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DESCRIPTION

METHOD AND SYSTEM FOR PRODUCING HIGH-PURITY HYDROGEN CHLORIDE

5 Technical Field

The present invention relates, in general, to a method and system for producing high-purity hydrogen chloride, and more particularly, to a method and system for producing high-purity hydrogen chloride, in which a high-purity hydrogen
10 chloride having a purity of 3 N (99.9%) to 6 N (99.9999%) can be produced with low energy using a simpler process by reacting purified hydrogen with purified chloride at a high temperature of about 1,200~1,400 °C to synthesize hydrogen chloride, converting the hydrogen chloride to a liquid state
15 and purifying the liquid-state hydrogen chloride.

Background Art

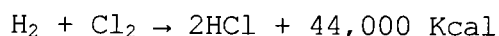
Anhydrous hydrogen chloride (HCl), also known as anhydrous hydrochloric acid, is a compound, which has a molecular weight
20 of 36.47, is present in a gaseous state at room temperature and atmospheric pressure and is liquefied at atmospheric pressure and -85 °C. Hydrogen chloride is used in the production of various chemicals, including medical drugs and dye intermediates, and particularly, high-purity hydrogen chloride
25 is advantageously used in semiconductor manufacturing

processes.

As used herein, the expression "hydrogen chloride" refers to a gaseous or liquid anhydrous hydrochloric acid, and the expression "hydrochloric acid" refers to a 35-37 wt% aqueous solution of hydrogen chloride. In addition, unless otherwise specified herein, the expression "high-purity hydrogen chloride" refers to a hydrogen chloride having a purity of 3 N (99.9%) or higher, preferably 3 N (99.9%) to 6 N (99.9999%). As used herein, the terms "crude hydrogen" and "crude chlorine" refer to unpurified hydrogen (H₂) and unpurified chlorine (Cl₂), respectively, and the terms "hydrogen" and "chlorine" refer to either purified hydrogen and chlorine, or hydrogen and chlorine elements in mixtures.

The synthesis of hydrogen chloride is generally performed by allowing crude chlorine (Cl₂) and crude hydrogen (H₂), produced by the electrolysis of brine, to react with each other at a high temperature of 1,200~1,300 °C.

[Reaction Equation 1]



When HCl gas obtained according to reaction equation 1 is cooled and absorbed into water, a 35-37 wt% aqueous solution of hydrochloric acid is produced. Conventionally, the production of anhydrous hydrochloric acid is performed by a wet process using hydrochloric acid. Specifically, liquid hydrogen chloride is produced by heating a 35-37 wt% aqueous solution of

hydrochloric acid in an evaporator to generate hydrogen chloride gas and dehydrating, drying, purifying and cooling the hydrogen chloride gas, followed by compression and cooling. This conventional production method has shortcomings in that a large amount of equipment maintenance cost is required because hydrochloric acid is treated at high temperature, and a large amount of energy cost is required because of the use of a large amount of steam.

If HCl gas produced according to reaction equation 1 can be compressed and cooled directly after the production thereof, anhydrous hydrogen chloride can be produced in a simple and easy manner. However, crude hydrogen (H_2) produced by the electrolysis of brine usually contains a large amount of water, and crude chlorine (Cl_2) produced in a general electrolytic cell contains oxygen (O_2), nitrogen (N_2), carbon dioxide (CO_2), water (H_2O) and metal components, and thus has a purity of about 99.8%. Among these impurities, water and oxygen interfere with the processes of compressing and liquefying hydrogen chloride. Specifically, water and oxygen which is converted to water during the synthesis of hydrogen chloride make it difficult to operate equipment such as a compressor. Thus, when water and oxygen are removed from the raw materials, a compressor for compressing hydrogen chloride can be used without difficulty, making it possible to produce a hydrogen chloride having a purity of 3 N or lower. However, in order to produce high-

purity (99.999% or higher) hydrogen chloride which is used in semiconductor manufacturing processes and the like, not only water and oxygen, but also other impurities, need to be removed. Particularly, carbon dioxide gas, once mixed with hydrogen chloride gas, is almost impossible to separate from the hydrogen chloride gas. For this reason, the production of hydrogen chloride is based on the wet process which is disadvantageous in terms of productivity and cost.

