

# CARRIER FOR OLEFIN POLYMERIZATION CATALYST, PREPARATION METHOD AND APPLICATION THEREOF

## **Cross Reference of Related Applications**

The present application claims the benefit of the Chinese Patent Application Nos. 200910235562.3, 200910235563.8, 200910235564.2 and CN200910235565.7, filed on October 16, 2009, which are incorporated herein by reference in their entirety and for all purposes.

## **Field of the Invention**

The present invention relates to a carrier useful in the preparation of catalyst components for olefin polymerization, a method for the preparation thereof and application thereof. More specifically, the present invention relates to a novel spheric magnesium compound carrier obtained by reacting a magnesium halide-alcohol adduct solution with an epoxy compound, a method for the preparation thereof and application thereof.

## **Background Art**

Active magnesium halide carriers useful in the preparation of catalyst components for olefin polymerization are well known in the art. A usual active magnesium halide carrier is an adduct of a magnesium halide and an alcohol, generally in the form of spheric particles. Spheric catalyst components are obtained by reacting the magnesium halide-alcohol adduct carrier with a titanium compound, such as a titanium halide, and an electron donor compound. When used in olefin polymerization, in particular in propylene polymerization, such catalyst components exhibit high polymerization activities and high stereospecificities, and the resulting polymers have good particle morphology.

The disclosed magnesium halide-alcohol adduct carriers comprise generally only magnesium dichloride and an alcohol. Some of the disclosed magnesium halide-alcohol adduct carriers further comprise a minor amount of water. Such magnesium halide-alcohol adduct carriers may be prepared by known processes, such as spray drying process, spray cooling process, high-pressure extruding process, or high-speed stirring process. See, for example, US4,421,674, US4,469,648, WO8707620, WO9311166, US5,100,849, US6,020,279, US4,399,054, EP0 395 383, EP-A-700936, CN1034736, CN1463990, CN1580136, US6,127,304 and US6,323,152.

WO98/44009 discloses an adduct of formula  $MgCl_2 \cdot mROH \cdot nH_2O$ , wherein R is a  $C_1$ - $C_{10}$  alkyl,  $2 \leq m \leq 4.2$ , and  $0 \leq n \leq 0.7$ . The X-ray diffraction pattern of the adduct is characterized in that, in the range of  $2\theta$  diffraction angles between  $5^\circ$  and  $15^\circ$ , there are three

main diffraction lines at diffraction angle  $2\theta$  of  $8.8\pm 0.2^\circ$ ,  $9.4\pm 0.2^\circ$  and  $9.8\pm 0.2^\circ$ , the intensest diffraction line being the one at  $2\theta$  of  $8.8 \pm 0.2^\circ$ , the intensity of the other two diffraction lines being at least 0.2 times the intensity of the intensest diffraction line. Besides the above-described X-ray diffraction pattern, the above adduct is further characterized by its differential scanning calorimetry (DSC) profile, in which no peaks are present at temperatures below  $90^\circ\text{C}$  or, even if peaks are present below said temperature, the fusion enthalpy associated with said peaks is less than 30% of the total fusion enthalpy.

WO2003/082930 discloses an adduct of formula  $\text{MgCl}_2 \cdot m\text{EtOH}$ , wherein  $2.5 \leq m \leq 3.2$ . The adduct optionally contains water up to 1 wt % based on the total weight of the adduct. The X-ray diffraction pattern of the adduct is characterized in that, in the range of  $2\theta$  diffraction angles between  $5^\circ$  and  $15^\circ$ , there are three main diffraction lines at diffraction angle  $2\theta$  of  $8.8\pm 0.2^\circ$ ,  $9.4\pm 0.2^\circ$  and  $9.8\pm 0.2^\circ$ , the intensest diffraction line being the one at  $2\theta$  of  $8.8 \pm 0.2^\circ$ , the intensity of the other two diffraction lines being at least 0.2 times the intensity of the intensest diffraction line. The DSC profile of the adduct is characterized by a highest melting temperature peak over  $109^\circ\text{C}$  and an associated fusion enthalpy of 103 J/g or lower.

WO2004/026920 discloses an adduct of formula  $\text{MgCl}_2 \cdot m\text{EtOH} \cdot n\text{H}_2\text{O}$ , wherein  $3.4 \leq m \leq 4.4$ , and  $0 \leq n \leq 0.7$ . The X-ray diffraction pattern of the adduct is characterized in that, in the range of  $2\theta$  diffraction angles between  $5^\circ$  and  $10^\circ$ , there are at least two diffraction lines at  $2\theta$  of  $9.3\pm 0.2^\circ$  and  $9.9\pm 0.2^\circ$ , the intensest diffraction line being the one at  $2\theta$  of  $9.3\pm 0.2^\circ$ , the intensity of the other diffraction line being lower than 0.4 times the intensity of the intensest diffraction line. The DSC profile of the adduct is characterized in that there is only one melting peak in a range of from  $90$  to  $105^\circ\text{C}$ .

Besides the above-mentioned magnesium halide-alcohol binary adduct carriers, the prior art also discloses other forms of active magnesium halide carriers. For example, CN1922212A discloses a carrier obtained by reacting a solution of a magnesium halide in a cyclic ether and an alcohol with a titanium halide. CN101190953A discloses a magnesium-containing adduct carrier formed by reacting a  $\text{C}_1\text{-C}_5$  alcohol with powdery magnesium in the presence of a methyl halide. CN1590415A discloses a complex carrier prepared by reacting a  $\text{C}_2\text{-C}_4$  lower alcohol with powdery magnesium in the presence of a methyl halide to form a homogeneous magnesium compound solution and supporting the formed magnesium compound on a spheric silica carrier. CN1016422B, CN1177868C, CN101056894A, US4,727,051, CN1255436C, US5,034,361, US5,849,655, US5,817,591 and US4,469,648 disclose active magnesium dichloride carriers prepared by using an alkoxy magnesium as a starting material.

## Summary of the Invention

After diligently studying, the inventors have found that a novel spheric magnesium compound can be obtained by reacting a magnesium halide-alcohol adduct solution with an epoxy compound. The spheric magnesium compound has a DSC profile and an X-ray diffraction pattern different from those of known magnesium dichloride-alcohol adduct carriers and magnesium dichloride carriers. The spheric magnesium compound may be used as a carrier to react with a titanium compound and an optional internal electron donor, thereby providing a spheric catalyst component for olefin polymerization with desired performance. On this basis, the present invention was made.

