

GLYCOPYRROLATE SALTS

[0001] Glycopyrrolate is a quaternary ammonium cation of the muscarinic anticholinergic group. Glycopyrrolate, typically as a bromide salt, has been used in the treatment of a variety of conditions including diarrhea (U.S. Pat. Nos. 6,214,792 and 5,919,760), urinary incontinence (U.S. Pat. Nos. 6,204,285 and 6,063,808), and anxiety (U.S. Pat. No. 5,525,347). Additionally, U.S. Pat. No. 5,976,499 discloses a method for diagnosing cystic fibrosis in a patient by, in part, stimulating sweat production through the injection of a glycopyrrolate solution into a patient. Glycopyrrolate has also been used for the treatment of hyperhidrosis in US 20100276329.

[0002] Hyperhidrosis affects 8.8 million individuals in the United States alone, of whom 50.8% are estimated to have axillary hyperhidrosis and 25-34% have palmar or plantar hyperhidrosis. Hyperhidrosis is often treated with aluminum salts. Applying aluminum salts, such as aluminum chloride, causes frequent skin irritation and is of only limited effectiveness. Topically applied glycopyrrolate bromide has been shown to cause less skin irritation and have increased effectiveness over aluminum chloride.

[0003] Glycopyrrolate has well-known pharmacology (anticholinergic) and acts as a muscarinic receptor antagonist. As with other anticholinergic agents, glycopyrrolate inhibits the action of acetylcholine on structures innervated by postganglionic cholinergic nerves, such as sweat glands. Under physiologic conditions, salts of glycopyrrolate are dissociated; therefore, the pharmacological activity of glycopyrrolate is mediated by the active cation moiety, also referred to as glycopyrronium. .

[0004] Glycopyrrolate has previously been made available as a bromide salt or an acetate salt. The bromide salt of glycopyrrolate is sold as Rubinol®. The term “glycopyrrolate” as used in the label for Rubinol® refers to the bromide salt which is more formally referred to as glycopyrronium bromide.

[0005] A drawback of using bromide salts of pharmaceutical compounds is the potential for inducing bromism which can result from too high an intake of bromide. Symptoms of bromism may include neurological abnormalities such as vision impairment and upper motor neuron disorders and dermatologic conditions such as papular and macular rashes. Symptoms more often develop due to chronic use rather than acute toxicity.

SUMMARY OF THE INVENTION

[0006] In one aspect of the invention, a salt of glycopyrrolate is provided wherein the anion is selected from benzoate, edisylate, oxalate, hydrogen sulfate, and tosylate.

[0007] In a further aspect of the invention, glycopyrrolate tosylate, including polymorphs, co-crystals, hydrates and solvates thereof, is provided.

[0008] In a further aspect of the invention, solid glycopyrrolate tosylate is provided, including polymorphs, solvates, hydrates and co-crystals thereof and amorphous glycopyrrolate tosylate.

[0009] In another aspect of the invention, glycopyrrolate tosylate monohydrate is provided.

[0010] In a further aspect of the invention, crystalline glycopyrrolate tosylate, including polymorphs, co-crystals, hydrates and solvates thereof, is provided.

[0011] In a yet another aspect of the invention, crystalline glycopyrrolate tosylate monohydrate and polymorphs thereof are provided.

[0012] In another aspect of the invention, Form C glycopyrrolate tosylate is provided.

[0013] In a further aspect of the invention, dehydrated crystalline glycopyrrolate tosylate monohydrate, hereinafter referred to as dehydrated Form D, is provided.

[0014] In further aspects of the invention, processes for making Forms C and D of glycopyrrolate tosylate are provided, as are Form C and Form D glycopyrrolate tosylate made by those processes.

[0015] In an additional aspect of the invention, processes for making threo glycopyrrolate tosylate are provided.

[0016] In another aspect of the invention, glycopyrrolate tosylate is provided.

[0017] In a further aspect of the invention threo glycopyrrolate tosylate is provided.

[0018] In another aspect of the invention, methods of treating hyperhidrosis using Forms C or D of glycopyrrolate tosylate are provided.

[0019] In another aspect of the invention, amorphous glycopyrrolate tosylate is provided.

[0020] In an additional aspect of the invention, solid dispersions comprising glycopyrrolate tosylate are provided.

[0021] In a further aspect of the invention, a topical comprising glycopyrrolate tosylate is provided.

[0022] In yet another aspect of the invention, processes for preparing an absorbent pad containing an aqueous solution of glycopyrrolate tosylate absorbed onto the pad is provided.

BRIEF DESCRIPTION OF THE DRAWINGS

[0023] Figure 1 is the ORTEP drawing of Form D glycopyrrolate tosylate monohydrate.

[0024] Figure 2 is an x-ray powder diffraction pattern of Form D glycopyrrolate tosylate monohydrate.

[0025] Figure 3 is an infrared ("IR") spectrum of Form D glycopyrrolate tosylate monohydrate.

[0026] Figure 4 is the x-ray powder diffraction pattern of Form C glycopyrrolate tosylate.

[0027] Figure 5 is the IR spectrum of Form C glycopyrrolate tosylate.

[0028] Figure 6 is the indexing solution for Form C glycopyrrolate tosylate with Pawley refinement.

[0029] Figure 7 includes a DSC thermogram and TGA for Form C glycopyrrolate tosylate.

[0030] Figure 8 is the x-ray powder diffraction for dehydrated Form D glycopyrrolate tosylate.

[0031] Figure 9 is an overlay of Form D and dehydrated Form D glycopyrrolate tosylate.

[0032] Figure 10 is the indexing solution for dehydrated Form D glycopyrrolate tosylate with Pawley refinement.

[0033] Figure 11 is the indexing solution for Form D glycopyrrolate tosylate monohydrate with Pawley refinement.

[0034] Figure 12 is the x-ray powder diffraction for crystalline glycopyrrolate benzoate.

[0035] Figure 13 is the DSC thermogram for crystalline glycopyrrolate benzoate.

[0036] Figure 14 is the x-ray powder diffraction for crystalline di-glycopyrrolate edisylate.

[0037] Figure 15 is DSC thermogram for crystalline di-glycopyrrolate edisylate.

[0038] Figure 16 is the x-ray powder diffraction for crystalline glycopyrrolate oxalate.

[0039] Figure 17 is the x-ray powder diffraction for crystalline glycopyrrolate hydrogen sulfate.

[0040] Figure 18 is the DSC thermogram for crystalline glycopyrrolate hydrogen sulfate.

[0041] Figure 19 is an x-ray amorphous diffraction pattern of glycopyrrolate tosylate.

[0042] Figure 20 is a DSC/TGA overlay of an x-ray amorphous glycopyrrolate tosylate.

- [0043] Figure 21 is the x-ray powder diffraction pattern for glycopyrrolate bromide.
- [0044] Figure 22 is the x-ray powder diffraction pattern for glycopyrrolate acetate.
- [0045] Figure 23 is the modulated DSC thermogram of an x-ray amorphous glycopyrrolate tosylate.
- [0046] Figure 24 is an overlay of a portion of the infrared spectrum of a solid dispersion of HPMCAS:glycopyrrolate tosylate (1:1) and its respective components.
- [0047] Figure 25 is the modulated DSC thermogram of a solid dispersion of HPMCAS:glycopyrrolate tosylate (1:1).
- [0048] Figure 26 is the modulated DSC thermogram of a solid dispersion of sucrose:glycopyrrolate tosylate (9:1).
- [0049] Figure 27a is an overlay of a portion of the infrared spectrum of a solid dispersion of Kollicoat® IR:glycopyrrolate tosylate (1:1) and its respective components.
- [0050] Figure 27b is an overlay of a portion of the infrared spectrum of a solid dispersion of Kollicoat® IR:glycopyrrolate tosylate (1:1) and its respective components.
- [0051] Figure 28 is the modulated DSC thermogram of a solid dispersion of Kollicoat® IR:glycopyrrolate tosylate (1:1).
- [0052] Figure 29 is the modulated DSC thermogram of a solid dispersion of Kollicoat® IR:glycopyrrolate tosylate (9:1).
- [0053] Figure 30a is an overlay of a portion of the infrared spectrum of a solid dispersion of Soluplus®:glycopyrrolate tosylate (1:1) and its respective components.
- [0054] Figure 30b is an overlay of a portion of the infrared spectrum of a solid dispersion of Soluplus®:glycopyrrolate tosylate (1:1) and its respective components.
- [0055] Figure 30c is an overlay of a portion of the infrared spectrum of a solid dispersion of Soluplus®:glycopyrrolate tosylate (1:1) and its respective components.
- [0056] Figure 31 is the modulated DSC thermogram of a solid dispersion of Soluplus®:glycopyrrolate tosylate (1:1).
- [0057] Figure 32a is an overlay of a portion of the infrared spectrum of a solid dispersion of PVP K29/32:glycopyrrolate tosylate (1:1) and its respective components.
- [0058] Figure 32b is an overlay of a portion of the infrared spectrum of a solid dispersion of PVP K29/32:glycopyrrolate tosylate (1:1) and its respective components.

[0059] Figure 32c is an overlay of a portion of the infrared spectrum of a solid dispersion of PVP K29/32:glycopyrrolate tosylate (1:1) and its respective components.

[0060] Figure 32d is an overlay of a portion of the infrared spectrum of a solid dispersion of PVP K90:glycopyrrolate tosylate (1:1) and its respective components

[0061] Figure 33 is the modulated DSC thermogram of a solid dispersion of PVP K29/32:glycopyrrolate tosylate (1:1).

[0062] Figure 34 is the modulated DSC thermogram of a solid dispersion of PVP K29/32:glycopyrrolate tosylate (8:1).

[0063] Figure 35 is the modulated DSC thermogram of a solid dispersion of PVP K90:glycopyrrolate tosylate (1:1).

[0064] Figure 36a is an overlay of a portion of the infrared spectrum of a solid dispersion of Kollidon® VA 64:glycopyrrolate tosylate (1:1) and its respective components.

[0065] Figure 36b is an overlay of a portion of the infrared spectrum of a solid dispersion of Kollidon® VA 64:glycopyrrolate tosylate (1:1) and its respective components.

[0066] Figure 36c is an overlay of a portion of the infrared spectrum of a solid dispersion of Kollidon® VA 64:glycopyrrolate tosylate (1:1) and its respective components.

[0067] Figure 37 is the modulated DSC thermogram of a solid dispersion of Kollidon® VA 64:glycopyrrolate tosylate (1:1).

DETAILED DESCRIPTION OF THE INVENTION

[0068] The term "solid form" is often used to refer to a class or type of solid-state material. One kind of solid form is a "polymorph" which refers to two or more compounds having the same chemical formula but differing in solid-state structure. Salts may be polymorphic. When polymorphs are elements, they are termed allotropes. Carbon possesses the well-known allotropes of graphite, diamond, and buckminsterfullerene. Polymorphs of molecular compounds, such as active pharmaceutical ingredients ("APIs"), are often prepared and studied in order to identify compounds meeting scientific or commercial needs including, but not limited to, improved solubility, dissolution rate, hygroscopicity, and stability.

[0069] Other solid forms include solvates and hydrates of compounds including salts. A solvate is a compound wherein a solvent molecule is present in the crystal structure together with

another compound, such as an API. When the solvent is water, the solvent is termed a hydrate. Solvates and hydrates may be stoichiometric or non-stoichiometric. A monohydrate is the term used when there is one water molecule, stoichiometrically, with respect to, for example, an API, in the unit cell.

[0070] In order to identify the presence of a particular solid form, one of ordinary skill typically uses a suitable analytical technique to collect data on the form for analysis. For example, chemical identity of solid forms can often be determined with solution-state techniques such as ^{13}C -NMR or ^1H -NMR spectroscopy and such techniques may also be valuable in determining the stoichiometry and presence of “guests” such as water or solvent in a hydrate or solvate, respectively. These spectroscopic techniques may also be used to distinguish, for example, solid forms without water or solvent in the unit cell (often referred to as “anhydrates”), from hydrates or solvates.

[0071] Solution-state analytical techniques do not provide information about the solid state as a substance and thus, for example, solid-state techniques may be used to distinguish among solid forms such as anhydrates. Examples of solid-state techniques which may be used to analyze and characterize solid forms, including anhydrates and hydrates, include single crystal x-ray diffraction, x-ray powder diffraction (“XRPD”), solid-state ^{13}C -NMR, Infrared (“IR”) spectroscopy, Raman spectroscopy, and thermal techniques such as Differential Scanning Calorimetry (DSC), melting point, and hot stage microscopy.

[0072] Polymorphs are a subset of crystalline forms that share the same chemical structure but differ in how the molecules are packed in a solid. When attempting to distinguish polymorphs based on analytical data, one looks for data which characterize the form. For example, when there are two polymorphs of a compound (*e.g.*, Form I and Form II), one can use x-ray powder diffraction peaks to characterize the forms when one finds a peak in a Form I pattern at angles where no such peak is present in the Form II pattern. In such a case, that single peak for Form I distinguishes it from Form II and may further act to characterize Form I. When more forms are present, then the same analysis is also done for the other polymorphs. Thus, to characterize Form I against the other polymorphs, one would look for peaks in Form I at angles where such peaks are not present in the x-ray powder diffraction patterns of the other polymorphs. The collection of peaks, or indeed a single peak, which distinguishes Form I from the other known polymorphs is a collection of peaks which may be used to characterize Form I.

If, for example, two peaks characterize a polymorph then those two peaks can be used to identify the presence of that polymorph and hence characterize the polymorph. Those of ordinary skill in the art will recognize that there are often multiple ways, including multiple ways using the same analytical technique, to characterize polymorphic polymorphs. For example, one may find that three x-ray powder diffraction peaks characterize a polymorph. Additional peaks could also be used, but are not necessary, to characterize the polymorph up to and including an entire diffraction pattern. Although all the peaks within an entire diffractogram may be used to characterize a crystalline form, one may instead, and typically does as disclosed herein, use a subset of that data to characterize such a crystalline form depending on the circumstances.

[0073] When analyzing data to distinguish an anhydrate from a hydrate, for example, one can rely on the fact that the two solid forms have different chemical structures – one having water in the unit cell and the other not. Thus, this feature alone may be used to distinguish the forms of the compound and it may not be necessary to identify peaks in the anhydrate, for example, which are not present in the hydrate or vice versa.