10 **Disclosure**

Technical Problem

Accordingly, the present invention has been made keeping in mind the above-described problems occurring in the prior art, and an object of the present invention is to provide a method and system of producing high-purity hydrogen chloride by a dry process in a more economical and simpler manner, which can substitute for the conventional wet process of producing high-purity hydrogen chloride using hydrochloric acid as a starting material.

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Technical Solution

In order to accomplish the above objects, the present invention provides a method for producing high-purity hydrogen chloride, comprising the steps of: purifying each of crude hydrogen and crude chlorine as raw materials to a purity of

99.999% or higher; reacting an excessive molar amount of the purified hydrogen with the purified chlorine at a temperature ranging from 1,200 °C to 1,400 °C to synthesize hydrogen chloride; converting the hydrogen chloride to a liquid state by
5 compression; and purifying the hydrogen chloride and separating unreacted hydrogen by fractional distillation.

In the inventive method for producing high-purity hydrogen chloride, purifying the crude hydrogen may be performed by removing water and oxygen from the crude hydrogen, produced by
10 electrolysis of brine, using a catalyst and an adsorbent to remove water and oxygen, and purifying the crude chloride may be performed by subjecting the crude chlorine gas to a first adsorption process to remove water, subjecting the crude chlorine to a first low-temperature distillation process to
15 remove metal components, and then subjecting the crude chlorine to a second low-temperature distillation process to remove gas components.

In the method of the present invention, the purified hydrogen is preferably used in an amount larger than the
20 purified chlorine by 10-20 mole%.

The present invention also provides a system for producing high-purity hydrogen chloride, comprising: hydrogen and chlorine supply pipes for supplying hydrogen and chlorine purified to a purity of 99.999% or higher, respectively; a
25 reactor in which hydrogen and chlorine, supplied through the

hydrogen and chlorine supply pipes, are reacted with each other to synthesize hydrogen chloride; a compressor for liquefying the hydrogen chloride by compression; and a distillation column for purifying the liquefied hydrogen chloride and separating
5 and removing unreacted hydrogen by fractional distillation.

In the inventive system for producing high-purity hydrogen chloride, a chiller is preferably provided in front or rear of the compressor.

The compressor or the distillation column preferably
10 comprises two or more stages.

Moreover, the inventive system for producing high-purity hydrogen chloride may further comprise a cooling/absorption column in which the hydrogen chloride resulting from the compressor is dissolved without purification to prepare
15 hydrochloric acid.

In addition, a chlorine purification system is provided in front of the chlorine supply pipe and may comprise: an adsorption column for removing water from the crude chlorine gas; a first low-temperature distillation column for removing
20 metal components; a cooler for cooling chlorine distilled in the first low-temperature distillation column; and a second low-temperature distillation column for removing gas components other than chlorine.

25 **Advantageous Effects**

According to the inventive method and system for producing high-purity hydrogen chloride, high-purity hydrogen chloride having a purity of 3N to 6N can be produced in a very simple and easy manner using a completely closed dry process
5 by reacting hydrogen directly with chlorine to synthesize hydrogen chloride, compressing and cooling the synthesized hydrogen chloride and removing unreacted hydrogen from the hydrogen chloride in a simple distillation column. In addition, according to the present invention, the production
10 process can be easily simplified and automated, and energy consumption can be significantly reduced.

Description of Drawings

FIG. 1 is a schematic view showing the configuration of a
15 system for producing high-purity hydrogen chloride according to one embodiment of the present invention.

FIG. 2 is a schematic view showing the configuration of a chlorine purification system for removing impurities from the raw material crude chlorine gas according to one embodiment of
20 the present invention.

Mode for Invention

The above objects, features and advantages of the present invention will be more apparent from the following embodiments
25 explained with respect to the accompanying drawings.

