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SPECIFICATION

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CATALYSTS FOR THE AMMOXIDATION OF ALKANES

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BACKGROUND OF THE INVENTION

Field of the Invention

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The present invention is directed to a catalyst and a process for the catalytic ammoxidation of propane and isobutane to their corresponding α,β -unsaturated mononitriles; i.e., acrylonitrile and methacrylonitrile, utilizing the disclosed catalyst.

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Because of the price differential between propylene and propane, an economic incentive exists for the development of a viable catalyst and catalytic process for the conversion of propane to acrylonitrile.

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Earlier attempts in the prior art to develop an efficient process for the ammoxidation of propane to acrylonitrile produced either insufficient yields or processes that necessitated adding halogen promoters to the feed. The latter procedure would require not only reactors made of special corrosion-resistant materials, but also a quantitative recovery of the promoter. The added costs thus eliminated the advantages of the propane/propylene price differential.

Recently, U.S. Patents 5,008,427 and 5,498,588 assigned to the assignee of the instant application have been directed to novel propane ammoxidation

catalysts and the process of utilizing these catalysts to produce acrylonitrile from propane. The present invention is directed to improvements in these catalysts.

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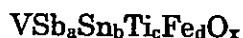
SUMMARY OF THE INVENTION

It is the object of the present invention to provide an improved catalyst for the ammoxidation of paraffins to unsaturated mononitriles, in particular, the ammoxidation of propane and isobutane to acrylonitrile or methacrylonitrile.

10 It is a further object of the present invention to provide an improved catalytic ammoxidation process for making unsaturated mononitriles from lower paraffins, in particular, the catalytic ammoxidation process for the production of acrylonitrile and methacrylonitrile from propane and isobutane, respectively.

Other objects as well as aspects, features and advantages of the present
15 invention will become apparent from the study of the accompanying disclosure and the appended claims.

According to one aspect of the invention, there is provided a catalyst comprising the elements in proportions indicated by the following empirical
20 formula:



where

$$1 \leq a \leq 1.8$$

$$0 \leq b \leq 0.35$$

25 $0 \leq c \leq 0.15$

$$0 < d \leq 0.8$$

$$0 < b + c \leq 0.5$$

$$1 \leq a - d < 1.8$$

$$\text{and when } a - d > 1.2$$

then $0 < d < 0.5$ and $0.3 \leq b + c \leq 0.5$

and when $a - d \leq 1.2$

then $0.2 < d \leq 0.8$ and $0 < b + c < 0.3$

and x is determined by the oxidation states of the cations present in the catalyst.

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In another aspect of the present invention, a process for making an α, β -unsaturated mononitrile comprises contacting in a reaction zone in the vapor phase a paraffin selected from the group consisting of propane and isobutane with oxygen and ammonia in the presence of a catalyst, the gaseous composition in the reaction zone containing a mole ratio of paraffin to ammonia in the range of from 2.5 to 16 and a mole ratio of paraffin to oxygen in the range of from about 1 to 10, said catalyst having the elements and proportions indicated by the empirical formula:

15



where

$1 \leq a \leq 1.8$

$0 \leq b \leq 0.35$

20 $0 \leq c \leq 0.15$

$0 < d \leq 0.8$

$0 < b + c \leq 0.5$

$1 \leq a - d < 1.8$

and when $a - d > 1.2$

25 then $0 < d < 0.5$ and $0.3 \leq b + c \leq 0.5$

and when $a - d \leq 1.2$

then $0.2 < d \leq 0.8$ and $0 < b + c < 0.3$

and x is determined by the oxidation states of the cations present in the catalyst.

In a further preferred embodiment of the present invention, the catalyst comprises the elements in proportions indicated by the following empirical formula:

5



where

$$1.3 \leq a \leq 1.8$$

$$0 \leq b \leq 0.2$$

10

$$0 \leq c \leq 0.1$$

$$0.2 < d \leq 0.6$$

$$0 < b + c < 0.3$$

$$1 \leq a - d \leq 1.2$$

and x is determined by the oxidation states of the cations present in the catalyst.