[0074] X-ray powder diffraction patterns are some of the most commonly used solid-state analytical techniques used to characterize solid forms. An x-ray powder diffraction pattern is an x-y graph with $^{\circ}2\theta$ (diffraction angle) on the x-axis and intensity on the y-axis. The peaks within this plot may be used to characterize a crystalline solid form. The data is often represented by the position of the peaks on the x-axis rather than the intensity of peaks on the y-axis because peak intensity can be particularly sensitive to sample orientation (see *Pharmaceutical Analysis*, Lee & Web, pp. 255-257 (2003)). Thus, intensity is not typically used by those skilled in the art to characterize solid forms.

[0075] As with any data measurement, there is variability in x-ray powder diffraction data. In addition to the variability in peak intensity, there is also variability in the position of peaks on the x-axis. This variability can, however, typically be accounted for when reporting the positions of peaks for purposes of characterization. Such variability in the position of peaks along the x-axis derives from several sources. One comes from sample preparation. Samples of the same crystalline material, prepared under different conditions may yield slightly different diffractograms. Factors such as particle size, moisture content, solvent content, and orientation may all affect how a sample diffracts x-rays. Another source of variability comes from instrument parameters. Different x-ray instruments operate using different parameters and these

may lead to slightly different diffraction patterns from the same crystalline solid form. Likewise, different software packages process x-ray data differently and this also leads to variability. These and other sources of variability are known to those of ordinary skill in the pharmaceutical arts.

[0076] Due to such sources of variability, it is common to recite x-ray diffraction peaks using the word "about" prior to the peak value in $^{\circ}2\theta$ which presents the data to within 0.1 or 0.2 $^{\circ}2\theta$ of the stated peak value depending on the circumstances. The x-ray powder diffraction data corresponding to the solid forms of glycopyrrolate including glycopyrrolate tosylate of the disclosure were collected on instruments which were routinely calibrated and operated by skilled scientists. Accordingly, the variability associated with these data would be expected to be closer to $\pm 0.1^{\circ}2\theta$ than to $\pm 0.2^{\circ}2\theta$ and indeed likely less than 0.1 with the instruments used herein. However, to take into account that instruments used elsewhere by those of ordinary skill in the art may not be so maintained, for example, all x-ray powder diffraction peaks cited herein have been reported with a variability on the order of $\pm 0.2^{\circ}2\theta$ and are intended to be reported with such a variability whenever disclosed herein and are reported in the specification to one significant figure after the decimal even though analytical output may suggest higher precision on its face.

[0077] Single-crystal x-ray diffraction provides three-dimensional structural information about the positions of atoms and bonds in a crystal. It is not always possible or feasible, however, to obtain such a structure from a crystal, due to, for example, insufficient crystal size or difficulty in preparing crystals of sufficient quality for single-crystal x-ray diffraction.

[0078] X-ray powder diffraction data may also be used, in some circumstances, to determine the crystallographic unit cell of the crystalline structure. The method by which this is done is called "indexing." Indexing is the process of determining the size and shape of the crystallographic unit cell consistent with the peak positions in a suitable x-ray powder diffraction pattern. Indexing provides solutions for the three unit cell lengths (a, b, c), three unit cell angles (α , β , γ), and three Miller index labels (h, k, l) for each peak. The lengths are typically reported in Angstrom units and the angles in degree units. The Miller index labels are unitless integers. Successful indexing indicates that the sample is composed of one crystalline phase and is therefore not a mixture of crystalline phases.

[0079] IR spectroscopy is another technique that may be used to characterize solid forms together with or separately from x-ray powder diffraction. In an IR spectrum, absorbed light is plotted on the x-axis of a graph in the units of "wavenumber"(cm^{-1}), with intensity on the y-axis. Variation in the position of IR peaks also exists and may be due to sample conditions as well as data collection and processing. The typical variability in IR spectra reported herein is on the order of plus or minus 2.0 cm^{-1} . Thus, the use of the word "about" when referencing IR peaks is meant to include this variability and all IR peaks disclosed herein are intended to be reported with such variability.

[0080] Thermal methods are another typical technique to characterize solid forms. Different polymorphs of the same compound often melt at different temperatures. Thus, the melting point of a polymorph, as measured by methods such as capillary melting point, DSC, and hot stage microscopy, alone or in combination with techniques such as x-ray powder diffraction, IR spectroscopy, or both, may be used to characterize polymorphs or other solid forms.

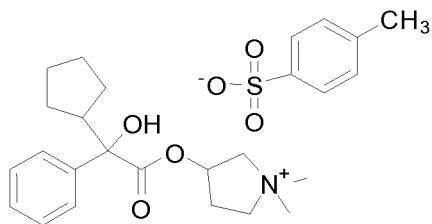
[0081] As with any analytical technique, melting point determinations are also subject to variability. Common sources of variability, in addition to instrumental variability, are due to colligative properties such as the presence of other solid forms or other impurities within a sample whose melting point is being measured.

[0082] As used herein, the term "glycopyrrolate" refers to the glycopyrronium cation of the same salt. In other words, as used herein, glycopyrrolate and glycopyrronium are used interchangeably. For example, glycopyrrolate tosylate and glycopyrronium tosylate refer to the same salt.

[0083] The present invention provides the tosylate salt of glycopyrrolate or a solvate thereof, including the solution and various solid forms thereof, the process of preparing glycopyrrolate tosylate, and the therapeutic use of glycopyrrolate tosylate.

[0084] By "glycopyrrolate tosylate," it is meant a tosylate salt of glycopyrrolate or a tosylate salt of glycopyrronium having the chemical name of 3-[(cyclopentylhydroxyphenylacetyl)oxy]-1,1-dimethyl-pyrrolidinium tosylate, also known as "3-(2-cyclopentyl-2-hydroxy-2-phenylacetox)-1,1-dimethylpyrrolidinium 4-methylbenzenesulfonate," and a structure as shown below:

[0085]



[0086] Furthermore, the term “glycopyrrolate tosylate,” as used herein, unless otherwise specified explicitly or implicitly, such as a glycopyrrolate tosylate resulting from a glycopyrrolate starting material with specific diastereomers (*e.g.*, glycopyrrolate bromide used herein which was a mixture of R,S and S,R diastereomers), includes any one of the four diastereomers listed below as well as any mixture of two, three, or four of the diastereomers: (R)-3-((S)-2-cyclopentyl-2-hydroxy-2-phenylacetoxy)-1,1-dimethylpyrrolidinium 4-methylbenzenesulfonate;

(S)-3-((R)-2-cyclopentyl-2-hydroxy-2-phenylacetoxy)-1,1-dimethylpyrrolidinium 4-methylbenzenesulfonate;

(R)-3-((R)-2-cyclopentyl-2-hydroxy-2-phenylacetoxy)-1,1-dimethylpyrrolidinium 4-methylbenzenesulfonate; and

(S)-3-((S)-2-cyclopentyl-2-hydroxy-2-phenylacetoxy)-1,1-dimethylpyrrolidinium 4-methylbenzenesulfonate.

[0087] In one embodiment, “glycopyrrolate tosylate” is (R)-3-((S)-2-cyclopentyl-2-hydroxy-2-phenylacetoxy)-1,1-dimethylpyrrolidinium 4-methylbenzenesulfonate. In another embodiment, the “glycopyrrolate tosylate” is (S)-3-((R)-2-cyclopentyl-2-hydroxy-2-phenylacetoxy)-1,1-dimethylpyrrolidinium 4-methylbenzenesulfonate. In another embodiment, the “glycopyrrolate tosylate” is (R)-3-((R)-2-cyclopentyl-2-hydroxy-2-phenylacetoxy)-1,1-dimethylpyrrolidinium 4-methylbenzenesulfonate. In another embodiment, the “glycopyrrolate tosylate” is (S)-3-((S)-2-cyclopentyl-2-hydroxy-2-phenylacetoxy)-1,1-dimethylpyrrolidinium 4-methylbenzenesulfonate. In yet another embodiment, the “glycopyrrolate tosylate” is a racemic mixture of (R)-3-((S)-2-cyclopentyl-2-hydroxy-2-phenylacetoxy)-1,1-dimethylpyrrolidinium 4-methylbenzenesulfonate and (S)-3-((R)-2-cyclopentyl-2-hydroxy-2-phenylacetoxy)-1,1-dimethylpyrrolidinium 4-methylbenzenesulfonate. In yet another embodiment, the “glycopyrrolate tosylate” is a racemic mixture of (R)-3-((R)-2-cyclopentyl-2-hydroxy-2-phenylacetoxy)-1,1-dimethylpyrrolidinium 4-methylbenzenesulfonate and (S)-3-((S)-2-cyclopentyl-2-hydroxy-2-phenylacetoxy)-1,1-di-

methylpyrrolidinium 4-methylbenzenesulfonate. The solvate, such as hydrate, of “glycopyrrolate tosylate”, can be a solvate, *e.g.*, a hydrate, of any one of the four diastereomers listed above or any mixture of two, three, or four of the diastereomers. When referencing “threo” glycopyrrolate tosylate, those of ordinary skill will recognize that it refers to a mixture of R,S and S, R diastereomers. Thus, threo glycopyrrolate tosylate refers to a racemic mixture of (R)-3-((S)-2-cyclopentyl-2-hydroxy-2-phenylacetoxy)-1,1-dimethylpyrrolidinium 4-methylbenzenesulfonate and (S)-3-((R)-2-cyclopentyl-2-hydroxy-2-phenylacetoxy)-1,1-dimethylpyrrolidinium 4-methylbenzenesulfonate.

[0088] It is to be understood that the invention further includes isotopic substitution. For example, deuterated glycopyrrolates are included within the definition of glycopyrrolate.

[0089] In one embodiment of the disclosure, a salt of glycopyrrolate is provided wherein the anion is selected from benzoate, edisylate, oxalate, hydrogen sulfate, and tosylate including hydrates and solvates thereof. In a further embodiment, a solid salt of glycopyrrolate is provided wherein the anion is selected from benzoate, edisylate, oxalate, hydrogen sulfate, and tosylate including polymorphs, hydrates, solvates, the corresponding amorphous forms of each salt, and co-crystals thereof.

[0090] In a further embodiment, a crystalline salt of glycopyrrolate benzoate is provided. An x-ray powder diffraction pattern substantially the same as the pattern of Figure 12 may be used to characterize one embodiment of crystalline glycopyrrolate benzoate. A smaller subset of the peaks may be used to characterize crystalline glycopyrrolate benzoate. For example, any one or more of the peaks, for example, at about 8.0, 11.8, 16.1, 17.8, 18.8, 20.1, or 23.8 °2 θ may be used to characterize crystalline glycopyrrolate benzoate. For example, the peaks at about 8.0 °2 θ and 16.0 °2 θ may be used to characterize glycopyrrolate benzoate. In another embodiment, a DSC endotherm at about 79 °C as shown in Figure 13 may be used to characterize crystalline glycopyrrolate benzoate. Combinations of x-ray data and DSC data may also be used to characterize glycopyrrolate benzoate. For example, one or more of the peaks at about 8.0, 11.8, 16.1, 17.8, 18.8, 20.1, or 23.8 °2 θ , such as the peaks at about 8.0 °2 θ and 18.8 °2 θ together with a DSC endotherm at about 79°C may be used to characterize glycopyrrolate benzoate.

[0091] In an additional embodiment, a crystalline salt of di-glycopyrrolate edisylate is provided. An x-ray powder diffraction pattern substantially the same as the pattern of Figure 14 may be used to characterize one embodiment of crystalline di-glycopyrrolate edisylate. A

smaller subset of the peaks may be used to characterize crystalline di-glycopyrrolate edisylate. For example, any one or more of the peaks, for example, at about 5.2, 9.2, 10.4, 11.2, 12.9, 15.3, 17.9, 18.6, 20.9, 22.3, or 23.7 °2 Θ may be used to characterize crystalline di-glycopyrrolate edisylate. For example, the peaks at about 11.2 and 17.9 °2 Θ may be used to characterize di-glycopyrrolate edisylate. In another embodiment, a DSC endotherm at about 103°C as shown in Figure 15 may be used to characterize crystalline di-glycopyrrolate edisylate. Combinations of x-ray data and DSC data may also be used to characterize di-glycopyrrolate edisylate. For example, in addition, one or more of the peaks at about 5.2, 9.2, 10.4, 11.2, 12.9, 15.3, 17.9, 18.6, 20.9, 22.3, or 23.7 °2 Θ , such as the peaks at about 11.2 and 17.9 °2 Θ together with a DSC endotherm at about 103 °C may be used to characterize di-glycopyrrolate edisylate.

[0092] In a further embodiment, a crystalline salt of glycopyrrolate oxalate is provided. An x-ray powder diffraction pattern substantially the same as the pattern of Figure 16 may be used to characterize one embodiment of crystalline glycopyrrolate oxalate. A smaller subset of the peaks may be used to characterize crystalline glycopyrrolate oxalate. For example, any one or more of the peaks, for example, at about 5.0, 8.4, 10.7, or 12.1 °2 Θ may be used to characterize crystalline glycopyrrolate oxalate. For example, the peaks at about 5.0 and 8.4 °2 Θ may be used to characterize glycopyrrolate oxalate.

[0093] In an additional embodiment, a crystalline salt of glycopyrrolate hydrogen sulfate is provided. An x-ray powder diffraction pattern substantially the same as the pattern of Figure 17 may be used to characterize one embodiment of crystalline glycopyrrolate hydrogen sulfate. A smaller subset of the peaks may be used to characterize crystalline glycopyrrolate hydrogen sulfate. For example, any one or more of the peaks, for example, at about 5.6, 13.1, 14.5, 17.2, 18.2, 19.9, 20.2, 21.4, 21.6, 22.7, or 28.9 °2 Θ may be used to characterize crystalline glycopyrrolate hydrogen sulfate. For example, the peaks at about 5.6 and 13.1 °2 Θ may be used to characterize glycopyrrolate sulfate. In another embodiment, a DSC endotherm at about 160 °C and/or a second endotherm at about 169 °C as shown in Figure 18 may be used to characterize crystalline glycopyrrolate hydrogen sulfate. Combinations of x-ray data and DSC data may also be used to characterize glycopyrrolate hydrogen sulfate. For example, in addition, one or more of the peaks at about 5.6, 13.1, 14.5, 17.2, 18.2, 19.9, 20.2, 21.4, 21.6, 22.7, or 28.9, such as the peaks at about 5.6 and 13.1 °2 Θ , together with a DSC endotherm at about 160 °C or a

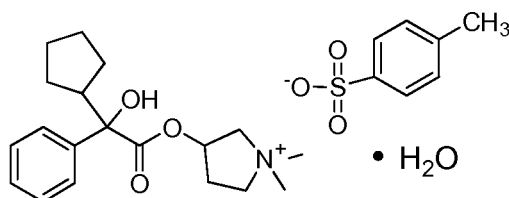
second endotherm at about 169 °C or both may be used to characterize glycopyrrolate hydrogen sulfate.

