

Glass for Memory Recording Media

CROSS-REFERENCE TO RELATED APPLICATION

[0001] This application claims the benefit of priority under 35 U.S.C. § 119 of U.S. Provisional Application Serial No. 63/308130 filed on February 9, 2022 and Netherlands Application No. 2030965 filed on February 16, 2022, the contents of each of which are relied upon and incorporated herein by reference in their entireties.

FIELD

[0002] The present specification generally relates to glass compositions suitable for use as substrate material for electronic devices. More specifically, the present specification is directed to substrate materials for memory recording disks suitable for the Heat Assisted Magnetic Recording (HAMR) process.

BACKGROUND

[0003] The industry is continuously requiring the ability to store more and more data. In order to accommodate this need for increased data storage, the industry is having to make shifts in the way they store data on disks. In order to get around physics barriers, they have moved to a new technology called Heat Assisted Magnetic Recording (HAMR). With this technology, data storage is accomplished by heating up small areas of the magnetic memory material to write the data and rapidly cooling the magnetic memory material to store the data in a more stable phase. The substrate used to support the magnetic memory material in current magnetic disk technology is aluminum. With the transition to HAMR technology, however, aluminum will no longer be a viable substrate because it does not have the ability to function at the temperatures, such as about 700°C or greater, necessary for writing data with the HAMR technology. Therefore, a need exists for new substrate materials, such as glass, that are resistant to high temperatures, have higher rigidity and smooth, flat surfaces, and high resistance to fast cooling and heating (or thermal shock resistance), which, in turn, requires low coefficient of thermal expansion.

SUMMARY

[0004] According to an embodiment of the present disclosure, a glass comprises a plurality of components, the glass having a composition of the components comprising greater than or equal to 60.0 mol.% and less than or equal to 75.0 mol.% SiO₂, greater than or equal to 10.0 mol.% and less than or equal to 18.0 mol.% Al₂O₃, greater than or equal to 0.0 mol.% and less than or equal to 10.0 mol.% Li₂O, greater than or equal to 0.0 mol.% and less than or equal to 5.0 mol.% Na₂O, greater than or equal to 0.0 mol.% and less than or equal to 3.8 mol.% B₂O₃, greater than or equal to 0.0 mol.% and less than or equal to 3.0 mol.% P₂O₅, greater than or equal to 0.0 mol.% and less than or

equal to 2.0 mol.% ZnO, greater than or equal to 0.0 mol.% and less than or equal to 0.7 mol.% La₂O₃, greater than or equal to 0.0 at.% and less than or equal to 3.0 at.% F, a sum of CaO + MgO greater than or equal to 5.0 mol.%, a sum of Li₂O + Na₂O greater than or equal to 0.5 mol.%, a sum of Li₂O + MgO greater than or equal to 0.0 mol.% and less than or equal to 10.0 mol.%, a sum of MgO + ZnO greater than or equal to 0.0 mol.% and less than or equal to 10.0 mol.%, a sum of CaO + SrO greater than or equal to 0.0 mol.% and less than or equal to 9.0 mol.%, and a sum of ZrO₂ + TiO₂ + FeO + Fe₂O₃ greater than or equal to 0.0 mol.% and less than or equal to 1.5 mol.%, wherein the glass has an aluminum-binding parameter P_{Al} that is greater than or equal to -2.8, a modifier-binding parameter P_{mod} that is less than or equal to 2.8 and an anorthite precipitation parameter P_{anort} that is less than or equal to 10, where P_{Al} is calculated from the glass composition in terms of mol.% of the components according to the Formula (VIII):

$$P_{Al} = R_2O + RO + P_2O_5 + 1.6 * RE_mO_n - Al_2O_3, \quad (VIII)$$

P_{mod} is calculated from the glass composition in terms of mol.% of the components according to the Formula (VII):

$$P_{mod} = R_2O + RO - Al_2O_3 - P_2O_5 - RE_mO_n, \quad (VII)$$

P_{anort} is calculated from the glass composition in terms of mol.% of the components according to the Formula (I):

$$P_{anort} = \min(CaO + SrO + 0.5 * BaO + Na_2O + 0.5 * K_2O, Al_2O_3), \quad (I)$$

where R₂O is a total sum of monovalent metal oxides, RO is a total sum of divalent metal oxides, RE_mO_n is a total sum of rare earth metal oxides in all redox states present, and an asterisk (*) means multiplication.

[0005] According to another embodiment of the present disclosure, a glass comprises a plurality of components, the glass having a composition of the components comprising greater than or equal to 60.0 mol.% and less than or equal to 80.0 mol.% SiO₂, greater than or equal to 10.0 mol.% and less than or equal to 18.0 mol.% Al₂O₃, greater than or equal to 0.5 mol.% and less than or equal to 7.5 mol.% Li₂O, greater than or equal to 0.0 mol.% and less than or equal to 0.5 mol.% ZrO₂, greater than or equal to 0.0 mol.% and less than or equal to 3.0 mol.% RE_mO_n, greater than or equal to 0.0 at.% and less than or equal to 0.5 at.% F and may optionally contain one or more components selected from P₂O₅, B₂O₃, MgO, CaO, BaO, ZnO, MnO, Na₂O, K₂O, Fe₂O₃, FeO, Cu₂O, Rb₂O, Ag₂O, Cs₂O, Au₂O, Hg₂O, Tl₂O, BeO, CoO, NiO, CuO, SrO, CdO, SnO, PbO and TiO₂, wherein the composition of the components satisfies the condition: $0.00 \leq \min(RE_mO_n, P_2O_5) [\text{mol.\%}] \leq 0.30$, and wherein the glass has a cristobalite precipitation parameter P_{crist} that is less than or equal to 28, an anorthite

precipitation parameter P_{anort} that is less than or equal to 10, a cordierite precipitation parameter P_{cord} that is less than or equal to 5.0 and a spodumene precipitation parameter P_{spod} that is less than or equal to 7.5, and wherein the glass satisfies the conditions: $P_{spm} > 32$, $P_{anpt} > 680$ and $P_{160kP} > 1150$, where P_{anpt} is an annealing point parameter, calculated from the glass composition in terms of mol.% of the components according to the Formula (IX):

$$P_{anpt} = 664.7 + 5.2303 * SiO_2 - 11.493 * B_2O_3 - 7.1742 * P_2O_5 + 8.3980 * ZrO_2 - 2.0585 * MgO - 2.1088 * CaO - 3.8995 * BaO - 10.323 * ZnO - 9.0727 * MnO - 23.455 * Li_2O - 33.819 * Na_2O - 25.204 * K_2O + 15.745 * Y_2O_3 + 8.9047 * La_2O_3 - 33.960 * (Fe_2O_3 + FeO) - 5.6704 * (R_2O + RO - Al_2O_3) - 4.2545 * (SiO_2 - (6 * K_2O + 6 * Na_2O + 4 * Li_2O + 2 * RO)) - 19.439 * Cu_2O, \quad (IX)$$

P_{spm} is a specific modulus parameter, calculated from the glass composition in terms of mol.% of the components according to the Formula (X):

$$P_{spm} = 32.10 + 0.47744 * SiO_2 - 1.6506 * Al_2O_3 - 0.11775 * B_2O_3 - 0.30166 * P_2O_5 + 0.20187 * TiO_2 + 0.20219 * ZrO_2 + 0.96268 * MgO + 0.83379 * CaO + 0.53685 * SrO + 0.41218 * BaO + 0.63264 * ZnO + 0.75365 * MnO + 0.62984 * CuO - 1.2496 * Na_2O - 1.5154 * K_2O + 1.4746 * Cu_2O - 0.037941 * Y_2O_3 - 0.75836 * La_2O_3 - 1.8052 * (R_2O + RO - Al_2O_3) - 0.47488 * (SiO_2 - (6 * K_2O + 6 * Na_2O + 4 * Li_2O + 2 * RO)), \quad (X)$$

P_{160kP} is a parameter predicting a temperature at which a viscosity of the glass is 160 kP, calculated from the glass composition in terms of mol.% of the components according to the Formula (XI):

$$P_{160kP} = 1058 + 2.5492 * SiO_2 - 25.725 * Al_2O_3 - 11.327 * B_2O_3 - 10.014 * P_2O_5 - 14.309 * TiO_2 - 11.594 * ZrO_2 + 30.559 * MgO + 29.290 * CaO + 30.592 * SrO + 30.079 * BaO + 23.323 * ZnO + 19.724 * MnO + 11.888 * PbO + 11.462 * CuO + 17.090 * Li_2O + 16.475 * Na_2O + 11.386 * K_2O + 14.422 * Y_2O_3 - 36.909 * La_2O_3 - 34.144 * (Fe_2O_3 + FeO) - 35.001 * (R_2O + RO - Al_2O_3), \quad (XI)$$

P_{crist} is calculated from the glass composition in terms of mol.% of the components according to the Formula (VI):

$$P_{crist} = SiO_2 - 6 * (Na_2O + K_2O) - 4 * Li_2O - 2 * (CaO + SrO + BaO) - 2.5 * MgO, \quad (VI)$$

P_{anort} is calculated from the glass composition in terms of mol.% of the components according to the Formula (I):

$$P_{anort} = \min(CaO + SrO + 0.5 * BaO + Na_2O + 0.5 * K_2O, Al_2O_3), \quad (I)$$

P_{cord} is calculated from the glass composition in terms of mol.% of the components according to the Formula (III):

$$P_{cord} = MgO + MnO + FeO, \quad (III)$$

P_{spod} is calculated from the glass composition in terms of mol.% of the components according to the Formula (IV):

$$P_{\text{spod}} = \min(\text{Li}_2\text{O}, \text{Al}_2\text{O}_3 - \text{K}_2\text{O} - 0.5 * \text{Na}_2\text{O}), \quad (\text{IV})$$

where $R_2\text{O}$ is a total sum of monovalent metal oxides, RO is a total sum of divalent metal oxides, RE_mO_n is a total sum of rare earth metal oxides in all redox states present, and an asterisk (*) means multiplication.

[0006] According to one more embodiment of the present disclosure, a glass comprises a plurality of components, the glass having a composition of the components comprising greater than or equal to 60.0 mol.% and less than or equal to 75.0 mol.% SiO_2 , greater than or equal to 10.5 mol.% and less than or equal to 18.0 mol.% Al_2O_3 , greater than or equal to 0.0 mol.% and less than or equal to 10.0 mol.% CaO , greater than or equal to 0.0 mol.% and less than or equal to 7.8 mol.% Li_2O , greater than or equal to 0.0 mol.% and less than or equal to 4.3 mol.% MgO , greater than or equal to 0.0 mol.% and less than or equal to 0.5 mol.% ZrO_2 , greater than or equal to 0.0 mol.% and less than or equal to 5.0 mol.% RE_mO_n and may optionally contain one or more components selected from B_2O_3 , P_2O_5 , TiO_2 , SrO , BaO , ZnO , MnO , CuO , Na_2O , K_2O , Cu_2O , Rb_2O , Ag_2O , Cs_2O , Au_2O , Hg_2O , Tl_2O , BeO , FeO , CoO , NiO , CdO , SnO , PbO and Fe_2O_3 , wherein the glass has a cristobalite precipitation parameter P_{crist} that is less than or equal to 28, an anorthite precipitation parameter P_{anort} that is less than or equal to 10 and a modifier-binding parameter P_{mod} that is greater than or equal to -3.0 and the glass satisfies the condition: $P_{\text{spm}} - (92.5 - 0.05 * P_{160\text{kP}}) > 0.000$, where P_{spm} is a specific modulus parameter, calculated from the glass composition in terms of mol.% of the components according to the Formula (X):

$$\begin{aligned} P_{\text{spm}} = & 32.10 + 0.47744 * \text{SiO}_2 - 1.6506 * \text{Al}_2\text{O}_3 - 0.11775 * \text{B}_2\text{O}_3 - 0.30166 * \text{P}_2\text{O}_5 + 0.20187 * \text{TiO}_2 \\ & + 0.20219 * \text{ZrO}_2 + 0.96268 * \text{MgO} + 0.83379 * \text{CaO} + 0.53685 * \text{SrO} + 0.41218 * \text{BaO} + 0.63264 \\ & * \text{ZnO} + 0.75365 * \text{MnO} + 0.62984 * \text{CuO} - 1.2496 * \text{Na}_2\text{O} - 1.5154 * \text{K}_2\text{O} + 1.4746 * \text{Cu}_2\text{O} - \quad (\text{X}) \\ & 0.037941 * \text{Y}_2\text{O}_3 - 0.75836 * \text{La}_2\text{O}_3 - 1.8052 * (\text{R}_2\text{O} + \text{RO} - \text{Al}_2\text{O}_3) - 0.47488 * (\text{SiO}_2 - (6 * \text{K}_2\text{O} + 6 * \\ & \text{Na}_2\text{O} + 4 * \text{Li}_2\text{O} + 2 * \text{RO})), \end{aligned}$$

$P_{160\text{kP}}$ is a parameter predicting a temperature at which the glass has a viscosity of 160 kP, calculated from the glass composition in terms of mol.% of the components according to the Formula (XI):

$$\begin{aligned} P_{160\text{kP}} = & 1058 + 2.5492 * \text{SiO}_2 - 25.725 * \text{Al}_2\text{O}_3 - 11.327 * \text{B}_2\text{O}_3 - 10.014 * \text{P}_2\text{O}_5 - 14.309 * \text{TiO}_2 - \\ & 11.594 * \text{ZrO}_2 + 30.559 * \text{MgO} + 29.290 * \text{CaO} + 30.592 * \text{SrO} + 30.079 * \text{BaO} + 23.323 * \text{ZnO} + \quad (\text{XI}) \\ & 19.724 * \text{MnO} + 11.888 * \text{PbO} + 11.462 * \text{CuO} + 17.090 * \text{Li}_2\text{O} + 16.475 * \text{Na}_2\text{O} + 11.386 * \text{K}_2\text{O} + \\ & 14.422 * \text{Y}_2\text{O}_3 - 36.909 * \text{La}_2\text{O}_3 - 34.144 * (\text{Fe}_2\text{O}_3 + \text{FeO}) - 35.001 * (\text{R}_2\text{O} + \text{RO} - \text{Al}_2\text{O}_3), \end{aligned}$$

P_{crist} is a value of a cristobalite precipitation parameter, calculated from the glass composition in terms of mol.% of the components according to the Formula (VI):

$$P_{\text{crist}} = \text{SiO}_2 - 6 * (\text{Na}_2\text{O} + \text{K}_2\text{O}) - 4 * \text{Li}_2\text{O} - 2 * (\text{CaO} + \text{SrO} + \text{BaO}) - 2.5 * \text{MgO}, \quad (\text{VI})$$

P_{anort} is a value of an anorthite precipitation parameter, calculated from the glass composition in terms of mol.% of the components according to the Formula (I):

$$P_{anort} = \min(\text{CaO} + \text{SrO} + 0.5 * \text{BaO} + \text{Na}_2\text{O} + 0.5 * \text{K}_2\text{O}, \text{Al}_2\text{O}_3), \quad (\text{I})$$

P_{mod} is a value of a modifier-binding parameter, calculated from the glass composition in terms of mol.% of the components according to the Formula (VII):

$$P_{mod} = \text{R}_2\text{O} + \text{RO} - \text{Al}_2\text{O}_3 - \text{P}_2\text{O}_5 - \text{RE}_m\text{O}_n, \quad (\text{VII})$$

where R_2O is a total sum of monovalent metal oxides, RO is a total sum of divalent metal oxides, RE_mO_n is a total sum of rare earth metal oxides in all redox states present, and an asterisk (*) means multiplication.

[0007] These and other aspects, objects, and features of the present disclosure will be understood and appreciated by those skilled in the art upon studying the following specification, claims, and appended drawings.

BRIEF DESCRIPTION OF THE DRAWINGS

[0008] FIG. 1 is a plot illustrating the relationship between the annealing point An.P. and the annealing point parameter P_{anpt} calculated by formula (IX) for some Comparative Glasses and some Exemplary Glasses according to an embodiment of the present disclosure.

[0009] FIG. 2 is a plot illustrating the relationship between the specific modulus E/d_{RT} and the specific modulus parameter P_{spm} calculated by formula (X) for some Comparative Glasses and some Exemplary Glasses according to an embodiment of the present disclosure.

[0010] FIG. 3 is a plot illustrating the relationship between T_{160kP} , the temperature at which the glass has a viscosity of 160 kP, and a parameter P_{160kP} calculated by formula (XI) that predicts the temperature at which the glass has a viscosity of 160 kP for some Comparative Glasses and some Exemplary Glasses according to an embodiment of the present disclosure.

[0011] FIG. 4 is a plot illustrating the relationship between the parameter P_{160kP} and the specific modulus parameter P_{spm} for some Comparative Glasses and some Exemplary Glasses according to an embodiment of the present disclosure.

[0012] FIG. 5 is a plot illustrating the relationship between the temperature T_{160kP} and the specific modulus E/d_{RT} for some Comparative Glasses and some Exemplary Glasses according to an embodiment of the present disclosure.

DETAILED DESCRIPTION

[0013] In the following detailed description, for purposes of explanation and not limitation, example embodiments disclosing specific details are set forth to provide a thorough understanding

of various principles of the present disclosure. However, it will be apparent to one having ordinary skill in the art, having had the benefit of the present disclosure, that the present disclosure may be practiced in other embodiments that depart from the specific details disclosed herein. Moreover, descriptions of well-known devices, methods and materials may be omitted so as not to obscure the description of various principles of the present disclosure. Finally, wherever applicable, like reference numerals refer to like elements.

[0014] Unless otherwise expressly stated, it is in no way intended that any method set forth herein be construed as requiring that its steps be performed in a specific order. Accordingly, where a method claim does not actually recite an order to be followed by its steps, or it is not otherwise specifically stated in the claims or descriptions that the steps are to be limited to a specific order, it is in no way intended that an order be inferred, in any respect. This holds for any possible non-express basis for interpretation, including, without limitation, matters of logic with respect to arrangement of steps or operational flow; plain meaning derived from grammatical organization or punctuation; the number or type of embodiments described in the specification.

[0015] As used herein, the term "and/or," when used in a list of two or more items, means that any one of the listed items can be employed by itself, or any combination of two or more of the listed items can be employed. For example, if a composition is described as containing components A, B, and/or C, the composition can contain A alone; B alone; C alone; A and B in combination; A and C in combination; B and C in combination; or A, B, and C in combination.

[0016] Modifications of the disclosure will occur to those skilled in the art and to those who make or use the disclosure. Therefore, it is understood that the embodiments shown in the drawings and described above are merely for illustrative purposes and not intended to limit the scope of the disclosure, which is defined by the following claims, as interpreted according to the principles of patent law, including the doctrine of equivalents.

[0017] As used herein, the term "about" means that amounts, sizes, formulations, parameters, and other quantities and characteristics are not and need not be exact, but may be approximate and/or larger or smaller, as desired, reflecting tolerances, conversion factors, rounding off, measurement error and the like, and other factors known to those skilled in the art. When the term "about" is used in describing a value or an end-point of a range, the disclosure should be understood to include the specific value or end-point referred to. Whether or not a numerical value or end-point of a range in the specification recites "about," the numerical value or end-point of a range is intended to include two embodiments: one modified by "about," and one not modified by "about." It will be further understood that the end-points of each of the ranges are significant both in relation to the other end-point, and independently of the other end-point.

[0018] The term “component” refers to a material or compound present in a glass composition. Components include oxides, including but not limited to those expressed in Formulas (I), and (II), and the claims. Representative components include B_2O_3 , P_2O_5 , Al_2O_3 , CuO , Cu_2O , RO , R_2O , SnO_2 , MnO_2 , RE_mO_n , SiO_2 , Ta_2O_5 , ZnO , WO_3 , Nb_2O_5 , TiO_2 , ZrO_2 , Bi_2O_3 , TeO_2 , etc. Other representative components include halogens (e.g. F, Br, Cl). Whenever a component is included as a term in a mathematical expression or formula, it is understood that the component refers to the amount of the component in units of mol.% in the composition of the glass. For example, the expression “ $B_2O_3 + P_2O_5$ ” refers to the sum of the amount of B_2O_3 in units of mol.% and the amount of P_2O_5 in units of mol.% in the composition of the glass. A mathematical expression or formula is any expression or formula that includes a mathematical operator such as “+”, “-”, “*”, “/”, “min”, or “max”.

[0019] The term “formed from” can mean one or more of comprises, consists essentially of, or consists of. For example, a component that is formed from a particular material can comprise the particular material, consist essentially of the particular material, or consist of the particular material.

[0020] The terms “free” and “substantially free” are used interchangeably herein to refer to an amount and/or an absence of a particular component in a glass composition that is not intentionally added to the glass composition. It is understood that the glass composition may contain traces of a particular constituent component as a contaminant or a tramp in an amount of less than 0.10 mol.%.

[0021] As used herein, the term “tramp”, when used to describe a particular constituent component in a glass composition, refers to a constituent component that is not intentionally added to the glass composition and is present in an amount of less than 0.10 mol.%. Tramp components may be unintentionally added to the glass composition as an impurity in another constituent component and/or through migration of the tramp component into the composition during processing of the glass composition.

[0022] Unless otherwise specified, the term “glass” is used to refer to a glass made from a glass composition disclosed herein.

[0023] The symbol “*” means multiplication when used in any formula herein.

[0024] The term “log” means logarithm in base 10.

[0025] Temperature is expressed herein in units of °C (degrees Celsius).

[0026] The term “room temperature” refers to a temperature in a range from 20 °C to 25 °C.

[0027] Viscosity is expressed herein in units of P (Poise). The term “kP” means kiloPoise.

[0028] The term “glass former” is used herein to refer to a component that, being solely present in the glass composition (i.e., without other components, except for tramps), is able to form a glass when cooling the melt at a rate of not greater than about 300 °C/min.

[0029] The term "modifier", as used herein, refers to the oxides of monovalent or divalent metals, i.e., R_2O or RO , where "R" stands for a cation. Modifiers can be added to a glass composition to change the atomic structure of the melt and the resulting glass. In some embodiments, the modifier may change the coordination numbers of cations present in the glass formers (e.g., boron in B_2O_3), which may result in forming a more polymerized atomic network and, as a result, may provide better glass formation.

[0030] As used herein, the term "RO" refers to a total content of divalent metal oxides (in mol.%), the term " R_2O " refers to a total content of monovalent metal oxides (in mol.%), and the term " Alk_2O " refers to a total content of alkali metal oxides (in mol.%). The term R_2O encompasses alkali metal oxides (Alk_2O), in addition to other monovalent metal oxides, such as Ag_2O , Tl_2O , and Hg_2O , for example. As discussed below, in the present disclosure, a rare earth metal oxide is referred to herein by its normalized formula (RE_2O_3) in which the rare earth metal RE has the redox state "+3," and thus rare earth metal oxides are not encompassed by the term RO.

[0031] As used herein, the term "rare earth metals" refers to the metals listed in the Lanthanide Series of the IUPAC Periodic Table, plus yttrium and scandium. As used herein, the term "rare earth metal oxides," is used to refer to the oxides of rare earth metals in different redox states, such as "+3" for lanthanum in La_2O_3 , "+4" for cerium in CeO_2 , "+2" for europium in EuO , etc. In general, the redox states of rare earth metals in oxide glasses may vary and, in particular, the redox state may change during melting, based on the composition and/or the redox conditions in the furnace where the glass is melted and/or heat-treated (e.g., annealed). Unless otherwise specified, a rare earth metal oxide component is referred to herein by its normalized formula in which the rare earth metal has the redox state "+3." Accordingly, in the case in which a rare earth metal having a redox state other than "+3" is added to the glass composition batch, the batch composition is recalculated by adding or removing some oxygen to maintain the stoichiometry. For example, when CeO_2 (with cerium in redox state "+4") is used as a component, the resulting as-batched composition is recalculated assuming that two moles of CeO_2 is equivalent to one mole of Ce_2O_3 , and the resulting as-batched composition is expressed in terms of Ce_2O_3 . As used herein, the term " RE_mO_n " is used to refer to the total content (in mol.%) of rare earth metal oxides in all redox states present, and the term " RE_2O_3 " is used to refer to the total content (in mol. %) of rare earth metal oxides in the "+3" redox state, also specified as "trivalent equivalent".

[0032] Unless otherwise specified, the composition of all components in a glass are expressed in terms of mole percent (mol.%) in the glass. The composition of components in the glasses reported herein were determined by chemical analysis of the glass. The composition of B_2O_3 was determined using ICP (inductively coupled plasma). The composition of Li_2O was determined using FES

(fluorescence excitation spectroscopy). The composition of all other components was determined using XRF (x-ray fluorescence spectroscopy). For the avoidance of doubt, "composition" or "glass composition" as reported herein refers to the composition of the components in final glass articles and is distinguished from batch composition. As will be understood by those having ordinary skill in the art, the composition of final glass articles may differ from the batch composition used to form the final glass article due, for example, to the fact that various melt constituents (e.g., fluorine, alkali metals, boron, etc.) may be subject to different levels of volatilization (e.g., as a function of vapor pressure, melt time and/or melt temperature) during melting of the constituents in the process of forming the final glass article. Notwithstanding the foregoing, the difference between the composition of final glass articles and the batch composition is expected to be small.

[0033] In the case when fluorine or other halogen (chlorine, bromine, and/or iodine) is added to or is present in an oxide glass, the molecular representation of the resulting glass composition may be expressed in different ways. In the present disclosure, the content of fluorine as a single term, when present, is expressed in terms of atomic percent (at.%), which is determined based on the fraction of fluorine in a total sum of all atoms in a glass composition multiplied by a factor of 100.

[0034] In the present disclosure, the following method of representation of fluorine-containing compositions and concentration ranges is used. The concentration limits for all oxides (e.g. SiO_2 , B_2O_3 , Na_2O , etc.) are presented under the assumption that the respective cations (such as, for example, silicon [Si^{4+}], boron [B^{3+}], sodium [Na^+], etc.) are initially presented in the form of the corresponding oxides. When fluorine is present, for the purposes of calculating the concentration of components of the composition, some part of the oxygen in the oxide is equivalently replaced with fluorine (i.e. one atom of oxygen is replaced with two atoms of fluorine). The fluorine is assumed to be present in the form of silicon fluoride (SiF_4); accordingly, the total sum of all oxides plus SiF_4 is assumed to be 100 mole percent or 100 weight percent in all compositions.

[0035] The density of the glasses at room temperature was determined using the buoyancy method of ASTM C693-93(2013). The estimated error of the density measurements was 0.001 g/cm^3 .

[0036] The term "liquidus temperature" (T_{liq}) is used herein to refer to a temperature above which the glass composition is completely liquid with no crystallization of constituent components of the glass. Unless otherwise specified, liquidus temperature was measured in accordance with ASTM C829-81 (2015), titled "Standard Practice for Measurement of Liquidus Temperature of Glass by the Gradient Furnace Method."

[0037] As used herein, the term "liquidus viscosity" (η_{liq}) refers to the viscosity of a molten glass at the liquidus temperature. Unless specified otherwise, a liquidus viscosity value disclosed in this application is determined by the following method. First, the liquidus temperature of the glass is

measured in accordance with ASTM C829-81 (2015), titled "Standard Practice for Measurement of Liquidus Temperature of Glass by the Gradient Furnace Method." Next, the viscosity of the glass at the liquidus temperature is measured in accordance with ASTM C965-96 (2012), titled "Standard Practice for Measuring Viscosity of Glass Above the Softening Point". The term "Vogel-Fulcher-Tammann (VFT) relation," as used herein, describes the temperature dependence of the viscosity and is represented by the following equation:

$$\log(\eta) = A + B / (T - T_0)$$

where η is viscosity, and A, B and T_0 are empirical parameters. To determine A, B, and T_0 , the viscosity of the glass composition is measured over a given temperature range. The raw data of viscosity versus temperature is then fit with the VFT equation by least-squares fitting to obtain A, B, and T_0 . With these values, a viscosity point (e.g., 200 P Temperature, 35000 P Temperature, and 200000 P Temperature) at any temperature above the softening point may be calculated.

[0038] The term " α ," or " α_{20-300} ," as used herein, refers to the average coefficient of linear thermal expansion (CTE) of the glass composition over a temperature range from 20 °C to 300 °C. CTE is expressed in units of $10^{-7}/^{\circ}\text{C}$ and is measured by using a horizontal dilatometer (push-rod dilatometer) in accordance with ASTM E228-11. The numeric measure of α is a linear average value in a specified temperature range ΔT (e.g., 20 °C to 300 °C) expressed as $\alpha = \Delta L / (L_0 \Delta T)$, where L_0 is the linear size of a sample at some temperature within or near the measured range, and L is the change in the linear size (ΔL) in the measured temperature range ΔT .

[0039] The terms "modulus" and "elastic modulus" refer to Young's modulus at room temperature. The Young's modulus E, shear modulus (G), and the Poisson's ratio (μ) are measured by Resonant Ultrasound Spectroscopy at room temperature, using a Quasar RUSpec 4000 available from ITW Indiana Private Limited, Magnaflux Division using the technique set forth in ASTM E2001-13, titled "Standard Guide for Resonant Ultrasound Spectroscopy for Defect Detection in Both Metallic and Non-metallic Parts."