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In a further preferred embodiment of the present invention the catalyst comprises the elements in proportions indicated by the following empirical formula:

20



where

$$1.5 \leq a \leq 1.8$$

$$0 \leq b \leq 0.1$$

25

$$0 \leq c \leq 0.05$$

$$0.3 \leq d \leq 0.6$$

$$0 < b + c \leq 0.15$$

$$1 \leq a - d \leq 1.2$$

and x is determined by the oxidation states of the cations present in the catalyst.

In still another preferred embodiment of the present invention, the
5 calcination temperature for the preparation of the catalyst is at least 780°C. Calcination temperatures as high as 1200°C may be utilized. However, calcination temperatures usually in the range of 790° to 1050°C are preferred. The optimum calcination temperature can vary from composition to composition, but the particular narrow optimum calcination temperature range for a given
10 composition can be determined easily by routine experimentation.

In the ammoxidation process of the present invention, the reaction is carried out in the gaseous phase by contacting a mixture of the paraffin, ammonia and molecular oxygen in the reaction zone. In addition, an inert diluent such as nitrogen, helium, carbon dioxide and water may be utilized in
15 the practice of the invention.

The reaction temperature range can vary from 350° to 700°C but is usually between 430° to 520°C. The latter temperature is especially useful in the case of propane ammoxidation to acrylonitrile.

The average contact time during the process can often be from 0.01 to
20 10 seconds, but is usually from 0.02 to 10 seconds, most preferably between 0.1 to 5 seconds.

The pressure in the reaction zone usually ranges from 2 to 75 psia, most preferably, from 2 to up to 50 psia.

For specific details as to the catalyst preparation preferred in the
25 practice of the present invention, reference is made to U.S. Patent 5,008,427, assigned to the assignee of the present application and herein incorporated by reference.

DETAILED DESCRIPTION OF THE INVENTION

The catalyst and process of the present invention are herein set forth in further detail. The examples are presented for illustration purposes only and are not considered limiting.

Examples of the Invention

Table I

Example No.	Catalyst Composition	WWH (hr ⁻¹)	Temp. (°C)	Pressure (psig)	Feed Ratios				C3H8 Conv. (%)	Acrylonitrile Selectivity (%)
					C3H8	NH3	O2	N2		
1	VSb1.5Sn.05Ti.05Fe.35Ox + 20% SiO2	0.82	480	15	3.00	0.79	2.01	2.00	20.29	59.14
2	VSb1.6Sn.02Ti.04Fe.6Ox + 20% SiO2	1.60	480	15	3.00	0.81	2.00	1.99	20.17	58.35
3	VSb1.8Sn.05Ti.05Fe.6Ox + 20% SiO2	1.23	480	15	3.00	0.79	2.00	2.01	20.63	57.86
4	VSb1.61Sn.05Ti.05Fe.6Ox + 20% SiO2	1.36	480	15	3.00	0.80	1.99	1.99	20.57	57.42
5	VSb1.38Sn.05Ti.1Fe.3Ox + 20% SiO2	0.94	475	15	3.00	0.79	2.00	2.01	19.91	57.33
6	VSb1.6Sn.02Ti.02Fe.5Ox + 20% SiO2	1.62	480	15	3.00	0.81	2.00	1.99	20.14	57.28
7	VSb1.5Sn.2Ti.1Fe.1Ox + 20% SiO2	1.46	480	15	3.00	0.79	2.00	2.01	20.07	57.19
8	VSb1.38Sn.05Ti.1Fe.3Ox + 20% SiO2	1.07	480	15	3.00	0.79	2.00	2.01	20.27	57.16
9	VSb1.4Ti.15Fe.35Ox + 20% SiO2	1.05	480	15	3.00	0.79	2.01	2.00	20.66	56.86
10	VSb1.5Sn.05Ti.0.1Fe.0.35Ox + 20% SiO2	0.84	475	15	3.00	0.81	1.97	2.01	19.54	56.08
11	VSb1.4Ti.15Fe.35Ox + 20% SiO2	1.05	475	15	3.00	0.79	2.01	2.00	19.65	56.05

