

PROCESS FOR METHACRYLIC ACID PRODUCTION

FIELD OF THE INVENTION

The invention relates to a method for preparing methacrylic acid from methacrolein
5 using a heterogeneous catalyst.

BACKGROUND OF THE INVENTION

Processes for preparing methacrylic acid are known. For example, WO 2014/146961
discloses a process for preparation of methacrylic acid or a methacrylic acid by the gas
10 phase oxidation of at least one C₄ compound.

EP 3144291 discloses a process for the preparation of alkyl methacrylates and
methacrylic acid in which methacrolein is synthesized in a first reactor, the methacrolein is
subjected to oxidative esterification in a second reactor to form an alkyl methacrylate, and at
least part of the alkyl methacrylate is reacted with water to form methacrylic acid in a third
15 reactor.

It is desirable to maximize selectivity and reduce the formation of all byproducts in
the efficient production of methacrylic acid.

SUMMARY OF THE INVENTION

20 The invention is directed to a process for producing methacrylic acid comprising:

a) producing methacrolein from propionaldehyde and formaldehyde;

b) producing methacrylic acid in an oxidative reaction from the methacrolein
produced in step a) and water;

wherein:

25 step b) is performed at a pressure above 1 bar;

step b) is performed in a reactor system in a liquid phase reaction in the presence of
a heterogeneous noble metal-containing catalyst, wherein the reactor system comprises an
oxygen-containing gas;

an average concentration of methacrolein in step b) is less than 40 wt% based on the
30 total weight of water and methacrolein; and

the reactor system of step b) has an average ratio of water to methacrolein less than
40:1 based on an average amount of water and methacrolein entering and exiting the
system.

DETAILED DESCRIPTION OF THE INVENTION

All percentage compositions are weight percentages (wt%), and all temperatures are in °C, unless otherwise indicated. Averages are arithmetic averages unless otherwise indicated. An “average concentration” is the arithmetic average of the concentration
5 entering a region and the concentration exiting the region, where the region is an individual reactor, a reactor system, or a zone within a reactor or reactor system. An “average ratio” is the ratio of the average concentration of one component relative to the average concentration of another component. For example, the average ratio of water to
10 methacrolein in a reactor system is calculated by dividing the average concentration of water entering and exiting the reactor system by the average concentration of methacrolein entering and exiting the reactor system.

A noble metal is any of gold, platinum, iridium, osmium, silver, palladium, rhodium and ruthenium. More than one noble metal may be present in the catalyst, in which case the limits apply to the total of all noble metals.

15 The “catalyst center” is the centroid of the catalyst particle, i.e., the mean position of all points in all coordinate directions. A diameter is any linear dimension passing through the catalyst center and the average diameter is the arithmetic mean of all possible diameters. The aspect ratio is the ratio of the longest to the shortest diameters.

A reactor system refers to one or more reactors where a designated reaction takes
20 place. For example, the oxidative reaction of methacrolein to produce methacrylic acid may be the designated reaction that takes place in the reactor system. The reactor system may comprise a single reactor or a plurality of reactors. Additionally, the reactor system may be subdivided into multiple zones, i.e., a multizone reactor system. Zones may be defined by physical separation, such as by walls or barriers that define separate areas, or by differences
25 in the reaction conditions, such as, for example, pressure, temperature, composition or concentration of the catalyst, reactants, or other reaction components such as inert materials, pH modifiers, etc. For example, the reactor system may comprise a single reactor comprising a single zone, a single reactor comprising multiple zones, multiple reactors comprising a single zone in each reactor, multiple reactors where one or more reactors has a
30 single zone and one or more reactors that comprise multiple zones, or multiple reactors each comprising multiple zones. By definition, a reactor system comprising multiple reactors would be considered a multizone reactor system. An example of a multizone reactor may be a continuous tubular reactor comprising multiple zones, including one or more mixing zones, a cooling zone, and one or more catalyst zones where the reaction takes place.

