

SPECIFICATION

COPPER ELECTROPLATING LIQUID, PRETREATMENT LIQUID FOR COPPER ELECTROPLATING AND METHOD OF COPPER ELECTROPLATING

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TECHNICAL FIELD

The present invention relates to a copper electroplating solution for forming copper wiring on a semiconductor wafer, and more specifically to a copper electroplating solution, 10 a pretreatment liquid for copper electroplating, and a method of copper electroplating suitable for filling fine via holes or trenches formed on a semiconductor wafer without producing defects such as voids or seams.

15 BACKGROUND ART

In processing semiconductor wafers, aluminum has chiefly been used as a wiring material; however, as the density of wirings integration increases, copper, which has higher electric conductivity, has substituted aluminum to prevent 20 increase in the delay time of signals.

For the formation of copper wirings, the Damascene method is used, in which after forming a wiring pattern on a silicon wafer, a barrier layer and a seed layer are formed using a CVD method or a sputtering method, then the wiring pattern 25 is filled using electroplating, and excessively deposited copper is removed using CMP (chemical-mechanical polishing).

In order to form a copper wiring on the semiconductor wafer, as described above, it is required to fill via holes or trenches. In recent times, however, since the wiring patterns have become finer, the coverage of the copper seed
5 layers in via holes or trenches formed using sputtering or the like have become poor, and the portions where the copper films is extremely thin are often produced.

Although an electroplating solution based on an aqueous solution containing copper sulfate acidified with sulfuric
10 acid has generally used for the above-described copper electroplating, the use of such an electroplating solution acidified with sulfuric acid has caused a phenomenon of easy dissolution of copper seed layers by sulfuric acid contained in the electroplating solution.

15 Therefore, there have been problems that copper is not deposited, and defects such as voids and seams occur inevitably on the area lacking the seed layer due to this dissolution.

Heretofore, copper electroplating solutions and methods of electroplating, which enable the filling without copper
20 seed layer dissolution, have not existed, and the method for solving the problems has been sought.

DISCLOSURE OF THE INVENTION

The object of the present invention is to develop a copper
25 electroplating solution and a method of electroplating for filling fine via holes or trenches of a wiring (LSI) pattern formed on a semiconductor wafer, by adding a copper dissolution

inhibiting component in the electroplating solution, or by pretreatment using a solution that contains a copper dissolution inhibiting component so as to inhibit the dissolution of the copper seed layer of poor coverage and prevent producing defects such as voids on seams.

In view of the above-described problems, the present invention provides:

1. A copper electroplating solution characterized by containing an azole or silane coupling agent;
- 10 2. The copper electroplating solution according to the above-described 1, characterized by containing 1 to 10000 mg/L of the azole or silane coupling agent;
3. The copper electroplating solution according to the above-described 1, characterized by containing 10 to 5000 mg/L of the azole or silane coupling agent;
- 15 4. The copper electroplating solution according to any of the above-described 1 to 3 characterized by a copper sulfate electroplating solution containing copper sulfate, sulfuric acid, chlorine, and additives as major components;
- 20 5. A pretreatment solution for copper electroplating characterized by containing an azole or silane coupling agent;
6. The pretreatment solution for copper electroplating according to the above-described 5, characterized by containing 1 to 10000 mg/L of the azole or silane coupling agent; and
- 25 7. The pretreatment solution for copper electroplating according to the above-described 5, characterized by

containing 10 to 5000 mg/L of the azole or silane coupling agent.

The present invention also provides:

8. A method of copper electroplating characterized by
5 immersion in a copper sulfate electroplating solution containing the azole or silane coupling agent and, as major components, copper sulfate, sulfuric acid, chlorine, and additives for 1 to 60 seconds;
9. The method of copper electroplating according to the
10 above-described 8, characterized by immersion for 3 to 10 seconds;
10. The method of copper electroplating according to the above-described 8 or 9, characterized by containing 1 to 10000 mg/L of the azole or silane coupling agent;
- 15 11. The method of copper electroplating according to the above-described 8 or 9, characterized by containing 10 to 5000 mg/L of the azole or silane coupling agent;
12. A method of copper electroplating characterized in that the electroplating is carried out using a copper sulfate
20 electroplating solution containing copper sulfate, sulfuric acid, chlorine, and additives as major components after immersion in an aqueous solution containing an azole or silane coupling agent;
13. The method of copper electroplating according to the
25 above-described 12, characterized by immersion for 1 to 60 seconds in an aqueous solution containing an azole or silane coupling agent;

