

**DEHYDRATION OF ION-EXCHANGE RESIN
CATALYZED REACTIONS**

BACKGROUND OF THE INVENTION

5 **Field of the Invention**

The invention relates to processes and apparatus for condensation of a bisphenol with a ketone, in the presence of acidic ion-exchange resins as catalysts.

10 **Brief Description of Related Art**

The preparation of bisphenols by the condensation of monocyclic phenols with a ketone in the presence of an acid catalyst is a widely known procedure; see for example the description in U.S. Patent 2,858,342. As recognized in this early
15 patent, the condensation reaction forms one mole of water co-product for each mole of reactants condensed. The presence of this water co-product during subsequent reaction of the phenols with the
20 ketones, acts to lower the reflux temperatures of the reaction mixture. Consequently, in the U.S. Patent 2,858,342 it is suggested that the water of condensation be removed during the reaction by distillation into a Dean-Stark trap.

25 Later processes employ as the condensation catalyst acidic ion-exchange resins; see for example the descriptions found in U.S. Patents 3,634,341 and 3,760,006. These descriptions point out that water content of the resin catalyst is an important
30 process variable. Catalyst efficiency is impaired if the water content of the reaction mixture is more than about 2-3 weight percent, especially in a fixed bed reactor. In the U.S. Patent 3,634,341 it is said that the resin catalyst should be dried, if

necessary, by passing anhydrous phenol through the bed of resin.

A catalyst utilized in the commercial processes is a macro-reticular or gelular, acidic ion-exchange resin; see for example the U.S. Patent 4,191,843 (Kwantes et al). These catalysts are the product of monoethylenically unsaturated monomers copolymerized with polyvinylidene monomers, as described for example in the U.S. Patent 3,037,052 (Bortnick). Preferred macro-reticular resins are sulfonated to position sulfonic acid groups, pendant from the aromatic nucleus of the vinylidene moiety; see the U.S. Patent 3,037,052.

The above-described acid catalyst (sulfocationites) contains up to 40 to 85 percent by weight of water and are generally made available by the manufacturer in this hydrated state. The presence of the water will inhibit the desired catalyst activity in condensing phenols with ketones, because the water will bind to the sulfonic acid groups in competition with reactant. For this reason, Bortnick (U.S. Patent 3,037,052) teaches one to dehydrate the catalyst prior to use. As a method of dehydration, Bortnick suggests "drying at 105°C. to about 125°C. at a pressure of 5 to 10 mm" or by azeotropic distillation with an organic liquid (column 4, lines 54-64).

Kwantes et al., in U.S. Patent 4,191,843 describing in particular the preparation of bisphenols, states that "the reactor may be filled with the acid ion exchanger by any known technique. Such techniques include adding the desired amount of dry ion exchanger, water wet ion exchanger or slurry of the ion exchanger in the reactor"; see column 2, lines 51-55. Details of the drying procedure are

not elaborated upon beyond stating that the reactor charged resin is "drained"; see the patent Example 1.

5 In the commercially practiced processes for preparing bisphenol-A, acid ion-exchange resins are dried in the plant reactor (usually a column reactor) prior to initiation of the reaction between phenol and acetone. Drying is effected by passing phenol over the catalyst in the absence of acetone,
10 to absorb and/or displace the water associated with the acid ion-exchange resin. Water that is absorbed or displaced by the phenol is removed from the resin bed as an effluent stream. The water is ultimately removed from the phenol effluent by employing
15 conventional separation techniques such as distillation and extraction.

The aforementioned dehydration procedure is time consuming, often requiring several weeks and results in substantial reactor down time. Where a
20 plant utilizes large amounts of catalyst resin, resin addition and drying can take several months. Furthermore, cycles of resin addition and dehydration are known to stress and damage macro-reticular or gelular resin beads. This
25 deleteriously affects reaction mixture flow in the reactor and pressure drop across the bed.