Comparative ExamplesTable II

Comp. Example No.	Catalyst Composition	WWH (hr ⁻¹)	Temp. (°C)	Pressure (psig)	Feed Ratios				C3H8 Conv. (%)	Acrylonitrile Selectivity (%)
					C3H8	NH3	O2	N2		
12	VSb1.38Ti1.1Fe.1Ox + 20% SiO2	0.48	485	15	3.00	0.79	2.00	2.01	20.19	55.75
13	VSb1.2Sn0.05Ti0.15Fe0.1Ox + 20%SiO2	0.67	480	15	3.00	0.80	1.96	2.06	19.59	55.70
14	VSb1.3Ti0.15Fe0.2Ox + 20 % SiO2	0.59	480	15	3.00	0.80	2.01	2.00	19.61	54.33
15	VSb1.6Fe0.5Ox + 20 % SiO2	0.67	480	15	3.00	1.02	2.02	1.99	20.10	53.80
16	VSb1.8Sn0.05Ti0.15Fe0.1Ox + 20%SiO2	0.37	480	15	3.00	0.81	1.96	2.02	20.24	53.42

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The following table provides additional examples of the invention and comparative examples. The "C" in front of the Example No. denotes a "comparative" example.

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Table III

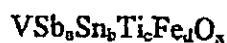
Example No.	Catalyst Composition	WWH (hr ⁻¹)	Temp. (°C)	Pressure (psig)	Feed Ratios				C3H8 Conv. (%)	Acrylonitrile Selectivity (%)
					C3H8	NH3	O2	N2		
17	VSb1.5Sn.05Ti.05Fe.35Ox + 20% SiO2	0.82	480	15	3.00	0.79	2.01	2.00	20.29	59.14
18	VSb1.5Sn.05Ti.05Fe.35Ox + 20% SiO2	0.76	480	15	3.00	0.79	2.01	2.00	21.38	59.06
19	VSb1.6Sn.02Ti.04Fe.6Ox + 20% SiO2	1.60	480	15	3.00	0.81	2.00	1.99	20.17	58.35
20	VSb1.38Ti.1Fe.3Ox + 20% SiO2	0.70	480	15	3.00	0.79	2.00	2.01	20.17	58.21
C21	VSb1.5Sn.125Ti.05Fe.225Ox + 20% SiO2	0.67	480	15	3.00	0.81	2.00	1.99	20.81	57.99
C22	VSb1.2Sn0.05Ti0.05Fe0.1Ox + 20%SiO2	0.53	480	15	3.00	0.80	1.97	2.09	19.64	57.89
23	VSb1.8Sn.05Ti.05Fe.6Ox + 20% SiO2	1.23	480	15	3.00	0.79	2.00	2.01	20.63	57.86
C24	VSb1.38Ti.1Fe.2Ox + 20% SiO2	0.66	485	15	3.00	0.79	2.00	2.01	20.14	57.85
25	VSb1.61Sn.05Ti.05Fe.6Ox + 20% SiO2	1.36	480	15	3.00	0.80	1.99	1.99	20.57	57.42
26	VSb1.38Sn.05Ti.1Fe.3Ox + 20% SiO2	0.94	475	15	3.00	0.79	2.00	2.01	19.91	57.33
27	VSb1.6Sn.02Ti.02Fe.5Ox + 20% SiO2	1.62	480	15	3.00	0.81	2.00	1.99	20.14	57.28
28	VSb1.5Sn.2Ti.1Fe.1Ox + 20% SiO2	1.46	480	15	3.00	0.79	2.00	2.01	20.07	57.19
29	VSb1.38Sn.05Ti.1Fe.3Ox + 20% SiO2	1.07	480	15	3.00	0.79	2.00	2.01	20.27	57.16
C30	VSb1.2Sn.35Ti.05Fe.1Ox + 20% SiO2	1.71	480	15	3.00	0.79	2.00	2.00	20.13	56.94
C31	VSb1.5Sn.125Ti.05Fe.225Ox + 20% SiO2	0.68	480	15	3.00	0.81	2.00	1.99	20.57	56.86
32	VSb1.4Ti.15Fe.35Ox + 20% SiO2	1.05	480	15	3.00	0.79	2.01	2.00	20.66	56.86
C33	VSb1.38Sn.05Ti.1Fe.2Ox + 20% SiO2	0.64	485	15	3.00	0.79	2.00	2.01	20.23	56.79
C34	VSb1.38Sn.05Ti.1Fe.1Ox + 20% SiO2	0.47	485	15	3.00	0.79	2.00	2.01	19.58	56.27
35	VSb1.5Sn0.05Ti0.15Fe0.35Ox + 20%SiO2	0.68	480	15	3.00	0.80	1.96	2.06	19.84	56.17
36	VSb1.5Sn0.05Ti0.1Fe0.35Ox + 20%SiO2	0.84	475	15	3.00	0.81	1.97	2.01	19.54	56.08
37	VSb1.4Ti.15Fe.35Ox + 20% SiO2	1.05	475	15	3.00	0.79	2.01	2.00	19.65	56.05
C38	VSb1.38Ti.1Fe.1Ox + 20% SiO2	0.48	485	15	3.00	0.79	2.00	2.01	20.19	55.75
C39	VSb1.2Sn0.05Ti0.15Fe0.1Ox + 20%SiO2	0.67	480	15	3.00	0.80	1.96	2.06	19.59	55.70
40	VSb1.5Sn0.05Ti0.1Fe0.35Ox + 20%SiO2	0.84	480	15	3.00	0.81	1.97	2.01	20.78	55.65
C41	VSb1.5Sn.35Ti.1Fe.35Ox + 20% SiO2	1.71	480	15	3.00	0.79	2.00	2.00	19.96	55.58