[0040] The glass transition temperature (T_g) is measured by differential scanning calorimeter (DSC) by heating glass samples initially at room temperature at a rate of 10 K/min.

[0041] The term "softening point" (T_{soft}) refers to the temperature at which the viscosity of the glass composition is $10^{7.6}$ Poise. The softening point of the glass compositions was determined using the fiber elongation method of ASTM C336-71(2015) or a parallel plate viscosity (PPV) method which measures the viscosity of inorganic glass from 10^7 to 10^9 poise as a function of temperature, similar to ASTM C1351M.

[0042] The term "annealing point" (An.P.) refers to the temperature determined according to ASTM C598-93(2013), at which the viscosity of a glass is approximately $10^{13.2}$ Poise.

[0043] The terms "strain point" and "Tstrain" refer to the temperature determined according to ASTM C598-93, at which the viscosity of a glass at a given glass composition is approximately $10^{14.7}$ Poise.

[0044] In the mathematical formulas used in the present disclosure, the term " $\min(A, B)$ " means the lesser of the values A and B, and the term " $\max(A, B)$ " means the greater of the quantities A and B, where "A" and "B" may be any quantities (concentrations of components, values of properties, etc.) The term " $\text{abs}(X)$ " means absolute value of a quantity X.

[0045] The glass composition may include silica (SiO_2). In embodiments of the glass compositions disclosed herein, SiO_2 is the primary constituent of the glass network. Pure SiO_2 has a relatively high specific modulus, a low CTE, and a high annealing point. However, pure SiO_2 also has a high melting point. If the concentration of SiO_2 in a glass composition is too high, the formability of the glass composition may be diminished as high concentrations of SiO_2 increase the difficulty in melting the glass, which, in turn, adversely impacts the formability of the glass. Accordingly, the content of silica is limited. In embodiments, the glass composition may contain silica (SiO_2) in an amount from greater than or equal to 54.0 mol.% to less than or equal to 84.0 mol.% and all ranges and sub-ranges between the foregoing values. In some embodiments, the glass composition may contain SiO_2 in an amount greater than or equal to 54.0 mol.%, greater than or equal to 55.0 mol.%, greater than or equal to 60.0 mol.%, greater than or equal to 63.9 mol.%, greater than or equal to 64.5 mol.%, greater than or equal to 69.0 mol.%, greater than or equal to 70.0 mol.%, greater than or equal to 74.0 mol.%, or greater than or equal to 79.0 mol.%. In some other embodiments, the glass composition may contain SiO_2 in an amount less than or equal to 84.0 mol.%, less than or equal to 80.0 mol.%, less than or equal to 79.0 mol.%, less than or equal to 75.0 mol.%, less than or equal to 74.4 mol.%, less than or equal to 74.0 mol.%, less than or equal to 70.0 mol.%, less than or equal to 69.0 mol.%, or less than or equal to 55.0 mol.%. In some more embodiments, the glass composition may contain SiO_2 in an amount greater than or equal to 60.0 mol.% and less than or equal to 80.0 mol.%, greater than or equal to 60.0 mol.% and less than or equal to 75.0 mol.%, greater than or equal to 63.9 mol.% and less than or equal to 74.4 mol.%, greater than or equal to 64.5 mol.% and less than or equal to 74.5 mol.%, greater than or equal to 54.0 mol.% and less than or equal to 84.0 mol.%, greater than or equal to 54.0 mol.% and less than or equal to 55.0 mol.%, greater than or equal to 55.0 mol.% and less than or equal to 69.0 mol.%, greater than or equal to 60.0 mol.% and less than or equal to 69.0 mol.%, greater than or equal to 63.9 mol.% and less than or equal to 69.0 mol.%, greater than or equal to 64.5 mol.% and less than or equal to 69.0 mol.%, greater than or equal to 69.0 mol.% and less than or equal to 84.0 mol.%, greater than or equal to 69.0 mol.% and less than or equal to 70.0 mol.%.

[0046] The glass composition may include boron oxide (B_2O_3). Like SiO_2 , Al_2O_3 , and P_2O_5 , B_2O_3 may be added to the glass composition as a network former. However, it has been found that additions of boron significantly reduce diffusivity of alkali ions in the glass, which, in turn, adversely impacts the ion exchange performance of the resultant glass. The addition of B_2O_3 in some amounts may help to decrease the viscosity of a glass-forming melt, therefore decreasing the temperature required for melting and forming. However, if too much B_2O_3 is added, volatility may occur at the surface of a glass melt, resulting in compositional inhomogeneity or localized reductions in the viscosity at the liquidus temperature for areas of the glass melt. Further, the addition of too much B_2O_3 may adversely affect the annealing point of the glass composition. Accordingly, the content of boron oxide is preferably limited, or glass compositions may be substantially free of B_2O_3 . In embodiments, the glass composition may contain boron oxide (B_2O_3) in an amount from greater than or equal to 0.0 mol.% to less than or equal to 10.0 mol.% and all ranges and sub-ranges between the foregoing values. In some embodiments, the glass composition may contain B_2O_3 in an amount greater than or equal to 0.0 mol.%, greater than or equal to 5.0 mol.%, greater than or equal to 7.0 mol.%, greater than or equal to 8.0 mol.%, or greater than or equal to 9.0 mol.%. In some other embodiments, the glass composition may contain B_2O_3 in an amount less than or equal to 10.0 mol.%, less than or equal to 9.0 mol.%, less than or equal to 8.0 mol.%, less than or equal to 7.0 mol.%, less than or equal to 5.0 mol.%, less than or equal to 3.8 mol.%, less than or equal to 3.0 mol.%, less than or equal to 0.8 mol.%, or less than or equal to 0.7 mol.%. In some more embodiments, the glass composition may contain B_2O_3 in an amount greater than or equal to 0.0 mol.% and less than or equal to 3.8 mol.%, greater than or equal to 0.0 mol.% and less than or equal to 3.0 mol.%, greater than or equal to 0.0 mol.% and less than or equal to 0.8 mol.%, greater than or equal to 0.0 mol.% and less than or equal to 0.7 mol.%, greater than or equal to 0.0 mol.% and less than or equal to 10.0 mol.%, greater than or equal to 5.0 mol.% and less than or equal to 7.0 mol.%, greater than or equal to 7.0 mol.% and less than or equal to 10.0 mol.%, greater than or equal to 8.0 mol.% and less than or equal to 10.0 mol.%, greater than or equal to 8.0 mol.% and less than or equal to 9.0 mol.%.

[0047] The glass composition may include phosphorus oxide (P_2O_5). Like SiO_2 and Al_2O_3 , P_2O_5 may be added to the glass composition as a network former, thereby reducing the meltability and formability of the glass composition. In some amounts, P_2O_5 can be added to the glass composition to increase the viscosity at the liquidus temperature, therefore limiting the propensity to crystallize during the cooling of the glass-forming melt. The addition of P_2O_5 may also increase the diffusivity of ions in the glass composition during ion-exchange treatments, thereby increasing the efficiency of such treatments. However, if too much P_2O_5 is added, the annealing point of the glass may be

reduced, the CTE may be increased, or phase separation may be induced upon cooling of the glass-forming melt. Accordingly, the content of phosphorus oxide is preferably limited, or glass compositions may be substantially free of P_2O_5 . In embodiments, the glass composition may contain phosphorus oxide (P_2O_5) in an amount from greater than or equal to 0.0 mol.% to less than or equal to 10.0 mol.% and all ranges and sub-ranges between the foregoing values. In some embodiments, the glass composition may contain P_2O_5 in an amount greater than or equal to 0.0 mol.%, greater than or equal to 1.1 mol.%, greater than or equal to 2.5 mol.%, greater than or equal to 5.0 mol.%, greater than or equal to 7.0 mol.%, greater than or equal to 8.0 mol.%, or greater than or equal to 9.0 mol.%. In some other embodiments, the glass composition may contain P_2O_5 in an amount less than or equal to 10.0 mol.%, less than or equal to 9.0 mol.%, less than or equal to 8.0 mol.%, less than or equal to 7.0 mol.%, less than or equal to 5.0 mol.%, less than or equal to 4.0 mol.%, less than or equal to 3.4 mol.%, or less than or equal to 3.0 mol.%. In some more embodiments, the glass composition may contain P_2O_5 in an amount greater than or equal to 0.0 mol.% and less than or equal to 4.0 mol.%, greater than or equal to 0.0 mol.% and less than or equal to 3.4 mol.%, greater than or equal to 0.0 mol.% and less than or equal to 3.0 mol.%, greater than or equal to 1.1 mol.% and less than or equal to 2.99 mol.%, greater than or equal to 0.0 mol.% and less than or equal to 10.0 mol.%, greater than or equal to 1.1 mol.% and less than or equal to 3.0 mol.%, greater than or equal to 5.0 mol.% and less than or equal to 7.0 mol.%.

[0048] The glass composition may have limitations for the amount of rare earth metal oxides. Rare earth metal oxides, especially yttrium oxide (Y_2O_3), can be added to the glass composition to increase the specific modulus, annealing and strain points, at the same time decreasing the high-temperature viscosity and, thus, allowing for melting of the glass at a lower temperature more compatible with the conventional refractories. However, when the content of RE_mO_n is high, it may adversely increase the liquidus temperature because of crystallization of refractory minerals, such as rare earth metal aluminates ($RE_3Al_5O_{12}$), silicates (RE_2SiO_5 , $RE_2Si_2O_7$) and others, which may, in turn, decrease the liquidus viscosity. Accordingly, the content of rare earth metal oxides is preferably limited, or glass compositions may be substantially free of RE_mO_n .

[0049] In some other embodiments, the glass composition may contain rare earth metal oxides RE_mO_n in an amount less than or equal to 5.0 mol.%, less than or equal to 3.0 mol.%, or less than or equal to 2.5 mol.%. In some more embodiments, the glass composition may contain RE_mO_n in an amount greater than or equal to 0.0 mol.% and less than or equal to 5.0 mol.%, greater than or equal to 0.0 mol.% and less than or equal to 3.0 mol.%, or greater than or equal to 0.0 mol.% and less than or equal to 2.5 mol.%.

[0050] The glass composition may include lanthanum oxide (La_2O_3). Lanthanum oxide can be added to the glass compositions to increase the annealing and strain points, while at the same time decreasing the melting temperature. However, adding lanthanum oxide may decrease the liquidus viscosity and specific modulus of the glass. Accordingly, the content of lanthanum oxide is preferably limited, or glass compositions may be substantially free of La_2O_3 . In embodiments, the glass composition may contain lanthanum oxide (La_2O_3) in an amount from greater than or equal to 0.0 mol.% to less than or equal to 5.0 mol.% and all ranges and sub-ranges between the foregoing values. In some other embodiments, the glass composition may contain La_2O_3 in an amount less than or equal to 5.0 mol.%, less than or equal to 2.5 mol.%, or less than or equal to 0.7 mol.%. In some more embodiments, the glass composition may contain La_2O_3 in an amount greater than or equal to 0.0 mol.% and less than or equal to 0.7 mol.%, greater than or equal to 0.0 mol.% and less than or equal to 5.0 mol.%.

[0051] Glass composition may include alkali metal oxides (Alk_2O). Alkali metal oxides can be added to the glass compositions to decrease the liquidus temperature and high-temperature viscosity, maintaining high liquidus viscosity. Also, adding Alk_2O can increase the solubility of high-modulus species, such as, for example, TiO_2 , ZrO_2 , Y_2O_3 and others, thus indirectly increasing the specific modulus. Also, in presence of boron oxide (B_2O_3), alkali metal oxides may transform the boron atoms from trigonal to tetrahedral form, which may also indirectly increase the specific modulus of glass.

[0052] In embodiments, the glass composition may contain alkali metal oxides (Alk_2O) in an amount from greater than or equal to 0.0 mol.% to less than or equal to 12.0 mol.% and all ranges and sub-ranges between the foregoing values. In some embodiments, the glass composition may contain alkali metal oxides Alk_2O in an amount greater than or equal to 0.0 mol.%, greater than or equal to 0.5 mol.%, or greater than or equal to 2.5 mol.%. In some other embodiments, the glass composition may contain Alk_2O in an amount less than or equal to 12.0 mol.%, less than or equal to 10.0 mol.%, less than or equal to 9.0 mol.%, less than or equal to 8.0 mol.%, less than or equal to 6.0 mol.%. In some more embodiments, the glass composition may contain Alk_2O in an amount greater than or equal to 0.0 mol.% and less than or equal to 12.0 mol.%, greater than or equal to 0.0 mol.% and less than or equal to 10.0 mol.%, greater than or equal to 0.0 mol.% and less than or equal to 8.0 mol.%, greater than or equal to 1.0 mol.% and less than or equal to 8.0 mol.%, greater than or equal to 2.0 mol.% and less than or equal to 7.0 mol.%.

[0053] The glass composition may include lithium oxide (Li_2O). Li_2O lowers the viscosity of a glass, which enhances the melting behavior, formability, and the Young's modulus, and may enable ion exchangeability. However, when too much Li_2O is added, the CTE of the glass increases and the strain point decreases. Further, if the amount of Li_2O is too high, Li_2O may cause crystallization of

undesired species such as spodumene ($\text{LiAl}(\text{SiO}_3)_2$) which may reduce the viscosity at the liquidus temperature and, therefore, increase the critical cooling rate, which may cause crystallization of the glass-forming melt when cooling. Accordingly, the content of lithium oxide is preferably limited, or glass compositions may be substantially free of Li_2O . In embodiments, the glass composition may contain lithium oxide (Li_2O) in an amount from greater than or equal to 0.0 mol.% to less than or equal to 10.0 mol.% and all ranges and sub-ranges between the foregoing values. In some embodiments, the glass composition may contain Li_2O in an amount greater than or equal to 0.0 mol.%, greater than or equal to 0.3 mol.%, greater than or equal to 0.5 mol.%, greater than or equal to 2.8 mol.%, greater than or equal to 3.0 mol.%, greater than or equal to 3.2 mol.%, greater than or equal to 5.0 mol.%, greater than or equal to 7.0 mol.%, greater than or equal to 8.0 mol.%, or greater than or equal to 9.0 mol.%. In some other embodiments, the glass composition may contain Li_2O in an amount less than or equal to 10.0 mol.%, less than or equal to 9.0 mol.%, less than or equal to 8.0 mol.%, less than or equal to 7.8 mol.%, less than or equal to 7.5 mol.%, less than or equal to 7.4 mol.%, less than or equal to 7.0 mol.%, less than or equal to 6.2 mol.%, less than or equal to 6.0 mol.%, or less than or equal to 5.0 mol.%. In some more embodiments, the glass composition may contain Li_2O in an amount greater than or equal to 0.0 mol.% and less than or equal to 10.0 mol.%, greater than or equal to 0.0 mol.% and less than or equal to 7.8 mol.%, greater than or equal to 0.0 mol.% and less than or equal to 7.4 mol.%, greater than or equal to 0.3 mol.% and less than or equal to 7.5 mol.%, greater than or equal to 0.5 mol.% and less than or equal to 7.5 mol.%, greater than or equal to 2.8 mol.% and less than or equal to 7.0 mol.%, greater than or equal to 3.0 mol.% and less than or equal to 6.32 mol.%, greater than or equal to 3.2 mol.% and less than or equal to 6.2 mol.%, greater than or equal to 0.0 mol.% and less than or equal to 5.0 mol.%, greater than or equal to 0.3 mol.% and less than or equal to 5.0 mol.%, greater than or equal to 0.5 mol.% and less than or equal to 5.0 mol.%, greater than or equal to 3.0 mol.% and less than or equal to 10.0 mol.%, greater than or equal to 3.0 mol.% and less than or equal to 5.0 mol.%, greater than or equal to 3.2 mol.% and less than or equal to 5.0 mol.%, greater than or equal to 5.0 mol.% and less than or equal to 6.0 mol.%.

[0054] The glass composition may include sodium oxide (Na_2O). Na_2O lowers the viscosity of a glass, which enhances the melting behavior, formability, and may enable ion exchangeability. However, when too much Na_2O is added, the CTE of the glass composition increases, the Young's modulus decreases, and the strain point decreases. Accordingly, the content of sodium oxide is preferably limited, or glass compositions may be substantially free of Na_2O . In embodiments, the glass composition may contain sodium oxide (Na_2O) in an amount from greater than or equal to 0.0 mol.% to less than or equal to 10.0 mol.% and all ranges and sub-ranges between the foregoing values. In

some embodiments, the glass composition may contain Na_2O in an amount greater than or equal to 0.0 mol.%, greater than or equal to 0.4 mol.%, greater than or equal to 0.9 mol.%, greater than or equal to 1.0 mol.%, greater than or equal to 5.0 mol.%, greater than or equal to 7.0 mol.%, greater than or equal to 8.0 mol.%, or greater than or equal to 9.0 mol.%. In some other embodiments, the glass composition may contain Na_2O in an amount less than or equal to 10.0 mol.%, less than or equal to 9.0 mol.%, less than or equal to 8.0 mol.%, less than or equal to 7.0 mol.%, less than or equal to 5.0 mol.%, less than or equal to 4.8 mol.%, less than or equal to 4.4 mol.%, or less than or equal to 4.0 mol.%. In some more embodiments, the glass composition may contain Na_2O in an amount greater than or equal to 0.0 mol.% and less than or equal to 5.0 mol.%, greater than or equal to 0.0 mol.% and less than or equal to 4.8 mol.%, greater than or equal to 0.4 mol.% and less than or equal to 4.4 mol.%, greater than or equal to 0.9 mol.% and less than or equal to 4.0 mol.%, greater than or equal to 1.0 mol.% and less than or equal to 3.98 mol.%, greater than or equal to 0.0 mol.% and less than or equal to 10.0 mol.%, greater than or equal to 0.0 mol.% and less than or equal to 4.0 mol.%, greater than or equal to 0.4 mol.% and less than or equal to 4.0 mol.%, greater than or equal to 0.9 mol.% and less than or equal to 10.0 mol.%, greater than or equal to 1.0 mol.% and less than or equal to 4.0 mol.%, greater than or equal to 5.0 mol.% and less than or equal to 10.0 mol.%, greater than or equal to 5.0 mol.% and less than or equal to 7.0 mol.%.

[0055] The glass composition may include zinc oxide (ZnO). Zinc oxide can be added to the glass composition to increase the specific modulus, similar to MgO . However, when the content of ZnO is high, it may cause crystallization of refractory minerals, such as, for example, gahnite (ZnAl_2O_4), or liquid-liquid phase separation of the glass forming melt. Accordingly, the content of zinc oxide is preferably limited, or glass compositions may be substantially free of ZnO . In embodiments, the glass composition may contain zinc oxide (ZnO) in an amount from greater than or equal to 0.0 mol.% to less than or equal to 5.0 mol.% and all ranges and sub-ranges between the foregoing values. In some other embodiments, the glass composition may contain ZnO in an amount less than or equal to 5.0 mol.%, less than or equal to 3.0 mol.%, less than or equal to 2.5 mol.%, or less than or equal to 2.0 mol.%. In some more embodiments, the glass composition may contain ZnO in an amount greater than or equal to 0.0 mol.% and less than or equal to 3.0 mol.%, greater than or equal to 0.0 mol.% and less than or equal to 2.0 mol.%, greater than or equal to 0.0 mol.% and less than or equal to 5.0 mol.%.

[0056] The glass composition may include zirconia (ZrO_2). Zirconia can be added to the glass compositions to increase the specific modulus. However, addition of even small amounts of zirconia to alumina-rich compositions may cause crystallization of refractory minerals, such as zircon (ZrSiO_4) or zirconia (ZrO_2), which may increase the liquidus temperature and decrease the liquidus viscosity.

Accordingly, the content of zirconia is preferably limited, or glass compositions may be substantially free of ZrO_2 . In embodiments, the glass composition may contain zirconia (ZrO_2) in an amount from greater than or equal to 0.0 mol.% to less than or equal to 5.0 mol.% and all ranges and sub-ranges between the foregoing values. In some other embodiments, the glass composition may contain ZrO_2 in an amount less than or equal to 5.0 mol.%, less than or equal to 2.5 mol.%, less than or equal to 1.5 mol.%, or less than or equal to 0.5 mol.%. In some more embodiments, the glass composition may contain ZrO_2 in an amount greater than or equal to 0.0 mol.% and less than or equal to 0.5 mol.%, greater than or equal to 0.0 mol.% and less than or equal to 5.0 mol.%.

[0057] The glass composition may include barium oxide (BaO). The addition of BaO increases the viscosity at the liquidus temperature, therefore limiting the propensity to crystallize during the cooling of the glass-forming melt. However, when too much BaO is added, the density and the CTE of the glass increases. Accordingly, the content of barium oxide is preferably limited, or glass compositions may be substantially free of BaO. In embodiments, the glass composition may contain barium oxide (BaO) in an amount from greater than or equal to 0.0 mol.% to less than or equal to 10.0 mol.% and all ranges and sub-ranges between the foregoing values. In some embodiments, the glass composition may contain BaO in an amount greater than or equal to 0.0 mol.%, or greater than or equal to 5.0 mol.%. In some other embodiments, the glass composition may contain BaO in an amount less than or equal to 10.0 mol.%, less than or equal to 5.0 mol.%, less than or equal to 3.0 mol.%, less than or equal to 1.0 mol.%, less than or equal to 0.7 mol.%, or less than or equal to 0.625 mol.%. In some more embodiments, the glass composition may contain BaO in an amount greater than or equal to 0.0 mol.% and less than or equal to 3.0 mol.%, greater than or equal to 0.0 mol.% and less than or equal to 1.0 mol.%, greater than or equal to 0.0 mol.% and less than or equal to 0.7 mol.%, greater than or equal to 0.0 mol.% and less than or equal to 0.625 mol.%, greater than or equal to 0.0 mol.% and less than or equal to 10.0 mol.%.

[0058] The glass composition may include strontium oxide (SrO). SrO lowers the viscosity of a glass, which enhances the melting behavior, formability, and the strain point of the glass composition. In some embodiments, the addition of SrO increases the viscosity at the liquidus temperature, therefore limiting the propensity to crystallize during the cooling of the glass-forming melt. However, when too much SrO is added, the density and the CTE of the glass increase. Accordingly, the content of strontium oxide is preferably limited, or glass compositions may be substantially free of SrO. In embodiments, the glass composition may contain strontium oxide (SrO) in an amount from greater than or equal to 0.0 mol.% to less than or equal to 10.0 mol.% and all ranges and sub-ranges between the foregoing values. In some embodiments, the glass composition may contain SrO in an amount greater than or equal to 0.0 mol.%, greater than or equal to 0.49 mol.%, greater than or

equal to 5.0 mol.%, greater than or equal to 7.0 mol.%, greater than or equal to 8.0 mol.%, or greater than or equal to 9.0 mol.%. In some other embodiments, the glass composition may contain SrO in an amount less than or equal to 10.0 mol.%, less than or equal to 9.0 mol.%, less than or equal to 8.0 mol.%, less than or equal to 7.0 mol.%, less than or equal to 5.0 mol.%, less than or equal to 1.5 mol.%, less than or equal to 1.19 mol.%, less than or equal to 1.1 mol.%, or less than or equal to 0.95 mol.%. In some more embodiments, the glass composition may contain SrO in an amount greater than or equal to 0.0 mol.% and less than or equal to 1.5 mol.%, greater than or equal to 0.0 mol.% and less than or equal to 1.1 mol.%, greater than or equal to 0.0 mol.% and less than or equal to 0.95 mol.%, greater than or equal to 0.49 mol.% and less than or equal to 1.19 mol.%, greater than or equal to 0.0 mol.% and less than or equal to 10.0 mol.%, greater than or equal to 0.49 mol.% and less than or equal to 0.95 mol.%, greater than or equal to 5.0 mol.% and less than or equal to 7.0 mol.%, greater than or equal to 8.0 mol.% and less than or equal to 9.0 mol.%.

[0059] The glass composition may include magnesia (MgO). MgO lowers the viscosity of a glass, which enhances the melting behavior, formability, the strain point, and the Young's modulus, and may improve ion exchangeability. However, when too much MgO is added, the density and the CTE of the glass increases. Further, if the amount of MgO is too high, MgO may cause crystallization of refractory species such as cordierite ($\text{Mg}_2\text{Al}_4\text{Si}_5\text{O}_{18}$) which may reduce the viscosity at the liquidus temperature and, therefore, increase the critical cooling rate, which may cause crystallization of the glass-forming melt when cooling. Accordingly, the content of magnesia is preferably limited, or glass compositions may be substantially free of MgO. In embodiments, the glass composition may contain magnesia (MgO) in an amount from greater than or equal to 0.0 mol.% to less than or equal to 10.0 mol.% and all ranges and sub-ranges between the foregoing values. In some embodiments, the glass composition may contain MgO in an amount greater than or equal to 0.0 mol.%, greater than or equal to 0.4 mol.%, greater than or equal to 1.0 mol.%, greater than or equal to 2.0 mol.%, greater than or equal to 5.0 mol.%, greater than or equal to 7.0 mol.%, greater than or equal to 8.0 mol.%, or greater than or equal to 9.0 mol.%. In some other embodiments, the glass composition may contain MgO in an amount less than or equal to 10.0 mol.%, less than or equal to 9.0 mol.%, less than or equal to 8.0 mol.%, less than or equal to 7.0 mol.%, less than or equal to 6.0 mol.%, less than or equal to 5.0 mol.%, less than or equal to 4.4 mol.%, less than or equal to 4.3 mol.%, or less than or equal to 4.0 mol.%. In some more embodiments, the glass composition may contain MgO in an amount greater than or equal to 0.0 mol.% and less than or equal to 6.0 mol.%, greater than or equal to 0.0 mol.% and less than or equal to 5.0 mol.%, greater than or equal to 0.0 mol.% and less than or equal to 4.3 mol.%, greater than or equal to 0.4 mol.% and less than or equal to 4.4 mol.%, greater than or equal to 1.0 mol.% and less than or equal to 4.0 mol.%, greater than or equal to 1.99 mol.%

and less than or equal to 4.31 mol.%, greater than or equal to 0.0 mol.% and less than or equal to 10.0 mol.%, greater than or equal to 0.0 mol.% and less than or equal to 4.0 mol.%, greater than or equal to 0.4 mol.% and less than or equal to 4.0 mol.%, greater than or equal to 5.0 mol.% and less than or equal to 6.0 mol.%.

[0060] The glass composition may include calcium oxide (CaO). CaO lowers the viscosity of a glass, which enhances the melting behavior, formability, the strain point, and the Young's modulus, and may improve ion exchangeability. However, when too much CaO is added, the density and the CTE of the glass increases. Further, if the amount of CaO is too high, CaO may cause crystallization of undesired species such as anorthite ($\text{CaAl}_2\text{Si}_2\text{O}_8$) which may reduce the viscosity at the liquidus temperature and, therefore, increase the critical cooling rate, which may cause crystallization of the glass-forming melt when cooling. Accordingly, the content of calcium oxide is preferably limited, or glass compositions may be substantially free of CaO. In embodiments, the glass composition may contain calcium oxide (CaO) in an amount from greater than or equal to 0.0 mol.% to less than or equal to 10.0 mol.% and all ranges and sub-ranges between the foregoing values. In some embodiments, the glass composition may contain CaO in an amount greater than or equal to 0.0 mol.%, greater than or equal to 0.3 mol.%, greater than or equal to 1.0 mol.%, greater than or equal to 1.4 mol.%, greater than or equal to 1.9 mol.%, greater than or equal to 3.0 mol.%, greater than or equal to 5.0 mol.%, greater than or equal to 7.0 mol.%, greater than or equal to 8.0 mol.%, or greater than or equal to 9.0 mol.%. In some other embodiments, the glass composition may contain CaO in an amount less than or equal to 10.0 mol.%, less than or equal to 9.0 mol.%, less than or equal to 8.0 mol.%, less than or equal to 7.5 mol.%, less than or equal to 7.0 mol.%, less than or equal to 6.0 mol.%, less than or equal to 5.6 mol.%, or less than or equal to 5.0 mol.%. In some more embodiments, the glass composition may contain CaO in an amount greater than or equal to 0.0 mol.% and less than or equal to 10.0 mol.%, greater than or equal to 0.3 mol.% and less than or equal to 7.5 mol.%, greater than or equal to 1.0 mol.% and less than or equal to 6.0 mol.%, greater than or equal to 1.4 mol.% and less than or equal to 5.6 mol.%, greater than or equal to 1.9 mol.% and less than or equal to 5.0 mol.%, greater than or equal to 2.5 mol.% and less than or equal to 5.17 mol.%, greater than or equal to 0.0 mol.% and less than or equal to 5.0 mol.%, greater than or equal to 0.3 mol.% and less than or equal to 5.0 mol.%, greater than or equal to 1.0 mol.% and less than or equal to 5.0 mol.%, greater than or equal to 1.4 mol.% and less than or equal to 5.0 mol.%, greater than or equal to 1.9 mol.% and less than or equal to 10.0 mol.%, greater than or equal to 5.0 mol.% and less than or equal to 5.6 mol.%.

