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MERCURY ADSORBENT

REPLACED SHEETS SG

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BACKGROUND OF THE INVENTION

Field of the Invention

5 This invention relates to an activated carbon based adsorbent for mercury or mercury compounds contained in liquid hydrocarbons and method for the same, more particularly to activated carbon based adsorbents and method for adsorbing and eliminating mercury or mercury compounds (hereinafter often simplified to "mercury") contained in a slight amount in liquid hydrocarbons, for instance naphtha and intermediates of oil products or petrochemical products, by contacting such with capability of substantially complete elimination thereof from the liquid hydrocarbons of interest.

15 Description of the Related Art

 Heretofore, alumina based catalyst carrying palladium, for instance, has been used for the hydrogenation process of reforming liquid hydrocarbons, such as naphtha, wherein the hydrogenation reaction suffers damage on the catalyst if impurity mercury or mercury compounds are present in the liquid hydrocarbons.

20 Then, mercury tends to readily form amalgams with many kinds of metals. For such reason, if an apparatus constructed from aluminum based alloys is involved in such process, there is harmful risk of corrosion due to amalgamation with mercury.

25 Accordingly there has been strong desire for progress in the elimination of mercury from such liquid hydrocarbons.

There has been reported an adsorbent for mercury which bases on a porous adsorbent carrying sulfur to eliminate mercury by chemical reaction between mercury and sulfur. Further physical adsorption involving no chemical reaction with use of a porous adsorbent such as activated carbon, zeolite, alumina is feasible to eliminate inorganic mercury in hydrocarbons, but this method has such problems as inferior performance in mercury elimination rate wherein the adsorption performance exceedingly lowers in the case of a mercury concentration less than 10 bbp.

In the art disclosed heretofore concerning the adsorbent carrying sulfur, the sulfur carrying activated carbon is, for example, prepared by mixing activated carbon with fine sulfur particles and heating such mixture at 110-400°C (Japanese Patent Application Laid-open No. 59-78915/1984); activated carbon carrying organic sulfur compound (Japanese Patent Application Laid-open No. 62-114632/1987). Therein, as for choice of a sulfur compound, the use of sulfur simple solid or organic sulfur compound such as thiophene is typical, and such porous adsorbents carrying sulfur compound have been intended to mainly eliminate mercury from a gaseous objective interest, instead of elimination from liquid hydrocarbons.

Further, such art does not inhibit dissolution or desorption of the sulfur carried by the adsorbents into the liquid hydrocarbon as contamination, in addition to elimination of mercury as subject purpose. Liquid hydrocarbons are mostly subjected to the hydrogenation at the stage of intermediate pro-

duct, wherein the contaminant or impurity sulfur contained in such hydrocarbon would give serious damage to the catalysts for hydrogenation. Thus the known art is undesirable to the present invention.

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Summary of the invention

An object of the invention is to provide an adsorbent and method for effectively adsorbing and eliminating mercury contained in a slight amount in liquid hydrocarbons including the oil product or the intermediate thereof, for instance naphtha, without dissolution of such a contaminant as sulfur into the hydrocarbons.

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My studies on the art heretofore disclosed concerning elimination of mercury contained in a slight amount in liquid hydrocarbons including the oil product or intermediate, for instance naphtha, proves that the activated carbon carrying sulfur or organic sulfur compounds does not inhibit dissolution of carried sulfur compounds into liquid hydrocarbons as far as the process is performed in liquid phase. A dissolution ranges 10 to 400 ppm. The naphtha containing such sulfur concentration would give serious damage to the catalysts used in subsequent process. Thus the art known is conclusively limited to the gaseous process only, but unacceptable for the liquid process.

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The studies based on the activated carbon not carrying either sulfur or sulfur compounds is made. Then, as is known, the activated carbon is a kind of unique material having remarkably developed micropore structure and is active as non-polar ad-

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sorbent to nearly all substances. The activated carbon exhibits a certain degree of adsorption to inorganic mercury dispersed in a solution, and the adsorption of such mercury is fairly successful in the case of activated carbon having a limited micropore structure, but the activated carbon having such structure with no carrier effect is proved to be unsuccessful to adsorb and eliminate inorganic or organic mercury contained in a slight amount in liquid hydrocarbons to such extent that no harm would substantially ensue in subsequent oil process.

My studies on activated carbon carrying various substances proves that carrying hydrochloric acid is found to be remarkably superior to the activated carbon with no carrier in the subject application, wherein superior adsorption is confirmed of inorganic and organic mercury in liquid phase, but also of such mercury present in a slight amount. Further it is confirmed that where hydrochloric acid is carried in excess in the preparation, such excess hydrochloric acid can be drained out in the preparation stage and undrained or carried hydrochloric acid will not dissolve or desorb into the liquid phase in the applied adsorption process.

