

PREPARATION OF EXO-OLEFIN TERMINATED POLYOLEFINS VIA QUENCHING WITH ALKOXYSILANES OR ETHERS

1. FIELD

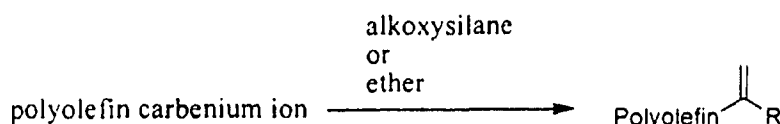
[0001] Provided herein are methods for preparing exo-olefin terminated polyolefins.

2. BACKGROUND

[0002] Exo-olefin terminated polyolefins, such as polyisobutylene (PIB), are useful precursors for the preparation of polymers containing specific functional end groups. Specifically, exo-olefin end groups may be transformed into other specific functional end groups. Polymers containing specific end groups have several useful purposes. For example, PIB-based succinimido amine dispersants are useful as engine lubricants. Thus, there is a need for methods of selectively or exclusively producing exo-olefin terminated polyolefins.

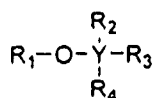
3. SUMMARY

[0003] In one embodiment, provided herein are methods for preparing a polyolefin containing one or more exo-olefinic end groups on the polyolefin chain, comprising quenching an ionized polyolefin with an alkoxy silane or an ether. In some embodiments, the method is represented by the following scheme:



wherein R is hydrogen or hydrocarbyl. In some embodiments, R is hydrocarbyl.

[0004] In some embodiments, the alkoxy silane or ether compound has the formula:



wherein Y is carbon or silicon;

R₁ is hydrocarbyl, and R₂-R₄ are each, independently, H or hydrocarbyl.

[0005] In some embodiments, the polyolefin containing one or more exo-olefinic end groups formed is at least 98 percent by mole of all products.

4. DETAILED DESCRIPTION

4.1 DEFINITIONS

[0006] Unless defined otherwise, all technical and scientific terms used herein have the same meaning as is commonly understood by one of ordinary skill in the art. In the event that there are a plurality of definitions for a term used herein, the definitions provided in this section prevail unless stated otherwise.

[0007] As used herein, "alcohol" refers to a compound of formula:



wherein R is aliphatic hydrocarbyl, and the -OH may be attached to a primary, secondary, or tertiary carbon.

[0008] As used herein, "alkyl" refers to a carbon chain or group containing from 1 to 20 carbons, or 1 to 16 carbons. Such chains or groups may be straight or branched. Exemplary alkyl groups herein include, but are not limited to, methyl, ethyl, propyl, isopropyl, isobutyl, n-butyl, sec-butyl, tert-butyl, isopentyl, neopentyl, tert-pentyl, or isohexyl. As used herein, "lower alkyl" refers to carbon chains or groups having from 1 carbon atom to about 6 carbon atoms.

[0009] As used herein, "alkenyl" refers to a carbon chain or group containing from 2 to 20 carbons, or 2 to 16 carbons, wherein the chain contains one or more double bonds. An example includes, but is not limited to, an allyl group. The double bond of an alkenyl carbon chain or group may be conjugated to another unsaturated group. An alkenyl carbon chain or group may be substituted with one or more heteroatoms. An alkenyl carbon chain or group may contain one or more triple bonds.

[0010] As used herein, "alkynyl" refers to a carbon chain or group containing from 2 to 20 carbons, or 2 to 16 carbons, wherein the chain contains one or more triple bonds. An example includes, but is not limited to, a propargyl group. The triple bond of an alkynyl carbon chain or group may be conjugated to another unsaturated group. An alkynyl carbon chain or group may be substituted with one or more heteroatoms. An alkynyl carbon chain or group may contain one or more double bonds.

[0011] As used herein, "aryl" refers to a monocyclic or polycyclic aromatic group containing from 6 to about 30 carbon atoms. Aryl groups include, but are not limited to, groups such as unsubstituted or substituted fluorenyl, unsubstituted or substituted phenyl, or unsubstituted or substituted naphthyl.

[0012] As used herein, "alkaryl" refers to an aryl group substituted with at least one alkyl, alkenyl, or alkynyl group.

[0013] As used herein, "aralkyl" refers to an alkyl, alkenyl, or alkynyl group substituted with at least one aryl group.

[0014] As used herein, "alkoxysilane" refers to a compound containing at least one silicon atom bonded to at least one alkoxy group.

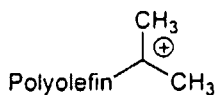
[0015] As used herein, "amine" refers to a compound of formula:



wherein R_1 , R_2 , and R_3 are each, independently, hydrogen or hydrocarbyl.

[0016] As used herein, "carbocation" and "carbenium ion" refer to a positively charged carbon atom.

[0017] As used herein, "carbocation terminated polyolefin" refers to a polyolefin containing at least one carbocation end group. Examples include, but are not limited to, compounds of the formula:



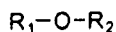
[0018] As used herein, "chain end concentration" refers to the sum of the molar concentration of carbocationic end groups and dormant end groups. When a mono-functional initiator is used, the chain end concentration is approximately equal to the initiator concentration. For a multi-functional initiator, when the functionality of the initiator equals x , then the chain end concentration is approximately equal to x times the initiator concentration.

[0019] As used herein, "coupled polyolefin" refers to the product of the addition of a carbocation terminated polyolefin to another polyolefin.

[0020] As used herein, "diluent" refers to a liquid diluting agent or compound. Diluents may be a single compound or a mixture of two or more compounds. Diluents may completely dissolve or partially dissolve the reaction components. Examples include, but are not limited to, hexane or methyl chloride, or mixtures thereof.

[0021] As used herein, "electron donor" refers to a molecule that is capable of donating a pair of electrons to another molecule. Examples include, but are not limited to, molecules capable of complexing with Lewis acids. Further examples include, but are not limited to, bases and/or nucleophiles. Further examples include, but are not limited to, molecules capable of abstracting or removing a proton. Further examples include, but are not limited to, pyridine derivatives.

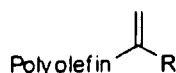
[0022] As used herein, "ether" refers to a compound of the formula:



wherein R_1 and R_2 are each, independently, hydrocarbyl.

[0023] As used herein “exo-olefin end group” or “exo-olefinic end group” refers to a terminal olefin moiety.

[0024] As used herein, “exo-olefin terminated polyolefin” refers to a polyolefin that contains at least one exo-olefin end group. Examples include, but are not limited to, compounds of the following formula:



wherein R is hydrogen or hydrocarbyl. In some embodiments, R is hydrocarbyl.

[0025] As used herein, “halide,” “halo,” or “halogen” refers to F, Cl, Br, or I.

[0026] As used herein, “hydrocarbyl” refers to a monovalent, linear, branched or cyclic group which contains only carbon and hydrogen atoms.

[0027] As used herein, “initiator” refers to a compound that provides a carbocation. Examples include, but are not limited to, compounds with one or more carbocation precursor groups. An initiator may be mono-functional or multi-functional. As used herein, “mono-functional initiator” refers to an initiator that provides approximately one stoichiometric equivalent of carbocation relative to initiator. As used herein, “multi-functional initiator” refers to an initiator that provides approximately x stoichiometric equivalents of carbocation relative to initiator, wherein x represents the functionality of the initiator. When a mono-functional initiator is used, the chain end concentration is approximately equal to the initiator concentration. For a multi-functional initiator, when the functionality of the initiator equals x, then the chain end concentration equals x times the initiator concentration.

[0028] As used herein, “ionized polyolefin” refers to a polyolefin containing at least one carbenium ion. An example includes, but is not limited to, a tert-halide terminated polyolefin that has been ionized into a cationic polyolefin. A further example includes, but is not limited to, a quasiliving carbocationic polyolefin. A further example includes, but is not limited to, an exo-olefin terminated polyolefin that has been ionized into an ionized polyolefin or quasiliving carbocationic polyolefin. A further example includes, but is not limited to, a polyolefin containing an olefin that has been ionized into a quasiliving carbocationic polyolefin or a cationic polyolefin.

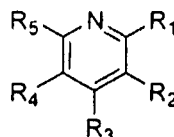
[0029] As used herein, "Lewis acid" refers to a chemical entity that is capable of accepting a pair of electrons. An example includes, but is not limited to, titanium tetrachloride.

[0030] As used herein, "monomer" refers to an olefin that is capable of combining with a carbocation to form another carbocation. An example includes, but is not limited to, isobutylene.

[0031] As used herein, "percent by mole of all products" refers to the proportion of the number of moles of a particular product of a reaction to the number of moles of all products of the reaction multiplied by one hundred.

[0032] As used herein, "polyolefin" refers to a polymer that comprises at least two olefin monomers. An example includes, but is not limited to, polyisobutylene.

[0033] As used herein, "pyridine derivative" refers to a compound of the formula:



wherein R₁, R₂, R₃, R₄, and R₅ are each, independently, hydrogen or hydrocarbyl; or R₁ and R₂, or R₂ and R₃, or R₃ and R₄, or R₄ and R₅ independently form a fused aliphatic ring of about 4 to about 7 carbon atoms or a fused aromatic ring of about 5 to about 7 carbon atoms. An example includes, but is not limited to, 2,6-lutidine.

[0034] As used herein, "quasiliving carbocationic polymerization conditions" refers to polymerization conditions that allow for the formation of quasiliving carbocationic polyolefins.

[0035] As used herein, "quasiliving carbocationic polyolefin" refers to a carbocationic polyolefin that has been formed under quasiliving polymerization conditions and is active in propagation in the presence of olefinic monomer.

[0036] As used herein, "quasiliving polymerization" refers to a carbocationic polymerization in which the rate of irreversible chain-breaking reactions is either zero or very low relative to the rate of propagation. Quasiliving polymerization proceeds by initiation and is followed by propagation, wherein propagating (active) polymer chains are in rapid equilibrium with non-propagating (dormant) polymer chains.

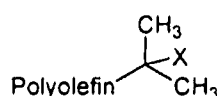
[0037] As used herein, "quasiliving polymerization conditions" refers to reaction conditions that allow quasiliving polymerization to occur.

[0038] As used herein, “quenching” refers to reacting a carbenium ion with a quenching agent.

[0039] As used herein, “quenching agent” refers to a compound that can, either alone or in combination with another compound, react with a carbenium ion.

[0040] As used herein, “terminator” refers to a compound that terminates polymerization via deactivation of the Lewis acid.

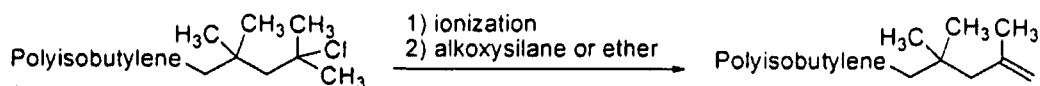
[0041] As used herein, “tert-halide terminated polyolefin” refers to a polyolefin that contains at least one tertiary halide end group. An example includes, but is not limited to, a compound of formula:



wherein X is a halogen.