[0094] In a further embodiment, a crystalline salt of glycopyrrolate acetate is provided. An x-ray powder diffraction pattern substantially the same as the pattern of Figure 22 may be used to characterize one embodiment of crystalline glycopyrrolate acetate. A smaller subset of the peaks may be used to characterize crystalline glycopyrrolate acetate. For example, any one or more of the peaks, for example, at about 5.2, 10.4, 10.8, 11.3, 12.6, 15.4, 17.5, 19.1, or 23.6 °2 θ may be used to characterize crystalline glycopyrrolate acetate. For example, the peaks at about 5.2 and 11.3 °2 θ may be used to characterize glycopyrrolate acetate.

[0095] In another embodiment crystalline glycopyrrolate tosylate monohydrate is provided, also referred to herein as Form D glycopyrrolate tosylate or Form D or crystalline glycopyrrolate tosylate monohydrate. Exemplary preparations of Form D glycopyrrolate tosylate include Examples 8 and 9 herein. The ORTEP drawing of Form D glycopyrrolate tosylate, based on its crystal structure, is set forth in Figure 1. The chemical structure of Form D glycopyrrolate tosylate is set forth below as Formula I:

Formula I

[0096]



[0097] The XRPD pattern corresponding to Form D glycopyrrolate tosylate is represented by Figure 1. The crystal structure of the monoclinic Form D glycopyrrolate tosylate is set forth herein with the crystal data and acquisition parameter provided in Table 1.

[0098] Table 1. Crystal Data and Data Collection Parameters for Glycopyrrolate**Tosylate Form D**

formula	C ₂₆ H ₃₇ N ₃ O ₇ S
formula weight	507.65
space group	P2 ₁ /n (No. 14)
<i>a</i>	8.8715(5) Å
<i>b</i>	11.5849(7) Å
<i>c</i>	25.5323(14) Å
β	96.9 deg
<i>V</i>	2604.9(3) Å ³
<i>Z</i>	4
<i>d</i> _{calc} , g cm ⁻³	1.294
crystal dimensions, mm	0.23 x 0.20 x 0.18
temperature, K	150.
radiation (wavelength, Å)	Cu K α (1.54184)
monochromator	confocal optics
linear abs coef, mm ⁻¹	1.479
absorption correction applied	empirical ^a
transmission factors: min, max	0.592, 0.766
diffractometer	Rigaku RAPID-II
<i>h</i> , <i>k</i> , <i>l</i> range	0 to 10 0 to 13 -31 to 30
2 θ range, deg	3.49-140.48
mosaicity, deg	0.76
programs used	SHELXTL
<i>F</i> ₀₀₀	1088.0
weighting	
$1/[\sigma^2(F_o^2) + (0.1231P)^2 + 0.8250P]$ where $P = (F_o^2 + 2F_c^2)/3$	
data collected	24514
unique data	4024
<i>R</i> _{int}	0.086
data used in refinement	4024
cutoff used in <i>R</i> -factor calculations	$F_o^2 > 2.0\sigma(F_o^2)$
data with $I > 2.0\sigma(I)$	3812
number of variables	331
largest shift/esd in final cycle	0.00
<i>R</i> (<i>F</i> _o)	0.064
<i>R</i> _w (<i>F</i> _o ²)	0.185
goodness of fit	1.098

^a Otwinowski, Z.; Minor, W. *Methods Enzymol.* **1997**, 276, 307.^b Flack, H. D. *Acta Cryst.*, **1983** A39, 876.^c Hooft, R. W. W., Straver, L. H., and Spek, A. L. *J. Appl. Cryst.*, **2008**, 41, 96-103.

[0099] Form D glycopyrrolate tosylate was found to be monoclinic with space group P2₁/n. At 150K, the calculated density was found to be 1.294 grams per cubic centimeter. To two significant figures after the decimal, the unit cell dimensions were determined to be: *a* equals

about 8.87Å; b equals about 11.58Å; and c equals about 25.53Å, with corresponding unit cell angles of $\alpha = 90.00^\circ$, $\beta = 96.9^\circ$, and $\gamma = 90.00^\circ$. The Form D unit cell was found to be racemic with both R,S and S,R diastereomers of glycopyrrolate in the unit cell.

[00100] A pattern substantially the same as the pattern of Figure 2 may be used to characterize Form D glycopyrrolate tosylate. A smaller subset of the peaks identified in Figure 2 may instead be used to characterize Form D glycopyrrolate tosylate. For example, any one or more of peaks at about 6.9, 10.3, 12.6, 13.7, 14.9, 15.3, 15.7, 16.4, 17.7, 18.2, or $20.6^\circ 2\theta$ may be used to characterize Form D glycopyrrolate tosylate. For example, the single peak at about 6.9 or 10.3 or 12.6, or $20.6^\circ 2\theta$ may be used to characterize Form D glycopyrrolate tosylate. In another example, peaks at about 6.9 and $10.3^\circ 2\theta$ may be used to characterize Form D glycopyrrolate tosylate. In a further example, the peaks at about 6.9, 10.3, and $12.6^\circ 2\theta$ may be used to characterize Form D glycopyrrolate tosylate. In still another example, the peaks at about 10.3 and $12.6^\circ 2\theta$ characterize Form D glycopyrrolate tosylate. Table 2 identifies selected peaks from Figure 2. Intensity is provided for completeness.

[00101] Table 2 – Selected Peaks from Figure 2

Diffraction angle $^\circ(2\theta)$	d spacing (Å)	Intensity (%)
6.87 ± 0.20	12.867 ± 0.385	100
10.26 ± 0.20	8.620 ± 0.171	16
12.55 ± 0.20	7.052 ± 0.114	85
13.72 ± 0.20	6.454 ± 0.095	15
14.91 ± 0.20	5.943 ± 0.080	29
15.31 ± 0.20	5.788 ± 0.076	18
15.68 ± 0.20	5.653 ± 0.073	17
16.43 ± 0.20	5.396 ± 0.066	14
17.73 ± 0.20	5.002 ± 0.057	19
18.15 ± 0.20	4.888 ± 0.054	25
18.60 ± 0.20	4.770 ± 0.051	53
18.82 ± 0.20	4.716 ± 0.050	28
19.59 ± 0.20	4.532 ± 0.046	16
20.21 ± 0.20	4.395 ± 0.043	26
20.62 ± 0.20	4.307 ± 0.042	63
21.09 ± 0.20	4.212 ± 0.040	19
21.63 ± 0.20	4.109 ± 0.038	19
23.50 ± 0.20	3.786 ± 0.032	14
25.15 ± 0.20	3.541 ± 0.028	27

[00102] Further, Form D glycopyrrolate tosylate is distinguishable from Form C glycopyrrolate tosylate and the dehydrated form of Form D glycopyrrolate tosylate by the presence of water in the unit cell of Form D and may be so characterized.

[00103] Form D glycopyrrolate tosylate may also be characterized by the IR spectrum in Figure 3. When considering just IR spectroscopy, the entire IR spectrum may be used to characterize Form D glycopyrrolate tosylate or a subset of the spectrum may be so used. For example, any one or more of peaks at about 1734, 1196, 1125, 1036, 1013, and 682 cm^{-1} or others may be used alone or in combination to characterize Form D glycopyrrolate tosylate. Selected peaks from the IR spectrum in Figure 3 are set forth below in Table 3.

[00104] **Table 3 – Selected Peaks in the IR Spectrum of Form D in from Figure 3 in cm^{-1}**

682	1230
703	1265
713	1281
735	1312
750	1320
801	1329
815	1361
850	1373
856	1382
880	1445
908	1464
934	1476
940	1488
954	1495
975	1599
1013	1636
1024	1734
1036	2868
1075	2954
1084	2967
1125	3033
1139	3057
1155	3422
1182	3568
1196	

[00105] Form D glycopyrrolate tosylate may be characterized by both the IR and XRPD data as set forth herein. For example, Form D glycopyrrolate tosylate may be characterized by one or more XRPD peaks selected from, for example, about 6.9, 10.3, 12.6, 13.7, 14.9, 15.3, 15.7, 16.4, 17.7, 18.2, or 20.6 $^{\circ}2\theta$ and one or more of the IR peaks selected from, for example, about 1734, 1196, 1125, 1036, 1013, and 682 cm^{-1} .

[00106] Form D may be prepared by several methods. In one method, glycopyrrolate bromide is treated with a metal salt such as silver salt, of tosylate to form a glycopyrrolate salt. In particular Form D glycopyrrolate tosylate may be prepared by treating Ag-tosylate with glycopyrrolate-X in a suitable solvent to form a slurry; removing the solids from the slurry to obtain a solution; lyophilizing the solution to form a solid; dissolving the solid in a crystallization solvent; and removing the crystallization solvent to form Form D glycopyrrolate tosylate, wherein X is a halide. Suitable solvents are those that will afford a slurry when treating Ag-tosylate with glycopyrrolate-X. An example of a suitable solvent is an alcohol such as isopropanol. A crystallization solvent is a solvent, or mixtures thereof, which will dissolve sufficient solid provided after the lyophilizing stage such that when the crystallization solvent is removed, Form D glycopyrrolate is the resulting solid. An example of a crystallization solvent is a mixture of acetonitrile and water. Embodiments include where X is a halide such as iodide or bromide.

[00107] In some embodiments, the crystallization solvent is removed by lowering the temperature of the solid obtained after lyophilizing in solution and decanting the solvent. In these and other embodiments, an anti-solvent, such as toluene, is added to the solution containing the dissolved solid.

[00108] Form D glycopyrrolate tosylate may also be prepared by treating glycopyrrolate-Y and *p*-toluenesulfonic acid in a suitable solvent; removal of the solvent to form a solid; re-dissolving the solid in a crystallization solvent to form a solution and removing the crystallization solvent to form Form D glycopyrrolate tosylate wherein Y is an organic anion. An example of Y is acetate.

[00109] In some embodiments, an anti-solvent, such as toluene, is added to the solution containing the dissolved solid.

[00110] As disclosed in US 20100276329, glycopyrrolate bromide may be used to treat hyperhidrosis such as by using a wipe containing a solution of glycopyrrolate bromide. It is the glycopyrrolate cation (glycopyrronium) of the bromide salt which is the active clinical moiety because glycopyrronium has equivalent binding affinity for the M3 muscarinic acetylcholine receptor in vitro when delivered as either the bromide or another salt such as the tosylate salt. In one study, patients suffering from hyperhidrosis were treated with formulations containing 2% and 4% glycopyrrolate based on a glycopyrrolate bromide preparation. It was observed that

axillary sweating was reduced in patients during this study and a dose dependent trend in efficacy responses was also observed. Such dose dependency is consistent with glycopyrrolate's anti-muscarinic activity. Accordingly, glycopyrrolate tosylate may also be used to treat hyperhidrosis in patients such as by administering a topical containing glycopyrrolate tosylate. By topical, what is meant is a material or formulation comprising or containing glycopyrrolate tosylate which may be used to deliver glycopyrrolate tosylate, including a pharmaceutically effective amount of glycopyrrolate tosylate, to a patient. In many embodiments, the glycopyrrolate tosylate is three glycopyrrolate tosylate. Examples of a topical include, but are not limited to, solutions, ointments, gels, lotions, powders, sprays, creams, cream bases, patches, pastes, washes, dressings, masks, gauzes, bandages, swabs, brushes, or pads. The application of the topical may be controlled by controlling the dose amount or the rate of release. The dose may be controlled by dissolving or dispensing three glycopyrrolate tosylate, for example, in the appropriate medium. These and other dose controlling formulations may be used to deliver controlled doses such as specific unit doses, metered doses, or multiple doses from the topical.

[00111] In one embodiment, the topical is an absorbent pad. In such embodiments, such an absorbent pad may contain another topical such as a solution. As used herein, absorbent pads and nonwoven wipes are interchangeable and have the same meaning. In another embodiment, an absorbent pad containing three glycopyrrolate tosylate in solution may be used to treat hyperhidrosis. Further, pads or wipes containing one or more of glycopyrrolate benzoate, edisylate, oxalate, or hydrogen sulfate in solution may similarly be used to treat hyperhidrosis in patients. In another embodiment, the pharmaceutically acceptable solution of three glycopyrrolate tosylate is a topical.

[00112] In another embodiment, crystalline glycopyrrolate tosylate anhydrate is disclosed, also referred to herein as Form C glycopyrrolate tosylate or Form C. Exemplary preparations of Form C glycopyrrolate tosylate include Examples 11, 12, and 13 herein.

[00113] The x-ray powder diffraction pattern corresponding to Form C glycopyrrolate tosylate is provided in Figure 4. The infrared spectrum corresponding to Form C glycopyrrolate tosylate is provided in Figure 5. Form C was indexed to determine unit cell dimensions and the indexing solution is presented as Figure 6.

[00114] An x-ray powder diffraction pattern substantially the same as the pattern of Figure 4 may be used to characterize Form C glycopyrrolate tosylate. A smaller subset of the peaks

identified in Figure 4 may be used to characterize Form C glycopyrrolate tosylate. For example, any one or more of the peaks at about 5.5, 11.0, 11.8, 13.9, 14.9, 17.8, 19.6, 20.4, 21.6 and 22.1 $^{\circ}2\theta$ may be used to characterize Form C glycopyrrolate tosylate. For example, the single peaks at about 5.5 or 11.0 or 14.9 $^{\circ}2\theta$ may be used to characterize Form C glycopyrrolate tosylate, or any combination of the three. In another example, peaks at about 5.5 and 11.0 $^{\circ}2\theta$ may be used to characterize Form C glycopyrrolate. In a further example, the peaks at about 5.5, 11.0, and 14.9 $^{\circ}2\theta$ may be used to characterize Form C glycopyrrolate tosylate. Table 4 identifies selected peaks from Figure 4. Further, Form C glycopyrrolate tosylate is distinguishable from Form D glycopyrrolate tosylate since Form C lacks water in the unit cell. Intensity is provided for completeness.