The studies confirm effectiveness of acids other than hydrochloric acid, and further added are studies on micropore volume as well as micropore distribution suitable to adsorbing and retaining the acids including hydrochloric acid, and these studies sustains the present invention.

The present invention provides an adsorbent carrying an acid for adsorbing mercury contained in liquid hydrocarbons.

wherein the activated carbon has preferably the micropore volume at more than 80 ml/g defined by the micropores whose radii are less than 8 angstroms. Herein, the micropore volume is defined as adsorption volume of nitrogen gas measured by the instrument
5 BEL SORP which is converted to a volume of gas at standard conditions. Such value will be referred to as micropore volume. Where the micropore volume defined by adsorbed gas volume may be expressed by denotation in liquid nitrogen, this conversion will be obtained by assuming that 1 ml of nitrogen gas at
10 standard condition, is equivalent to liquid nitrogen 0.001555 ml.

As for choice of acid carried by the activated carbon, preferred is hydrochloric acid, sulfuric acid, or phosphoric acid, and hydrochloric acid is the most preferable. As for the acti-
15 vated carbon, preferred preparation is such that carbonaceous material is activated under a condition including water vapor content at less than 30 volume % (hereafter a gas composition is expressed by volumetric percentages and often denoted by simply %) and cooled down under 300 °C in the same atmosphere, or the
20 activated carbon obtained by usual process is subjected to treatment under nitrogen gas or carbon dioxide gas with substantially no oxygen and/or water vapor at a temperature higher than 500°C and cooled down under 300 °C in the same atmosphere.

The present invention includes a preferable method of ad-
25 sorbing and eliminating mercury contained in a slight amount in hydrocarbons including oil product or intermediate thereof, such as naphtha, by containing the activated carbon carrying an

inorganic acid in liquid phase.

Detailed Description of the Invention

5 The mercury adsorbent of the invention comprises the activated carbon carrying an acid, and the base material of activated carbon is provided with a specific surface of not less than several hundreds square meters per gram and broadly speaking, a carbonaceous material having high adsorptivity is acceptable. Source materials include carbonized product of coconut
10 shell, phenol resin, and the like or coal. Activation methods include such a process as employing a high temperature with water vapor or carbon dioxide, or zinc chloride, phosphoric acid, concentrated sulfuric acid treatment.

15 As for the form, any form in crushed, particle, granule is effective, but in view of pressure drop, adsorption capacity, and handling convenience in exchange of charged loads in field work, particle or granule form is preferable. The granules are manufactured as usual by blending 100 parts of the carbonaceous material with 30-60 parts of oil pitch, coal tar or resin, and
20 by kneading, molding followed by activation. The activated carbon thus obtained is a unique material as non-polar adsorbent exhibiting superior adsorption to nearly all liquid or gaseous objectives.

25 The present invention employs the activated carbon carrying an acid wherein the acid is not limitative and broad inorganic and organic acids exhibit the subject effects wherein preferred is one which adsorbs mercury largely and will not let the acid

dissolve or desorb into liquid hydrocarbons during application or will not cause any chemical changes. These considerations endorse hydrochloric acid, sulfuric acid, phosphoric acid, and the most preferable is hydrochloric acid.

5 Improvement in the mercury adsorption attained by the activated carbon carrying hydrochloric acid as compared with the activated carbon with no carrier is shown in Table 1. appearing in the following. Example 1 and Comparative Example 1, appearing in the following, show mercury concentrations contained in
10 light naphtha and mercury adsorptions nearly at equilibrium adsorption. The data shows that, in the case of mercury concentrations at 100, 10, 1 μ g/kg, adsorbed mercury is 0.95, 0.64, 0.38 mg/g by the activated carbon carrying hydrochloric acid and 0.14, 0.03, 0.01 mg/g by the activated carbon carrying no hydrochloric acid. This shows the mercury adsorption by carrying
15 hydrochloric acid is improved 7, 21, 38 times. Based on the same carrier, the superiority in mercury adsorption by the art of carrying hydrochloric acid is made clear as compared with carrying nothing.

20 Therein remarkable is that the improvement in the adsorption is superior in the case of a lower mercury concentration in naphtha, and this result shows the inventive adsorbent is suitable to the object of the present invention in view of the fact that mercury concentrations contained in
25 liquid hydrocarbons are exceeding low.