Another example of a multizone single reactor may be a stirred bed reactor comprising internal walls containing the catalyst that defines a catalyst zone through which liquid reactants are circulated, and a feed/removal zone outside of the catalyst zone where the reactants enter the reactor and products exit the reactor. When referring to the average concentration or any ratio of the reactor system, the average concentration or ratio is calculated based on what enters the reactor system and what exits the reactor system.

The reactor system may comprise a reactor configured as a fluidized bed reactor, a fixed bed reactor, a trickle bed reactor, a packed bubble column reactor, or a stirred bed reactor. Preferably, the reactor system comprises a packed bubble column reactor.

The catalyst may be present in the form of a slurry or a fixed bed depending on the reactor in which the catalyst is present. For example, a slurry catalyst can be used in a stirred bed reactor or a fluidized bed reactor, whereas a fixed bed catalyst can be used in a fixed bed reactor, trickle bed reactor, or a packed bubble column reactor. Preferably, the catalyst is in the form of a fixed bed reactor.

The size of the catalyst can be selected based on the type of reactor. For example, a slurry catalyst may have an average particle diameter less than 200 μm , such as, for example, from 10 μm to 200 μm . A fixed bed catalyst may have an average particle diameter 200 μm or greater, such as, for example, from 200 μm to 30 mm. Preferably, the average diameter of the catalyst particle is at least 60 μm , preferably at least 100 μm , preferably at least 200 μm , preferably at least 300 μm , preferably at least 400 μm , preferably at least 500 μm , preferably at least 600 μm , preferably at least 700 μm , preferably at least 800 μm ; preferably no more than 30 mm, preferably no more than 20 mm, preferably no more than 10 mm, preferably no more than 5 mm, preferably no more than 4 mm, preferably no more than 3 mm.

The noble metal-containing catalyst comprises particles of a noble metal. Preferably, the noble metal comprises palladium or gold, and more preferably the noble metal comprises gold.

The particles of a noble metal preferably have an average diameter of less than 15 nm, preferably less than 12 nm, more preferably less than 10 nm, and even more preferably less than 8 nm. The standard deviation of the average diameter of the noble metal particles is +/- 5 nm, preferably +/- 2.5 nm, and more preferably +/- 2 nm. As used herein, the standard deviation is calculated by the following equation:

$$\text{standard deviation} = \sqrt{\frac{\sum_{i=1}^n |x - \bar{x}|}{n}}$$

where x is the size of each particle, $\bar{x}$ is the mean of the n number of particles, and n is at least 500.

Preferably, the noble metal-containing catalyst further comprises titanium-containing particles.

5 The titanium-containing particles may comprise elemental titanium or a titanium oxide, TiO_x . Preferably, the titanium-containing particles comprise a titanium oxide.

 The titanium-containing particles preferably have an average diameter of less than 5 times the average diameter of the noble metal-containing particles, more preferably an average diameter of less than 4 times the average diameter of the noble metal-containing
10 particles, even more preferably an average particle diameter of less than 3 times the average diameter of the noble metal-containing particles, still more preferably an average particle diameter of less than 2 times the average diameter of the noble metal-containing particles, and yet more preferably an average particle diameter of less than 1.5 times the average diameter of the noble metal-containing particles.

15 The amount by weight of the noble metal-containing particles with respect to the amount of the titanium-containing particles may range from 1:1 to 1:20. Preferably, the weight ratio of noble metal-containing particles to titanium-containing particles ranges from 1:2 to 1:15, more preferably from 1:3 to 1:10, even more preferably from 1:4 to 1:9, and still more preferably from 1:5 to 1:8.

20 Preferably, the noble metal particles are evenly distributed among the titanium-containing particles. As used herein, the term “evenly distributed” means the noble metal particles are randomly dispersed among the titanium-containing particles with substantially no agglomeration of the noble metal particles. Preferably, at least 80% of the total number of the noble metal particles are present in a particle having an average diameter less than 15
25 nm. More preferably, at least 90% of the total number of the noble metal particles are present in a particle having an average diameter less than 15 nm. Even more preferably, at least 95% of the total number of noble metal particles are present in a particle having an average diameter less than 15 nm..