[00115] Table 4 Selected Peaks from Figure 4

Diffraction angle °(2θ)	d spacing (Å)	Intensity (%)
5.47 ± 0.20	16.168 ± 0.614	100
10.98 ± 0.20	8.057 ± 0.149	34
11.82 ± 0.20	7.489 ± 0.128	13
13.87 ± 0.20	6.384 ± 0.093	20
14.86 ± 0.20	5.963 ± 0.081	82
17.75 ± 0.20	4.997 ± 0.056	67
17.92 ± 0.20	4.951 ± 0.055	53
18.12 ± 0.20	4.897 ± 0.054	35
19.60 ± 0.20	4.528 ± 0.046	51
20.39 ± 0.20	4.356 ± 0.043	42
21.59 ± 0.20	4.116 ± 0.038	27
22.14 ± 0.20	4.014 ± 0.036	26

[00116] Form C glycopyrrolate tosylate may also be characterized by the IR spectrum in Figure 5. When considering just IR spectroscopy, the entire IR spectrum may be used to characterize Form C glycopyrrolate tosylate or a subset of the spectrum may be so used. For example, any one or more of the peaks at about 1733, 1236, 1211, 1198, 1186, 1177, 1120, 1032, 1008, and 682 cm⁻¹ or others may be used alone or in combination to characterize Form C glycopyrrolate tosylate. Selected peaks from the IR spectrum in Figure 5 are set forth below in Table 5.

[00117] Table 5 Selected Peaks from Figure 5 in cm⁻¹

682	1120
706	1177
714	1186
742	1198
755	1211
786	1236
801	1293
821	1317
849	1446
886	1464
929	1475
938	1485
956	1597
980	1733
1008	2867
1032	2961
1075	3032

[00118] Form C glycopyrrolate tosylate may be characterized by both the IR and XRPD data as set forth herein. For example, Form C glycopyrrolate tosylate may be characterized by one or more XRPD peaks selected from, for example, about 5.5, 11.0, 11.8, 13.9, 14.9, 17.8, 19.6, 20.4, 21.6, and 22.1 °2 Θ and one or more of the IR peaks selected from, for example, 1733, 1236, 1211, 1198, 1186, 1177, 1120, 1032, 1008, and 682 cm⁻¹.

[00119] Form C may also be characterized by its thermal characteristics. For example, Form C exhibits a melting endotherm at about 168°C when measured with a Tzero™ pan type configuration at a heating rate of 10°C per minute from -30 °C to 250 °C.

[00120] Form C may be characterized by its DSC thermogram alone or in combination with the x-ray powder diffraction data, IR data, or both. For example, Form C glycopyrrolate tosylate may be characterized by a DSC thermogram having an endotherm at about 168°C and the x-ray powder diffraction pattern of Figure 4 and the IR spectrum of Figure 5. However, it is not necessary to use all of these data to characterize Form C when using DSC. For example, the single peak at about 5.5 °2 Θ and the DSC endotherm at about 168°C may be used to characterize Form C glycopyrrolate tosylate (see Figure 7). In another example, the peak at about 168°C and the IR peak at about 1733 cm⁻¹ may be used to characterize Form C glycopyrrolate tosylate. In yet another example, the endotherm at 168 °C, the x-ray powder diffraction peak at about 5.5 °2 Θ , and the IR peak at about 1733 cm⁻¹ may be used to characterize Form C glycopyrrolate tosylate.

[00121] Form C may be prepared by dehydrating Form D. Alternatively, Form C may be prepared by dissolving a glycopyrrolate salt such as, for example, at elevated temperatures such as about 50 °C. Slow cooling of the solution to room temperature followed by vacuum filtration and washing in a suitable organic solvent such as acetone results in the formation of Form C.

[00122] In a further embodiment, dehydrated forms of Form D are provided. An exemplary preparation of dehydrated Form D includes Example 10 herein. In one such embodiment, a dehydrated form of Form D, hereinafter referred to as dehydrated Form D, is provided wherein there is no water in the unit cell. An x-ray powder diffraction pattern of dehydrated Form D is provided in Figure 8. An overlay of the diffraction pattern showing dehydrated Form D and Form D is provided in Figure 9.

[00123] The indexing solution, with a Pawley refinement, to dehydrated Form D is presented in Figure 10 and indicates a unit cell which is of the same proportions, within experimental

variation, as with the indexing solution of Form D, also with a Pawley refinement (Figure 11) except for a loss of volume, which is consistent with water loss, and which results in a smaller unit cell. The indexing solution from Figure 11 presents a, b, and c parameters which correspond, respectively, to the c, b, and a parameters of the single crystal study (performed at 150 K) as set forth in Table 1.

[00124] The overlay pattern from Form D and dehydrated Form D show that there are some shifts between the two forms and that can also be seen in the comparison of the peak positions for selected Miller indices as set forth in Table 6 below. The differences in the Miller indices between Form D and dehydrated Form D confirm that they are different solid forms.

[00125] Table 6 – Select Miller Indices and Peak Comparisons between Form D and Dehydrated Form D

h	k	l	Form D (2 Θ)	Dehydrated (2 Θ)	Δ
2	0	0	6.84873	6.74897	-0.09976
1	1	0	8.16348	8.21407	0.05059
2	1	0	10.09605	10.08663	-0.00942
1	0	-1	10.22338	10.50517	0.28179
1	0	1	11.02323	11.37050	0.34727
0	1	1	12.50656	12.83560	0.32904
1	-1	-1	12.63742	12.91262	0.2752
2	0	2	22.15015	22.85492	0.70477
1	1	2	22.21449	22.92323	0.70874

[00126] Dehydrated Form D is further distinguishable from Form D since it lacks water of crystallization whereas Form D is a monohydrate and from Form C because the peaks of dehydrated Form D (an anhydrate) differ substantially from those in Form C (anhydrate). For example, as Table 6 indicates, dehydrated Form D has a peak at about 6.75° 2 Θ whereas the closest peak from Form C is at about 6.30 °2 Θ , a difference of 0.45°2 Θ . In addition, the indexing solution for Form C shows the unit cell to be triclinic whereas the unit cell of dehydrated Form D is monoclinic.

[00127] In another series of embodiments, variable hydrates, each with different water content in between dehydrated Form D and monohydrate Form D is provided. Such embodiments provide for a continuum of water content in between dehydrated Form D and Form D as illustrated with one example in Figure 9. One would expect that other materials with an

intermediate water content to generally exhibit x-ray powder diffraction pattern yielding peaks which are intermediate between Form D and dehydrated Form D.

[00128] In a further embodiment, the amorphous glycopyrrolate tosylate has an x-ray powder diffraction pattern exhibiting a figure substantially the same as Figure 19. In another embodiment, the amorphous glycopyrrolate tosylate of the invention has a glass transition temperature onset of about 11.6 °C. In yet another embodiment, the amorphous glycopyrrolate tosylate of the invention has an x-ray powder diffraction pattern substantially the same as in Figure 19 and a glass transition onset temperature of about 11.6 °C. In still an additional embodiment, the amorphous glycopyrrolate tosylate of the invention has an x-ray powder diffraction pattern exhibiting an amorphous halo but that is not substantially similar to that of Figure 19.

[00129] The amorphous glycopyrrolate tosylate of the invention was observed to be amorphous by X-ray diffraction in that it had contained the “amorphous halo” associated with amorphous solids. Such a material is often called “x-ray amorphous.” As used herein, “amorphous” when describing glycopyrrolate tosylate means amorphous as determined by x-ray powder diffraction such as, for example, as shown in Figure 19. DSC and thermogravimetric data for an x-ray amorphous form are shown in Figure 20 whereas the modulated DSC thermogram is set forth in Figure 23.

[00130] Amorphous glycopyrrolate tosylate was found to be extremely hygroscopic which deliquesced readily when exposed to standard atmospheric conditions. In addition, the extremely low glass transition renders the amorphous form of glycopyrrolate challenging to formulate successfully. However, applicants have been able to raise the glass transition temperature, and reduce the likelihood of deliquescence by preparing solid dispersions of glycopyrrolate tosylate.

[00131] Solid dispersions can be prepared in a number of different methods known in the art including lyophilization and spray drying. The solid dispersions herein were all created by lyophilization. The solid dispersions prepared herein are set forth in the Examples and may be prepared by combining a solution of glycopyrrolate tosylate with a solution of an excipient in one or more solvents where both components are soluble. The solutions may be filtered and are then cooled so that the solutions freeze. After freezing, the solutions are dried, such as in a lyophilizer, so as to form dispersions. The presence of a solid dispersion can be verified by comparing, for example, spectra of the starting materials with the purported dispersion or by

observing a glass temperature different than either of the components. A mixture would be evident by a simply linear combination of the peaks of the two starting materials whereas in a dispersion, peak shifts indicate the preparation of a different material, namely, a solid dispersion. A solid dispersion is also evident by the presence of a single glass transition temperature.

[00132] A solid dispersion comprising glycopyrrolate tosylate and excipients including monosaccharides, disaccharides, and pharmaceutically acceptable polymers containing cyclic ether moieties may be formed under suitable conditions such as by lyophilization. In some embodiments, such solid dispersions have a glass transition temperature of at least about 25°C including at least about 40 °C and at least about 60 °C. In these and other embodiments, the weight ratio of sucrose to glycopyrrolate tosylate is about 9:1. In other embodiments, the cyclic ethers are six-membered rings, such as in hypromellose acetate succinate (HPMCAS) and such solid dispersions have a glass transition temperature of at least about 25 °C including at least about 40 °C and at least about 60 °C. In these and other embodiments, the weight ratio of HPMCAS to glycopyrrolate tosylate is about 1:1.

[00133] A solid dispersion comprising glycopyrrolate tosylate and excipients including pharmaceutically acceptable polymers containing polyethylene glycol moieties such as a polyvinyl alcohol – polyethylene glycol graft copolymer, such as Kollicoat® IR, or a polyvinyl caprolactam-polyvinyl acetate-polyethylene glycol graft copolymer, such as Soluplus®, may be formed under suitable conditions such as by lyophilization. In some embodiments, such solid dispersions have a glass transition temperature of at least about 30°C including at least about 40°C.

[00134] A solid dispersion comprising glycopyrrolate tosylate and excipients including pharmaceutically acceptable polymers containing vinyl pyrrolidone moieties such as polyvinyl pyrrolidone or vinyl pyrrolidone-vinyl acetate copolymer may be formed under suitable conditions such as by lyophilization. In some embodiments, such solid dispersions have a glass transition temperature of at least about 25 °C including at least about 35 °C and further including about 60°C. Examples of polyvinyl pyrrolidone polymers used herein include PVP K29/32 and PVP K90. Examples of a vinyl pyrrolidone – vinyl acetate copolymer used herein include Kollidon® VA 64.

[00135] As used herein, the term “a pharmaceutically acceptable polymer” means a polymer approved for use in humans in pharmaceutical formulations including those polymers not yet approved but for whom approval is pending.

[00136] HPMCAS may be used to form a solid dispersion with glycopyrrolate tosylate in, for example, a ratio of about 1 to 1 of HPMCAS to glycopyrrolate by weight. An example of such a preparation can be found in Example 19. Figure 24 is an overlay infrared spectrum of a region of the spectrum showing differences between the dispersion and the component parts. For example, there is a peak at about 1211 cm^{-1} in the glycopyrrolate tosylate spectrum and a peak at about 1235 cm^{-1} in the HPMCAS spectrum. By comparison, in the solid dispersion spectrum, a single peak appears at about 1228 cm^{-1} indicating the material is not a physical mixture. This is confirmed with Figure 25 which shows a single glass transition temperature (also sometimes referred to as Tg) at about 42°C .

[00137] The solid 1:1 dispersion of HPMCAS and glycopyrrolate tosylate may be characterized by either its infrared spectrum, glass transition temperature or both. For example, a 1:1 solid dispersion of HPMCAS:glycopyrrolate tosylate may be characterized by a peak at about 1228 cm^{-1} , a glass transition temperature of about 42°C , or both.

[00138] Sucrose may be used to form a solid dispersion with glycopyrrolate tosylate in, for example, a ratio of about 9 to 1 of sucrose to glycopyrrolate by weight. An example of such a preparation can be found in Example 20. Figure 26 shows a single glass transition temperature at about 62°C confirming the presence of a solid dispersion. This glass transition temperature may be used to characterize the dispersion.

[00139] A polyvinyl alcohol-polyethylene glycol copolymer may be used to form a solid dispersion with glycopyrrolate tosylate, in, for example, ratios of between about 1:1 and 9:1. An example of the 1:1 dispersion preparation can be found in Example 21 and a 9:1 dispersion in Example 22. Figures 27a and 27b are overlay infrared spectra of two regions of the spectrum showing differences between the 1:1 dispersion and the component parts. For example, there are peaks at about 1107 cm^{-1} and 1322 cm^{-1} in the glycopyrrolate tosylate spectrum and peaks at about 1092 cm^{-1} and 1331 cm^{-1} in the polyvinyl alcohol-polyethylene glycol copolymer spectrum. By comparison, in the solid dispersion spectrum, a single peak now appears at about 1099 cm^{-1} and 1324 cm^{-1} respectively indicating the material is not a physical mixture. This is

confirmed with Figure 28 which shows a single glass transition temperature (T_g) at about 32 °C. Figure 29 shows a single glass transition temperature of the 9:1 dispersion to be at about 35 °C.

[00140] The 1:1 solid dispersion of a polyvinyl alcohol-polyethylene glycol copolymer and glycopyrrolate tosylate may be characterized by its infrared spectrum, glass transition temperature, or both. For example, one or more peaks in the infrared spectrum of the dispersion at about 1099 cm^{-1} and 1324 cm^{-1} , a glass transition temperature at about 32 °C, or a combination thereof may be used to characterize the solid dispersion. The 9:1 solid dispersion may be characterized by a glass transition temperature at about 35 °C.

[00141] A polyvinyl caprolactam-polyvinyl acetate-polyethylene glycol graft copolymer may be used to form a solid dispersion with glycopyrrolate tosylate in, for example, a ratio of about 1 to 1 of the polymer to glycopyrrolate by weight. An example of such a preparation can be found in Example 23. Figures 30a, 30b, and 30c are overlay infrared spectra of regions of the spectrum showing differences between the dispersion and the component parts. For example, there are peaks at about 938 cm^{-1} , about 1190 cm^{-1} , and about 1448 cm^{-1} in the glycopyrrolate tosylate spectrum and peaks at about 947 cm^{-1} , about 1197 cm^{-1} , and about 1442 cm^{-1} in the polyvinyl caprolactam-polyvinyl acetate-polyethylene glycol graft copolymer spectrum. By comparison, in the solid dispersion spectrum, a single peak appears at about 942 cm^{-1} , about 1195 cm^{-1} , and about 1445 cm^{-1} respectively indicating the material is not a physical mixture. This is confirmed with Figure 31 which shows a single glass transition temperature at about 40 °C.