4.2 METHODS

[0042] Provided herein are methods for preparing a polyolefin containing one or more exo-olefinic end groups on the polyolefin chain, comprising quenching an ionized polyolefin with an alkoxy silane or an ether. In some embodiments, the polyolefin prepared by the methods provided herein contains one exo-olefinic end group on the polyolefin chain. In some embodiments, the polyolefin prepared by the methods provided herein contains more than one exo-olefinic end group on the polyolefin chain. In some embodiments, the method is represented by the following scheme:

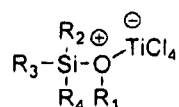


[0043] In some embodiments, the method is performed in a diluent. In some embodiments, multiple alkoxy silanes are used. In some embodiments, multiple ethers are used.

[0044] In some embodiments of the methods described herein, the quenching agent is an alkoxy silane or an ether. Without being limited to any theory, in some embodiments, the quenching agent is a complex formed between two compounds or atoms.

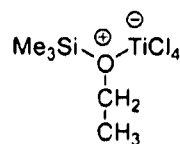
[0045] Without being limited to any theory, in some embodiments, the quenching agent is a complex between the alkoxy silane and a Lewis acid. Without being limited to any theory, in some embodiments, the quenching agent is a complex between the alkoxy silane and a

titanium tetrahalide. Without being limited to any theory, in some embodiments, the quenching agent is a complex between the alkoxysilane and titanium tetrachloride. Without being limited to any theory, in some embodiments, the quenching agent is a complex between allyloxytrimethylsilane, ethoxytrimethylsilane, or methoxytrimethylsilane and a Lewis acid. Without being limited to any theory, in some embodiments, the quenching agent is a complex between allyloxytrimethylsilane, ethoxytrimethylsilane, or methoxytrimethylsilane, and a titanium tetrahalide. Without being limited to any theory, in some embodiments, the quenching agent is a complex between allyloxytrimethylsilane, ethoxytrimethylsilane, or methoxytrimethylsilane, and titanium tetrachloride. Without being limited to any theory, in some embodiments, the quenching agent is a complex between ethoxytrimethylsilane and titanium tetrachloride. Without being limited to any theory, in some embodiments, the quenching agent is a 1:1 stoichiometric complex between an alkoxysilane and a Lewis acid. Without being limited to any theory, in some embodiments, the quenching agent is a 1:1 stoichiometric complex between an alkoxysilane and a titanium tetrahalide. Without being limited to any theory, in some embodiments, the quenching agent is a 1:1 stoichiometric complex between an alkoxytrimethylsilane and titanium tetrachloride. Without being limited to any theory, in some embodiments, the quenching agent is a 1:1 stoichiometric complex between ethoxytrimethylsilane and titanium tetrachloride. Without being limited to any theory, in some embodiments, the quenching agent is a compound of the following formula:



wherein R₁-R₄ are hydrocarbyl.

[0046] Without being limited to any theory, in some embodiments, the quenching agent is the following compound:



[0047] Without being limited to any theory, in some embodiments, the quenching agent is the alkoxysilane. Without being limited to any theory, in some embodiments, the quenching agent is allyloxytrimethylsilane, ethoxytrimethylsilane, or methoxytrimethylsilane.

[0048] Without being limited to any theory, in some embodiments, the quenching agent is a complex between the ether and Lewis acid. Without being limited to any theory, in some

embodiments, the quenching agent is a complex between the ether and a titanium tetrahalide. Without being limited to any theory, in some embodiments, the quenching agent is a complex between the ether and titanium tetrachloride. Without being limited to any theory, in some embodiments, the quenching agent is a complex between tert-butyl ethyl ether or tert-butyl methyl ether and a titanium tetrahalide. Without being limited to any theory, in some embodiments, the quenching agent is a complex between tert-butyl ethyl ether or tert-butyl methyl ether and titanium tetrachloride.

[0049] Without being limited to any theory, in some embodiments, more than one quenching agent act to quench the carbenium ion. For example, in the case of an alkoxysilane, the carbenium ion may be quenched with a mix of free alkoxysilane and complexed alkoxysilane.

[0050] Without being limited to any theory, in some embodiments, an electron donor is the quenching agent. Without being limited to any theory, in some embodiments, a pyridine derivative is the quenching agent. Without being limited to any theory, in some embodiments, 2,6-lutidine is the quenching agent.

[0051] In some embodiments, the quenching time is from about 1 minute to about 90 minutes. In some embodiments, the quenching time is greater than 60 minutes. In some embodiments, the quenching time is about 5 minutes. In some embodiments, the quenching time is about 10 minutes. In some embodiments, the quenching time is about 15 minutes. In some embodiments, the quenching time is about 20 minutes. In some embodiments, the quenching time is about 25 minutes. In some embodiments, the quenching time is about 30 minutes. In some embodiments, the quenching time is about 35 minutes. In some embodiments, the quenching time is about 40 minutes. In some embodiments, the quenching time is about 45 minutes. In some embodiments, the quenching time is about 50 minutes. In some embodiments, the quenching time is about 55 minutes. In some embodiments, the quenching time is about 60 minutes.

[0052] In some embodiments, the reaction is quenched at a temperature from about -150°C to about -30°C. In some embodiments, the reaction is quenched at a temperature from about -150°C to about -40°C. In some embodiments, the reaction is quenched at a temperature from about -100°C to about -30°C. In some embodiments, the reaction is quenched at a temperature from about -100°C to about -40°C. In some embodiments, the reaction is quenched at a temperature from about -90°C to about -60°C. In some embodiments, the reaction is quenched at a temperature of about -100°C. In some embodiments, the reaction is quenched at a temperature of about -90°C. In some

embodiments, the reaction is quenched at a temperature of about -80°C. In some embodiments, the reaction is quenched at a temperature of about -70°C. In some embodiments, the reaction is quenched at a temperature of about -60°C. In some embodiments, the reaction is quenched at a temperature of about -55°C. In some embodiments, the reaction is quenched at a temperature of about -50°C. In some embodiments, the reaction is quenched at a temperature of about -45°C. In some embodiments, the reaction is quenched at a temperature of about -40°C. In some embodiments, the reaction is quenched at a temperature of about -35°C. In some embodiments, the reaction is quenched at a temperature of about -30°C.

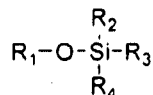
[0053] In some embodiments, the reaction mixture and isolated polyolefin products are largely free of colored precipitates.

4.2.1 QUENCHING AGENTS

[0054] Various embodiments of the quenching agents for use herein are described herein. All combinations of such embodiments are within the scope of this disclosure.

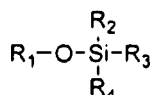
(a) Alkoxysilanes

[0055] In some embodiments, the alkoxysilane has the formula:



wherein R_1 is hydrocarbyl; and R_2 - R_4 are each, independently, hydrogen or hydrocarbyl.

[0056] In some embodiments, the alkoxysilane has the formula:



wherein:

[0057] R_1 is alkyl, alkenyl, alkynyl, aryl, alkaryl, or aralkyl; and

[0058] R_2 , R_3 , and R_4 are each, independently, hydrogen, alkyl, alkenyl, alkynyl, aryl, alkaryl, or aralkyl, or R_2 and R_3 , or R_2 and R_4 , or R_3 and R_4 independently form a fused aliphatic ring of about 3 to about 7 carbon atoms.

[0059] In some embodiments, R_1 is alkyl, alkenyl, alkynyl, aryl, alkaryl, or aralkyl. In some embodiments, R_1 is alkyl, alkenyl, alkynyl, or aralkyl. In some embodiments, R_1 is

alkyl, alkenyl, or alkynyl. In some embodiments, R_1 is alkyl. In some embodiments, R_1 is lower alkyl.

[0060] In some embodiments, R_1 is methyl, ethyl, or allyl. In some embodiments, R_1 is methyl. In some embodiments, R_1 is ethyl. In some embodiments, R_1 is allyl.

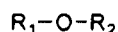
[0061] In some embodiments, R_2 , R_3 , and R_4 are each, independently, hydrogen, alkyl, alkenyl, alkynyl, aryl, alkaryl, or aralkyl, or R_2 and R_3 , or R_2 and R_4 , or R_3 and R_4 independently form a fused aliphatic ring of about 3 to about 7 carbon atoms. In some embodiments, R_2 , R_3 , and R_4 are each, independently, hydrogen, alkyl, alkenyl, alkynyl, or aralkyl, or R_2 and R_3 , or R_2 and R_4 , or R_3 and R_4 independently form a fused aliphatic ring of about 3 to about 7 carbon atoms. In some embodiments, R_2 , R_3 , and R_4 are each, independently, hydrogen, alkyl, alkenyl, or alkynyl. In some embodiments, R_2 , R_3 , and R_4 are each, independently, hydrogen or alkyl. In some embodiments, R_2 , R_3 , and R_4 are each, independently, alkyl. In some embodiments, R_2 , R_3 , and R_4 are each, independently, lower alkyl. In some embodiments, R_2 , R_3 , and R_4 are each methyl. In some embodiments, R_2 , R_3 , and R_4 are each hydrogen. In some embodiments, one of R_2 - R_4 is hydrogen, and the other two are both methyl.

[0062] In some embodiments, the alkoxysilane is a monoalkoxysilane, dialkoxysilane, trialkoxysilane, or tetraalkoxysilane. In some embodiments, the alkoxysilane is a monoalkoxytrialkylsilane, dialkoxydialkylsilane, trialkoxymonoalkylsilane, or a tetraalkoxysilane. In some embodiments, the alkoxysilane is a monoalkoxytrialkylsilane or dialkoxydialkylsilane. In some embodiments, the alkoxysilane is a monoalkoxytrialkylsilane. In some embodiments, the alkoxysilane is a monoalkoxytrimethylsilane, monoalkoxytripropylsilane, or monoalkoxytributylsilane. In some embodiments, the alkoxysilane is a monoalkoxytrimethylsilane.

[0063] In some embodiments, the alkoxysilane is allyloxytrimethylsilane, ethoxytrimethylsilane, or methoxytrimethylsilane. In some embodiments, the alkoxysilane is allyloxytrimethylsilane. In some embodiments, the alkoxysilane is ethoxytrimethylsilane. In some embodiments, the alkoxysilane is methoxytrimethylsilane.

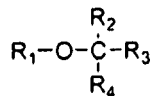
(b) Ethers

[0064] In some embodiments, the ether has the formula:



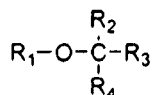
wherein R_1 and R_2 are each, independently, hydrocarbyl.

[0065] In some embodiments, the ether has the formula:



wherein R₁ is hydrocarbyl and R₂-R₄ are each, independently, hydrogen or hydrocarbyl.

[0066] In some embodiments, the ether has the formula:



wherein:

[0067] R₁ is alkyl, alkenyl, alkynyl, aryl, alkaryl, or aralkyl; and

[0068] R₂, R₃, and R₄ are each, independently, hydrogen, alkyl, alkenyl, alkynyl, aryl, alkaryl, or aralkyl, or R₂ and R₃, or R₂ and R₄, or R₃ and R₄ independently form a fused aliphatic ring of about 3 to about 7 carbon atoms.