[0061] The glass composition may include alumina (Al_2O_3). Al_2O_3 may serve as a glass network former, similar to SiO_2 . Al_2O_3 may increase the viscosity of the glass composition due to its

tetrahedral coordination in a glass melt formed from a glass composition, decreasing the formability of the glass compositions when the amount of Al_2O_3 is too high. In some embodiments when the concentration of Al_2O_3 is balanced against the concentration of SiO_2 and the concentration of alkali oxides in the glass composition, Al_2O_3 can reduce the liquidus temperature of the glass melt, thereby reducing the glass liquidus viscosity and improving the compatibility with certain forming process. However if the amount of Al_2O_3 is too high, particularly in excess of the total modifier content (alkali oxides and alkaline earth oxides), Al_2O_3 may cause crystallization of refractory species such as corundum (Al_2O_3) or aluminosilicates like mullite ($3\text{Al}_2\text{O}_3 \cdot 2\text{SiO}_2$) which may reduce the viscosity at the liquidus temperature and, therefore, increase the critical cooling rate, which may cause crystallization of the glass-forming melt when cooling. Accordingly, the content of alumina is preferably limited. In embodiments, the glass composition may contain alumina (Al_2O_3) in an amount from greater than or equal to 10.0 mol.% to less than or equal to 19.0 mol.% and all ranges and sub-ranges between the foregoing values. In some embodiments, the glass composition may contain Al_2O_3 in an amount greater than or equal to 10.0 mol.%, greater than or equal to 10.5 mol.%, greater than or equal to 11.0 mol.%, greater than or equal to 12.0 mol.%, greater than or equal to 12.34 mol.%, greater than or equal to 15.0 mol.%, greater than or equal to 16.0 mol.%, greater than or equal to 17.0 mol.%, or greater than or equal to 18.0 mol.%. In some other embodiments, the glass composition may contain Al_2O_3 in an amount less than or equal to 19.0 mol.%, less than or equal to 18.0 mol.%, less than or equal to 17.0 mol.%, less than or equal to 16.0 mol.%, less than or equal to 15.0 mol.%, or less than or equal to 11.0 mol.%. In some more embodiments, the glass composition may contain Al_2O_3 in an amount greater than or equal to 10.0 mol.% and less than or equal to 20.0 mol.%, greater than or equal to 10.0 mol.% and less than or equal to 18.0 mol.%, greater than or equal to 10.0 mol.% and less than or equal to 17.0 mol.%, greater than or equal to 10.5 mol.% and less than or equal to 18.0 mol.%, greater than or equal to 11.0 mol.% and less than or equal to 18.0 mol.%, greater than or equal to 12.0 mol.% and less than or equal to 18.0 mol.%, greater than or equal to 12.34 mol.% and less than or equal to 18.0 mol.%, greater than or equal to 10.0 mol.% and less than or equal to 11.0 mol.%, greater than or equal to 10.5 mol.% and less than or equal to 20.0 mol.%, greater than or equal to 10.5 mol.% and less than or equal to 11.0 mol.%, greater than or equal to 11.0 mol.% and less than or equal to 20.0 mol.%, greater than or equal to 11.0 mol.% and less than or equal to 15.0 mol.%, greater than or equal to 12.0 mol.% and less than or equal to 20.0 mol.%, greater than or equal to 12.0 mol.% and less than or equal to 15.0 mol.%,

[0062] Glass composition may include fluorine (F). Fluorine can be added in a small amount to the glass compositions of the present disclosure as a fining agent. However, fluorine may cause

environmental concern. Accordingly, the content of fluorine is preferably limited, or glass compositions may be substantially free of fluorine. In embodiments, the glass composition may contain fluorine (F) in an amount from greater than or equal to 0.0 at.% to less than or equal to 3.0 at.% and all ranges and sub-ranges between the foregoing values. In some other embodiments, the glass composition may contain F in an amount less than or equal to 3.0 at.%, less than or equal to 2.0 at.%, less than or equal to 1.0 at.%, less than or equal to 0.5 at.%, or less than or equal to 0.05 at.%. In some more embodiments, the glass composition may contain F in an amount greater than or equal to 0.0 at.% and less than or equal to 3.0 at.%, greater than or equal to 0.0 at.% and less than or equal to 0.5 at.%, greater than or equal to 0.0 at.% and less than or equal to 0.05 at.%.

[0063] In some embodiments, the glass composition may have a sum of CaO+MgO greater than or equal to 0.0 mol.%, or greater than or equal to 5.0 mol.%. In some other embodiments, the glass composition may have a sum of CaO+MgO less than or equal to 9.0 mol.% or less than or equal to 5.0 mol.%. In some more embodiments, the glass composition may have a sum of CaO+MgO greater than or equal to 0.0 mol.% and less than or equal to 9.0 mol.%, or greater than or equal to 0.0 mol.% and less than or equal to 5.0 mol.%.

[0064] In some embodiments, the glass composition may have a sum of CaO+SrO greater than or equal to 0.0 mol.%, greater than or equal to 2.0 mol.%, or greater than or equal to 5.0 mol.%. In some other embodiments, the glass composition may have a sum of CaO+SrO less than or equal to 9.0 mol.% or less than or equal to 5.0 mol.%. In some more embodiments, the glass composition may have a sum of CaO+SrO greater than or equal to 0.0 mol.% and less than or equal to 9.0 mol.%, or greater than or equal to 0.0 mol.% and less than or equal to 5.0 mol.%, greater than or equal to 2.0 mol.% and less than or equal to 9.0 mol.%, or greater than or equal to 2.0 mol.% and less than or equal to 5.0 mol.%.

[0065] In some embodiments, the glass composition may have a sum of Li₂O+MgO greater than or equal to 0.0 mol.%, or greater than or equal to 5.0 mol.%. In some other embodiments, the glass composition may have a sum of Li₂O+MgO less than or equal to 10.0 mol.%, less than or equal to 8.0 mol.%, or less than or equal to 5.0 mol.%. In some more embodiments, the glass composition may have a sum of Li₂O+MgO greater than or equal to 0.0 mol.% and less than or equal to 10.0 mol.%, greater than or equal to 0.0 mol.% and less than or equal to 8.0 mol.%, or greater than or equal to 0.0 mol.% and less than or equal to 5.0 mol.%, greater than or equal to 5.0 mol.% and less than or equal to 10.0 mol.%, or greater than or equal to 5.0 mol.% and less than or equal to 8.0 mol.%.

[0066] In some embodiments, the glass composition may have a sum of Li₂O+Na₂O greater than or equal to 0.0 mol.%, greater than or equal to 0.5 mol.%, greater than or equal to 4.0 mol.%, or greater than or equal to 5.0 mol.%. In some other embodiments, the glass composition may have a

sum of $\text{Li}_2\text{O}+\text{Na}_2\text{O}$ less than or equal to 9.0 mol.% or less than or equal to 5.0 mol.%. In some more embodiments, the glass composition may have a sum of $\text{Li}_2\text{O}+\text{Na}_2\text{O}$ greater than or equal to 0.0 mol.% and less than or equal to 9.0 mol.%, or greater than or equal to 0.0 mol.% and less than or equal to 5.0 mol.%, greater than or equal to 0.5 mol.% and less than or equal to 9.0 mol.%, or greater than or equal to 0.5 mol.% and less than or equal to 5.0 mol.%, greater than or equal to 4.0 mol.% and less than or equal to 9.0 mol.%.

[0067] In some embodiments, the glass composition may have a sum of $\text{MgO}+\text{CaO}+\text{SrO}+\text{BaO}+\text{ZnO}$ greater than or equal to 0.0 mol.%, greater than or equal to 0.5 mol.%, or greater than or equal to 1.0 mol.%.

[0068] In some embodiments, the glass composition may have a sum of $\text{MgO}+\text{ZnO}$ greater than or equal to 0.0 mol.%, greater than or equal to 1.0 mol.%, or greater than or equal to 5.0 mol.%. In some other embodiments, the glass composition may have a sum of $\text{MgO}+\text{ZnO}$ less than or equal to 10.0 mol.%, less than or equal to 5.0 mol.%, or less than or equal to 4.0 mol.%. In some more embodiments, the glass composition may have a sum of $\text{MgO}+\text{ZnO}$ greater than or equal to 0.0 mol.% and less than or equal to 10.0 mol.%, greater than or equal to 0.0 mol.% and less than or equal to 5.0 mol.%, or greater than or equal to 0.0 mol.% and less than or equal to 4.0 mol.%, greater than or equal to 1.0 mol.% and less than or equal to 10.0 mol.%, greater than or equal to 1.0 mol.% and less than or equal to 5.0 mol.%, or greater than or equal to 1.0 mol.% and less than or equal to 4.0 mol.%.

[0069] In some other embodiments, the glass composition may have a sum of $\text{ZrO}_2+\text{TiO}_2+\text{FeO}+\text{Fe}_2\text{O}_3$ less than or equal to 1.5 mol.%, less than or equal to 1.0 mol.%, or less than or equal to 0.021 mol.%. In some more embodiments, the glass composition may have a sum of $\text{ZrO}_2+\text{TiO}_2+\text{FeO}+\text{Fe}_2\text{O}_3$ greater than or equal to 0.0 mol.% and less than or equal to 1.5 mol.%, greater than or equal to 0.0 mol.% and less than or equal to 1.0 mol.%, or greater than or equal to 0.0 mol.% and less than or equal to 0.021 mol.%.

[0070] In some embodiments, the glass composition may have limitations for $\min(\text{RE}_m\text{O}_n, \text{P}_2\text{O}_5)$. It was empirically found that in alumina-rich glasses, especially when the content of alumina exceeds the total content of alkali metal oxides, adding even small amounts of rare earth metal oxides (RE_mO_n), especially yttrium oxide (Y_2O_3), together with P_2O_5 may cause crystallization of rare earth metal phosphates, such as yttrium phosphate (YPO_4), at high temperatures, which may raise the liquidus temperature and decrease the liquidus viscosity of glass. Accordingly, in some embodiments, it is preferable that the glass composition either contains only one of P_2O_5 or RE_mO_n , or is substantially free of both P_2O_5 and RE_mO_n . In some embodiments, the glass may have a value of $\min(\text{RE}_m\text{O}_n, \text{P}_2\text{O}_5)$ greater than or equal to 0.00 mol.%, greater than or equal to 0.10 mol.%, or greater

than or equal to 0.20 mol.%. In some other embodiments, the glass may have a value of $\min(\text{RE}_m\text{O}_n, \text{P}_2\text{O}_5)$ less than or equal to 0.30 mol.%, less than or equal to 0.20 mol.%, less than or equal to 0.15 mol.%, or less than or equal to 0.10 mol.%. In some more embodiments, the glass may have a $\min(\text{RE}_m\text{O}_n, \text{P}_2\text{O}_5)$ greater than or equal to 0.00 mol.% and less than or equal to 0.30 mol.%, greater than or equal to 0.00 mol.% and less than or equal to 0.20 mol.%, greater than or equal to 0.00 mol.% and less than or equal to 0.15 mol.%, or greater than or equal to 0.00 mol.% and less than or equal to 0.10 mol.%, greater than or equal to 0.10 mol.% and less than or equal to 0.30 mol.%, greater than or equal to 0.10 mol.% and less than or equal to 0.20 mol.%, or greater than or equal to 0.10 mol.% and less than or equal to 0.15 mol.%.

[0071] In some embodiments, the glass may have the specific modulus E/d_{RT} from greater than or equal to 32.0 GPa·cm³/g to less than or equal to 35.1 GPa·cm³/g and all ranges and sub-ranges between the foregoing values. In some embodiments, the glass may have the E/d_{RT} greater than or equal to 32.0 GPa·cm³/g, greater than or equal to 32.5 GPa·cm³/g, greater than or equal to 33.0 GPa·cm³/g, greater than or equal to 33.9 GPa·cm³/g, greater than or equal to 34.0 GPa·cm³/g, greater than or equal to 34.5 GPa·cm³/g, greater than or equal to 34.7 GPa·cm³/g, greater than or equal to 34.9 GPa·cm³/g, or greater than or equal to 35.0 GPa·cm³/g. In some other embodiments, the glass may have the E/d_{RT} less than or equal to 35.1 GPa·cm³/g, less than or equal to 35.0 GPa·cm³/g, less than or equal to 34.9 GPa·cm³/g, less than or equal to 34.7 GPa·cm³/g, less than or equal to 34.5 GPa·cm³/g, less than or equal to 34.0 GPa·cm³/g, less than or equal to 33.0 GPa·cm³/g, or less than or equal to 32.5 GPa·cm³/g. In some more embodiments, the glass may have the E/d_{RT} greater than or equal to 32.0 GPa·cm³/g and less than or equal to 35.1 GPa·cm³/g, greater than or equal to 32.5 GPa·cm³/g and less than or equal to 35.1 GPa·cm³/g, greater than or equal to 33.0 GPa·cm³/g and less than or equal to 34.0 GPa·cm³/g.

[0072] In some embodiments, the glass may have an annealing point (An.P.) from greater than or equal to 650 °C to less than or equal to 800 °C and all ranges and sub-ranges between the foregoing values. In some embodiments, the glass may have an An.P. greater than or equal to 650 °C, greater than or equal to 680 °C, greater than or equal to 700 °C, greater than or equal to 703 °C, greater than or equal to 730 °C, greater than or equal to 740 °C, greater than or equal to 750 °C, greater than or equal to 760 °C, or greater than or equal to 780 °C. In some other embodiments, the glass may have an An.P. less than or equal to 800 °C, less than or equal to 780 °C, less than or equal to 760 °C, less than or equal to 750 °C, less than or equal to 740 °C, less than or equal to 729 °C, or less than or equal to 700 °C. In some more embodiments, the glass may have an An.P. greater than or equal to 650 °C and less than or equal to 800 °C, greater than or equal to 650 °C and less than or equal to 700 °C, greater than or equal to 700 °C and less than or equal to 729 °C, greater than or

equal to 703 °C and less than or equal to 729 °C, greater than or equal to 740 °C and less than or equal to 800 °C.

[0073] In some embodiments, the glass may have a temperature corresponding to a viscosity of 160,000 Poise (T_{160kP}) from greater than or equal to 1093 °C to less than or equal to 1300 °C and all ranges and sub-ranges between the foregoing values. In some embodiments, the glass may have the T_{160kP} greater than or equal to 1093 °C, greater than or equal to 1100 °C, greater than or equal to 1150 °C, greater than or equal to 1240 °C, greater than or equal to 1250 °C, greater than or equal to 1260 °C, or greater than or equal to 1280 °C. In some other embodiments, the glass may have the T_{160kP} less than or equal to 1300 °C, less than or equal to 1280 °C, less than or equal to 1260 °C, less than or equal to 1250 °C, less than or equal to 1240 °C, less than or equal to 1209 °C, less than or equal to 1150 °C, or less than or equal to 1100 °C. In some more embodiments, the glass may have the T_{160kP} greater than or equal to 1093 °C and less than or equal to 1300 °C, greater than or equal to 1150 °C and less than or equal to 1209 °C, greater than or equal to 1250 °C and less than or equal to 1300 °C.

[0074] In some embodiments, the glass may have a logarithm of liquidus viscosity ($\log(\eta_{liq})$), where liquidus viscosity η_{liq} is expressed in units of Poise, greater than or equal to 5. or greater than or equal to 5.2.

[0075] In some embodiments, the glass may have an average linear thermal expansion coefficient of glass in the range 20-300°C $\alpha_{20-300} \times 10^7$ less than or equal to 40 K⁻¹.

[0076] In some embodiments, the glass may have a temperature corresponding to a viscosity of 200 Poise (T_{200P}) less than or equal to 1750 °C. In some other embodiments, the glass may have a T_{200P} less than or equal to 1700 °C or less than or equal to 1650 °C and less than or equal to 1600 °C.

[0077] In some embodiments, the glass may have an aluminum-binding parameter P_{Al} greater than or equal to -2.8.

[0078] In some embodiments, the glass may have a modifier-binding parameter P_{mod} greater than or equal to -3, greater than or equal to -0.2, greater than or equal to 0.8, greater than or equal to 1.8, or greater than or equal to 2. In some other embodiments, the glass may have a modifier-binding parameter P_{mod} less than or equal to 2.8 or less than or equal to 2. In some more embodiments, the glass may have a P_{mod} greater than or equal to -3 and less than or equal to 2.8, or greater than or equal to -3 and less than or equal to 2, greater than or equal to -0.2 and less than or equal to 2.8, or greater than or equal to -0.2 and less than or equal to 2, greater than or equal to 0.8 and less than or equal to 2.8, or greater than or equal to 0.8 and less than or equal to 2, greater than or equal to 1.8 and less than or equal to 2.8.

[0079] In some other embodiments, the glass may have an anorthite precipitation parameter P_{anort} less than or equal to 10.

[0080] In some other embodiments, the glass may have a cristobalite precipitation parameter P_{crist} less than or equal to 28.

[0081] In some other embodiments, the glass may have a cordierite precipitation parameter P_{cord} less than or equal to 5.

[0082] In some other embodiments, the glass may have a spodumene precipitation parameter P_{spod} less than or equal to 7.5.

[0083] In some embodiments, the glass may have a quantity $E/d_{\text{RT}} - (92.5 - 0.05 * T_{160\text{Kp}})$ greater than or equal to 0.000.

[0084] Anorthite precipitation parameter P_{anort} is a quantity calculated by the following formula (I):

$$P_{\text{anort}} = \min(\text{CaO} + \text{SrO} + 0.5 * \text{BaO} + \text{Na}_2\text{O} + 0.5 * \text{K}_2\text{O}, \text{Al}_2\text{O}_3), \quad (\text{I})$$

where chemical formulas mean the amounts of corresponding components in the glass composition in mol.%.

[0085] The anorthite precipitation parameter P_{anort} , as well as other precipitation parameters (P_{cord} , P_{spod} and P_{crist}) described below, relate to empirically observed correlations between the glass composition and the liquidus temperature in cases when a specific crystalline phase (such as anorthite, cordierite, spodumene, cristobalite) was found to precipitate at the liquidus temperature. Without being bound to any specific theory, it is believed that these correlations are a consequence of chemical binding or association of specific oxides with each other in the glass forming melts, which, in turn, may form an atomic structure similar in structure to a crystalline phase with a tendency to precipitate from the melt.

[0086] Specifically, the anorthite precipitation parameter P_{anort} accounts for oxides that form the mineral anorthite ($\text{CaAl}_2\text{Si}_2\text{O}_8$), including oxides that can be soluble in anorthite in the solid state, thus forming solid solutions, also known as plagioclases. It is known from the literature (see, for example: G.W.MOREY, Data of Geochemistry, 6th Edition, Chapter L: Phase Equilibrium Relations of the Common Rock-Forming Oxides Except Water, United States Government Printing Office, Washington, 1964, 173 pages) and has been empirically observed in certain glasses of the present disclosure, that anorthite or plagioclase crystalline precipitates can contain (besides CaO , Al_2O_3 and/or SiO_2), sodium oxide (Na_2O), potassium oxide K_2O , strontium oxide (SrO) and/or barium oxide (BaO), so that embodiments of potential crystalline precipitates related to anorthite can be represented by the chemical formulas $\text{M}_2\text{O} * \text{Al}_2\text{O}_3 * 3\text{SiO}_2$, where M_2O refers to one or more of Na_2O and K_2O , or $\text{MO} * \text{Al}_2\text{O}_3 * 2\text{SiO}_2$, where MO refers to one or more of CaO , SrO and BaO . It has also been

empirically observed that the solubility of BaO and K₂O in these crystalline precipitates is, in most cases, somewhat less than the solubility of CaO, SrO and Na₂O.

[0087] Without being bound to any specific theory, it is hypothesized that melts of some embodiments of the aluminosilicate glasses of the present disclosure may contain structural units with compositions similar to the compositions of anorthite and related crystalline precipitates. The greater the amount of such structural units in the melt is, the more likely it is that precipitation may occur, and, accordingly, the higher the temperature at which the corresponding crystalline precipitates may exist in equilibrium with the liquid phase.

[0088] Then, in embodiments of the present disclosure in which the content of SiO₂ in the glass composition is high enough to form anorthite or a related crystalline precipitate, it is hypothesized that the amount of such structural units (M₂O*Al₂O₃*3SiO₂ and MO*Al₂O₃*2SiO₂) that form in the melt of the glass composition can be roughly correlated with the lesser of the content of alumina (Al₂O₃) in mol.% and the total content of M₂O+MO in mol.% in the glass composition, i.e. as min(M₂O+MO, Al₂O₃), where chemical formulas and abbreviations mean the content of corresponding oxides in glass in terms of mol.%.

[0089] Regarding the limited solubility of K₂O and BaO relative to CaO, SrO and Na₂O in anorthite and related crystalline precipitates as noted above, an empirical factor of 0.5 was applied to the amounts of K₂O and BaO when formulating the anorthite precipitation parameter P_{anort}. Accordingly, the sum (M₂O+MO), as stated above in the context of precipitation of anorthite and related crystalline phases, can be expressed as

$$M_2O+MO \approx (CaO+SrO+0.5*BaO+Na_2O+0.5*K_2O) \quad (II).$$

[0090] Finally, the lesser of the sum (M₂O+MO), as calculated by the formula (II), and Al₂O₃ gives an estimate of the amount of structural units that, when present in a glass forming melt, may precipitate as anorthite or a related crystalline phase (e.g. plagioclase).

[0091] Based on the above reasoning, it is hypothesized that a higher value of the parameter P_{anort} may indicate a higher probability of precipitating anorthite or related crystalline phases (e.g. plagioclase) at high temperatures, thus causing higher liquidus temperature. This means that in some embodiments of the present disclosure, to reduce the liquidus temperature of glass, it is preferable that the value of P_{anort} is small.

[0092] Cordierite precipitation parameter P_{cord} is a quantity calculated by the following formula (III):

$$P_{cord} = MgO+MnO+FeO, \quad (III)$$

where chemical formulas mean the amounts of corresponding components in the glass composition in mol.%.

[0093] The cordierite precipitation parameter P_{cord} is designed to predict the probability of precipitation of cordierite and related crystalline phases from the glass forming melts of the present disclosure at high temperatures.

[0094] The chemical formula of cordierite is commonly known as $2\text{MgO} \cdot 2\text{Al}_2\text{O}_3 \cdot 5\text{SiO}_2$, or $\text{Mg}_2\text{Al}_4\text{Si}_5\text{O}_{18}$. However, it has been empirically observed that when precipitating from the glass forming melts, variations of cordierite in which MnO and/or FeO isostructurally replace MgO (in whole or in part) can form. Accordingly, assuming the amount of alumina and silica in the glass composition of the is sufficient to bind all MgO, MnO and FeO, the value of P_{cord} , as an estimate for predicting the tendency of precipitation of cordierite and related crystalline phases, can be presented as a sum of the concentrations of MgO, MnO, and FeO in the glass composition in mol.%, which gives the above-presented formula (III).

[0095] It is hypothesized that if the value of P_{cord} is high, cordierite (or a related crystalline phase) may possibly precipitate from the glass forming melts at high temperatures. Accordingly, in some embodiments of the present disclosure, the value of P_{cord} is preferably limited.

[0096] Spodumene precipitation parameter P_{spod} is a quantity calculated by the following formula (IV):

$$P_{\text{spod}} = \min(\text{Li}_2\text{O}, \text{Al}_2\text{O}_3 - \text{K}_2\text{O} - 0.5 \cdot \text{Na}_2\text{O}), \quad (\text{IV})$$

where chemical formulas mean the amounts of corresponding components in mol.% in the glass composition.

[0097] The spodumene precipitation parameter P_{spod} is designed to estimate the probability of precipitation of spodumene ($\text{LiAlSi}_2\text{O}_6$, or $\text{Li}_2\text{O} \cdot \text{Al}_2\text{O}_3 \cdot 4\text{SiO}_2$) from the glass forming melts at high temperatures.

[0098] Without being bound to a specific theory, it is hypothesized that the probability of precipitating spodumene from the glass forming melts at high temperatures may correlate to the amount of structural units with atomic compositions similar to that of spodumene. If the amount of silica in an embodiment of a glass composition of the present disclosure is sufficient for the formation of such structural units, the amount of such structural units in the melt can be roughly correlated to the lesser of the concentrations of Li_2O and Al_2O_3 , in mol.%, in the glass composition, mathematically expressed as

$$P_{\text{spod}} \approx \min(\text{Li}_2\text{O}, \text{Al}_2\text{O}_3). \quad (\text{V})$$

[0099] In addition, it has been empirically observed that if the glass composition contains significant amounts of Na_2O and, especially, K_2O , precipitation of spodumene may be suppressed. Without being bound to a specific theory, it is hypothesized that this effect may be caused by chemical binding of the alumina in the glass forming melt with sodium and potassium oxides, thus making the

bound alumina inactive in the reaction with Li_2O and suppressing the precipitation of spodumene. To approximately account for this effect, the concentration of K_2O and half of the concentration of Na_2O are deducted from the concentration of Al_2O_3 when estimating the tendency of spodumene to precipitate from the melt. The term " Al_2O_3 " in the formula (V) above is thus replaced by $(\text{Al}_2\text{O}_3 - \text{K}_2\text{O} - 0.5*\text{Na}_2\text{O})$, where the chemical formulas of oxides mean the concentrations of these oxides in the glass composition in mol.%. After this substitution, formula (V) is modified to give formula (IV) above for P_{spod} .

[00100] It is hypothesized that if the value of P_{spod} is high, spodumene may possibly precipitate from the glass forming melts at high temperatures. Accordingly, in some embodiments of the present disclosure, the value of P_{spod} is preferably limited.

[00101] Cristobalite precipitation parameter P_{crist} is a quantity calculated by the following formula (VI):

$$P_{\text{crist}} = \text{SiO}_2 - 6*(\text{Na}_2\text{O} + \text{K}_2\text{O}) - 4*\text{Li}_2\text{O} - 2*(\text{CaO} + \text{SrO} + \text{BaO}) - 2.5*\text{MgO}, \quad (\text{VI})$$

where chemical formulas mean the amounts of corresponding components in the glass composition in mol.%.

[00102] The cristobalite precipitation parameter P_{crist} is introduced to estimate the probability of precipitation of the mineral cristobalite (SiO_2) from the glass forming melts of the present disclosure at high temperatures.

[00103] Without being bound to a specific theory, it is hypothesized that the amount of SiO_2 in the glass composition that may potentially precipitate in the form of cristobalite correlates to the residual amount of silica remaining after binding with other oxides (i.e. alumina, alkali metal oxides and alkaline earth metal oxides) in structural units with compositions of other crystalline phases having a tendency to precipitate. Such crystalline phases include $\text{Li}_2\text{O}*\text{Al}_2\text{O}_3*4\text{SiO}_2$ (spodumene), $\text{Na}_2\text{O}*\text{Al}_2\text{O}_3*6\text{SiO}_2$ (albite), $\text{K}_2\text{O}*\text{Al}_2\text{O}_3*6\text{SiO}_2$ (K-feldspar), $\text{MgO}*\text{Al}_2\text{O}_3*2.5\text{SiO}_2$ (cordierite, which can also be presented as $2\text{MgO}*\text{Al}_2\text{O}_3*5\text{SiO}_2$), $\text{CaO}*\text{Al}_2\text{O}_3*2\text{SiO}_2$ (anorthite), $\text{SrO}*\text{Al}_2\text{O}_3*2\text{SiO}_2$ (Sr-feldspar) and $\text{BaO}*\text{Al}_2\text{O}_3*2\text{SiO}_2$ (celsian).

[00104] If the amount of silica and alumina in the glass composition is sufficient for forming all of the listed crystalline phases, the residual amount of silica can be obtained by deducting the amounts of alkali and alkaline earth metal oxides, multiplied by the stoichiometric coefficients for silica in the formulas of the listed crystalline phases ($\text{Li}_2\text{O}*\text{Al}_2\text{O}_3*4\text{SiO}_2$ for spodumene, $\text{Na}_2\text{O}*\text{Al}_2\text{O}_3*6\text{SiO}_2$ for albite, $\text{CaO}*\text{Al}_2\text{O}_3*2\text{SiO}_2$ for anorthite, etc.), from the total amount of silica in the glass composition in mol.%. Thus, the residual amount of silica can be evaluated by the formula (VI) above.

[00105] It is hypothesized that if the value of P_{crist} is high, cristobalite may possibly precipitate from the glass forming melts at high temperatures. Accordingly, in some embodiments of the present disclosure, the value of P_{crist} is preferably limited.

[00106] The modifier-binding parameter P_{mod} is a quantity calculated by formula (VII):

$$P_{\text{mod}} = R_2O + RO - Al_2O_3 - P_2O_5 - RE_mO_n \quad (\text{VII})$$

where R_2O is total sum of monovalent metal oxides, RO is total sum of divalent metal oxides, RE_mO_n is total sum of rare earth metal oxides, and chemical formulas mean the amounts of corresponding components in the glass composition in mol.%.

