

Description

This invention relates to glass substrates for flat panel display devices.

Two recent technological advances have made large, flat-panel display devices a distinct practical possibility: (1) the preparation of liquid crystals exhibiting improved characteristics; and (2) the production of a surface layer of fine-grained polysilicon.

The development of liquid crystals has progressed to the point where miniature color television receiver sets have been produced from them and much larger information display panels are technically possible. Liquid crystals inherently demonstrate slow responses to electrical signals and, therefore, a "switch" to rapidly respond to an electrical stimulus is required to set up an active matrix display. The thin film transistor (TFT) performs that function.

It is well-known that a TFT can be fabricated from single crystal silicon; however, by the very nature of single crystal silicon, there is a limitation on the size of a TFT that can be fashioned therefrom. It has been discovered that a fine-grained polysilicon layer which has been deposited onto a substrate can be recrystallized to large-grained polysilicon by scanning with a heat source such as a laser. It has been observed that a TFT prepared from large-grained polysilicon evidences only slight deterioration in electrical characteristics when compared with those displayed by a TFT fabricated from single crystal silicon, and operates very satisfactorily in multiplexing a liquid crystal display.

A large-grained polysilicon active matrix display requires a substrate which is transparent, flat, smooth, inert, compatible with silicon, in terms of thermal expansion, and capable of withstanding processing temperatures of at least 850°C. Accordingly, a glass will meet those requirements if it is free from alkali metal ions, it displays a linear coefficient of thermal expansion (25°-300°C) of $30\text{-}40 \times 10^{-7}/^\circ\text{C}$, and an annealing point greater than 875°C. In order to utilize the glass sheet forming method described in United States Patent No. 3,338,696, the glass must exhibit a liquidus viscosity of at least 10kPa.s (100,000 poises), preferably over 25kPa.s (250,000 poises). Also, transparent, homogeneous glasses with annealing points in the vicinity of 900°C are very difficult to obtain by melting batches at temperatures no higher than 1800°C. 1800°C represents the practical limit for melting materials in contact with platinum-rhodium and most refractories used in an oxidizing atmosphere. Finally, to permit glass forming to be carried out utilizing conventional equipment and techniques, the liquidus temperature should be no more than 1400°C.

Therefore, the primary objective of the present invention is to prepare transparent, homogeneous glasses demonstrating a linear coefficient of thermal expansion (25°-300°C) of $30\text{-}40 \times 10^{-7}/^\circ\text{C}$, an annealing point greater than 875°C, a liquidus temperature no higher than 1400°C, a liquidus viscosity of at least 10kPa.s (100,000 poises), which can be melted at temperatures not exceeding 1800°C.

The present invention provides a clear, transparent, homogeneous strontium aluminosilicate glass capable of being melted at temperatures no higher than 1800°C which exhibits an annealing point greater than 850°C, a linear coefficient of thermal expansion over the temperature range of 25°-300°C of $30\text{-}40 \times 10^{-7}/^\circ\text{C}$, a liquidus temperature not exceeding 1400°C, and a liquidus viscosity of at least 10kPa.s (10^5 poises), said glass being essentially free from alkali metal oxides, MgO, CaO, ZnO, La_2O_3 , MnO, and consisting essentially, expressed in terms of mole percent on the oxide basis, of

9-12% SrO
9-12% Al_2O_3
77-82% SiO_2

The limits of the component regimes are extremely critical. To illustrate, where the level of SrO is too high, the annealing point falls below the desired target and the linear coefficient of thermal expansion may be raised to an unacceptably high value. Conversely when the SrO content is too low, the glass becomes very difficult to melt and/or the liquidus temperature is raised to an unacceptable level. Al_2O_3 concentrations lower than those specified lead to annealing points that are too low. Al_2O_3 contents greater than those specified result in undesirably high liquidus temperatures. When the amount of SiO_2 exceeds the prescribed maximum, the glasses become too difficult to melt; at SiO_2 levels below the stated minimum, either the annealing point becomes too low or the liquidus temperature becomes too high.

Whereas a precise conversion of composition ranges expressed in mole percent to ranges in weight percent is not possible, an approximation of the ternary system expressed in weight percent comprises 13-18% SrO, 13-18% Al_2O_3 , and 66-72% SiO_2 .