 The noble metal particles in the catalyst may be disposed on a surface of a support
30 material. Preferably, the support material is a particle of an oxide material; preferably γ -, δ -, or θ -alumina, silica, magnesia, titania, zirconia, hafnia, vanadia, niobium oxide, tantalum oxide, ceria, yttria, lanthanum oxide or a combination thereof. Preferably, in portions of the catalyst comprising the noble metal, the support has a surface area greater than $10 \text{ m}^2/\text{g}$, preferably greater than $30 \text{ m}^2/\text{g}$, preferably greater than $50 \text{ m}^2/\text{g}$, preferably greater than 100

m²/g, preferably greater than 120 m²/g. In portions of the catalyst which comprise little or no noble metal, the support may have a surface area less than 50 m²/g, preferably less than 20 m²/g. The average diameter of the support and the average diameter of the final catalyst particle are not significantly different.

5 Preferably, the aspect ratio of the catalyst particle is no more than 10:1, preferably no more than 5:1, preferably no more than 3:1, preferably no more than 2:1, preferably no more than 1.5:1, preferably no more than 1.1:1. Preferred shapes for the catalyst particle include spheres, cylinders, rectangular solids, rings, multi-lobed shapes (e.g., cloverleaf cross section), shapes having multiple holes and “wagon wheels;” preferably spheres.

10 Irregular shapes may also be used.

The noble metal particles can be dispersed throughout the catalyst or have varying concentration densities, such as, for example, a gradient concentration or layered structure. Preferably, at least 90 wt% of the noble metal(s) is in the outer 70% of catalyst volume (i.e., the volume of an average catalyst particle), preferably the outer 60% of catalyst volume, 15 preferably the outer 50%, preferably the outer 40%, preferably the outer 35%, preferably in the outer 30%, preferably in the outer 25%. Preferably, the outer volume of any particle shape is calculated for a volume having a constant distance from its inner surface to its outer surface (the surface of the particle), measured along a line perpendicular to the outer surface. For example, for a spherical particle the outer x% of volume is a spherical shell 20 whose outer surface is the surface of the particle and whose volume is x% of the volume of the entire sphere. Preferably, at least 95 wt% of the noble metal is in the outer volume of the catalyst, preferably at least 97 wt%, preferably at least 99 wt%. Preferably, at least 90 wt% (preferably at least 95 wt%, preferably at least 97 wt%, preferably at least 99 wt%) of the noble metal(s) is within a distance from the surface that is no more than 30% of the 25 catalyst diameter, preferably no more than 25%, preferably no more than 20%, preferably no more than 15%, preferably no more than 10%, preferably no more than 8%. Distance from the surface is measured along a line which is perpendicular to the surface.

Preferably, the catalyst comprises gold particles and titanium-containing particles on a support material comprising silica. Preferably, the gold particles and titanium-containing 30 particles form an eggshell structure on the support particles. The eggshell layer may have a thickness of 500 microns or less, preferably 250 microns or less, and more preferably 100 microns or less.

Preferably, at least 0.1% by weight of the total weight of the gold particles are exposed on a surface of the catalyst, where the surface includes both the outer surface and

pores of the catalyst. As used herein, the term “exposed” means that at least a portion of the gold particle is not covered by another gold particle or titanium-containing particle, i.e., the reactants can directly contact the gold particle. The gold particles may therefore be disposed within a pore of the support material and still be exposed by virtue of the reactant being able to directly contact the gold particle within the pore. More preferably, at least 0.25% by weight of the total weight of the gold particles are exposed on the surface of the catalyst, even more preferably, at least 0.5% by weight of the total weight of the gold particles are exposed on the surface of the catalyst, and still more preferably, at least 1% by weight of the total weight of the gold particles are exposed on the surface of the catalyst.

