate is dried in the stream of an inert gas or in the current of air at a temperature in the range of 150° - 250°C, preferably 180° - 220 °C, for a period in the

range of 5 - 14 hours, preferably 12 - 20 hours, to obtain an ion exchanger such as, for example VOM_mPO_4 (wherein m represents a numerical value in the range of 0.01 - 0.5, preferably 0.1 - 0.5).

5 (5) Step for calcination

This step comprises calcining the dried precipitate after the step for drying. This step is performed optionally.

The dried ion exchanger such as, for example, $\text{VOM}_{0.5}\text{PO}_4$, is pulverized or molded and then calcined in the stream of an
 10 inert gas or in the stream of a hydrocarbon gas such as butane, preferably in the stream of a mixed gas of hydrocarbon gas with air, at a temperature in the range of $350^\circ - 600^\circ\text{C}$, preferably $400^\circ - 600^\circ\text{C}$, for a period in the approximate range of 5 - 24 hours, preferably 10 - 20 hours. A concentration of
 15 hydrocarbon gas is in the range of 0.5 - 10% by volume, preferably 5 - 10% by volume calculated, as butane to total of the hydrocarbon gas and air.

Now, Method 2 for preparation will be explained. The content of the water of crystallization, H_2O , in a compound is
 20 represented by a specific amount for the sake of convenience. The method 2 for preparation comprises refluxing a vanadium compound and a phosphorus compound in an aqueous solvent thereby forming $\text{VOPO}_4 \cdot 2.0\text{H}_2\text{O}$, reducing the produced α - $\text{VOPO}_4 \cdot 2.0\text{H}_2\text{O}$ further in an organic solvent thereby producing
 25 a vanadyl hydrogen orthophosphate hydrate such as, for example, $\text{VoHPO}_4 \cdot 1.5\text{H}_2\text{O}$, and treating this compound with an aqueous solution of a metal salt thereby inducing an ion exchange. This preparation is carried out more specifically by the following procedure.

30 (1) Step for reaction of vanadium compound with phosphorus compound

This step comprises refluxing a vanadium compound and a phosphorus compound in an aqueous solvent thereby forming

$\text{VOPO}_4 \cdot 2.0\text{H}_2\text{O}$.

The ratio of the amounts of the phosphorus source and the vanadium source, i.e. phosphorus/vanadium (atomic ratio), is in the range of 5/1 - 12/1, preferably 7/1 - 10/1.

5 The $\text{VOPO}_4 \cdot 2.0\text{H}_2\text{O}$ can be obtained by adding the vanadium source and the phosphorus source to distilled water and refluxing the resultant mixture at a temperature in the range of $60^\circ - 100^\circ\text{C}$, preferably $80^\circ - 100^\circ\text{C}$, for a period in the range of 2 - 12 hours, preferably 6 - 10 hours.

10 (2) Step for reducing $\text{VOPO}_4 \cdot 2.0\text{H}_2\text{O}$

This step comprises reducing the $\alpha\text{-VOPO}_4 \cdot 2.0\text{H}_2\text{O}$ obtained at the step mentioned above in an organic solvent thereby obtaining $\text{VOHPO}_4 \cdot 1.5\text{H}_2\text{O}$. The solvents usable for reducing $\text{VOPO}_4 \cdot 2.0\text{H}_2\text{O}$ include such 1-alkanols as 1-propanol, 1-butanol, 15 and 1-pentanol. preferably 1-butanol. The $\text{VOHPO}_4 \cdot 1.5\text{H}_2\text{O}$ is obtained by treating $\text{VOPO}_4 \cdot 2.0\text{H}_2\text{O}$ in the solvent mentioned above at a temperature in the range of $60^\circ - 150^\circ\text{C}$, preferably $80^\circ - 110^\circ\text{C}$, for a period in the range of 2 - 24 hours, preferably 8 - 12 hours.

20 (3) Step for metal ion exchange

This step comprises treating the $\text{VOHPO}_4 \cdot 1.5\text{H}_2\text{O}$ having such a structure as shown in Fig. 3 with the aqueous solution of a metal salt thereby inducing an ion-exchange reaction. The ion-exchange reaction is carried out by combining the produced 25 $\text{VOHPO}_4 \cdot 1.5\text{H}_2\text{O}$ with the aqueous solution of a metal salt and refluxing the resultant mixture at a temperature in the range of $60^\circ - 100^\circ\text{C}$, preferably $80^\circ - 100^\circ\text{C}$, for a period in the range of 5 - 60 hours, preferably 15 - 60 hours. The metal salts available for this treatment include nitrates, sulfates, 30 carbonates, and organic acid salts. Preferably, divalent metal salts of such organic acids as acetic acid are used. Appropriately, the metal is used in an amount in the range of 0.5 - 5 mols, preferably 3 - 5 mols per mol of the $\text{VOHPO}_4 \cdot 0.5\text{H}_2\text{O}$.

As typical examples of the divalent metal, cobalt, nickel, copper, zinc, manganese, iron, magnesium, palladium, platinum, and germanium may be cited. Among other divalent metals mentioned above, cobalt, nickel, copper, and zinc prove particularly advantageous. These divalent metals may be used either singly or in the form of a mixture of two or more members. The product of the ion-exchange reaction is filtered and washed to remove the excess metal salt.