An object of the invention is to provide a novel spheric magnesium compound useful as a carrier used in the preparation of a catalyst component for olefin polymerization, comprising a reaction product of at least the following components: (a) a magnesium halide as defined hereinbelow; (b) an alcohol compound; and (c) an epoxy compound as defined hereinbelow.

Another object of the invention is to provide a novel spheric magnesium compound useful as a carrier used in the preparation of a catalyst component for olefin polymerization, having a characteristic DSC profile.

Another object of the invention is to provide a novel spheric magnesium compound useful as a carrier used in the preparation of a catalyst component for olefin polymerization, having a characteristic X-ray diffraction pattern.

Another object of the invention is to provide a process for preparing the spheric magnesium compound according to the invention.

Another object of the invention is to provide use of the spheric magnesium compound according to the invention as a carrier in the preparation of a catalyst component for olefin polymerization.

The spheric magnesium compound useful as a carrier used in the preparation of a catalyst component for olefin polymerization has a good particle morphology and narrow particle size distribution, and the process for the preparation thereof is simple and feasible, and consumes less energy. When used in olefin polymerization, in particular in propylene (co)polymerization, the solid catalyst component prepared by using said compound as a carrier achieves at least one of the following desired effects: high polymerization activity of catalyst, high stereospecificity of catalyst, good hydrogen response of catalyst, high stereoregularity of polymer having high melt index, and low content of polymer fines.

### **Brief Description of the Drawings**

Figure 1 shows a DSC profile of the magnesium compound carrier prepared in Example 1.

Figure 2 shows a DSC profile of a known magnesium dichloride-ethanol adduct of formula

MgCl<sub>2</sub>·2.7C<sub>2</sub>H<sub>5</sub>OH.

Figure 3 shows an X-ray diffraction pattern of the magnesium compound carrier prepared in Example 1.

Figure 4 shows an X-ray diffraction pattern of the known magnesium dichloride-ethanol adduct of formula MgCl<sub>2</sub>·2.7C<sub>2</sub>H<sub>5</sub>OH.

Figure 5 shows X-ray diffraction patterns of several carriers, wherein a is one for MgCl<sub>2</sub>; b is one for MgCl<sub>2</sub>·2.7C<sub>2</sub>H<sub>5</sub>OH; c is one for diethoxy magnesium; and d is one for the magnesium compound carrier of the invention.

Figure 6 shows a micrograph of the magnesium compound carrier prepared in Example 1.

### Detailed Description of the Preferred Embodiments

As used herein, the term “catalyst component” intends to mean main catalyst component or procatalyst, which, together with a conventional cocatalyst such as an alkyl aluminum compound and an optional external electron donor, constitutes a catalyst for olefin polymerization.

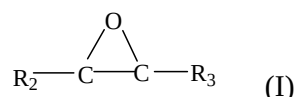
As used herein, the term “spheric magnesium compound” or “spheric carrier” means that the magnesium compound or carrier has a spheroid-like particle morphology, but does not require that the particles of the magnesium compound or carrier are in the form of perfect spheroid. Similarly, as used herein, the term “spheric catalyst component” means that the catalyst component has a spheroid-like particle morphology, but does not require that the particles of the catalyst component are in the form of perfect spheroid.

In a first aspect, the present invention provides a magnesium compound useful as a spheric carrier used in the preparation of a catalyst component for olefin polymerization, the magnesium compound comprising a reaction product of at least the following components:

(a) a magnesium halide represented by a general formula of MgX<sub>2-n</sub>R<sub>n</sub>, wherein X is independently chloride or bromide, R is a C<sub>1</sub>-C<sub>14</sub> alkyl, a C<sub>6</sub>-C<sub>14</sub> aryl, a C<sub>1</sub>-C<sub>14</sub> alkoxy, or a C<sub>6</sub>-C<sub>14</sub> aryloxy, and n is 0 or 1;

(b) an alcohol compound, preferably an alcohol compound represented by a general formula of R<sub>1</sub>OH, wherein R<sub>1</sub> is a C<sub>1</sub>-C<sub>12</sub> alkyl, a C<sub>3</sub>-C<sub>10</sub> cycloalkyl, a C<sub>7</sub>-C<sub>12</sub> aralkyl or a C<sub>6</sub>-C<sub>10</sub> aryl, and preferably a C<sub>1</sub>-C<sub>8</sub> alkyl; and

(c) an epoxy compound represented by a general formula (I):



wherein R<sub>2</sub> and R<sub>3</sub> are independently hydrogen, a C<sub>1</sub>-C<sub>5</sub> linear or branched alkyl, or a C<sub>1</sub>-C<sub>5</sub> linear or branched haloalkyl, and preferably hydrogen, a C<sub>1</sub>-C<sub>3</sub> alkyl or a C<sub>1</sub>-C<sub>3</sub> haloalkyl.

Examples of the magnesium halide compound of the general formula  $\text{MgX}_{2-n}\text{R}_n$  include, but are not limited to, magnesium dichloride, magnesium dibromide, phenoxy magnesium chloride, isopropoxy magnesium chloride, and butoxy magnesium chloride, with magnesium dichloride being preferred. The magnesium halides may be used alone or in combination.

The alcohol compound is preferably ones of the general formula of  $\text{R}_1\text{OH}$ , wherein  $\text{R}_1$  is a  $\text{C}_1\text{-C}_{12}$  alkyl, a  $\text{C}_3\text{-C}_{10}$  cycloalkyl, a  $\text{C}_7\text{-C}_{12}$  aralkyl or a  $\text{C}_6\text{-C}_{10}$  aryl, and preferably a  $\text{C}_1\text{-C}_8$  alkyl. The alcohol compound can also be glycols. Examples of the alcohol compound useful in the invention include, but are not limited to, methanol, ethanol, propanol, isopropanol, n-butanol, isobutanol, pentanol, isopentanol, n-hexanol, n-octanol, 2-ethylhexanol, ethylene glycol and propylene glycol. The alcohol compounds may be used alone or in combination.

Examples of the epoxy compound of the general formula (I) include, but are not limited to, epoxy ethane, epoxy propane, epoxy butane, epoxy chloropropane, epoxy chlorobutane, epoxy bromopropane, and epoxy bromobutane. The epoxy compounds may be used alone or in combination.