The catalyst is preferably produced by precipitating the noble metal from an aqueous solution of metal salts in the presence of the support. Suitable noble metal salts may include, but are not limited to, tetrachloroauric acid, sodium aurothiosulfate, sodium aurothiomalate, gold hydroxide, palladium nitrate, palladium chloride and palladium acetate. In one preferred embodiment, the catalyst is produced by an incipient wetness technique in which an aqueous solution of a suitable noble metal precursor salt is added to a porous inorganic oxide such that the pores are filled with the solution and the water is then removed by drying. The resulting material is then converted into a finished catalyst by calcination, reduction, or other treatments known to those skilled in the art to decompose the noble metal salts into metals or metal oxides. Preferably, a C₂-C₁₈ thiol comprising at least one hydroxyl or carboxylic acid substituent is present in the solution. Preferably, the C₂-C₁₈ thiol comprising at least one hydroxyl or carboxylic acid substituent has from 2 to 12 carbon atoms, preferably 2 to 8, preferably 3 to 6. Preferably, the thiol compound comprises no more than 4 total hydroxyl and carboxylic acid groups, preferably no more than 3, preferably no more than 2. Preferably, the thiol compound has no more than 2 thiol groups, preferably no more than one. If the thiol compound comprises carboxylic acid substituents, they may be present in the acid form, conjugate base form or a mixture thereof. The thiol component also may be present either in its thiol (acid) form or its conjugate base (thiolate) form. Especially preferred thiol compounds include thiomalic acid, 3-mercaptopropionic acid, thioglycolic acid, 2-mercaptoethanol and 1-thioglycerol, including their conjugate bases.

In one embodiment of the invention, the catalyst is produced by deposition precipitation in which a porous inorganic oxide is immersed in an aqueous solution containing a suitable noble metal precursor salt and that salt is then made to interact with the surface of the inorganic oxide by adjusting the pH of the solution. The resulting treated

solid is then recovered (e.g. by filtration) and then converted into a finished catalyst by calcination, reduction, or other pre-treatments known to those skilled in the art to decompose the noble metal salts into metals or metal oxides.

5 The catalyst bed may further comprise inert or acidic materials. Preferred inert or acidic materials include, e.g., alumina, clay, glass, silica carbide and quartz. Preferably, the inert or acidic materials located before and/or after the catalyst bed have an average diameter equal to or greater than that of the catalyst, preferably 1 to 30 mm; preferably at least 2 mm; preferably no greater than 30 mm, preferably no greater than 10 mm, preferably no greater than 7 mm.

10 This invention is useful in a process for producing an methacrylic acid which comprises reacting methacrolein and water in the presence of an oxygen-containing gas in an oxidative reactor system containing a catalyst bed.

The catalyst bed, which may comprise a slurry bed or fixed bed, comprises the catalyst particles. The oxidative reactor system further comprises a liquid phase comprising
15 methacrolein, water and methacrylic acid and a gaseous phase comprising oxygen. The liquid phase may further comprise byproducts, e.g., methacrolein dimethyl acetal (MDA).

Preferably, an average concentration of methacrolein in the oxidative reactor system is less than 40 wt% based on the total weight of water and methacrolein. Preferably, the oxidative reactor system has an average ratio of water to methacrolein less than 40:1 based
20 on an average amount of water and methacrolein entering and exiting the system.

Preferably, oxygen concentration in a gas stream exiting the oxidative reactor system is at least 1 mol%, more preferably at least 2 mol%, even more preferably at least 2.5 mol%, still more preferably at least 3 mol%, yet more preferably at least 3.5 mol%, even yet more preferably at least 4 mol %, and most preferably at least 4.5 mol%, based on the
25 total volume of the gas stream exiting the oxidative reactor system. Preferably, the oxygen concentration in a gas stream exiting the oxidative reactor system is no more than 7.5 mol%, preferably no more than 7.25 mol%, preferably no more than 7 mol%, based on the total amount of the gas stream exiting the oxidative reactor system.