(4) Step for drying

This step comprises drying the ion exchanger obtained in the form of a precipitate. The precipitate is dried in the stream of an inert gas such as nitrogen or in a stream of air at a temperature in the range of 150° - 250°C, preferably 180° - 220 °C, for a period in the range of 5 - 24 hours, preferably 12 - 20 hours to obtain VOM_mPO_4 . In this formula, m is a numerical value in the range of 0.01 - 0.5, preferably 0.1 - 0.5.

(5) Step for calcination

This step comprises calcining the dried precipitate. It is carried out optionally. The dried ion exchanger such as, for example, VOM_mPO_4 , is pulverized or molded and then calcined in the stream of an inert gas or in the stream of a hydrocarbon gas such as butane, preferably in the stream of a mixed gas of hydrocarbon gas with air, at a temperature in the range of 350° - 600°C, preferably 400 - 500 °C, for a period in the approximate range of 5 - 24 hours, preferably 10 - 20 hours. A concentration of hydrocarbon gas is in the range of 0.5 - 10% by volumes, preferably 5 - 10% by volume calculated as butane to total of the hydrocarbon gas and air.

By the steps mentioned above, the vanadium-phosphorus oxide of this invention is obtained.

When the metal ion-exchanged phosphorus-vanadium compound obtained by the method 1 for preparation or the method

2 for preparation mentioned above is used as a catalyst, the compound can be used as molded in a specific form. The molding can be carried out in the presence of a molding aid. As typical examples of the molding air to be effectively used herein, 5 inorganic substances such as silica gel, alumina sol, and talc and organic substances such as graphite and fatty acid salts may be cited. Further, the molding may be carried out in the presence of fibers of an inorganic substance. The catalyst of this invention for gas-phase oxidation can be used either 10 singly or can be used as molded together with such a carrier as silica, alumina, titania, silicon carbide, or ceramic substance or deposited on such a carrier. It has no particular limit to impose as to shape. It can be molded in the form of powder, spheres, cylinders, arcs, or saddles by such known 15 methods as tableting or extrusion molding.

As typical examples of the catalytic gas-phase oxidation reaction, the reactions for the production of maleic anhydride by the oxidation of butane, the production of methacrolein and methacrylic acid by the oxidation of isobutane, 20 the production of methacrylic acid by the oxidation of methacrolein, the production of acrylonitrile by the ammoxidation of propane, and the production of methacrylic acid by the oxydehydrogenation of isobutyric acid may be cited. The catalyst can be particularly used for the selective oxidation 25 of normal butane into maleic anhydride in the presence of molecular oxygen.

Now, this invention will be described more specifically below with reference to working examples.

Example 1

30 In 100 ml of distilled water, 6.0 g of $\text{VOHPO}_4 \cdot 0.5\text{H}_2\text{O}$ which was prepared by the reaction of vanadium pentoxide (V_2O_5) in isobutanol with 99% orthophosphoric acid at 80°C was stirred. The resultant mixture was adjusted to pH 6.5 by gradual addition

of an aqueous 1.0 mol/liter sodium hydroxide solution. The mixture was immediately filtered and repeatedly washed to remove excess sodium ions. This precipitate and 250 ml of an aqueous 0.2 mol/liter divalent cobalt acetate solution were
 5 together refluxed at 80°C for 48 hours. The refluxed mixture was filtered and washed to remove excess metal ions. The cleaned mixture was dried overnight at room temperature and further dried in a stream of argon at 200°C for 3 hours. The dried mixture was confirmed to be a divalent cobalt ion
 10 exchanger of $\text{VOM}_{0.5}\text{PO}_4$. The cobalt ion exchanger thus obtained was calcined in the stream of air-mixed gas having a normal butane concentration of 2.0 vol% at 480°C for 12 hours.

The X-ray diffraction spectrum of a dry mass of the produced cobalt ion-exchanged phosphorus-vanadium oxide
 15 ($\text{VOC}_{0.5}\text{PO}_4$) was found to have such main peaks as shown in Table 1. The X-ray diffraction spectrum of $\text{VOHPO}_4 \cdot 1.5\text{H}_2\text{O}$ was also shown in Table 1 for reference. Further, lattice constants of $\text{VOHPO}_4 \cdot 1.5\text{H}_2\text{O}$ and the cobalt ion-exchanged phosphorous-vanadium oxide are shown in Table 2. The X-ray diffraction
 20 chart of this compound was as shown in Fig. 4.

The X-ray diffraction spectrum of a sintered mass of the cobalt-exchanged phosphorus-vanadium oxide ($\text{VOC}_{0.5}\text{PO}_4$) was found to have such main peaks as shown in Table 3.

Example 2

25 $\text{VOHPO}_4 \cdot 1.5\text{H}_2\text{O}$ was prepared by refluxing vanadium pentoxide (V_2O_5) and 85% orthophosphoric acid in distilled water to form $\text{VOPO}_4 \cdot 2.0\text{H}_2\text{O}$ and refluxing the produced compound in 1-butanol at 110°C. In 250 ml of an aqueous 0.2 mol/liter divalent cobalt acetate solution, 6.0 g of the ($\text{VOHPO}_4 \cdot 1.5\text{H}_2\text{O}$)
 30 was refluxed at 80°C for 48 hours. The mixture resulting from the reflux was filtered and washed to remove excess metal ions. The cleaned mixture was dried overnight at room temperature. It was further dried in a stream of argon at 200°C for 3 hours.