In the formation of the spheric magnesium compound, relative to one mole of the magnesium halide, the amount of the alcohol compound used may range from 4 to 40 moles, preferably from 4 to 30 moles, more preferably from 6 to 25 moles, and still more preferably from 6 to 20 moles, and the amount of the epoxy compound used may range from 1 to 10 moles, and preferably from 2 to 6 moles.

In a second aspect, the present invention provides a magnesium compound useful as a spheric carrier used in the preparation of a catalyst component for olefin polymerization, which magnesium compound has a characteristic DSC profile characterized by having a distinct exothermal peak in a temperature range of from 70 to 250 °C, said exothermal peak having a peak maximum at a temperature of from 100 to 220 °C and an associated exothermal enthalpy of larger than 40J/g.

In a preferred embodiment, the DSC profile of the magnesium compound is characterized in that the peak maximum of the exothermal peak appears at a temperature of from 100 to 200 °C.

In another preferred embodiment, the DSC profile of the magnesium compound is characterized in that the peak maximum of the exothermal peak appears at a temperature of from 130 to 210 °C.

In still another preferred embodiment, the DSC profile of the magnesium compound is characterized in that the peak maximum of the exothermal peak appears at a temperature of from 130 to 200 °C.

In a preferred embodiment, the DSC profile of the magnesium compound is characterized in that the exothermal peak has an associated exothermal enthalpy of larger than 100J/g.

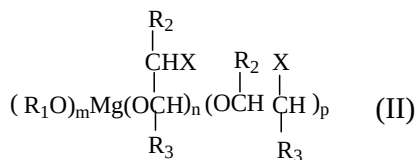
In a third aspect, the present invention provides a magnesium compound useful as a spheric carrier used in the preparation of a catalyst component for olefin polymerization, which magnesium compound has a characteristic X-ray diffraction pattern characterized in that in a  $2\theta$  angle range of from  $5^\circ$  to  $15^\circ$ , there are at least two diffraction lines, wherein the intensest diffraction line appears at a diffraction angle  $2\theta$  of  $10.0\pm0.4^\circ$ , and the secondary intensest diffraction line appears at a diffraction angle  $2\theta$  of from  $10.5$  to  $12.5^\circ$ , for example at a diffraction angle  $2\theta$  of  $11.5\pm0.4^\circ$ , and has an intensity of at least 0.2 times the intensity of the intensest diffraction line.

In an embodiment, the X-ray diffraction pattern of the magnesium compound is further characterized in that diffraction lines appeared in the  $2\theta$  angle range of from  $5^\circ$  to  $15^\circ$  other than the intensest and the secondary intensest diffraction lines have intensities of less than 0.2 times the intensity of the intensest diffraction line.

In an embodiment, the X-ray diffraction pattern of the magnesium compound is further characterized in that in a  $2\theta$  angle range of from  $15$  to  $32^\circ$ , there is a broad diffraction peak with a peak maximum in a  $2\theta$  angle range of from  $20$  to  $21^\circ$ .

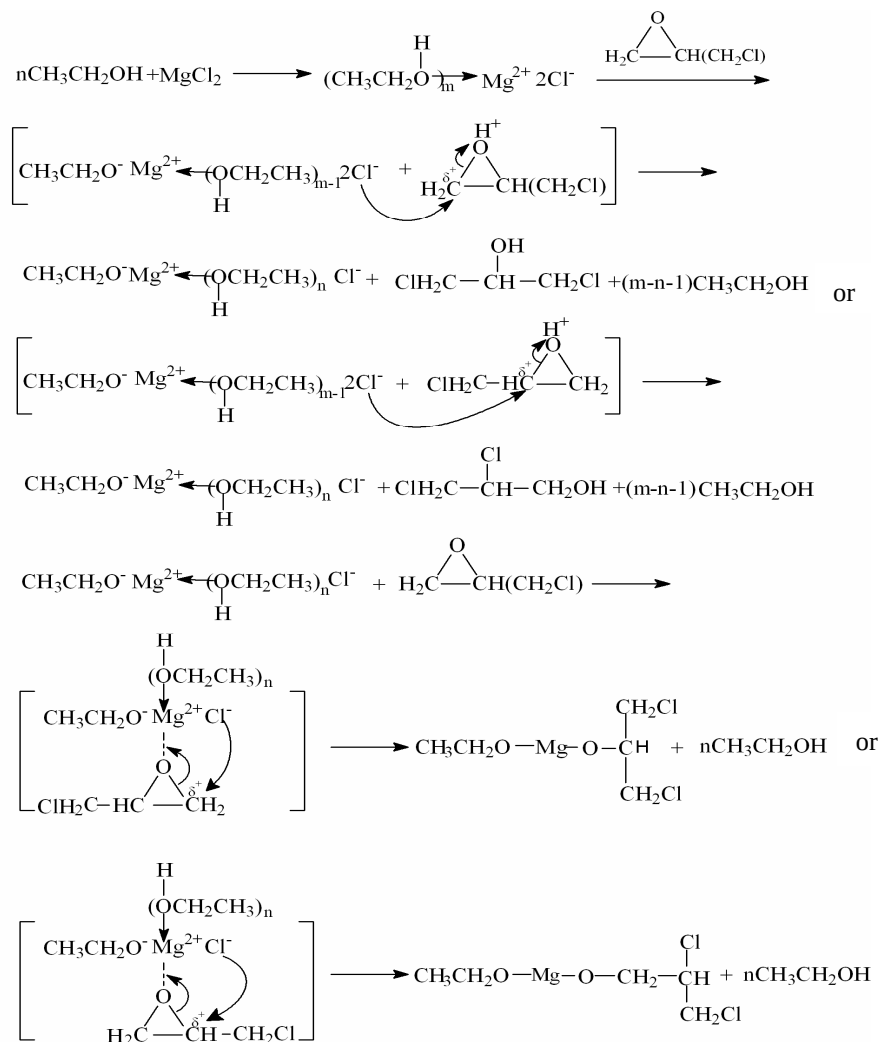
In another embodiment, the X-ray diffraction pattern of the magnesium compound is further characterized in that in a  $2\theta$  angle range of from  $15$  to  $32^\circ$ , there is a broad diffraction peak with a peak maximum in a  $2\theta$  angle range of from  $20$  to  $21^\circ$  and at least one shoulder peak at a  $2\theta$  angle of  $16.5\pm0.4^\circ$  and/or  $25.6\pm0.4^\circ$ .

Without being limited to any specific theory, it is believed that the magnesium compound of the invention prepared from the  $MgX_2$ , the  $R_1OH$  and the epoxy compound of the formula (I) has a formula:



wherein  $p + m + n = 2$ .