Improvement brought by carrying an acid other than hydrochloric acid, for instance, sulfuric acid, phosphoric acid,

is shown in Example 11 (sulfuric acid) and Example 13 (phosphoric acid) in Table 3.

With respect to prevention of the carried acid from desorption, the prevention by washing out or removal of excess acid after adsorbing an inorganic acid including hydrochloric acid, sulfuric acid, phosphoric acid is feasible. such inorganic acid molecules adsorbed inside micropores of the activated carbon structure are so strongly combined that such acid molecules are not readily to desorb or become free. This feature is so remarkable with molecules of hydrochloric acid having such a small molecular size that this acid is highly resistant to desorption or liberation into liquid hydrocarbons. On the other hand, it is necessary to wash out excess acid after adsorption thereof to avoid the acid desorption which will cause possible troubles in the subsequent process. Hydrochloric acid is the most preferable in consideration of a minimal tendency of desorption and superior performance in the mercury adsorption.

Turning to the requirement for the activated carbon, broad range in properties is acceptable, however, some difference in the performance is found wherein the mechanism taken by the activated carbon carrying hydrochloric acid in adsorbing mercury is unknown and wherein the studies on relationship between the specific surface area, micropore volume, distribution of the micropores of the activated carbon and the mercury adsorption proves that the mercury adsorption is sharply correlated to a proportion of the micropore volume having smaller micropore radii.

In Examples 1, 2, and 3, employed are the activated carbons having nearly the same micropore volume but different proportions of the micropore volume having radii less than 8 angstroms which ranges from Example 1 (largest proportion) to 2 and 3. The results indicate that the proportion of the micropore having radii less than 8 angstroms is significant. That is, it is presumed that mercury is mainly adsorbed into micropores having less than 8 angstroms. In view of other results, preferred is the activated carbon provided with the micropore volume at more than 80 ml/g having the micropore radii less than 8 angstroms.

Turning to the descriptions about more detailed operations related to the present invention.

As for the activated carbon base carried with hydrochloric acid, broad range is acceptable and usual manufacture process therefor is already noted. Therein, a general tendency is such that a degree of activation is made higher, a volume of micropore having smaller radii lowers. The activation conditions are influential to increase in the total micropore volume and to accompanying decrease in the smaller micropore proportion.

As noted before, preferred is the activated carbon provided with the micropore volume at more than 80 ml/g having the micropore radii less than 8 angstroms. In the present invention, measurements on the micropore distribution and micropore volume of activated carbon bases are done by the instrument commercially named: BEL SORP, Type 28 SA manufactured by JAPAN BBL

INC. and calculated based on the adsorption isotherm line for nitrogen gas. Thereby the measurable lower limit for the micropore radii is at 3-4 angstroms and upper limit is at 60-70 angstroms. Normally the activated carbon has the total micropore volume at 100-500 ml/g and the average radius at 12-14 angstroms, and accordingly the invention prefers the activated carbon having a high proportion of the micropore volume provided with smallest radius range. Therefore, advisable activation conditions are those for depressing the tendency of enlarging the average micropore radii with enhancing the total micropore volume.

The activation gas for manufacturing the activated carbon normally contains water vapor and carbon dioxide gas and the present invention does not limit the content of carbon dioxide, but prefers the water vapor content less than 30%. Normally the water vapor not less than 40-60% is employed, since the activation rate brought by water vapor is largely higher than that by carbon dioxide and thus a gas composition is set to have more water vapor content. Compared with the normal condition, the present invention prefers a mild condition with slower activation rate.

Table 1 and Examples 1-5 in the following show that the activated carbon activated under the condition including a higher water vapor content contains less micropore volume having micropore radii less than 8 angstroms and in turn performs inferior mercury adsorption.

Detailed mechanism is unknown of why the activation con-

dition containing less water vapor content improves the mercury
adsorption, but it is presumed that such activation condition
should provide for such internal structure that the micropore
having smaller radii accounts for large proportion, and such
5 micropore portion contributes to the mercury adsorption. In
order to increase the micropore volume having micropore radii
less than 8 angstroms, preferred is to set the condition in-
cluding the water vapor content at less than 15% in the acti-
vation gas and to depress lowering the micropore volume having
10 radii less than 8 angstroms, and further preferred is the
condition to increase the activation degree.

The present invention prefers, after the activation,
cooling the activated carbon under the gas composition unchanged
or similar to that for the activation until under 300 °C and
15 then taking out from the system. Herein, the cooling gas com-
position unchanged or similar to that necessary to the acti-
vation means nitrogen gas, carbon dioxide gas or mixture thereof
or such as substantially contains no oxygen and/or water vapor,
and accordingly this requirement does not means that the acti-
20 vation gas and the cooling gas should have the same composition,
and "substantially contains no oxygen and/or water vapor" means
such gaseous condition as does not permit existence of oxygen
atoms combined on the micropore surface of the activated car-
bon, in other words, condition containing oxygen and water
25 vapor at less than 1-2%.