One solution considered for the problems associated with the commercially practiced procedure of drying the acidic ion-exchange resin is to follow
30 the teachings of the aforementioned U.S. Patent 3,037,052 and pre-dry the resin to an anhydrous state before it is charged to the column reactor. However, this procedure also has certain disadvantages. For example, if the resin is not
35 rapidly dried under elevated temperature conditions

(circa 150°C.), Zundel et al. has reported (Physik. Chem. (Frankfurt) 59, 225 [1968]) that the water is hydrogen bonded to three sulfonic acid groups and cannot be removed even under severe drying

5 conditions. However, if drying is carried out at too rapid a rate, the resin beads may be subjected to abrasion and damage with a resultant loss of operating life. Clearly, the rate of drying is critical to an economic process.

10 The present invention provides apparatus and a method of drying and removing water of condensation during the course of the condensation reaction for maximum efficiency and long (extended) catalyst life.

15 **SUMMARY OF THE INVENTION**

The invention comprises a process for the condensation of a monocyclic phenol with a ketone, which comprises;

20 reacting a phenol with a ketone in the presence of a catalytic proportion of an acidic ion-exchange resin, in a reaction zone, whereby a mole of water of condensation is obtained as a co-product of the condensation for each mole of reacted ketone; and

25 continuously removing the water of condensation from the reaction zone by entrainment in a drying gas passed through the reaction zone.

Brief Description of the Drawing

30 The accompanying drawing is a side view of apparatus, in cross-section, for carrying out the method and process of the invention.

DETAILED DESCRIPTION OF THE PREFERRED EMBODIMENTS OF THE INVENTION

35 Processes for the synthesis of bisphenols by the condensation of a ketone with a stoichiometric excess of a phenol compound, are generally well

known; see for example the descriptions given in the U.S. Patents 3,634,341; 3,760,006; 4,045,379; 4,107,218; 4,294,995; 4,346,247; 4,396,728; 4,400,555; 4,424,283; 4,584,416; and 5,075,511, all of which are incorporated herein by reference thereto. The processes may be carried out continuously or in batch procedures.

In general, the processes may be carried out by bringing the ketone and a stoichiometric excess of the phenol together in a reaction zone, at a temperature of from about 40°C. to 95°C. in the presence of a cationic exchange resin under ambient atmospheric conditions. In some processes, higher than atmospheric pressures may be employed, generally within the range of 1-10 atmospheres. The bisphenol product can then be separated from the reaction mixture by conventional procedures well known in the art, including by distillation, crystallization and like methods.

In the commercial process the cationic exchange resin catalyst generally employed is a sulfonated aromatic resin, optionally modified to have mercaptan groups covalently or ionically bound to the resin as promoter sites. Representative of such cationic exchange resins are those described in the U.S. Patents 3,760,006; 3,634,341; 4,045,379; 4,294,995; 4,346,247; and 4,820,740, all of which are incorporated herein by reference thereto.

In general, the resins employed are chemically modified sulfonated aromatic resins. The resins are hydrocarbon polymers having a plurality of pendant sulfonic acid groups. Examples include sulfonated polystyrene or poly(styrene-divinyl-benzene) copolymer and sulfonated phenol-formaldehyde resins. The sulfonated resins are commercially available in

water swollen form as micro-recticular and macro-recticular types. Specific examples of suitable resins are Amberlite IR-120H, Amberlyst 15H, Amberlyst 31, Dowex 50-X-4, Dowex MSC-1H, Duolite C-26, Permutit QH, Chempro C-2 and Imac C8P/H (Amberlite, Amberlyst, Dowex, Duolite, Permutit, Chempro and Imac are registered U.S. Trademarks). Further examples of such ion exchangers as well as methods for preparing such ion exchangers are described in the Encyclopedia of Polymer Science and Technology, 1967, vol. 7, pages 695 to 708. The exchange capacity of the acidic resin is preferably at least 2.0 meq. H^+ /g of dry resin, with exchange capacities in the range of from 3.0 to 5.5 meq. H^+ /g (dry resin) particularly preferred. Sulfonation may be by the process described in U.S. Patent 2,366,007 which is incorporated herein by reference thereto.