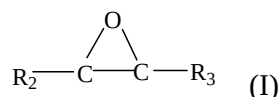
Taking a magnesium compound prepared from magnesium dichloride, epoxy chloropropane, and ethanol as an example, it is possible that the magnesium compound is formed through the following reaction mechanism:



In a fourth aspect, the present invention provides a process for preparing the spheric magnesium compound of the invention, the process comprising

a) mixing a magnesium halide of a general formula of  $\text{MgX}_{2-n}\text{R}_n$ , an alcohol compound and an optional inert liquid medium in a vessel, preferably in a closed vessel, heating the resultant mixture to a temperature of from 30 to 160 °C and allowing it to react, to form a magnesium halide-alcohol adduct solution; and

b) reacting the magnesium halide-alcohol adduct solution with an epoxy compound of the general formula (I):



at a temperature of from 30 to 160 °C, to form a particulate spheric magnesium compound,

wherein R, X, R<sub>2</sub> and R<sub>3</sub> are as defined above.

In the above process, relative to one mole of the magnesium halide, the amount of the alcohol used may range from 4 to 40 moles, preferably from 4 to 30 moles, more preferably from 6 to 25 moles, and still more preferably from 6 to 20 moles, and the amount of the epoxy compound used may range from 1 to 10 moles, and preferably from 2 to 6 moles.

The inert liquid medium may be chosen from liquid aliphatic, aromatic or alicyclic hydrocarbons, silicone oils, and mixtures thereof. Examples include, but are not limited to, hexanes, heptanes, octanes, nonanes, decanes, dodecanes, kerosenes, paraffin oils, vaseline oils, white oils, methylsilicone oils, and mixtures thereof. If the inert liquid medium is used, there is not a specific limitation to the amount thereof. However, it is preferred that the inert liquid medium is used in an amount of from 1/3 to 20 L, and preferably from 2/3 to 10 L, relative to one mole of the magnesium halide.

Examples and preferences of the magnesium halide, the alcohol compound and the epoxy compound are as described above for the first aspect.

In the above process, a trace amount of water present in the magnesium halide and/or the alcohol can be involved in the reaction to form the magnesium halide-alcohol adduct solution.

In step a) of the above process, the individual materials may be added into the vessel in any order.

The particulate spheric magnesium compound formed in step b) of the above process may be washed with an inert hydrocarbon solvent such as hexane and then dried (for example, dried under reduced pressure), as well known by those skilled in the art.

In an embodiment, the preparation process may be carried out as follows:

- 1) the magnesium halide-alcohol adduct solution is prepared by heating a mixture of the magnesium halide, the alcohol and the optional inert liquid medium in a closed vessel with stirring to a temperature of from 30 to 160 °C, and preferably from 60 to 120 °C, and allowing the mixture to react sufficiently; and

- 2) the particulate spheric magnesium compound is formed by adding the epoxy compound into the magnesium halide-alcohol adduct solution while stirring and allowing the resultant mixture to react at a temperature of from 30 to 160 °C, and preferably from 60 to 120 °C.

In another embodiment, the preparation process may be carried out as follows:

- 1) the magnesium halide-alcohol adduct solution is prepared by heating a mixture of the magnesium halide, the alcohol and the optional inert liquid medium in a closed vessel with stirring to a temperature of from 30 to 160 °C, and preferably from 60 to 120 °C, and allowing the mixture to react sufficiently; and

- 2) the particulate spheric magnesium compound is formed by adding the magnesium halide-alcohol adduct solution into a mixture of the epoxy compound and the inert liquid medium

and allowing the resultant mixture to react at a temperature of from 30 to 160 °C, and preferably from 60 to 120 °C.

The total amount of the inert liquid medium used in steps 1) and 2) ranges from 1/3 to 20 L, and preferably from 2/3 to 10 L, relative to one mole of the magnesium halide. The inert liquid medium may be distributed at any suitable ratio between steps 1) and 2). For example, the ratio of the inert liquid medium used in step 1) to that used in step 2) may range from 1:10-5:1.

In another embodiment, the preparation process may be carried out as follows:

1) the magnesium halide-alcohol adduct solution is formed by reacting the magnesium halide with the alcohol in the inert liquid medium in a closed vessel at a temperature below 60 °C with stirring; and

2) the particulate spheric magnesium compound is formed by adding the epoxy compound into the magnesium halide-alcohol adduct solution, heating the resultant mixture with stirring to a temperature of from 60 to 160 °C, and preferably from 60 to 120 °C, and allowing the mixture to react sufficiently.

In this embodiment, the amount of the alcohol used ranges preferably from 10 to 30 moles, and more preferably from 15 to 25 moles, relative to one mole of the magnesium halide.

The spheric magnesium compound of the invention is useful as a spheric carrier used in the preparation of a catalyst component for olefin polymerization. Thus, in a fifth aspect, the present invention provides use of the spheric magnesium compound as a carrier in the preparation of a catalyst component for olefin polymerization.

According to the use of the invention, a catalyst component for olefin polymerization is formed by contacting the particulate spheric magnesium compound with a titanium compound and an optional internal electron donor compound. This can be performed according to a process known per se, such as the process described in Chinese patent CN1091748A.

The titanium compound and optional internal electron donor compound used in the preparation of a catalyst component for olefin polymerization as well as their amounts are well known by those skilled in the art.

In an embodiment, the catalyst component is prepared by a process comprising the steps of: suspending the particulate spheric magnesium compound in chilled titanium tetrachloride or a mixture of titanium tetrachloride and an inert solvent, with the temperature of the liquid being generally in a range of from -30°C to 0°C, and preferably from -20°C to -10°C; heating the resulting mixture to a temperature of from 40°C to 130°C, and preferably from 80°C to 130°C, and maintaining at that temperature for 0.5 to 2.0 hours; then recovering the solids by filtration; optionally, repeating the above treatment with titanium tetrachloride one or more times, and preferably 1 to 4 times; and finally, washing the resultant solid catalyst component with an inert

solvent several times, for example, 2 to 5 times. The inert solvent is preferably an aliphatic or aromatic hydrocarbon, such as hexane, heptane, octane, decane, toluene, and the like.

Before, during or after the reaction between the particulate spheric magnesium compound and the titanium compound, at least one internal electron donor compound may be used to treat the particulate spheric magnesium compound. In particular, when the catalyst component is one intended to use in propylene polymerization, the addition of such an internal electron donor compound may be crucial in order to obtain a propylene polymer with a high isotacticity.