14. The method of copper electroplating according to the above-described 12, characterized by immersion for 3 to 10 seconds in an aqueous solution containing an azole or silane coupling agent;
- 5 15. The method of copper electroplating according to any of the above-described 12 to 14, characterized by containing 1 to 10000 mg/L of the azole or silane coupling agent; and
16. The method of copper electroplating according to any of the above-described 12 to 14, characterized by containing 10
10 to 5000 mg/L of the azole or silane coupling agent.

BEST MODE FOR CARRYING OUT THE INVENTION

A via hole or a trench for filling copper wiring is formed on the surface of a semiconductor (such as silicon) wafer,
15 and is covered with a barrier metal selected from a group consisting of titanium (Ti), tantalum (Ta), tungsten (W), the nitrides thereof, and the like, of a thickness of about 0.01 to 0.1 μm , for preventing the diffusion of copper (Cu) into silicon (Si) on the surface by a covering method such as vapor
20 deposition sputtering and CVD method or the like.

On the barrier metal layer is deposited a thin copper layer (seed layer) using a covering method such as vapor deposition, sputtering, and CVD, similarly to the above.

Since the above-described barrier metal has generally
25 a high electric resistance, and in subsequent copper electroplating to be covered, difference in current densities between the vicinity of the contact provided in the

circumferential portion of the wafer and the central portion becomes large, copper, which has a low electric resistance, is previously applied (thin coating).

The thickness of the copper layer is 0.01 to 0.1 μm .

5 However, the thickness of the copper layer may be determined optionally in semiconductor processing, and is not limited to the above-described range.

In general, the copper wiring (LSI etc.) is formed by using a process, in which after forming a wiring pattern on
10 a silicon wafer using the Damascene method, a barrier layer and a seed layer (a thin copper layer) are formed using a CVD method or a sputtering method, then the wiring pattern is filled using electroplating, and excessively deposited copper is removed using CMP (chemical-mechanical polishing).

15 The present invention uses a copper electroplating solution containing an azole or silane coupling agent, in place of conventional copper electroplating methods for filling.

The addition of an azole or silane coupling agent to the copper electroplating solution is a significant feature of
20 the present invention, and effectively inhibits the dissolution of the seed layer formed by sputtering.

As a result, the problems of the occurrence of defects such as voids and seams in conventional copper electroplating methods for filling can be solved at a stroke.

25 The azole added to the above-described electroplating solution is the general term for heterocyclic compounds with

five atoms having two or more different atoms, among which at least one is a nitrogen atom.

The representative example of the azoles include imidazole, thiazole, 1,2,3-triazole, 1,2,3-thiazole,
5 1,2,3,4-tetrazole, 1,2,3,4-thiatriazole, bendazole, benzimidazole, 1,2,3-benzotriazole, and 5-methyl-1-H-benzotriazole.

These azoles have an excellent effect to form complex compound with copper by the lone electron pair of nitrogen
10 atom to form a durable film on the surface of copper, which inhibits dissolution by an acid.

Also, the silane coupling agent is an organo-silicic compound (silane) having the action to chemically bond an organic material with an inorganic material (coupling), and
15 has an organic functional group (X) having affinity (or reactivity) with the organic material and a hydrolyzing group (OR) having affinity (or reactivity) with the inorganic material. The chemical structure thereof is represented by a general formula, XSi(OR)_3 .

20 The representative example of the silane coupling agents includes imidazolesilane as indicated by the following general formulas (1), (2), and (3).