In embodiments of the present invention disclosed in the specification of the present invention, specific structural or functional descriptions are exemplified to merely describe the embodiments of the present invention, and the embodiments of
5 the present invention can be implemented in various forms and should not be interpreted as being limited to the embodiments described in the specification of the present invention.

The present invention can be modified variously and can have various forms, and specific embodiments will be
10 illustrated in the drawings and will be described in detail in the specification. However, the present invention is not limited to the specific embodiments and should be construed as including all the changes, equivalents and substitutions included in the spirit and scope of the present invention.

15 Terms, such as "first" and/or "second," can be used to describe various components, but the components are not limited by the terms. The terms are used only for the purpose of distinguishing a component from other components. For example, the first component can be designated as the second component
20 without departing from the scope of the present invention, and, similarly, the second component can also be designated as the first component.

When it is stated that a specific component is "connected" or "coupled" to another component, it should be understood that
25 the specific component, can be directly connected or linked,

but other components may be interposed between the specific component and the other component. In contrast, when it is stated that a specific component is "directly connected" or "directly coupled" to another component, it should be understood that no other components are interposed between the specific component and the other component. Other expressions for describing the relationship between components, that is, "between ~", and "immediately between ~", or "adjacent to ~", and "immediately adjacent to ~", should be interpreted in the same manner.

The terms used in the present specification are used only to describe specific embodiments, and are not intended to limit the present invention. Singular expressions may include the meaning of plural expressions unless otherwise clearly specified. In the present application, it should be understood that terms such as "comprises" or "has", are intended to indicate that proposed features, numbers, steps, operations, components, parts, or combinations thereof exist, and the probability of existence or addition of one or more other features, steps, operations, components, parts or combinations thereof is not excluded thereby.

Unless otherwise defined, all terms used herein, including technical or scientific terms, are not defined otherwise, have the same meaning as terms generally understood by those skilled in the art. The terms, such as those defined in generally used

dictionaries, should be interpreted as having the same meaning as the terms in the context of related arts, and are not to be interpreted to have meanings that are ideal or are excessively formal, when the terms are not explicitly defined in the present specification.

Hereinafter, preferred embodiments of the present invention will be described in detail with reference to the accompanying drawings. Like reference numbers in each of the drawings indicate like members.

The inventive method for producing high-purity hydrogen chloride comprises the steps of: purifying each of crude hydrogen and crude chlorine as raw materials to a purity of 99.999% or higher; reacting an excessive molar amount of the purified hydrogen with the purified chlorine at a temperature ranging from 1,200 °C to 1,400 °C to synthesize hydrogen chloride; converting the hydrogen chloride to a liquid state by compression; and purifying the hydrogen chloride and separating unreacted hydrogen by fractional distillation.

As described above, crude hydrogen (H_2) gas produced by the electrolysis of brine has a purity of only 95-96%, and crude (Cl_2) gas in a general electrolytic cell contains oxygen (O_2), nitrogen (N_2), carbon dioxide (CO_2), water (H_2O) and metal components, and thus has a purity of about 99.8%. In the present invention, hydrogen having a purity of 99.9999% or higher can be provided by removing water and oxygen from crude

hydrogen using a catalyst and an adsorbent, and chlorine having a purity of 99.9999% or higher can be provided by removing water and other impurities from crude chlorine using a chlorine purification system to be described later.

5 FIG. 1 is a schematic view showing the configuration of a system for producing high-purity hydrogen chloride according to one embodiment of the present invention. As shown in FIG. 1, the inventive system for producing high-purity hydrogen chloride may comprise: hydrogen and chlorine supply pipes for
10 supplying hydrogen and chlorine purified to a purity of 99.999% or higher, respectively; a reactor in which hydrogen and chlorine, supplied through the hydrogen and chlorine supply pipes, are reacted with each other to synthesize hydrogen chloride; a compressor for liquefying the hydrogen chloride by
15 compression; and a distillation column for purifying the liquefied hydrogen chloride and separating and removing unreacted hydrogen by fractional distillation.