One preferred catalyst used in the process of the present invention is the Amberlyst[®] gelular types, which are styrene cross-linked with divinylbenzene or like cross-linking monomer and having pendant sulfonic acid groups attached to the aromatic nucleus of the styrene moiety. Most preferred are modified acid ion-exchange resins, wherein the modification comprises esterifying the resin with a mercapto alcohol or by partially neutralizing the resin with an alkyl mercaptoamine. As an alternative to such modification, the bisphenol preparation with the ion-exchange resin may be carried out in the presence of sulfur compound promoters such as alkyl mercaptans, especially 3-mercaptopropionic acid; see for example U.S. Patent 4,191,843 incorporated herein by reference thereto.

The catalyst is employed in a proportion effective for catalytic condensation of the phenol with the ketone. In general, a catalytic proportion is achieved by passage of the mixed reactants
5 through a bed of the resin under conditions for condensation.

The process of the invention is advantageously carried out in the synthesis of a wide variety of bisphenols such as 1-phenyl-1,1-bis(4-hydroxy-
10 phenyl) ethane; 1-methyl-1,1-bis(4-hydroarylphenyl) ethane and the like. It is a preferred improvement in the synthesis of 2,2-bis-(4-hydroxyphenyl) propane (bisphenol-A).

The commercial preparation of bisphenol-A may
15 be carried out in a number of ways. The reaction zone may comprise a single reactor or two or more reactors in series or in parallel. In the case of a multi-reactor reaction zone, suitably all of the phenol reactant is fed to the first reactor and the
20 acetone reactant is either fed all to the first reactor or divided between the first and second and possibly further reactors.

The molar ratio of phenol to acetone is at least 2 to 1 with a stoichiometric excess of phenol
25 being preferred. Molar ratios are preferably from 3:1 to 40:1, with molar ratios of from 10:1 to 30:1 being most preferred. The optimum ratio depends on reaction conditions, e.g., temperature of reaction and desired conversion.

30 The reaction temperature in the reactor zone may vary from 40°C. to 95°C. with reaction temperatures in the range of from 55°C. to 90°C. being preferred.

The reaction time in the reactor zone may also
35 vary and depends on reaction temperature and the

type of reactor. For example, in a continuous feed, tubular reactor, the liquid hour space velocity (LHSV) of the feed may vary between wide limits with velocities in the range of from 0.2 to
5 40 liters feedstream/liter catalyst⁻¹/hour⁻¹.

The partially dehydrated catalyst may be charged to an appropriate reactor, for example a fixed bed column reactor and saturated with anhydrous phenol to complete dehydration.
10 Alternatively, the wet ion exchange resin may be charged to the reaction zone and subjected to partial drying as described above in-situ, before saturation with phenol. The phenol addition may be carried out in the reactor or in a different vessel
15 and then transferred to the reactor or reaction zone. The catalyst is further dried by the addition of phenol to the desired level of water content in the catalyst. This level of water will affect the rate of reaction between the phenol and acetone.
20 Essentially about 4 weight percent water in the phenol effluent would bring about a marginally acceptable reaction rate while a level of about 2 weight percent in the effluent or less is preferred.

A level of less than about 1 weight percent is most preferred.
25 The reaction zone is then ready to receive acetone and a molar excess of phenol for condensation following known procedures.

The improved process of the invention will now be described in detail with reference to the
30 accompanying drawing of embodiment apparatus for carrying out the process. As shown in cross-section, a column reactor vessel 12 supports an ion-exchange resin bed 14 on a layer 16 of sand and a layer 18 of glass beads. The vessel 12 is jacketed
35 and warmed by a fluid medium circulated through the

jacket between inlet 20 and outlet 22. Mixed reactants enter the vessel 12 through feed inlet 24 and pass by gravity through the catalyst bed 14, which is a reaction zone for condensing the reactants. The product bisphenol drains from the resin bed 14 and is carried from the vessel 12 through conduit 26 for eventual recovery by conventional means. An inert gas, such as nitrogen is inserted into the vessel 12 through inlet 28 and distributed to the resin bed 14 through sparger 30. Bubbles of the drying gas, released from sparger 30 rises and passes through the resin bed 14, counter-current to the reactants and product, continuously entraining and desorbing the co-product water of condensation from the resin bed 14. The gas, with entrained water of condensation is vented from vessel 12 through vent 34 and recycled for re-use by separating the entrained water. The result is the maintenance of a near anhydrous reaction zone environment.