This dried mixture was found to be a divalent cobalt ion exchanger of $(VO_{0.5}PO_4)$. The cobalt ion exchanger was calcined in a stream of air-mixed gas having a normal butane concentration of 2.0 vol% at 480°C for 12 hours.

5 The X-ray diffraction spectrum of a dried mass of the produced cobalt ion-exchanged phosphorus-vanadium oxide ($VOCo_{0.6}PO_4$) was found to have such main peaks as shown in Table 1. The X-ray diffraction chart of this compound was as shown in Fig. 5.

10 The X-ray diffraction spectrum of a sintered mass of the produced cobalt-exchanged phosphorus-vanadium oxide ($VOCo_{0.6}PO_4$) was found to have such main peaks as shown in Table 3.

Example 3

15 In 100 ml of distilled water, 6.0 g of $VOHPO_4 \cdot 0.5H_2O$ which was prepared by the reaction of vanadium pentoxide (V_2O_5) in 2-butanol with 99% orthophosphoric acid at 80°C was stirred. The resultant mixture was adjusted to pH 6.5 by gradual addition of an aqueous 1.0 mol/liter sodium hydroxide solution. The
20 mixture was immediately filtered and repeatedly washed to remove excess sodium ions. This precipitate and 250 ml of an aqueous 0.2 mol/liter divalent nickel acetate solution were together refluxed at 90°C for 55 hours. The refluxed mixture was filtered and washed to remove excess metal ions. The
25 cleaned mixture was dried overnight at room temperature and further dried in a stream of argon at 200°C for 3 hours. The dried mixture was confirmed to be a divalent nickel ion exchanger of $VO_{0.5}PO_4$. The nickel ion exchanger thus obtained was calcined in the steam of air-mixed gas having a normal butane
30 concentration of 2.0 vol% at 460°C for 12 hours.

The X-ray diffraction spectrum of a dry mass of the produced nickel ion-exchanged phosphorus-vanadium oxide ($VONi_{0.5}PO_4$) was found to have such main peaks as shown in Table

1. Lattice constant of the nickel ion exchanged phosphorous-vanadium oxide is shown in Table 2.

The X-ray diffraction spectrum of a sintered mass of the nickel-exchanged phosphorus-vanadium oxide ($\text{VONi}_{0.5}\text{PO}_4$) was found to have such main peaks as shown in Table 3.

Example 4

$\text{VOHPO}_4 \cdot 1.5\text{H}_2\text{O}$ was prepared by refluxing vanadium pentoxide (V_2O_5) and 85% orthophosphoric acid in distilled water to form $\text{VOPO}_4 \cdot 2.0\text{H}_2\text{O}$ and refluxing the produced compound in 1-butanol at 110°C . In 250 ml of an aqueous 0.2 mol/liter divalent nickel acetate solution, 6.0 g of the ($\text{VOHPO}_4 \cdot 1.5\text{H}_2\text{O}$) was refluxed at 85°C for 60 hours. The mixture resulting from the reflux was filtered and washed to remove excess metal ions. The cleaned mixture was dried overnight at room temperature. It was further dried in a stream of argon at 200°C for 3 hours. This dried mixture was found to be a divalent nickel ion exchanger of ($\text{VOM}_{0.5}\text{PO}_4$). The nickel ion exchanger was calcined in a stream of air-mixed gas having a normal butane concentration of 2.0 vol% at 480°C for 12 hours.

The X-ray diffraction spectrum of a dried mass of the produced nickel-exchanged phosphorus-vanadium oxide ($\text{VOC}_{0.5}\text{PO}_4$) was found to have such main peaks as shown in Table 1.

The X-ray diffraction spectrum of a sintered mass of the produced nickel-exchanged phosphorus-vanadium oxide ($\text{VOC}_{0.6}\text{PO}_4$) was found to have such main peaks as shown in Table 3.

Example 5

In 100 ml of distilled water, 6.0 g of $\text{VOHPO}_4 \cdot 0.5\text{H}_2\text{O}$ which was prepared by the reaction of vanadium pentoxide (V_2O_5) in 2-butanol with 99% orthophosphoric acid at 80°C was stirred. The resultant mixture was adjusted to pH 6.5 by gradual addition of an aqueous 1.0 mol/liter sodium hydroxide solution. The

mixture was immediately filtered and repeatedly washed to remove excess sodium ions. This precipitate and 250 ml of an aqueous 0.2 mol/liter divalent copper acetate solution were together refluxed at 90°C for 55 hours. The refluxed mixture
 5 was filtered and washed to remove excess metal ions. The cleaned mixture was dried overnight at room temperature and further dried in a stream of argon at 200°C for 3 hours. The dried mixture was confirmed to be a divalent copper ion exchanger of $\text{VOM}_{0.5}\text{PO}_4$. The copper ion exchanger thus obtained
 10 was calcined in the steam of air-mixed gas having a normal butane concentration of 2.0 vol% at 480°C for 12 hours.