[00107] The modifier-binding parameter P_{mod} evaluates the total amount of modifiers (monovalent metal oxides R_2O and divalent metal oxides RO) that can be chemically bounded with the network formers - alumina (Al_2O_3), phosphorus oxide (P_2O_5) and rare earth metal oxides (RE_mO_n) – of the glass compositions disclosed herein. Assuming one mole of modifier (R_2O and RO) binds with one mole of network former, the difference between the total content of modifiers and the total content of network formers gives the value of P_{mod} as presented in the formula (VII) above.

[00108] Without being bound to a specific theory, the authors believe that if the parameter P_{mod} has a large positive value, it may mean that the glass structure may contain a significant amount of modifiers that may form non-bridging oxygen atoms, which, in turn, may cause a decrease in the low-temperature viscosity and, in particular, a decrease in the annealing and strain points, which may reduce the possibility of using the glass as a material for the memory recording media suitable to the HAMR process as specified above.

[00109] Accordingly, in some embodiments of the present disclosure, the glass compositions are preferably characterized by negative, zero or small positive values of P_{mod} .

[00110] The aluminum-binding parameter P_{Al} is a quantity calculated by formula (VIII):

$$P_{\text{Al}} = R_2O + RO + P_2O_5 + 1.6 * RE_mO_n - Al_2O_3, \quad (\text{VIII})$$

where R_2O is total sum of monovalent metal oxides, RO is total sum of divalent metal oxides, RE_mO_n is total sum of rare earth metal oxides, and chemical formulas mean the amounts of corresponding components in the glass composition in mol.%.

[00111] The aluminum-binding parameter P_{Al} estimates the amount of alumina that can be chemically bounded with other oxides, such as modifiers (R_2O and RO), phosphorus oxide (P_2O_5) and rare earth metal oxides (RE_mO_n). Without being bound to a specific theory, it is hypothesized that one mole of alumina can possibly react or bind with either one mole of modifiers, one mole of P_2O_5 , or 1.6 moles of RE_mO_n (the figure for RE_mO_n is selected on the basis of known compositions of rare earth aluminates, such as yttrium alumina garnet (YAG, $Y_3Al_5O_{12}$), lanthanum aluminate ($La_3Al_5O_{12}$) and others). Under this hypothesis, the difference between the total sum of the oxides in mol.%

(weighted by the stoichiometric coefficients noted above) and the concentration of alumina in mol.% in the glass composition characterizes the lack or excess of alumina comparing to the oxides.

[00112] Without being bound to a specific theory, it is hypothesized that an excess of alumina relative to the above-enumerated oxides (R_2O , RO , P_2O_5 , RE_mO_n) may potentially cause precipitation of alumina or alumina-rich crystalline phases, such as, for example, beta-alumina ($X \cdot 6Al_2O_3$, where "X" refers to one or more of monovalent or divalent metal oxides), mullite ($3Al_2O_3 \cdot 2SiO_2$), corundum (Al_2O_3) and others, from the glass forming melts at high temperatures, thus increasing the liquidus temperature and reducing the liquidus viscosity. Accordingly, in some embodiments of the present disclosure, the glass compositions are preferably characterized by negative, zero or small positive values of P_{Al} .

[00113] The annealing point $An.P.$, specific modulus E/d_{RT} and T_{160kP} are properties of glass that can be predicted from the glass composition. A linear regression analysis of the Exemplary Glasses of the present disclosure in the EXAMPLES section below and other glass compositions reported in the literature was performed to determine equations that can predict the composition dependences of the annealing point, specific modulus, and temperature at which the glass has a viscosity of 160 kP.

[00114] The training dataset of glass compositions satisfying the criteria specified in Table 1 below and having measured values of the properties of interest, about 100 glass compositions for each property ($(An.P., E/d_{RT} \text{ and } T_{160kP})$), was randomly selected from the literature data presented in the publicly available SciGlass Information System database and from the Exemplary Glasses from the embodiments presented herein. The linear regression analysis on the above-specified dataset (excluding outliers) was used to determine the formulas (IX), (X), and (XI) below for the predictive parameters P_{anpt} , P_{spm} , and P_{160kP} . In Table 1, RO is a total sum of divalent metal oxides, R_2O is a total sum of monovalent metal oxides and RE_mO_n is a total sum of rare earth metal oxides.

[00115] Another set of glass compositions satisfying the criteria of Table 1 was used as a validation set to evaluate the ability to interpolate within predefined compositional limits. Based on the validation analysis, interpolation to within standard deviations specified in the Table 2 was possible. An external dataset of prior art glass compositions, also randomly selected from the SciGlass Information System database, was used to evaluate the ability to predict the properties outside of the specified compositional limits with a reasonable accuracy. Multiple iterations of this process were performed in order to determine the best variant for each property, corresponding to the regression formulas for the parameters specified in Table 2.

[00116] The data for the Comparative Glass compositions used in the linear regression modeling, including the training dataset, validation dataset and external dataset were obtained from

the publically available SciGlass Information System database. Formulas (IX), (X) and (XI) below were obtained from the linear regression analysis and used to predict the annealing point, specific modulus, and temperature at which the viscosity is 160 kP, respectively, of the glasses:

$$P_{\text{anpt}} = 664.7 + 5.2303 * \text{SiO}_2 - 11.493 * \text{B}_2\text{O}_3 - 7.1742 * \text{P}_2\text{O}_5 + 8.3980 * \text{ZrO}_2 - 2.0585 * \text{MgO} - 2.1088 * \text{CaO} - 3.8995 * \text{BaO} - 10.323 * \text{ZnO} - 9.0727 * \text{MnO} - 23.455 * \text{Li}_2\text{O} - 33.819 * \text{Na}_2\text{O} - 25.204 * \text{K}_2\text{O} + 15.745 * \text{Y}_2\text{O}_3 + 8.9047 * \text{La}_2\text{O}_3 - 33.960 * (\text{Fe}_2\text{O}_3 + \text{FeO}) - 5.6704 * (\text{R}_2\text{O} + \text{RO} - \text{Al}_2\text{O}_3) - 4.2545 * (\text{SiO}_2 - (6 * \text{K}_2\text{O} + 6 * \text{Na}_2\text{O} + 4 * \text{Li}_2\text{O} + 2 * \text{RO})) - 19.439 * \text{Cu}_2\text{O}, \quad (\text{IX})$$

$$P_{\text{spm}} = 32.10 + 0.47744 * \text{SiO}_2 - 1.6506 * \text{Al}_2\text{O}_3 - 0.11775 * \text{B}_2\text{O}_3 - 0.30166 * \text{P}_2\text{O}_5 + 0.20187 * \text{TiO}_2 + 0.20219 * \text{ZrO}_2 + 0.96268 * \text{MgO} + 0.83379 * \text{CaO} + 0.53685 * \text{SrO} + 0.41218 * \text{BaO} + 0.63264 * \text{ZnO} + 0.75365 * \text{MnO} + 0.62984 * \text{CuO} - 1.2496 * \text{Na}_2\text{O} - 1.5154 * \text{K}_2\text{O} + 1.4746 * \text{Cu}_2\text{O} - 0.037941 * \text{Y}_2\text{O}_3 - 0.75836 * \text{La}_2\text{O}_3 - 1.8052 * (\text{R}_2\text{O} + \text{RO} - \text{Al}_2\text{O}_3) - 0.47488 * (\text{SiO}_2 - (6 * \text{K}_2\text{O} + 6 * \text{Na}_2\text{O} + 4 * \text{Li}_2\text{O} + 2 * \text{RO})), \quad (\text{X})$$

$$P_{160\text{kP}} = 1058 + 2.5492 * \text{SiO}_2 - 25.725 * \text{Al}_2\text{O}_3 - 11.327 * \text{B}_2\text{O}_3 - 10.014 * \text{P}_2\text{O}_5 - 14.309 * \text{TiO}_2 - 11.594 * \text{ZrO}_2 + 30.559 * \text{MgO} + 29.290 * \text{CaO} + 30.592 * \text{SrO} + 30.079 * \text{BaO} + 23.323 * \text{ZnO} + 19.724 * \text{MnO} + 11.888 * \text{PbO} + 11.462 * \text{CuO} + 17.090 * \text{Li}_2\text{O} + 16.475 * \text{Na}_2\text{O} + 11.386 * \text{K}_2\text{O} + 14.422 * \text{Y}_2\text{O}_3 - 36.909 * \text{La}_2\text{O}_3 - 34.144 * (\text{Fe}_2\text{O}_3 + \text{FeO}) - 35.001 * (\text{R}_2\text{O} + \text{RO} - \text{Al}_2\text{O}_3). \quad (\text{XI})$$

[00117] In Formulas (IX), (X) and (XI) and Tables 1 and 2, annealing point parameter P_{anpt} is a parameter that predicts the annealing point [°C], calculated from the components of the glass composition expressed in mol.%; specific modulus parameter P_{spm} is a parameter that predicts the specific modulus E/d_{RT} [GPa·cm³/g], calculated from the components of the glass composition expressed in mol.%; and the parameter $P_{160\text{kP}}$ is a parameter that predicts the glass temperature corresponding to a viscosity of 160 kP [°C], calculated from the components of the glass composition expressed in mol.%.

[00118] In Formulas (IX), (X) and (XI), each component of the glass composition is listed in terms of its chemical formula, where the chemical formula refers to the concentration of the component expressed in mol.%. For example, for purposes of Formulas (IX), (X) and (XI), SiO_2 refers to the concentration of SiO_2 , expressed in mol.%, in the glass composition. It is understood that not all components listed in Formulas (IX), (X) and (XI) are necessarily present in a particular glass composition and that Formulas (IX), (X) and (XI) are equally valid for glass compositions that contain less than all of the components listed in the formulas. It is further understood that Formulas (IX), (X) and (XI) are also valid for glass compositions within the scope and claims of the present disclosure that contain components in addition to the components listed in the formulas. If a component listed in Formulas (IX), (X) and (XI) is absent in a particular glass composition, the concentration of the component in the glass composition is 0 mol.% and the contribution of the component to the value calculated from the formulas is zero.

Table 1. Composition Space Used for Modeling

Property	An.P., °C		E/d _{RT} , GPa·cm ³ /g		T _{160kP} , °C	
Component	Min, mol.%	Max, mol.%	Min, mol.%	Max, mol.%	Min, mol.%	Max, mol.%
SiO ₂	60	75	60	75	60	75
Al ₂ O ₃	10	20	10	20	10	20
Li ₂ O	0	10	0	10	0	10
Na ₂ O	0	10	0	10	0	10
K ₂ O	0	5	0	5	0	5
BaO	0	5	0	5	0	5
B ₂ O ₃	0	3	0	3	0	3
La ₂ O ₃	0	3	0	3	0	3
RO	3	Not limited	3	Not limited	3	Not limited
R ₂ O	2.5	Not limited	2.5	Not limited	2.5	Not limited
F	0	0.05 [at.%]	0	1 [at.%]	0	1 [at.%]
ZrO ₂ + TiO ₂ + FeO + Fe ₂ O ₃	0	5	0	5	0	5
R ₂ O + RO + P ₂ O ₅ + 1.6 * RE _m O _n - Al ₂ O ₃	-5	Not limited	-5	Not limited	-5	Not limited
R ₂ O + RO - Al ₂ O ₃ - P ₂ O ₅ - RE _m O _n	Not limited	5	Not limited	5	Not limited	5
R ₂ O - Al ₂ O ₃	Not limited	0	Not limited	0	Not limited	0
Other species	0	Not limited	0	Not limited	0	Not limited

Table 2. Property prediction models

Property	Abbreviation	Unit	Predicting Parameter	Regression Formula	Composition Unit	Standard error
Annealing point	An.P.	°C	P _{anpt}	Formula (IX)	mol.%	22
Specific modulus	E/d _{RT}	GPa·cm ³ /g	P _{spm}	Formula (X)	mol.%	0.32
Temperature corresponding to a viscosity of 160 kP	T _{160kP}	°C	P _{160kP}	Formula (XI)	mol.%	37

[00119] FIG. 1 is a plot of the parameter P_{anpt} calculated by Formula (IX) as a function of measured annealing point An.P. for some Literature Glasses ("Comp. Glasses") and some Exemplary Glasses ("Ex. Glasses"). As illustrated by the data in FIG. 1, the compositional dependence of the parameter P_{anpt} had an error within a range of ± 22 °C of the measured An.P. for the majority of glasses, which corresponds to the standard error specified in Table 2.

[00120] FIG. 2 is a plot of the parameter P_{spm} calculated by Formula (X) as a function of measured specific modulus E/d_{RT} for some Literature Glasses ("Comp. Glasses") and some Exemplary

Glasses ("Ex. Glasses"). As illustrated by the data in FIG. 2, the compositional dependence of the parameter P_{spm} had an error within a range of $\pm 0.32 \text{ GPa}\cdot\text{cm}^3/\text{g}$ of the measured E/d for the majority of glasses, which corresponds to the standard error specified in Table 2.

[00121] FIG. 3 is a plot of the parameter $P_{160\text{kP}}$ calculated by Formula (XI) as a function of the measured temperature $T_{160\text{kP}}$ at which the glass has a viscosity of 160 kP for some Literature Glasses ("Comp. Glasses") and some Exemplary Glasses ("Ex. Glasses"). As illustrated by the data in FIG. 3, the compositional dependence of the parameter $P_{160\text{kP}}$ had an error within a range of $\pm 37^\circ\text{C}$ of the measured $T_{160\text{kP}}$ for the majority of glasses, which corresponds to the standard error specified in Table 2.

[00122] Table 3 identifies the combination of components and their respective amounts according to some embodiments of the present disclosure. The Exemplary Glasses A in Table 3 may include additional components according to any aspects of the present disclosure as described herein.

Table 3: Exemplary Glasses A

Component	Amount (mol.%)
SiO_2	60.0 to 75.0 mol.%
Al_2O_3	10.0 to 18.0 mol.%
Li_2O	0.0 to 10.0 mol.%
Na_2O	0.0 to 5.0 mol.%
Sum of ($\text{CaO}+\text{MgO}$)	≤ 5.0 mol.%
Sum of ($\text{Li}_2\text{O}+\text{Na}_2\text{O}$)	≤ 0.5 mol.%
Sum of ($\text{Li}_2\text{O}+\text{MgO}$)	0.0 to 10.0 mol.%
Sum of ($\text{MgO}+\text{ZnO}$)	0.0 to 10.0 mol.%
Sum of ($\text{CaO}+\text{SrO}$)	0.0 to 9.0 mol.%

[00123] Exemplary Glasses A according to embodiments of the present disclosure may optionally fluorine (F) in an amount 0.0 to 3.0 at.%.

[00124] According to some embodiments of the present disclosure, Exemplary Glasses A may also have an aluminum-binding parameter P_{Al} of greater than or equal to -2.8.

[00125] According to some embodiments of the present disclosure, Exemplary Glasses A may also have a modifier-binding parameter P_{mod} of less than or equal to 2.8.

[00126] According to some embodiments of the present disclosure, Exemplary Glasses A may also have an anorthite precipitation parameter P_{anort} of less than or equal to 10.

[00127] Table 4 identifies the combination of components and their respective amounts according to some embodiments of the present disclosure. The Exemplary Glasses B in Table 4 may include additional components according to any aspects of the present disclosure as described herein.

Table 4: Exemplary Glasses B

Component	Amount (mol.%)
SiO ₂	60.0 to 80.0 mol.%
Al ₂ O ₃	10.0 to 18.0 mol.%
Li ₂ O	0.5 to 7.5 mol.%

[00128] Exemplary Glasses B according to embodiments of the present disclosure may optionally fluorine (F) in an amount 0.0 to 0.5 at.%,

[00129] According to some embodiments of the present disclosure, Exemplary Glasses B may also satisfy the following condition:

$$0.00 \leq \min(\text{RE}_m\text{O}_n, \text{P}_2\text{O}_5) \leq 0.30$$

where chemical formulas refer to the amounts of components in glass, expressed in mol.%.

[00130] According to some embodiments of the present disclosure, Exemplary Glasses B may also have a cristobalite precipitation parameter P_{crist} of less than or equal to 28.

[00131] According to some embodiments of the present disclosure, Exemplary Glasses B may also have an anorthite precipitation parameter P_{anort} of less than or equal to 10.

[00132] According to some embodiments of the present disclosure, Exemplary Glasses B may also have a cordierite precipitation parameter P_{cord} of less than or equal to 5.

[00133] According to some embodiments of the present disclosure, Exemplary Glasses B may also have a spodumene precipitation parameter P_{spod} of less than or equal to 7.5.

[00134] According to some embodiments of the present disclosure, Exemplary Glasses B may also have a specific modulus E/d_{RT} of greater than or equal to 32 GPa·cm³/g.

[00135] According to some embodiments of the present disclosure, Exemplary Glasses B may also have an annealing point An.P. of greater than or equal to 680 °C.

[00136] According to some embodiments of the present disclosure, Exemplary Glasses B may also have a temperature $T_{160\text{kPa}}$ corresponding to a viscosity of 160 kPa of greater than or equal to 1150 °C.

[00137] Table 5 identifies the combination of components and their respective amounts according to some embodiments of the present disclosure. The Exemplary Glasses C in Table 5 may include additional components according to any aspects of the present disclosure as described herein.

Table 5: Exemplary Glasses C

Component	Amount (mol.%)
SiO ₂	60.0 to 75.0 mol.%
Al ₂ O ₃	10.5 to 18.0 mol.%
CaO	0.0 to 10.0 mol.%
Li ₂ O	0.0 to 7.8 mol.%
MgO	0.0 to 4.3 mol.%

Component	Amount (mol.%)
ZrO ₂	0.0 to 0.5 mol.%
Total sum of rare earth metal oxides RE _m O _n	0.0 to 5.0 mol.%

[00138] Exemplary Glasses C according to embodiments of the present disclosure may have a cristobalite precipitation parameter P_{crist} of less than or equal to 28.

[00139] According to some embodiments of the present disclosure, Exemplary Glasses C may also have an anorthite precipitation parameter P_{anort} of less than or equal to 10.

[00140] According to some embodiments of the present disclosure, Exemplary Glasses C may also have a modifier-binding parameter P_{mod} of greater than or equal to -3.

[00141] According to some embodiments of the present disclosure, Exemplary Glasses C may also satisfy the following formula:

$$E/d_{\text{RT}} - (92.5 - 0.05 * T_{160\text{kP}}) > 0.000,$$

where E/d_{RT} is a specific modulus and $T_{160\text{kP}}$ is a temperature corresponding to a glass viscosity of 160 kP.

EXAMPLES

[00142] The following examples describe various features and advantages provided by the disclosure, and are in no way intended to limit the invention and appended claims.

[00143] Glass compositions were prepared and analyzed. The analyzed glass compositions included the components listed in Table 6 below. Example glasses were melted with conventional raw materials, such as sand, aluminum oxide, alkali carbonates, alkali nitrates, spodumene, nepheline syenite, borax, boric acid, aluminum metaphosphate, disodium phosphate, magnesium oxide, rare earth metal oxides, tin oxide, and various combinations. The glasses were melted in platinum crucibles between 1500 °C and 1575 °C for 5 to 6 hours, poured as a relatively thin stream into a water bath to quench the molten glass into small particles (drigaged) and then re-melted at a higher temperature of 1650 °C for 5 to 6 hours to improve homogeneity and melt quality. The glasses were then cast onto a steel plate, annealed for 1 hour near the anneal temperatures given in Table 6 below, and cooled to room temperature to form final glass articles.

Table 6. Exemplary Glass Compositions

Exemplary Glass		1	2	3	4	5	6	7	8
Composition - mol. %									
SiO ₂	mol. %	65.22	65.41	64.53	63.99	66.66	69.12	65.87	65.83
Al ₂ O ₃	mol. %	17.41	17.43	17.90	17.87	16.89	15.95	17.87	17.87
Li ₂ O	mol. %	5.94	6.18	6.27	4.87	5.37	5.85	6.64	6.54
CaO	mol. %	1.99	4.02	4.01	3.98	3.98	4.01	3.01	4.02
MgO	mol. %	3.97	1.98	1.98	1.98	1.98	1.98	0.97	0.09
Na ₂ O	mol. %	3.36	3.48	3.42	4.61	3.68	2.73	3.44	3.46
P ₂ O ₅	mol. %	0.85	1.14	1.53	2.34	1.08	0	1.84	1.83
B ₂ O ₃	mol. %	0.90	0	0	0	0	0	0	0
K ₂ O	mol. %	0.20	0.20	0.20	0.20	0.20	0.20	0.20	0.20
SnO ₂	mol. %	0.15	0.15	0.15	0.15	0.15	0.15	0.15	0.15
Fe ₂ O ₃	mol. %	0.01	0.01	0.01	0.01	0.01	0.01	0.01	0.01
Composition constraints									
CaO + SrO	mol. %	1.990	4.020	4.010	3.979	3.980	4.010	3.010	4.020
CaO + MgO	mol. %	5.961	6.000	5.990	5.960	5.959	5.990	3.980	4.110
Li ₂ O + Na ₂ O	mol. %	9.300	9.659	9.689	9.478	9.051	8.580	10.08	10.00
Li ₂ O + MgO	mol. %	9.911	8.159	8.249	6.849	7.350	7.829	7.610	6.629
MgO + ZnO	mol. %	3.970	1.980	1.980	1.981	1.979	1.980	0.9700	0.09000
ZrO ₂ + TiO ₂ + FeO + Fe ₂ O ₃	mol. %	0.010000	1.000E-02	9.998E-03	9.997E-03	1.000E-02	0.010001	9.999E-03	0.010001
Measured properties									
$\alpha_{20-300} \times 10^7$	K ⁻¹	47.100	50.400	50.000	50.100	47.800	46.200	49.500	50.200
Str.P.	°C	651	664	661	664	671	679	665	673
An.P.	°C	699	712	710	713	721	728	716	723
T _{soft}	°C	930	942	939	948	956	967	955	960
T _{160kP}	°C	1107	1124	1123	1137	1140	1151	1143	1151
T _{35kP}	°C	1181	1198	1198	1214	1216	1228	1220	1228
T _{200 P}	°C	1571	1590	1591	1611	1615	1630	1624	1628
T _{liq}	°C	>1350	>1305	>1300	1335	1310	1295	>1350	1345
log(η_{liq})	P	3.38	3.76	3.81	3.69	3.86	4.05	3.64	3.72
E	GPa	85.640	85.290	85.080	83.020	84.670	85.840	83.500	83.150
G	GPa	34.890	34.890	34.750	33.990	34.680	35.230	34.270	34.200
E/d _R T	GPa·cm ³ /g	35.000	34.800	34.700	33.900	34.600	35.100	34.300	34.200
d _R T	g/cm ³	2.444	2.452	2.453	2.447	2.447	2.443	2.431	2.434
Predicted and calculated properties									
P _{anort}		5.5833	7.6455	7.6252	8.8092	7.8656	6.9109	6.6386	7.6568
P _{cord}		3.9877	1.9934	1.9938	1.9939	1.9943	1.9941	0.99785	0.11337
P _{spod}		5.9599	6.2501	6.3182	4.7043	5.3023	5.8515	6.7991	6.6516
P _{crist}		5.1090	4.7247	3.6404	2.5914	8.4910	14.524	7.3384	8.0376
P _{mod}		-2.6835	-2.6250	-3.4590	-4.7228	-2.7830	-1.0930	-5.2663	-5.2561
P _{Al}		-0.94912	-0.29191	-0.34876	0.04274	-0.58979	-1.0930	-1.5574	-1.5515
P _{spm} [for E/d _R T]	GPa·cm ³ /g	34.74	34.51	34.47	33.83	34.32	34.75	34.35	34.21
P _{anpt} [for An.P.]	°C	693.3	696.8	695.3	692.0	704.1	715.5	686.9	688.1
P _{160kP} [for T _{160kP}]	°C	1162	1160	1157	1155	1172	1188	1159	1159
P _{spm} - (92.5 - 0.05 * P _{160kP})		0.4714	0.1683	0.0318	-0.9212	0.4787	1.7666	0.0406	-0.1057

Table 6 Continued

Exemplary Glass		9	10	11	12	13	14	15	16
Composition - mol.%									
SiO ₂	mol.%	67.09	64.37	67.14	64.47	63.75	63.89	63.94	63.93
Al ₂ O ₃	mol.%	16.85	17.84	16.81	17.80	18.04	17.96	17.97	17.99
Li ₂ O	mol.%	6.01	5.59	5.80	5.38	7.12	7.12	7.10	7.08
CaO	mol.%	4.00	4.01	6.00	6.00	4.03	3.00	2.00	1.00
MgO	mol.%	1.97	1.98	0.12	0.12	0.98	1.98	2.97	3.96
Na ₂ O	mol.%	3.03	3.93	3.09	3.96	2.11	2.12	2.10	2.10
P ₂ O ₅	mol.%	0.69	1.92	0.68	1.91	3.71	3.67	3.66	3.68
K ₂ O	mol.%	0.20	0.20	0.20	0.20	0.10	0.10	0.10	0.10
SnO ₂	mol.%	0.15	0.15	0.15	0.15	0.15	0.15	0.15	0.15
Fe ₂ O ₃	mol.%	0.01	0.01	0.01	0.01	0.01	0.01	0.01	0.01
Composition constraints									
CaO + SrO	mol.%	4.000	4.010	6.000	6.001	4.029	3.000	2.000	0.9999
CaO + MgO	mol.%	5.970	5.990	6.120	6.121	5.009	4.981	4.970	4.960
Li ₂ O + Na ₂ O	mol.%	9.039	9.520	8.890	9.341	9.230	9.239	9.199	9.179
Li ₂ O + MgO	mol.%	7.979	7.570	5.919	5.501	8.100	9.100	10.07	11.04
MgO + ZnO	mol.%	1.970	1.980	0.1200	0.1200	0.9799	1.981	2.970	3.960
ZrO ₂ + TiO ₂ + FeO + Fe ₂ O ₃	mol.%	1.000E-02	9.998E-03	0.010001	0.010001	9.999E-03	9.997E-03	1.000E-02	9.999E-03
Measured properties									
$\alpha_{20-300} \times 10^7$	K ⁻¹	47.300	49.200	49.400	51.300	45.700	44.600	43.700	42.200
Str.P.	°C	670	663	688	678	659	656	653	652
An.P.	°C	720	712	737	728	710	706	705	703
T _{soft}	°C	955	947	967	961	947	946	945	943
T _{160kP}	°C	1134	1136	1144	1137	1137		1133	1127
T _{35kP}	°C	1209	1212	1219	1211	1214	1212	1209	1203
T _{200 P}	°C	1607	1604	1616	1599	1612	1609	1602	1597
T _{liq}	°C	1315	1325	1300	1300	1280	>1320	>1335	>1300
log(η_{liq})	P	3.78	3.73	3.94	3.88	4.05		3.65	3.85
E	GPa	85.640	84.260	84.670	83.360	82.400	82.600	83.080	83.290
G	GPa	35.030	34.480	34.680	34.060	33.850	33.920	33.990	34.270
E/d _{RT}	GPa·cm ³ /g	35.000	34.400	34.500	33.900	34.000	34.100	34.400	34.500
d _{RT}	g/cm ³	2.448	2.451	2.453	2.456	2.423	2.420	2.417	2.415
Predicted and calculated properties									
P _{anort}		7.2275	8.1340	9.2570	10.163	6.2238	5.2277	4.2299	3.2340
P _{cord}		1.9937	1.9950	0.15720	0.15605	0.99683	1.9942	2.9919	3.9882
P _{spod}		6.0307	5.4717	5.7513	5.2056	6.9861	6.9860	6.9881	6.9875
P _{crist}		9.7140	3.9111	11.308	5.4590	11.675	11.171	10.667	10.167
P _{mod}		-2.2927	-4.1964	-2.3415	-4.2330	-7.5157	-7.5129	-7.5094	-7.5136
P _{Al}		-0.89687	-0.28859	-0.96807	-0.35083	-0.0591	-0.05621	-0.05267	-0.05724
P _{spm} [for E/d _{RT}]	GPa·cm ³ /g	34.64	34.17	34.37	33.89	34.11	34.24	34.37	34.50
P _{anpt} [for An.P.]	°C	707.0	695.4	710.1	698.1	690.3	690.2	690.2	690.5
P _{160kP} [for T _{160kP}]	°C	1175	1159	1177	1160	1155	1156	1157	1159
P _{spm} - (92.5 - 0.05 * P _{160kP})		1.0443	-0.3727	0.7290	-0.6402	-0.6555	-0.4635	-0.2129	-0.0080