Preferably, the liquid phase in the oxidative reactor system is at a temperature from
30 40 to 120 °C; preferably at least 50 °C, and preferably at least 55 °C. The temperature of the liquid phase in the oxidative reactor system is preferably no more than 110 °C, and preferably no more than 100 °C. When the oxidative reactor system comprises more than one reactor and/or more than one zone, the temperature in each reactor and/or zone may be

the same or different. For example, a reaction mixture exiting a reactor or zone may be cooled prior to entering the next reactor or zone.

Preferably, the catalyst bed in the oxidative reactor system is at a pressure from 1 to 150 bar (100 to 15000 kPa). The pressure in the catalyst bed of the oxidative reactor system may be at least 10 bar, preferably at least 20 bar, preferably at least 30 bar, preferably at least 40 bar, or preferably at least 60 bar. For example, the pressure in the catalyst bed of the oxidative reactor system may be at least 100 bar. When the oxidative reactor system comprises more than one reactor and/or zone, the pressure in each reactor and/or zone may be the same or different.

The heterogeneous noble metal-containing catalyst in the oxidative reactor system may be present in an amount ranging from 0.02 kg to 2 kg of catalyst for every gram-mole of methacrylic acid exiting the reactor system over the course of 1 hour. Preferably, the heterogeneous noble metal-containing catalyst in the oxidative reactor system is present in an amount of at least 0.02 kg to 0.5 kg of catalyst, for every gram-mole of methacrylic acid exiting the reactor system over the course of 1 hour. Preferably, the heterogeneous noble metal-containing catalyst in the oxidative reactor system is present in an amount of less than 0.4 kg of catalyst, more preferably less than 0.3 kg of catalyst, still more preferably less than 0.25 kg of catalyst, and even more preferably less than 0.2 kg of catalyst for every gram-mole of methacrylic acid exiting the reactor system over the course of 1 hour.

The amount of methacrylic acid exiting the reactor is dependent on the conversion of methacrolein in the oxidative reactor system. For example, at 50% conversion of methacrolein entering the oxidative reactor system, 2 moles of methacrolein would be required for every mole of methacrylic acid produced. In this example, the heterogeneous noble metal-containing catalyst in the oxidative reactor system may be present in an amount ranging from 0.01 to 1 kg of catalyst for every gram-mole of methacrolein entering the reactor system over the course of 1 hour. At 25% conversion of methacrolein entering the oxidative reactor system, 4 moles of methacrolein would be required for every mole of methacrylic acid produced, and the heterogeneous noble metal-containing catalyst in the oxidative reactor system may be present in an amount ranging from 0.005 to 0.5 kg of catalyst for every gram-mole of methacrolein entering the reactor system over the course of 1 hour. At 75% conversion of methacrolein entering the oxidative reactor system, 1.33 moles of methacrolein would be required for every mole of methacrylic acid produced, and the heterogeneous noble metal-containing catalyst in the oxidative reactor system may be present in an amount ranging from 0.015 to 1.5 kg of catalyst for every gram-mole of

methacrolein entering the reactor system over the course of 1 hour. Disregarding any external recycle streams, the oxidative reactor system preferably exhibits at least 25% conversion of methacrolein to methacrylic acid, more preferably at least 35% conversion, and even more preferably at least 40% conversion of methacrolein to methacrylic acid in the oxidative reactor system. Addition of an external recycle stream that recycles unreacted methacrolein to the oxidative reactor system can also be used to improve the overall conversion efficiency of the process.

When the noble metal-containing catalyst comprises gold, the gold may be present in an amount ranging from 0.0001 kg to 0.1 kg for every gram-mole of methacrylic acid exiting the reactor system over the course of 1 hour. Preferably, the gold is present in an amount of at least 0.0001 kg to 0.005 kg for every gram-mole of methacrylic acid exiting the reactor system over the course of 1 hour. Preferably, the gold is present in an amount less than 0.004 kg for every gram-mole of methacrylic acid exiting the reactor system over the course of 1 hour.