The X-ray diffraction spectrum of a dry mass of the produced copper ion-exchanged phosphorus-vanadium oxide ($\text{VONi}_{0.5}\text{PO}_4$) was found to have such main peaks as shown in Table
 15 1. Lattice constant of the copper ion-exchanged phosphorous-vanadium oxide is shown in Table 2.

Example 6

$\text{VOHPO}_4 \cdot 1.5\text{H}_2\text{O}$ was prepared by refluxing vanadium pentoxide (V_2O_5) and 85% orthophosphoric acid in distilled water
 20 to form $\text{VOPO}_4 \cdot 2.0\text{H}_2\text{O}$ and refluxing the produced compound in 1-butanol at 110°C. In 250 ml of an aqueous 0.2 mol/liter divalent zinc acetate solution, 6.0 g of the ($\text{VOHPO}_4 \cdot 1.5\text{H}_2\text{O}$) was refluxed at 80°C for 30 hours. The mixture resulting from the reflux was filtered and washed to remove excess metal ions.
 25 The cleaned mixture was dried overnight at room temperature. It was further dried in a stream of argon at 200°C for 3 hours. This dried mixture was found to be a divalent zinc ion exchanger of ($\text{VOM}_{0.4}\text{PO}_4$). The zinc ion exchanger was calcined in a stream of air-mixed gas having a normal butane concentration of 2.0
 30 vol% at 480°C for 12 hours.

The X-ray diffraction spectrum of a dried mass of the produced zinc-exchanged phosphorus-vanadium oxide ($\text{VOZn}_{0.4}\text{PO}_4$) was found to have such main peaks as shown in Table 1. Lattice

constant of the zinc ion-exchanged phosphorus oxide is shown in Table 2.

Table 1

VOHPO ₄ ·1.5H ₂ O			Ni		Co		Cu		Zn	
hkl	d value (Å)	strength	d value (Å)	strength	d value (Å)	strength	d value (Å)	strength	d value (Å)	strength
001	7.97	S	7.9	W	7.85	W	7.31	VS	7.25	VS
			6.71	W	6.64	W-M			6.56	W
020	4.81	W	4.87	W	4.89	VS	4.85	S	4.6	W
			3.81	W	3.85	W	3.77	M		W
112	3.30	W	3.3	W	3.2	W				W
220	2.94	VS	2.94	W	2.98	W	2.93	S	2.98	W
202	2.72	W	2.7	W	2.71	W				W
			2.61	W	2.62	W	2.64	S	2.63	W
040	2.41	W	2.39	W	2.42	M	2.43	W	2.47	W

Table 2

Lattice Const. (Å)	a	b	c
$\text{VOHPO}_4 \cdot 1.5\text{H}_2\text{O}$	7.43	9.62	7.97
$\text{VONi}_{0.5}\text{PO}_4 \cdot 1.5\text{H}_2\text{O}$	7.43	9.74	7.90
$\text{VOCu}_{0.5}\text{PO}_4 \cdot 1.5\text{H}_2\text{O}$	7.43	9.78	7.84
$\text{VOCu}_{0.5}\text{PO}_4 \cdot 1.5\text{H}_2\text{O}$	7.43	9.70	7.31
$\text{VOZn}_{0.5}\text{PO}_4 \cdot 1.5\text{H}_2\text{O}$	7.43	9.88	7.25

- 5 A factor showing the interlayer distance of the phosphorous-vanadium compound is $hkl=1$ in Table 1 and the value of c axis about the lattice constant.

Table 3

Ni		Co	
d value (Å)	Strength	d value (Å)	Strength
5.86	W	5.79	W
4.46	W	4.43	W~M
3.45	W	3.4	W~M
3.07	S	3.04	VS
2.95	W	2.89	W~M
2.57	VS	2.55	W~M
2.14	W	2.12	W

10

Examples 7 - 10

Synthesis of maleic anhydride by oxidation of n-butane

The terms "conversion," "selectivity," and "yield" are defined as follows.

Conversion (mol%) = (Number of mols of butane consumed
5 by reaction/number of mols of butane fed to the reaction) x
100

Selectivity (mol%) = (Number of mols of formed maleic
anhydride/number of mols of butane consumed by reaction) x 100

Yield (mol%) = (Conversion Selectivity) x 100

10 A sample, 10 g in weight, of each metal ion-exchanged
vanadium-phosphorus oxides obtained in Example 1 and Examples
4 - 6 was molded in the form of pellets, 5 mm in diameter and
5 mm in height. A flow type reaction vessel was packed with
the pellets. A mixed gas comprising n-butane with air having
15 a n-butane concentration of 1.5 vol% was introduced at a spatial
velocity of 2000 hr⁻¹ into the reaction vessel to induce
gaseous-phase oxidation of n-butane. The results are shown in
Table 4.