In the above process, relative to one mole of magnesium in the spheric magnesium compound, the amount of the internal electron donor compound used may range from 0 to 0.5 moles, and preferably from 0.05 to 0.3 moles; and the amount of the titanium compound used may range from 5 to 50 moles, and preferably from 8 to 30 moles.

When used in olefin polymerization, in particular in propylene (co)polymerization, the solid catalyst component prepared by using the magnesium compound of the invention as a carrier achieves at least one of the following desired effects: high polymerization activity of catalyst, high stereospecificity of catalyst, good hydrogen response of catalyst, high stereoregularity of polymer having high melt index, and low content of polymer fines.

### **Examples**

The following examples are provided to further illustrate the present invention and by no means intend to limit the scope thereof.

Testing methods:

1. Melt index of polymers: measured according to ASTM D1238-99, at 230 °C and 2.16kg load.
2. Isotacticity of polymers: measured by heptane extraction method carried out as follows: 2 g of dry polymer sample was extracted with boiling heptane in an extractor for 6 hours, then the residual substance was dried to constant weight, and the ratio of the weight of the residual polymer (g) to 2 (g) was regarded as isotacticity.
3. Particle size distribution: average particle size and particle size distribution of the particulate magnesium halide adducts were measured on Masters Sizer Model 2000 (manufactured by Malvern Instruments Co., Ltd.).
4. DSC profile: acquired on a DSC 7 instrument available from Perkin Elmer Co. by raising the temperature from 25 to 300 °C at a rate of 10 °C/min under nitrogen atmosphere.
5. X-ray diffraction pattern: acquired on an X'Pert MPD Model multifunctional X-ray diffractometer with a graphite monochromator and a scintillation counter available from Philips

Co., Netherlands, under the following conditions: CuK $\alpha$  ( $\lambda=1.5406 \text{ \AA}$ ), tube voltage of 40kV, tube current of 40mA, DS=SS=1° slot system, receiving slot of 0.3mm, scanning speed of 3° (2 $\theta$ )/min., and scanning range (2 $\theta$ ) of from 5° to 75°. The sample was sealed in a 50 micron-thickness polyethylene plastic bag.

#### Example 1

##### A. Preparation of spheric magnesium compound

To a 500ml reactor were charged successively 7.2g of magnesium dichloride, 180 ml of white oil and 82 ml of ethanol, and the contents were heated with stirring to 90 °C. After the contents were allowed to react at that temperature for 1 hour, 24ml of epoxy chloropropane was added to the reactor, and the reaction was allowed to continue at that temperature for 0.5 hours. After removing the liquid by filtration, the residual solids were washed with hexane five 5 times and then dried under vacuum, to give a spheric magnesium compound.

A DSC profile of the spheric magnesium compound is shown in Figure 1.

An X-ray diffraction pattern of the spheric magnesium compound is shown in Figure 3. In this X-ray diffraction pattern, in the 2 $\theta$  angle range of from 5 to 15°, there are three diffraction lines at 2 $\theta$  angles of 9.95° (100%), 11.1° (15.7%) and 11.41° (36%), and in the 2 $\theta$  angle range of from 15 to 32°, there is a broad peak with a peak maximum at 2 $\theta$  angle of 20.45° (11.3%) and shoulder peaks at 2 $\theta$  angles of 16.37° (7.7%), 25.35° (7.7%) and 30.07° (7.1%). The numerical values in the parentheses represent intensities relative to the intensest diffraction line ( $I/I_0$ ).

Figure 2 shows a DSC profile of a known magnesium dichloride-ethanol adduct of formula  $\text{MgCl}_2 \cdot 2.7\text{C}_2\text{H}_5\text{OH}$ , and Figure 4 shows an X-ray diffraction pattern of this magnesium dichloride-ethanol adduct. Figure 5 further shows X-ray diffraction patterns of several carriers, wherein a is one for  $\text{MgCl}_2$ ; b is one for  $\text{MgCl}_2 \cdot 2.7\text{C}_2\text{H}_5\text{OH}$ ; c is one for diethoxy magnesium; and d is one for the present carrier. By comparing said DSC profiles and X-ray diffraction patterns, it is apparent that the magnesium compound carrier of the invention is different from the magnesium dichloride-ethanol adduct carrier and the magnesium dichloride carrier known in the art.

##### B. Preparation of spheric catalyst component

100 ml of titanium tetrachloride was added to a 300ml glass reactor and cooled to -20 °C. Then 8g of the above-prepared spheric magnesium compound was added to the reactor, and the contents were heated to 110 °C, with 1.5ml of diisobutyl phthalate being added to the reactor during the heating. After the liquid was removed through filtration, the residual solids were washed with titanium tetrachloride twice and with hexane thrice, and then dried under vacuum to give a spheric catalyst component.

### C. Propylene polymerization

Under nitrogen atmosphere, to a 5L stainless steel autoclave were charged successively 2.5L of propylene, 1mmol of triethyl aluminum in 10 ml of hexane, 0.05mmol of methyl cyclohexyl dimethoxy silane (CHMMS) in 1 ml of hexane, 10mg of the above-prepared catalyst component and 1.5L (standard volume) of hydrogen gas. The contents were heated to 70°C, and polymerization was allowed to continue at 70°C for 1 hour. The autoclave was cooled and then the pressure was vented. The autoclave was opened and the resulting propylene polymer was recovered. The results are shown in Table 2 below.

### Example 2

A spheric magnesium compound was prepared according to the procedure described in Example 1, step A, except that the reaction temperature was 100°C.

A DSC profile of the spheric magnesium compound has an exothermal peak at a temperature range of from 75.6 to 249 °C, said exothermal peak having a peak maximum at 161.5 °C and an associated exothermal enthalpy of 304.2J/g.

The spheric magnesium compound has an X-ray diffraction pattern, wherein in the 2 $\theta$  angle range of from 5 to 15°, there are two diffraction lines at 2 $\theta$  angles of 10.1° (100%) and 11.59° (39.3%), and in the 2 $\theta$  angle range of from 15 to 32°, there is a broad peak with a peak maximum at 2 $\theta$  angle of 20.2° (43.5%) and shoulder peaks at 2 $\theta$  angles of 16.46° (9.5%), 25.40° (11%), 27.43 (7.3%) and 30.17° (9.1%). The numerical values in the parentheses represent intensities relative to the intensest diffraction line ( $I/I_0$ ).