In addition, the inventive system for producing high-purity hydrogen chloride may further comprise a chlorine
20 purification system provided in front of the chlorine supply pipe. FIG. 2 shows an embodiment of the chlorine purification system.

As shown in FIG. 2, the chlorine purification system may comprise: an adsorption column for removing water from chlorine
25 gas having a purity of 99.8%; a first low-temperature

distillation column for removing metal components from the chlorine gas; a cooler for cooling chlorine distilled in the first low-temperature distillation column; and a second low-temperature distillation column for removing gas components
5 from the chlorine. This chlorine purification system can be connected in-line with the above system for producing high-purity hydrogen chloride such that it can supply purified high-purity chlorine to the hydrogen chloride production system. Alternatively, the chlorine purification system can also be
10 present separately from the hydrogen chloride production system such that purified high-purity chlorine, purified in the chlorine purification system and stored in a tank, can be supplied to the hydrogen chloride production system.

Using the chlorine purification system, high-purity
15 chlorine having a purity of 99.9999% or higher can be obtained by passing crude chlorine gas having a purity of 99-99.9% through an adsorption column to remove water, passing the crude chlorine through a first low-temperature distillation column (temperature: -25 °C to 15 °C) to remove metal components such
20 as iron, chromium and nickel, and then passing the crude chlorine through a second low-temperature distillation column (temperature: -35 °C to 5 °C) to remove gas components such as carbon dioxide, nitrogen and oxygen.

In the inventive system for producing high-purity hydrogen
25 chloride, the flow rates of chlorine and hydrogen, which are

raw materials, are controlled by a flow control valve (FVC). For the reaction of hydrogen with chlorine, hydrogen is preferably added in an amount larger than chlorine. Theoretically, hydrogen and chlorine should be allowed to react
5 at a molar ratio of 1:1 in order to produce hydrogen chloride. However, when unreacted chlorine remains in hydrogen chloride, it will not be easy to separate from the hydrogen chloride, and the toxicity of the remaining chlorine can cause damage to the reaction system. For this reason, for the reaction of hydrogen
10 with chlorine, hydrogen is preferably added in an amount larger than chlorine by 10-20 mole%.

The reactor is preferably made of graphite which is not influenced by the raw material chlorine or hydrogen chloride at high temperature, and the compressor is preferably made of a
15 material which can resist hydrogen chloride. The compressor is preferably a reciprocating compressor comprising two or more stages. In addition, in order to increase compression efficiency, a chiller is preferably provided in front or rear of the compressor. The operating temperature of the reactor is
20 1,200~1,400 °C, and preferably $1,300 \pm 50$ °C. In order to maintain this temperature, hydrogen is heated by combustion with air, and water produced by this heating is absorbed by HCl gas produced in the initial stage of synthesis and is removed with hydrochloric acid. After the initial reaction, the
25 temperature of the reactor can be maintained by reaction heat.

After the reaction, a portion of unreacted hydrogen is suitably vented before or after passage through the chiller, thus reducing cooling efficiency, the liquefied hydrogen chloride is subjected to a purification process of removing metal components and the like by fractional distillation and a process of separating and removing unreacted hydrogen. In this way, high-purity hydrogen chloride having a purity of 6N or higher can be produced by passing the liquefied hydrogen chloride through the multi-stage distillation column and removing impurities such as hydrogen through the top of the column. The liquefied hydrogen chloride contains a very small amount of hydrogen due to partial pressure, and this hydrogen can act as an impurity in some processes. For this reason, the liquefied hydrogen chloride is preferably distilled in a distillation column at low temperature to completely remove the remaining hydrogen. In the inventive system for producing high-purity hydrogen chloride, the compressor or the distillation column preferably comprises two or more stages which provide higher efficiency. The hydrogen chloride, subjected to fractional distillation in the distillation column, is stored in a hydrogen chloride tank which stores purified liquid hydrogen chloride.

In addition, in order to increase economical efficiency, the inventive system for producing high-purity hydrogen chloride may further comprise a cooling/absorption column which

can produce a 37-38 wt% aqueous solution of hydrochloric acid having a high purity of 5N (99.999%) or higher by dissolving a portion of the synthesized gas in ultrapure water before liquefaction.