Minor substitutions of MgO, CaO, ZnO, La_2O_3 , and MnO for SrO in the ternary system undesirably affect the linear coefficient of expansion and/or the annealing point and/or the liquidus temperature of the glass. Therefore, only very small additions of those materials can be tolerated when the desired glass properties are to be maintained. Strong fluxes such as the alkali metal oxides will desirably be essentially absent.

The substitution of BaO for part of the SrO causes a slight rise in the linear coefficient of expansion and has the very desirable capability of reducing the liquidus temperature without adversely affecting the viscosity of the glass to any substantial extent. Hence, the substitution of BaO for SrO does not alter the annealing point appreciably. That capability of BaO is of practical significance in that, by balancing the SrO and BaO concentrations, the coefficient of thermal expansion can be varied and carefully controlled. Whereas substitutions of BaO for SrO up to 8 mole percent, preferably no more than 6 mole percent, can be useful, volatilization of BaO is a recognized problem when melting temperatures of 1800°C are employed. SrO is much less subject to volatilization at 1800°C.

Furthermore, and even more importantly from the standpoint of glass melting, the inclusion of large amounts of BaO leads to the development of seeds in the glass. For example, where BaO is substituted in total for SrO, it is virtually impossible to produce at melting temperatures no higher than 1800°C a glass containing no more than a few seeds. In contrast, laboratory experience has demonstrated that essentially seed-free glasses can be prepared from glasses within the above-prescribed ranges of the ternary system SrO-Al₂O₃-SiO₂. Accordingly, whereas the incorporation of BaO can impart useful properties to the glass, its presence can lead to environment polluting volatilization and renders the glass more difficult to melt to good quality.

Prior Art

Kh. Sh. Iskhakov, "Region of Glass Formation in a Strontium Oxide-Aluminum Oxide-Silicon Dioxide System", *Uzb. Khim. Zh.* 15 [1], 10-12 (1971) describes the preparation of glasses composed, in mole percent, of 25-60% SrO, 5-30% Al₂O₃, and 35-65% SiO₂ by melting at 1500°-1550°C for 1-1.5 hours. The SrO contents are quite apparently far removed from those of the present invention.

Kh. Sh. Iskhakov, "Properties of Glasses in the Strontia-Alumina-Silica System", *Uzb. Khim. Zh.*, 15 [2], 79-81 (1971) discusses several physical properties determined on glasses having compositions within the ranges of the above Iskhakov literature reference. It was observed that the glasses manifested coefficients of thermal expansion of $64-97 \times 10^{-7}/^{\circ}\text{C}$. It was stated that as the SrO content is raised, the coefficient of thermal expansion rises.

G. I. Zhuravlev, A. I. Kuznetsov, T. I. Semenova, and N. G. Suikovskaya, *Glass*, USSR SU870,365, January 7, 1984, discloses the preparation of glasses demonstrating higher softening points and special electrical resistivities containing, in weight percent, 25-35% SrO, 11-20% Al₂O₃, and 41-63% SiO₂. The SrO and SiO₂ levels are generally higher and lower, respectively, when compared with their concentrations in the present inventive glasses.

G. I. Zhuravlev, V. A. Malov, A. I. Kuznetsov, and T. I. Draunina, "Application of Dispersed Glass of Silicon Dioxide-Aluminum Oxide-Strontium Oxide System on a Niobium Substrate by Electrophoretic Deposition", *Zh. Prikl. Khim. (Leningrad)* 54 [7] 1601-4 (1981) describes the application of a protective electric insulating coating on niobium parts of high pressure lamps through electrophoretic deposition of glass having a composition within the ranges disclosed in the above Zhuravlev et al. literature reference.

United States Patent No. 4,180,618 reports the fabrication of electronic devices comprised of a thin film of silicon deposited upon a substrate, wherein the substrate is composed of a glass consisting essentially, in weight percent, of 55-75% SiO₂, 5-25% Al₂O₃, and at least one alkaline earth oxide selected from the group in the indicated proportions of 9-15% CaO, 14-20% SrO, and 18-26% BaO. The CaO and BaO contents are higher than can be tolerated in the present inventive glasses. Whereas the broad ranges of SrO, Al₂O₃, and SiO₂ overlap those of the inventive glasses, there is no recognition of the matrix of superior properties which can be secured from compositions coming within the narrowly circumscribed ranges of the present invention; nor is there a working example provided which falls within the limits of the present inventive glasses.