In terms of the amount of heterogeneous noble metal-containing catalyst in the oxidative reactor system with respect to the amount of methacrolein entering the reactor system, at 50% conversion of methacrolein entering the oxidative reactor system, the gold in the heterogeneous noble metal-containing catalyst in the oxidative reactor system may be present in an amount ranging from 0.00005 to 0.05 kg of gold for every gram-mole of methacrolein entering the reactor system over the course of 1 hour. At 25% conversion of methacrolein entering the oxidative reactor system, the gold in the heterogeneous noble metal-containing catalyst in the oxidative reactor system may be present in an amount ranging from 0.000025 to 0.025 kg of catalyst for every gram-mole of methacrolein entering the reactor system over the course of 1 hour. At 75% conversion of methacrolein entering the oxidative reactor system, the gold in the heterogeneous noble metal-containing catalyst in the oxidative reactor system may be present in an amount ranging from 0.000075 to .075 kg of catalyst for every gram-mole of methacrolein entering the reactor system over the course of 1 hour.

The pH in the catalyst bed may range from 2 to 10. Some catalysts may be deactivated in acidic conditions. Therefore, when the catalyst is not acid resistant, the pH in the catalyst bed is from 4 to 10; preferably at least 5, preferably at least 5.5; preferably no greater than 9, preferably no greater than 8, preferably no greater than 7.5.

To increase the pH in the reactor system, a base material may be added. The base material may comprise an Arrhenius base (i.e., a compound that dissociates in water to form

- hydroxide ions), a Lewis base (i.e., a compound capable of donating a pair of electrons), or a Bronsted-Lowry base (i.e., a compound capable of accepting a proton). Examples of Arrhenius bases include, but are not limited to, hydroxides of alkali and alkali earth metals. Examples of Lewis bases include, but are not limited to, amines, sulfates, and phosphines.
- 5 Examples of Bronsted-Lowry bases include, but are not limited to, halides, nitrates, nitrites, chlorites, chlorates, etc. Ammonia can be either a Lewis base or a Bronsted-Lowry base.

- The present inventors have discovered that high localized concentrations of base materials in the reactor system can cause the formation of unwanted Michael adduct as byproducts. Therefore, to aid in the minimization of the formation of Michael adducts, the
- 10 base material is preferably mixed with at least one other material prior to entering the reactor system. Preferably, the base material is introduced at a position external to the reactor system and mixed with one or more reactants or diluents to form a base-containing stream. For example, the base material may be mixed with water or a non-reactive solvent, i.e., a solvent that does not negatively impact the formation of methacrylic acid in the
- 15 reactor system. The position external to the reactor system may be a mixing vessel. Alternatively, the position external to the reactor may be a line through which components travel to the reactor system, such as a feed line or a recycle line, in which sufficient mixing occurs, such as by turbulent flow, baffles, jet mixer, or other mixing method.

- Preferably, the amount of the base material in the base-containing stream is 50 wt%
- 20 or less based on the total weight of the base-containing stream, preferably 25 wt% or less, preferably 20 wt% or less, preferably 15 wt% or less, preferably 10 wt% or less, preferably 5 wt% or less, or preferably 1 wt% or less. The base material is preferably diluted by a factor of less than 1:2, such as, less than 1:3, less than 1:4, less than 1:5, less than 1:10, less than 1:20, or less than 1:100, relative to the total weight of the base-containing stream prior
- 25 to entering the reactor system.

- Preferably, the base-containing stream is sufficiently mixed to avoid localized spikes in the concentration of the base material within the base-containing stream before it is added to the reactor system. For example, it is preferred that the base-containing stream reach at least 95% degree of homogeneity, i.e., variations in the concentration of the base material
- 30 deviate within +/- 5% of the average concentration of base material for the base-containing stream prior to entering the reactor system. Preferably, the base-containing stream reaches 95% degree of homogeneity within 4 minutes of introduction of the base material, more preferably within 2 minutes, and even more preferably within 1 minute of introduction of the base material.