Further, the activated carbon preferred in the invention is
obtained by conversion from the one obtained by the normal or

usual activation condition. That is, such normal activated carbon is heated in a gas composition used for above-mentioned preferable activation condition, to a temperature higher than 500°C and cooled under 300 °C in the same gas. Herein, the upper limit of the heat treatment temperature is not limitative, however, it is preferred less than activation temperature, usually more than 850 °C. Then, "activated carbon obtained by normal or usual condition" means the one which was activated in such a atmosphere as contains water vapor at fairly higher than 15% and thereafter cooled and taken out into air before enough cooling down.

In the previous description about the conversion, treatment time of the normal activated carbon depends upon a temperature wherein 500 °C prefers 20-180 min. and 800 °C prefers several min. This conversion is thought to bring about such effect due to sintering or shrinkage of micropores which comes from the heat effect to carbonaceous structure wherein the employment of the heat treatment under such circumstance as noted bases on the thought that such would prevent loss by oxidation on the carbonaceous surface and enhance the shrinkage in micropore radii.

As for the form of the activated carbon, any form in powder, crushed, pellet, particle, fibrous, or honeycomb is acceptable. The granular or molded form is manufactured as usual by blending 100 parts of the carbonaceous material with 30-60 parts of oil pitch or coal tar as binder, and by kneading, molding followed by activation.

The step of carrying or adsorbing an acid to the activated carbon thus obtained is feasible by immersing the activated carbon mass into an acid solution of the interest and letting the acid adsorb into micropores and thereafter by washing out excess acid component by water or organic solvent for removal. In the case of an inorganic acid, such as hydrochloric acid, concentration of the solution is not limitative, but 0.1 N to 3 N (normality) is suitable. As for adsorption method, in addition to immersing into a solution as noted, impregnation by spraying of such acid solution like shower may be employed.

It is requisite to remove unadsorbed or excess inorganic acid, such as hydrochloric acid, from the activated carbon base after the carrying process. The presence of unadsorbed acid would invite dissolution thereof into liquid hydrocarbons and give damage to the catalysts used in the subsequent oil process or cause harmful corrosion due to amalgamation. The washing out is feasible by mild stirring a water or organic solvent based solution, wherein the activated carbon is immersed. Thereafter, the activated carbon carrying acid may be packed into an adsorption tower after sufficient drying or at the state of containing water at about 50 %. Otherwise, after packing such activated carbon into the adsorption tower, back washing by an organic solvent is permitted to remove the excess acid.

As noted hereinbefore, the importance in the invention lies in the micropores having radii less than 8 angstroms. Acid molecules adsorbed in such micropores will not desorb or dissolve out by washing in water or organic solvent, further such

acid molecules at the state adsorbed in the micropores will perform the superior adsorption for mercury and scarcely desorb through contact with liquid hydrocarbon.

5 In the process of adsorbing hydrochloric acid, for instance, immersing the activated carbon mass into 1-2 N. hydrochloric acid solution for about 1 hour is enough to perform the adsorption noted and in turn the hydrochloric acid molecules are so strongly combined by the adsorption that such molecules after the washing noted would not desorb through contact with liquid hydrocarbons. Test result concerning the desorption during immersing into distilled water for long time is shown at the level of trace or less than 1 mg/g. as Examples 1-5 show. It is thought that hydrochloric acid for the reason of its small molecular size is suited to be caught in exceedingly small pores. In view of these results, hydrochloric acid is the best in the inorganic acids.

15 When the activated carbon carrying acid is immersed in distilled water, though acid molecules are combined strongly, pH value reading lowers a little, and this pH lowering indicates that a carried substance is an acid. This lowering becomes clearer if temperature is raised more than 60 °C. This phenomenon is thought to be a result of the nature of a strong acid if there is only a trace amount thereof. Normally the activated carbon is manufactured from carbonaceous material including coconut shell, coal or charcoal, and plant tissues generally contain a trace amount of potassium, sodium and other metallic compounds, and these are changed to oxides, for instance

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potassium oxide, through carbonization and activation processes, and such will produce metal hydroxide in contact with water and as a result, pH value would be led to about 9-10.