Table 1 records a number of glass compositions, expressed in terms of mole percent on the oxide basis, illustrating the parameters of the instant invention.

Thus Examples 1,6,16,19,32,36,37,40,41,47 and 48 are compositions of the invention whereas Examples 2-5,7-15,17,18, 20-31, 33-35, 38, 39 and 42-46 are comparative examples.

The actual batch ingredients can comprise any materials, either the oxides or other compounds, which, when melted together, will be converted into the desired oxides in the proper proportions. In the laboratory work described below, the batch materials consisted of high purity sand, Al₂O₃, SrCO₃, BaCO₃, MgO, CaO, ZnO, MnO, and La₂O₃.

The batch materials were compounded, ballmilled to assist in achieving a homogeneous melt, and charged into platinum or platinum-rhodium crucibles. Those batches containing 77% SiO₂ and less could be melted at 1600-1650°C for 4-16 hours in an electrically-heated furnace. The batches of higher SiO₂

contents were melted at 1700°-1800°C for 16 hours in a gas-oxygen fired furnace. In each case, the melts were cast into an iron mold to form circularly-shaped slabs having a diameter of about 30.5cm (12") and a thickness of about 1.9cm (0.75"), and those slabs immediately transferred to an annealer.

Table IA reports the compositions of Table I in terms of weight percent on the oxide basis.

TABLE I

	<u>1</u>	<u>2</u>	<u>3</u>	<u>4</u>	<u>5</u>	<u>6</u>
SiO ₂	80	80	80	80	80	80
Al ₂ O ₃	10	10	10	10	10	10
SrO	10	8	6	8	6	4
MgO	-	2	4	-	-	-
CaO	-	-	-	2	4	-
BaO	-	-	-	-	-	6
	<u>7</u>	<u>8</u>	<u>9</u>	<u>10</u>	<u>11</u>	<u>12</u>
SiO ₂	80	80	80	80	80	80
Al ₂ O ₃	10	10	10	10	10	10
SrO	8	6	8	6	8	6
ZnO	2	4	-	-	-	-
LaO _{1.5}	-	-	2	4	-	-
MnO	-	-	-	-	2	4
	<u>13</u>			<u>14</u>	<u>15</u>	<u>16</u>
SiO ₂	80			80	80	79
Al ₂ O ₃	10			10	10	10
SrO	8			4	-	11
BaO+CaO	2			6	10	-
	<u>17</u>	<u>18</u>	<u>19</u>	<u>20</u>	<u>21</u>	<u>22</u>
SiO ₂	77	77	77	75	75	75
Al ₂ O ₃	8	10	12	8	10	12
SrO	15	13	11	17	15	13

Table I (Cont.)

5		<u>23</u>	<u>24</u>	<u>25</u>	<u>26</u>	<u>27</u>	<u>28</u>
	SiO ₂	73	73	73	70	71	80
	Al ₂ O ₃	8	10	12	15	12	7
10	SrO	19	17	15	15	17	13
		<u>29</u>	<u>30</u>	<u>31</u>	<u>32</u>		
15	SiO ₂	78	76	80	79		
	Al ₂ O ₃	9	11	8	9		
	SrO	13	13	12	12		
20		<u>33</u>	<u>34</u>			<u>35</u>	<u>36</u>
	SiO ₂	82	81			82	81
	Al ₂ O ₃	7	8			8	9
25	SrO	11	11			10	10
			<u>37</u>	<u>38</u>	<u>39</u>		
30	SiO ₂		78	77	83		
	Al ₂ O ₃		12	13	8		
	SrO		10	10	9		
35		<u>40</u>	<u>41</u>	<u>42</u>	<u>43</u>	<u>44</u>	<u>45</u>
	SiO ₂	80	79	76	84	83	82
40	Al ₂ O ₃	11	12	13	8	9	10
	SrO	9	9	9	8	6	8
		<u>46</u>	<u>47</u>	<u>48</u>			
45	SiO ₂	83	80	80			
	Al ₂ O ₃	10	10	10			
	SrO	7	3	2			
50	BaO	-	7	8			