5 As described above, in the inventive method and system for producing high-purity hydrogen chloride, hydrogen chloride can be produced with a purity ranging from 3 N (99.9%) to 6 N (99.9999%) depending on the degree of purification of the raw materials and the reaction product. In addition, the
10 production process can be simplified and energy consumption can be significantly reduced, compared to the conventional wet process. Thus, according to the present invention, a large amount of high-purity hydrogen chloride can be produced in a more cost-effective manner.

15 Hereinafter, the present invention will be described in further detail with reference to examples. It is to be understood, however, that these examples are for illustrative purposes only and are not intended to limit the scope of the present invention.

20 **Example**

In the example of the present invention, a system for producing high-purity hydrogen chloride was used, which comprises: a reactor 10 for reacting purified high-purity hydrogen with purified high-purity chlorine; a compressor 20
25 for cooling and compressing the hydrogen chloride gas obtained

in the reactor; a chiller 21 for the hydrogen chloride passed through the compressor; a hydrochloric acid tank 60 for dissolving the hydrogen chloride, passed through the compressor, in deionized water, to prepare high-purity
5 hydrochloric acid, and storing the prepared hydrochloric acid; a two-stage distillation column (i.e., a first distillation column 40 and a second distillation column 50) for fractionally distilling the hydrogen chloride, liquefied in the compressor, to remove unreacted hydrogen and the like; and a hydrogen
10 chloride tank 30 for storing the hydrogen chloride purified in the distillation column. Using this hydrogen chloride production system, hydrogen chloride was produced. Specifically, hydrogen and chlorine were introduced into the reactor at flow rates of about 80 m³/hr and about 70 m³/hr,
15 respectively, such that the amount of hydrogen introduced was larger than that of chlorine by about 15 mole%. The reactor was maintained at about 1,300 °C. The temperature of the synthesized hydrogen chloride at the outlet of the compressor was about 60~165 °C, and the synthesized hydrogen chloride was
20 liquefied by cooling to about -20 °C using the chiller, and the liquefied hydrogen chloride was cooled to about -40 °C while it was passed through the distillation column.

Table 1 below shows the results of analysis of purities and impurities of crude hydrogen and crude chlorine as raw
25 materials, hydrogen and chlorine after purification, and

hydrogen chloride after purification in a compressor and a distillation column, as carried out according to the present invention. Table 2 below shows the results of analysis of purity and impurities of an aqueous hydrochloric acid solution
5 formed in a cooling/absorption column from a hydrogen chloride produced using the inventive system for producing high-purity hydrogen chloride. As can be seen in Tables 1 and 2, hydrogen chloride produced using the inventive system for producing high-purity hydrogen chloride had a purity of 5 N-6 N (99.999-
10 99.9999%).

[Table 1]

		Raw materials		After purification of raw materials		After synthesis	After purification	Storage tank
		Crude H ₂	Crude Cl ₂	H ₂	Cl ₂	HCl	HCl	HCl
Purity		95-96%	99.8%	99.999%	99.9995%	99.995%	99.999-99.9999%	99.999-99.9999%
Impurities	O ₂	≤10 ppm	≤500 ppm	≤ 2 ppm	≤0.5 ppm	≤1 ppm	≤1 ppm	≤1 ppm
	N ₂	-		≤1 ppm	≤1 ppm	≤1 ppm	≤1 ppm	≤1 ppm
	CO	-		≤1 ppm	≤0.5 ppm	≤0.5 ppm	≤0.5 ppm	≤0.5 ppm
	CO ₂	-		≤ 1 ppm	≤0.5 ppm	≤0.5 ppm	≤1 ppm	≤1 ppm
	CH ₄	-		≤1 ppm	≤0.5 ppm	≤1 ppm	-	-
	H ₂ O	≤40,000 ppm	≤5ppm	≤4 ppm	≤1 ppm	≤1 ppm	≤1 ppm	≤1 ppm

[Table 2]