[0069] In some embodiments, R₁ is alkyl, alkenyl, alkynyl, aryl, alkaryl, or aralkyl. In some embodiments, R₁ is alkyl, alkenyl, alkynyl, or aralkyl. In some embodiments, R₁ is alkyl, alkenyl, or alkynyl. In some embodiments, R₁ is alkyl. In some embodiments, R₁ is lower alkyl. In some embodiments, R₁ is methyl, ethyl, or allyl. In some embodiments, R₁ is methyl. In some embodiments, R₁ is ethyl. In some embodiments, R₁ is allyl.

[0070] In some embodiments, R₂, R₃, and R₄ are each, independently, hydrogen, alkyl, alkenyl, alkynyl, aryl, alkaryl, or aralkyl, or R₂ and R₃, or R₂ and R₄, or R₃ and R₄ independently form a fused aliphatic ring of about 3 to about 7 carbon atoms. In some embodiments, R₂, R₃, and R₄ are each, independently, hydrogen, alkyl, alkenyl, alkynyl, or aralkyl, or R₂ and R₃, or R₂ and R₄, or R₃ and R₄ independently form a fused aliphatic ring of about 3 to about 7 carbon atoms. In some embodiments, R₂, R₃, and R₄ are each, independently, hydrogen, alkyl, alkenyl, or alkynyl. In some embodiments, R₂, R₃, and R₄ are each, independently, hydrogen or alkyl. In some embodiments, R₂, R₃, and R₄ are each, independently, alkyl. In some embodiments, R₂, R₃, and R₄ are each, independently, lower alkyl. In some embodiments, R₂, R₃, and R₄ are each methyl. In some embodiments, R₂, R₃, and R₄ are each hydrogen.

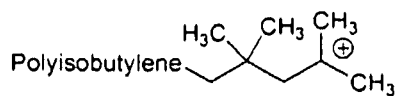
[0071] In some embodiments, the ether is an isopropyl alkyl ether, a sec-butyl alkyl ether, an isobutyl alkyl ether, or a tert-butyl alkyl ether. In some embodiments, the ether is a dialkyl ether. In some embodiments, the ether is diethyl ether or dipropyl ether. In some embodiments, the ether is a tert-butyl alkyl ether.

[0072] In some embodiments, the ether is tert-amyl methyl ether, isopropyl methyl ether, isopropyl ethyl ether, sec-butyl methyl ether, sec-butyl ethyl ether, isobutyl methyl ether, isobutyl ethyl ether, tert-butyl methyl ether, or tert-butyl ethyl ether. In some embodiments, the ether is tert-butyl ethyl ether or tert-butyl methyl ether. In some embodiments, the ether is tert-butyl ethyl ether. In some embodiments, the ether is tert-butyl methyl ether.

4.2.2 IONIZED POLYOLEFINS

[0073] Ionized polyolefins may be made by any method known to those of skill in the art. Examples include, but are not limited to, ionizing a tert-halide with a Lewis acid; ionizing a preformed polyolefin with a Lewis acid; or polymerizing an olefin monomer under quasiliving carbocationic polymerization conditions.

[0074] In some embodiments, the ionized polyolefin is a carbocation terminated polyolefin. In some embodiments, the ionized polyolefin contains one or more carbocation end groups. In some embodiments, the ionized polyolefin contains one carbocation end group. In some embodiments, the ionized polyolefin contains two carbocation end groups. In some embodiments, the ionized polyolefin contains three carbocation end groups. In some embodiments, the ionized polyolefin is a polyisobutylene with a cationic end group. In some embodiments, the ionized polyolefin is a compound of the following formula:



(a) **Ionized Polyolefins from tert-halides**

[0075] In some embodiments, the ionized polyolefin is derived from a tert-halide terminated polyolefin. In some embodiments, the ionized polyolefin is derived from a tert-chloride terminated polyolefin, tert-bromide terminated polyolefin, or tert-iodide terminated polyolefin. In some embodiments, the ionized polyolefin is derived from a tert-chloride terminated polyolefin or tert-bromide terminated polyolefin. In some embodiments, the ionized polyolefin is derived from a tert-chloride polyolefin.

[0076] In some embodiments, the ionized polyolefin is generated by contacting a tert-halide terminated polyolefin with a Lewis acid. In some embodiments, the ionized polyolefin is generated by contacting a tert-chloride terminated polyolefin, tert-bromide terminated polyolefin, or tert-iodide terminated polyolefin with a Lewis acid. In some embodiments, the ionized polyolefin is generated by contacting a tert-chloride terminated polyolefin with a Lewis acid.

[0077] In some embodiments, such Lewis acid is a non-protic acid. In some embodiments, such Lewis acid is a metal halide or non-metal halide. In some embodiments, such Lewis acid is a metal halide. In some embodiments, such Lewis acid is a titanium (IV) halide, a zinc (II) halide, a tin (IV) halide, or an aluminum (III) halide. In some embodiments, such Lewis acid is a titanium (IV) halide. In some embodiments, such Lewis acid is a tin (IV) halide. In some embodiments, such Lewis acid is an aluminum (III) halide. In some embodiments, such Lewis acid is titanium tetrabromide or titanium tetrachloride. In some embodiments, such Lewis acid is titanium tetrachloride. In some embodiments, such Lewis acid is zinc chloride. In some embodiments, such Lewis acid is ethyl aluminum dichloride. In some embodiments such Lewis acid is a non-metal halide. In some embodiments, the Lewis acid is an antimony (VI) halide, a gallium (III) halide, or a boron (III) halide. In some embodiments, the Lewis acid is boron trichloride. In some embodiments, the Lewis acid is a titanium tetrahalide, a boron trihalide, aluminum trichloride, tin tetrachloride, zinc chloride, or ethyl aluminum dichloride. In some embodiments, the Lewis acid is a titanium tetrachloride, titanium tetrabromide, or boron trichloride.

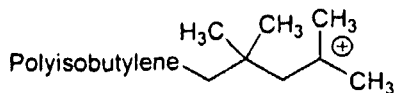
(b) **Ionized Polyolefins from preformed polyolefins**

[0078] In some embodiments, the ionized polyolefin is a preformed polyolefin. In some embodiments, such preformed polyolefin contains one or more double bonds. In some embodiments, such preformed polyolefin contains one double bond. In some embodiments, such preformed polyolefin is a polyisobutylene derivative. In some embodiments, such preformed polyolefin contains one or more endo olefins.

[0079] In some embodiments, the ionized polyolefin is generated by contacting a Lewis acid with a preformed polyolefin. In some embodiments, the ionized polyolefin is generated by contacting a preformed polyolefin containing one or more double bonds with a Lewis acid. In some embodiments, the ionized polyolefin is generated by contacting a preformed polyolefin containing one double bond with a Lewis acid. In some embodiments, the ionized polyolefin is generated by contacting a polyisobutylene derivative with a Lewis acid. In some embodiments, the ionized polyolefin is generated by contacting a preformed polyolefin containing one or more endo olefins with a Lewis acid.

(c) **Ionized Polyolefins from olefinic monomers under quasiliving carbocationic polymerization conditions**

[0080] In some embodiments, the ionized polyolefin is derived from olefinic monomers under quasiliving carbocationic conditions. Under such conditions, a quasiliving carbocationic polyolefin is generated. Such conditions may be achieved by any method known to those of skill in the art. In some embodiments, a monomer, an initiator, and a Lewis acid are used. In some embodiments, the ionized polyolefin is a quasiliving carbocationic polyisobutylene. In some embodiments, such quasiliving carbocationic polyolefin is a compound of the following formula:



(i) Initiators

[0081] In some embodiments, the initiator is a compound or polyolefin with one or more tertiary end groups. In some embodiments, the initiator has one tertiary end group. In some embodiments, the initiator has more than one tertiary end group.

[0082] In some embodiments, the initiator is a compound of formula $(X'-CR_aR_b)_nR_c$ wherein R_a , R_b and R_c independently comprise at least one of alkyl, aromatic, alkyl aromatic groups, and can be the same or different, and X' is an acetate, etherate, hydroxyl group, or a halogen. In some embodiments, R_c has a valence of n , and n is an integer of one to 4. In some embodiments, R_a , R_b and R_c are hydrocarbon groups containing one carbon atom to about 20 carbon atoms. In some embodiments, R_a , R_b and R_c are hydrocarbon groups containing one carbon atom to about 8 carbon atoms. In some embodiments, X' is a halogen. In some embodiments, X' is chloride. In some embodiments, the structure of R_a , R_b and R_c mimics the growing species or monomer. In some embodiments, such structure is a 1-phenylethyl derivative for polystyrene or a 2,4,4-trimethyl pentyl derivative for polyisobutylene. In some embodiments, the initiator is a cumyl, dicumyl or tricumyl halide. In some embodiments, chlorides are used. In some embodiments, the initiator is 2-chloro-2-phenylpropane, *i.e.*, cumyl chloride; 1,4-di(2-chloro-2-propyl)benzene, *i.e.*, di(cumylchloride); 1,3,5-tri(2-chloro-2-propyl)benzene, *i.e.*, tri(cumylchloride); 2,4,4-trimethyl-2-chloropentane; 2-acetyl-2-phenylpropane, *i.e.*, cumyl acetate; 2-propionyl-2-phenyl propane, *i.e.*, cumyl propionate; 2-methoxy-2-phenylpropane, *i.e.*, cumylmethyl ether; 1,4-di(2-methoxy-2-propyl)benzene, *i.e.*, di(cumylmethyl ether); or 1,3,5-tri(2-methoxy-2-propyl)benzene, *i.e.*, tri(cumylmethyl ether). In some embodiments, the initiator is 2-chloro-2,4,4-trimethyl pentane (TMPCI), 1,3-di(2-chloro-2-propyl)benzene,

1,3,5 tri(2-chloro-2-propyl)benzene, or 1,3-di(2-chloro-2-propyl)-5-tert-butylbenzene (bDCC).

[0083] In some embodiments, the initiator is mono-functional, bi-functional, or multi-functional. In some embodiments, the initiator is mono-functional. In some embodiments, the initiator is 2-chloro-2-phenylpropane, 2-acetyl-2-phenylpropane, 2-propionyl-2-phenylpropane, 2-methoxy-2-phenylpropane, 2-ethoxy-2-phenylpropane, 2-chloro-2,4,4-trimethylpentane, 2-acetyl-2,4,4-trimethylpentane, 2-propionyl-2,4,4-trimethylpentane, 2-methoxy-2,4,4-trimethylpentane, or 2-ethoxy-2,4,4-trimethylpentane. In some embodiments, the initiator is 2-chloro-2,4,4-trimethylpentane.

[0084] In some embodiments, the initiator is bi-functional. In some embodiments, the initiator is 1,3-di(2-chloro-2-propyl)benzene, 1,3-di(2-methoxy-2-propyl)benzene, 1,4-di(2-chloro-2-propyl)benzene, 1,4-di(2-methoxy-2-propyl)benzene, or 5-*tert*-butyl-1,3-di(2-chloro-2-propyl) benzene. In some embodiments, the initiator is 5-*tert*-butyl-1,3-di(2-chloro-2-propyl) benzene.

[0085] In some embodiments, the initiator is multi-functional. In some embodiments, the initiator is 1,3,5-tri(2-chloro-2-propyl)benzene or 1,3,5-tri(2-methoxy-2-propyl)benzene.