25 (1)

(2)

5

(3)

10

(where R^1 denotes a hydrogen atom, a vinyl group, or an alkyl group having 1 to 5 carbon atoms; R^2 denotes a hydrogen atom, or an alkyl group having 1 to 20 carbon atoms; each of R^3 and R^4 denotes an alkyl group having 1 to 3 carbon atoms; and n denotes an integer from 1 to 3.);

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aminosilanes (e.g., γ -aminopropyltrimethoxysilane, γ -aminopropyltriethoxysilane,

N- β -(aminoethyl)- γ -aminopropyltrimethoxysilane, and N- β -(aminoethyl)- γ -aminopropyltriethoxysilane);

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epoxysilanes (e.g., γ -glycidoxypropyltrimethoxysilane, γ -glycidoxypropylmethyldiethoxysilane, and γ -glycidoxypropyltriethoxysilane); vinylsilanes (e.g., vinyltrichlorosilane, vinyltrimethoxysilane, and vinyltriethoxysilane); methacrylsilanes (e.g.,

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γ -methacryloxypropylmethyldimethoxy silane, γ -methacryloxypropyltrimethoxysilane, γ -methacryloxypropylmethyldiethoxysilane, and

γ -methacryloxypropyltriethoxysilane); and mercaptosilanes (e.g., γ -mercaptopropyltrimethoxysilane),

Alkoxysilyl (Si-OR) groups in these silane coupling agents are hydrolyzed by water or moisture to form silanol groups, which react with the surface of copper to form SiO-Cu bonds.

Furthermore, remaining silanol groups, which have not bonded with copper, are hydrolyzed to form siloxane bonds with each other, resulting in the formation of a film on the surface of copper. This film has an excellent effect to protect copper from dissolution due to an acid.

The quantity of the azole or silane coupling agent added to the copper electroplating solution (content in the electroplating solution) is preferably 1 to 10000 mg/L. This is because if the content is less than 1 mg/L, the effect of preventing the dissolution of the thin copper film due to addition is small; and if the content exceeds 10000 mg/L, since the effect of preventing the dissolution of the thin copper film due to addition is saturated, the excessive addition is wasteful, and tends to have an adverse effect on filling characteristics.

Therefore, the addition quantity is preferably 1 to 10000 mg/L. The preferable range for effectively preventing the dissolution of the thin copper film is the content between 10 and 5000 mg/L.

As the copper electroplating solution, a copper sulfate electroplating solution containing copper sulfate, sulfuric

acid, chlorine, and additives as major components is used. The addition of an azole or silane coupling agent according to the present invention is particularly effective for the copper sulfate electroplating solution.

5 The time of immersion in the electroplating solution according to the present invention is preferably 1 to 60 seconds. The more preferable range of the immersion time is 3 to 10 seconds. However, this immersion time can be changed depending on the design of filling, and it is not required
10 to limit in this time range.

 In the above description, although the azole or silane coupling agent is contained in the copper electroplating solution, the dissolution of the thin copper film due to an acid can be effectively inhibited similarly when
15 electroplating is carried out using the above-described copper sulfate electroplating solution, after immersion in a pretreatment solution consisting of an aqueous solution containing the azole or silane coupling agent.

 In this case, it is preferable that the content of the
20 azole or silane coupling agent in the pretreatment solution (in the aqueous solution) is similarly 1 to 10000 mg/L. This is because if the content is less than 1 mg/L, the effect of preventing the dissolution of the thin copper film due to addition is small; and if the content exceeds 10000 mg/L, since
25 the effect of preventing the dissolution of the thin copper film due to addition is saturated, the excessive addition is wasteful. The more preferable content range is 10 to 5000

mg/L, similarly in the case of addition in the electroplating solution.

The time of immersion in the pretreatment solution according to the present invention is preferably 1 to 60 seconds.

5 If it is less than 1 second, the effect is small; and if it exceeds 60 seconds, the effect is no longer changed. The more preferable range of the immersion time is 3 to 10 seconds.

The method of copper electroplating is carried out as described below, although it is not limited to a particular
10 method. For example, a semiconductor wafer, which is a material to be electroplated, is disposed in an electroplating bath so as to face the anode. The back surface of the semiconductor wafer is sealed so as not to make contact with the electroplating solution leaving the top surface to be
15 electroplated, and the contact for feeding electricity is disposed in the vicinity of the edge of the semiconductor wafer.

As the anode, a phosphorus-containing copper anode (P content: 0.04 - 0.06%), or an insoluble anode is used. As the insoluble anode, the use of Pt or Pt-electroplated Ti is
20 preferable. A commercially available dimension stable anode (DSA) can also be used. When a phosphorus-containing copper anode is used, the copper consumed for electroplating is automatically replenished by the dissolution of the anode.