Table 6 Continued

Exemplary Glass		17	18	19	20	21	22	23	24
Composition - mol.%									
SiO ₂	mol.%	64.94	64.11	74.16	74.46	74.40	74.47	74.53	74.36
Al ₂ O ₃	mol.%	17.64	17.96	12.34	12.39	12.41	12.39	12.41	12.40
Li ₂ O	mol.%	7.13	6.41	4.12	4.06	3.97	3.07	3.04	3.10
CaO	mol.%	3.99	3.99	4.12	2.99	2.02	3.99	2.98	2.01
MgO	mol.%	0.98	2.42	2.02	2.94	4.01	1.96	2.93	3.97
Na ₂ O	mol.%	2.11	3.25	2.95	2.88	2.91	3.84	3.83	3.88
P ₂ O ₅	mol.%	2.95	1.50	0	0	0	0	0	0
K ₂ O	mol.%	0.10	0.20	0.14	0.13	0.14	0.13	0.13	0.13
SnO ₂	mol.%	0.15	0.15	0.14	0.14	0.13	0.14	0.14	0.14
Fe ₂ O ₃	mol.%	0.01	0.01	0.01	0.01	0.01	0.01	0.01	0.01
Composition constraints									
CaO + SrO	mol.%	3.990	3.990	4.120	2.990	2.021	3.990	2.981	2.010
CaO + MgO	mol.%	4.970	6.411	6.141	5.931	6.031	5.949	5.911	5.980
Li ₂ O + Na ₂ O	mol.%	9.238	9.661	7.069	6.941	6.880	6.909	6.870	6.980
Li ₂ O + MgO	mol.%	8.109	8.831	6.140	7.001	7.981	5.029	5.970	7.071
MgO + ZnO	mol.%	0.9801	2.420	2.021	2.941	4.011	1.960	2.931	3.970
ZrO ₂ + TiO ₂ + FeO + Fe ₂ O ₃	mol.%	9.999E-03	0.010000	0.010001	0.010002	0.010003	9.999E-03	0.010002	0.010000
Measured properties									
$\alpha_{20-300} \times 10^7$	K ⁻¹	45.600	48.400	43.400	42.100	40.000	44.500	42.300	39.200
Str.P.	°C	664	660	670	674	672	681	680	678
An.P.	°C	715	709	726	729	728	738	736	733
T _{soft}	°C	951	938	989	999	999	1010	1008	1006
T _{160kP}	°C	1133	1121	1210	1216	1210	1225	1230	1222
T _{35kP}	°C	1209	1195	1301	1304	1298	1315	1321	1310
T _{200 P}	°C	1605	1587	1741	1754	1744	1772	1777	1757
T _{liq}	°C	>1280	>1330	1225	1295	1200	1225	1180	1200
log(η_{liq})	P	4.01	3.59	5.09	4.61	5.29	5.20	5.61	5.38
E	GPa	83.220	85.770	82.740	83.150	83.500	81.910	82.330	82.670
G	GPa	34.060	35.030	34.200	34.410	34.540	33.920	33.990	34.200
E/d _{RT}	GPa·cm ³ /g	34.300	34.900	34.300	34.600	34.800	34.000	34.200	34.400
d _{RT}	g/cm ³	2.426	2.457	2.409	2.404	2.402	2.409	2.405	2.404
Predicted and calculated properties									
P _{anort}		6.2241	7.4271	7.0084	6.0142	5.0215	7.9985	7.0053	6.0127
P _{cord}		0.99752	2.4431	1.9876	2.9801	3.9728	1.9863	2.9809	3.9737
P _{spod}		7.0082	6.1689	3.9722	3.9728	3.9717	2.9794	2.9786	2.9798
P _{crist}		12.644	4.1157	26.975	26.485	25.991	25.015	24.517	24.013
P _{mod}		-6.4085	-3.4428	0.61597	0.61503	0.61291	0.61340	0.61273	0.61359
P _{Al}		-0.42705	-0.37205	0.61597	0.61503	0.61291	0.61340	0.61273	0.61359
P _{spm} [for E/d _{RT}]	GPa·cm ³ /g	34.29	34.57	34.04	34.17	34.30	33.75	33.87	34.00
P _{anpt} [for An.P.]	°C	694.5	700.5	721.5	721.6	721.8	719.6	719.9	719.8
P _{160kP} [for T _{160kP}]	°C	1162	1162	1202	1203	1204	1201	1203	1203
P _{spm} - (92.5 - 0.05 * P _{160kP})		-0.1365	0.0720	1.3885	1.8421	2.0447	1.3565	1.6112	1.6558

Table 6 Continued

Exemplary Glass		25	26	27	28	29	30	31	32
Composition - mol. %									
SiO ₂	mol. %	73.91	73.49	74.38	73.99	73.12	72.43	74.61	74.52
Al ₂ O ₃	mol. %	12.84	13.33	12.44	12.85	13.34	13.81	12.47	12.46
Li ₂ O	mol. %	3.03	3.04	3.39	3.38	3.04	3.03	2.47	2.51
CaO	mol. %	3.05	3.03	3.11	3.08	3.05	3.04	3.50	3.02
MgO	mol. %	2.98	2.98	2.51	2.51	3.00	2.97	2.97	3.48
Na ₂ O	mol. %	3.92	3.86	3.91	3.92	3.89	3.86	3.86	3.88
Y ₂ O ₃	mol. %	0	0	0	0	0.30	0.60	0	0
K ₂ O	mol. %	0.15	0.15	0.15	0.16	0.15	0.15	0.01	0.01
SnO ₂	mol. %	0.10	0.10	0.09	0.10	0.09	0.09	0.09	0.10
TiO ₂	mol. %	0.01	0.01	0.01	0.01	0.01	0.01	0.01	0.01
Fe ₂ O ₃	mol. %	0.01	0.01	0.01	0.01	0.01	0.01	0.01	0.01
Composition constraints									
CaO + SrO	mol. %	3.050	3.030	3.110	3.080	3.051	3.041	3.500	3.019
CaO + MgO	mol. %	6.030	6.010	5.621	5.589	6.050	6.011	6.470	6.498
Li ₂ O + Na ₂ O	mol. %	6.950	6.901	7.301	7.297	6.931	6.890	6.330	6.388
Li ₂ O + MgO	mol. %	6.009	6.021	5.900	5.888	6.039	6.000	5.439	5.988
MgO + ZnO	mol. %	2.980	2.980	2.510	2.509	2.999	2.971	2.970	3.479
ZrO ₂ + TiO ₂ + FeO + Fe ₂ O ₃	mol. %	0.02000	0.02000	0.02001	0.02000	0.02000	0.02000	0.02000	0.02000
Measured properties									
$\alpha_{20-300} \times 10^7$	K ⁻¹	42.800	42.600	44.900	44.600	42.900	43.500	41.300	41.420
Str.P.	°C	680	685	671	676	682	684	690	688
An.P.	°C	736	740	727	730	736	738	746	744
T _{soft}	°C	1007	1008	1003	1003	998	993	1018	1016
T _{35kP}	°C	1310	1304	1306	1297	1294	1277		
T _{200 P}	°C	1749	1738	1753	1741	1721	1700		
T _{liq}	°C	1205	1235	1205	1205	1195	1205	1220	1215
E	GPa	82.740	83.290	82.190	82.400	83.640	84.460	82.330	82.600
G	GPa	34.200	34.340	33.990	34.130	34.480	34.680	34.060	34.130
E/d _{RT}	GPa·cm ³ /g	34.330	34.500	34.190	34.230	34.390	34.420	34.190	34.300
d _{RT}	g/cm ³	2.410	2.414	2.404	2.407	2.432	2.454	2.408	2.408
Predicted and calculated properties									
P _{anort}		7.0449	7.0455	7.0444	7.0451	7.0451	7.0456	7.4690	6.9696
P _{cord}		2.9926	2.9927	2.4942	2.4943	2.9933	2.9933	2.9933	3.4921
P _{spod}		2.9920	2.9908	3.4903	3.4908	2.9922	2.9919	2.4931	2.4941
P _{crist}		23.709	23.210	23.366	22.960	22.908	22.110	26.152	25.898
P _{mod}		0.23759	-0.26460	0.63876	0.23711	-0.55723	-1.3549	0.48917	0.48719
P _{Al}		0.23759	-0.26460	0.63876	0.23711	0.22086	0.20089	0.48917	0.48719
P _{spm} [for E/d _{RT}]	GPa·cm ³ /g	33.93	34.01	33.86	33.92	33.99	34.06	33.88	33.95
P _{anpt} [for An.P.]	°C	721.3	723.8	713.2	715.0	728.1	734.7	726.9	726.7
P _{160kP} [for T _{160kP}]	°C	1204	1207	1195	1197	1211	1218	1212	1212
P _{spm} - (92.5 - 0.05 * P _{160kP})		1.6374	1.9497	1.2074	1.4013	2.0717	2.5249	2.1109	2.1288

Table 6 Continued

Exemplary Glass		33	34	35	36	37	38	39	40
Composition - mol.%									
SiO ₂	mol.%	74.64	74.29	74.48	74.68	68.75	69.62	69.62	69.87
Al ₂ O ₃	mol.%	12.49	12.80	12.47	12.50	13.38	12.38	12.39	12.25
Li ₂ O	mol.%	3.86	3.92	3.56	4.32	4.31	3.98	3.76	3.75
CaO	mol.%	2.51	2.51	3.00	2.00	4.19	5.18	4.67	5.04
MgO	mol.%	3.51	3.49	3.47	3.49	4.40	3.40	3.96	4.33
Na ₂ O	mol.%	2.88	2.87	2.90	2.89	0.02	0.02	1.94	1.88
P ₂ O ₅	mol.%	0	0	0	0	2.84	2.84	2.83	2.78
BaO	mol.%	0	0	0	0	0.77	0.76	0	0
SrO	mol.%	0	0	0	0	1.21	1.21	0.70	0
K ₂ O	mol.%	0.01	0.01	0.01	0.01	0.01	0.49	0.01	0
SnO ₂	mol.%	0.09	0.09	0.09	0.09	0.10	0.10	0.10	0.10
TiO ₂	mol.%	0.01	0.01	0.01	0.01	0.01	0.01	0.01	0
Fe ₂ O ₃	mol.%	0.01	0.01	0.01	0.01	0.01	0.01	0.01	0
Composition constraints									
CaO + SrO	mol.%	2.509	2.510	3.000	2.000	5.400	6.389	5.370	5.040
CaO + MgO	mol.%	6.018	6.000	6.470	5.490	8.590	8.579	8.631	9.372
Li ₂ O + Na ₂ O	mol.%	6.738	6.790	6.459	7.209	4.331	4.001	5.700	5.631
Li ₂ O + MgO	mol.%	7.368	7.410	7.030	7.810	8.711	7.381	7.721	8.082
MgO + ZnO	mol.%	3.509	3.490	3.470	3.490	4.400	3.400	3.961	4.331
ZrO ₂ + TiO ₂ + FeO + Fe ₂ O ₃	mol.%	0.02000	0.02000	0.02001	0.02000	0.02000	0.02000	0.02000	0
Measured properties									
$\alpha_{20-300 \times 10^7}$	K ⁻¹	40.300	40.500	39.680	40.900	36.800	39.100	40.700	41.000
Str.P.	°C	680	680	684	673	671	675	661	660
An.P.	°C	735	735	739	728	723	726	713	711
T _{soft}	°C	1004	1002	1004	998	978	978	972	968
T _{160kP}	°C					1170	1185	1180	
T _{35kP}	°C			1304	1302	1251	1270	1264	
T _{200 P}	°C			1751	1752	1669	1697	1699	
T _{liq}	°C	1220	1235	1210	1215	1200	1200	1115	1140
log(η_{liq})	P					4.95	5.08		
E	GPa	83.640	83.770	83.640	83.500	82.740	80.530	81.360	81.840
G	GPa	34.610	34.610	34.540	34.540	34.060	33.160	33.580	33.790
E/d _{RT}	GPa·cm ³ /g	34.830	34.860	34.760	34.850	33.500	32.700	33.600	33.860
d _{RT}	g/cm ³	2.401	2.403	2.406	2.396	2.467	2.461	2.424	2.417
Predicted and calculated properties									
P _{anort}		5.4766	5.4777	5.9747	4.9777	5.7691	7.0232	7.3540	6.9446
P _{cord}		3.4913	3.4918	3.4920	3.4919	4.4594	3.4596	4.0009	4.2998
P _{spod}		3.9891	3.9884	3.4905	4.4873	3.9987	3.4987	3.2494	3.7495
P _{crist}		26.882	26.580	27.880	25.888	29.371	29.844	24.238	22.339
P _{mod}		0.48826	0.19029	0.48851	0.48940	-1.7299	-0.72350	-0.74642	-0.17671
P _{Al}		0.48826	0.19029	0.48851	0.48940	4.0088	5.0152	4.9933	5.4630
P _{spm} [for E/d _{RT}]	GPa·cm ³ /g	34.31	34.35	34.25	34.36	33.53	32.97	33.35	33.65
P _{anpt} [for An.P.]	°C	722.4	723.9	728.6	715.9	747.7	746.2	723.9	716.7
P _{160kP} [for T _{160kP}]	°C	1207	1209	1213	1200	1205	1194	1180	1171
P _{spm} - (92.5 - 0.05 * P _{160kP})		2.363	2.474	2.4028	2.1532	0.9443	-0.2710	-0.5716	-0.2505

Table 6 Continued

Exemplary Glass		41	42	43	44	45	46	47	48
Composition - mol.%									
SiO ₂	mol. %	70.53	70.58	70.58	70.68	71.03	70.03	70.86	70.89
Al ₂ O ₃	mol. %	12.38	12.32	12.32	12.30	12.34	12.67	12.74	12.76
Li ₂ O	mol. %	3.76	4.70	3.76	4.70	4.70	0	0	0
CaO	mol. %	5.03	5.04	5.99	4.99	5.04	6.20	5.71	6.21
MgO	mol. %	4.35	4.32	4.29	4.78	4.84	5.52	5.54	5.56
Na ₂ O	mol. %	1.88	0.97	0.97	0.97	0.96	0.96	0.96	0.96
P ₂ O ₅	mol. %	1.97	1.97	1.99	1.49	0.99	0	0	0
BaO	mol. %	0	0	0	0	0	2.67	2.68	1.67
SrO	mol. %	0	0	0	0	0	0.52	0.06	0.52
B ₂ O ₃	mol. %	0	0	0	0	0	0.56	0.57	0.57
Y ₂ O ₃	mol. %	0	0	0	0	0	0.72	0.73	0.72
SnO ₂	mol. %	0.10	0.10	0.10	0.10	0.10	0.12	0.11	0.11
TiO ₂	mol. %	0	0	0	0	0	0.02	0.03	0.02
Fe ₂ O ₃	mol. %	0	0	0	0	0	0.01	0.01	0.01
Composition constraints									
CaO + SrO	mol. %	5.029	5.040	5.990	4.990	5.040	6.720	5.769	6.730
CaO + MgO	mol. %	9.379	9.360	10.28	9.769	9.880	11.72	11.25	11.77
Li ₂ O + Na ₂ O	mol. %	5.641	5.670	4.730	5.669	5.659	0.9600	0.9599	0.9601
Li ₂ O + MgO	mol. %	8.110	9.020	8.050	9.479	9.540	5.521	5.538	5.561
MgO + ZnO	mol. %	4.349	4.320	4.290	4.780	4.841	5.521	5.538	5.561
ZrO ₂ + TiO ₂ + FeO + Fe ₂ O ₃	mol. %	0	0	0	0	0	0.03000	0.04000	0.03000
Measured properties									
$\alpha_{20-300 \times 10^7}$	K ⁻¹	41.100	40.400	39.400	40.400	40.500			
Str.P.	°C	664	660	670	661	660			
An.P.	°C	716	711	722	710	711			791
T _{soft}	°C	966							
T _{35kP}	°C	1253	1240		1233	1237	1262	1277	1274
T _{200 P}	°C	1685	1671		1662	1663	1636	1653	1652
T _{liq}	°C	1115	1165	1170	1160	1175	1210	1235	1205
E	GPa	83.020	83.570	83.710	84.390	85.020	85.840	85.700	86.050
G	GPa	34.270	34.410	34.480	34.750	34.890			
E/d _{RT}	GPa·cm ³ /g	34.260	34.520	34.460	34.760	34.960	32.900	33.000	33.500
d _{RT}	g/cm ³	2.423	2.421	2.429	2.428	2.432	2.610	2.595	2.571
Predicted and calculated properties									
P _{anort}		6.9441	6.0061	6.9464	6.0057	6.0061	8.8396	7.9701	8.4187
P _{cord}		4.3007	4.3001	4.3002	4.8000	4.7991	5.4077	5.4589	5.4615
P _{spod}		3.7501	4.6876	3.7490	4.6891	4.6887	0	0	0
P _{crist}		23.153	25.039	26.906	23.779	24.282	32.678	34.944	35.071
P _{mod}		0.64405	0.64939	0.64949	1.6474	2.1449	2.3223	1.3941	1.3349
P _{Al}		4.6434	4.6484	4.6480	4.6466	4.1441	4.1495	3.2397	3.1816
P _{spm} [for E/d _{RT}]	GPa·cm ³ /g	33.90	34.18	34.07	34.38	34.53	33.05	33.21	33.50
P _{anpt} [for An.P.]	°C	723.3	725.1	737.2	729.0	733.3	802.0	801.9	804.7
P _{160kP} [for T _{160kP}]	°C	1181	1182	1193	1184	1191	1265	1272	1272
P _{spm} - (92.5 - 0.05 * P _{160kP})		0.5605	0.8067	1.2564	1.1535	1.6179	3.7767	4.3127	4.6183

Table 6 Continued

Exemplary Glass		49	50	51	52	53	54	55	56
Composition - mol.%									
SiO ₂	mol.%	71.20	71.09	71.11	71.07	71.24	71.50	73.94	73.91
Al ₂ O ₃	mol.%	12.48	12.33	12.23	12.27	12.27	12.25	12.16	12.18
Li ₂ O	mol.%	3.80	3.79	3.64	3.79	3.78	3.75	4.00	3.99
CaO	mol.%	4.87	5.91	5.72	5.51	5.44	5.47	1.99	2.50
MgO	mol.%	4.68	4.81	4.60	4.43	4.35	4.42	3.97	3.46
Na ₂ O	mol.%	1.85	0.95	0.76	0.97	0.97	0.95	2.89	2.89
P ₂ O ₅	mol.%	0.99	0.99	0.99	0.99	0.99	0.70	0	0
B ₂ O ₃	mol.%	0	0	0.81	0.84	0.83	0.82	0.80	0.81
K ₂ O	mol.%	0.01	0.01	0.01	0.01	0.01	0.01	0.14	0.14
SnO ₂	mol.%	0.10	0.10	0.10	0.10	0.10	0.11	0.09	0.10
TiO ₂	mol.%	0.01	0.01	0.01	0.01	0.01	0.01	0.01	0.01
Fe ₂ O ₃	mol.%	0.01	0.01	0.01	0.01	0.01	0.01	0.01	0.01
Composition constraints									
CaO + SrO	mol.%	4.871	5.910	5.721	5.509	5.441	5.470	1.990	2.500
CaO + MgO	mol.%	9.551	10.72	10.32	9.939	9.792	9.890	5.961	5.958
Li ₂ O + Na ₂ O	mol.%	5.650	4.741	4.400	4.760	4.752	4.700	6.890	6.879
Li ₂ O + MgO	mol.%	8.480	8.600	8.241	8.220	8.132	8.169	7.971	7.448
MgO + ZnO	mol.%	4.681	4.809	4.601	4.430	4.351	4.419	3.970	3.459
ZrO ₂ + TiO ₂ + FeO + Fe ₂ O ₃	mol.%	0.02000	0.02000	0.02001	0.02000	0.02000	0.02000	0.02000	0.02000
Measured properties									
$\alpha_{20-300} \times 10^7$	K ⁻¹	41.500	39.800	37.400	38.400	38.100	37.600	40.900	41.700
Str.P.	°C	665	675	668	664	665	666	659	659
An.P.	°C	717	725	720	716	716	718	715	714
T _{soft}	°C	963	968	964	963	961	963	980	980
T _{35kP}	°C	1239	1245	1241	1242	1241	1247	1286	1281
T _{200 P}	°C	1666	1666	1659	1662	1664	1674	1733	1728
T _{liq}	°C	>1275	1190	1170	1155	1160	1190	1205	1165
E	GPa	84.530	85.360	84.330	84.190	84.190	84.600	82.530	82.530
G	GPa	34.820	35.030	34.610	34.540	34.540	34.750	34.200	34.130
E/d _{RT}	GPa·cm ³ /g	34.760	34.970	34.700	34.690	34.700	34.840	34.490	34.450
d _{RT}	g/cm ³	2.432	2.441	2.430	2.427	2.426	2.428	2.393	2.396
Predicted and calculated properties									
P _{anort}		6.7259	6.8653	6.4862	6.4844	6.4161	6.4253	4.9500	5.4589
P _{cord}		4.6808	4.8094	4.6014	4.4298	4.3506	4.4194	3.9704	3.4588
P _{spod}		3.7993	3.7905	3.6395	3.7905	3.7814	3.7499	4.0001	3.9897
P _{crist}		23.399	26.324	28.988	27.935	28.474	28.748	25.856	26.134
P _{mod}		1.7434	2.1504	1.5072	1.4498	1.2944	1.6465	0.83248	0.79928
P _{Al}		3.7235	4.1305	3.4875	3.4298	3.2743	3.0465	0.83248	0.79928
P _{spm} [for E/d _{RT}]	GPa·cm ³ /g	34.29	34.44	34.33	34.29	34.29	34.38	34.18	34.11
P _{anpt} [for An.P.]	°C	732.8	745.4	739.8	734.5	734.8	737.9	711.3	711.4
P _{160kP} [for T _{160kP}]	°C	1194	1202	1201	1196	1197	1201	1192	1192
P _{spm} - (92.5 - 0.05 * P _{160kP})		1.4684	2.0544	1.8622	1.5844	1.6495	1.9442	1.2885	1.2018

Table 6 Continued

Exemplary Glass		57	58	59	60	61	62	63	64	65
Composition - mol. %										
SiO ₂	mol. %	73.98	73.74	73.22	68.62	68.50	68.59	68.79	68.88	68.93
Al ₂ O ₃	mol. %	12.18	12.85	12.59	12.26	12.57	12.28	12.41	12.34	12.27
Li ₂ O	mol. %	4.39	3.99	4.03	4.40	4.21	4.17	4.24	4.22	4.21
CaO	mol. %	2.00	2.01	1.99	5.26	5.36	5.02	5.53	5.33	5.06
MgO	mol. %	3.48	3.98	3.95	4.31	4.23	4.30	4.82	4.83	4.86
Na ₂ O	mol. %	2.91	2.88	2.88	2.05	2.03	2.06	2.05	2.05	2.06
P ₂ O ₅	mol. %	0	0	0	1.96	1.96	1.95	1.00	0.99	0.98
B ₂ O ₃	mol. %	0.81	0	0.79	1.01	1.02	1.00	1.04	0.98	1.01
Y ₂ O ₃	mol. %	0	0.30	0.30	0	0	0	0	0	0
ZnO	mol. %	0	0	0	8.00E-4	0	0.50	0	0.25	0.49
K ₂ O	mol. %	0.14	0.14	0.14	0.011	0.011	0.011	0.0102	0.0109	0.0102
SnO ₂	mol. %	0.09	0.09	0.09	0.0972	0.0994	0.0995	0.10	0.0999	0.10
TiO ₂	mol. %	0.01	0.01	0.01	0.0065	0.0057	0.0049	0.0064	0.0056	0.0056
Fe ₂ O ₃	mol. %	0.01	0.01	0.01	0.0138	0.0138	0.0134	0.014	0.014	0.0141
Composition constraints										
CaO + SrO	mol. %	2.000	2.010	1.990	5.256	5.356	5.020	5.525	5.326	5.063
CaO + MgO	mol. %	5.479	5.990	5.939	9.564	9.589	9.320	10.34	10.16	9.922
Li ₂ O + Na ₂ O	mol. %	7.298	6.869	6.910	6.448	6.241	6.236	6.290	6.276	6.261
Li ₂ O + MgO	mol. %	7.868	7.969	7.979	8.711	8.445	8.474	9.054	9.054	9.064
MgO + ZnO	mol. %	3.479	3.980	3.949	4.309	4.233	4.796	4.816	5.081	5.354
ZrO ₂ + TiO ₂ + FeO + Fe ₂ O ₃	mol. %	0.01999	0.02000	0.02000	0.02028	0.01950	0.01830	0.02042	0.01965	0.01967
Measured properties										
$\alpha_{20-300 \times 10^7}$	K ⁻¹	42.700	41.200	41.100	44.700	44.200	43.200	44.300	43.300	43.000
Str.P.	°C	655	676	662	631	636	634	635	635	634
An.P.	°C	709	730	716	680	686	683	684	684	683
T _{soft}	°C	976	988	975	927	927	927	925	922	923
T _{35kP}	°C	1283	1286	1272	1207	1212	1214	1200	1200	1204
T _{200 P}	°C	1741	1723	1706	1642	1644	1642	1629	1626	1634
T _{liq}	°C	1155	1210	1205						
E	GPa	82.190	84.390	83.500	82.188	82.326	82.395	83.774	83.981	83.912
G	GPa	33.990	34.820	34.410	33.854	33.923	33.923	34.406	34.475	34.475
E/d _{RT}	GPa·cm ³ /g	34.370	34.850	34.550	33.878	33.907	33.866	34.334	34.390	34.348
d _{RT}	g/cm ³	2.391	2.422	2.417	2.426	2.428	2.433	2.440	2.442	2.443
Predicted and calculated properties										
P _{anort}		4.9792	4.9599	4.9397	7.3063	7.3912	7.0868	7.5816	7.3848	7.1238
P _{cord}		3.4789	3.9799	3.9490	4.3084	4.2331	4.2999	4.8162	4.8320	4.8593
P _{spod}		4.3891	3.9891	4.0297	4.4023	4.2117	4.1740	4.2381	4.2223	4.2047
P _{cris}		25.434	25.700	25.132	17.391	18.110	18.669	16.380	16.873	17.437
P _{mod}		0.73767	-0.14829	0.09976	1.7981	1.3128	1.8309	3.2374	3.3593	3.4366
P _{Al}		0.73767	0.63166	0.88001	5.7266	5.2259	5.7259	5.2288	5.3414	5.3934
P _{spm} [for E/d _{RT}]	GPa·cm ³ /g	34.16	34.37	34.23	33.80	33.82	33.68	34.14	34.09	34.03
P _{anpt} [for An.P.]	°C	706.0	729.2	717.6	700.2	704.5	699.3	710.4	709.1	706.5
P _{160kP} [for T _{160kP}]	°C	1187	1212	1199	1149	1155	1149	1159	1158	1156
P _{spm} - (92.5 - 0.05 *		1.0202	2.4542	1.663	-1.263	-0.9322	-1.3887	-0.4037	-0.5173	-0.6790

Exemplary Glass	57	58	59	60	61	62	63	64	65
P _{160kP})									

[00144] Table 7 below lists the compositions and properties for Comparative Glasses C1-C5.

Table 7. Compositions and Properties of Comparative Example Glasses

Comparative Examples		C1	C2	C3	C4	C5
Reference		[1]	[2]	[3]	[4]	[3]
Composition - mol. %						
SiO ₂	mol. %	69.10	71.64	63.85	68.73	65.06
Al ₂ O ₃	mol. %	13.26	13.07	16.86	11.39	16.29
B ₂ O ₃	mol. %	3.98	2.00	0	4.72	0
Na ₂ O	mol. %	3.62	1.42	9.25	4.45	7.95
Li ₂ O	mol. %	7.09	6.98	7.48	4.39	7.68
P ₂ O ₅	mol. %	1.22	0	2.50	1.85	2.46
CaO	mol. %	0.81	0	0	0	0
MgO	mol. %	0.85	3.00	0	3.26	0.51
SnO ₂	mol. %	0.0473	0.10	0.049	0.17	0.0487
ZrO ₂	mol. %	0.0105	0	0	0.16	0
Fe ₂ O ₃	mol. %	0.0041	0	0	0	0
ZnO	mol. %	0	1.80	0	0.40	0
Ce ₂ O ₃	mol. %	0	0	0	0.096	0
K ₂ O	mol. %	0	0	0	0.21	0
SiF ₄	mol. %	0	0	0	0.0883	0
TiO ₂	mol. %	0	0	0	0.0822	0
Measured properties						
$\alpha_{20-300} \times 10^7$	K ⁻¹				60.000	
Str.P.	°C	572.00	613.00	608.00		606.00
An.P.	°C	626.00	666.00	661.00		661.00
T _{soft}	°C			924.00	900.00	926.40
T _{160kP}	°C			1138.3	1168.8	1139.6
T _{35kP}	°C			1223.1	1283.3	1226.0
T _{200P}	°C			1659.0	2008.0	1661.0
T _{liq}	°C	1145.0	1280.0	1095.0	1100.0	1105.0
log(η_{liq})	P	5.0810	4.1430	5.5850	5.6820	5.4990
E	GPa	75.700	82.600	77.400	73.000	77.800
G	GPa		34.130	32.000		32.100
E/d _{RT}	GPa-cm ³ /g			32.160	30.290	32.420
d _{RT}	g/cm ³			2.407	2.410	2.400
Predicted and calculated properties						
P _{spm} [for E/d _{RT}]	GPa-cm ³ /g	41.04	41.59	41.70	38.24	42.03
P _{anpt} [for An.P.]	°C	639.0	689.8	612.5	624.5	625.9
P _{160kP} [for T _{160kP}]	°C	1251	1305	1217	1209	1238

[00145] The reference key for each of the Comparative Glasses listed in Table 7 is as follows:

[1] US2019127265; [2] US2020199019; [3] US2019161390; [4] US2020199013.