Table 4

Example	Catalyst	Reaction temperature (°C)	Conversion (%)	Selectivity(%)
Example 7	Catalyst of Ex.1 (Co)	450	75	58
Example 8	Catalyst of Ex.4 (Ni)	450	80	52
Example 9	Catalyst of Ex.5 (Cu)	450	85	49
Example 10	Catalyst of Ex.6 (Zn)	450	80	57

20

Examples 11 - 14

Synthesis of methacrolein and methacrylic acid by oxidation of 1-butane

The terms "conversion," "selectivity," and "yield" are defined as follows.

5 Conversion (mol%) = (Number of mols of butane consumed by reaction/number of mols of butane fed to the reaction) x 100

 Selectivity of methacrolein (mol%) = (Number of mols of formed methacrolein/number of mols of butane consumed by
10 reaction) x 100

 Selectivity of methacrylic acid (mol%) = (Number of mols of formed methacrylic acid/number of mols of butane consumed by reaction) x 100

 Yield (mol%) = (Conversion x Selectivity) x 100

15 A sample, 10 g in weight, of each metal ion-exchanged vanadium-phosphorus oxides obtained in Example 1 and Examples 4 - 6 was molded in the form of pellets, 5 mm in diameter and 5 mm in height. A flow type reaction vessel was packed with the pellets. A mixed gas comprising 1-butane and steam with
20 air having a 1-butane concentration of 20 vol% and a steam concentration of 10 vol% was introduced at a spatial velocity of 2000 hr⁻¹ into the reaction vessel to induce gaseous-phase oxidation of n-butane. The results are shown in Table 5.

Table 5

Example	Catalyst	Reaction temperature (°C)	Conversion (%)	Selectivity(%)	
				Methacrolein	Methacrylic acid
Example 11	Catalyst of Ex.1 (Co)	450	15	11	3.5
Example 12	Catalyst of Ex.4 (Ni)	450	10	14	2.5
Example 13	Catalyst of Ex.5 (Cu)	450	17	19	4.0
Example 14	Catalyst of Ex.6 (Zn)	450	19	10	1.5

The entire disclose of Japanese Patent Application Nos.
5 8-52001 and 8-52002 both filed on March 8, 1996 including specification, claims, drawing and summary are incorporated herein by reference in its entirety.



NON METAL ION-EXCHANGED PHOSPHORUS-VANADIUM COMPOUND AND
SOLID ACID CATALYST USING THE COMPOUND

WHAT IS CLAIMED IS:

1. A metal ion-exchanged phosphorus-vanadium compound obtained by treating vanadyl hydrogen orthophosphate hydrate represented by the formula (1)



wherein n fulfills the expression, $0 \leq n \leq 2.0$, with an aqueous divalent metal salt solution thereby effecting the exchange of H^+ present between the layers of said vanadyl hydrogen orthophosphate hydrate and drying the resultant ion
10 exchanger.

2. A compound according to claim 1, wherein the interlayer distance is in the range of 7.0 to 8.2 Å.

3. A metal ion-exchanged phosphorus-vanadium compound obtained by treating vanadyl hydrogen orthophosphate hydrate
15 represented by the formula (1)



wherein n fulfills the expression, $0 \leq n \leq 2.0$, with an aqueous divalent metal salt solution thereby effecting the exchange of H^+ present between the layers of said vanadyl
20 hydrogen orthophosphate hydrate, drying the resultant ion exchanger, and calcining the dried ion exchanger.

4. A compound according to claim 1 or claim 3, wherein said divalent metal is at least one member selected from the group consisting of cobalt, nickel, copper, zinc, manganese,
25 iron, magnesium, palladium, platinum, and germanium.

5. A method according to claim 1 or claim 3, wherein said drying is carried out at a temperature in the range of 150° to 250 °C.

6. A method according to claim 3, wherein said
30 calcination is carried out in the presence of a hydrocarbon under an reducing atmosphere at a temperature in the range of 350° to 600 °C.

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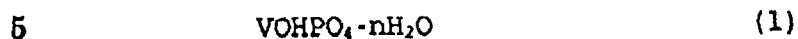
7. A solid catalyst formed of a metal ion-exchanged phosphorus-vanadium compound set forth in any of claims 1 - 6.

5 8. A catalyst for the gaseous-phase partial oxidation of a hydrocarbon, formed of a metal ion-exchanged phosphorus-vanadium compound set forth in any of claims 1 - 6.

NON METAL ION-EXCHANGED PHOSPHORUS-VANADIUM COMPOUND AND
SOLID ACID CATALYST USING THE COMPOUND

ABSTRACT OF THE DISCLOSURE

A metal ion-exchanged phosphorus-vanadium compound obtained by treating vanadyl hydrogen orthophosphate hydrate represented by the formula (1)



wherein n fulfills the expression, $0 \leq n \leq 2.0$, with an aqueous divalent metal salt solution thereby effecting the exchange of H^+ present between the layers of said vanadyl hydrogen orthophosphate hydrate, drying the resultant ion
10 exchanger, and if necessary calcining the dried ion exchanger, a solid acid and a partial oxidation catalyst for hydrocarbons containing the compound.

No drawing to accompany the abstract upon publication.