TABLE IA

5		<u>1</u>	<u>2</u>	<u>3</u>	<u>4</u>	<u>5</u>	<u>6</u>
	SiO ₂	70	71.3	72.7	71.0	72.0	67.1
	Al ₂ O ₃	14.9	15.2	15.4	15.1	15.3	14.3
10	SrO	15.1	12.3	9.4	12.3	9.3	5.8
	MgO	-	1.2	2.5	-	-	-
	CaO	-	-	-	1.6	3.4	-
15	BaO	-	-	-	-	-	12.8
		<u>7</u>	<u>8</u>	<u>9</u>	<u>10</u>	<u>11</u>	<u>12</u>
	SiO ₂	70.5	70.9	70.5	71.1	70.7	71.4
20	Al ₂ O ₃	15.0	15.1	15.0	15.1	15.0	15.2
	SrO	12.2	9.2	12.2	9.2	12.2	9.2
	ZnO	2.3	4.8	-	-	-	-
25	LaO _{1.5}	-	-	2.3	4.6	-	-
	MnO	-	-	-	-	2.1	4.2
		<u>13</u>			<u>14</u>	<u>15</u>	<u>16</u>
30	SiO ₂	69.0			69.9	65.3	68.7
	Al ₂ O ₃	14.7			14.9	13.9	14.8
35	SrO	11.9			6.0	-	16.5
	BaO+CaO	4.4			9.2	20.8	-
		<u>17</u>	<u>18</u>	<u>19</u>	<u>20</u>	<u>21</u>	<u>22</u>
40	SiO ₂	66.1	66.1	66.3	63.6	63.6	63.8
	Al ₂ O ₃	11.7	14.6	17.3	11.5	14.4	17.1
	SrO	22.2	19.3	16.4	24.9	22.0	19.1
45		<u>23</u>	<u>24</u>	<u>25</u>	<u>26</u>	<u>27</u>	<u>28</u>
	SiO ₂	61.1	61.2	61.4	57.8	58.9	70.0
50	Al ₂ O ₃	11.4	14.2	16.8	20.8	16.7	10.4
	SrO	27.5	24.6	21.8	21.4	24.4	19.6

Table IA (Cont.)

5		<u>29</u>	<u>30</u>	<u>31</u>	<u>32</u>		
	SiO ₂	67.4	64.9	70.1	66.8		
	Al ₂ O ₃	13.2	15.9	11.9	13.3		
10	SrO	19.4	19.2	18.0	17.9		
		<u>33</u>	<u>34</u>		<u>35</u>	<u>36</u>	
15	SiO ₂	72.6	71.3		72.7	71.3	
	Al ₂ O ₃	10.5	12.0		12.0	13.5	
	SrO	16.9	16.7		15.3	15.2	
20			<u>37</u>	<u>38</u>	<u>39</u>		
	SiO ₂		67.6	66.3	74.0		
	Al ₂ O ₃		17.4	18.8	12.1		
25	SrO		15.0	14.9	13.9		
		<u>40</u>	<u>41</u>	<u>42</u>	<u>43</u>	<u>44</u>	<u>45</u>
30	SiO ₂	70.0	68.9	67.6	75.4	74.0	72.7
	Al ₂ O ₃	16.4	17.5	18.9	12.2	13.7	15.1
	SrO	13.6	13.6	13.5	12.4	12.3	12.2
35		<u>46</u>	<u>47</u>	<u>48</u>			
	SiO ₂	74.0	66.7	66.2			
	Al ₂ O ₃	15.2	14.1	14.0			
40	SrO	10.8	4.3	2.9			
	BaO	-	14.9	16.9			