Example No.	Catalyst Composition	WWH (hr ⁻¹)	Temp. (°C)	Pressure (psig)	Feed Ratios				C3H8 Conv. (%)	Acrylonitrile Selectivity (%)
					C3H8	NH3	O2	N2		
C42	VSb1.2Sn.35Ti.15Fe.1Ox + 20% SiO2	1.79	480	15	3.00	0.79	2.00	2.00	20.04	55.39
43	V1.0Sb1.5Sn0.2Ti0.05Fe0.35Ox + 20%SiO2	1.30	480	15	3.00	0.82	1.97	2.04	20.96	55.35
C44	VSb1.8Sn.04Ti.02Fe.5Ox + 20% SiO2	1.56	480	15	3.00	0.81	2.00	1.99	20.29	54.92
45	VSb1.8Sn.02Ti.02Fe.6Ox + 20% SiO2	1.28	480	15	3.00	0.82	2.01	2.00	20.09	54.81
C46	VSb1.3Ti0.15Fe0.2Ox + 20 % SiO2	0.59	480	15	3.00	0.80	2.01	2.00	19.61	54.33
C47	VSb1.5Sn0.2Ti0.15Fe0.35Ox + 20%SiO2	1.29	480	15	3.00	0.82	1.97	2.04	19.99	54.00
C48	VSb1.5Sn.125Ti.05Fe.1Ox + 20% SiO2	0.45	480	15	3.00	0.80	1.99	1.99	20.51	53.83
C49	VSb1.6Fe0.5Ox + 20 % SiO2	0.67	480	15	3.00	1.02	2.02	1.99	20.10	53.80
C50	VSb1.8Sn.35Ti.15Fe.1Ox + 20% SiO2	1.26	480	15	3.00	0.79	2.00	2.00	19.50	53.79
C51	VSb1.5Sn.125Ti.05Fe.1Ox + 20% SiO2	0.44	485	15	3.00	0.80	1.99	1.99	21.03	53.72
C52	VSb1.8Sn.04Ti.02Fe.5Ox + 20% SiO2	1.59	480	15	3.00	0.81	2.00	1.99	20.09	53.67
53	VSb1.7Ti0.1Fe0.6 + 20% SiO2	1.34	480	16	3.00	0.80	2.00	1.98	20.10	53.60
C54	VSb1.8Sn0.05Ti0.15Fe0.1Ox + 20%SiO2	0.37	480	15	3.00	0.81	1.96	2.02	20.24	53.42
C55	VSb1.5Sn.125Ti.05Fe.1Ox + 20% SiO2	0.44	488	15	3.00	0.80	1.99	1.99	21.36	53.28
C56	VSb1.5Sn0.2Ti0.1Fe0.35Ox + 20%SiO2	1.49	475	15	3.00	0.81	1.97	2.01	20.35	53.10
57	VSb1.5Ti.15Fe.4Ox + 20% SiO2	1.08	475	15	3.00	0.79	2.01	2.00	20.71	52.76
58	VSb1.8Sn.02Ti.02Fe.6Ox + 20% SiO2	1.14	480	15	3.00	0.82	2.01	2.00	20.71	52.64
C59	VSb1.8Sn.35Ti.05 Fe.6Ox + 20% SiO2	2.32	480	15	3.00	0.79	2.00	2.01	19.85	51.27