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FIG. 1

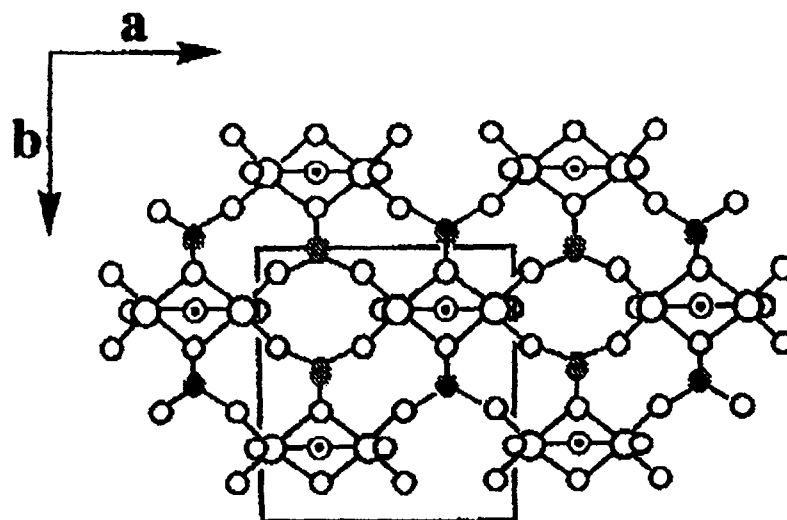


FIG. 2

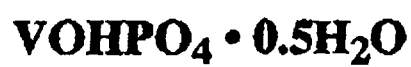
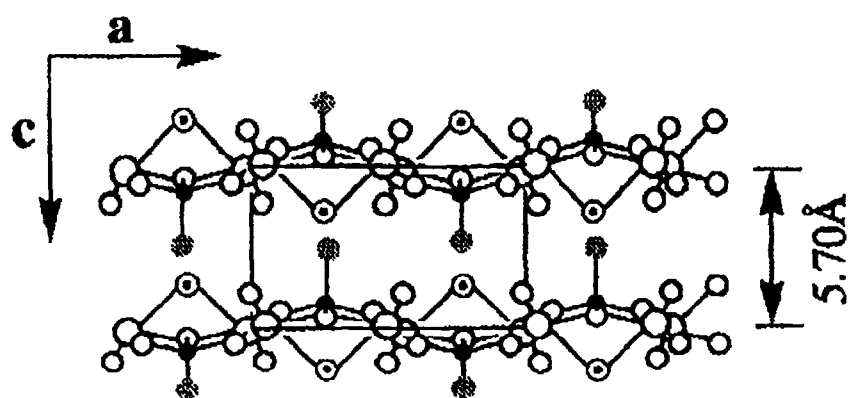


FIG. 3

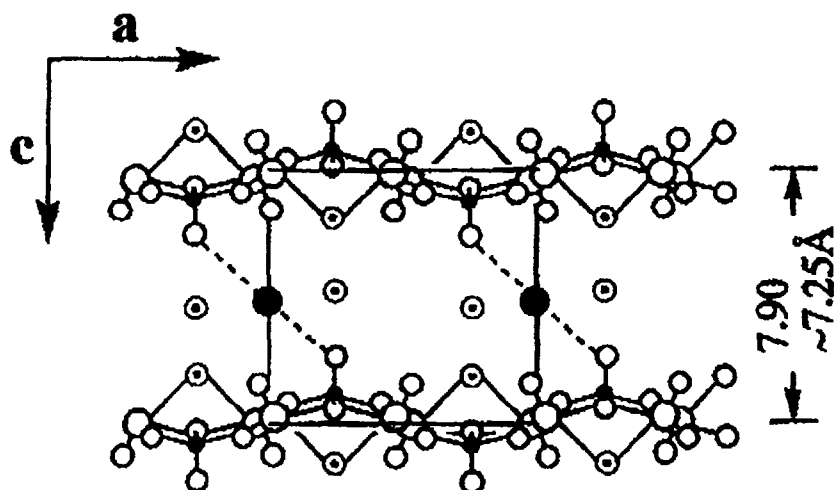
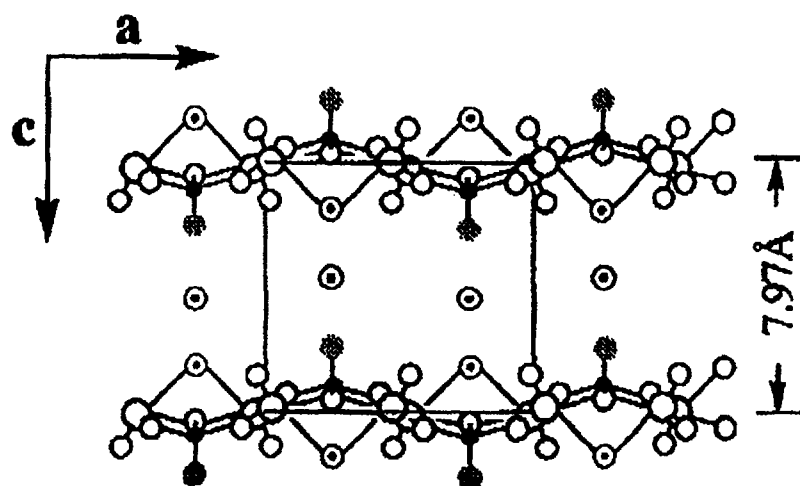