Preparation of spheric catalyst component and propylene polymerization were conducted according to the procedures described in Example 1.

### Example 3

To a 500ml reactor were charged successively 10.5g of magnesium dichloride, 180 ml of white oil and 120 ml of ethanol, and the contents were heated with stirring to 85 °C. After the contents were allowed to react at that temperature for 1 hour, 35ml of epoxy chloropropane was added to the reactor, and the reaction was allowed to continue at that temperature for 0.5 hours. After removing the liquid by filtration, the residual solids were washed with hexane five 5 times and then dried under vacuum, to give a spheric magnesium compound.

A DSC profile of the spheric magnesium compound has an exothermal peak at a temperature range of from 77.87 to 209.83 °C, said exothermal peak having a peak maximum at 151.37 °C and an associated exothermal enthalpy of 199.31J/g.

The spheric magnesium compound has an X-ray diffraction pattern, wherein in the 2 $\theta$  angle

range of from 5 to 15°, there are two diffraction lines at 2θ angles of 10.05° (100%) and 11.55° (28.8%), and in the 2θ angle range of from 15 to 32°, there is a broad peak with a peak maximum at 2θ angle of 20.71° (13.1%) and shoulder peaks at 2θ angles of 16.36° (6.7%), 19.62 (6.3%), 25.40° (15%) and 30.0° (3.8%). The numerical values in the parentheses represent intensities relative to the intensest diffraction line (I/I<sub>0</sub>).

Preparation of spheric catalyst component and propylene polymerization were conducted according to the procedures described in Example 1.

#### Example 4

To a 300ml reactor were charged successively 4.8g of magnesium dichloride, 100 ml of decane and 30 ml of ethanol, and the contents were heated with stirring to 75 °C. After the contents were allowed to react at that temperature for 1 hour, 8ml of epoxy chloropropane was added to the reactor, and the reaction was allowed to continue at that temperature for 1 hour. After removing the liquid by filtration, the residual solids were washed with hexane five 5 times and then dried under vacuum, to give a spheric magnesium compound.

Preparation of spheric catalyst component and propylene polymerization were conducted according to the procedures described in Example 1.

#### Example 5

To a 500ml reactor were charged successively 24g of magnesium dichloride, 150 ml of methyl silicone oil and 90 ml of ethanol, and the contents were heated with stirring to 100 °C. After the contents were allowed to react at that temperature for 2 hours, the reaction mixture was transferred into an epoxy chloropropane/methyl silicone oil (40ml/350ml) mixture preheated to 100 °C, and the reaction was allowed to continue for 1 hour. After removing the liquid by filtration, the residual solids were washed with hexane five 5 times and then dried under vacuum, to give a spheric magnesium compound.

A DSC profile of the spheric magnesium compound has an exothermal peak at a temperature range of from 95.6 to 178.7 °C, said exothermal peak having a peak maximum at 137.67 °C and an associated exothermal enthalpy of 43.6J/g.

Preparation of spheric catalyst component and propylene polymerization were conducted according to the procedures described in Example 1.

#### Example 6

To a 300ml reactor were charged successively 4.8g of magnesium dichloride, 150 ml of decane and 54 ml of ethanol, and the contents were heated with stirring to 55 °C. After the

contents were allowed to react at that temperature for 1 hour, 8ml of epoxy chloropropane was added to the reactor, and the reaction mixture was then heated to 80 °C and allowed to react for 0.5 hours. After removing the liquid by filtration, the residual solids were washed with hexane five 5 times and then dried under vacuum, to give a spheric magnesium compound.

A DSC profile of the spheric magnesium compound has an exothermal peak at a temperature range of from 90.2 to 192.7 °C, said exothermal peak having a peak maximum at 137.2 °C and an associated exothermal enthalpy of 102.5J/g.

The spheric magnesium compound has an X-ray diffraction pattern, wherein in the  $2\theta$  angle range of from 5 to 15°, there are two diffraction lines at  $2\theta$  angles of 10.14° (100%) and 11.55° (31.9%), and in the  $2\theta$  angle range of from 15 to 32°, there is a broad peak with a peak maximum at  $2\theta$  angle of 20.41° (53.3%) and shoulder peaks at  $2\theta$  angles of 16.72° (11.4%), 25.44° (16.3%) and 30.15° (13.3%). The numerical values in the parentheses represent intensities relative to the intensest diffraction line ( $I/I_0$ ).

Preparation of spheric catalyst component and propylene polymerization were conducted according to the procedures described in Example 1.

#### Example 7

To a 500ml reactor were charged successively 7.2g of magnesium dichloride, 180 ml of white oil, 20ml of 2-ethylhexanol and 70 ml of ethanol, and the contents were heated with stirring to 90 °C. After the contents were allowed to react at that temperature for 1 hour, 20ml of epoxy chloropropane was added to the reactor, and the reaction was allowed to continue at that temperature for 0.5 hours. After removing the liquid by filtration, the residual solids were washed with hexane five 5 times and then dried under vacuum, to give a spheric magnesium compound.

A DSC profile of the spheric magnesium compound has an exothermal peak at a temperature range of from 73.2 to 229.3 °C, said exothermal peak having a peak maximum at 180.67 °C and an associated exothermal enthalpy of 420.4J/g.

The spheric magnesium compound has an X-ray diffraction pattern, wherein in the  $2\theta$  angle range of from 5 to 15°, there are three diffraction lines at  $2\theta$  angles of 10.0° (100%), 11.0° (17.5%) and 11.45° (23.4%), and in the  $2\theta$  angle range of from 15 to 32°, there is a broad peak with a peak maximum at  $2\theta$  angle of 20.8° (21.3%) and shoulder peaks at  $2\theta$  angles of 16.26° (5.3%), 25.3° (4.2%) and 26.4° (6.1%). The numerical values in the parentheses represent intensities relative to the intensest diffraction line ( $I/I_0$ ).

Preparation of spheric catalyst component and propylene polymerization were conducted according to the procedures described in Example 1.

#### Example 8

To a 300ml reactor were charged successively 4.8g of magnesium dichloride, 100 ml of decane and 30 ml of ethanol, and the contents were heated with stirring to 80 °C. After the contents were allowed to react at that temperature for 1 hour, 7ml of epoxy propane was added to the reactor, and the reaction was allowed to continue at that temperature for 1 hour. After removing the liquid by filtration, the residual solids were washed with hexane five 5 times and then dried under vacuum, to give a spheric magnesium compound.