When an insoluble anode is used, since the copper content
25 in the solution is decreased due to electroplating, a copper sulfate solution is supplied to maintain the copper content.

A typical composition of a copper electroplating solution (excluding the above-described azole or silane coupling agent) is as follows:

- Copper sulfate (as copper): 0.1 to 100 g/L (preferably 10 to 50 g/L)
- Sulfuric acid: 0.1 to 500 g/L (preferably 10 to 300 g/L)
- Chlorine: 0.1 to 500 mg/L (preferably 30 to 100 mg/L)
- Organic sulfur compound: 1 to 500 $\mu\text{mol/L}$ (preferably 4 to 200 $\mu\text{mol/L}$)
- Polyether compound: 0.1 to 5000 $\mu\text{mol/L}$ (preferably 0.5 to 500 $\mu\text{mol/L}$)
- Balance: water

Furthermore, a quaternary ammonium salt addition compound consisting of a tertiary alkyl amine and polyepichlorohydrine, and polyalkylethyleneimine or the like is added within the above-described range as required.

An example of electroplating conditions is as follows:

- Current density: 0.1 to 100 A/dm^2 (preferably 0.5 to 5 A/dm^2)
- Solution temperature: 10 to 80°C (preferably 15 to 30°C)

The current density, solution temperature, and the flow rate of the solution (relative rate between the surface to be electroplated and the solution bulk) in electroplating has dependent relationship on each other, and by imparting a proper flow rate to the solution, a desired deposition rate and copper deposition (crystalline state) can be obtained.

The methods for imparting the flow rate to the solution include a method to swing or rotate the wafer to be electroplated, or a method to agitate the ambient air.

Also, as the current to be supplied in electroplating
5 not only a direct current, but also a pulse current or PR (periodic reverse) current can be used.

When the pulse current is used, after a current flows to deposit copper for a certain time (on time), the current stops flowing for a certain time (off time), to release the
10 copper-ion depleted state in the vicinity of the electrode, where deposition takes place, thereby enabling current density during on time to be higher than ordinary direct current.

On the other hand, when the PR current is used, copper deposited in a certain time is dissolved by supplying a reverse
15 current for a certain time. Thereby, the deposition of copper on the portions where the current concentrates easily, such as the corner of the trench, can be inhibited. This enables the deposits that cannot be obtained by a direct current to be obtained.

20 The copper electroplating solution may contain an organic sulfur compound. The organic sulfur compounds include, for example, bis-sulfo disulfide disodium salt, bis-(1-sulfomethyl) disulfide disodium salt, bis-(2-sulfoethyl) disulfide disodium salt,
25 bis-(3-sulfopropyl) disulfide disodium salt, bis-(4-sulfobutyl) disulfide disodium salt, tetramethylthiuram disulfide, tetraethylthiuram disulfide,

sodium 3-mercapto-1-propanesulfonate, and sodium 2,3-dimercapto-1-propanesulfonate.

These are evenly adsorbed on the entire surface without depending on the place of the surface to be electroplated in the initial stage of electroplating, and accelerate the reaction. However, as the reaction proceeds, the adsorption density in the depressed portion in the via hole or trench elevates, resulting in the increase of the deposition rate at this portion.

10 In contrast, the adsorption density in the protruded portion in the vicinity of the entrance of the via hole or trench is lowered, resulting in the decrease of the deposition rate at this portion. Thus, the copper electroplating solution containing an organic sulfur compound has a significant effect to make the thickness of the deposited film even. The quantity of the organic sulfur compound is preferably 1 to 500 $\mu\text{mol/L}$.

Also, a polyether compound can be added to the copper electroplating solution. Polyether compounds that can be used include, for example, polyethyleneglycol (molecular weight: 100 to 50000) and polypropyleneglycol (molecular weight: 100 to 50000).

20 These are preferentially adsorbed on the surface of the wafer with a via hole or a trench (places other than the via hole or trench) at the time of electroplating, and inhibits the deposition of copper on the surface portion. The quantity of the polyether compound is preferably 0.1 to 5000 $\mu\text{mol/L}$.