FIG. 4

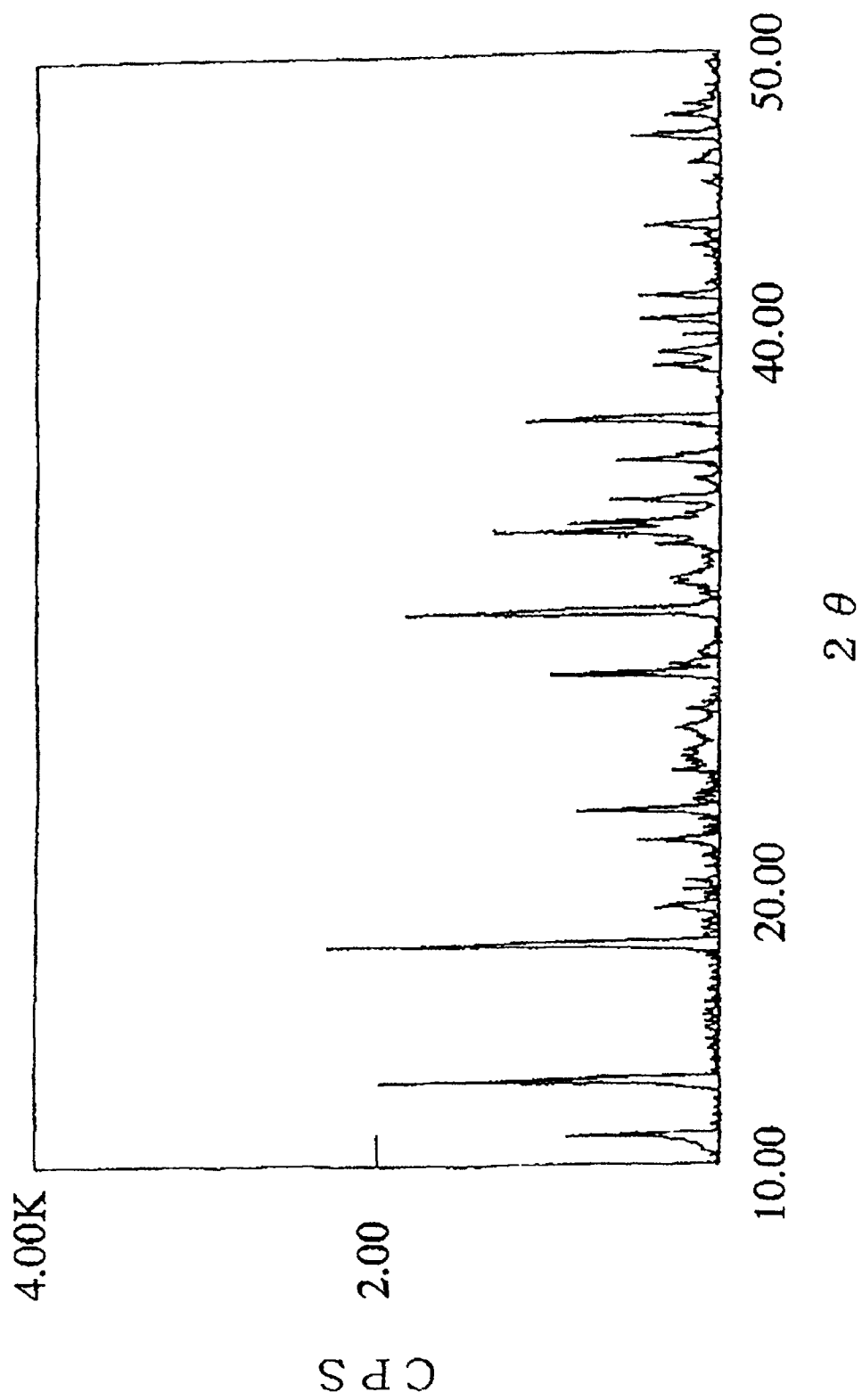
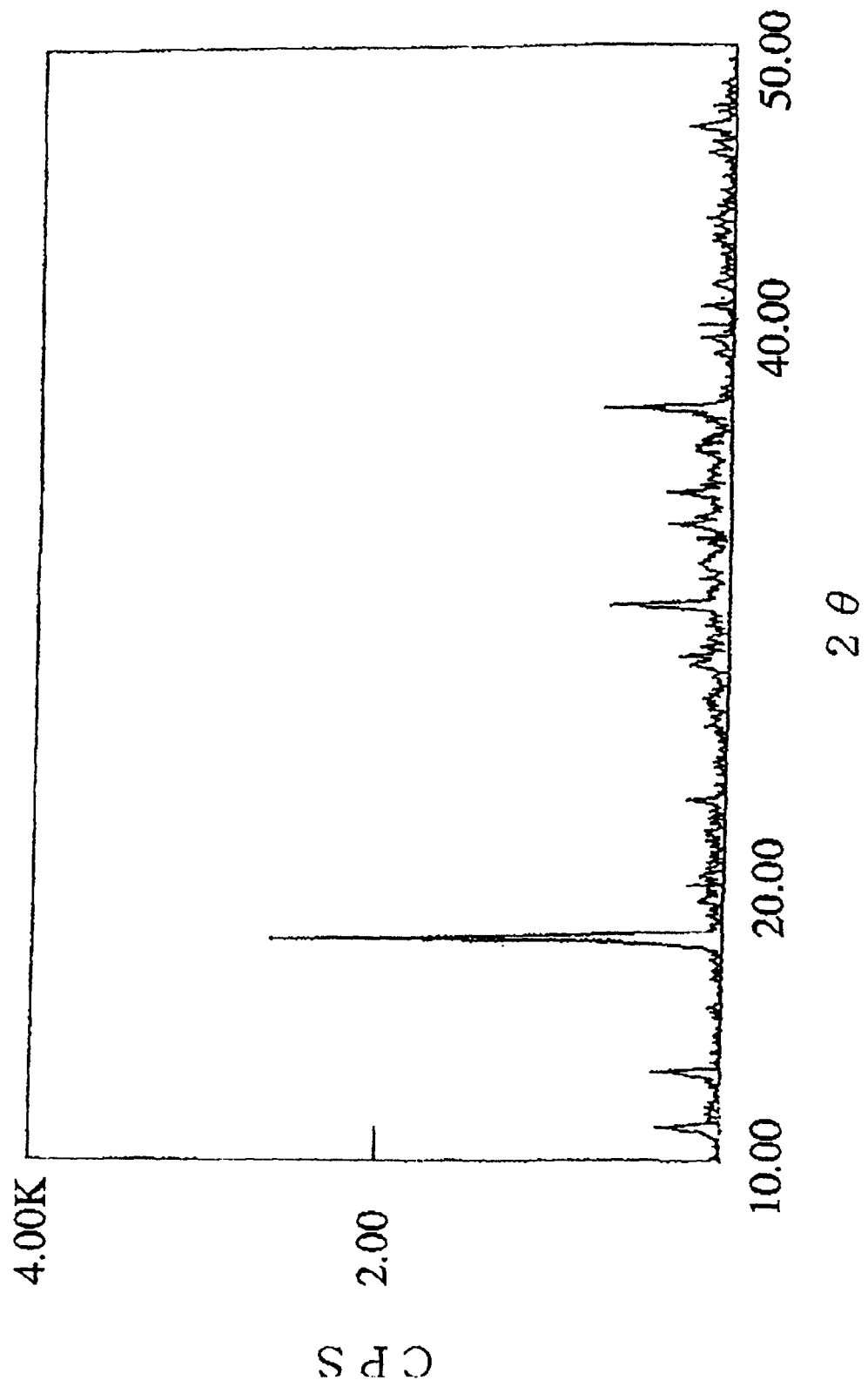