A DSC profile of the spheric magnesium compound has an exothermal peak at a temperature range of from 57.5 to 236.4 °C, said exothermal peak having a peak maximum at 198.37 °C and an associated exothermal enthalpy of 265.7J/g.

Preparation of spheric catalyst component and propylene polymerization were conducted according to the procedures described in Example 1.

#### Example 9

In a 500ml reactor, a spheric magnesium compound carrier was prepared according to the procedure of Example 1. At the end of the last times of hexane washing, the liquid was removed by filtration. To the reactor was directly added 120 ml of titanium tetrachloride cooled to -20 °C, and the contents were then heated to 110 °C while stirring, with 2ml of diisobutyl phthalate being added to the reactor during the heating. After the liquid was removed through filtration, the residual solids were wished with titanium tetrachloride twice and with hexane thrice, and then dried under vacuum to give a spheric catalyst component. The obtained catalyst component has an average particle size (D50) of 60.6 microns and a particle size distribution, SPAN ((D90-D10)/D50), of 0.54.

#### Example 10

To a 500ml reactor were charged successively 7.2g of magnesium dichloride, 180 ml of white oil and 82 ml of ethanol, and the contents were heated with stirring to 95 °C. After the contents were allowed to react at that temperature for 1 hour, 30ml of epoxy chloropropane was added to the reactor, and the reaction was allowed to continue at that temperature for 0.5 hours. After removing the liquid by filtration, the residual solids were washed with hexane five 5 times and then dried under vacuum, to give a spheric magnesium compound.

The spheric magnesium compound has an X-ray diffraction pattern, wherein in the 2 $\theta$  angle range of from 5 to 15°, there are two diffraction lines at 2 $\theta$  angles of 9.8° (100%) and 10.7° (50%), and in the 2 $\theta$  angle range of from 15 to 32°, there is a broad peak with a peak maximum at 2 $\theta$  angle of 20.3° (24%) and shoulder peaks at 2 $\theta$  angles of 16. 6° (12.2%), 25.9° (8.0%), 27.1°

(5.2%), 27.86° (5.2%), and 29.85° (11.8%). The numerical values in the parentheses represent intensities relative to the intensest diffraction line ( $I/I_0$ ).

Table 1 Particle size distribution of magnesium compounds

| Example No. | Particle size distribution of spheric magnesium compounds |           |           |                                    |
|-------------|-----------------------------------------------------------|-----------|-----------|------------------------------------|
|             | D10<br>μm                                                 | D50<br>μm | D90<br>μm | Particle size distribution<br>SPAN |
| Example 1   | 58.5                                                      | 79.0      | 128.8     | 0.9                                |
| Example 2   | 62.4                                                      | 91.5      | 117.2     | 0.6                                |
| Example 3   | 75.2                                                      | 87.5      | 136.5     | 0.7                                |
| Example 4   | 152.3                                                     | 215.0     | 281.3     | 0.6                                |
| Example 5   | 14.4                                                      | 28.5      | 61.9      | 1.6                                |
| Example 6   | 25.7                                                      | 95.5      | 162.3     | 1.4                                |
| Example 7   | 39.0                                                      | 72.4      | 108.5     | 1.0                                |
| Example 8   | 124.5                                                     | 149.1     | 201.3     | 0.6                                |
| Example 10  | -                                                         | 99.7      | -         | 0.76                               |

It can be seen from the results shown in Table 1 that the spheric magnesium compound carriers of the invention have narrow particle size distribution.

Table 2 Catalyst performance

| Example No. | Polymerization activity<br>(KgPP/gCat.) | Isotacticity index of<br>polymer<br>(wt%) | Polymer melt<br>index<br>(g/10min) |
|-------------|-----------------------------------------|-------------------------------------------|------------------------------------|
| Example 1   | 37.8                                    | 97.3                                      | 12                                 |
| Example 3   | 34.6                                    | 96.8                                      | 8.1                                |
| Example 4   | 40.2                                    | 98.0                                      | 5.6                                |
| Example 5   | 19.8                                    | 97.3                                      | 7.7                                |
| Example 7   | 51.3                                    | 97.7                                      | 6.0                                |
| Example 8   | 42.8                                    | 97.6                                      | 4.8                                |
| Example 9   | 41.6                                    | 97.6                                      | 8.0                                |

It can be seen from the results shown in Table 2 that, when used in propylene polymerization, the catalysts prepared by using the spheric magnesium compounds of the invention as a carrier exhibit high polymerization activities and high stereospecificities.

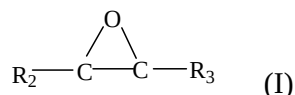
The patents, patent applications and testing methods cited in the specification are incorporated herein by reference.

While the invention has been described with reference to exemplary embodiments, it will be understood by those skilled in the art that various changes and modifications may be made without departing from the spirit and scope of the invention. Therefore, the invention is not limited to the particular embodiments disclosed as the best mode contemplated for carrying out this invention, but the invention will include all embodiments falling within the scope of the appended claims.

**What is claimed is:**

1. A particulate spheric magnesium compound comprising a reaction product of at least the following components:

- (a) a magnesium halide represented by a general formula of  $\text{MgX}_{2-n}\text{R}_n$ , wherein X is independently chloride or bromide, R is a  $\text{C}_1\text{-C}_{14}$  alkyl, a  $\text{C}_6\text{-C}_{14}$  aryl, a  $\text{C}_1\text{-C}_{14}$  alkoxy, or a  $\text{C}_6\text{-C}_{14}$  aryloxy, and n is 0 or 1;
- (b) an alcohol compound; and
- (c) an epoxy compound represented by a general formula (I):



wherein  $\text{R}_2$  and  $\text{R}_3$  are independently hydrogen, a  $\text{C}_1\text{-C}_5$  linear or branched alkyl, or a  $\text{C}_1\text{-C}_5$  linear or branched haloalkyl.

2. The magnesium compound of claim 1, wherein the alcohol compound is at least one represented by a general formula of  $\text{R}_1\text{OH}$ , wherein  $\text{R}_1$  is a  $\text{C}_1\text{-C}_{12}$  alkyl, a  $\text{C}_3\text{-C}_{10}$  cycloalkyl, a  $\text{C}_7\text{-C}_{12}$  aralkyl or a  $\text{C}_6\text{-C}_{10}$  aryl.