[00146] In some embodiments of the present disclosure, the glass articles can be formed from the melts with the fusion draw process at a viscosity that is close to 160 kP. To enable use of the fusion draw process, the temperature at which the viscosity of the glass is 160 KP (T_{160kP}) should be greater than or about equal to the liquidus temperature of glass. At the same time, to be suitable with the HAMR process, the glass should have a high specific modulus.

[00147] However, it has been empirically observed that for alumina-rich silicate glasses, it is difficult to achieve a high specific modulus together with a high T_{160kP} , while still keeping the liquidus temperature acceptably low to enable high volume manufacturing. Without being bound to a specific theory, it is hypothesized that this difficulty may be caused competing effects of different components on the liquidus temperature, high-temperature viscosity, and specific modulus. Components that increase the specific modulus (such as, for example, MgO, Li₂O, TiO₂, Y₂O₃, ZrO₂ and others) tend to either decrease the high-temperature viscosity (as, for example, Li₂O), increase the liquidus temperature (as, for example, MgO, TiO₂, ZrO₂), or both (as, for example, in the case of Y₂O₃), thus establishing a trade-off between the three attributes: specific modulus, liquidus temperature and T_{160kP} . To balance this trade-off, the glass compositions of some embodiments of the present disclosure contain the components in the combinations and proportions that are specified below.

[00148] FIG. 4 is a plot showing the relationship between P_{160kP} , the parameter that predicts T_{160kP} , and the specific modulus parameter P_{spm} for some of the Exemplary Glasses and some of the Comparative Glasses. The Exemplary Glasses (filled circles) are the Examples 7 to 29 from Table 6. The Comparative Glasses (open circles) are the Examples C1 and C2 from Table 7. P_{160kP} was determined according to Formula (XI). The parameter P_{spm} that predicts specific modulus was determined according to Formula (X). All of the Exemplary Glasses and Comparative Glasses shown in FIG. 4 have the features specified in Table 8. In Table 8, the specification "Not limited" refers to a limitation that was not considered when selecting the compositions.

Table 8. Limitations for glass compositions shown in FIG.4

Quantity	Unit	Min	Max
SiO ₂	mol.%	60	75
Al ₂ O ₃	mol.%	10.5	18
CaO	mol.%	0	10
Li ₂ O	mol.%	0	7.8
MgO	mol.%	0	4.3
ZrO ₂	mol.%	0	0.5
RE _m O _n	mol.%	0	5
P_{spm}	GPa-cm ³ /g	25	Not limited
P_{mod}		-3	Not limited
P_{crist}		Not limited	28
P_{anort}		Not limited	10

[00149] The Comparative Glasses C1 and C2 depicted in FIG. 4 were selected from among the Comparative Glasses as having the highest specific modulus parameter P_{spm} with a value of P_{160kP} closest to the value of P_{160kP} of the Exemplary Glasses shown in FIG. 4.

[00150] The line corresponding to the formula $Y = 92.5 - 0.05 * X$ shown in FIG. 4 (where Y represents P_{spm} and X represents P_{160kP}) provides a visual representation of the differences between the Comparative Glasses C_1 and C_2 and the Exemplary Glasses 7 to 29. As can be seen in FIG. 4, the Exemplary Glasses (filled circles) and none of the Comparative Glasses (open circles) represented in FIG. 4 fall above the line $Y = 92.5 - 0.05 * X$. In other words, all of the Exemplary Glasses and none of the Comparative Glasses represented in FIG. 4 satisfy the following formula (XII)(a):

$$P_{spm} - (92.5 - 0.05 * P_{160kP}) > 0.00 \quad (XII)(a)$$

[00151] As can also be seen in FIG. 4, some of Exemplary Glasses and none of the Comparative Glasses represented in FIG. 4 fall above the line $Y = 94 - 0.05 * X$, where Y corresponds to P_{spm} and x corresponds to P_{160kP} . In other words, some of the Exemplary Glasses and none of the Comparative Glasses represented in FIG. 4 satisfy the following formula (XII)(b):

$$P_{spm} - (94 - 0.05 * P_{160kP}) > 0.00 \quad (XII)(b)$$

[00152] The Exemplary Glasses represented in FIG. 4 are, by prediction, superior in terms of the combination of T_{160kP} and E/d_{RT} to the best known Comparative Glasses that have the features specified in Table 8.

[00153] FIG. 5 is a plot showing the relationship between the temperature T_{160kP} , the temperature at which the glass has a viscosity of 160 kP, and the specific modulus E/d_{RT} for some of the Exemplary Glasses and some of the Comparative Glasses. The Exemplary Glasses (filled circles) are the Examples 7 to 12, 15 and 39 from Table 6. The Comparative Glasses (open circles) are the Examples C3 to C5 from Table 7. All of the Exemplary Glasses and Comparative Glasses shown in FIG. 5 have the features specified in Table 9. In Table 9, the specification "Not limited" refers to a limitation that was not considered when selecting the compositions. In FIG. 5, some of the above-enumerated compositions may be labeled for better visibility, some others may not, and some more glasses may not be shown, which does not affect the further conclusions.

Table 9. Limitations for glass compositions shown in FIG. 5

Quantity	Unit	Min	Max
SiO ₂	mol. %	60	75
Al ₂ O ₃	mol. %	10.5	18
CaO	mol. %	0	10
Li ₂ O	mol. %	0	7.8
MgO	mol. %	0	4.3
ZrO ₂	mol. %	0	0.5
RE _m O _n	mol. %	0	5
E/d	GPa·cm ³ /g	25	Not limited
P _{mod}		-3	Not limited
P _{crist}		Not limited	28
P _{anort}		Not limited	10

[00154] The Comparative Glasses C3, C4, and C5 shown in FIG. 5 were selected from among the Comparative Glasses as having the highest measured values of the specific modulus E/d_{RT} with a value of T_{160kP} closest to the value of T_{160kP} of the Exemplary Glasses shown in FIG. 5.

[00155] The line corresponding to the formula $Y = 92.5 - 0.05 * X$ (where Y corresponds to E/d_{RT} and X corresponds to T_{160kP}) shown in FIG. 5 provides a visual representation of the differences between the Comparative Glasses C3, C4, and C5 and the Exemplary Glasses 7 to 12, 15 and 39. As can be seen in FIG. 5, all of the Exemplary Glasses (filled circles) and none of the Comparative Glasses (open circles) represented in FIG. 5 fall above the line $Y = 92.5 - 0.05 * X$. In other words, some of the Exemplary Glasses and none of the Comparative Glasses represented in FIG. 5 satisfy the following formula (XIII)(a):

$$E/d - (92.5 - 0.05 * T_{160kP}) > 0.00 \quad (XIII)(a)$$

[00156] As can also be seen in FIG. 5, some of the Exemplary Glasses and none of the Comparative Glasses represented in FIG. 5 fall above the line $Y = 94 - 0.05 * X$. In other words, some of the Exemplary Glasses and none of the Comparative Glasses represented in FIG. 5 satisfy the following formula (XIII)(b):

$$E/d_{RT} - (94 - 0.05 * T_{160kP}) > 0.00 \quad (XIII)(b)$$

[00157] The Exemplary Glasses represented in FIG. 5 that satisfy the formula (XIII)(b) are characterized by the highest values of E/d_{RT} at comparable values of T_{160kP} among the glasses that have the features specified in Table 9.

[00158] This means that, under the conditions specified in Table 9, some of the Exemplary Glasses have higher measured values of the specific modulus E/d_{RT} at comparable measured values of the temperature T_{160kP} than the best of the Comparative Glasses satisfying the same conditions. The Exemplary Glasses, according to measurements, have higher values of E/d_{RT} at comparable values of T_{160kP} than the best of the Comparative Glasses that have the features specified in Table 9.

[00159] The values of all attributes specified in Tables 8 and 9 and Formulas (XII)(a), (XII)(b), (XIII)(a) and (XIII)(b) for the Comparative Glasses C1 to C5 plotted in FIGS 4 and 5 are presented in Table 10 below.

Table 10. Attributes of Comparative Example Glasses Having the Features of Tables 8 and 9

Ex. #		C1	C2	C3	C4	C5
Composition						
Al ₂ O ₃	mol. %	13.26	13.07	16.86	11.39	16.29
CaO	mol. %	0.81	0	0	0	0
Li ₂ O	mol. %	7.10	6.98	7.47	4.39	7.68
RE _m O _n	mol. %	0	0	0	0.0953	0
MgO	mol. %	0.85	3.00	0	3.26	0.51
ZrO ₂	mol. %	0.011	0	0	0.16	0
Measured properties						
P _{mod}				-2.6347	-0.6289	-2.6044
P _{cris}				-21.58	15.08	-14.68

Ex. #		C1	C2	C3	C4	C5
P_{anort}				9.2567	4.5523	7.9535
T_{160kP}	°C			1138	1169	1140
E/d	GPa·cm ³ /g			32.16	30.29	32.42
$E/d - (92.5 - 0.05 * T_{160kP})$				-3.4275	-3.769	-3.101
$E/d - (94 - 0.05 * T_{160kP})$				-4.9275	-5.269	-4.601
Predicted and calculated properties						
P_{160kP}		1097	1154	1047	1079	1063
P_{spm}		33.49	34.36	32.92	32.63	33.19
$P_{spm} - (92.5 - 0.05 * P_{160kP})$		-4.1728	-0.4556	-7.2483	-5.9271	-6.1570
$P_{spm} - (94 - 0.05 * P_{160kP})$		-5.6728	-1.9556	-8.7483	-7.4271	-7.6570

[00160] As follows from FIGS. 4 and 5, both predicted and measured property data confirms that some of the Exemplary Glasses have better combination of modifier-binding parameter P_{mod} , temperature T_{160kP} and specific modulus E/d_{RT} than the best of the Comparative Glasses that have the features specified in Tables 8 and 9 accordingly.

[00161] The following non-limiting aspects are encompassed by the present disclosure. To the extent not already described, any one of the features of the first through the seventy-third aspect may be combined in part or in whole with features of any one or more of the other aspects of the present disclosure to form additional aspects, even if such a combination is not explicitly described.

[00162] According to a first aspect, the glass comprises a plurality of components, the glass having a composition of the components comprising greater than or equal to 60.0 mol.% and less than or equal to 75.0 mol.% SiO₂, greater than or equal to 10.0 mol.% and less than or equal to 18.0 mol.% Al₂O₃, greater than or equal to 0.0 mol.% and less than or equal to 10.0 mol.% Li₂O, greater than or equal to 0.0 mol.% and less than or equal to 5.0 mol.% Na₂O, greater than or equal to 0.0 mol.% and less than or equal to 3.8 mol.% B₂O₃, greater than or equal to 0.0 mol.% and less than or equal to 3.0 mol.% P₂O₅, greater than or equal to 0.0 mol.% and less than or equal to 2.0 mol.% ZnO, greater than or equal to 0.0 mol.% and less than or equal to 0.7 mol.% La₂O₃, greater than or equal to 0.0 at.% and less than or equal to 3.0 at.% F, a sum of CaO + MgO greater than or equal to 5.0 mol.%, a sum of Li₂O + Na₂O greater than or equal to 0.5 mol.%, a sum of Li₂O + MgO greater than or equal to 0.0 mol.% and less than or equal to 10.0 mol.%, a sum of MgO + ZnO greater than or equal to 0.0 mol.% and less than or equal to 10.0 mol.%, a sum of CaO + SrO greater than or equal to 0.0 mol.% and less than or equal to 9.0 mol.% and a sum of ZrO₂ + TiO₂ + FeO + Fe₂O₃ greater than or equal to 0.0 mol.% and less than or equal to 1.5 mol.%, wherein the glass has an aluminum-binding parameter P_{Al} that is greater than or equal to -2.8, a modifier-binding parameter P_{mod} that is less than or equal to 2.8 and an anorthite precipitation parameter P_{anort} that is less than or equal to 10, where P_{Al} is calculated from the glass composition in terms of mol.% of the components according to the following formula:

$$P_{Al} = R_2O + RO + P_2O_5 + 1.6 * RE_mO_n - Al_2O_3,$$

P_{mod} is calculated from the glass composition in terms of mol.% of the components according to the following formula:

$$P_{mod} = R_2O + RO - Al_2O_3 - P_2O_5 - RE_mO_n,$$

P_{anort} is calculated from the glass composition in terms of mol.% of the components according to the following formula:

$$P_{anort} = \min(CaO + SrO + 0.5 * BaO + Na_2O + 0.5 * K_2O, Al_2O_3),$$

where R_2O is a total sum of monovalent metal oxides, RO is a total sum of divalent metal oxides, RE_mO_n is a total sum of rare earth metal oxides in all redox states present, and an asterisk (*) means multiplication.

[00163] According to a second aspect, the glass of the first aspect, wherein the composition of the components comprises greater than or equal to 10.0 mol.% and less than or equal to 17.0 mol.% Al_2O_3 , greater than or equal to 0.0 mol.% and less than or equal to 7.4 mol.% Li_2O , greater than or equal to 0.0 mol.% and less than or equal to 4.8 mol.% Na_2O , greater than or equal to 0.0 mol.% and less than or equal to 4.3 mol.% MgO , greater than or equal to 0.0 mol.% and less than or equal to 3.0 mol.% B_2O_3 , greater than or equal to 0.0 mol.% and less than or equal to 0.5 mol.% ZrO_2 , greater than or equal to 2.5 mol.% Alk_2O , a sum of $MgO + CaO + SrO + BaO + ZnO$ greater than or equal to 1.0 mol.%, wherein the composition of the components is substantially free of fluorine, and wherein P_{Al} is less than or equal to 3.0, and modifier-binding parameter, P_{mod} is greater than or equal to -3.0, and where the glass has a cordierite precipitation parameter, P_{cord} that is less than or equal to 8.0 and a cristobalite precipitation parameter, P_{crist} that is less than or equal to 28 where

P_{cord} is calculated from the glass composition in terms of mol.% of the components according to the following formula:

$$P_{cord} = MgO + MnO + FeO,$$

P_{crist} is calculated from the glass composition in terms of mol.% of the components according to the following formula:

$$P_{crist} = SiO_2 - 6 * (Na_2O + K_2O) - 4 * Li_2O - 2 * (CaO + SrO + BaO) - 2.5 * MgO,$$

where Alk_2O is a total sum of alkali metal oxides.

[00164] According to a third aspect, the glass of any one of aspects 1-2, wherein the composition of the components comprises greater than or equal to 0.3 mol.% and less than or equal

to 7.5 mol.% CaO, greater than or equal to 0.3 mol.% and less than or equal to 7.5 mol.% Li₂O, greater than or equal to 0.0 mol.% and less than or equal to 5.0 mol.% MgO, greater than or equal to 0.0 mol.% and less than or equal to 0.5 mol.% SnO₂ and greater than or equal to 0.0 mol.% and less than or equal to 0.3 mol.% Fe₂O₃.

[00165] According to a fourth aspect, the glass of aspect 1, wherein the composition of the components comprises one or more of the following components: greater than or equal to 63.9 mol.% and less than or equal to 74.4 mol.% SiO₂, greater than or equal to 12.34 mol.% and less than or equal to 18.0 mol.% Al₂O₃, greater than or equal to 2.8 mol.% and less than or equal to 7.0 mol.% Li₂O, greater than or equal to 1.4 mol.% and less than or equal to 5.6 mol.% CaO, greater than or equal to 0.4 mol.% and less than or equal to 4.4 mol.% MgO, greater than or equal to 0.4 mol.% and less than or equal to 4.4 mol.% Na₂O, greater than or equal to 0.0 mol.% and less than or equal to 1.1 mol.% SrO, greater than or equal to 0.0 mol.% and less than or equal to 0.8 mol.% B₂O₃, greater than or equal to 0.0 mol.% and less than or equal to 0.7 mol.% BaO and greater than or equal to 0 mol.% and less than or equal to 0.55 mol.% Y₂O₃.

[00166] According to a fifth aspect, the glass of any one of aspects 1 and-4, wherein the composition of the components comprises greater than or equal to 64.5 mol.% and less than or equal to 74.4 mol.% SiO₂, greater than or equal to 12.34 mol.% and less than or equal to 18.0 mol.% Al₂O₃, greater than or equal to 3.2 mol.% and less than or equal to 6.2 mol.% Li₂O, greater than or equal to 1.9 mol.% and less than or equal to 5.0 mol.% CaO, greater than or equal to 1.0 mol.% and less than or equal to 4.0 mol.% MgO, greater than or equal to 0.9 mol.% and less than or equal to 4.0 mol.% Na₂O, greater than or equal to 0 mol.% and less than or equal to 0.95 mol.% SrO, greater than or equal to 0.0 mol.% and less than or equal to 0.7 mol.% B₂O₃, greater than or equal to 0 mol.% and less than or equal to 0.625 mol.% BaO and greater than or equal to 0.0 mol.% and less than or equal to 0.5 mol.% Y₂O₃.

[00167] According to a sixth aspect, the glass of aspect 1, wherein the composition of the components comprises greater than or equal to 11.0 mol.% and less than or equal to 18.0 mol.% Al₂O₃, greater than or equal to 0.5 mol.% and less than or equal to 7.5 mol.% Li₂O, greater than or equal to 0.0 mol.% and less than or equal to 3.0 mol.% B₂O₃, greater than or equal to 0.0 mol.% and less than or equal to 0.5 mol.% ZrO₂, greater than or equal to 0.5 mol.% Alk₂O, greater than or equal to 0.0 at.% and less than or equal to 0.5 at.% F, a sum of MgO + CaO + SrO + BaO + ZnO greater than or equal to 0.5 mol.% , wherein the composition of the components satisfies the conditions: $\min(\text{RE}_m\text{O}_n, \text{P}_2\text{O}_5) [\text{mol.}\%] \leq 0.15$, and wherein the glass has an annealing point An.P. that is greater than or equal to 680 °C, a temperature at which the viscosity is 160 kP T_{160kP} that is greater than or equal to 1150 °C, a specific modulus, E/d_{RT} that is greater than or equal to 32 GPa·cm³/g, a cordierite

precipitation parameter P_{cord} that is less than or equal to 5.0, a cristobalite precipitation parameter P_{crist} that is less than or equal to 28 and a spodumene precipitation parameter P_{spod} that is less than or equal to 7.5, where E is a Young's modulus, d_{RT} is a density, P_{cord} is calculated from the glass composition in terms of mol.% of the components according to the following formula:

$$P_{\text{cord}} = \text{MgO} + \text{MnO} + \text{FeO},$$

P_{crist} is calculated from the glass composition in terms of mol.% of the components according to the following formula:

$$P_{\text{crist}} = \text{SiO}_2 - 6 * (\text{Na}_2\text{O} + \text{K}_2\text{O}) - 4 * \text{Li}_2\text{O} - 2 * (\text{CaO} + \text{SrO} + \text{BaO}) - 2.5 * \text{MgO},$$

P_{spod} is calculated from the glass composition in terms of mol.% of the components according to the following formula:

$$P_{\text{spod}} = \min(\text{Li}_2\text{O}, \text{Al}_2\text{O}_3 - \text{K}_2\text{O} - 0.5 * \text{Na}_2\text{O}),$$

where Alk_2O is a total sum of alkali metal oxides.

[00168] According to a seventh aspect, the glass of any one of aspects 1-4, wherein the composition of the components comprises greater than or equal to 0.0 mol.% and less than or equal to 1.0 mol.% BaO and greater than or equal to 0.0 mol.% and less than or equal to 1.0 mol.% K_2O , wherein the composition of the components is substantially free of fluorine and substantially free of PbO and wherein the composition of the components satisfies the conditions: $0.00 \leq \min(\text{RE}_m\text{O}_n, \text{P}_2\text{O}_5) [\text{mol.\%}] \leq 0.30$.

[00169] According to an eighth aspect, the glass of any one of aspects 1-7, wherein the glass has a logarithm of liquidus viscosity, $\log(\eta_{\text{liq}} [P])$ that is greater than or equal to 5.0.

[00170] According to a ninth aspect, the glass of the eighth aspect, wherein the glass has a logarithm of liquidus viscosity, $\log(\eta_{\text{liq}} [P])$ that is greater than or equal to 5.2.

[00171] According to a tenth aspect, the glass of any one of aspects 1-9, wherein the glass satisfies the condition: $P_{\text{anpt}} > 710$, where P_{anpt} is an annealing point parameter, calculated from the glass composition in terms of mol.% of the components according to the Formula (I):

$$P_{\text{anpt}} = 664.7 + 5.2303 * \text{SiO}_2 - 11.493 * \text{B}_2\text{O}_3 - 7.1742 * \text{P}_2\text{O}_5 + 8.3980 * \text{ZrO}_2 - 2.0585 * \text{MgO} - 2.1088 * \text{CaO} - 3.8995 * \text{BaO} - 10.323 * \text{ZnO} - 9.0727 * \text{MnO} - 23.455 * \text{Li}_2\text{O} - 33.819 * \text{Na}_2\text{O} - 25.204 * \text{K}_2\text{O} + 15.745 * \text{Y}_2\text{O}_3 + 8.9047 * \text{La}_2\text{O}_3 - 33.960 * (\text{Fe}_2\text{O}_3 + \text{FeO}) - 5.6704 * (\text{R}_2\text{O} + \text{RO} - \text{Al}_2\text{O}_3) - 4.2545 * (\text{SiO}_2 - (6 * \text{K}_2\text{O} + 6 * \text{Na}_2\text{O} + 4 * \text{Li}_2\text{O} + 2 * \text{RO})) - 19.439 * \text{Cu}_2\text{O}, \quad (\text{I}).$$

[00172] According to an eleventh aspect, the glass of any one of aspects 1-10, wherein the glass has an annealing point An.P. that is greater than or equal to 710 °C.

[00173] According to a twelfth aspect, the glass of the eleventh aspect, wherein the annealing point An.P. is greater than or equal to 730 °C.

[00174] According to a thirteenth aspect, the glass of any one of aspects 1-12, wherein the glass has a cordierite precipitation parameter, P_{cord} that is less than or equal to 5.0, a cristobalite precipitation parameter, P_{crist} that is less than or equal to 28 and a spodumene precipitation parameter, P_{spod} that is less than or equal to 7.5, where P_{cord} is calculated from the glass composition in terms of mol.% of the components according to the following formula:

$$P_{\text{cord}} = \text{MgO} + \text{MnO} + \text{FeO},$$

P_{crist} is calculated from the glass composition in terms of mol.% of the components according to the following formula:

$$P_{\text{crist}} = \text{SiO}_2 - 6 * (\text{Na}_2\text{O} + \text{K}_2\text{O}) - 4 * \text{Li}_2\text{O} - 2 * (\text{CaO} + \text{SrO} + \text{BaO}) - 2.5 * \text{MgO},$$

P_{spod} is calculated from the glass composition in terms of mol.% of the components according to the following formula:

$$P_{\text{spod}} = \min(\text{Li}_2\text{O}, \text{Al}_2\text{O}_3 - \text{K}_2\text{O} - 0.5 * \text{Na}_2\text{O}).$$

[00175] According to a fourteenth aspect, the glass of any one of aspects 1-13, wherein the glass satisfies the condition: $P_{\text{spm}} > 32$, where P_{spm} is a specific modulus parameter, calculated from the glass composition in terms of mol.% of the components according to the Formula (II):

$$\begin{aligned} P_{\text{spm}} = & 32.10 + 0.47744 * \text{SiO}_2 - 1.6506 * \text{Al}_2\text{O}_3 - 0.11775 * \text{B}_2\text{O}_3 - 0.30166 * \text{P}_2\text{O}_5 + 0.20187 * \\ & \text{TiO}_2 + 0.20219 * \text{ZrO}_2 + 0.96268 * \text{MgO} + 0.83379 * \text{CaO} + 0.53685 * \text{SrO} + 0.41218 * \text{BaO} + \\ & 0.63264 * \text{ZnO} + 0.75365 * \text{MnO} + 0.62984 * \text{CuO} - 1.2496 * \text{Na}_2\text{O} - 1.5154 * \text{K}_2\text{O} + 1.4746 * \\ & \text{Cu}_2\text{O} - 0.037941 * \text{Y}_2\text{O}_3 - 0.75836 * \text{La}_2\text{O}_3 - 1.8052 * (\text{R}_2\text{O} + \text{RO} - \text{Al}_2\text{O}_3) - 0.47488 * (\text{SiO}_2 - (6 * \\ & \text{K}_2\text{O} + 6 * \text{Na}_2\text{O} + 4 * \text{Li}_2\text{O} + 2 * \text{RO})), \end{aligned} \quad (\text{II}).$$

[00176] According to a fifteenth aspect, the glass of any one of aspects 1-14, wherein the glass has specific modulus E/d_{RT} that is greater than or equal to 32 GPa·cm³/g, where E is a Young's modulus and d_{RT} is a density.

[00177] According to a sixteenth aspect, the glass of the fifteenth aspect, wherein the specific modulus E/d_{RT} is greater than or equal to 33 GPa·cm³/g.

[00178] According to a seventeenth aspect, the glass of the sixteenth aspect, wherein the specific modulus E/d_{RT} is greater than or equal to 34 GPa·cm³/g.

[00179] According to an eighteenth aspect, the glass of the seventeenth aspect, wherein the specific modulus E/d_{RT} is greater than or equal to 35 GPa·cm³.

[00180] According to a nineteenth aspect, the glass of any one of aspects 1-18, wherein the glass satisfies the conditions $P_{160kP} > 1150$, where P_{160kP} is a parameter predicting the temperature at which the glass has a viscosity of 160 kP, calculated from the glass composition in terms of mol.% of the components according to the Formula (III):

$$P_{160kP} = 1058 + 2.5492 * SiO_2 - 25.725 * Al_2O_3 - 11.327 * B_2O_3 - 10.014 * P_2O_5 - 14.309 * TiO_2 - 11.594 * ZrO_2 + 30.559 * MgO + 29.290 * CaO + 30.592 * SrO + 30.079 * BaO + 23.323 * ZnO + 19.724 * MnO + 11.888 * PbO + 11.462 * CuO + 17.090 * Li_2O + 16.475 * Na_2O + 11.386 * K_2O + 14.422 * Y_2O_3 - 36.909 * La_2O_3 - 34.144 * (Fe_2O_3 + FeO) - 35.001 * (R_2O + RO - Al_2O_3), \quad (III).$$

[00181] According to a twentieth aspect, the glass of any one of aspects 1-19, wherein the glass has temperature T_{160kP} at which the glass has a viscosity of 160 kP that is greater than or equal to 1150 °C.

[00182] According to a twenty-first aspect, the glass of the twentieth aspect, wherein the temperature T_{160kP} is greater than or equal to 1200 °C.

[00183] According to a twenty-second aspect, the glass of any one of aspects 1-21, wherein the glass has an average linear thermal expansion coefficient over a temperature range 20-300°C $\alpha_{20-300} \times 10^7$ that is less than or equal to 40 K⁻¹.

[00184] According to a twenty-third aspect, the glass of any one of aspects 1-22, wherein the glass has a temperature T_{200P} at which the glass has a viscosity of 200 P that is less than or equal to 1700 °C.

[00185] According to a twenty-fourth aspect, the glass of the twenty-third aspect, wherein the temperature T_{200P} is less than or equal to 1650 °C.

[00186] According to a twenty-fifth aspect, the glass of any one of aspects 1-24, wherein P_{Al} is greater than or equal to -2.0 and P_{mod} is less than or equal to 2.0.