45 Table II records a number of physical properties measured on the above glasses. The annealing point (Ann. Pt.) and strain point (Str. Pt.) in terms of °C were determined in accordance with the beam bending method described in ASTM C598; the linear coefficient of thermal expansion (Coef. Exp.) in terms of $\times 10^{-7}/^{\circ}\text{C}$ over the range of 25°-300°C was ascertained in accordance with ASTM E228 utilizing a fused silica dilatometer; the internal liquidus temperature (Liquidus) in terms of °C was determined by placing the
50 glass in a platinum boat, introducing the boat into a furnace having a temperature gradient spanning the liquidus temperature, and maintaining the boat in that furnace for 24 hours; the viscosity of the glass at the liquidus (Liq. Vis.) in terms of 10 kPa.s (10^5 poises) was delineated by using a rotational viscometer while cooling the glass at 2°C/minute; the D.C. resistivity ($\log \rho$ at 250°C and 350°C) in terms of ohm-cm, dielectric constant (Die. Con.) at ambient temperatures and 1KHz, and the loss tangent (Loss Tan.) at
55 ambient temperature and 1KHz were defined following the techniques set out in ASTM D150, D257, and C657; and the chemical durability of the glass upon individual exposures to 5% aqueous HCl for 24 hours, to 5% aqueous NaOH for 6 hours, and to 0.02N aqueous Na₂CO₃ for 6 hours, each exposure being made

at 95° C and each loss being reported in terms of mg/cm², was ascertained on polished plates of glass both in terms of visual appearance (Appear.) and in terms of weight loss (Wt. Loss). With respect to appearance, N.C. indicates no change, S. Frost means slight frosting, and S. Haze reflects slight haze.

Glasses 17-26 were melted for 16 hours at 1650° C in an electrically-fired furnace. Glasses 23 and 24 were clear, viscous, and free from seeds; glasses 18, 20, and 22 were clear, but quite seedy; and the remainder of the glasses contained some seeds and were clear. The rest of the exemplary compositions were melted in a gas-oxygen furnace for 16 hours at 1800° C. At SiO₂ levels up to 81 mole percent, the glasses were clear, viscous, and free from seeds. At a concentration of 82 mole percent SiO₂, the glasses were clear, extremely viscous, and surprisingly, free from seeds. At a content of 83 mole percent SiO₂, the glasses appeared to be clear, but so viscous that they barely moved, and contained some seeds.

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TABLE II

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	<u>1</u>	<u>2</u>	<u>3</u>	<u>4</u>	<u>5</u>	<u>6</u>
Ann. Pt.	892	863	848	880	871	888
Str. Pt.	839	848	791	822	812	827
Coef. Exp.	36.2	31.9	27.8	34.0	33.9	36.9
Liquidus	1363	1408	1427	1428	1446	1319
Liq. Vis.	1.1	-	-	-	-	14
Log ρ						
250	12.74	-	-	-	-	11.97
350	10.76	-	-	-	-	10.09
Die. Con.	5.43	-	-	-	-	6.00
Loss Tan.	-	-	-	-	-	0.0022
5% HCl						
Appear.	N.C.	-	-	-	-	N.C.
Wt. Loss	0.02	-	-	-	-	0.01
5% NaOH						
Appear.	S. Haze	-	-	-	-	S.Frost
Wt. Loss	1.2	-	-	-	-	1.3
0.02N Na_2CO_3						
Appear.	N.C.	-	-	-	-	S.Haze
Wt. Loss	0.02	-	-	-	-	0.02

Table II (Cont.)

	<u>7</u>	<u>8</u>	<u>9</u>	<u>10</u>	<u>11</u>	<u>12</u>
Ann. Pt.	846	821	870	861	857	841
Str. Pt.	789	763	811	804	799	785
Coef. Exp.	30.2	26.5	33.3	32.0	31.7	27.8
Liquidus	1408	1443	1418	1436	1486	1417
	<u>13</u>			<u>14</u>	<u>15</u>	<u>16</u>
Ann. Pt.	888			871	873	884
Str. Pt.	829			811	810	826
Coef. Exp.	36.6			34.5	39.3	38
Liquidus	1414			1399	1336	1365
	<u>17</u>	<u>18</u>	<u>19</u>	<u>20</u>	<u>21</u>	<u>22</u>
Ann. Pt.	809	848	901	794	829	880
Str. Pt.	755	790	837	743	769	820
Coef. Exp.	45.1	41.2	37.0	48.3	44.3	41.2
Liquidus	1320	1346	1375	1301	1313	1426
	<u>23</u>	<u>24</u>	<u>25</u>	<u>26</u>	<u>27</u>	<u>28</u>
Ann. Pt.	785	807	849	885	820	811
Str. Pt.	736	750	797	832	772	758
Coef. Exp.	51.6	48.0	44.1	43.9	48.3	38.1
Liquidus	1209	1289	1433	1573	1421	1441
	<u>29</u>	<u>30</u>	<u>31</u>	<u>32</u>		
Ann. Pt.	865	-	834	893		
Str. Pt.	804	-	778	832		
Coef. Exp.	37.0	39.5	38.4	35.3		
Liquidus	1369	1416	1408	1382		

Table II (Cont.)