For a mixing vessel, the time required for an additive to reach a 95% degree of homogeneity is defined at Θ_{95} , which can be calculated by the method disclosed by Grenville and Nienow, The Handbook of Industrial Mixing, Pages 507-509, which gives the following expression for a stirred tank in turbulent flow:

$$\theta_{95} = 5.20 \frac{T^{1.5} H^{0.5}}{D^2 N_p^{1/3} N}$$

where T is the tank diameter, H is the liquid height, D is the impeller diameter, N_p is the characteristic power number of the impeller(s), and N is the impeller speed. Similar expressions exist for static mixers, jet mixed vessels, etc.

Preferably, no base material is added to the reactor system, either internal or external to the reactor system. Preferably, when no base material is added to the reactor system, the noble metal-containing catalyst comprises an acid-resistant catalyst such as a catalyst comprised of gold and titanium-containing particles. Operating the oxidative reactor system in the absence of a base material may provide several advantages. One advantage is the increased selectivity and space time yield (STY) due to lower production of Michael adducts. Another advantage is the reduction in cost due to the reduced cost to treat aqueous waste. Aqueous waste exiting an oxidative process in which a base material was used can produce large quantities of inorganic salts, which can be difficult or impossible to treat with biological water treatment processes. This in turn, may require the use of other waste treatment process, such as incineration.

The methacrolein used in the oxidative reaction is preferably produced by either an aldol condensation or Mannich condensation. Preferably, the methacrolein is formed by the Mannich condensation of propionaldehyde and formaldehyde in the presence of a suitable catalyst. The molar ratio of propionaldehyde to formaldehyde may range from 1:20 to 20:1, preferably from 1:1.5 to 1.5:1, more preferably 1:1.25 to 1.25:1, and even more preferably from 1:1.1 to 1.1:1.

Examples of catalysts that may be used in a Mannich condensation process include, for example, amine-acid catalysts. Acids of the amine-acid catalysts may include, but are not limited to, inorganic acids (e.g., sulfuric acid and phosphoric acid) and organic mono-, di-, or polycarboxylic acids (e.g., aliphatic C_1 - C_{10} monocarboxylic acids, C_2 - C_{10} dicarboxylic acids, C_2 - C_{10} polycarboxylic acids). Amines of the amine-acid catalysts may include, but are not limited to compounds of formula NHR^1R^2 , where R^1 and R^2 are each independently C_1 - C_{10} alkyl, which are optionally substituted with an ether, hydroxyl, secondary amino or tertiary amino group, or R^1 and R^2 , together with the adjacent nitrogen,

may form a C₅-C₇ heterocyclic ring, optionally containing a further nitrogen atom and/or an oxygen atom, and which are optionally substituted by a C₁-C₄ alkyl or C₁-C₄ hydroxyalkyl.

The Mannich condensation reaction is preferably carried out in the liquid phase by reacting propionaldehyde, formaldehyde, and methanol in the presence of an amine-acid catalyst in a reactor at a temperature of at least 20 °C and at a pressure greater than 1 bar. The temperature of the reactor may range from 20 °C to 220 °C, preferably from 80 °C to 220 °C, and more preferably from 120 °C to 220 °C. The pressure of the reactor may range from greater than 1 bar to 120 bar.

Inhibitors can be added to the reactor to prevent the formation of unwanted products. For example, 4-hydroxy-2,2,6,6-tetramethylpiperidin-1-oxyl (4-hydroxy-TEMPO) can be added to the reactor.

The propionaldehyde used to prepare the methacrolein can be prepared by the hydroformylation of ethylene. The hydroformylation process is known in the art, and is disclosed, for example, in U.S. Patent No. 4,427,486, U.S. Patent No. 5,087,763, U.S. Patent No. 4,716,250, U.S. Patent No. 4,731,486, and U.S. Patent No. 5,288,916. The hydroformylation of ethylene to propionaldehyde comprises contacting ethylene with carbon monoxide and hydrogen in the presence of a hydroformylation catalyst. Examples of hydroformylation catalysts include, for example, metal-organophosphorus ligand complexes, such as organophosphines, organophosphites, and organophosphoramidites. The ratio of carbon monoxide to hydrogen may range from 1:10 to 100:1, preferably from 1:10 to 10:1. The hydroformylation process may be conducted at a temperature ranging from -25 °C to 200 °C, preferably from 50 °C to 120 °C.