	37% HCl mixing		37% HCl storage tank		
	HCl	DI water	37% HCl		
Purity	99.999-99.9999%		18 Ω.cm or more	36-38%	
Impurities	O ₂	≤1 ppm	-	NH4	≤0.5 ppm
	N ₂	≤1 ppm		SO4	≤0.5 ppm

	CO	≤0.5 ppm		PO4	≤0.05 ppm
	CO ₂	≤1 ppm		Residue on	≤3 ppm
	CH ₄	-		Total	<1,000 ppb
	H ₂ O	≤1 ppm			

The exemplary embodiment of the present invention, which is described as above and shown in the drawings, should not be interpreted as limiting the technical spirit of the present invention. The scope of the present invention is limited only by matters set forth in the claims and those skilled in the art can modify and change the technical subjects of the present invention in various forms. Therefore, as long as these improvements and changes are apparent to those skilled in the art, they are included in the protective scope of the present invention.

[Description of Reference Numerals in the Drawings]

10: reactor for HCl synthesis; 20: compressor;
21: chiller; 30: hydrogen chloride tank;
15 40: first HCl distillation column;
50: second HCl distillation column;
60: hydrochloric acid tank;
70: first low-temperature distillation column;
80: Cl₂ cooler;
20 90: second low-temperature distillation column.

Industrial Applicability

As described above, in the inventive method and system for producing high-purity hydrogen chloride, hydrogen chloride can be produced with a purity ranging from 3 N (99.9%) to 6 N (99.9999%) depending on the degree of purification of the raw materials and the reaction product. In addition, the production process can be simplified and energy consumption can be significantly reduced, compared to the conventional wet process. Thus, according to the present invention, a large amount of high-purity hydrogen chloride can be produced in a more cost-effective manner.

CLAIMS

1. A method for producing high-purity hydrogen chloride, comprising the steps of:

purifying each of crude hydrogen and crude chlorine as
5 raw materials to a purity of 99.999% or higher;

reacting an excessive molar amount of the purified hydrogen with the purified chlorine at a temperature ranging from 1,200 °C to 1,400 °C to synthesize hydrogen chloride;

converting the hydrogen chloride to a liquid state by
10 compression; and

purifying the hydrogen chloride and separating unreacted hydrogen by fractional distillation.

2. The method of claim 1, wherein purifying the crude
15 hydrogen is performed by removing water and oxygen from the crude hydrogen using a catalyst and an adsorbent, and purifying the crude chloride is performed by subjecting the crude chlorine gas to a first adsorption process to remove water, subjecting the crude chlorine to a first low-temperature
20 distillation process to remove metal components, and then subjecting the crude chlorine to a second low-temperature distillation process to remove gas components other than chlorine.

25 3. The method of claim 1, wherein the purified hydrogen

is used in an amount larger than the purified chlorine by 10-20 mole% in the step of reacting.

4. A system for producing high-purity hydrogen chloride,
5 comprising:

hydrogen and chlorine supply pipes for supplying hydrogen and chlorine purified to a purity of 99.999% or higher, respectively;

a reactor in which hydrogen and chlorine, supplied through
10 the hydrogen and chlorine supply pipes, are reacted with each other to synthesize hydrogen chloride;

a compressor for liquefying the hydrogen chloride by compression; and

a distillation column for purifying the liquefied hydrogen
15 chloride and separating and removing unreacted hydrogen by fractional distillation.

5. The system of claim 4, wherein a chiller is provided in front or rear of the compressor.

20 6. The system of claim 4, wherein the compressor or the distillation column comprises two or more stages.

7. The system of claim 4, wherein the system further comprises a cooling/absorption column in which the hydrogen
25 chloride resulting from the compressor is dissolved without

purification to prepare hydrochloric acid.

8. The system of claim 4, wherein a chlorine purification system comprises an adsorption column for removing
5 water from the crude chlorine gas; a first low-temperature distillation column for removing metal components; a cooler for cooling chlorine distilled in the first low-temperature distillation column; and a second low-temperature distillation column for removing gas components other than chlorine.

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