(ii) **Monomers**

[0086] In some embodiments, the monomer is a hydrocarbon monomer, *i.e.*, a compound containing only hydrogen and carbon atoms, including but not limited to, olefins and diolefins, and those having from about 2 to about 20 carbon atoms. In some embodiments, such compounds have from about 4 to about 8 carbon atoms.

[0087] In some embodiments, the methods described herein can be employed for the polymerization of such monomers to produce polymers of different molecular weights. In some embodiments, the methods described herein can be employed for the polymerization of such monomers to produce polymers of different, but uniform molecular weights. In some embodiments, such molecular weight is from about 300 to in excess of a million g/mol. In some embodiments, such polymers are low molecular weight liquid or viscous polymers having a molecular weight of from about 200 to 10,000 g/mol, or solid waxy to plastic, or elastomeric materials having molecular weights of from about 100,000 to 1,000,000 g/mol, or more.

[0088] In some embodiments, the monomer is isobutylene, styrene, beta pinene, isoprene, butadiene, or substituted compounds of the preceding types. In some embodiments, the monomer is isobutylene, 2-methyl-1-butene, 3-methyl-1-butene, 4-methyl-1-pentene, or beta-

pinene. In some embodiments, the monomer is isobutylene, 2-methyl-1-butene, 3-methyl-1-butene, or 4-methyl-1-pentene. In some embodiments, the monomer is isobutylene.

[0089] In some embodiments, mixtures of monomers may be used, *i.e.*, more than one monomer is used.

4.2.3 LEWIS ACIDS

[0090] In some embodiments, the Lewis acid is a non-protic acid. In some embodiments, the Lewis acid is a metal halide or non-metal halide. In some embodiments, the Lewis acid is a metal halide. In some embodiments, the Lewis acid is a titanium (IV) halide, a zinc (II) halide, a tin (IV) halide, or an aluminum (III) halide. In some embodiments, the Lewis acid is a titanium (IV) halide. In some embodiments, the Lewis acid is a tin (IV) halide. In some embodiments, the Lewis acid an aluminum (III) halide. In some embodiments, the Lewis acid is titanium tetrabromide or titanium tetrachloride. In some embodiments, the Lewis acid is titanium tetrachloride. In some embodiments, the Lewis acid is zinc chloride. In some embodiments, the Lewis acid is ethyl aluminum dichloride. In some embodiments the Lewis acid is a non-metal halide. In some embodiments, the Lewis acid is an antimony (VI) halide, a gallium (III) halide, or a boron (III) halide. In some embodiments, the Lewis acid is boron trichloride. In some embodiments, the Lewis acid is a titanium tetrahalide, a boron trihalide, aluminum trichloride, tin tetrachloride, zinc chloride, or ethyl aluminum dichloride. In some embodiments, the Lewis acid is a titanium tetrachloride, titanium tetrabromide, or boron trichloride.

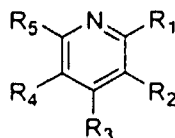
4.2.4 ELECTRON DONORS

[0091] As is understood to one of ordinary skill in the art, some electron donors are capable of converting traditional polymerization systems into quasiliving polymerization systems. In some embodiments, the methods described herein are performed in the presence of an electron donor.

[0092] In some embodiments, the electron donor is capable of complexing with Lewis acids. In some embodiments, the electron donor is a base and/or nucleophile. In some embodiments, the electron donor is capable of abstracting or removing a proton. In some embodiments, the electron donor is an organic base. In some embodiments, the electron donor is an amide. In some embodiments, the electron donor is N,N-dimethylformamide, N,N-dimethylacetamide, or N,N-diethylacetamide. In some embodiments, the electron donor is a sulfoxide. In some embodiments, the electron donor is dimethyl sulfoxide. In some embodiments, the electron donor is an ester. In some embodiments, the electron donor is

methyl acetate or ethyl acetate. In some embodiments, the electron donor is a phosphate compound. In some embodiments, the electron donor is trimethyl phosphate, tributyl phosphate, or triamide hexamethylphosphate. In some embodiments, the electron donor is an oxygen-containing metal compound. In some embodiments, the electron donor is tetraisopropyl titanate.

[0093] In some embodiments, the electron donor is pyridine or a pyridine derivative. In some embodiments, the electron donor is a compound of the formula:



wherein R₁, R₂, R₃, R₄, and R₅ are each, independently, hydrogen or hydrocarbyl; or R₁ and R₂, or R₂ and R₃, or R₃ and R₄, or R₄ and R₅ independently form a fused aliphatic ring of about 3 to about 7 carbon atoms or a fused aromatic ring of about 5 to about 7 carbon atoms. In some embodiments, R₁ and R₅ are each, independently, hydrocarbyl, and R₂-R₄ are hydrogen.

[0094] In some embodiments, the electron donor is 2,6-di-*tert*-butylpyridine, 2,6-lutidine, 2,4-dimethylpyridine, 2,4,6-trimethylpyridine, 2-methylpyridine, or pyridine. In some embodiments, the electron donor is N,N-dimethylaniline or N,N-dimethyltoluidine. In some embodiments, the electron donor is 2,6-lutidine.

[0095] In some embodiments, more than one electron donor is used.

4.2.5 DILUENTS

[0096] In some embodiments of the methods described herein, the methods are performed in a diluent. In some embodiments, the diluent is a single compound or a mixture of two or more compounds. In some embodiments, the diluent completely dissolves the reaction components or partially dissolves the reaction components. In some embodiments, the diluent completely or nearly completely dissolves the reaction components. In some embodiments, the diluent completely dissolves the reaction components. In some embodiments, the diluent nearly completely dissolves the reaction components.

[0097] In some embodiments, the diluent has a low boiling point and/or low freezing point. In some embodiments, the diluent is an alkane, an alkyl monohalide, or an alkyl polyhalide. In some embodiments, the diluent is a normal alkane. In some embodiments, the diluent is propane, normal butane, normal pentane, normal hexane, normal heptane, normal octane, normal nonane or normal decane. In some embodiments, the diluent is a branched

alkane. In some embodiments, the alkane is isobutane, isopentane, neopentane, isohexane, 3-methylpentane, 2,2-dimethylbutane, or 2,3-dimethylbutane. In some embodiments, the diluent is a halogenated alkane. In some embodiments, the diluent is chloroform, ethylchloride, n-butyl chloride, methylene chloride, methyl chloride, 1,2-dichloroethane, 1,1,2,2-tetrachloroethane, carbon tetrachloride, 1,1-dichloroethane, n-propyl chloride, isopropyl chloride, 1,2-dichloropropane, or 1,3-dichloropropane. In some embodiments, the diluent is an alkene or halogenated alkene. In some embodiments, the diluent is vinyl chloride, 1,1-dichloroethene, or 1,2-dichloroethene. In some embodiments, the diluent is carbon disulfide, sulfur dioxide, acetic anhydride, acetonitrile, benzene, toluene, methylcyclohexane, chlorobenzene, or nitroalkane. In some embodiments, the diluent is a mixture of the compounds in this paragraph.

[0098] In some embodiments, the diluent is a mixture of hexane and methyl chloride. In some embodiments, such mixture is from about 30/70 to about 70/30 hexane/methyl chloride by volume. In some embodiments, such mixture is from about 50/50 to about 100/0 hexane/methyl chloride by volume. In some embodiments, such mixture is about 60/40 hexane/methyl chloride by volume.

4.2.6 TEMPERATURE

[0099] In some embodiments, the methods described herein are performed at a temperature from about -150°C to about -30°C. In some embodiments, the temperature is from about -150°C to about -40°C. In some embodiments, the temperature is from about -100°C to about -30°C. In some embodiments, the temperature is from about -100°C to about -40°C. In some embodiments, the temperature is from about -90°C to about -60°C. In some embodiments, the temperature is about -100°C. In some embodiments, the temperature is about -90°C. In some embodiments, the temperature is about -80°C. In some embodiments, the temperature is about -70°C. In some embodiments, the temperature is about -60°C. In some embodiments, the temperature is about -55°C. In some embodiments, the temperature is about -50°C. In some embodiments, the temperature is about -45°C. In some embodiments, the temperature is about -40°C. In some embodiments, the temperature is about -35°C. In some embodiments, the temperature is about -30°C.

[00100] Without being bound to any theory, in some embodiments, the addition of an electron donor has been found to increase the yield of exo-olefin terminated polyolefin. Without being limited to any theory, in some embodiments, when the quenching occurs at higher temperatures (e.g. greater than -50°C) the addition of an electron donor has been found

to increase the yield of *exo*-terminated polyolefin. Without being limited to any theory, in some embodiments, when the quenching occurs at higher temperatures (*e.g.* greater than -50°C) the addition of 2,6-lutidine has been found to increase the yield of *exo*-terminated polyolefin. In some embodiments, the reaction temperature is from about -50°C to about -30°C, the 2,6-lutidine concentration is about 0.05 M, the titanium tetrachloride concentration is about 0.2 M, the chain end concentration is about 0.035 M, the ethoxytrimethylsilane concentration is about 0.1 M, and the *exo*-olefin terminated polyolefin formed is at least 90 percent by mole of all products.

[00101] In some embodiments, the methods described herein are performed in the presence of 2,6 lutidine, at a temperature of about -70°C to about -30°C, and the polyolefin containing one or more *exo*-olefinic end groups formed is at least 90 percent by mole of all products.

[00102] In some embodiments, the methods described herein are performed in the presence of 2,6 lutidine, at a temperature of about -70°C to about -30°C, and the polyolefin containing one or more *exo*-olefinic end groups formed is at least 80 percent by mole of all products.

[00103] In some embodiments, the methods described herein are performed in the presence of 2,6 lutidine, at a temperature of about -70°C to about -30°C, and the polyolefin containing one or more *exo*-olefinic end groups formed is at least 70 percent by mole of all products.

4.2.7 CONCENTRATIONS

[00104] In some embodiments, the alkoxysilane or ether is present in stoichiometric excess relative to chain end. In some embodiments, the alkoxysilane or ether is present at a concentration less than the chain end concentration.

[00105] In some embodiments, the alkoxysilane is present at a concentration of from about 1.1 to about 10 times the chain end concentration. In some embodiments, the alkoxysilane is present at a concentration of about 0.7 times to 7.5 times the chain end concentration. In some embodiments, the alkoxysilane is present at a concentration of about 0.85 to 6 times the chain end concentration. In some embodiments, the alkoxysilane is present at a concentration of from about 1.5 times to about 7.5 times the chain end concentration. In some embodiments, the alkoxysilane is present at a concentration of from about 2 times to about 6 times the chain end concentration. In some embodiments, the alkoxysilane is present at a concentration of from about 2 times to about 4 times the chain end concentration. In some

embodiments, the alkoxysilane is present at a concentration of from about 2 times to about 3 times the chain end concentration. In some embodiments, the alkoxysilane is present at a concentration of about 1.5 times the chain end concentration. In some embodiments, the alkoxysilane is present at a concentration of about 2 times the chain end concentration. In some embodiments, the alkoxysilane is present at a concentration of about 3 times the chain end concentration. In some embodiments, the alkoxysilane is present at a concentration of about 4 times the chain end concentration. In some embodiments, the alkoxysilane is present at a concentration of about 5 times the chain end concentration. In some embodiments, the alkoxysilane is present at a concentration of about 6 times the chain end concentration. In some embodiments, the alkoxysilane is present at a concentration of about 7 times the chain end concentration. In some embodiments, the alkoxysilane is present at a concentration of about 0.7 times the chain end concentration. In some embodiments, the alkoxysilane is present at a concentration of about 0.8 times the chain end concentration. In some embodiments, the alkoxysilane is present at a concentration of about 0.85 times the chain end concentration. In some embodiments, the alkoxysilane is present at a concentration of about 0.9 times the chain end concentration. In some embodiments, the alkoxysilane is present at a concentration of about 0.95 times the chain end concentration. In some embodiments, the alkoxysilane concentration is about equal to the chain end concentration.