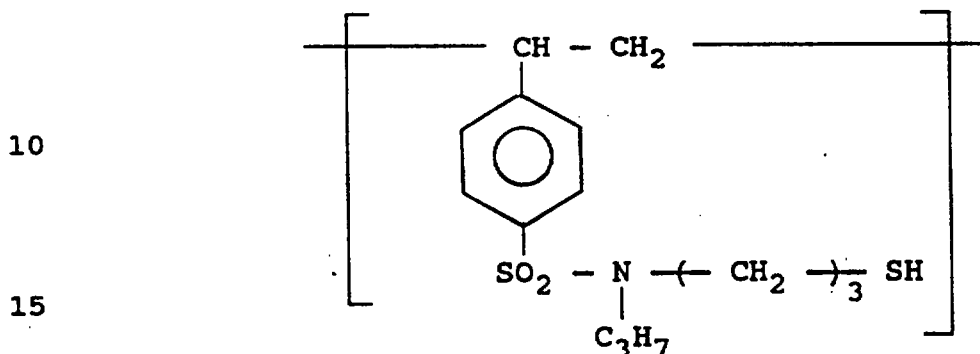
A further benefit results from the counter-current passage of the drying gas through the resin bed, which seems to "agitate" the reaction mixture, improving acetone conversion rates.

The following example describes the manner and process of making and using the invention and sets forth the best mode contemplated by the inventors for carrying out the invention, but should not be construed as limiting.

EXAMPLE 1

A tubular reactor similar to that shown in the accompanying drawing with a feed inlet and a product outlet was maintained at a temperature of about 65° C. The reactor was charged with a sulfonated polystyrene polymer catalyst prepared in accordance

with the procedure described in Example 1 of the U.S. Patent 4,396,728 and having approximately 80% pendant sulfonyl styryl groups and 20% of chemically combined N-propylaminopropyl-mercaptan groups of formula:-



catalyst and fed a reaction mixture of phenol and acetone (6% by weight acetone in pure phenol). At a weighted hour space velocity (WHSV) of 0.5 grams of feed per gram of dry catalyst per hour. During the reaction period, dry nitrogen gas was passed through the catalyst resin bed at a rate of about 2 cc/min./gm. dry catalyst to 20 cc/min./gm. dry catalyst. The gas was vented from the reaction vessel, with entrained water of the condensation reaction in admixture with some phenol and some acetone. The product bisphenol-A was collected over a period of time, with periodic analysis to determine the percentage of conversion and selectivity effected by the catalyst. The analytical results and time of collection is set forth in Table 1, below.

TABLE 1

Example 1	% H ₂ O	Acetone	Phenol	IPP	PP BPA	OP BPA	Dimers	BPX-1	CR-1	Spiro	BPXR	Unk
Start	0.22	4.3	82.57	0.00	8.27	2.38	2.36	0.70	0.96	0.04	0.45	2.27
Feed Tank			76.69	0.00	11.69	3.00	3.02	0.89	1.17	0.26	0.58	2.71
WHSV=0.5												
2 Hours	1.3	0.79	85.59	0.00	6.58	1.67	1.65	0.50	0.66	0.15	0.33	2.59
.3 Hours	1.2	1.33										
4 Hours	0.94	0.64	77.04	0.00	13.23	2.49	2.36	0.68	0.95	0.19	0.46	2.61
5 Hours	0.59	0.44	73.42	0.00	16.91	2.76	2.56	0.73	1.01	0.21	0.31	2.09
6 Hours	0.45	0.48	71.00	0.00	18.64	2.79	2.55	0.71	1.02	0.22	0.50	2.57
7 Hours	0.4	0.44	68.00	0.00	21.11	2.95	2.71	0.77	1.08	0.24	0.52	2.63
Control *	1.37	0.17	68.24	0.00	22.84	2.35	2.31	0.74	0.91	0.04	0.45	2.16

FOOTNOTE: The control run was carried out at a
WHSV = 0.25, without nitrogen gas purging.