Further, in the case of phenol resin as starting material, such activated carbon shows a weak acidity in pH value when immersed in distilled water, but such acidity extent is nearly equal to such resultant reading that distilled water is left in air for long time and its pH tends to acidic region, due to dissolution of carbon dioxide contained in air, and the case from the phenol resin will not drop further in pH value. However, the activated carbon carrying hydrochloric acid or sulfuric acid tends to drop under pH 6.5 when immersed in distilled water for long time. Accordingly, the identification of carrying acid can be readily done by measuring the pH value immersed in distilled water.

There is no limitation about the choice of acids carried on the activated carbon base, but preferred is as noted before an inorganic acid including hydrochloric acid, sulfuric acid or phosphoric acid, or mixture thereof, and the most preferred is hydrochloric acid.

The present invention includes the method of eliminating mercury contained in hydrocarbons. Herein, the hydrocarbon means broad range of hydrocarbons which will be subjected to elimination of mercury by solid-liquid contact process with mercury adsorbent in solid form, such objective is mainly naphtha and intermediates of oil products or petrochemical products. Typical is naphtha or other intermediate which is composed of hy-

drocarbons having about 6-15 carbon atoms and present in liquid phase at ambient temperature, and in addition, liquefied hydrocarbons derived from oil (petroleum) or coal may be applied to the method of the present invention.

5 Further, hydrocarbons having not more than 5 carbon atoms including natural gas, ethylene, propylene are in gas form at ambient temperature, but these resources may be handled in liquefied form under pressure. Such liquefied hydrocarbons may also be applied to the present method, and such hydrocarbons
10 present in gas form at ambient temperature may be converted to liquid state and then such may be applied to the present method.

 In particular, liquefied hydrocarbons having not more than 5 carbon atoms including liquefied natural gas (LNG), liquefied
15 petroleum gas (LPG), liquefied ethylene, liquefied propylene, and naphtha, are commercially handled in liquid phase and such liquid hydrocarbons may be applied to the adsorbent of the present invention. Thus remarkable industrial merit will be awarded. The objective liquid hydrocarbon may be simple one com-
20 ponent or blend of a plurality of such components.

 The method for eliminating mercury may be applied to any form of mercury contained in liquid hydrocarbons, and such includes mercury simple solid, inorganic mercury compounds, organic mercury compounds, and there is no limitation concerning
25 concentration, that is, the present invention may be applied to hydrocarbons containing mercury from the level of an amount of, to a trace of or exceedingly small amount of. The main objective

is naphtha which usually contains mercury at 0.002-10 mg/kg and such a slight level is suited to the present invention which employs the method of adsorption. In field practice of eliminating mercury, preferred is such pretreatment as filters the objective liquid hydrocarbon to remove the sludge contained therein together with mercury removable by the filtration as sludge component.

The present method for eliminating mercury is advantageously feasible by the process of employing an adsorption tower packed with the activated carbon in the form of fixed bed, wherein preferred particle size ranges from 4.75 to 0.15 mm, more preferably 1.70 to 0.50 mm.

When mercury concentration in hydrocarbon is at $100 \mu\text{g/kg}$, necessary mass of the activated carbon depends upon the allowable concentration at outlet part and the kind of adsorbent. However, roughly stated, 0.1-10 g of mercury may be eliminated with use of 1 kg of the adsorbent.

As described above, the inventive activated carbon based adsorbent is capable of nearly complete elimination of mercury contained in a slight amount in hydrocarbons by solid-liquid contact with the liquid hydrocarbons, and therein hydrochloric acid carried on the activated carbon would scarcely desorb into the objective hydrocarbons. Thus the invention is suitable to treatment of naphtha and intermediates of oil products or petrochemical products.

EXAMPLE

The invention will be further described by examples in the following.

Example 1

Coconut shells were carbonized to carbonaceous substance
5 and then crushed to 4-10 mesh (particle size ranges from 1.7 mm to 4.75 mm) which was used as raw material to prepare the activated carbon base in particle form. This carbonized material was activated at 900 °C by combustion gas of liquefied petroleum gas (gas composition: nitrogen 70%, oxygen 0.2%, carbon dioxide
10 19.8%, water vapor 10%) and cooled in the same gas circumstance under 300 °C. The activated carbon was crushed to the particle size from 10 to 32 mesh (particle size ranges from 0.5 mm to 1.7 mm). The activated carbon thus obtained had ash component (residue after strong incineration) at 2.5 wt.%.