[00142] The solid 1:1 dispersion of a polyvinyl caprolactam-polyvinyl acetate-polyethylene glycol graft copolymer and glycopyrrolate tosylate may be characterized by either its infrared spectrum, glass transition temperature or both. For example, a 1:1 solid dispersion of a polyvinyl caprolactam-polyvinyl acetate-polyethylene glycol graft copolymer:glycopyrrolate tosylate may be characterized by one or more peaks at about 942 cm^{-1} , about 1195 cm^{-1} , or about 1445 cm^{-1} , a glass transition temperature of about 40 °C, or a combination thereof.

[00143] A polyvinyl pyrrolidone polymer may be used to form a solid dispersion with glycopyrrolate tosylate in, for example, a ratio of about 1 to 1 of the polymer to glycopyrrolate by weight or a ratio of about 8 to 1 by weight. An example of such a preparation can be found in Examples 24, 25 and 26. Figures 32a, 32b, and 32c are overlay infrared spectra of regions of the spectrum showing differences between the dispersion and the component parts for Example 24.

For example, there are peaks at about 1283 cm^{-1} and 1651 cm^{-1} in the glycopyrrolate tosylate spectrum and peaks at about 1294 cm^{-1} , about 1465 cm^{-1} , and about 1641 cm^{-1} in the polyvinyl pyrrolidone spectrum of Example 24. By comparison, in the solid dispersion spectrum, peaks occur at about 1288 cm^{-1} , about 1461 cm^{-1} , and about 1664 cm^{-1} respectively indicating the material is not a physical mixture. Further, the dispersion exhibits a peak at about 1438 cm^{-1} , which has no counterpart peaks in either the polymer or the glycopyrrolate tosylate. In addition, Figure 33 indicates a solid dispersion showing a single glass transition temperature at about $38\text{ }^{\circ}\text{C}$ for the dispersion of Example 24. The 8:1 dispersion of Example 25 shows a single glass transition temperature of about $26\text{ }^{\circ}\text{C}$ as seen in Figure 34.

[00144] The 1:1 solid dispersion using the polyvinyl pyrrolidone of Example 24 may be characterized by its infrared spectrum, glass transition temperature, or both. For example, it may be characterized by one or more peaks in the infrared spectrum at 1288 cm^{-1} , about 1461 cm^{-1} , about 1664 cm^{-1} , or about 1438 cm^{-1} , a glass transition temperature of about $38\text{ }^{\circ}\text{C}$, or a combination thereof. The 8:1 solid dispersion using the polymer of Example 25 may be characterized by a glass transition temperature of about $26\text{ }^{\circ}\text{C}$.

[00145] The polyvinyl pyrrolidone of Example 26 was used to prepare an approximately 1 to 1 solid dispersion of polyvinyl pyrrolidone to glycopyrrolate tosylate. Figure 32d is an overlay infrared spectrum of a region of the spectrum showing differences between the dispersion and the component parts. For example, there is a peak at about 1650 cm^{-1} in the glycopyrrolate tosylate spectrum and a peak at about 1658 cm^{-1} in the polyvinyl pyrrolidone spectrum. By comparison, in the solid dispersion spectrum, a single peak appears at about 1664 cm^{-1} indicating the material is not a physical mixture. This is confirmed with Figure 35 which shows a single glass transition temperature at about $36\text{ }^{\circ}\text{C}$.

[00146] The solid 1:1 dispersion of a polyvinyl pyrrolidone polymer of Example 26 and glycopyrrolate tosylate may be characterized by either its infrared spectrum, glass transition temperature or both. For example, the solid dispersion may be characterized by a peak at about 1664 cm^{-1} a glass transition temperature of about $36\text{ }^{\circ}\text{C}$, or both.

[00147] A vinyl pyrrolidone – vinyl acetate copolymer may be used to form a solid dispersion with glycopyrrolate tosylate in, for example, a ratio of about 1 to 1 of a compound of the copolymer to glycopyrrolate by weight. An example of such a preparation can be found in Example 27. Figures 36a, 36b, and 36c are overlay infrared spectra of regions of the spectrum

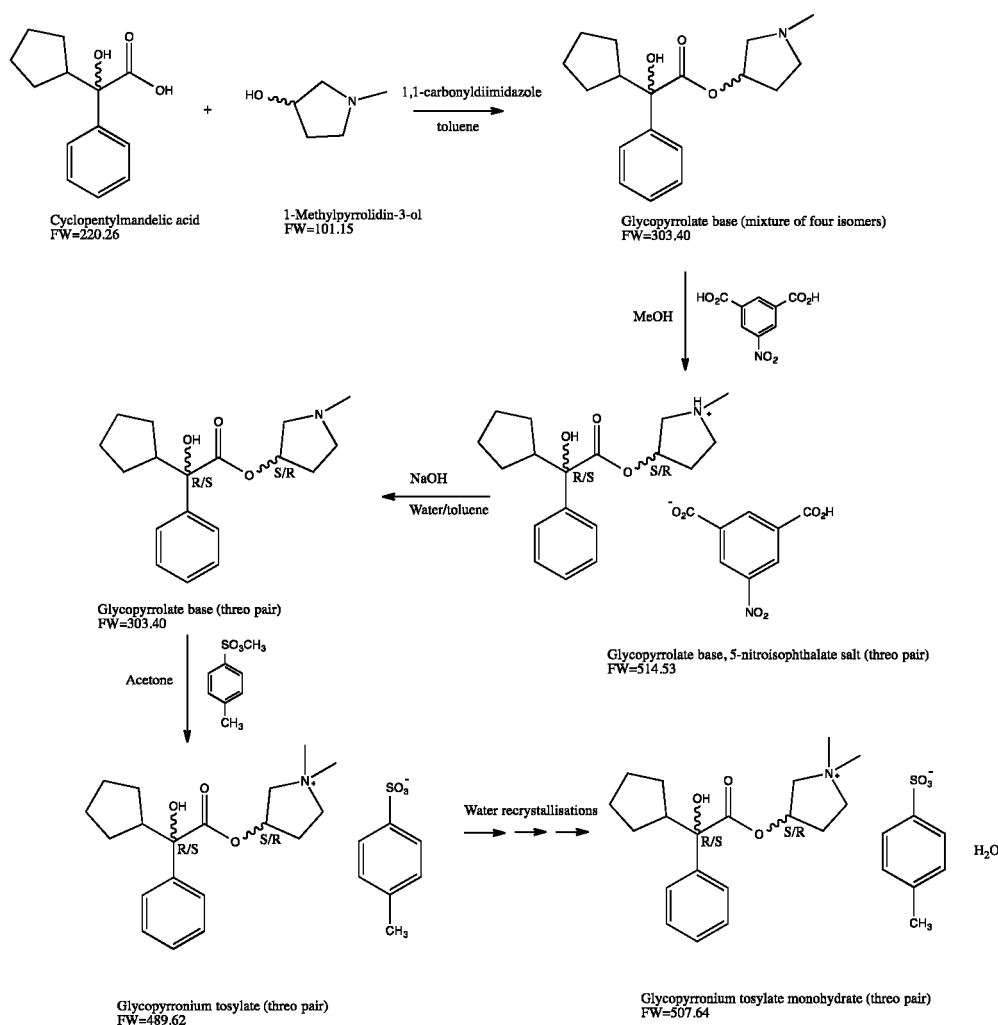
showing differences between the dispersion and the component parts. For example, there are peaks at about 1381 cm^{-1} , 1300 cm^{-1} , 1283 cm^{-1} , and 1651 cm^{-1} in the glycopyrrolate tosylate spectrum and peaks at about 1377 cm^{-1} , 1293 cm^{-1} , and 1641 cm^{-1} in the copolymer spectrum. By comparison, in the solid dispersion spectrum, a single peak appears at about 1371 cm^{-1} , 1287 cm^{-1} , and 1673 cm^{-1} respectively, indicating the material is not a physical mixture. This is confirmed with Figure 37 which shows a glass transition temperature at about 64°C .

[00148] The solid 1:1 dispersion of a compound of a vinyl pyrrolidone – vinyl acetate copolymer and glycopyrrolate tosylate may be characterized by either its infrared spectrum, glass transition temperature or both. For example, a 1:1 solid dispersion of vinyl pyrrolidone – vinyl acetate copolymer glycopyrrolate tosylate may be characterized by one or more peaks at about 1371 cm^{-1} , 1287 , and 1673 cm^{-1} , a glass transition temperature of about 64°C , or a combination thereof.

[00149] Threo glycopyrrolate tosylate may be prepared by treating racemic cyclopentylmandelic acid with racemic 1-methylpyrrolidin-3-ol and 1,1' carbonyldiimidazole in a suitable solvent, such as an organic solvent, to form glycopyrrolate base; treating the glycopyrrolate base in a suitable solvent, such as an alcohol, with a resolving acid to form a salt of threo glycopyrrolate; treating the salt of the threo glycopyrrolate salt with a suitable base in a suitable solvent, such as a mixture of organic solvents and water, to form a threo glycopyrrolate base; and treating the threo glycopyrrolate base with *p*-toluenesulfonic acid methyl ester, also known as methyl tosylate or methyl 4-benzenesulfonate in a suitable solvent, such as an organic solvent, to form threo glycopyrrolate tosylate. Subsequent treatment as disclosed herein may then be used to prepare, for example, Forms C, D, dehydrated D, or amorphous glycopyrrolate tosylate. In the case of Form D, such further treatment may include recrystallization in water. Care should be given when working with tosylate compounds since it is known in the art that aryl sulfonic acids, for example tosylic acid, may react with alcohols to form sulfonate esters, which are alkylating agents. Further, the equilibrium of the reaction is so surprisingly significantly displaced toward the dissociated tosylate anion that, even after spiking such a low amount as 1 ppm of ethyl tosylate into a 3% threo glycopyrrolate tosylate formulation, the ethyl tosylate levels diminish over time under long-term ($25^{\circ}\text{C}/60\%$ relative humidity) and accelerated ($40^{\circ}\text{C}/75\%$ relative humidity) stability conditions and are no longer detectable within weeks of the spike.

[00150] Suitable solvents for the preparation of the glycopyrrolate base include those where the cyclopentylmandelic acid, 1-methylpyrrolidin-3-ol and 1,1' carbonyldiimidazole are soluble such as toluene. The resolving acid is chosen so that the glycopyrrolate base formed, which is a mixture of four isomers, when treated with the resolving acid results in a salt which produces substantially the threo pair of isomers. The resolving acid may be dissolved in solvents such as an alcohol with one example being methanol. This reaction relies on the different solubilities of the salts produced. One such resolving acid is 5-Nitroisophthalic acid and it may be dissolved in methanol. Suitable bases for forming the threo glycopyrrolate base from the salt of the threo glycopyrrolate base and the resolving acid include hydroxides such as sodium hydroxide and such treatment may be done in a mixture of, for example, toluene and water. Treatment of the resulting base with *p*-toluenesulfonic acid methyl ester results in the desired threo glycopyrrolate tosylate. Suitable solvents include acetone and ethyl acetate. Water recrystallizations may then be used to form Form D of glycopyrrolate tosylate monohydrate. In some embodiments, seeding with Form D may assist in the formation of Form D. The Form D may then be dried in some embodiments. A general scheme for the synthesis of threo glycopyrrolate tosylate can be found in Scheme 1 which shows the ultimate formation of Form D.

Scheme 1 – Flow Diagram synthesis for Threo glycopyrrolate tosylate



[00151] In another embodiment, the present invention provides three glycopyrrolate tosylate in a topical which is not a solution such as ointment or a cream. An example of such a cream would be cetomacrogol cream. In another embodiment, the topical is a gel.

[00152] In one embodiment, the topical comprises three glycopyrrolate tosylate. In some of these embodiments, the topical further comprises buffers and/or may be in an aqueous solution. When buffers are used, said buffers may be, for example, citric acid and sodium citrate. The buffered topical may further comprise an alcohol such as ethanol.

[00153] In another embodiment, the present invention provides a pharmaceutically acceptable solution comprising three glycopyrrolate tosylate or a solvate thereof and one or more pharmaceutically acceptable additives. Such additives may include such co-solvents as ethanol and one or more pharmaceutically acceptable excipients.

[00154] In another embodiment, the pharmaceutically acceptable solution comprising threo glycopyrrolate tosylate or a solvate thereof is aqueous and further comprises one or more buffers. In many embodiments, the pharmaceutically acceptable solution is aqueous. Examples of buffers include, but are not limited to citric acid and sodium citrate dihydrate. The citric acid includes anhydrous citric acid. The solution may also contain one or more alcohols such as ethanol. Dehydrated ethanol is an alcohol that may be used. In one embodiment of the invention, the pharmaceutically acceptable aqueous solution comprising glycopyrrolate tosylate comprises about 0.15% by weight anhydrous citric acid, about 0.06% sodium citrate dihydrate by weight, between about 57 to about 59.5% by weight of dehydrated ethanol, and between about 1% and about 6% by weight glycopyrrolate tosylate.

[00155] In yet another embodiment, the topical is prepared so as to be film-forming. In such embodiments, a binding agent used. Examples of binding agents include povidones such as povidone K90. Such film-forming solutions further comprise one or more film-forming agents. Examples of film forming agents include butyl esters of a polyvinylmethylether/maleic anhydride acid copolymer. An example of such a film forming agent is the Gantrez™ ES-425 butyl ester copolymer.

[00156] In yet another embodiment, the aqueous pharmaceutically acceptable solution is prepared so as to be film-forming. In such embodiments, a binding agent used. Examples of binding agents include povidones such as povidone K90. Such film-forming solutions further comprise one or more film-forming agents. Examples of film forming agents include butyl esters of a polyvinylmethylether/maleic anhydride acid copolymer. An example of such a film forming agent is the Gantrez™ ES-425 butyl ester copolymer.

[00157] In some embodiments, the pharmaceutically acceptable solution is absorbed onto a carrier. For example, such a carrier may be a pad such as an absorbent pad or nonwoven wipe suitable for holding such solution when in storage as well as for application of the solution to desired areas of skin.

[00158] According to the present invention, the absorbent pad can be based on cotton fabric or non-cotton fabric. In one embodiment, the absorbent pad is based on synthetic nonwoven fabric, such as nonwoven rayon and polypropylene fabric. In one embodiment, the absorbent pad is a 75:25 rayon and polypropylene pad.

[00159] In some embodiments the absorbent pad material comprises polypropylene. In other

embodiments, the absorbent pad is substantially all polypropylene and in others, the pad is 100% polypropylene. Such pads may be nonwoven fabric with the following characteristics:

Table 7 Pad Properties

Physical Property	Characteristics
Basis Weight	1.231-1.254 ounces/yard ²
Machine Direction Grab Tensile	15.495-18.862 lbf (pounds-force)
Cross Direction Grab Tensile	14.425-16.190 lbf
Fiber Denier	2.443-2.569 dpf (denier per filament)

[00160] The pH of a topical such as a solution of glycopyrrolate tosylate, absorbed onto a pad is between 3.5 and 5.5 and often between about 4.0 and 5.0, including about 4 to 4.7 and about 4.1 to 4.6. For a glycopyrrolate tosylate monohydrate topical such as for a pad, the amount of glycopyrrolate tosylate monohydrate solution used in a pad is typically between about 2 g and 4 g including about 2.8 g or other pharmaceutically acceptable amounts.