Also as described above, a quaternary ammonium salt addition compound consisting of a tertiary alkyl amine and polyepichlorohydrine can be added as required. This substance effectively inhibits the deposition of copper on the surface portion of the wafer in the same manner as the polyether compound. The quantity of the quaternary ammonium salt addition compound is preferably 0.1 to 200 $\mu\text{mol/L}$.

Also, polyalkylethylenimine can be added to the copper electroplating solution. This substance inhibits the deposition of copper on the surface portion of the wafer in the same manner as the polyether compound. The quantity of the polyalkylethylenimine is preferably 0.1 to 7 $\mu\text{mol/L}$.

Although the above-described quaternary ammonium salt addition compound consisting of a tertiary alkyl amine and polyepichlorohydrine, and polyalkylethylenimine are not necessarily required to be contained as the components of the electroplating solution when the content of the organic sulfur compound and the polyether compound is within the most optimal range, the effect on the filling characteristics of the fine via hole or trench is more stabilized when contained within the above-described range.

By adding the combination of the above-described additives within the above-described ranges of content of each component to the copper electroplating solution, the deposition rate of copper can be changed from place to place when copper electroplating is carried out on a semiconductor wafer with a fine via hole or trench.

That is, the deposition rate on the depressed portion in the via hole or trench can be increased, and the deposition rate on the protruded portion in the vicinity of the entrance of the via hole or trench, and on the surface of the wafer
5 can be decreased, thereby improving the filling characteristics in the via hole or trench.

The film thickness of copper electroplating may be a thickness for filling the via hole or trench on the surface of the semiconductor wafer, and for forming wiring by
10 planarization using subsequent chemical-mechanical polishing (CMP). In general, the thickness is 0.5 to 2 μm , but the thickness may be optionally selected without being limited to the above-described range.

15 Examples and Comparative Examples

Next, the examples and the comparative examples of the present invention will be described. These examples illustrate only preferable examples, but the present invention is not limited to these examples. The present invention
20 inevitably includes modifications and other examples within the range of the technical concept thereof.

Electroplating was carried out using electroplating solutions shown in Examples 1 to 7. Also, after immersion treatment by pretreatment solutions shown in Examples 8 and
25 9, electroplating was carried out using electroplating solutions.

Also electroplating was carried out using electroplating solutions shown in Comparative Examples 1 to 3. As the material to be electroplated, a silicon wafer with a fine via pattern on which TaN (30 nm)/Cu (100 nm) is sputtered was used.

- 5 Electroplating of about 1 μm was carried out at a solution temperature of 20°C, and a cathode current density of 1A/dm². The via pattern had a depth of 1 μm , and a hole diameter of 0.18 μm .

(Example 1)

- 10 Copper 15 g/L, sulfuric acid 180 g/L, chlorine 70 mg/L, polyethyleneglycol (molecular weight: 3350) 50 $\mu\text{mol/L}$, bis-(3-sulfopropyl) disulfide disodium salt 30 $\mu\text{mol/L}$, quaternary epichlorohydrine (a quaternary ammonium salt addition compound consisting of a tertiary alkyl amine and
15 polyepichlorohydrine) 6 $\mu\text{mol/L}$, polybenzylethylenimine 2 $\mu\text{mol/L}$, an aqueous solution of imidazolesilane (a mixed aqueous solution of 20 g/L of a silane coupling agent obtained by the equimolar reaction of imidazole and
20 γ -glycidoxypyriltrimethoxysilane or the like, and 20 mL/L of methanol) 50 mL/L

(Example 2)

- Copper 20 g/L, sulfuric acid 200 g/L, chlorine 50 mg/L, polyethyleneglycol (molecular weight: 10000) 10 $\mu\text{mol/L}$, sodium 3-mercapto-1-propanesulfonate 30 $\mu\text{mol/L}$, quaternary
25 epichlorohydrine 30 $\mu\text{mol/L}$, γ -aminopropyltrimethoxysilane 2 g/L

(Example 3)

Copper 30 g/L, sulfuric acid 180 g/L, chlorine 100 mg/L,
polyethyleneglycol (molecular weight: 15000) 10 $\mu\text{mol/L}$,
bis-(2-sulfoethyl) disulfide disodium salt 20 $\mu\text{mol/L}$,
5 quaternary epichlorohydrine 15 $\mu\text{mol/L}$,
polyallylethyleneimine 1 $\mu\text{mol/L}$,
 γ -glycidoxypolytriethoxysilane 500 mg/L