[00106] In some embodiments, the ether is present at a concentration of from about 1.1 to about 10 times the chain end concentration. In some embodiments, the ether is present at a concentration of about 0.7 times to 7.5 times the chain end concentration. In some embodiments, the ether is present at a concentration of about 0.85 to 6 times the chain end concentration. In some embodiments, the ether is present at a concentration of from about 1.5 times to about 7.5 times the chain end concentration. In some embodiments, the ether is present at a concentration of from about 2 times to about 6 times the chain end concentration. In some embodiments, the ether is present at a concentration of from about 2 to about 4 times the chain end concentration. In some embodiments, the ether is present at a concentration of from about 2 to about 3 times the chain end concentration. In some embodiments, the ether is present at a concentration of about 1.5 times the chain end concentration. In some embodiments, the ether is present at a concentration of about 2 times the chain end concentration. In some embodiments, the ether is present at a concentration of about 3 times the chain end concentration. In some embodiments, the ether is present at a concentration of about 4 times the chain end concentration. In some embodiments, the ether is present at a concentration of about 5 times the chain end concentration. In some embodiments, the ether

is present at a concentration of about 6 times the chain end concentration. In some embodiments, the ether is present at a concentration of about 7 times the chain end concentration. In some embodiments, the ether is present at a concentration of about 0.7 times the chain end concentration. In some embodiments, the ether is present at a concentration of about 0.8 times the chain end concentration. In some embodiments, the ether is present at a concentration of about 0.85 times the chain end concentration. In some embodiments, the ether is present at a concentration of about 0.9 times the chain end concentration. In some embodiments, the ether is present at a concentration of about 0.95 times the chain end concentration. In some embodiments, the ether concentration is about equal to the chain end concentration.

[00107] In some embodiments, the alkoxysilane is present at a concentration of from about 0.15 to about 4 times the Lewis acid concentration. In some embodiments, the alkoxysilane is present at a concentration of from about 0.5 to about 3 times the Lewis acid concentration. In some embodiments, the alkoxysilane is present at a concentration of from about 0.75 to about 2 times the Lewis acid concentration. In some embodiments, the alkoxysilane concentration is about 0.5 times the Lewis acid concentration. In some embodiments, the alkoxysilane concentration is about 0.75 times the Lewis acid concentration. In some embodiments, the alkoxysilane concentration is about equal to the Lewis acid concentration. In some embodiments, the alkoxysilane concentration is about 0.2 times the Lewis acid concentration. In some embodiments, the alkoxysilane concentration is about 0.3 times the Lewis acid concentration. In some embodiments, the alkoxysilane concentration is about 0.4 times the Lewis acid concentration. In some embodiments, the alkoxysilane concentration is about 0.5 times the Lewis acid concentration. In some embodiments, the alkoxysilane concentration is about 0.6 times the Lewis acid concentration. In some embodiments, the alkoxysilane concentration is about 0.7 times the Lewis acid concentration. In some embodiments, the alkoxysilane concentration is about 0.8 times the Lewis acid concentration. In some embodiments, the alkoxysilane concentration is about 0.9 times the Lewis acid concentration. In some embodiments, the alkoxysilane concentration is about 1.0 times the Lewis acid concentration. In some embodiments, the alkoxysilane concentration is about 1.5 times the Lewis acid concentration. In some embodiments, the alkoxysilane concentration is about 2 times the Lewis acid concentration. In some embodiments, the alkoxysilane concentration is about 2.5 times the Lewis acid concentration. In some embodiments, the alkoxysilane concentration is about 3 times the Lewis acid concentration.

[00108] In some embodiments, the ether is present at a concentration of from about 0.15 to about 4 times the Lewis acid concentration. In some embodiments, the ether is present at a concentration of from about 0.5 to about 3 times the Lewis acid concentration. In some embodiments, the ether is present at a concentration of from about 0.75 to about 2 times the Lewis acid concentration. In some embodiments, the ether concentration is about 0.5 times the Lewis acid concentration. In some embodiments, the ether concentration is about 0.75 times the Lewis acid concentration. In some embodiments, the ether concentration is about equal to the Lewis acid concentration. In some embodiments, the ether concentration is about 0.2 times the Lewis acid concentration. In some embodiments, the ether concentration is about 0.3 times the Lewis acid concentration. In some embodiments, the ether concentration is about 0.4 times the Lewis acid concentration. In some embodiments, the ether concentration is about 0.5 times the Lewis acid concentration. In some embodiments, the ether concentration is about 0.6 times the Lewis acid concentration. In some embodiments, the ether concentration is about 0.7 times the Lewis acid concentration. In some embodiments, the ether concentration is about 0.8 times the Lewis acid concentration. In some embodiments, the ether concentration is about 0.9 times the Lewis acid concentration. In some embodiments, the ether concentration is about 1.0 times the Lewis acid concentration. In some embodiments, the ether concentration is about 1.5 times the Lewis acid concentration. In some embodiments, the ether concentration is about 2 times the Lewis acid concentration. In some embodiments, the ether concentration is about 2.5 times the Lewis acid concentration. In some embodiments, the ether concentration is about 3 times the Lewis acid concentration.

[00109] In some embodiments, the methods described herein are performed at a chain end concentration from about 0.005 M to about 0.2 M. In some embodiments, the methods described herein are performed at a chain end concentration from about 0.010 M to about 0.14 M. In some embodiments, the methods described herein are performed at a chain end concentration from about 0.014 M to about 0.10 M. In some embodiments, the methods described herein are performed at a chain end concentration from about 0.014 M to about 0.040 M.

[00110] In some embodiments, the chain end concentration is from about 0.10 M to about 0.15 M, and the exo-olefin terminated polyolefin formed is at least 70 percent by mole of all products. In some embodiments, the chain end concentration is from about 0.10 M to about 0.15 M, and the exo-olefin terminated polyolefin formed is at least 80 percent by mole of all products. In some embodiments, the chain end concentration is from about 0.10 M to about

0.15 M, and the exo-olefin terminated polyolefin formed is at least 90 percent by mole of all products.

[00111] In some embodiments, the Lewis acid is added in two or more aliquots. In some embodiments, the Lewis acid is added in two aliquots. In some embodiments, a first aliquot of Lewis acid is added during the polymerization reaction, and a second aliquot of Lewis acid is added during the quenching reaction. In some embodiments, a first aliquot of titanium tetrachloride is added during the polymerization reaction, and a second aliquot is added during the quenching reaction. Without being bound to any theory, if alkoxysilane quenching is performed in situ from a direct polymerization (as opposed to re-initiation of tert-chloride polyisobutylene), the optimum TiCl_4 concentration may be too high for the prior polymerization reaction; this often occurs for reactions with relatively high chain-end concentrations. When this situation occurs, it is manifested in a polymerization that is too fast and/or too exothermic. In such case, in some embodiments, a lower concentration first aliquot of Lewis acid may be used for the polymerization, and after the alkoxysilane or ether has been added, a second aliquot of Lewis acid may be added such that the total Lewis acid concentration has been increased to that which is optimum for quenching.

In some embodiments, the alkoxysilane is added after the addition of Lewis acid. In some embodiments, the ether is added after the addition of Lewis acid. In some embodiments, the alkoxysilane is added prior to the addition of Lewis acid. In some embodiments, the ether is added prior to the addition of Lewis acid.

4.2.8 EXO-OLEFIN SELECTIVITY

[00112] In some embodiments, the methods described herein selectively provide exo-olefin terminated polyolefins. In some embodiments, exo-olefin terminated polyolefin, polyolefins containing endo olefins, tert-halide polyolefins, and coupled polyolefins are reaction products. In some embodiments, the exo-olefin terminated polyolefin is the major product, and polyolefins containing endo olefins, tert-halide polyolefins, and coupled polyolefins are the minor products.

[00113] In some embodiments, the exo-olefin terminated polyolefin formed is at least 40 percent by mole of all products. In some embodiments, the exo-olefin terminated polyolefin formed is at least 50 percent by mole of all products. In some embodiments, the exo-olefin terminated polyolefin formed is at least 60 percent by mole of all products. In some embodiments, the exo-olefin terminated polyolefin formed is at least 75 percent by mole of all products. In some embodiments, the exo-olefin terminated polyolefin formed is at least 85

percent by mole of all products. In some embodiments, the exo-olefin terminated polyolefin formed is at least 90 percent by mole of all products. In some embodiments, the exo-olefin terminated polyolefin formed is at least 95 percent by mole of all products. In some embodiments, the exo-olefin terminated polyolefin formed is at least 97 percent by mole of all products. In some embodiments, the exo-olefin terminated polyolefin formed is at least 98 percent by mole of all products.

4.2.9 TERMINATORS

[00114] In some embodiments, the terminator is a compound capable of deactivating a Lewis acid. In some embodiments, the terminator is a base and/or a nucleophile. In some embodiments, the terminator is a base. In some embodiments, the terminator is an electron donor. In some embodiments, the terminator is an organic base. In some embodiments, the terminator is an alcohol or amine. In some embodiments, the terminator is an alcohol. In some embodiments, the terminator is a pyridine derivative.

[00115] In some embodiments, the terminator is methanol, ethanol, or isopropanol. In some embodiments, the terminator is methanol. In some embodiments, the terminator is water. In some embodiments, the terminator is diethylamine, triethylamine, pyridine, 2,6-lutidine, n-butylamine, or tert-amylamine.

[00116] In some embodiments, the terminator is added after the quenching step.

5. EXAMPLES

5.1 DEFINITIONS OF ABBREVIATIONS

- [00117] As used herein, "allyloxyTMS" and "AOTMS" refer to allyloxytrimethylsilane.
- [00118] As used herein, "coupled" refers to coupled polyolefin.
- [00119] As used herein, "CE" refers to chain end.
- [00120] As used herein, "DI" refers to deionized.
- [00121] As used herein, "endo" refers to a polyolefin containing an endo-olefin.
- [00122] As used herein, "ethoxyTMS" and "EtOTMS" refer to ethoxytrimethylsilane.
- [00123] As used herein, "exo" refers to exo-olefin terminated polyolefin.
- [00124] As used herein, "GPC" refers to gel permeation chromatography.
- [00125] As used here, "GPC-MALLS" refers to gel permeation chromatography-multi-angle laser light scattering.
- [00126] As used here, "hex" refers to hexane.
- [00127] As used herein, "IB" refers to isobutylene.