	<u>33</u>	<u>34</u>	<u>35</u>	<u>36</u>
Ann. Pt.	843	863	857	896
Str. Pt.	781	800	798	824
Coef. Exp.	36.1	35.4	35.5	33.7
Liquidus	1486	1457	1427	1400

	<u>37</u>	<u>38</u>	<u>39</u>
Ann. Pt.	889	877	904
Str. Pt.	831	818	841
Coef. Exp.	33.1	31.6	30.4
Liquidus	1390	1446	1457

	<u>40</u>	<u>41</u>	<u>42</u>	<u>43</u>	<u>44</u>	<u>45</u>
Ann. Pt.	893	-	884	908	907	901
Str. Pt.	834	827	826	846	846	838
Coef. Exp.	-	31.8	30.4	29.4	27.4	31.3
Liquidus	1381	-	1487	1436	1460	1461

	<u>46</u>	<u>47</u>	<u>48</u>
Ann. Pt.	897	887	885
Str. Pt.	835	824	822
Coef. Exp.	27.1	38.0	38.0
Liquidus	1450	1315	1342

A conjunctural study of Tables I and II clearly illustrates the criticality of maintaining the SrO, Al₂O₃, and SiO₂ concentrations within the prescribed ranges of the ternary system to prepare glasses capable of being melted at temperatures no higher than 1800°C, and which demonstrate annealing points greater than 875°C, linear coefficients of thermal expansion over the temperature range of 25°-300°C of 30-40x10⁻⁷/°C, liquidus temperatures not exceeding 1400°C, and liquidus viscosities of at least 10 kPa.s (100,000 poises). That study also indicates the generally unfavorable effects which the inclusion of MgO, CaO, ZnO, La₂O₃ and/or MnO can have upon the resulting glasses with respect to the above properties. In contrast, the addition of BaO can be useful in reducing the liquidus temperature without significantly affecting the viscosity of the glass, but care must be exercised in such additions because of volatilization of BaO and the development of seeds in the glass.

Claims

1. A clear, transparent, homogeneous, strontium aluminosilicate glass capable of being melted at temperatures no higher than 1800°C which exhibits an annealing point greater than 875°C, a linear coefficient of thermal expansion over the temperature range of 25°-300°C of $30-40 \times 10^{-7}/^{\circ}\text{C}$, a liquidus temperature not exceeding 1400°C, and a liquidus viscosity of at least 10 kPa.s (10^5 poises), said glass being essentially free from alkali metal oxides, MgO, CaO, ZnO, La_2O_3 , and MnO, and consisting essentially expressed in terms of mole percent on the oxide basis, of

9-12% SrO
9-12% Al_2O_3
77-82% SiO_2

2. A glass according to claim 1 wherein up to 8 mole percent BaO is substituted for SrO.

3. In a flat panel display service containing a flat transparent substrate for growing large crystal polysilicon semiconductor films thereon, the improvement wherein said substrate is a strontium aluminosilicate glass capable of being melted at temperatures no higher than 1800°C and which exhibits an annealing point greater than 875°C, a linear coefficient of thermal expansion over the temperature range of 25°-300°C of $30-40 \times 10^{-7}/^{\circ}\text{C}$, a liquidus temperature not exceeding 1400°C, and a liquidus viscosity of at least 10 kPa.s (10^5 poises), said glass being essentially free from alkali metal oxides, MgO, CaO, ZnO, La_2O_3 , and MnO, and consisting essentially, in terms of mole percent on the oxide basis, of

9-12% SrO
9-12% Al_2O_3
77-82% SiO_2

4. A device according to claim 3 wherein up to 8 mole percent BaO is substituted for SrO.

Revendications

1. Un verre d'aluminosilicate de strontium, homogène, transparent, clair, pouvant être fondu à des températures non supérieures à 1800°C, qui présente une température de recuit supérieure à 875°C, un coefficient de dilatation thermique linéaire sur l'intervalle de température de 25° à 300°C de 30 à $40 \times 10^{-7}/^{\circ}\text{C}$, une température de liquidus ne dépassant pas 1400°C et une viscosité au liquidus d'au moins 10 kPa.s (10^5 poises), ledit verre étant essentiellement exempt d'oxydes de métaux alcalins, MgO, CaO, ZnO, La_2O_3 et MnO, et étant essentiellement constitué, comme exprimé en pourcentages molaires sur la base des oxydes, de

9 à 12 % de SrO
9 à 12 % de Al_2O_3
77 à 82 % de SiO_2 .