[00187] According to a twenty-sixth aspect, the glass comprises a plurality of components, the glass having a composition of the components comprising greater than or equal to 60.0 mol.% and less than or equal to 80.0 mol.% SiO_2 , greater than or equal to 10.0 mol.% and less than or equal to 18.0 mol.% Al_2O_3 , greater than or equal to 0.5 mol.% and less than or equal to 7.5 mol.% Li_2O , greater than or equal to 0.0 mol.% and less than or equal to 0.5 mol.% ZrO_2 , greater than or equal to 0.0 mol.% and less than or equal to 3.0 mol.% RE_mO_n , greater than or equal to 0.0 at.% and less than or equal to 0.5 at.% F and may optionally contain one or more components selected from P_2O_5 , B_2O_3 , MgO, CaO, BaO, ZnO, MnO, Na_2O , K_2O , Fe_2O_3 , FeO, Cu_2O , Rb_2O , Ag_2O , Cs_2O , Au_2O , Hg_2O , Tl_2O , BeO, CoO, NiO, CuO, SrO, CdO, SnO, PbO and TiO_2 , wherein the composition of the components satisfies the conditions: $0.00 \leq \min(RE_mO_n, P_2O_5) [\text{mol.\%}] \leq 0.30$, and wherein the glass has a cristobalite precipitation parameter P_{crist} that is less than or equal to 28, an anorthite precipitation parameter

P_{anort} that is less than or equal to 10, a cordierite precipitation parameter P_{cord} that is less than or equal to 5.0 and a spodumene precipitation parameter P_{spod} that is less than or equal to 7.5, where

P_{crist} is calculated from the glass composition in terms of mol.% of the components according to the following formula:

$$P_{crist} = SiO_2 - 6 * (Na_2O + K_2O) - 4 * Li_2O - 2 * (CaO + SrO + BaO) - 2.5 * MgO,$$

P_{anort} is calculated from the glass composition in terms of mol.% of the components according to the following formula:

$$P_{anort} = \min(CaO + SrO + 0.5 * BaO + Na_2O + 0.5 * K_2O, Al_2O_3),$$

P_{cord} is calculated from the glass composition in terms of mol.% of the components according to the following formula:

$$P_{cord} = MgO + MnO + FeO,$$

P_{spod} is calculated from the glass composition in terms of mol.% of the components according to the following formula:

$$P_{spod} = \min(Li_2O, Al_2O_3 - K_2O - 0.5 * Na_2O),$$

and wherein the glass satisfies the conditions: $P_{spm} > 32$, $P_{anpt} > 680$ and $P_{160kP} > 1150$, where P_{anpt} is an annealing point parameter, calculated from the glass composition in terms of mol.% of the components according to the Formula (I):

$$P_{anpt} = 664.7 + 5.2303 * SiO_2 - 11.493 * B_2O_3 - 7.1742 * P_2O_5 + 8.3980 * ZrO_2 - 2.0585 * MgO - 2.1088 * CaO - 3.8995 * BaO - 10.323 * ZnO - 9.0727 * MnO - 23.455 * Li_2O - 33.819 * Na_2O - 25.204 * K_2O + 15.745 * Y_2O_3 + 8.9047 * La_2O_3 - 33.960 * (Fe_2O_3 + FeO) - 5.6704 * (R_2O + RO - Al_2O_3) - 4.2545 * (SiO_2 - (6 * K_2O + 6 * Na_2O + 4 * Li_2O + 2 * RO)) - 19.439 * Cu_2O, \quad (I)$$

P_{spm} is a specific modulus parameter, calculated from the glass composition in terms of mol.% of the components according to the Formula (II):

$$P_{spm} = 32.10 + 0.47744 * SiO_2 - 1.6506 * Al_2O_3 - 0.11775 * B_2O_3 - 0.30166 * P_2O_5 + 0.20187 * TiO_2 + 0.20219 * ZrO_2 + 0.96268 * MgO + 0.83379 * CaO + 0.53685 * SrO + 0.41218 * BaO + 0.63264 * ZnO + 0.75365 * MnO + 0.62984 * CuO - 1.2496 * Na_2O - 1.5154 * K_2O + 1.4746 * Cu_2O - 0.037941 * Y_2O_3 - 0.75836 * La_2O_3 - 1.8052 * (R_2O + RO - Al_2O_3) - 0.47488 * (SiO_2 - (6 * K_2O + 6 * Na_2O + 4 * Li_2O + 2 * RO)), \quad (II)$$

P_{160kP} is a parameter predicting the temperature at which the glass has a viscosity of 160 kP, calculated from the glass composition in terms of mol.% of the components according to the Formula (III):

$$P_{160kP} = 1058 + 2.5492 * SiO_2 - 25.725 * Al_2O_3 - 11.327 * B_2O_3 - 10.014 * P_2O_5 - 14.309 * TiO_2 - 11.594 * ZrO_2 + 30.559 * MgO + 29.290 * CaO + 30.592 * SrO + 30.079 * BaO + 23.323 * ZnO + 19.724 * MnO + 11.888 * PbO + 11.462 * CuO + 17.090 * Li_2O + 16.475 * Na_2O + 11.386 * K_2O + 14.422 * Y_2O_3 - 36.909 * La_2O_3 - 34.144 * (Fe_2O_3 + FeO) - 35.001 * (R_2O + RO - Al_2O_3), \quad (III)$$

where R_2O is a total sum of monovalent metal oxides, RO is a total sum of divalent metal oxides, RE_mO_n is a total sum of rare earth metal oxides in all redox states present, and an asterisk (*) means multiplication.

[00188] According to a twenty-seventh aspect, the glass of the twenty-sixth aspect, wherein the glass has a Young's modulus E at room temperature, a density d_{RT} at room temperature, a specific modulus, E/d_{RT} that is greater than or equal to $32 \text{ GPa}\cdot\text{cm}^3/\text{g}$, an annealing point, $An.P.$ that is greater than or equal to 680°C and a temperature T_{160kP} at which the glass has a viscosity of 160 kP that is greater than or equal to 1150°C .

[00189] According to a twenty-eighth aspect, the glass of any one of aspects 26-27, wherein the composition of the components comprises greater than or equal to 60.0 mol.% and less than or equal to 75.0 mol.% SiO_2 , greater than or equal to 10.0 mol.% and less than or equal to 17.0 mol.% Al_2O_3 , greater than or equal to 0.5 mol.% and less than or equal to 7.4 mol.% Li_2O , greater than or equal to 0.0 mol.% and less than or equal to 10.0 mol.% CaO , greater than or equal to 0.0 mol.% and less than or equal to 4.8 mol.% Na_2O , greater than or equal to 0.0 mol.% and less than or equal to 4.3 mol.% MgO , greater than or equal to 0.0 mol.% and less than or equal to 3.0 mol.% B_2O_3 , greater than or equal to 0.0 mol.% and less than or equal to 3.0 mol.% TiO_2 , greater than or equal to 0.0 mol.% and less than or equal to 3.0 mol.% ZnO , greater than or equal to 2.5 mol.% Alk_2O , a sum of $MgO + CaO + SrO + BaO + ZnO$ greater than or equal to 1.0 mol.%, wherein the composition of the components is substantially free of fluorine, and wherein the glass has an aluminum-binding parameter P_{Al} that is less than or equal to 3.0 and a modifier-binding parameter P_{mod} that is greater than or equal to -3.0, where P_{Al} is calculated from the glass composition in terms of mol.% of the components according to the following formula:

$$P_{Al} = R_2O + RO + P_2O_5 + 1.6 * RE_mO_n - Al_2O_3,$$

and P_{mod} is calculated from the glass composition in terms of mol.% of the components according to the following formula:

$$P_{mod} = R_2O + RO - Al_2O_3 - P_2O_5 - RE_mO_n,$$

where Alk_2O is a total sum of alkali metal oxides.

[00190] According to a twenty-ninth aspect, the glass of any one of aspects 26-27, wherein the composition of the components comprises greater than or equal to 60.0 mol.% and less than or

equal to 75.0 mol.% SiO₂, greater than or equal to 0.3 mol.% and less than or equal to 7.5 mol.% CaO, greater than or equal to 0.0 mol.% and less than or equal to 5.0 mol.% MgO, greater than or equal to 0.0 mol.% and less than or equal to 5.0 mol.% Na₂O, greater than or equal to 0.0 mol.% and less than or equal to 4.0 mol.% P₂O₅, greater than or equal to 0.0 mol.% and less than or equal to 0.5 mol.% SnO₂ and greater than or equal to 0.0 mol.% and less than or equal to 0.3 mol.% Fe₂O₃.

[00191] According to a thirtieth aspect, the glass of any one of aspects 26-27 and 29, wherein the composition of the components comprises one or more of the following components: greater than or equal to 63.9 mol.% and less than or equal to 74.4 mol.% SiO₂, greater than or equal to 12.34 mol.% and less than or equal to 18.0 mol.% Al₂O₃, greater than or equal to 2.8 mol.% and less than or equal to 7.0 mol.% Li₂O, greater than or equal to 1.4 mol.% and less than or equal to 5.6 mol.% CaO, greater than or equal to 0.4 mol.% and less than or equal to 4.4 mol.% MgO, greater than or equal to 0.4 mol.% and less than or equal to 4.4 mol.% Na₂O, greater than or equal to 0.0 mol.% and less than or equal to 3.4 mol.% P₂O₅, greater than or equal to 0.0 mol.% and less than or equal to 1.1 mol.% SrO, greater than or equal to 0.0 mol.% and less than or equal to 0.8 mol.% B₂O₃, greater than or equal to 0.0 mol.% and less than or equal to 0.7 mol.% BaO and greater than or equal to 0 mol.% and less than or equal to 0.550 mol.% Y₂O₃.

[00192] According to a thirty-first aspect, the glass of any one of aspects 26, 27, and 29-30, wherein the composition of the components comprises greater than or equal to 64.5 mol.% and less than or equal to 74.4 mol.% SiO₂, greater than or equal to 12.34 mol.% and less than or equal to 18.0 mol.% Al₂O₃, greater than or equal to 3.2 mol.% and less than or equal to 6.2 mol.% Li₂O, greater than or equal to 1.9 mol.% and less than or equal to 5.0 mol.% CaO, greater than or equal to 1.0 mol.% and less than or equal to 4.0 mol.% MgO, greater than or equal to 0.9 mol.% and less than or equal to 4.0 mol.% Na₂O, greater than or equal to 0.0 mol.% and less than or equal to 3.0 mol.% P₂O₅, greater than or equal to 0 mol.% and less than or equal to 0.95 mol.% SrO, greater than or equal to 0.0 mol.% and less than or equal to 0.7 mol.% B₂O₃, greater than or equal to 0 mol.% and less than or equal to 0.625 mol.% BaO and greater than or equal to 0.0 mol.% and less than or equal to 0.5 mol.% Y₂O₃.

[00193] According to a thirty-second aspect, the glass of any one of aspects 26, 27, and 29, wherein the composition of the components comprises greater than or equal to 60.0 mol.% and less than or equal to 75.0 mol.% SiO₂, greater than or equal to 11.0 mol.% and less than or equal to 18.0 mol.% Al₂O₃, greater than or equal to 0.0 mol.% and less than or equal to 3.0 mol.% B₂O₃, greater than or equal to 0.0 mol.% and less than or equal to 3.0 mol.% TiO₂, greater than or equal to 0.5 mol.% Alk₂O, greater than or equal to 0.5 mol.% MgO + CaO + SrO + BaO + ZnO and wherein the

composition of the components satisfies the conditions $\min(\text{RE}_m\text{O}_n, \text{P}_2\text{O}_5) [\text{mol.}\%] \leq 0.15$, where Alk_2O is a total sum of alkali metal oxides.

[00194] According to a thirty-third aspect, the glass of any one of aspects 26-29, wherein the composition of the components comprises greater than or equal to 0.0 mol.% and less than or equal to 1.0 mol.% BaO and greater than or equal to 0.0 mol.% and less than or equal to 1.0 mol.% K₂O, wherein the composition of the components is substantially free of fluorine and substantially free of PbO and wherein the composition of the components satisfies the conditions: $0.00 \leq \min(\text{RE}_m\text{O}_n, \text{P}_2\text{O}_5) [\text{mol.}\%] \leq 0.30$.

[00195] According to a thirty-fourth aspect, the glass of any one of aspects 26-33, wherein the glass has a logarithm of liquidus viscosity, $\log(\eta_{\text{liq}} [\text{P}])$ that is greater than or equal to 5.0.

[00196] According to a thirty-fifth aspect, the logarithm of liquidus viscosity, $\log(\eta_{\text{liq}} [\text{P}])$ is greater than or equal to 5.2.

[00197] According to a thirty-sixth aspect, the glass of any one of aspects 26-35, wherein $P_{\text{anpt}} > 710$.

[00198] According to a thirty-seventh aspect, the glass of any one of aspects 26-36, wherein the glass has an annealing point An.P. that is greater than or equal to 710 °C.

[00199] According to a thirty-eighth aspect, the glass of the thirty-seventh aspect, wherein the annealing point An.P. is greater than or equal to 730 °C.

[00200] According to a thirty-ninth aspect, the glass of any one of aspects 26-38, wherein $P_{\text{spm}} > 33$.

[00201] According to a fortieth aspect, the glass of any one of aspects 26-39, wherein the glass has specific modulus, E/d_{RT} that is greater than or equal to 33 GPa·cm³/g, where E is a Young's modulus and d_{RT} is a density.

[00202] According to a forty-first aspect, the glass of the fortieth aspect, wherein the specific modulus E/d_{RT} is greater than or equal to 34 GPa·cm³/g.

[00203] According to a forty-second aspect, the glass of any one of aspects 26-41, wherein the glass satisfies the conditions: $P_{160\text{KP}} > 1200$, where $P_{160\text{KP}}$ is a parameter predicting a temperature at which the glass has a viscosity of 160 kP, calculated from the glass composition in terms of mol.% of the components according to the Formula (III):

$$P_{160\text{KP}} = 1058 + 2.5492 * \text{SiO}_2 - 25.725 * \text{Al}_2\text{O}_3 - 11.327 * \text{B}_2\text{O}_3 - 10.014 * \text{P}_2\text{O}_5 - 14.309 * \text{TiO}_2 - 11.594 * \text{ZrO}_2 + 30.559 * \text{MgO} + 29.290 * \text{CaO} + 30.592 * \text{SrO} + 30.079 * \text{BaO} + 23.323 * \text{ZnO} + 19.724 * \text{MnO} + 11.888 * \text{PbO} + 11.462 * \text{CuO} + 17.090 * \text{Li}_2\text{O} + 16.475 * \text{Na}_2\text{O} + 11.386 * \text{K}_2\text{O} + 14.422 * \text{Y}_2\text{O}_3 - 36.909 * \text{La}_2\text{O}_3 - 34.144 * (\text{Fe}_2\text{O}_3 + \text{FeO}) - 35.001 * (\text{R}_2\text{O} + \text{RO} - \text{Al}_2\text{O}_3), \quad (\text{III}).$$

[00204] According to a forty-third aspect, the glass of any one of aspects 26-42, wherein the glass has a temperature T_{160kP} at which the glass has a viscosity of 160 kP that is greater than or equal to 1200 °C.

[00205] According to a forty-fourth aspect, the glass of any one of aspects 26-43, wherein the glass has an average linear thermal expansion coefficient over a range of temperature from 20-300°C $\alpha_{20-300} \times 10^7$ that is less than or equal to 40 K⁻¹.

[00206] According to a forty-fifth aspect, the glass of any one of aspects 26-44, wherein the glass has a temperature T_{200P} at which the glass has a viscosity of 200 P that is less than or equal to 1700 °C.

[00207] According to a forty-sixth aspect, the glass of the forty-fifth aspect, wherein the temperature T_{200P} is less than or equal to 1650 °C.

[00208] According to a forty-seventh aspect, the glass of any one of aspects 26-46, wherein the glass has an aluminum-binding parameter P_{Al} that is greater than or equal to -2.0 and a modifier-binding parameter P_{mod} that is less than or equal to 2.0, where P_{Al} is calculated from the glass composition in terms of mol.% of the components according to the following formula:

$$P_{Al} = R_2O + RO + P_2O_5 + 1.6 * RE_mO_n - Al_2O_3,$$

and P_{mod} is calculated from the glass composition in terms of mol.% of the components according to the following formula:

$$P_{mod} = R_2O + RO - Al_2O_3 - P_2O_5 - RE_mO_n.$$

[00209] According to a forty-eighth aspect, the glass comprises a plurality of components, the glass having a composition of the components comprising greater than or equal to 60.0 mol.% and less than or equal to 75.0 mol.% SiO₂, greater than or equal to 10.5 mol.% and less than or equal to 18.0 mol.% Al₂O₃, greater than or equal to 0.0 mol.% and less than or equal to 10.0 mol.% CaO, greater than or equal to 0.0 mol.% and less than or equal to 7.8 mol.% Li₂O, greater than or equal to 0.0 mol.% and less than or equal to 4.3 mol.% MgO, greater than or equal to 0.0 mol.% and less than or equal to 0.5 mol.% ZrO₂, greater than or equal to 0.0 mol.% and less than or equal to 5.0 mol.% RE_mO_n and may optionally contain one or more components selected from B₂O₃, P₂O₅, TiO₂, SrO, BaO, ZnO, MnO, CuO, Na₂O, K₂O, Cu₂O, Rb₂O, Ag₂O, Cs₂O, Au₂O, Hg₂O, Tl₂O, BeO, FeO, CoO, NiO, CdO, SnO, PbO and Fe₂O₃, wherein the glass has a cristobalite precipitation parameter P_{crist} that is less than or equal to 28, an anorthite precipitation parameter P_{anort} that is less than or equal to 10 and a modifier-binding parameter P_{mod} that is greater than or equal to -3.0, where

P_{crist} is calculated from the glass composition in terms of mol.% of the components according to the following formula:

$$P_{\text{crist}} = \text{SiO}_2 - 6 * (\text{Na}_2\text{O} + \text{K}_2\text{O}) - 4 * \text{Li}_2\text{O} - 2 * (\text{CaO} + \text{SrO} + \text{BaO}) - 2.5 * \text{MgO},$$

P_{anort} is calculated from the glass composition in terms of mol.% of the components according to the following formula:

$$P_{\text{anort}} = \min(\text{CaO} + \text{SrO} + 0.5 * \text{BaO} + \text{Na}_2\text{O} + 0.5 * \text{K}_2\text{O}, \text{Al}_2\text{O}_3),$$

P_{mod} is calculated from the glass composition in terms of mol.% of the components according to the following formula:

$$P_{\text{mod}} = \text{R}_2\text{O} + \text{RO} - \text{Al}_2\text{O}_3 - \text{P}_2\text{O}_5 - \text{RE}_m\text{O}_n,$$

and wherein the glass satisfies the condition: $P_{\text{spm}} - (92.5 - 0.05 * P_{160\text{kP}}) > 0.000$, where P_{spm} is a specific modulus parameter, calculated from the glass composition in terms of mol.% of the components according to the Formula (II):

$$\begin{aligned} P_{\text{spm}} = & 32.10 + 0.47744 * \text{SiO}_2 - 1.6506 * \text{Al}_2\text{O}_3 - 0.11775 * \text{B}_2\text{O}_3 - 0.30166 * \text{P}_2\text{O}_5 + 0.20187 * \text{TiO}_2 \\ & + 0.20219 * \text{ZrO}_2 + 0.96268 * \text{MgO} + 0.83379 * \text{CaO} + 0.53685 * \text{SrO} + 0.41218 * \text{BaO} + 0.63264 \\ & * \text{ZnO} + 0.75365 * \text{MnO} + 0.62984 * \text{CuO} - 1.2496 * \text{Na}_2\text{O} - 1.5154 * \text{K}_2\text{O} + 1.4746 * \text{Cu}_2\text{O} - \quad (\text{II}) \\ & 0.037941 * \text{Y}_2\text{O}_3 - 0.75836 * \text{La}_2\text{O}_3 - 1.8052 * (\text{R}_2\text{O} + \text{RO} - \text{Al}_2\text{O}_3) - 0.47488 * (\text{SiO}_2 - (6 * \text{K}_2\text{O} + 6 * \\ & \text{Na}_2\text{O} + 4 * \text{Li}_2\text{O} + 2 * \text{RO})), \end{aligned}$$

$P_{160\text{kP}}$ is a parameter predicting a temperature at which a viscosity of the glass is 160 kP, calculated from the glass composition in terms of mol.% of the components according to the Formula (III):

$$\begin{aligned} P_{160\text{kP}} = & 1058 + 2.5492 * \text{SiO}_2 - 25.725 * \text{Al}_2\text{O}_3 - 11.327 * \text{B}_2\text{O}_3 - 10.014 * \text{P}_2\text{O}_5 - 14.309 * \text{TiO}_2 - \\ & 11.594 * \text{ZrO}_2 + 30.559 * \text{MgO} + 29.290 * \text{CaO} + 30.592 * \text{SrO} + 30.079 * \text{BaO} + 23.323 * \text{ZnO} + \quad (\text{III}) \\ & 19.724 * \text{MnO} + 11.888 * \text{PbO} + 11.462 * \text{CuO} + 17.090 * \text{Li}_2\text{O} + 16.475 * \text{Na}_2\text{O} + 11.386 * \text{K}_2\text{O} + \\ & 14.422 * \text{Y}_2\text{O}_3 - 36.909 * \text{La}_2\text{O}_3 - 34.144 * (\text{Fe}_2\text{O}_3 + \text{FeO}) - 35.001 * (\text{R}_2\text{O} + \text{RO} - \text{Al}_2\text{O}_3), \end{aligned}$$

where R_2O is a total sum of monovalent metal oxides, RO is a total sum of divalent metal oxides, RE_mO_n is a total sum of rare earth metal oxides in all redox states present, and an asterisk (*) means multiplication.

[00210] According to a forty-ninth aspect, the glass of the forty-eighth aspect, wherein the glass satisfies the conditions $E/d_{\text{RT}} - (92.5 - 0.05 * T_{160\text{kP}}) > 0.000$, where E/d_{RT} [GPa·cm³/g] is a specific modulus, E is a Young's modulus, d_{RT} is a density, $T_{160\text{kP}}$ is a temperature at which the glass has a viscosity of 160 kP.

[00211] According to a fiftieth aspect, the glass of any one of aspects 48-49, wherein the composition of the components comprises greater than or equal to 10.5 mol.% and less than or equal to 17.0 mol.% Al_2O_3 , greater than or equal to 0.0 mol.% and less than or equal to 7.4 mol.% Li_2O , greater than or equal to 0.0 mol.% and less than or equal to 4.8 mol.% Na_2O , greater than or equal to 0.0 mol.% and less than or equal to 3.0 mol.% B_2O_3 , greater than or equal to 0.0 mol.% and

less than or equal to 3.0 mol.% TiO_2 , greater than or equal to 0.0 mol.% and less than or equal to 3.0 mol.% ZnO , greater than or equal to 2.5 mol.% Alk_2O , a sum of $\text{MgO} + \text{CaO} + \text{SrO} + \text{BaO} + \text{ZnO}$ greater than or equal to 1.0 mol.%, wherein the composition of the components is substantially free of fluorine, and wherein the glass has an aluminum-binding parameter P_{Al} that is less than or equal to 3.0 and a cordierite precipitation parameter P_{cord} that is less than or equal to 8.0, where P_{Al} is calculated from the glass composition in terms of mol.% of the components according to the following formula:

$$P_{\text{Al}} = \text{R}_2\text{O} + \text{RO} + \text{P}_2\text{O}_5 + 1.6 * \text{RE}_m\text{O}_n - \text{Al}_2\text{O}_3,$$

and P_{cord} is calculated from the glass composition in terms of mol.% of the components according to the following formula:

$$P_{\text{cord}} = \text{MgO} + \text{MnO} + \text{FeO},$$

where Alk_2O is a total sum of alkali metal oxides.

[00212] According to a fifty-first aspect, the glass of any one of aspects 48-49, wherein the composition of the components comprises greater than or equal to 0.3 mol.% and less than or equal to 7.5 mol.% CaO , greater than or equal to 0.3 mol.% and less than or equal to 7.5 mol.% Li_2O , greater than or equal to 0.0 mol.% and less than or equal to 5.0 mol.% Na_2O , greater than or equal to 0.0 mol.% and less than or equal to 4.0 mol.% P_2O_5 , greater than or equal to 0.0 mol.% and less than or equal to 0.5 mol.% SnO_2 and greater than or equal to 0.0 mol.% and less than or equal to 0.3 mol.% Fe_2O_3 .

[00213] According to a fifty-second aspect, the glass of any one of aspects 48-49 and 51, wherein the composition of the components comprises one or more of the following components: greater than or equal to 63.9 mol.% and less than or equal to 74.4 mol.% SiO_2 , greater than or equal to 12.34 mol.% and less than or equal to 18.0 mol.% Al_2O_3 , greater than or equal to 2.8 mol.% and less than or equal to 7.0 mol.% Li_2O , greater than or equal to 1.4 mol.% and less than or equal to 5.6 mol.% CaO , greater than or equal to 0.4 mol.% and less than or equal to 4.4 mol.% Na_2O , greater than or equal to 0.4 mol.% and less than or equal to 4.3 mol.% MgO , greater than or equal to 0.0 mol.% and less than or equal to 3.4 mol.% P_2O_5 , greater than or equal to 0.0 mol.% and less than or equal to 1.1 mol.% SrO , greater than or equal to 0.0 mol.% and less than or equal to 0.8 mol.% B_2O_3 , greater than or equal to 0.0 mol.% and less than or equal to 0.7 mol.% BaO and greater than or equal to 0 mol.% and less than or equal to 0.550 mol.% Y_2O_3 .

[00214] According to a fifty-third aspect, the glass of any one of aspects 48-49 and 51-52, wherein the composition of the components comprises greater than or equal to 64.5 mol.% and less than or equal to 74.5 mol.% SiO_2 , greater than or equal to 12.34 mol.% and less than or equal to 18.0

mol.% Al_2O_3 , greater than or equal to 3.2 mol.% and less than or equal to 6.2 mol.% Li_2O , greater than or equal to 1.9 mol.% and less than or equal to 5.0 mol.% CaO , greater than or equal to 1.0 mol.% and less than or equal to 4.0 mol.% MgO , greater than or equal to 0.9 mol.% and less than or equal to 4.0 mol.% Na_2O , greater than or equal to 0.0 mol.% and less than or equal to 3.0 mol.% P_2O_5 , greater than or equal to 0 mol.% and less than or equal to 0.95 mol.% SrO , greater than or equal to 0.0 mol.% and less than or equal to 0.7 mol.% B_2O_3 , greater than or equal to 0 mol.% and less than or equal to 0.625 mol.% BaO and greater than or equal to 0.0 mol.% and less than or equal to 0.5 mol.% Y_2O_3 .

[00215] According to a fifty-fourth aspect, the glass of any one of aspects 48, 49, and 51, wherein the composition of the components comprises greater than or equal to 11.0 mol.% and less than or equal to 18.0 mol.% Al_2O_3 , greater than or equal to 0.5 mol.% and less than or equal to 7.5 mol.% Li_2O , greater than or equal to 0.0 mol.% and less than or equal to 3.0 mol.% B_2O_3 , greater than or equal to 0.0 mol.% and less than or equal to 3.0 mol.% TiO_2 , greater than or equal to 0.5 mol.% Alk_2O , greater than or equal to 0.0 at.% and less than or equal to 0.5 at.% F, a sum of $\text{MgO} + \text{CaO} + \text{SrO} + \text{BaO} + \text{ZnO}$ greater than or equal to 0.5 mol.%, wherein the composition of the components satisfies the condition: $\min(\text{RE}_m\text{O}_n, \text{P}_2\text{O}_5) [\text{mol.}\%] \leq 0.15$, and wherein the glass has an annealing point An.P. that is greater than or equal to 680°C , a temperature corresponding to the at which the glass has a viscosity of 160 kP $T_{160\text{kP}}$ that is greater than or equal to 1150°C , a Young's modulus E , a density d_{RT} , a specific modulus E/d_{RT} that is greater than or equal to $32 \text{ GPa}\cdot\text{cm}^3/\text{g}$, a cordierite precipitation parameter P_{cord} that is less than or equal to 5.0 and a spodumene precipitation parameter P_{spod} that is less than or equal to 7.5, where P_{cord} is calculated from the glass composition in terms of mol.% of the components according to the following formula:

$$P_{\text{cord}} = \text{MgO} + \text{MnO} + \text{FeO},$$

and P_{spod} is calculated from the glass composition in terms of mol.% of the components according to the following formula:

$$P_{\text{spod}} = \min(\text{Li}_2\text{O}, \text{Al}_2\text{O}_3 - \text{K}_2\text{O} - 0.5 * \text{Na}_2\text{O}),$$

where Alk_2O is a total sum of alkali metal oxides.

[00216] According to a fifty-fifth aspect, the glass of any one of aspects 48-52 and 54, wherein the composition of the components comprises greater than or equal to 0.0 mol.% and less than or equal to 1.0 mol.% BaO and greater than or equal to 0.0 mol.% and less than or equal to 1.0 mol.% K_2O , wherein the composition of the components is substantially free of fluorine and substantially free of PbO and wherein the composition of the components satisfies the conditions: $0.00 \leq \min(\text{RE}_m\text{O}_n, \text{P}_2\text{O}_5) [\text{mol.}\%] \leq 0.30$.

[00217] According to a fifty-sixth aspect, the glass of any one of aspects 48-55, wherein the glass has a logarithm of liquidus viscosity $\log(\eta_{\text{liq}} [P])$ that is greater than or equal to 5.0.

[00218] According to a fifty-seventh aspect, the glass of the fifty-sixth aspect, wherein the logarithm of liquidus viscosity $\log(\eta_{\text{liq}} [P])$ is greater than or equal to 5.2.

[00219] According to a fifty-eighth aspect, the glass of any one of aspects 48-57, wherein the glass satisfies the condition: $P_{\text{anpt}} > 710$, where P_{anpt} is annealing point parameter, calculated from the glass composition in terms of mol.% of the components according to the Formula (I):

$$P_{\text{anpt}} = 664.7 + 5.2303 * \text{SiO}_2 - 11.493 * \text{B}_2\text{O}_3 - 7.1742 * \text{P}_2\text{O}_5 + 8.3980 * \text{ZrO}_2 - 2.0585 * \text{MgO} - 2.1088 * \text{CaO} - 3.8995 * \text{BaO} - 10.323 * \text{ZnO} - 9.0727 * \text{MnO} - 23.455 * \text{Li}_2\text{O} - 33.819 * \text{Na}_2\text{O} - 25.204 * \text{K}_2\text{O} + 15.745 * \text{Y}_2\text{O}_3 + 8.9047 * \text{La}_2\text{O}_3 - 33.960 * (\text{Fe}_2\text{O}_3 + \text{FeO}) - 5.6704 * (\text{R}_2\text{O} + \text{RO} - \text{Al}_2\text{O}_3) - 4.2545 * (\text{SiO}_2 - (6 * \text{K}_2\text{O} + 6 * \text{Na}_2\text{O} + 4 * \text{Li}_2\text{O} + 2 * \text{RO})) - 19.439 * \text{Cu}_2\text{O}, \quad (I).$$

[00220] According to a fifty-ninth aspect, the glass of any one of aspects 48-58, wherein the glass has annealing point An.P. that is greater than or equal to 710 °C.