15 The activated carbon particles obtained above were vacuum deaerated and measured of the adsorption isotherm line in nitrogen gas by the instrument (BBL SORP Type 28 SA, JAPAN BEL INC.) and thereon the total micropore volume and radius distribution of the micropores were calculated. The total micropore volume was 279 ml/g and proportion of micropore volume
20 having radii less than 8 angstroms was 43 %. That is, the micropore volume is 120 ml/g. As noted before, herein the micropore volume is always denoted by conversion from the nitrogen gas adsorbed amount to the volume at the standard condition. If converted to the volume of liquid nitrogen, the total micropore volume is 0.434 ml/g and the micropore volume
25 having radii less than 8 angstroms is 0.187 ml/g.

Then, the activated carbon particles were immersed in 1N hydrochloric acid aqueous solution for 1.5 hr at room temperature and then washed with distilled water and dried at 110 °C for 12 hrs, and then washed with light naphtha (C₆ to C₉ hydrocarbons).

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Mercury adsorption was measured at various mercury concentrations by contacting the activated carbon carrying hydrochloric acid with light naphtha containing mercury and mercury compounds, wherein organic mercury accounts for 20 % of the total mercury contained in the light naphtha. The adsorbent active carbon 10 g was immersed in the light naphtha with mild stirring and after 2 hr. residual mercury concentration in the naphtha and mercury adsorption by the active carbon were measured. The mercury adsorption performance was evaluated by measuring mercury adsorbed by the activated carbon at naphtha mercury concentrations: 100, 10, 1 μ g/kg. Quantitative determination of mercury was made by atomic adsorption analysis.

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Table 1 shows description of the activated carbon carrying hydrochloric acid and result of mercury adsorption. Under the condition which was thought to be at nearly equilibrium, the adsorbed mercury is: 0.95, 0.38 mm/g for mercury concentration: 100, 1 μ g/kg. This result shows that the subject activated carbon has superior mercury adsorption. Further, no organic mercury was found in the subject naphtha after the adsorption, that is, all organic mercury was adsorbed by the subject activated carbon.

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Note: In the Table, organic mercury adsorption is ranked:○

is good, × is fail, and overall mercury adsorption is ranked:

◎ is superior, ○ is good, △ is employable, × is fail.

Table 1

	Material of Activated Carbon	Activation Gas (Vol %)	Total Micropore volume (ml/g)	Micropores having radii less than 8 angstroms		Substance Carried on Activated Carbon	Mercury Adsorption, by Adsorbent (mg/g)			Adsorptivity of Organic Mercury	Substance Dissolved into Naphtha (mg/Kg)	Overall Adsorption Performance
				Volume Proportion (%)	Volume (ml/g)		Mercury Content in Naphtha (100 μ g/Kg)	Same as left (10 ")	Same as left (1 ")			
Example 1	Coconut Shell	10:19.8:70:0.2	279	43	120	HCl	0.95	0.64	0.38	○	Cl ⁻ 1>	◎
" 2	"	30:19.8:50:0.2	282	38	107	HCl	0.73	0.51	0.28	○	Cl ⁻ 1>	◎
" 3	"	45:19.8:35:0.2	284	25	71	HCl	0.55	0.41	0.19	○	Cl ⁻ 1>	○
" 4	"	55:19.8:25:0.2	280	15	42	HCl	0.31	0.19	0.08	○	Cl ⁻ 5	△
" 5	"	65:19.8:15:0.2	275	8	22	HCl	0.20	0.08	0.04	○	Cl ⁻ 30	△
Comparative Example 1	Coconut Shell	10:19.8:70:0.2	279	43	120	—	0.14	0.03	0.01	×	— —	×

Examples 2 to 5

The activation process herein employed activation gases having different water vapor contents: 30 % (Example 2), 45 % (Example 3), 55 % (Example 4), 65 % (Example 5). Otherwise all conditions were kept to be the same as those employed in Example 1, and thereby the activated carbon was prepared and subjected to the step of carrying hydrochloric acid and washed with distilled water and light naphtha. Thus the adsorbent, wherein activated carbon carrying hydrochloric acid, was obtained, and these were applied to the same mercury adsorption test as noted in Example 1.

Table 1 shows description of the activated carbons carrying hydrochloric acid and result of mercury adsorption. The result indicates that the water vapor content in the activation gas increases with decrease in the micropore volume of the micropores having radii less than 8 angstroms and its proportion in the total micropore volume. As for the mercury adsorption, each adsorbent performs mercury adsorption, wherein the adsorbed amount decreases with decrease in the volume of micropores having radii less than 8 angstroms.

Comparative Example 1

The adsorbent composed of the activated carbon prepared in Example 1 and with no carrier was employed to test the mercury adsorption. The result is shown in Table 1, wherein pH value when the subject activated carbon was immersed in distilled water was 8.8.