[00161] A topical such as a solution may contain varying weight percents of glycopyrrolate tosylate such as glycopyrrolate tosylate monohydrate. In some embodiments, the weight percent of the glycopyrrolate tosylate, such as glycopyrrolate tosylate monohydrate, is between about 1 % and about 4%, including between 1.25% and about 4%, including between 2.5% and 3.75% and including each of about 1.25%, 2.5% and about 3.75%. The weight percents of glycopyrrolate tosylate, including glycopyrrolate tosylate monohydrate, may also be expressed in glycopyrronium weight percent only. For these weight percents, the weight percents may vary between about 0.6% and about 3.2%, including between about 1.6% and about 2.4% and including each of about 0.6%, 1.6% and about 2.4%. These weights are readily converted into weight percents of Form D. For example, 1.6% of glycopyrronium ion translates into 2.5% of Form D. The glycopyrrolate tosylate in any of the embodiments wherein they are absorbed onto the pads or are contained or comprised within the other topicals may be three glycopyrrolate tosylate. The topicals such as the absorbent pad containing a pharmaceutically acceptable solution can be applied to the area of the body to be treated.

[00162] Processes for making aqueous solutions of glycopyrrolate tosylate include treating solid glycopyrrolate tosylate in solution with water so as to dissolve the solid glycopyrrolate tosylate in solution. One may also add one or more buffers and/or alcohol, to the solution. The solution so obtained may then be wetted onto an absorbent pad so that a pharmaceutically

acceptable amount of glycopyrrolate tosylate has been absorbed onto the pad.. The alcohol may be ethanol such as dehydrated ethanol. The buffers may be citric acid and sodium citrate. In some embodiments, the glycopyrrolate tosylate or a solvate to be dissolved is in a crystalline form. Examples of such crystalline forms include Form C or Form D. In some embodiments, the glycopyrrolate tosylate or a solvate thereof is in an x-ray amorphous form. In other embodiments, pads containing a pharmaceutically acceptable aqueous solution of glycopyrrolate tosylate made by such processes are provided. The wetting may be done while the pad is in a pouch. In many embodiments, the pouch is heat-sealed after wetting. A typical pouch material is laminate containing aluminum foil as a layer. The glycopyrrolate tosylate of the processes herein may be threo glycopyrrolate tosylate.

[00163] In another embodiment, a pharmaceutically acceptable aqueous solution of glycopyrrolate tosylate may be prepared by dissolving glycopyrrolate tosylate in a mixture of water with ethanol. One or more pharmaceutically acceptable excipients can be added either prior to or after the addition of the glycopyrrolate tosylate or a solvate thereof and the aqueous solvent. Said glycopyrrolate tosylate may be threo glycopyrrolate tosylate.

[00164] The pharmaceutically acceptable solution of glycopyrrolate tosylate or a solvate thereof is therapeutically useful. For example, the pharmaceutically acceptable solution can be used for treating hyperhidrosis or reducing sweating in mammals. The pharmaceutically acceptable solution is typically applied from a pad on which the solution is absorbed. In one embodiment, the present invention provides a method of treating hyperhidrosis in a mammal by topically administering to the skin of the mammal a therapeutically effective amount of a pharmaceutically acceptable solution of glycopyrrolate tosylate or a solvate thereof. In one embodiment, the mammal is a human. The pharmaceutically acceptable solution can be applied to one or several areas or even the whole body including, but not limited to, the hands, *e.g.*, palms; axillae; feet, *e.g.*, soles; groin; face, *e.g.*, cheeks and forehead; and trunk, *e.g.*, back and abdomen, or scalp. In some embodiments, methods of treating primary axillary hyperhidrosis with glycopyrrolate tosylate or a solvate thereof comprising topically administering a therapeutically effective amount of an aqueous glycopyrrolate tosylate solution to the skin of a mammal in need thereof. In many embodiments, such administration may be with an absorbent pad. In other embodiments, methods of treating palmar or plantar hyperhidrosis with glycopyrrolate tosylate or a solvate thereof are provided. Dosing of glycopyrrolate tosylate may

be daily. Said pharmaceutically acceptable solution of glycopyrrolate tosylate may be three glycopyrrolate tosylate.

[00165] Instrumental Techniques Used in the Examples

[00166] X-ray Powder Diffraction (XRPD)

[00167] X-ray Powder Diffraction (XRPD) – Reflection Geometry

[00168] XRPD patterns were collected with a PANalytical X'Pert PRO MPD diffractometer using an incident beam of Cu K α radiation produced using a long, fine-focus source and a nickel filter. The diffractometer was configured using the symmetric Bragg–Brentano geometry. Prior to the analysis, a silicon specimen (NIST SRM 640d) was analyzed to verify the observed position of the Si 111 peak is consistent with the NIST-certified position. A specimen of the sample was prepared as a thin, circular layer centered on a silicon zero-background substrate. Antiscatter slits (SS) were used to minimize the background generated by air. Soller slits for the incident and diffracted beams were used to minimize broadening from axial divergence. Diffraction patterns were collected using a scanning position-sensitive detector (X'Celerator) located 240 mm from the sample and Data Collector software v. 2.2b. The data acquisition parameters for each pattern were: Cu (1.54059Å) x-ray tube, 45 kV voltage, 40 mA amperage, 3.50 – 40.00 °2 θ scan range, 0.017 or 0.08 °2 θ step size, 1835 - 1947 s collection time, 1.1 or 1.2 °/min scan speed, 1/8 ° divergence slit (DS), 1/4 ° incident-beam antiscatter slit (SS), 0.0 null revolution time.

[00169] X-ray Powder Diffraction (XRPD) – Transmission Geometry

[00170] XRPD patterns were collected with a PANalytical X'Pert PRO MPD diffractometer using an incident beam of Cu radiation produced using an Optix long, fine-focus source. An elliptically graded multilayer mirror was used to focus Cu K α x-rays through the specimen and onto the detector. Prior to the analysis, a silicon specimen (NIST SRM 640d) was analyzed to verify the observed position of the Si 111 peak is consistent with the NIST-certified position. A specimen of the sample was sandwiched between 3- μ m-thick films and analyzed in transmission geometry. A beam stop, short antiscatter extension, and antiscatter knife edge were used to minimize the background generated by air. Soller slits for the incident and diffracted beams were used to minimize broadening from axial divergence. Diffraction patterns were collected using a scanning position-sensitive detector (X'Celerator) located 240 mm from the specimen and Data Collector software v. 2.2b. The data acquisition parameters for each pattern were: Cu

(1.54059Å) x-ray tube, 45 kV voltage, 40 mA amperage, 1.0 – 39.99 °2θ scan range, 0.017 °2θ step size, 717 – 721 s collection time, 3.3 or 3.2 °/min scan speed, 1/2 ° divergence slit (DS), null incident-beam antiscatter slit (SS), 1.0 null revolution time.

[00171] Variable Temperature X-ray Powder Diffraction (VT-XRPD)

[00172] XRPD patterns were collected with a PANalytical X'Pert PRO MPD diffractometer using an incident beam of Cu Kα radiation produced using a long, fine-focus source and a nickel filter. The diffractometer was configured using the symmetric Bragg–Brentano geometry. Data were collected and analyzed using Data Collector software v. 2.2b. Prior to the analysis, a silicon specimen (NIST SRM 640d) was analyzed to verify the Si 111 peak position. A specimen of the sample was packed into a nickel-coated copper well. Antiscatter slits (SS) were used to minimize the background generated by air scattering. Soller slits for the incident and diffracted beams were used to minimize broadening from axial divergence. Diffraction patterns were collected using a scanning position-sensitive detector (X'Celerator) located 240 mm from the sample. The data acquisition parameters for each pattern were: Cu (1.54059Å) x-ray tube, 45 kV voltage, 40 mA amperage, 3.50 – 26.00 °2θ scan range, 0.008 °2θ step size, 1869 s collection time, 0.7 °/min scan speed, 1/8 ° divergence slit (DS), 1/4 ° incident-beam antiscatter slit (SS), 0.0 null revolution time.

[00173] An Anton Paar TTK 450 stage was used to collect *in situ* XRPD patterns at non-ambient temperatures. The sample was heated with a resistance heater located directly under the sample holder, and the temperature was monitored with a platinum 100 resistance sensor located directly under the specimen. The power to the heater was supplied and controlled by an Anton Paar TCU 100 interfaced with Data Collector.

[00174] Infrared Spectroscopy (IR)

[00175] IR spectra were acquired on Nicolet 6700 Fourier transform infrared (FT-IR) spectrophotometer (Thermo Nicolet) equipped with an Ever-Glo mid/far IR source, an extended range potassium bromide (KBr) beamsplitter, and a deuterated triglycine sulfate (DTGS) detector. Wavelength verification was performed using NIST SRM 1921b (polystyrene). An attenuated total reflectance (ATR) accessory (Thunderdome™, Thermo Spectra-Tech), with a germanium (Ge) crystal was used for data acquisition. Each spectrum represents 256 co-added scans collected at a spectral resolution of 2 cm⁻¹. A background data set was acquired with a

clean Ge crystal. A Log 1/R (R = reflectance) spectrum was obtained by taking a ratio of these two data sets against each other.

[00176] Differential Scanning Calorimetry (DSC)

[00177] DSC was performed using a TA Instruments 2920 differential scanning calorimeter. Temperature calibration was performed using NIST-traceable indium metal. The sample was placed into an aluminum DSC pan, covered with a lid, and the weight was accurately recorded. A weighed aluminum pan configured as the sample pan was placed on the reference side of the cell. Modulated DSC data (*see, e.g.*, Figure 23) were obtained on a TA Instruments Q2000 differential scanning calorimeter equipped with a refrigerated cooling system (RCS). Temperature calibration was performed using NIST-traceable indium metal. The sample was placed into an aluminum DSC pan, and the weight was accurately recorded. The pan was covered with a lid perforated with a laser pinhole, and the lid was hermetically sealed. A weighed, crimped aluminum pan was placed on the reference side of the cell. Data were obtained using a modulation amplitude of ± 1 °C and a 60 second period with an underlying heating rate of 2 °C/minute from -50 to 220 °C. The reported glass transition temperatures are obtained from the inflection point of the step change in the reversing heat flow versus temperature curve.

[00178] Proton Nuclear Magnetic Resonance (¹H NMR)

[00179] The solution NMR spectra were acquired with a Varian UNITYINOVA-400 spectrometer. The samples were prepared by dissolving a small amount of sample in DMSO-d₆ containing TMS.

[00180] Pawley Refinement

[00181] Indexing and subsequent Pawley refinement provides the most accurate determination of unit cell volume and cell parameters from XRPD data. These computations were performed using TOPAS 4.2, 2009, Bruker AXS GmbH, Karlsruhe, Germany. The background was modeled using a 3rd order Chebychev polynomial. Peak shape was modeled using Lorentzian crystallite size broadening and axial divergence was modeled using the full axial model. Peak positions were allowed to vary by fitting the unit cell parameters. Whole pattern Pawley refinement was performed on all parameters simultaneously to a convergence of 0.001 in χ^2 .

[00182] Thermogravimetric Analysis (TGA)

[00183] TG analyses were performed using a TA Instruments 2950 thermogravimetric analyzer. Temperature calibration was performed using nickel and Alumel™. Each sample was placed in an aluminum pan and inserted into the TG furnace. The furnace was heated under a nitrogen purge. The data acquisition parameters are displayed above each thermogram in the Data section of this report. The method code on the thermogram is an abbreviation for the start and end temperature as well as the heating rate; *e.g.*, 25-350-10 means "from 25 to 350 °C, at 10 °C/min."

[00184] Examples

[00185] Example 1 --Salt Screen

[00186] Fourteen salts were targeted; however, only six glycopyrrolate salts were successfully isolated and characterized: acetate, benzoate, edisylate, oxalate, hydrogen sulfate, and tosylate. These salts were formed either by (1) reaction of glycopyrrolate bromide with silver salts of salt formers, or (2) reaction of glycopyrrolate acetate with salt former acids.

[00187] Example 2 – Glycopyrrolate Benzoate

[00188] The glycopyrrolate benzoate salt was prepared only once using route (1) from Example 1. Glycopyrrolate benzoate was generated on reacting saturated aqueous solutions of each glycopyrrolate bromide with silver benzoate at approximately 92 °C, followed by filtration and subsequent lyophilization of the mother liquor. The material was then recrystallized in acetone/MTBE (1/2, vol/vol) and sonicated to form white crystalline solids. An XRPD pattern associated with this material is in Figure 12. Proton NMR showed the presence of equimolar amounts of the glycopyrrolate and benzoate species, as well as water. Thermal analysis of the sample showed a single endotherm with a peak maximum of 79 °C in the DSC thermogram concomitant with a 3.5 wt % loss between 25 and 80 °C in the TG trace. The weight loss was equivalent to approximately one mole of water indicating the formation of a monohydrate.

[00189] Example 3 – Di-Glycopyrrolate edisylate

[00190] Di-glycopyrrolate Edisylate salt was formed using process (2) from Example 1. A second molar equivalent of glycopyrrolate acetate was added to the reaction mixture of glycopyrrolate acetate and a minor amount of silver acetate and one molar equivalent of 1,2-ethanedisulfonic acid in ethyl acetate/isopropanol (83/17, vol/vol). The mixture was stirred for approximately five minutes before the resulting grey solids were isolated and dried under vacuum at ambient temperature for one day. The dried solids were crystalline with a minor

amount of silver acetate by XRPD (Figure 14). The XRPD pattern was successfully indexed which indicated that the material was composed of a single crystalline phase. Proton NMR spectroscopy confirmed the presence of two moles of glycopyrrolate per mole of edisylate, and water. Thermal analysis of the sample showed a 3.8 wt % loss between 25 and 95 °C in the TG trace and an endotherm with a peak maximum at 103 °C in the DSC thermogram. The mass loss equates to approximately two moles water indicating a dihydrate.

[00191] Example 4 – Glycopyrrolate Oxalate

[00192] Glycopyrrolate oxalate was prepared using process (2) from Example 1. Equimolar amounts of oxalic acid and glycopyrrolate acetate were dissolved in methanol then fast evaporated and dried under vacuum. The resulting glassy, gel-like material was recrystallized by slurrying in ethyl acetate to produce grey solids that were then dried under vacuum before analysis by XRPD and proton NMR spectroscopy. The XRPD pattern can be found in Figure 16.