[00221] According to a sixtieth aspect, the glass of the fifty-ninth aspect, wherein the annealing point An.P. is greater than or equal to 730 °C.

[00222] According to a sixty-first aspect, the glass of any one of aspects 48-60, wherein the glass has cordierite precipitation parameter, P_{cord} that is less than or equal to 5.0 and spodumene precipitation parameter, P_{spod} that is less than or equal to 7.5, where P_{cord} is a value of cordierite precipitation parameter, calculated from the glass composition in terms of mol.% of the components according to the following formula:

$$P_{\text{cord}} = \text{MgO} + \text{MnO} + \text{FeO},$$

P_{spod} is a value of spodumene precipitation parameter, calculated from the glass composition in terms of mol.% of the components according to the following formula:

$$P_{\text{spod}} = \min(\text{Li}_2\text{O}, \text{Al}_2\text{O}_3 - \text{K}_2\text{O} - 0.5 * \text{Na}_2\text{O}).$$

[00223] According to a sixty-second aspect, the glass of any one of aspects 48-61, wherein $P_{\text{spm}} > 32$.

[00224] According to a sixty-third aspect, the glass of any one of aspects 48-62, wherein the glass has a specific modulus E/d_{RT} that is greater than or equal to 32 GPa·cm³/g, where E is a Young's modulus and d_{RT} is a density.

[00225] According to a sixty-fourth aspect, the glass of the sixty-third aspect, wherein the specific modulus E/d_{RT} is greater than or equal to 33 GPa·cm³/g.

[00226] According to a sixty-fifth aspect, the glass of the sixty-fourth aspect, wherein the specific modulus, E/d_{RT} is greater than or equal to 34 GPa·cm³/g.

[00227] According to a sixty-sixth aspect, the glass of the sixty-fifth aspect, wherein the specific modulus E/d_{RT} is greater than or equal to 35 GPa.cm³/g.

[00228] According to a sixty-seventh aspect, the glass of any one of aspects 48-66, wherein $P_{160kP} > 1150$.

[00229] According to a sixty-eighth aspect, the glass of any one of aspects 48-67, wherein the glass has a temperature T_{160kP} at which the viscosity of the glass is 160 kP that is greater than or equal to 1150 °C.

[00230] According to a sixty-ninth aspect, the glass of the sixty-eighth aspect, wherein the temperature T_{160kP} is greater than or equal to 1200 °C.

[00231] According to a seventieth aspect, the glass of any one of aspects 48-69, wherein the glass has an average linear thermal expansion coefficient over a temperature range from 20-300°C $\alpha_{20-300} \times 10^7$ that is less than or equal to 40 K⁻¹.

[00232] According to a seventy-first aspect, the glass of any one of aspects 48-70, wherein the glass has a temperature T_{200P} at which the glass has a viscosity of 200 P that is less than or equal to 1700 °C.

[00233] According to a seventy-second aspect, the glass of the seventy-first aspect, wherein the temperature T_{200P} is less than or equal to 1650 °C.

[00234] According to a seventy-third aspect, the glass of any one of aspects 48-72, wherein the glass has an aluminum-binding parameter P_{Al} that is greater than or equal to -2.0 and a modifier-binding parameter P_{mod} that is less than or equal to 2.0, where P_{Al} is calculated from the glass composition in terms of mol.% of the components according to the following formula:

$$P_{Al} = R_2O + RO + P_2O_5 + 1.6 * RE_mO_n - Al_2O_3,$$

and P_{mod} is calculated from the glass composition in terms of mol.% of the components according to the following formula:

$$P_{mod} = R_2O + RO - Al_2O_3 - P_2O_5 - RE_mO_n,$$

[00235] Many variations and modifications may be made to the above-described embodiments of the disclosure without departing substantially from the spirit and various principles of the disclosure. All such modifications and variations are intended to be included herein within the scope of this disclosure and protected by the following claims.

[00236] To the extent not already described, the different features of the various aspects of the present disclosure may be used in combination with each other as desired. That a particular feature is not explicitly illustrated or described with respect to each aspect of the present disclosure is not meant to be construed that it cannot be, but it is done for the sake of brevity and conciseness

of the description. Thus, the various features of the different aspects may be mixed and matched as desired to form new aspects, whether or not the new aspects are expressly disclosed.

CLAIMS:

1. A glass comprising a plurality of components, the glass having a composition of the components comprising:

- greater than or equal to 60.0 mol.% and less than or equal to 75.0 mol.% SiO_2 ,
- greater than or equal to 10.0 mol.% and less than or equal to 18.0 mol.% Al_2O_3 ,
- greater than or equal to 0.0 mol.% and less than or equal to 10.0 mol.% Li_2O ,
- greater than or equal to 0.0 mol.% and less than or equal to 5.0 mol.% Na_2O ,
- greater than or equal to 0.0 mol.% and less than or equal to 3.8 mol.% B_2O_3 ,
- greater than or equal to 0.0 mol.% and less than or equal to 3.0 mol.% P_2O_5 ,
- greater than or equal to 0.0 mol.% and less than or equal to 2.0 mol.% ZnO ,
- greater than or equal to 0.0 mol.% and less than or equal to 0.7 mol.% La_2O_3 ,
- greater than or equal to 0.0 at.% and less than or equal to 3.0 at.% F,
- a sum of $\text{CaO} + \text{MgO}$ greater than or equal to 5.0 mol.%,
- a sum of $\text{Li}_2\text{O} + \text{Na}_2\text{O}$ greater than or equal to 0.5 mol.%,
- a sum of $\text{Li}_2\text{O} + \text{MgO}$ greater than or equal to 0.0 mol.% and less than or equal to 10.0 mol.%,
- a sum of $\text{MgO} + \text{ZnO}$ greater than or equal to 0.0 mol.% and less than or equal to 10.0 mol.%,
- a sum of $\text{CaO} + \text{SrO}$ greater than or equal to 0.0 mol.% and less than or equal to 9.0 mol.%, and
- a sum of $\text{ZrO}_2 + \text{TiO}_2 + \text{FeO} + \text{Fe}_2\text{O}_3$ greater than or equal to 0.0 mol.% and less than or equal to 1.5 mol.%,

wherein the glass has

- an aluminum-binding parameter P_{Al} that is greater than or equal to -2.8,
- a modifier-binding parameter P_{mod} that is less than or equal to 2.8 and
- an anorthite precipitation parameter P_{anort} that is less than or equal to 10,

where

- P_{Al} is calculated from the glass composition in terms of mol.% of the components according to the following formula:

$$P_{Al} = R_2O + RO + P_2O_5 + 1.6 * RE_mO_n - Al_2O_3,$$

- P_{mod} is calculated from the glass composition in terms of mol.% of the components according to the following formula:

$$P_{mod} = R_2O + RO - Al_2O_3 - P_2O_5 - RE_mO_n, \text{ and}$$

- P_{anort} is calculated from the glass composition in terms of mol.% of the components according to the following formula:

$$P_{anort} = \min(CaO + SrO + 0.5 * BaO + Na_2O + 0.5 * K_2O, Al_2O_3),$$

where R_2O is a total sum of monovalent metal oxides, RO is a total sum of divalent metal oxides, RE_mO_n is a total sum of rare earth metal oxides in all redox states present, and an asterisk (*) means multiplication.

2. The glass of claim 1, wherein the composition of the components comprises:

- greater than or equal to 0.3 mol.% and less than or equal to 7.5 mol.% CaO ,
- greater than or equal to 0.3 mol.% and less than or equal to 7.5 mol.% Li_2O ,
- greater than or equal to 0.0 mol.% and less than or equal to 5.0 mol.% MgO ,
- greater than or equal to 0.0 mol.% and less than or equal to 0.5 mol.% SnO_2 , and
- greater than or equal to 0.0 mol.% and less than or equal to 0.3 mol.% Fe_2O_3 .

3. The glass of any one of claims 1-2, wherein the glass has

- a logarithm of liquidus viscosity $\log(\eta_{liq} [P])$ that is greater than or equal to 5.0.

4. The glass of any one of claims 1-3, wherein the glass has

- a cordierite precipitation parameter P_{cord} that is less than or equal to 5.0,
- a cristobalite precipitation parameter P_{crist} that is less than or equal to 28 and
- spodumene precipitation parameter P_{spod} that is less than or equal to 7.5,

where

- P_{cord} is calculated from the glass composition in terms of mol.% of the components according to the following formula:

$$P_{\text{cord}} = \text{MgO} + \text{MnO} + \text{FeO},$$

- P_{crist} is calculated from the glass composition in terms of mol.% of the components according to the following formula:

$$P_{\text{crist}} = \text{SiO}_2 - 6 * (\text{Na}_2\text{O} + \text{K}_2\text{O}) - 4 * \text{Li}_2\text{O} - 2 * (\text{CaO} + \text{SrO} + \text{BaO}) - 2.5 * \text{MgO},$$

- P_{spod} is calculated from the glass composition in terms of mol.% of the components according to the following formula:

$$P_{\text{spod}} = \min(\text{Li}_2\text{O}, \text{Al}_2\text{O}_3 - \text{K}_2\text{O} - 0.5 * \text{Na}_2\text{O}).$$

5. The glass of any one of claims 1-4, wherein the glass satisfies the condition:

- $P_{\text{spm}} > 32$,

where

- P_{spm} is a specific modulus parameter, calculated from the glass composition in terms of mol.% of the components according to the Formula (II):

$$\begin{aligned} P_{\text{spm}} = & 32.10 + 0.47744 * \text{SiO}_2 - 1.6506 * \text{Al}_2\text{O}_3 - 0.11775 * \text{B}_2\text{O}_3 - 0.30166 * \text{P}_2\text{O}_5 + 0.20187 * \\ & \text{TiO}_2 + 0.20219 * \text{ZrO}_2 + 0.96268 * \text{MgO} + 0.83379 * \text{CaO} + 0.53685 * \text{SrO} + 0.41218 * \text{BaO} + \\ & 0.63264 * \text{ZnO} + 0.75365 * \text{MnO} + 0.62984 * \text{CuO} - 1.2496 * \text{Na}_2\text{O} - 1.5154 * \text{K}_2\text{O} + 1.4746 * \\ & \text{Cu}_2\text{O} - 0.037941 * \text{Y}_2\text{O}_3 - 0.75836 * \text{La}_2\text{O}_3 - 1.8052 * (\text{R}_2\text{O} + \text{RO} - \text{Al}_2\text{O}_3) - 0.47488 * (\text{SiO}_2 - (6 * \\ & \text{K}_2\text{O} + 6 * \text{Na}_2\text{O} + 4 * \text{Li}_2\text{O} + 2 * \text{RO})), \end{aligned} \quad (\text{II}).$$

6. The glass of any one of claims 1-5, wherein the glass satisfies the condition:

- $P_{160\text{kP}} > 1150$,

where

- $P_{160\text{kP}}$ is a parameter predicting the temperature at which the glass has a viscosity of 160 kP, calculated from the glass composition in terms of mol.% of the components according to the Formula (III):

$$\begin{aligned} P_{160\text{kP}} = & 1058 + 2.5492 * \text{SiO}_2 - 25.725 * \text{Al}_2\text{O}_3 - 11.327 * \text{B}_2\text{O}_3 - 10.014 * \text{P}_2\text{O}_5 - 14.309 * \text{TiO}_2 - \\ & 11.594 * \text{ZrO}_2 + 30.559 * \text{MgO} + 29.290 * \text{CaO} + 30.592 * \text{SrO} + 30.079 * \text{BaO} + 23.323 * \text{ZnO} + \\ & 19.724 * \text{MnO} + 11.888 * \text{PbO} + 11.462 * \text{CuO} + 17.090 * \text{Li}_2\text{O} + 16.475 * \text{Na}_2\text{O} + 11.386 * \text{K}_2\text{O} \\ & + 14.422 * \text{Y}_2\text{O}_3 - 36.909 * \text{La}_2\text{O}_3 - 34.144 * (\text{Fe}_2\text{O}_3 + \text{FeO}) - 35.001 * (\text{R}_2\text{O} + \text{RO} - \text{Al}_2\text{O}_3), \end{aligned} \quad (\text{III}).$$

7. The glass of any one of claims 1-6, wherein the glass has

- a temperature T_{200P} at which the glass has a viscosity of 200 Poise that is less than or equal to 1700 °C.

8. A glass comprising a plurality of components, the glass having a composition of the components comprising:

- greater than or equal to 60.0 mol.% and less than or equal to 80.0 mol.% SiO_2 ,
- greater than or equal to 10.0 mol.% and less than or equal to 18.0 mol.% Al_2O_3 ,
- greater than or equal to 0.5 mol.% and less than or equal to 7.5 mol.% Li_2O ,
- greater than or equal to 0.0 mol.% and less than or equal to 0.5 mol.% ZrO_2 ,
- greater than or equal to 0.0 mol.% and less than or equal to 3.0 mol.% RE_mO_n ,
- greater than or equal to 0.0 at.% and less than or equal to 0.5 at.% F and
- optionally comprising one or more components selected from P_2O_5 , B_2O_3 , MgO , CaO , BaO , ZnO , MnO , Na_2O , K_2O , Fe_2O_3 , FeO , Cu_2O , Rb_2O , Ag_2O , Cs_2O , Au_2O , Hg_2O , Tl_2O , BeO , CoO , NiO , CuO , SrO , CdO , SnO , PbO and TiO_2 ,

wherein the composition of the components satisfies the condition:

- $0.00 \leq \min(\text{RE}_m\text{O}_n, \text{P}_2\text{O}_5) [\text{mol.}\%] \leq 0.30$,

and wherein the glass has

- a cristobalite precipitation parameter P_{crist} that is less than or equal to 28,
- an anorthite precipitation parameter P_{anort} that is less than or equal to 10,
- a cordierite precipitation parameter P_{cord} that is less than or equal to 5.0 and
- a spodumene precipitation parameter P_{spod} that is less than or equal to 7.5,

where

- P_{crist} is calculated from the glass composition in terms of mol.% of the components according to the following formula:

$$P_{\text{crist}} = \text{SiO}_2 - 6 * (\text{Na}_2\text{O} + \text{K}_2\text{O}) - 4 * \text{Li}_2\text{O} - 2 * (\text{CaO} + \text{SrO} + \text{BaO}) - 2.5 * \text{MgO},$$

- P_{anort} is calculated from the glass composition in terms of mol.% of the components according to the following formula:

$$P_{\text{anort}} = \min(\text{CaO} + \text{SrO} + 0.5 * \text{BaO} + \text{Na}_2\text{O} + 0.5 * \text{K}_2\text{O}, \text{Al}_2\text{O}_3),$$

- P_{cord} is calculated from the glass composition in terms of mol.% of the components according to the following formula:

$$P_{\text{cord}} = \text{MgO} + \text{MnO} + \text{FeO},$$

- P_{spod} is calculated from the glass composition in terms of mol.% of the components according to the following formula:

$$P_{\text{spod}} = \min(\text{Li}_2\text{O}, \text{Al}_2\text{O}_3 - \text{K}_2\text{O} - 0.5 * \text{Na}_2\text{O})$$

and wherein the glass satisfies the conditions:

- $P_{\text{spm}} > 32$,
- $P_{\text{anpt}} > 680$ and
- $P_{160\text{kP}} > 1150$,

where

- P_{anpt} is an annealing point parameter, calculated from the glass composition in terms of mol.% of the components according to the Formula (I):

$$\begin{aligned} P_{\text{anpt}} = & 664.7 + 5.2303 * \text{SiO}_2 - 11.493 * \text{B}_2\text{O}_3 - 7.1742 * \text{P}_2\text{O}_5 + 8.3980 * \text{ZrO}_2 - 2.0585 * \text{MgO} - \\ & 2.1088 * \text{CaO} - 3.8995 * \text{BaO} - 10.323 * \text{ZnO} - 9.0727 * \text{MnO} - 23.455 * \text{Li}_2\text{O} - 33.819 * \text{Na}_2\text{O} - \\ & 25.204 * \text{K}_2\text{O} + 15.745 * \text{Y}_2\text{O}_3 + 8.9047 * \text{La}_2\text{O}_3 - 33.960 * (\text{Fe}_2\text{O}_3 + \text{FeO}) - 5.6704 * (\text{R}_2\text{O} + \text{RO} - \\ & \text{Al}_2\text{O}_3) - 4.2545 * (\text{SiO}_2 - (6 * \text{K}_2\text{O} + 6 * \text{Na}_2\text{O} + 4 * \text{Li}_2\text{O} + 2 * \text{RO})) - 19.439 * \text{Cu}_2\text{O}, \end{aligned} \quad (\text{I})$$

- P_{spm} is a specific modulus parameter, calculated from the glass composition in terms of mol.% of the components according to the Formula (II):

$$\begin{aligned} P_{\text{spm}} = & 32.10 + 0.47744 * \text{SiO}_2 - 1.6506 * \text{Al}_2\text{O}_3 - 0.11775 * \text{B}_2\text{O}_3 - 0.30166 * \text{P}_2\text{O}_5 + 0.20187 * \text{TiO}_2 \\ & + 0.20219 * \text{ZrO}_2 + 0.96268 * \text{MgO} + 0.83379 * \text{CaO} + 0.53685 * \text{SrO} + 0.41218 * \text{BaO} + 0.63264 \\ & * \text{ZnO} + 0.75365 * \text{MnO} + 0.62984 * \text{CuO} - 1.2496 * \text{Na}_2\text{O} - 1.5154 * \text{K}_2\text{O} + 1.4746 * \text{Cu}_2\text{O} - \\ & 0.037941 * \text{Y}_2\text{O}_3 - 0.75836 * \text{La}_2\text{O}_3 - 1.8052 * (\text{R}_2\text{O} + \text{RO} - \text{Al}_2\text{O}_3) - 0.47488 * (\text{SiO}_2 - (6 * \text{K}_2\text{O} + 6 * \\ & \text{Na}_2\text{O} + 4 * \text{Li}_2\text{O} + 2 * \text{RO})), \end{aligned} \quad (\text{II})$$

- $P_{160\text{kP}}$ is a parameter predicting the temperature at which the glass has a viscosity of 160 kP, calculated from the glass composition in terms of mol.% of the components according to the Formula (III):

$$\begin{aligned} P_{160\text{kP}} = & 1058 + 2.5492 * \text{SiO}_2 - 25.725 * \text{Al}_2\text{O}_3 - 11.327 * \text{B}_2\text{O}_3 - 10.014 * \text{P}_2\text{O}_5 - 14.309 * \text{TiO}_2 - \\ & 11.594 * \text{ZrO}_2 + 30.559 * \text{MgO} + 29.290 * \text{CaO} + 30.592 * \text{SrO} + 30.079 * \text{BaO} + 23.323 * \text{ZnO} + \\ & 19.724 * \text{MnO} + 11.888 * \text{PbO} + 11.462 * \text{CuO} + 17.090 * \text{Li}_2\text{O} + 16.475 * \text{Na}_2\text{O} + 11.386 * \text{K}_2\text{O} + \\ & 14.422 * \text{Y}_2\text{O}_3 - 36.909 * \text{La}_2\text{O}_3 - 34.144 * (\text{Fe}_2\text{O}_3 + \text{FeO}) - 35.001 * (\text{R}_2\text{O} + \text{RO} - \text{Al}_2\text{O}_3), \end{aligned} \quad (\text{III})$$

where R_2O is a total sum of monovalent metal oxides, RO is a total sum of divalent metal oxides, RE_mO_n is a total sum of rare earth metal oxides in all redox states present, and an asterisk (*) means multiplication.

9. The glass of claim 8, wherein

- P_{crist} is less than or equal to 28,
- P_{anort} is less than or equal to 10,
- P_{cord} is less than or equal to 5.0,
- P_{spod} is less than or equal to 7.5,

and wherein the glass has

- a specific modulus E/d_{RT} that is greater than or equal to $32 \text{ GPa}\cdot\text{cm}^3/\text{g}$,
- an annealing point An.P. that is greater than or equal to 680°C and
- a temperature $T_{160\text{kP}}$ at which the glass has a viscosity of 160 kP that is greater than or equal to 1150°C ,

where E is a Young's modulus and d_{RT} is a density.

10. The glass of any one of claims 8-9, wherein the composition of the components comprises:

- greater than or equal to 60.0 mol.% and less than or equal to 75.0 mol.% SiO_2 ,
- greater than or equal to 0.3 mol.% and less than or equal to 7.5 mol.% CaO ,
- greater than or equal to 0.0 mol.% and less than or equal to 5.0 mol.% MgO ,
- greater than or equal to 0.0 mol.% and less than or equal to 5.0 mol.% Na_2O ,
- greater than or equal to 0.0 mol.% and less than or equal to 4.0 mol.% P_2O_5 ,
- greater than or equal to 0.0 mol.% and less than or equal to 0.5 mol.% SnO_2 and
- greater than or equal to 0.0 mol.% and less than or equal to 0.3 mol.% Fe_2O_3 .

11. The glass of any one of claims 8-10, wherein the glass has

- a logarithm of liquidus viscosity $\log(\eta_{\text{liq}} [P])$ that is greater than or equal to 5.0.

12. The glass of any one of claims 8-11, wherein

- $P_{\text{spm}} > 33$.

13. The glass of any one of claims 8-12, wherein the glass has

- a temperature T_{200P} at which the glass has a viscosity of 200 Poise that is less than or equal to 1700 °C.

14. A glass comprising a plurality of components, the glass having a composition of the components comprising:

- greater than or equal to 60.0 mol.% and less than or equal to 75.0 mol.% SiO_2 ,
- greater than or equal to 10.5 mol.% and less than or equal to 18.0 mol.% Al_2O_3 ,
- greater than or equal to 0.0 mol.% and less than or equal to 10.0 mol.% CaO ,
- greater than or equal to 0.0 mol.% and less than or equal to 7.8 mol.% Li_2O ,
- greater than or equal to 0.0 mol.% and less than or equal to 4.3 mol.% MgO ,
- greater than or equal to 0.0 mol.% and less than or equal to 0.5 mol.% ZrO_2 ,
- greater than or equal to 0.0 mol.% and less than or equal to 5.0 mol.% RE_mO_n and
- optionally comprising one or more components selected from B_2O_3 , P_2O_5 , TiO_2 , SrO , BaO , ZnO , MnO , CuO , Na_2O , K_2O , Cu_2O , Rb_2O , Ag_2O , Cs_2O , Au_2O , Hg_2O , Tl_2O , BeO , FeO , CoO , NiO , CdO , SnO , PbO and Fe_2O_3 ,

wherein the glass has

- a cristobalite precipitation parameter P_{crist} that is less than or equal to 28,
- an anorthite precipitation parameter P_{anort} that is less than or equal to 10 and
- a modifier-binding parameter P_{mod} that is greater than or equal to -3.0 and

wherein the glass satisfies the condition:

- $P_{\text{spm}} - (92.5 - 0.05 * P_{160kP}) > 0.000$,

where

- P_{spm} is a specific modulus parameter, calculated from the glass composition in terms of mol.% of the components according to the Formula (II):

$$P_{\text{spm}} = 32.10 + 0.47744 * \text{SiO}_2 - 1.6506 * \text{Al}_2\text{O}_3 - 0.11775 * \text{B}_2\text{O}_3 - 0.30166 * \text{P}_2\text{O}_5 + 0.20187 * \text{TiO}_2 + 0.20219 * \text{ZrO}_2 + 0.96268 * \text{MgO} + 0.83379 * \text{CaO} + 0.53685 * \text{SrO} + 0.41218 * \text{BaO} + 0.63264 * \text{ZnO} + 0.75365 * \text{MnO} + 0.62984 * \text{CuO} - 1.2496 * \text{Na}_2\text{O} - 1.5154 * \text{K}_2\text{O} + 1.4746 * \text{Cu}_2\text{O} - 0.037941 * \text{Y}_2\text{O}_3 - 0.75836 * \text{La}_2\text{O}_3 - 1.8052 * (\text{R}_2\text{O} + \text{RO} - \text{Al}_2\text{O}_3) - 0.47488 * (\text{SiO}_2 - (6 * \text{K}_2\text{O} + 6 * \text{Na}_2\text{O} + 4 * \text{Li}_2\text{O} + 2 * \text{RO})), \quad (\text{II})$$

- P_{160kP} is a parameter predicting a temperature at which the glass has a viscosity of 160 kP, calculated from the glass composition in terms of mol.% of the components according to the Formula (III):

$$P_{160kP} = 1058 + 2.5492 * SiO_2 - 25.725 * Al_2O_3 - 11.327 * B_2O_3 - 10.014 * P_2O_5 - 14.309 * TiO_2 - 11.594 * ZrO_2 + 30.559 * MgO + 29.290 * CaO + 30.592 * SrO + 30.079 * BaO + 23.323 * ZnO + 19.724 * MnO + 11.888 * PbO + 11.462 * CuO + 17.090 * Li_2O + 16.475 * Na_2O + 11.386 * K_2O + 14.422 * Y_2O_3 - 36.909 * La_2O_3 - 34.144 * (Fe_2O_3 + FeO) - 35.001 * (R_2O + RO - Al_2O_3), \quad (III)$$

- P_{crist} is a value of a cristobalite precipitation parameter, calculated from the glass composition in terms of mol.% of the components according to the following formula:

$$P_{crist} = SiO_2 - 6 * (Na_2O + K_2O) - 4 * Li_2O - 2 * (CaO + SrO + BaO) - 2.5 * MgO,$$

- P_{anort} is a value of an anorthite precipitation parameter, calculated from the glass composition in terms of mol.% of the components according to the following formula:

$$P_{anort} = \min(CaO + SrO + 0.5 * BaO + Na_2O + 0.5 * K_2O, Al_2O_3),$$

- P_{mod} is a value of a modifier-binding parameter, calculated from the glass composition in terms of mol.% of the components according to the following formula:

$$P_{mod} = R_2O + RO - Al_2O_3 - P_2O_5 - RE_mO_n,$$

where R_2O is a total sum of monovalent metal oxides, RO is a total sum of divalent metal oxides, RE_mO_n is a total sum of rare earth metal oxides in all redox states present, and an asterisk (*) means multiplication.

15. The glass of claim 14, wherein

- P_{crist} is less than or equal to 28,
- P_{anort} is less than or equal to 10 and
- P_{mod} is greater than or equal to -3.0 and

wherein the glass satisfies the condition:

- $E/d_{RT} - (92.5 - 0.05 * T_{160kP}) > 0.000,$

where

- E/d_{RT} [GPa·cm³/g] is a specific modulus of the glass, E is a Young's modulus and d_{RT} is a density, and
- T_{160kP} [°C] is a temperature at which the glass has a viscosity of 160 kP.

16. The glass of any one of claims 14-15, wherein the composition of the components comprises:

- greater than or equal to 0.3 mol.% and less than or equal to 7.5 mol.% CaO,
- greater than or equal to 0.3 mol.% and less than or equal to 7.5 mol.% Li₂O,
- greater than or equal to 0.0 mol.% and less than or equal to 5.0 mol.% Na₂O,
- greater than or equal to 0.0 mol.% and less than or equal to 4.0 mol.% P₂O₅,
- greater than or equal to 0.0 mol.% and less than or equal to 0.5 mol.% SnO₂ and
- greater than or equal to 0.0 mol.% and less than or equal to 0.3 mol.% Fe₂O₃.

17. The glass of any one of claims 14-16, wherein the glass has

- a logarithm of liquidus viscosity, $\log(\eta_{\text{liq}} [P])$ that is greater than or equal to 5.0.

18. The glass of any one of claims 14-17, wherein the glass has

- a cordierite precipitation parameter P_{cord} that is less than or equal to 5.0 and
- a spodumene precipitation parameter P_{spod} that is less than or equal to 7.5

where

- P_{cord} is calculated from the glass composition in terms of mol.% of the components according to the following formula:

$$P_{\text{cord}} = \text{MgO} + \text{MnO} + \text{FeO},$$

- P_{spod} is calculated from the glass composition in terms of mol.% of the components according to the following formula:

$$P_{\text{spod}} = \min(\text{Li}_2\text{O}, \text{Al}_2\text{O}_3 - \text{K}_2\text{O} - 0.5 * \text{Na}_2\text{O}).$$

19. The glass of any one of claims 14-18, wherein

- $P_{\text{spm}} > 32$.

20. The glass of any one of claims 14-19, wherein the glass has a temperature T_{200P} at which the glass has a viscosity of 200 Poises that is less than or equal to 1700 °C.