FIG. 5





TEXT OF THE ABSTRACT

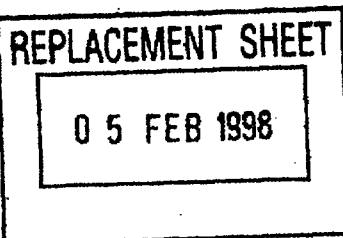
NOVEL METAL ION-EXCHANGED PHOSPHORUS-VANADIUM
COMPOUND AND SOLID ACID CATALYST USING THE COMPOUND

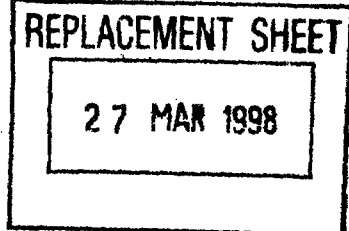
A metal ion-exchanged phosphorus-vanadium compound obtained by treating vanadyl hydrogen orthophosphate hydrate represented by the formula (1)



wherein n fulfills the expression, $0 \leq n \leq 2.0$, with an aqueous divalent metal salt solution thereby effecting the exchange of H^+ present between the layers of said vanadyl hydrogen orthophosphate hydrate, drying the resultant ion exchanger, and if necessary calcining the dried ion exchanger, a solid acid and a partial oxidation catalyst for hydrocarbons containing the compound.

No drawing to accompany the abstract upon publication.





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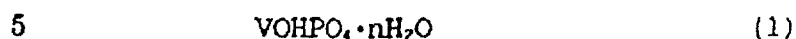
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NON METAL ION-EXCHANGED PHOSPHORUS-VANADIUM COMPOUND AND
SOLID ACID CATALYST USING THE COMPOUND



WHAT IS CLAIMED IS:

1. A metal ion-exchanged phosphorus-vanadium compound obtained by treating vanadyl hydrogen orthophosphate hydrate represented by the formula (1)



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3. A metal ion-exchanged phosphorus-vanadium compound obtained by treating vanadyl hydrogen orthophosphate hydrate
15 represented by the formula (1)



wherein n fulfills the expression, $0 \leq n \leq 2.0$, with an aqueous divalent metal salt solution thereby effecting the exchange of H^+ present between the layers of said vanadyl
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5. A compound according to claim 1 or claim 3, wherein said drying is carried out at a temperature in the range of 150° to 250 °C.

6. A compound according to claim 3, wherein said
30 calcination is carried out in the presence of a hydrocarbon under an reducing atmosphere at a temperature in the range of 350° to 600 °C.