2. Un verre selon la revendication 1, dans lequel BaO est substitué à SrO en une proportion d'au plus 8 moles pour cent.

3. Dans un dispositif d'affichage à panneau plat contenant un substrat plat transparent pour y développer des couches minces de semi-conducteurs en silicium polycristallin à grands cristaux, le perfectionnement dans lequel ledit substrat est un verre d'aluminosilicate de strontium pouvant être fondu à des températures non supérieures à 1800°C et qui présente une température de recuit supérieure à 875°C, un coefficient de dilatation thermique linéaire sur l'intervalle de température de 25° à 300°C de 30 à $40 \times 10^{-7}/^{\circ}\text{C}$, une température de liquidus ne dépassant pas 1400°C et une viscosité au liquidus d'au moins 10 kPa.s (10^5 poises), ledit verre étant essentiellement exempt d'oxydes de métaux alcalins, MgO, CaO, ZnO, La_2O_3 et MnO, et étant constitué essentiellement, comme exprimé en pourcentages molaires sur la base des oxydes, de

9 à 12 % de SrO
9 à 12 % de Al_2O_3
77 à 82 % de SiO_2 .



4. Un dispositif selon la revendication 3, dans lequel BaO est substitué à SrO en une proportion d'au plus 8 moles pour cent.

Patentansprüche

1. Klares, transparentes, homogenes Strontiumaluminiumsilicat-Glas, schmelzfähig bei Temperaturen nicht höher als 1800° C, welches eine obere Kühl- bzw. Entspannungstemperatur höher als 875° C, einen linearen thermischen Ausdehnungskoeffizient über den Temperaturbereich von 25 - 300° C von $30-40 \times 10^{-7}/^{\circ}\text{C}$, eine Liquidustemperatur, die 1400° C nicht überschreitet, und eine Liquidusviskosität von wenigstens 10 kPa.s (10^5 Poises) aufweist, wobei das Glas im wesentlichen frei ist von Alkalimetalloxiden, MgO, CaO, ZnO, La_2O_3 und MnO und im wesentlichen besteht aus, ausgedrückt in Mol.% auf der Basis der Oxide,

9 - 12 % SrO
9 - 12 % Al_2O_3
77 - 82 % SiO_2

2. Glas gemäß Anspruch 1, dadurch gekennzeichnet, daß bis zu 8 Mol.-% BaO für SrO substituiert ist.

3. Flache Anzeigeeinrichtung bzw. Plattenanzeigeeinrichtung mit einem flachen transparenten Substrat, zum Aufwachsen großer Kristalle von Polysilicon-Halbleiterfilmen darauf dadurch gekennzeichnet, daß das Substrat ein Strontiumaluminiumsilicat-Glas ist, das bei einer Temperatur nicht höher als 1800° C schmelzfähig ist und welches einen oberen Kühl- bzw. Entspannungspunkt, der höher als 875° C ist, einen linearen thermischen Ausdehnungskoeffizient über den Temperaturbereich von 25° - 300° C von $30-40 \times 10^{-7}/^{\circ}\text{C}$, eine Liquidustemperatur, die 1400° C nicht überschreitet, und eine Liquidusviskosität von wenigstens 10 kPa.s (10^5 Poises) aufweist, wobei das Glas im wesentlichen frei ist von Alkalimetalloxiden, MgO, CaO, ZnO, La_2O_3 und MnO und im wesentlichen besteht aus, ausgedrückt in Mol.% auf der Basis der Oxide,

9 - 12 % SrO
9 - 12 % Al_2O_3
77 - 82 % SiO_2

4. Vorrichtung gemäß Anspruch 3, dadurch gekennzeichnet, daß bis zu 8 Mol.% BaO für SrO substituiert ist.