- [00128] As used herein, "26Lut" refers to 2,6-lutidine.
- [00129] As used herein, "MeCl" refers to methyl chloride.
- [00130] As used herein, "MeOH" refers to methyl alcohol.
- [00131] As used herein, "methoxyTMS" and "MeOTMS" refer to methoxytrimethylsilane.
- [00132] As used herein, " M_n " refers to number average molecular weight.
- [00133] As used herein, " M_w " refers to weight average molecular weight.
- [00134] As used herein, "NMR" refers to nuclear magnetic resonance.
- [00135] As used herein, "PDI" refers to polydispersity index.
- [00136] As used herein, "PIB" refers to polyisobutylene.
- [00137] As used herein, "RT" refers to room temperature.
- [00138] As used herein, "tert-halide" refers to tert-halide polyolefins.
- [00139] As used herein, " γ -Cl" refers to tert-chloride terminated polyolefin.
- [00140] As used herein, "TMPCl" refers to 2-chloro-2,4,4-trimethylpentane.
- [00141] As used herein "x CE" refers to times chain end concentration. For example, "2 x CE" refers to 2 times chain end concentration.

5.2 EXAMPLE 1

[00142] This example involved a direct polymerization of IB from TMPCl with *in situ* quenching in a single reactor. The quencher was allyloxytrimethylsilane. The conditions used were as follows:

[TMPCl] = 0.014 M

[IB] = 0.5 M

[TiCl₄] = 0.083 M

[allyloxyTMS] = 0.019 M

-60°C

60/40 (v/v) Hex/MeCl

[00143] For this polymerization and all subsequent polymerizations described herein, real-time, attenuated total reflectance (ATR) Fourier Transform infrared (FTIR) spectroscopy was used to monitor monomer conversion. spectra were gathered in real time using the ReactIR 4000 reaction analysis system (light conduit type) (ASI Applied Systems, Millersville, MD) equipped with a DiComp (diamond-composite) insertion probe and a general purpose type PR-11 platinum resistance thermometer. Reaction conversion was determined by monitoring the area, above a two point baseline, of the absorbance centered at 887 cm⁻¹, associated with the =CH₂ wag of IB.

[00144] The polymerization procedure was as follows: A 250 mL glass reactor, equipped with an overhead mechanical stirrer and platinum resistance thermometer, was fitted to the end of the DiComp probe of the ReactIR 4000, and this assembly was immersed into a -60°C constant-temperature heptane bath within the glove box. To the reactor were added 114.0 mL hexane, 76.0 mL MeCl, 0.48 mL (2.8 mmol, 0.014 M) TMPCl, and 0.23 mL (2.0 mmol, 0.01 M) 26Lut. The mixture was stirred to achieve thermal equilibrium, and a solvent background spectrum was acquired and subtracted from all subsequent spectra. At this point, FTIR data were continuously acquired to establish a solvent reference absorbance area, A_r , prior to monomer addition. Next, IB, 8.18 mL (0.5 M), was charged to the reactor, thermal equilibration was established, and data were again acquired prior to initiation, to establish the initial absorbance area, A_o , associated with the initial monomer concentration. The coinitiator, 1.82 mL (0.017 mol, 0.083 M, 6xCE) TiCl_4 (neat and at room temperature), was then charged to the reactor. FTIR data were acquired continuously, and monomer conversion was calculated in real time as shown with Equation 1,

$$\frac{[\text{IB}]_o - [\text{IB}]_t}{[\text{IB}]_o} = 1 - \frac{A_t - A_r}{A_o - A_r} \quad \text{Equation 1}$$

where A_t is the instantaneous absorbance area at 887 cm^{-1} . The polymerization reaction proceeded until complete IB conversion, and a pre-quench aliquot was removed. Then 0.66 mL (3.9 mmol, 0.019 M, 1.4xCE) AOTMS (room temp) was added and aliquots were taken periodically and after 60 min quenching time, prechilled MeOH was added to terminate the reaction. Upon removal from the dry box, the terminated reaction mixture was allowed to warm to room temperature. Hexane (1-2 mL) was added to the prequench and other aliquot samples and a sample of the final polymer, and then the polymers were precipitated into MeOH. The recovered PIB was agitated (vortexed) with fresh MeOH to remove any remaining contaminants, and the MeOH was decanted. Samples were dried with gaseous nitrogen before NMR analysis.

[00145] NMR analysis of the pre-quench sample yielded a molar end group composition (%) as follows:

Exo	2
<i>r</i> -Cl	98
Coupled	0

[00146] NMR analysis of the final post-quench sample yielded a molar end group composition (%) as follows:

Exo	96
<i>i</i> -Cl	2
Coupled	2

5.3 EXAMPLE 2

[00147] This example involved a direct polymerization of IB from TMPCl with *in situ* quenching in a single reactor. The quencher was methoxytrimethylsilane. The conditions used were as follows:

[TMPCl] = 0.014 M
 [IB] = 0.5 M
 [TiCl₄] = 0.084 M
 [methoxyTMS] = 0.028 M
 -60°C
 60/40 (v/v) Hex/MeCl

[00148] The procedure was as follows: A stirred 250 mL glass reactor was equilibrated to -60.8°C within the glove box. To the reactor were added 90.9 mL hexane, 60.6 mL MeCl, 0.38 mL (2.2 mmol, 0.014 M) TMPCl, 6.54 mL (0.5 M) IB, 0.06 mL (0.003 M) 26Lut, and 1.47 mL (0.013 mol, 0.084 M, 6xCE) TiCl₄. The polymerization reaction proceeded for 20 min, a pre-quench aliquot was removed, and then 0.62 mL (4.5 mmol, 0.028 M, 2xCE) MeOTMS (room temp) was added. The temperature of the reaction mixture rose to a peak of -59.8°C and then gradually fell. Aliquots were taken periodically and after 60 min quenching time, prechilled MeOH was added to terminate the reaction. Upon removal from the dry box, the terminated reaction mixture was allowed to warm to room temperature. Hexane (1-2 mL) was added to the prequench and other aliquot samples and a sample of the final polymer, and then the polymers were precipitated into MeOH. The recovered PIB was agitated (vortexed) with fresh MeOH to remove any remaining contaminants, and the MeOH was decanted. Samples were dried with gaseous nitrogen before NMR analysis.

[00149] NMR analysis of the pre-quench sample yielded a molar end group composition (%) as follows:

Exo	1
<i>i</i> -Cl	99
Coupled	0

[00150] NMR analysis of the final post-quench sample yielded a molar end group composition (%) as follows:

Exo	98
<i>i</i> -Cl	1
Coupled	1

5.4 **EXAMPLE 3**

[00151] This example involved re-ionization of an isolated and purified PIB sample carrying *tert*-chloride end groups. The quencher was ethoxytrimethylsilane.

[00152] The *tert*-chloride-terminated PIB was produced according to the following procedure: A stirred 1000 mL glass reactor was immersed into a -75°C constant-temperature heptane/hexane bath within the glove box. Into the flask were charged 654.5 mL MeCl, 6.12 mL (0.036 mol, 0.045 M) TMPCl, 0.58 mL (4.9 mmol, 0.0062 M) 26Lut, 97.2 mL (1.2 mol, 1.5 M) IB, and 41.6 mL (0.53 mol, 0.68 M) BCl₃ (neat). The polymerization reaction reached high conversion of the IB (~99%) in about 5 h, at which time the reaction was terminated by careful addition of excess, prechilled MeOH. The terminated reaction mixture was transferred to a 1000mL beaker, and the solvents were allowed to evaporate overnight. The polymer and remaining non-volatiles were diluted with hexane and were washed several times with MeOH and slightly acidified water. The polymer solution was stirred over MgSO₄ in order to remove any residual water, and the polymer was isolated by filtration and dried at room temperature in a vacuum oven. The theoretical M_n for the masterbatch polymer was 2,049 g/mol and GPC-MALLS analysis of the product using a dn/dc of 0.110 yielded M_n = 1,890 g/mol and M_w/M_n = 1.01. A stock solution of the masterbatch polymer was created by dissolving 9.54 g of the polymer in 60 mL hexane. The molar concentration of this solution was 0.067 mol/L, based on the theoretical M_n of the polymer.

NMR analysis of the end group composition (%) yielded the following:

Exo	9.1
Endo	0
<i>i</i> -Cl	90.8
Coupled	0.1

[00153] The conditions that were used for quenching were as follows:

[CE] = 0.037 M
[ethoxyTMS] = 0.28 M

[TiCl₄] = 0.23 M
 -70°C
 60/40 (v/v) Hex/MeCl

[00154] Quenching was carried out as follows: A 50 mL glass test tube was equilibrated to -70°C within the glove box. To the tube were added 6.0 mL of a 0.067 M solution of PIB masterbatch, 4.0 mL MeCl, and 0.47 mL (3.0 mmol, 0.28 M, 7.5 x[CE]) EtOTMS, and the contents were allowed to equilibrate. Then 0.27 mL (2.5 mmol, 0.23 M, 6.1 x[CE]) TiCl₄ was added and the tube was shaken by hand periodically. After 60 min quenching time, prechilled MeOH was added to terminate the reaction. Upon removal from the dry box, the terminated reaction mixture was allowed to warm to room temperature. Hexane (1-2 mL) was added to the polymer, and then the polymer was precipitated into MeOH. The recovered PIB was agitated (vortexed) with fresh MeOH to remove any remaining contaminants, and the MeOH was decanted. Samples were dried with gaseous nitrogen before NMR analysis.

[00155] NMR analysis of the final post-quench sample yielded a molar end group composition (%) as follows:

Exo	97.1
Endo	1.9
<i>t</i> -Cl	0
Coupled	1.0

5.5 EXAMPLE 4

[00156] This example involved re-ionization of an isolated and purified PIB sample carrying *tert*-chloride end groups. The quencher was ethoxytrimethylsilane.

[00157] The *tert*-chloride-terminated PIB was produced according to the following procedure: A stirred 1000 mL glass reactor was equilibrated to -60°C within the glove box. To the reactor were added 361.3 mL hexane, 240.9 mL MeCl, 5.72 mL (0.034 mol, 0.045 M) TMPCl, 122.7 mL (2.0 M) IB, 0.87 mL (0.01 M) 26Lut, and 18.5 mL (0.17 mol, 0.22 M, 5xCE) TiCl₄. The polymerization reaction proceeded for 15 min, reaching a maximum temperature of -37.4°C. After full conversion of the IB, prechilled MeOH was added to terminate the reaction. Upon removal from the dry box, the terminated reaction mixture was allowed to warm to room temperature. Hexane was added to the polymer and then the solution was washed several times each with MeOH, acidic DI water, and DI water. The final sample was dried in vacuo at room temp to remove any remaining solvents prior to

NMR and GPC analysis. The theoretical M_n for the masterbatch polymer was 2,650 g/mol and GPC-MALLS analysis of the product using a dn/dc of 0.111 yielded $M_n = 2,850$ g/mol and $PDI = 1.03$. A stock solution of the masterbatch polymer was created by dissolving 11.553 g of the polymer in 60 mL hexane. The molar concentration of this solution was 0.061 mol/L, based on the theoretical M_n of the polymer.