While the invention has been described in conjunction with the specific embodiments thereof, it is evident that many alternatives, modifications and variations will be apparent to those skilled in the art in light of the foregoing description. Accordingly, it is intended to embrace all such alternatives, modifications and variations as fall within the spirit and scope of the appended claims.

CLAIMS:

1. An ammoxidation catalyst characterized by the following empirical formula comprising:



where

$$1 \leq a \leq 1.8$$

$$0 \leq b \leq 0.35$$

$$0 \leq c \leq 0.15$$

$$0 < d \leq 0.8$$

$$0 < b + c \leq 0.5$$

$$1 \leq a - d < 1.8$$

and when $a - d > 1.2$

then $0 < d < 0.5$ and $0.3 \leq b + c \leq 0.5$

and when $a - d \leq 1.2$

then $0.2 < d \leq 0.8$ and $0 < b + c < 0.3$

and x is determined by the oxidation states of the cations present in the catalyst.

2. The catalyst of claim 1 wherein $1.3 \leq a \leq 1.8$.
3. The catalyst of claim 2 wherein $0.2 < d \leq 0.6$.
4. The catalyst of claim 3 wherein $1 \leq a - d \leq 1.2$.
5. The catalyst of claim 1 wherein $1.5 \leq a \leq 1.8$.
6. The catalyst of claim 3 wherein $0.3 \leq d \leq 0.6$.
7. The catalyst of claim 1 wherein the calcination temperature for the preparation of the catalyst is at least 780°C.
8. The catalyst of claim 2 wherein the calcination temperature for the preparation of the catalyst is at least 780°C.

9. The catalyst of claim 3 wherein the calcination temperature for the preparation of the catalyst is at least 780°C.

10. The catalyst of claim 4 wherein the calcination temperature for the preparation of the catalyst is at least 780°C.

11. The catalyst of claim 5 wherein the calcination temperature for the preparation of the catalyst is at least 780°C.

12. The catalyst of claim 6 wherein the calcination temperature for the preparation of the catalyst is at least 780°C.

13. An ammoxidation catalyst characterized by the following empirical formula comprising:



where

$$1.3 \leq a \leq 1.8$$

$$0 \leq b \leq 0.2$$

$$0 \leq c \leq 0.1$$

$$0.2 < d \leq 0.6$$

$$0 < b + c < 0.3$$

$$1 \leq a - d \leq 1.2$$

and x is determined by the oxidation states of the cations present in the catalyst.

14. The catalyst of claim 13 wherein the calcination temperature for the preparation of the catalyst is at least 780°C.

15. An ammoxidation catalyst characterized by the following empirical formula:



where

$$1.5 \leq a \leq 1.8$$

$$0 \leq b \leq 0.1$$

$$0 \leq c \leq 0.05$$

$$0.3 \leq d \leq 0.6$$

$$0 < b + c \leq 0.15$$

$$1 \leq a - d \leq 1.2$$

and x is determined by the oxidation states of the cations present in the catalyst.

16. The catalyst of claim 15 wherein the calcination temperature for the preparation of the catalyst is at least 780°C.