Further, adsorbed mercury amount is exceedingly low as com-

pared with Example 1. The adsorbents employed in Examples 2-5 have smaller volume of the micropores having radii less than 8 angstroms as compared with the adsorbent employed in Comparative Example 1, but the mercury adsorption performed by Comparative
5 Example 1 is lower than that by Examples 2-5. This shows that the carrying hydrochloric acid has large effect in the mercury adsorption.

Examples 6 and 7

Herein the same carbonaceous material as in Example 1 was
10 employed to prepare the activated carbon base, wherein in the activation process different conditions were employed. That is, Example 7 was subjected to milder or lower activation to prepare the activated carbon base provided with the micropore volume having radii less than 8 angstroms at 82 ml/g. And Exam-
15 ple 6 was subjected to stronger or higher activation to prepare the activated carbon base provided with the micropore volume having radii less than 8 angstroms at 108 ml/g. The steps taken subsequently were carrying hydrochloric acid and washing, drying to finish the activated carbon carrying hydrochloric
20 acid, and the mercury adsorption test was carried out in the same way as in Example 1.

Description of the activated carbons carrying hydrochloric acid and result of the mercury adsorption are shown in Table 2. Then, Example 7 employed lower activation than Example 1,
25 wherein the activated carbon of Example 7 had: the total micropore vacuum 117 ml/g, proportion of the micropore volume having radii less than 8 angstroms was increased to 70 % whose

micropore volume is 82 ml/g. Then Example 6 employed higher activation than Example 1, wherein the activated carbon of Example 6 had: the total micropore volume 327 ml/g, proportion of the micropore volume having radii less than 8 angstroms was decreased to 33 % whose micropore volume is 108 ml/g. It is to be noticed that this value is lower than that in Example 1.

Table 2

	Material of Activat- ed Carbon	Activation Gas (Vol %)	Total Micro- pore volume (ml/g)	Micropores having radii less than 8 angstroms		Subst- ance Carri- ed on Activ- ated Carbon	Mercury Adsorption, by Adsorbent (mg/g)			Adsorpt- ivity of Organic Mercury	Substance Dissolved into Naphtha (mg/Kg)	Overall Adsorpt- ion Perform- ance
				Volume Propor- tion (%)	Vol- ume (ml/ g)		Mercury Content in Naphtha (100 μ g/ Kg)	Same as left (10 ")	Same as left (1 ")			
		H ₂ O:CO ₂ :N ₂ : O ₂										
Example 6	Coconut Shell	10:19. 8:70:0. 2	327	33	108	HCl	0. 83	0. 58	0. 33	○	Cl ⁻ 1>	◎
" 7	"	10:19. 8:70:0. 2	117	70	82	HCl	0. 31	0. 15	0. 09	○	Cl ⁻ 1>	○
" 8	Phenol resin	10:19. 8:70:0. 2	564	25	141	HCl	0. 55	0. 35	0. 22	○	Cl ⁻ 1>	○
" 9	"	45:19. 8:35:0. 2	562	8	45	HCl	0. 25	0. 15	0. 09	○	Cl ⁻ 12	△
" 10	Phenol resin Fiber	30:19. 8:50:0. 2	329	45	148	HCl	1. 53	1. 10	0. 82	○	Cl ⁻ 1	◎

Further, the mercury adsorption amount in Example 7 is ranked good, and the same in Example 6 is superior, wherein in either example no organic mercury was found in naphtha after the adsorption.

5 Examples 8-10

As carbonaceous material leading to the activated carbon, carbonized particle of phenol resin (Examples 8 and 9) and carbonized fiber of phenol resin (Example 10) were employed to prepare the activated carbon base and the steps of carrying
10 hydrochloric acid, washing, drying were carried out to finish the activated carbon carrying hydrochloric acid and then the mercury adsorption test was conducted in the same way as in Example 1.

Table 2 shows description of the activated carbon carrying
15 hydrochloric acid and result of the mercury adsorption test. The adsorbent prepared in Example 10 from fibrous phenol resin had very large micropore volume having radii less than 8 angstroms at 148 ml/g, and the mercury adsorption amounts were 1.53, 0.82 mg/g for mercury concentrations in light naphtha at 100, 1 μ
20 g/kg. So superior result.

Examples 11-13

Sulfuric acid (Examples 11, 12) and phosphoric acid (Example 13) were employed as acid carried by the activated carbon obtained in Example 1 and the mercury adsorption test
25 was carried out in the same way as in Example 1. Table 3 shows description of such adsorbents carrying such inorganic acids and result of the mercury adsorption test.