(Example 4)

Copper 10 g/L, sulfuric acid 240 g/L, chlorine 50 mg/L,
10 polyethyleneglycol (molecular weight: 1000) 250 $\mu\text{mol/L}$,
tetramethylthiuram disulfide 60 $\mu\text{mol/L}$, quaternary
epichlorohydrine 45 $\mu\text{mol/L}$,
 N - β -(aminoethyl)- γ -aminopropyltrimethoxysilane 1 g/L

(Example 5)

15 Copper 25 g/L, sulfuric acid 200 g/L, chlorine 70 mg/L,
polypropyleneglycol (molecular weight: 8000) 15 $\mu\text{mol/L}$,
sodium 2,3-dimercapto-1-propanesulfonate 30 $\mu\text{mol/L}$,
quaternary epichlorohydrine 25 $\mu\text{mol/L}$,
polybenzylethyleneimine 2 $\mu\text{mol/L}$,
20 γ -mercaptopropyltrimethoxysilane 2 g/L

(Example 6)

Copper 15 g/L, sulfuric acid 300 g/L, chlorine 50 mg/L,
polypropyleneglycol (molecular weight: 20000) 2.5 $\mu\text{mol/L}$,
bis-(3-sulfopropyl) disulfide disodium salt 45 $\mu\text{mol/L}$,
25 quaternary epichlorohydrine 60 $\mu\text{mol/L}$,
polyallylethyleneimine 1 $\mu\text{mol/L}$, 5-methyl-1H-benzotriazole
200 mg/L

(Example 7)

Copper 25 g/L, sulfuric acid 200 g/L, chlorine 60 mg/L,
polyethyleneglycol (molecular weight: 5000) 40 μ mol/L,
bis-(2-sulfoethyl) disulfide disodium salt 15 μ mol/L,
5 quaternary epichlorohydrine 20 μ mol/L, 1,2,3-benzotriazole
100 mg/L

(Example 8)

Pretreatment solution: An aqueous solution of imidazole
(a mixed aqueous solution of 20 g/L of a silane coupling agent
10 obtained by the equimolar reaction of imidazolesilane and
 γ -glycidoxypyrpyltrimethoxysilane or the like, and 20 mL/L
of methanol) 100 mL/L (10 sec immersion)

Electroplating solution: Copper 20 g/L, sulfuric acid
300 g/L, chlorine 80 mg/L, polypropyleneglycol (molecular
15 weight: 10000) 5 μ mol/L, bis-(3-sulfopropyl) disulfide
disodium salt 30 μ mol/L, quaternary epichlorohydrine 25 μ mol/L,
polyallylethyleneimine 1 μ mol/L

(Example 9)

Pretreatment solution: 1,2,3-benzotriazole 200 mg/L (10
20 sec immersion)

Electroplating solution: Copper 25 g/L, sulfuric acid
200 g/L, chlorine 100 mg/L, polyethyleneglycol (molecular
weight: 5000) 30 μ mol/L, bis-(2-sulfoethyl) disulfide
disodium salt 15 μ mol/L, quaternary epichlorohydrine 15 μ mol/L

25 (Comparative Example 1)

Copper 25 g/L, sulfuric acid 150 g/L, chlorine 50 mg/L,
polyethyleneglycol (molecular weight: 10000) 10 μ mol/L,

bis-(3-sulfopropyl) disulfide disodium salt 60 $\mu\text{mol/L}$,
quaternary epichlorohydrine 15 $\mu\text{mol/L}$,
polybenzylethyleneimine 2 $\mu\text{mol/L}$

(Comparative Example 2)

5 Copper 10 g/L, sulfuric acid 300 g/L, chlorine 100 mg/L,
polyethyleneglycol (molecular weight: 20000) 2.5 $\mu\text{mol/L}$,
sodium 3-mercapto-1-propanesulfonate 60 $\mu\text{mol/L}$, quaternary
epichlorohydrine 30 $\mu\text{mol/L}$

(Comparative Example 3)

10 Copper 30 g/L, sulfuric acid 200 g/L, chlorine 80 mg/L,
polyethyleneglycol (molecular weight: 5000) 30 $\mu\text{mol/L}$,
tetramethylthiuram disulfide 20 $\mu\text{mol/L}$, quaternary
epichlorohydrine 60 $\mu\text{mol/L}$, polyallylethyleneimine 1 $\mu\text{mol/L}$

The filling characteristics of obtained deposits into
15 fine via patterns were checked by cleavage cross-sectional
SEM observation, and results as shown in Table 1 were obtained.