[00158] NMR analysis of the end group composition (%) yielded the following:

Exo	17.1
Endo	8.0
<i>r</i> -Cl	73.6
Coupled	0.7

[00159] The conditions that were used for quenching were as follows:

[CE] = 0.035 M
 [ethoxyTMS] = 0.096 M
 [TiCl₄] = 0.19 M
 -60°C
 60/40 (v/v) Hex/MeCl

[00160] Quenching was carried out as follows: A 50 mL glass test tube was equilibrated to -70°C within the glove box. To the tube were added 6.0 mL of a 0.061 M solution of PIB masterbatch, 4.0 mL MeCl, and 0.156 mL (1.0 mmol, 0.096 M, $2.7 \times [CE]$) EtOTMS, and the contents were allowed to equilibrate. Then 0.22 mL (2.0 mmol, 0.19 M, $5.5 \times [CE]$) TiCl₄ was added and the tube was shaken by hand periodically. After 30 min quenching time, prechilled MeOH was added to terminate the reaction. Upon removal from the dry box, the terminated reaction mixture was allowed to warm to room temperature. Hexane (1-2 mL) was added to the polymer, and then the polymer was precipitated into MeOH. The recovered PIB was agitated (vortexed) with fresh MeOH to remove any remaining contaminants, and the MeOH was decanted. Samples were dried with gaseous nitrogen before NMR analysis.

[00161] NMR analysis of the final post-quench sample yielded a molar end group composition (%) as follows:

Exo	90.1
Endo	9.9
<i>r</i> -Cl	0.0
Coupled	0.0

5.6 EXAMPLE 5

[00162] This example involved a direct polymerization of IB with *in situ* quenching in a single reactor. The quencher was *tert*-butyl ethyl ether. The conditions used were as follows:

$[TMPCl] = 0.014\text{ M}$
 $[IB] = 0.5\text{ M}$
 $[TiCl_4] = 0.084\text{ M}$
 $[t\text{-butyl ethyl ether}] = 0.028\text{ M}$
 -60°C
 60/40 (v/v) Hex/MeCl

[00163] Quenching was carried out as follows:

[00164] A stirred 250 mL glass reactor was equilibrated to -60°C within the glove box. To the reactor were added 90.9 mL hexane, 60.6 mL MeCl, 0.38 mL (2.2 mmol, 0.014 M) TMPCl, 6.54 mL (0.5 M) IB, 0.06 mL (0.003 M) 26Lut, and 1.47 mL (0.013 mol, 0.084 M, 6xCE) $TiCl_4$. The polymerization reaction proceeded for 20 min and then 0.62 mL (4.5 mmol, 0.028 M, 2xCE) *t*-butyl ethyl ether (room temp) was added. Aliquots were taken periodically and after 60 min quenching time, prechilled MeOH was added to terminate the reaction. Upon removal from the dry box, the terminated reaction mixture was allowed to warm to room temperature. Hexane (1-2 mL) was added to the aliquot samples and a sample of the final polymer, and then the polymers were precipitated into MeOH. The recovered PIB was agitated (vortexed) with fresh MeOH to remove any remaining contaminants, and the MeOH was decanted. Samples were dried with gaseous nitrogen before NMR analysis.

[00165] NMR analysis of the final post-quench sample yielded a molar end group composition (%) as follows:

Exo	83
<i>t</i> -Cl	14
Coupled	3

5.7 EXAMPLE 6

[00166] This example involved a direct polymerization of IB with *in situ* quenching in a single reactor. The quencher was ethoxytrimethylsilane. The conditions used were as follows:

Polymerization:
 $[TMPCl] = 0.13\text{ M}$
 $[IB] = 5.2\text{ M}$
 $[TiCl_4] = 0.032\text{ M}$

Quench:
 $[CE] = 0.10\text{ M}$
 $[TiCl_4] = 0.83\text{ M}$
 $[EthoxyTMS] = 0.63\text{ M}$

[00167] A stirred 1000 mL glass reactor was equilibrated to -77.3°C within the glove box. To the reactor were added 253.9 mL hexane, 169.3 mL MeCl, 16.4 mL (0.096 mol, 0.13 M) TMPCl, 306.7 mL (5.2 M) IB, 1.05 mL (0.012 M) 26Lut, and 2.6 mL (0.024 mol, 0.032 M, 0.25xCE) TiCl₄. The polymerization reaction proceeded for 100 min, reaching a maximum temperature of -61.2°C. When the temperature of the reactor contents had fallen back to -70.4°C, a pre-quench aliquot was removed, and then 90.0 mL (0.58 mol, 6xCE) EtOTMS (slightly chilled) was added, followed immediately by addition of 81.7 mL (0.75 mol, 7.72xCE) TiCl₄ (neat, RT). The temperature of the reaction mixture rose to a peak of -52.1°C and then gradually fell. Aliquots were taken periodically and after 30 min quenching time, prechilled MeOH was added to terminate the reaction. The reaction was quite viscous, but seemed to terminate completely after addition of MeOH. Upon removal from the dry box, the terminated reaction mixture was allowed to warm to room temperature. Hexane (1-2 mL) was added to the prequench and other aliquot samples and a sample of the final polymer, and then the polymers were precipitated into MeOH. The recovered PIB was agitated (vortexed) with fresh MeOH to remove any remaining contaminants, and the MeOH was decanted. Samples were dried with gaseous nitrogen before NMR analysis, and then in vacuo at room temp to remove any remaining solvents prior to GPC analysis.

[00168] NMR analysis of the pre-quench sample yielded a molar end group composition (%) as follows:

Exo	9.5
Endo	2.1
<i>l</i> -Cl	85.8
Coupled	2.7

[00169] NMR analysis of the final post-quench sample yielded a molar end group composition (%) as follows:

Exo	94.6
Endo	1.9
<i>l</i> -Cl	0.0
Coupled	3.5

[00170] GPC traces (RI detector) of pre-quench and post-quench aliquots were identical. GPC-MALLS yielded $M_n = 3,300$ g/mol and $M_w/M_n = 1.1$.

5.8 EXAMPLE 7

[00171] This example involved a direct polymerization of IB from TMPCl with *in situ* quenching in a single reactor. The quencher was allyloxytrimethylsilane. The conditions used were as follows:

[TMPCl] = 0.014 M
 [IB] = 0.5 M
 [TiCl₄] = 0.083 M
 [26Lut] = 0.01 M
 [allyloxyTMS] = 0.012 M
 -60°C
 60/40 (v/v) Hex/MeCl

[00172] The procedure was as follows: A stirred 250 mL glass reactor was equilibrated to -60.0°C within the glove box. To the reactor were added 114.0 mL hexane, 76.0 mL MeCl, 0.48 mL (2.8 mmol, 0.014 M) TMPCl, 8.18 mL (0.5 M) IB, 0.23 mL (2.0 mmol, 0.01 M) 26Lut, and 1.82 mL (0.017 mol, 0.083 M, 6xCE) TiCl₄. The polymerization reaction proceeded until complete IB conversion. Then 0.41 mL (2.4 mmol, 0.012 M, 0.86xCE) AOTMS (room temp) was added and aliquots were taken periodically and after 62 min quenching time, prechilled MeOH was added to terminate the reaction. Upon removal from the dry box, the terminated reaction mixture was allowed to warm to room temperature. Hexane (1-2 mL) was added to the aliquot samples and a sample of the final polymer, and then the polymers were precipitated into MeOH. The recovered PIB was agitated (vortexed) with fresh MeOH to remove any remaining contaminants, and the MeOH was decanted. Samples were dried with gaseous nitrogen before NMR analysis.

[00173] NMR analysis of the final post-quench sample yielded a molar end group composition (%) as follows:

Exo	79
<i>t</i> -Cl	13
Coupled	8

5.9 EXAMPLE 8

[00174] This example involved a direct polymerization of IB from TMPCl with *in situ* quenching in a single reactor. The quencher was allyloxytrimethylsilane. The conditions used were as follows:

[TMPCl] = 0.14 M
 [IB] = 5.0 M
 [TiCl₄] = 0.28 M
 [allyloxyTMS] = 0.28 M

[26Lut] = 0.10 M
 -60°C
 60/40 (v/v) Hex/MeCl

[00175] The procedure was as follows: A stirred 250 mL glass reactor was equilibrated to -60.0°C within the glove box. To the reactor were added 31.5 mL hexane, 21.0 mL MeCl, 2.38 mL (0.014 mol, 0.14 M) TMPCl, 40.9 mL (5.0 M) IB, 1.16 mL (0.01 mol, 0.1 M) 26Lut, and 3.07 mL (0.028 mol, 0.28 M, 2xCE) TiCl₄. The polymerization reaction proceeded until complete IB conversion. Then 4.6 mL (0.27 mol, 0.026 M, 1.9xCE) AOTMS (room temp) was added and aliquots were taken periodically and after 105 min quenching time, prechilled MeOH was added to terminate the reaction. Upon removal from the dry box, the terminated reaction mixture was allowed to warm to room temperature. Hexane (1-2 mL) was added to the aliquot samples and a sample of the final polymer, and then the polymers were precipitated into MeOH. The recovered PIB was agitated (vortexed) with fresh MeOH to remove any remaining contaminants, and the MeOH was decanted. Samples were dried with gaseous nitrogen before NMR analysis.

[00176] NMR analysis of the final post-quench sample yielded a molar end group composition (%) as follows:

Exo	85
<i>t</i> -Cl	15
Coupled	1

5.10 EXAMPLE 9

[00177] This example involved re-ionization of an isolated and purified PIB sample carrying *tert*-chloride end groups. The quencher was ethoxytrimethylsilane. The PIB masterbatch that was used for this example is described in Example 4 above.

[00178] The conditions that were used for quenching were as follows:

[CE] = 0.036 M
 [ethoxyTMS] = 0.097 M
 [TiCl₄] = 0.097 M
 [26Lut] = 0.024 M
 -60°C
 60/40 (v/v) Hex/MeCl

[00179] Quenching was carried out as follows: A 50 mL glass test tube was equilibrated to -60°C within the glove box. To the tube were added 6.0mL of a 0.061M solution of PIB masterbatch, 4.0 mL MeCl, 0.029 mL (0.25 mmol, 0.024 M, 0.68 x[CE]) 26Lut, and 0.156

mL (1.0 mmol, 0.097 M, 2.7 x[CE]) EtOTMS, and the contents were allowed to equilibrate. Then 0.11 mL TiCl₄ (1.0 mmol, 0.097 M, 2.7 x[CE]) was added and the tube was shaken by hand periodically. After 30 min quenching time, prechilled MeOH was added to terminate the reaction. Upon removal from the dry box, the terminated reaction mixture was allowed to warm to room temperature. Hexane (1-2 mL) was added to the polymer, and then the polymer was precipitated into MeOH. The recovered PIB was agitated (vortexed) with fresh MeOH to remove any remaining contaminants, and the MeOH was decanted. Samples were dried with gaseous nitrogen before NMR analysis.

[00180] NMR analysis of the final post-quench sample yielded a molar end group composition (%) as follows:

Exo	89.3
Endo	8.9
<i>t</i> -Cl	0.0
Coupled	1.8

5.11 EXAMPLE 10

[00181] This example involved re-ionization of an isolated and purified PIB sample carrying *tert*-chloride end groups. The quencher was ethoxytrimethylsilane. The PIB masterbatch that was used for this example is described in Example 4 above.