The advantages of the improved method of the invention include the following. The reduction of the co-product water as it is produced avoids retardation of catalyst activity. there is a consequent improved conversion rate, and catalyst life. the improved conversion rates seem to accompany cooler reaction temperatures, which enhances selectivity. There appears to be improved color in the bisphenol product. More importantly 79-80 percent conversion of the acetone occurred at a throughput equal to 4 times that used in standard preparations.

The improved process of the invention has a potential promise of commissioning of new catalyst could be done faster and with much less energy

EXAMPLE 2 .

The general procedure of Example 1, supra., is repeated except that the reaction temperature is 80°C. The test results are shown in Table 2, below.

TABLE 2

Example 2	% H ₂ O	Acetone	Phenol	IPP	PP BPA	OP BPA	Dimers	BPX-1	CR-1	Spiko	BPXR	Unk
WHSV=1.0												
4 Hours	0.361	0.231										
5 Hours	0.192	0.189										
6 Hours	0.183	0.188										
7 Hours	0.236	0.139	73.02	0.00	16.92	2.15	2.39	0.56	0.91	0.08	0.42	2.53
8 Hours	0.223	0.146	70.39	0.00	18.54	2.34	3.75	0.58	1.01	0.09	0.47	2.54
9 Hours		0.197	67.54	0.00	20.24	2.52	4.13	0.86	1.09	0.09	0.52	3.22
Control**	1.37	0.17	68.24	0.00	22.84	2.35	2.31	0.71	0.91	0.04	0.45	2.16

FOOTNOTE: .. The control run was carried out at a
WHSV = 0.25, without nitrogen gas purging.

**WHAT IS CLAIMED IS:**

1. A process for the condensation of a monocyclic phenol with a ketone, which comprises;
reacting a phenol with a ketone in the presence of a catalytic proportion of an acidic ion-exchange resin, in a reaction zone, whereby a mole of water of condensation is obtained as a co-product of the condensation for each mole of reacted ketone; and continuously removing the water of condensation from the reaction zone by entrainment in a drying gas passed through the reaction zone.
2. A process of claim 1 wherein the phenol is phenol and the ketone is acetone.
3. A process of claim 1 wherein the drying gas is nitrogen.
4. A process of claim 1 wherein the reaction zone comprises a bed of the resin contained in a reactor.
5. A process for the preparation of bisphenol-A, which comprises;
providing a hydrated acidic ion-exchange resin catalyst for catalyzing the condensation of phenol with acetone;
vaporizing a portion of the water associated with the hydrated catalyst;
separating sufficient of the water vapor from the catalyst to remove from about 25 to 80 percent by weight of said water, whereby a partially dehydrated catalyst is obtained;
contacting the partially dehydrated catalyst with phenol so as to remove additional water from the catalyst until the water content of the effluent phenol contains less than about 4 weight percent water;





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condensing phenol and acetone in the presence of the dehydrated catalyst, whereby bisphenol-A is formed with water of the condensation; and
20 passing a drying gas through the dehydrated catalyst to entrain and remove the water of condensation as it is produced.

6. The process of claim 5 wherein the drying gas is nitrogen.

7. The process of claim 6 wherein the nitrogen gas is passed counter-current to the flow of reactants at a rate of from 2 cc to 20 cc/min./gm. dry catalyst.

8. The process of claim 7 carried out at a temperature of from 40°C. to 95°C.



DEHYDRATION OF ION-EXCHANGE RESIN CATALYZED
REACTIONS

ABSTRACT OF THE DISCLOSURE

The preparation of bisphenols by condensation of a phenol compound with a ketone is improved by employing an ion-exchange resin as the acid catalyst, and removing the co-product water of condensation as it forms in the catalyst resin bed by counter-current passage of a drying gas to remove the water. The improvement results in better ketone conversion rates and catalyst life.