Table 3

	Material of Activated Carbon	Activation Gas (Vol %)	Total Micropore volume (ml/g)	Micropores having radii less than 8 angstroms		Substance Carried on Activated Carbon	Mercury Adsorption, by Adsorbent (mg/g)			Adsorptivity of Organic Mercury	Substance Dissolved into Naphtha (mg/Kg)	Overall Adsorption Performance
				Volume Proportion (%)	Volume (ml/g)		Mercury Content in Naphtha (100 µg/Kg)	Same as left (10 ")	Same as left (1 ")			
Example 11	Coconut Shell	30:19.8:50:0.2	282	38	107	H ₂ SO ₄	0.42	0.24	0.15	○	SO ₄ ²⁻ 1	○
" 12	"	55:19.8:25:0.2	280	15	42	H ₂ SO ₄	0.20	0.08	0.02	○	SO ₄ ²⁻ 20	△
" 13	"	30:19.8:50:0.2	282	38	107	H ₃ PO ₄	0.35	0.22	0.13	○	PO ₄ ³⁻ 1	○
Comparative Example 2	Coconut Shell	10:19.8:70:0.2	279	43	120	sulfur	0.58	0.15	0.08	○	sulfur 380	×

The activated carbons employed in Examples 11 and 13 were provided with the micropore volume having radii less than 8 angstroms at 107 ml/g and performed good adsorption, but as compared with the adsorbent carrying hydrochloric acid, the mercury adsorption amount is fairly decreased. Further dissolution of acidic ions is increased. The activated carbon employed in Example 12 was provided with fairly small micropore volume having radii less than 8 angstroms at 42 ml/g, and exhibited fairly decrease in the mercury adsorption.

Comparative Example 2

Adsorbent carrying sulfur was compared. Particle activated carbon obtained in Example 1 at 100 g was mixed with powder sulfur 1 g and heated to prepare the adsorbent carrying sulfur at 1% in amount, which was subjected to the mercury adsorption test in the same way in Example 1. Table 3 shows description of the adsorbent carrying sulfur and result of mercury adsorption test.

This adsorbent carrying sulfur exhibits considerably good mercury adsorption, but exhibits a large dissolution of sulfur into naphtha as shown in Table 3. Accordingly, in view of such possible trouble or damage to catalysts used in oil process, this adsorbent is out of the consideration as mercury eliminator for naphtha.

Example 14

This illustrates a prototype field operation. A quantity of particle activated carbon carrying hydrochloric acid in the form of 10-32 mesh as prepared in Example 1 was packed in a adsorption tower (diameter: 25.4 cm, height 1 m) and light

naphtha containing mercury at 7 $\mu\text{g/kg}$ was flown or passed at a flow rate [LV value (linear velocity): 0.25 m/min]. The result of analysis showed that the light naphtha output from the tower was nearly completely freed from mercury including organic mercury. Further chlorine ions dissolved into the naphtha was less than 0.1 mg/kg or at a trace level. That is, dissolution of chlorine ion was hardly found.



What is claimed is:

1. Activated carbon based adsorbent carrying acid for adsorbing mercury or mercury compounds contained in liquid hydrocarbons.
2. Activated carbon based adsorbent carrying acid for adsorbing mercury or mercury compounds contained in liquid hydrocarbons as claimed in claim 1, wherein the activated carbon has more than 80 ml/g of micropore volume having radii less than 8 angstroms.
3. Activated carbon based adsorbent carrying acid for adsorbing mercury or mercury compounds contained in liquid hydrocarbons as claimed in claim 1, wherein an acid is one or more than one selected from the group consisting of hydrochloric acid, sulfuric acid, and phosphoric acid.
4. Activated carbon based adsorbent carrying acid for adsorbing mercury or mercury compounds contained in liquid hydrocarbons as claimed in claim 1, wherein an acid is hydrochloric acid.
5. Activated carbon based adsorbent carrying acid for adsorbing mercury or mercury compounds contained in liquid hydrocarbons as claimed in claim 1, wherein the activated carbon is manufactured by activating a carbonaceous material in condition containing water vapor at less than 30 vol.% and then cooling in the same atmosphere to less than 300°C, or wherein the activated carbon is manufactured by heat treating an activated carbon normally prepared, at a temperature over 500 °C in a gas composition comprised of nitrogen gas and/or carbon dioxide gas with no substantial presence of oxygen or water vapor, and then cooling in the same atmosphere to less than 300 °C.

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6. Method of adsorbing for elimination of mercury or mercury compounds comprising the steps of contacting activated carbon based adsorbent carrying inorganic acid with hydrocarbons containing mercury or mercury compounds in liquid phase, said hydrocarbons including naphtha and intermediates of oil products or petrochemical products.
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