[00182] The conditions that were used for quenching were as follows:

[CE] = 0.035 M
 [ethoxyTMS] = 0.096 M
 [TiCl₄] = 0.19 M
 [26Lut] = 0.048 M
 -60°C
 60/40 (v/v) Hex/MeCl

[00183] Quenching was carried out as follows: A 50 mL glass test tube was equilibrated to -60°C within the glove box. To the tube were added 6.0mL of a 0.061M solution of PIB masterbatch, 4.0 mL MeCl, 0.058 mL (0.50 mmol, 0.048 M, 1.4 x[CE]) 26Lut, and 0.156 mL (1.0 mmol, 0.096 M, 2.7 x[CE]) EtOTMS, and the contents were allowed to equilibrate. Then 0.22 mL (2.0 mmol, 0.19 M, 5.5 x[CE]) TiCl₄ was added and the tube was shaken by hand periodically. After 30 min quenching time, prechilled MeOH was added to terminate the reaction. Upon removal from the dry box, the terminated reaction mixture was allowed to warm to room temperature. Hexane (1-2 mL) was added to the polymer, and then the polymer was precipitated into MeOH. The recovered PIB was agitated (vortexed) with fresh

MeOH to remove any remaining contaminants, and the MeOH was decanted. Samples were dried with gaseous nitrogen before NMR analysis.

[00184] NMR analysis of the final post-quench sample yielded a molar end group composition (%) as follows:

Exo	87.7
Endo	8.8
<i>i</i> -Cl	0.0
Coupled	3.5

5.12 EXAMPLE 11

[00185] This example involved re-ionization of an isolated and purified PIB sample carrying *tert*-chloride end groups. The quencher was ethoxytrimethylsilane. The PIB masterbatch that was used for this example is described in Example 4 above.

[00186] The conditions that were used for quenching were as follows:

[CE] = 0.035 M
[ethoxyTMS] = 0.096 M
[TiCl₄] = 0.19 M
[26Lut] = 0.05 M
-60°C
60/40 (v/v) Hex/MeCl

[00187] Quenching was carried out as follows: A 50 mL glass test tube was equilibrated to -60°C within the glove box. To the tube were added 6.0 mL of a 0.061 M solution of PIB masterbatch, 4.0 mL MeCl, and 0.06 mL (0.52 mmol, 0.05 M, 1.4 x[CE]) 26Lut, and the contents were allowed to equilibrate. Then 0.22 mL (2.0 mmol, 0.19 M, 5.5 x[CE]) TiCl₄ was added and the tube was shaken by hand. Then 0.156 mL (1.0 mmol, 0.096 M, 2.7 x[CE]) EtOTMS was added and the tube was shaken by hand periodically. After 30 min quenching time, prechilled MeOH was added to terminate the reaction. Upon removal from the dry box, the terminated reaction mixture was allowed to warm to room temperature. Hexane (1-2 mL) was added to the polymer, and then the polymer was precipitated into MeOH. The recovered PIB was agitated (vortexed) with fresh MeOH to remove any remaining contaminants, and the MeOH was decanted. Samples were dried with gaseous nitrogen before NMR analysis.

[00188] NMR analysis of the final post-quench sample yielded a molar end group composition (%) as follows:

Exo	85.5
Endo	11.1
<i>t</i> -Cl	0.0
Coupled	3.4

5.13 EXAMPLE 12

[00189] This example involved re-ionization of an isolated and purified PIB sample carrying *tert*-chloride end groups. The quencher was ethoxytrimethylsilane. The PIB masterbatch that was used for this example is described in Example 4 above.

[00190] The conditions that were used for quenching were as follows:

[CE] = 0.035 M
 [ethoxyTMS] = 0.096 M
 [TiCl₄] = 0.19 M
 [26Lut] = 0.006 M.
 -30°C
 60/40 (v/v) Hex/MeCl

[00191] Quenching was carried out as follows: A 50 mL glass test tube was equilibrated to -30°C within the glove box. To the tube were added 6.0mL of a 0.061M solution of PIB masterbatch, 4.0 mL MeCl, 0.007 mL (0.06 mmol, 0.006 M, 0.16 x[CE]) 26Lut, and 0.156 mL (1.0 mmol, 0.096 M, 2.7 x[CE]) EtOTMS, and the contents were allowed to equilibrate. Then 0.22 mL (2.0 mmol, 0.19 M, 5.5 x[CE]) TiCl₄ was added and the tube was shaken by hand periodically. After 30 min quenching time, prechilled MeOH was added to terminate the reaction. Upon removal from the dry box, the terminated reaction mixture was allowed to warm to room temperature. Hexane (1-2 mL) was added to the polymer, and then the polymer was precipitated into MeOH. The recovered PIB was agitated (vortexed) with fresh MeOH to remove any remaining contaminants, and the MeOH was decanted. Samples were dried with gaseous nitrogen before NMR analysis.

[00192] NMR analysis of the final post-quench sample yielded a molar end group composition (%) as follows:

Exo	40.2
Endo	18.1
<i>t</i> -Cl	36.0
Coupled	5.6

5.14 EXAMPLE 13

[00193] This example involved re-ionization of an isolated and purified PIB sample carrying *tert*-chloride end groups. The quencher was ethoxytrimethylsilane. The PIB masterbatch that was used for this example is described in Example 4 above.

[00194] The conditions that were used for quenching were as follows:

[CE] = 0.035 M
[ethoxyTMS] = 0.096 M
[TiCl₄] = 0.19 M
[26Lut] = 0.05 M
-30°C
60/40 (v/v) Hex/MeCl

[00195] Quenching was carried out as follows: A 50 mL glass test tube was equilibrated to -30°C within the glove box. To the tube were added 6.0 mL of a 0.061 M solution of PIB masterbatch, 4.0 mL MeCl, 0.06 mL (0.52 mmol, 0.05 M, 1.4 x[CE]) 26Lut, and 0.156 mL (1.0 mmol, 0.096 M, 2.7 x[CE]) EtOTMS, and the contents were allowed to equilibrate. Then 0.22 mL TiCl₄ (2.0 mmol, 0.19 M, 5.5 x[CE]) was added and the tube was shaken by hand periodically. After 30 min quenching time, prechilled MeOH was added to terminate the reaction. Upon removal from the dry box, the terminated reaction mixture was allowed to warm to room temperature. Hexane (1-2 mL) was added to the polymer, and then the polymer was precipitated into MeOH. The recovered PIB was agitated (vortexed) with fresh MeOH to remove any remaining contaminants, and the MeOH was decanted. Samples were dried with gaseous nitrogen before NMR analysis.

[00196] NMR analysis of the final post-quench sample yielded a molar end group composition (%) as follows:

Exo	90.1
Endo	9.0
<i>t</i> -Cl	0.0
Coupled	0.9

5.15 **EXAMPLE 14**

[00197] This example involved re-ionization of an isolated and purified PIB sample carrying *tert*-chloride end groups. The quencher was ethoxytrimethylsilane. The PIB masterbatch that was used for this example is described in Example 3 above.

[00198] The conditions that were used for quenching were as follows:

[CE] = 0.038 M

[ethoxyTMS] = 0.28 M
[TiCl₄] = 0.095 M
-70°C
60/40 (v/v) Hex/MeCl

[00199] Quenching was carried out as follows: A 50 mL glass test tube was equilibrated to -70°C within the glove box. To the tube were added 6.0mL of a 0.067M solution of PIB masterbatch, 4.0 mL MeCl, and 0.47 mL (3.0 mmol, 0.28 M, 7.5 x[CE]) EtOTMS, and the contents were allowed to equilibrate. Then 0.11 mL TiCl₄ (1.0 mmol, 0.095 M, 2.5 x[CE]) was added and the tube was shaken by hand periodically. After 60 min quenching time, prechilled MeOH was added to terminate the reaction. Upon removal from the dry box, the terminated reaction mixture was allowed to warm to room temperature. Hexane (1-2 mL) was added to the polymer, and then the polymer was precipitated into MeOH. The recovered PIB was agitated (vortexed) with fresh MeOH to remove any remaining contaminants, and the MeOH was decanted. Samples were dried with gaseous nitrogen before NMR analysis.

[00200] NMR analysis of the final post-quench sample yielded a molar end group composition (%) as follows:

Exo	66.7
Endo	1.3
<i>r</i> -Cl	31.3
Coupled	0.7

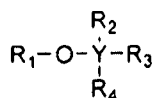
[00201] Since modifications would be apparent to those of skill in the art, this disclosure is intended to be limited only by the scope of the appended claims.

CLAIMS

What is claimed is:

1. A method for preparing a polyolefin containing one or more exo-olefinic end groups on the polyolefin chain, comprising quenching an ionized polyolefin with one or more alkoxysilane or ether compounds.

2. The method of claim 1, wherein the alkoxysilane or ether compound has the formula:



wherein Y is carbon or silicon;

R₁ is hydrocarbyl, and R₂-R₄ are each, independently, hydrogen or hydrocarbyl.

3. The method of claim 2, wherein:
R₁ is alkyl, alkenyl, alkynyl, aryl, alkaryl, or aralkyl; and
R₂, R₃, and R₄ are each, independently, hydrogen, alkyl, alkenyl, alkynyl, aryl, alkaryl, or aralkyl, or R₂ and R₃, or R₂ and R₄, or R₃ and R₄ independently form a fused aliphatic ring of about 3 to about 7 carbon atoms.

4. The method of claims 2 or 3, wherein R₁ is methyl, ethyl, or allyl.

5. The method of any of claims 2-4, wherein R₂, R₃, and R₄ are each methyl.

6. The method of any of claims 2-5, wherein the alkoxysilane or ether compound is allyloxytrimethylsilane, ethoxytrimethylsilane, methoxytrimethylsilane, tert-butyl ethyl ether or tert-butyl methyl ether.

7. The method of any of claims 1-6, wherein the ionized polyolefin is a quasiliving carbocationic polyolefin and the method is performed under quasiliving polymerization conditions.

8. The method of claim 7 wherein the quasiliving carbocationic polyolefin is prepared by adding a Lewis acid and one or more monomers to an initiator.

9. The method of claim 8, wherein the initiator is 2-chloro-2,4,4-trimethylpentane.
10. The method of claims 8 or 9, wherein the monomer is isobutylene.
11. The method of any of claims 1-6, wherein the ionized polyolefin is formed by adding a Lewis acid to a tert-halide terminated polyolefin.
12. The method of any of claims 8-11, wherein the Lewis acid is titanium tetrachloride.
13. The method of any of claims 1-12, wherein the method is performed at a temperature from about -150°C to about -30°C, from about -90°C to about -60°C, or about -60°C.
14. The method of any of claims 1-13, wherein the method is performed in the presence of one or more electron donors.
15. The method of any of claims 1-14, wherein the polyolefin containing one or more exo-olefinic end groups formed is at least 40 percent, at least 60 percent, at least 80 percent, at least 85 percent, at least 90 percent, at least 95 percent, at least 97 percent, or at least 98 percent by mole of all products.