[00193] Example 5 – Glycopyrrolate Hydrogen Sulfate

[00194] Glycopyrrolate hydrogen sulfate was prepared as a mixture with a trace amount of silver sulfate using process (2) from Example 1. Equimolar amounts of glycopyrrolate acetate and sulfuric acid were stirred in anhydrous ethyl acetate for approximately one day before the resulting material was isolated and dried under vacuum. The solids were characterized by XRPD, proton NMR spectroscopy, thermal techniques and elemental analysis. The XRPD pattern was unique and contained a trace amount of silver sulfate (Figure 17). The XRPD pattern was successfully indexed except for the silver sulfate peak at $28.35^{\circ}2\theta$, indicating that the glycopyrrolate hydrogen sulfate salt was composed of a single crystalline phase. The silver sulfate was likely formed from the silver acetate present in the glycopyrrolate acetate starting material. The NMR spectrum was consistent with a 1:1 ratio of a glycopyrrolate and hydrogen sulfate. Thermal analysis showed a major sharp endotherm with a peak maximum at 160 °C and a second endotherm with a peak maximum at 169°C, and a negligible weight loss of 0.2 wt % between 25 and 180 °C. Elemental analysis confirmed the anhydrous salt stoichiometry.

[00195] Example 6 – Glycopyrrolate tosylate

[00196] In a dark room, silver tosylate (3.5 g) was dissolved in water (~100 mL) by sonication. The solution was heated to approximately 40 °C and additional water was added (~15 mL). An equimolar amount of glycopyrrolate bromide (5 g) (mixture of R,S and S,R diastereomers) was added and immediately resulted in a yellow precipitate. The slurry was

stirred at approximately 40 °C overnight, and then slowly cooled while stirring to ambient temperature. At ambient temperature, the solids were vacuum filtered and the wet cake was washed three times with approximately 10 mL of water. The mother liquor was collected and filtered two times through a 0.2 µm nylon filter with glass microfiber (GMF). A clear solution was observed after filtration and was lyophilized at approximately -50 °C. After 6 days, a mixture of white, needle-like and slightly sticky, glassy solids was observed. Toluene (~20 mL) was added, and the slurry was briefly sonicated and then stirred at ambient temperature. Additional toluene (~80 mL) was added for easier stirring, and the mixture was allowed to stand at ambient conditions for 1 day. Solids of glycopyrrolate tosylate were collected by vacuum filtration and vacuum drying at ambient temperature for 1 day.

[00197] Example 7 – Preparation of Glycopyrrolate Tosylate

[00198] A slurry of equimolar amounts of glycopyrrolate acetate and *p*-toluenesulfonic acid was prepared in isopropanol (1 mL). The mixture was stirred at ambient temperature. Additional isopropanol (0.5 mL) was added to improve stirring, and the mixture was stirred overnight. Solids of glycopyrrolate tosylate were isolated by vacuum filtration and analyzed.

[00199] Example 8 –Preparation of Glycopyrrolate Tosylate Form D

[00200] Glycopyrrolate tosylate (1.0569 g) made from Example 6 was dissolved in 4 mL ACN/H₂O (50/50 vol/vol) by sonication. The solution was filtered through 0.2 µm nylon filter into a clean vial. The solvent was allowed to partially evaporate from an open vial under ambient conditions. Further evaporation was subsequently performed under nitrogen gas flow. A gel resulted which was vacuum dried at 40 °C for 1 day. Toluene (5 mL) was added and the mixture was sonicated for approximately 10 minutes causing white solids to precipitate. The mixture was stirred at ambient temperature for 1 day. The solids were isolated by vacuum filtration and the wet cake was washed with approximately 10 mL of toluene. The solids were vacuum dried at ambient temperature for 1 day. After vacuum drying the solids were placed in a vial which remained uncapped and placed inside a relative humidity chamber (~97 %). The chamber was placed inside an oven at 41 °C. After 6 days, the solids were analyzed by XRPD showing Form D.

[00201] Example 9 – Single Crystal Preparation of Form D

[00202] Glycopyrrolate tosylate (54.9 mg) made from Example 6 was dissolved in EtOAc/DMF (87/13 vol/vol) at approximately 55 °C at 24 mg/ml. The solution was hot filtered

through a 0.2 μm nylon filter into a pre-warmed vial. The vial containing the solution was first placed in a dry ice/acetone bath and then in a freezer (approximately -25 to -10 $^{\circ}\text{C}$). After 3 days, the solution was re-heated to approximately 50 $^{\circ}\text{C}$ and additional EtOAc was added for 96/4 EtOAc/DMF (vol/vol) at 7 mg/ml. The solution was quickly removed from elevated temperature and placed in the freezer. Solids were isolated by decanting the solvent and drying the solids under ambient conditions.

[00203] Single Crystal Data Collection

[00204] A colorless chunk of $\text{C}_{26}\text{H}_{37}\text{NO}_7\text{S}$ [$\text{C}_7\text{H}_7\text{O}_3\text{S}$, $\text{C}_{19}\text{H}_{28}\text{NO}_3$, H_2O] having approximate dimensions of $0.23 \times 0.20 \times 0.18$ mm, was mounted on a fiber in random orientation. Preliminary examination and data collection were performed with Cu $\text{K}\alpha$ radiation ($\lambda = 1.54184$ $\text{\AA}$) on a Rigaku Rapid II diffractometer equipped with confocal optics. Refinements were performed using SHELX97.

[00205] Example 10 – Preparation of Dehydrated Form D

[00206] A mixture of glycopyrrolate tosylate solids, including Form C and Form D, and a trace amount of silver tosylate was kept over P_2O_5 at ambient temperature for 18 days. The resulting solids were composed of a mixture of dehydrated Form D with a trace of silver tosylate as shown by XRPD analysis.

[00207] Example 11 – Preparation of Form C Glycopyrrolate Tosylate

[00208] Glycopyrrolate tosylate Form D, containing trace amounts of Form C and silver tosylate, was heated on an Anton Paar TTK 450 stage and XRPD patterns were collected *in situ* in the range 3.5 – 26° (2θ). All heating steps were at approximately 10 $^{\circ}\text{C}/\text{min}$. The stage was heated in incremental steps of 20 $^{\circ}\text{C}$ from 25 to 125 $^{\circ}\text{C}$. At each step, an XRPD pattern was collected over approximately 4 minutes. The stage was then heated to 135 $^{\circ}\text{C}$ and an XRPD pattern was collected over approximately 16 minutes and after heating further to 145 $^{\circ}\text{C}$, a pattern was collected in approximately 31 minutes. The sample was subsequently cooled to 25 $^{\circ}\text{C}$ at approximately 24 $^{\circ}\text{C}/\text{min}$, upon which a final XRPD pattern was collected over approximately 16 min. The XRPD pattern of this final pattern was indexed as Form C.

[00209] Example 12 – Preparation of Form C Glycopyrrolate Tosylate

[00210] Glycopyrrolate tosylate Form D from Example 6 was heated to an approximate temperature in the range 143 – 149 $^{\circ}\text{C}$ under a continuous nitrogen purge for approximately 3.3 hours. The vial containing the solids was capped, placed on a lab bench and allowed to cool

down to room temperature. At room temperature, the vial was placed in a jar containing P_2O_5 . The sample was prepared for XRPD analysis under nitrogen which confirmed production of Form C.

[00211] Example 13 – Preparation of Form C Glycopyrrolate Tosylate

[00212] Glycopyrrolate tosylate (59.5 mg) from Example 6 was dissolved in acetone at approximately 50 °C at 27 mg/ml. The solution was hot filtered through a 0.2 µm nylon filter into a pre-warmed vial. The vial was capped and left on the hot plate which was subsequently turned off to allow the sample to cool slowly to ambient temperature. At ambient temperature the solution was stirred causing white solids to precipitate. The solids were isolated by vacuum filtration and the wet cake was washed with approximately 2 ml of acetone. XRPD analysis resulted in Form C.

[00213] Example 14 – Amorphous glycopyrrolate tosylate

[00214] Glycopyrrolate tosylate from Example 6 was melted and cooled repeatedly until the majority of the solids had the appearance of a glass by microscopy. XRPD analysis indicated that the "glassy" sample was observed to be amorphous. A 2.2% weight loss was observed by TGA from 25 to 250 °C of the amorphous glycopyrrolate tosylate. The onset of the glass transition temperature was measured at 11.6 °C.

[00215] Example 15 – Preparing Crude Threo Glycopyrrolate Tosylate

[00216] Cyclopentylmandelic acid is combined with 1,1'-carbonyldiimidazol in toluene and is heated and stirred. N-methyl-3-pyrrolidinol is added while stirring and heated in toluene. The reaction mixture is then cooled and washed with purified water. The isolate toluene layer is then reduced to a concentrate of the glycopyrrolate base.

[00217] 5-Nitroisophthalic acid (1 eq.) is dissolved in methanol (20 vol) at room temperature with moderate agitation. The glycopyrrolate base (1 eq.) obtained above is then added. Once crystallization is initiated, the mixture is stirred at room temperature. The solids are then recovered in a filtration centrifuge and washed with methanol. The crude product is then suspended in methanol and stirred at approximately 65 °C for one hour, then cooled to 20 °C and stirred for a further 4 hours. The product is again recovered, washed with methanol, partially dried and discharged as wet glycopyrrolate 5-nitroisophthalate. The ratio of threo:erythro diastereomeric pairs is typically 96:4. The threo-glycopyrrolate base is obtained by treatment of the wet 5-nitroisophthalate salt with aqueous sodium hydroxide and toluene.

[00218] The threo-glycopyrrolate base is dissolved in acetone and treated with a slight excess of methyl-*p*-toluenesulfonate. The completion of reaction is monitored by TLC until the remaining base is NMT 2%. The crude glycopyrrolonium tosylate is recovered and washed twice with acetone. The wet cake obtained is dried under vacuum at elevated temperature.

[00219] Example 16 Pure Threo Glycopyrrolate Tosylate

[00220] The product of Example 15 is triturated in purified water and recovered and washed with cold purified water. The wet cake is then dissolved in water with agitation. The solution obtained is cooled and held until crystallization begins. The mixture is then further cooled and agitated and the product is recovered and washed with cold purified water. The product then undergoes a second recrystallization under similar conditions. The product is tray dried at not more than 40 °C without vacuum for a minimum time until the water content is between about 2.5%-4.0%.

[00221] Example 17a – Preparing an Aqueous Solution of Glycopyrrolate Tosylate

[00222] To a vessel of appropriate size, add purified water, citric acid and sodium citrate dihydrate and dissolve by mixing. Add dehydrated alcohol; initiate mixing and continue to mix until a homogenous clear solution is obtained. Continue mixing and add solid glycopyrrolate tosylate and mix until the glycopyrrolate tosylate is dissolved and the solution is homogenous. The solution should be clear and colorless or pale yellow with a pH of between about 4.0 and about 5.0 at about 25°C.

[00223] Example 17b – Preparing an Aqueous Solution of Threo Glycopyrrolate Tosylate using Form D

[00224] To a vessel of appropriate size, add purified water, citric acid and sodium citrate dihydrate and dissolve by mixing. Add dehydrated alcohol; initiate mixing and continue to mix until a homogenous clear solution is obtained. Continue mixing and add Form D glycopyrrolate tosylate and mix until the glycopyrrolate tosylate is dissolved and the solution is homogenous. The solution should be clear and colorless or pale yellow with a pH of between about 4.0 and about 5.0 at about 25°C.

[00225] Example 18 – Filling a Pouch and Pad

[00226] Each pouch is formed, heat sealed on three sides; bottom, and outer edges. A pad is folded and cut to size, and with a final fold in half, one pad is inserted into each preformed pouch through the open top. About 2.8g of the glycopyrrolate tosylate product of Example 15 is added

through the open top of each pouch wetting the enclosed pad. The top side of the pouch is heat sealed.

[00227] General Preparation of Solid Dispersions

[00228] Solutions of excipient and glycopyrrolate tosylate Form D were dissolved in water, ethanol/water or dioxane/water, filtered through a 0.2- μ m nylon fiber membrane, dropwise, into a vial submerged in a liquid nitrogen bath. The addition rate of the solution was monitored so that each drop of the sample was frozen prior to the next drop being added. The samples were placed on dry ice and immediately transported to a LABCONCO Triad Series lyophilizer and dried. After drying, the solids were isolated and stored over desiccant in a freezer. All samples were removed from the freezer and warmed to ambient temperature in a desiccator prior to analysis. Attempts were made to limit the amount of time the sample experienced at ambient humidity prior to analysis. Excipients were purchased from commercial suppliers and used as received including: PVP K-29/32 ISP Technologies, Inc. Wayne N.J; Kollicoat IR, Kollidon VA 64: BASF SE, Ludwigshafen, Germany; HPMCAS: Shin-Etsu Chemical Company Ltd., Tokyo, Japan; PVP K-90: Sigma-Aldrich, Inc., St. Louis MO., USA. This general procedure was followed in the examples set forth Table 8 below using the weights of excipient, glycopyrrolate tosylate and solvent choice as indicated.

Table 8 Solid Dispersion Examples

Example Number	Excipient (Loading)	Wt. Glycopyrrolate Tosylate (mg)	Wt. Excipient (mg)	Solvent
19	HPMCAS (1:1)	79.1	79.6	H ₂ O/ Dioxane [1:3]
20	Sucrose (9:1)	10.3	89.7	H ₂ O
21	Kollicoat® IR (1:1)	82.0	80.1	H ₂ O
22	Kollicoat® IR (9:1)	10.5	90.2	H ₂ O
23	Soluplus® (1:1)	78.2	79.5	H ₂ O/ Dioxane [1:1]
24	PVP K-29/32 (1:1)	80.4	80.5	H ₂ O
25	PVP K-29/32 (8:1)	11.3	89.5	H ₂ O
26	PVP K-90 (1:1)	79.5	80.4	H ₂ O/ EtOH [5:1]

27	Kollidon® VA 64 (1:1)	78.5	79.8	H ₂ O
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[00229] All examples presented are representative and non-limiting. The above-described embodiments may be modified or varied, without departing from the invention, as appreciated by those skilled in the art in light of the above teachings. It is therefore to be understood that, within the scope of the claims and their equivalents, the invention may be practiced otherwise than as specifically described.