Ethylene used to prepare propionaldehyde may be prepared from the dehydration of ethanol. For example, ethylene can be prepared by the acid-catalyzed dehydration of ethanol. Ethanol dehydration is known in the art and is disclosed, for example, in U.S. Patent No. 9,249,066. Preferably, ethanol is sourced from renewable resources, such as plant materials or biomass, as opposed to ethanol prepared from petroleum based sources. Using bio-resourced ethanol alone in the process for producing methacrylic acid can result in up to 50% of the carbon atoms of the methacrylic acid (i.e., 2 of the 4 carbon atoms in the methacrylic acid) coming from renewable resources.

To further increase the renewable carbon content in the methacrylic acid, additional starting materials can also be prepared from renewable resources. For example, formaldehyde can be prepared from syngas, where the syngas can be prepared from

biomass. Carbon monoxide, which can also be used in the preparation of propionaldehyde, can also be prepared from renewable resources, as disclosed by Li et al., ACS Nano, 2020, 14, 4, 4905-4915. Using these additional bio-resourced can further increase the amount of renewable carbon.

5 Alternatively, starting materials to produce the methacrylic acid can be prepared from recycled materials. For example, recycled carbon dioxide can be used to produce methanol, and the methanol can be used to produce formaldehyde.

 Preferably, at least 50% of the carbon atoms in the methacrylic acid are derived from renewable or recycled content, more preferably at least 75%, and even more preferably
10 100%.

CLAIMS:

1. A process for producing methacrylic acid comprising:
 - a) producing methacrolein from propionaldehyde and formaldehyde;
 - b) producing methacrylic acid in an oxidative reaction from the methacrolein produced in step a) and water;wherein:
 - step b) is performed at a pressure above 1 bar;
 - step b) is performed in a reactor system in a liquid phase reaction in the presence of a heterogeneous noble metal-containing catalyst, wherein the reactor system comprises an oxygen-containing gas;
 - an average concentration of methacrolein in step b) is less than 40 wt% based on the total weight of water and methacrolein; and
 - the reactor system of step b) has an average ratio of water to methacrolein less than 40:1 based on an average amount of water and methacrolein entering and exiting the system.
2. The process of any one of the preceding claims, wherein oxygen in a gas phase exiting the reactor system of step b) is present in an amount ranging from 1 mol% and 7.5 mol% oxygen based on the total amount of the gas phase.
3. The process of claim 2, wherein oxygen in the gas phase exiting the reactor system of step b) is present in an amount ranging from 2 mol% and 7.25 mol% based on the total amount of the gas phase.
4. The process of claim 3, wherein oxygen in the gas phase exiting the reactor system of step c) is present in an amount ranging from 4 mol% and less than 7 mol% based on the total amount of the gas phase.
5. The process of any one of the preceding claims, wherein the heterogeneous noble metal-containing catalyst is in the form of a slurry or fixed bed.
6. The process of any one of the preceding claims, wherein the heterogeneous noble metal-containing catalyst comprises gold.
7. The process of any one of the preceding claims, wherein the heterogeneous noble metal-containing catalyst is present in an amount ranging from 0.02 kg to 2 kg catalyst for every gram-mole of methacrylic acid exiting the reactor system over the course of 1 hour.
8. The process of any one of the preceding claims, wherein the heterogeneous noble metal-containing catalyst is present in an amount ranging from 0.0001 kg to 0.1 kg gold for every gram-mole of methacrylic acid exiting the reactor system over the course of 1 hour.

9. The process of any one of the preceding claims, the reactor of step b) comprises a multizone reactor.
10. The process of any one of the preceding claims, wherein the reactor system of step b) comprises a single reactor.
11. The process of any one of claims 1 to 9, wherein the reactor system of step b) comprises more than one reactor.
12. The process of any one of the preceding claims, further comprising producing the propionaldehyde from ethylene.