3. The magnesium compound of claim 2, wherein  $\text{R}_1$  is a  $\text{C}_1\text{-C}_8$  alkyl.

4. The magnesium compound of any of claims 1 to 3, wherein the magnesium halide is magnesium dichloride.

5. The magnesium compound of any of claims 1 to 4, wherein  $\text{R}_2$  and  $\text{R}_3$ , which are the same or different, represent hydrogen, a  $\text{C}_1\text{-C}_3$  alkyl or a  $\text{C}_1\text{-C}_3$  haloalkyl.

6. The magnesium compound of any of claims 1 to 5, wherein in the formation of the magnesium compound, the amount of the component (2) used ranges from 4 to 40 moles, and the amount of the component (3) used ranges from 1 to 10 moles, relative to one mole of the magnesium halide.

7. The magnesium compound of claim 6, wherein the amount of the component (2) used ranges from 6 to 20 moles, and the amount of the component (3) used ranges from 2 to 6 moles, relative to one mole of the magnesium halide.

8. A magnesium compound having a characteristic X-ray diffraction pattern characterized in that in a  $2\theta$  angle range of from  $5^\circ$  to  $15^\circ$ , there are at least two diffraction lines, wherein the intensest diffraction line appears at a diffraction angle  $2\theta$  of  $10.0 \pm 0.4^\circ$ , and the secondary intensest diffraction line appears at a diffraction angle  $2\theta$  of from  $10.5$  to  $12.5^\circ$ , and has an intensity of at least 0.2 times the intensity of the intensest diffraction line.

9. The magnesium compound of claim 8, wherein the X-ray diffraction pattern of the

magnesium compound is further characterized in that diffraction lines appeared in the  $2\theta$  angle range of from  $5^\circ$  to  $15^\circ$  other than the intensest and the secondary intensest diffraction lines have intensities of less than 0.2 times the intensity of the intensest diffraction line.

10. The magnesium compound of claim 8 or 9, wherein the X-ray diffraction pattern of the magnesium compound is further characterized in that in a  $2\theta$  angle range of from  $15^\circ$  to  $32^\circ$ , there is a broad diffraction peak with a peak maximum in a  $2\theta$  angle range of from  $20^\circ$  to  $21^\circ$ .

11. The magnesium compound of claim 8 or 9, wherein the X-ray diffraction pattern of the magnesium compound is further characterized in that in a  $2\theta$  angle range of from  $15^\circ$  to  $32^\circ$ , there is a broad diffraction peak with a peak maximum in a  $2\theta$  angle range of from  $20^\circ$  to  $21^\circ$  and at least one shoulder peak at a  $2\theta$  angle of  $16.5 \pm 0.4^\circ$  and/or  $25.6 \pm 0.4^\circ$ .

12. The magnesium compound of claim 8, wherein the secondary intensest diffraction line appears at a diffraction angle  $2\theta$  of  $11.5 \pm 0.4^\circ$ .

13. A magnesium compound having a characteristic DSC profile characterized by having a distinct exothermal peak in a temperature range of from  $70^\circ\text{C}$  to  $250^\circ\text{C}$ , said exothermal peak having a peak maximum at a temperature of from  $100^\circ\text{C}$  to  $220^\circ\text{C}$  and an associated exothermal enthalpy of larger than  $40\text{J/g}$ .

14. The magnesium compound of claim 13, wherein the DSC profile of the magnesium compound is characterized in that the peak maximum of the exothermal peak appears at a temperature of from  $100^\circ\text{C}$  to  $200^\circ\text{C}$ .

15. The magnesium compound of claim 13, wherein the DSC profile of the magnesium compound is characterized in that the peak maximum of the exothermal peak appears at a temperature of from  $130^\circ\text{C}$  to  $210^\circ\text{C}$ .

16. The magnesium compound of claim 13, wherein the DSC profile of the magnesium compound is characterized in that the peak maximum of the exothermal peak appears at a temperature of from  $130^\circ\text{C}$  to  $200^\circ\text{C}$ .

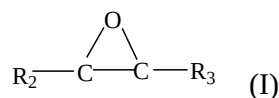
17. The magnesium compound of claim 13, wherein the DSC profile of the magnesium compound is characterized in that the exothermal peak has an associated exothermal enthalpy of larger than  $100\text{J/g}$ .

18. A process for preparing the magnesium compound of any of claims 1 to 17, comprising

a) mixing a magnesium halide of a general formula of  $\text{MgX}_{2-n}\text{R}_n$ , an alcohol compound and an optional inert liquid medium in a vessel, preferably in a closed vessel, heating the resultant mixture to a temperature of from  $30^\circ\text{C}$  to  $160^\circ\text{C}$  and allowing it to react, to form a magnesium halide-alcohol adduct solution; and

b) reacting the magnesium halide-alcohol adduct solution with an epoxy compound of the

general formula (I):



at a temperature of from 30 to 160 °C, to form a particulate spheric magnesium compound,

wherein R, X, R<sub>2</sub> and R<sub>3</sub> are as defined in claim 1.

19. The process of claim 18, having at least one of the following features:

- the alcohol compound is at least one represented by a general formula of R<sub>1</sub>OH, wherein R<sub>1</sub> is a C<sub>1</sub>-C<sub>12</sub> alkyl, a C<sub>3</sub>-C<sub>10</sub> cycloalkyl, a C<sub>7</sub>-C<sub>12</sub> aralkyl or a C<sub>6</sub>-C<sub>10</sub> aryl;

- the magnesium halide is magnesium dichloride;

- in the general formula (I), R<sub>2</sub> and R<sub>3</sub>, which are the same or different, represent hydrogen, a C<sub>1</sub>-C<sub>3</sub> alkyl or a C<sub>1</sub>-C<sub>3</sub> haloalkyl;

- the inert liquid medium is used in step a), and the amount thereof is from 1/3 L to 20 L, relative to one mole of the magnesium halide; and

- the amount of the alcohol compound used ranges from 4 to 40 moles, and the amount of the epoxy compound used ranges from 1 to 10 moles, relative to one mole of the magnesium halide.

20. The process of claim 18, having at least one of the following features:

- the inert liquid medium is used in step a), and the amount thereof is from 2/3 L to 10 L, relative to one mole of the magnesium halide; and

- the amount of the alcohol compound used ranges from 6 to 20 moles, and the amount of the epoxy compound used ranges from 2 to 6 moles, relative to one mole of the magnesium halide.