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10 The products have improved color.

Figure 1

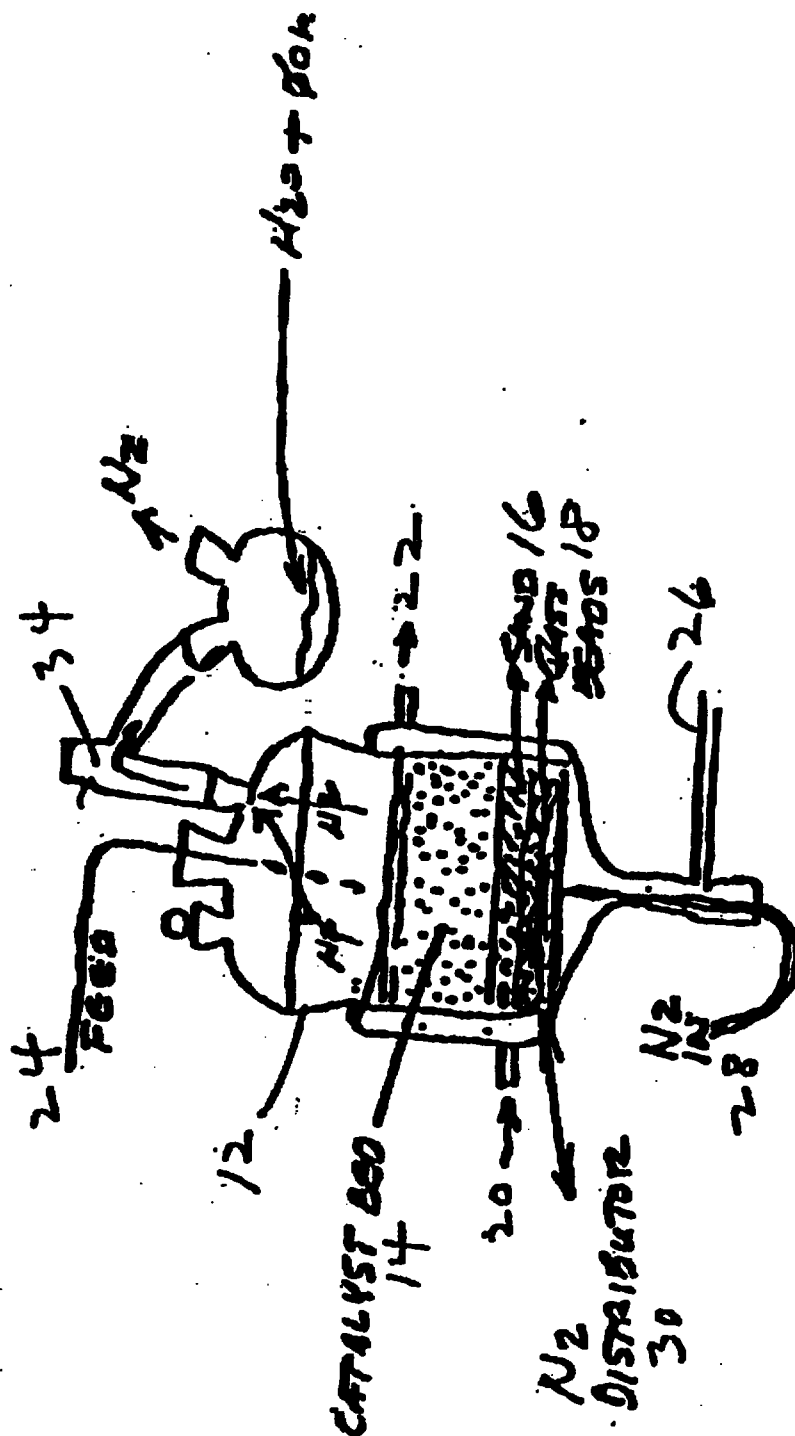


Figure 1





REPLACEMENT SHEET

12 DEC 1997

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WHAT IS CLAIMED IS :