As Table 1 shows, no voids and seams were produced in
any of Examples 1 to 9. However, voids and seams were produced
in Comparative Examples 1 to 3 treated with electroplating
20 solutions containing no azole or silane coupling agent of the
present invention.

Although the representative Examples and Comparative
Examples are shown above, the effect was confirmed in all the
compositions according to the claims of the present invention.

Table 1

	Occurrence of voids or seams
Example 1	No voids and seams
Example 2	No voids and seams
Example 3	No voids and seams
Example 4	No voids and seams
Example 5	No voids and seams
Example 6	No voids and seams
Example 7	No voids and seams
Example 8	No voids and seams
Example 9	No voids and seams
Comparative Example 1	Voids and seams occurred
Comparative Example 2	Voids and seams occurred
Comparative Example 3	Voids and seams occurred

INDUSTRIAL APPLICABILITY

The use of the copper electroplating solution and the pretreatment solution of the present invention has an excellent effect to remove the phenomenon of dissolution of a copper seed layer in sulfuric acid contained in the electroplating solution, when the formation of a copper film in the hole or on the bottom of a fine via or a trench was insufficient in the past, and to enable the copper to be filled without producing defects such as voids and seams in a fine via holes or trench formed on a semiconductor wafer.



CLAIMS

1. A copper electroplating solution, characterized by containing an azole or silane coupling agent.
- 5 2. The copper electroplating solution according to Claim 1, characterized by containing 1 to 10000 mg/L of the azole or silane coupling agent.
3. The copper electroplating solution according to Claim 1, characterized by containing 10 to 5000 mg/L of the azole
- 10 or silane coupling agent.
4. The copper electroplating solution according to any of Claims 1 to 3, characterized by a copper sulfate electroplating solution containing copper sulfate, sulfuric acid, chlorine, and additives as major components.
- 15 5. A pretreatment solution for copper electroplating, characterized by containing an azole or silane coupling agent.
6. The pretreatment solution for copper electroplating according to Claim 5, characterized by containing 1 to 10000 mg/L of the azole or silane coupling agent.
- 20 7. The pretreatment solution for copper electroplating according to Claim 5, characterized by containing 10 to 5000 mg/L of the azole or silane coupling agent.
8. A method of copper electroplating, characterized by immersion in a copper sulfate electroplating solution
- 25 containing the azole or silane coupling agent, and, as major components, copper sulfate, sulfuric acid, chlorine, and additives for 1 to 60 seconds.

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REPLACED



9. The method of copper electroplating according to Claim 8, characterized by immersion for 3 to 10 seconds.
10. The method of copper electroplating according to Claims 8 or 9, characterized by containing 1 to 10000 mg/L of the
5 azole or silane coupling agent.
11. The method of copper electroplating according to Claims 8 or 9, characterized by containing 10 to 5000 mg/L of the azole or silane coupling agent.
12. A method of copper electroplating, characterized in that
10 the electroplating is carried out using a copper sulfate electroplating solution containing copper sulfate, sulfuric acid, chlorine, and additives as major components after immersion in an aqueous solution containing an azole or silane coupling agent.
13. The method of copper electroplating according to Claim 12, characterized by immersion for 1 to 60 seconds in an aqueous
15 solution containing an azole or silane coupling agent.
14. The method of copper electroplating according to Claim 12, characterized by immersion for 3 to 10 seconds in an aqueous
20 solution containing an azole or silane coupling agent.
15. The method of copper electroplating according to any of Claims 12 to 14, characterized by containing 1 to 10000 mg/L of the azole or silane coupling agent.
16. The method of copper electroplating according to any of
25 Claims 12 to 14, characterized by containing 10 to 5000 mg/L of the azole or silane coupling agent.