1. A process for the condensation of a monocyclic phenol with a ketone comprising reacting a phenol with a ketone in the presence of a catalytic proportion of an acidic ion-exchange resin, in a resin bed supported by a layer of sand and a layer of glass beads, to form a reaction mixture in a reaction zone, whereby a mole of water of condensation is obtained as a co-product of the condensation for each mole of reacted ketone, and
 wherein the improvement comprises continuously removing the water of condensation from the reaction zone by entrainment in a drying gas passed through the reaction zone countercurrent to the phenol and the ketone, venting the gas, separating the entrained water from the drying gas and recycling the drying gas for re-use whereby the drying gas agitates the reaction mixture, maintains a near anhydrous reaction zone environment and improves acetone conversion rates.
2. A process of claim 1 wherein the phenol is phenol and the ketone is acetone.
3. A process of claim 1 wherein the drying gas is nitrogen.
4. A process of claim 1 wherein the reaction zone comprises a bed of the resin contained in a reactor.
5. An improved process for the preparation of bisphenol-A comprising
 providing a hydrated acidic ion-exchange resin catalyst, in a resin bed supported by a layer of sand and a layer of glass beads, for catalyzing the condensation of phenol with acetone;
 vaporizing a portion of the water associated with the hydrated catalyst;

This Sheet Has Been Replaced





12 DEC 1997

separating sufficient of the water vapor from the catalyst to remove from about 25 to 80 percent by weight of said water, whereby a partially dehydrated catalyst is obtained;

contacting the partially dehydrated catalyst with phenol so as to remove additional water from the catalyst until the water content of the effluent phenol contains less than about 4 weight percent water; and

condensing phenol and acetone in a reaction mixture in the presence of the dehydrated catalyst, whereby bisphenol-A is formed with water of the condensation; and

wherein the improvement comprises passing a drying gas through the dehydrated catalyst countercurrent to the phenol and acetone to entrain and remove the water of condensation as it is produced, venting the gas with the entrained water through a vent, separating the entrained water from the gas and recycling the gas for re-use whereby the drying gas agitates the reaction mixture, maintains a near anhydrous reaction zone environment and improves acetone conversion rate.

6. The process of claim 5 wherein the drying gas is nitrogen.
7. The process of claim 6 wherein the nitrogen gas is passed counter-current to the flow of reactants at a rate of from 2 cc to 20 cc/min./gm. dry catalyst.
8. The process of claim 7 carried out at a temperature of from 40°C. to 95°C.



REPLACEMENT SHEET

25 APR 1998

WHAT IS CLAIMED IS :

1. A process for the condensation of a monocyclic phenol with a ketone comprising reacting a phenol with a ketone in the presence of a catalytic proportion of an acidic ion-exchange resin, in a single resin bed supported by a layer of sand and a layer of glass beads, to form a reaction mixture in a reaction zone, whereby a mole of water of condensation is obtained as a co-product of the condensation for each mole of reacted ketone, and
wherein the improvement comprises continuously removing the water of condensation from the reaction zone by entrainment in a drying gas passed through the reaction zone countercurrent to the phenol and the ketone, venting the gas, separating the entrained water from the drying gas and recycling the drying gas for re-use whereby the drying gas agitates the reaction mixture, maintains a near anhydrous reaction zone environment and improves acetone conversion rates.
2. A process of claim 1 wherein the phenol is phenol and the ketone is acetone.
3. A process of claim 1 wherein the drying gas is nitrogen.
4. A process of claim 1 wherein the reaction zone comprises a bed of the resin contained in a reactor.
5. An improved process for the preparation of bisphenol-A comprising
providing a hydrated acidic ion-exchange resin catalyst, in a single resin bed supported by a layer of sand and a layer of glass beads, for catalyzing the condensation of phenol with acetone;
vaporizing a portion of the water associated with the hydrated catalyst;

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FIG. 1