CLAIMS

1. A racemic mixture of (R)-3-((S)-2-cyclopentyl-2-hydroxy-2-phenylacetoxy)-1,1-dimethylpyrrolidinium 4-methylbenzenesulfonate and (S)-3-((R)-2-cyclopentyl-2-hydroxy-2-phenylacetoxy)-1,1-dimethylpyrrolidinium 4-methylbenzenesulfonate.
2. A pharmaceutically acceptable solution comprising a compound of claim 1 or solvate thereof and or more pharmaceutically acceptable additives.
3. The solution of claim 2 wherein an additive is ethanol.
4. The solution of claim 3 further comprising one or pharmaceutically acceptable excipients.
5. A carrier comprising a pharmaceutically acceptable solution of claim 2.
6. The carrier of claim 5 wherein the carrier is an absorbent pad.
7. The solution of claim 2 for topical administration.
8. A method of treating hyperhidrosis comprising topically administering a therapeutically effective amount of a pharmaceutically acceptable solution of glycopyrrolate tosylate or solvate thereof to the skin of a mammal.
9. The method of claim 8, wherein the mammal is human.
10. The method of claim 9, wherein the solution is applied to the axilla.
11. The method of claim 9, wherein the solution is applied to the hands.
12. The method of claim 9, wherein the solution is applied to the feet.
13. The method of claim 9, wherein the solution is applied to the groin, face, back or abdomen.
14. The method of claims 8, 9, 10, 11, 12, or 13 where the administration is with an absorbent pad comprising an absorbed pharmaceutically acceptable solution of claim 1.
15. The method of claim 12 where the administration is with an absorbent pad comprising an absorbed pharmaceutically acceptable solution of claim 1.
16. A method of making threo glycopyrrolate tosylate comprising:
 - a) treating racemic cyclopentylmandelic acid with racemic 1-methylpyrrolidin-3-ol and 1,1'-carbonyldiimidazole in a suitable solvent to form glycopyrrolate base;
 - b) treating the glycopyrrolate base in a suitable solvent with a resolving acid to form a salt of threo glycopyrrolate;

- c) treating the salt of the threo glycopyrrolate salt with a suitable base in a suitable solvent to form a base of threo glycopyrrolate; and
 - d) treating the base of threo glycopyrrolate with *p*-toulenesulfonic acid in a suitable solvent to form threo glycopyrrolate tosylate.
17. The method of claim 16 further comprising the step of treating the threo glycopyrrolate tosylate in water to form Form D glycopyrrolate tosylate monohydrate.
18. The method of claim 16, wherein the resolving acid in step b) is 5-Nitroisophthalic acid.
19. The method of claim 16, wherein the resulting solution in a) is heated.
20. The method of claim 16, wherein the suitable solvent in step a) comprises an organic solvent.
21. The method of claim 20, wherein the organic solvent is toluene.
22. The method of claim 16, wherein the suitable solvent in step b) comprises an alcohol.
23. The method of claim 22, wherein the alcohol is methanol.
24. The method of claim 16, wherein the suitable solvent in step c) comprises a solution comprising water and an organic solvent.
25. The method of claim 24, wherein the organic solvent is toluene.
26. The method of claim 16, wherein the suitable solvent in step d) comprises an organic solvent.
27. The method of claim 26, wherein the organic solvent is selected from acetone and ethyl acetate.
28. The method of claim 27, wherein the organic solvent is acetone.
29. The method of claim 17, wherein the treatment with water comprises crystallizing Form D glycopyrrolate tosylate monohydrate in water.
30. The method of claim 29, wherein the crystallization is performed with seed crystals of Form D glycopyrrolate tosylate monohydrate.
31. The method of claim 30 further comprising drying the crystallized Form D glycopyrrolate tosylate monohydrate.
32. The method of claims 20, 22, 24, or 26, wherein the organic solvent of step a) is toluene; the alcohol of step b) is methanol; the organic solvent of step c) is toluene; and the organic solvent of step d) is acetone.
33. Crystalline glycopyrrolate tosylate.

34. Crystalline glycopyrrolate tosylate monohydrate.
35. Form D crystalline glycopyrrolate tosylate monohydrate of claim 34 characterized by an x-ray powder diffraction pattern comprising a peak at about $6.9^{\circ}2\theta$.
36. Form D crystalline glycopyrrolate tosylate monohydrate of claim 34 characterized by an x-ray powder diffraction pattern comprising a peak at about $12.6^{\circ}2\theta$.
37. Form D crystalline glycopyrrolate tosylate monohydrate of claim 34 characterized by an x-ray powder diffraction pattern comprising one peak at $6.9^{\circ}2\theta$ or $12.6^{\circ}2\theta$ and further comprising one or more peaks at about 10.3, 13.7, 14.9, 15.3, 15.7, 16.4, 17.7, or $18.2^{\circ}2\theta$.
38. Form D crystalline glycopyrrolate tosylate monohydrate of claim 34 characterized by an x-ray powder diffraction pattern comprising one or more peaks at about 10.3, 13.7, 14.9, 15.3, 15.7, 16.4, 17.7, or $18.2^{\circ}2\theta$.
39. Form D crystalline glycopyrrolate tosylate monohydrate of claim 34 characterized by an infrared spectrum comprising one or more peaks at about 1734, 1196, 1125, 1036, 1013, or 682 cm^{-1} .
40. Form D crystalline glycopyrrolate tosylate monohydrate of claim 34 characterized by (i) an x-ray powder diffraction pattern comprising one or more peaks at about 6.9, 10.3, 12.6, 13.7, 14.9, 15.3, 15.7, 16.4, 17.7, or $18.2^{\circ}2\theta$; and (ii) an infrared spectrum comprising one or more peaks at about 1734, 1196, 1125, 1036, 1013, or 682 cm^{-1} .
41. Form D crystalline glycopyrrolate tosylate monohydrate of claim 2 characterized by substantially the same x-ray powder diffraction pattern as in FIG. 2.
42. Form D crystalline glycopyrrolate tosylate monohydrate of claim 34 characterized by substantially the same infrared spectrum as in FIG. 3.
43. Form C crystalline glycopyrrolate tosylate of claim 33.
44. Form C crystalline glycopyrrolate tosylate of claim 43 characterized by x-ray powder diffraction pattern comprising a peak at about $5.5^{\circ}2\theta$.
45. Form C crystalline glycopyrrolate tosylate of claim 43 characterized by an x-ray powder diffraction pattern comprising a peak at about $11.0^{\circ}2\theta$.
46. Form C crystalline glycopyrrolate tosylate of claim 43 comprising an x-ray powder diffraction pattern comprising two or more peaks at about 5.5, 11.0, 11.8, 13.9, 14.9, 17.8, 19.6, 20.4, 21.6 or $22.1^{\circ}2\theta$.

47. Form C crystalline glycopyrrolate tosylate of claim 43 characterized by an infrared spectrum comprising one or more peaks at about 1733, 1236, 1211, 1198, 1186, 1177, 1120, 1032, 1008, or 682 cm^{-1} .
48. Form C crystalline glycopyrrolate tosylate of claim 43 characterized by (i) an x-ray powder diffraction pattern comprising two or more peaks at about 5.5, 11.0, 11.8, 13.9, 14.9, 17.8, 19.6, 20.4, 21.6, or 22.1 $^{\circ}2\Theta$; and (ii) an infrared spectrum comprising one or more peaks at about 1733, 1236, 1211, 1198, 1186, 1177, 1120, 1032, 1008, or 682 cm^{-1} .
49. Form C crystalline glycopyrrolate tosylate of claim 43 characterized by substantially the same x-ray powder diffraction pattern as FIG. 4.
50. Form C crystalline glycopyrrolate tosylate of claim 43 characterized by substantially the same infrared spectrum as FIG. 5.
51. Crystalline dehydrated Form D glycopyrrolate tosylate.
52. The dehydrated Form D glycopyrrolate tosylate of claim 51 wherein the x-ray powder diffraction is substantially the same as that of Figure 8.
53. Form D crystalline glycopyrrolate tosylate monohydrate characterized by an x-ray powder diffraction pattern comprising one or more peaks at about 6.9 $^{\circ}$ 2Θ and about 12.6 $^{\circ}$ 2Θ , prepared by treating glycopyrrolate bromide with a metal salt in a suitable solvent to form a glycopyrrolate salt slurry; removing the solids in the slurry to form a solution; lyophilizing the solution to form a solid; dissolving the solid in a crystallization solvent; and removing the crystallization solvent to form Form D.
54. The crystalline glycopyrrolate tosylate monohydrate of claim 53 further characterized by an x-ray powder diffraction pattern comprising one or more peaks at about 10.3, 13.7, 14.9, 15.3, 15.7, 16.4, 17.7, or 18.2 $^{\circ}2\Theta$.
55. The crystalline glycopyrrolate tosylate monohydrate of claims 53 or 54, wherein the metal salt is a silver salt.
56. The crystalline glycopyrrolate tosylate monohydrate of claim 55 wherein the suitable solvent is an alcohol.
57. The crystalline glycopyrrolate tosylate monohydrate of claim 56, wherein the alcohol is isopropanol.
58. The crystalline glycopyrrolate tosylate monohydrate of claims 54, 55, 56, or 57 wherein the crystallization solvent comprises acetonitrile and water.

59. The crystalline glycopyrrolate tosylate monohydrate of claim 58, wherein the crystallization solvent is removed by lowering the temperature of the solid in the crystallization solvent and decanting the solvent.
60. The crystalline glycopyrrolate tosylate monohydrate of claim 59 further comprising the addition of an anti-solvent.
61. The crystalline glycopyrrolate tosylate monohydrate of claim 60, wherein the antisolvent is toluene.
62. Form D crystalline glycopyrrolate tosylate monohydrate characterized by an x-ray powder diffraction pattern comprising one or more peaks at about $6.9^{\circ} 2\theta$ and about $12.6^{\circ} 2\theta$, prepared by treating glycopyrrolate-Y, wherein Y is an organic ion, and *p*-toluenesulfonic acid in a suitable solvent; removing the solvent to form a solid; re-dissolving the solid in a crystallization solvent to form a solution; and removing the crystallization solvent to form Form D glycopyrrolate tosylate.
63. The crystalline glycopyrrolate tosylate monohydrate of claim 62 wherein Y is acetate.
64. The crystalline glycopyrrolate tosylate monohydrate of claim 63 wherein the suitable solvent is isopropanol and the suitable solvent comprises acetonitrile and water.
65. The crystalline glycopyrrolate tosylate monohydrate of claim 64 wherein the crystallization solvent is removed by evaporation.
66. Glycopyrrolate tosylate, co-crystals, polymorphs, hydrates and solvates thereof.
67. Solid glycopyrrolate tosylate.
68. X-ray amorphous glycopyrrolate tosylate.
69. The glycopyrrolate of claim 68 characterized by an x-ray powder diffraction pattern comprising substantially the same as that of Figure 19.
70. The glycopyrrolate tosylate of claims 68 or 69 characterized by a glass transition onset temperature of about 11.6°C .
71. A salt of glycopyrrolate selected from the benzoate, edisylate, oxalate, and hydrogen sulfate salts and hydrates and solvates thereof.
72. Glycopyrrolate benzoate of claim 71.
73. Glycopyrrolate benzoate of claim 72 in crystalline form.
74. Crystalline glycopyrrolate benzoate of claim 73 having a DSC endotherm at about 79°C .

75. Crystalline glycopyrrolate benzoate of claims 73 or 74 characterized by an x-ray powder diffraction pattern comprising substantially the same as that of Figure 12.
76. Crystalline glycopyrrolate benzoate of claims 73 or 74 characterized by an x-ray powder diffraction pattern comprising peaks at about 8.0 and 18.8 °2 Θ .
77. Glycopyrrolate edisylate of claim 71.
78. The glycopyrrolate edisylate of claim 77 wherein the molar ratio of glycopyrrolate to edisylate is 2:1.
79. Glycopyrrolate edisylate of claim 78 in crystalline form.
80. Glycopyrrolate edisylate of claim 79 having a DSC endotherm at about 103°C.
81. Glycopyrrolate edisylate of claim 79 or 80 characterized by an x-ray powder diffraction pattern comprising substantially the same as that of Figure 14.
82. Crystalline glycopyrrolate edisylate of claims 79 or 80 characterized by an x-ray powder diffraction pattern comprising peaks at about 11.2 and 17.9 °2 Θ .
83. Glycopyrrolate oxalate of claim 71.
84. Crystalline glycopyrrolate oxalate of claim 83.
85. Crystalline glycopyrrolate oxalate of claim 84 characterized by an x-ray powder diffraction pattern comprising substantially the same as that of Figure 16.
86. Crystalline glycopyrrolate oxalate of claim 84 characterized by an x-ray powder diffraction pattern comprising peaks at about 5.0 and 8.4 °2 Θ .
87. Glycopyrrolate hydrogen sulfate of claim 71.
88. Crystalline glycopyrrolate hydrogen sulfate of claim 87.
89. Crystalline glycopyrrolate hydrogen sulfate of claim 88 having a DSC endotherm at about 160°C.
90. The crystalline glycopyrrolate hydrogen sulfate of claim 89 further comprising a DSC endotherm at about 169°C.
91. Crystalline glycopyrrolate hydrogen sulfate of claims 88, 89, or 90 characterized by an x-ray powder diffraction pattern comprising substantially the same as that of Figure 15.
92. Crystalline glycopyrrolate hydrogen sulfate of claims 88, 89, or 90 characterized by an x-ray powder diffraction pattern comprising peaks at about 5.6 and 13.1 °2 Θ .
93. Crystalline glycopyrrolate acetate.

94. Crystalline glycopyrrolate acetate of claim 93 characterized by an x-ray powder diffraction pattern comprising peaks at about 5.2 and 11.3 °2 θ .
95. Crystalline glycopyrrolate acetate of claim 93 characterized by an x-ray powder diffraction pattern comprising substantially the same as that of Figure 22.
96. An absorbent pad containing an aqueous solution comprising glycopyrrolate tosylate absorbed onto the pad.
97. The pad of claim 96, wherein the solution further comprises one or more buffers.
98. The pad of claim 97, wherein the solution further comprises an alcohol.
99. The pad of claim 98, wherein the one or more buffers are citric acid and sodium citrate and the alcohol is ethanol.
100. The pad of claim 99, wherein the ethanol is dehydrated ethanol.
101. The pad of claim 100 further comprising a binding agent.
102. The pad of claim 101, wherein the binding agent is povidone.
103. The pad of claim 102, wherein the binding agent is povidone K90.
104. The pad of claims 99, 100, 101, 102, or 103 further comprising a film-forming agent.
105. The pad of claim 104, wherein the film forming agent is a butyl ester of a polyvinylmethylether/maleic anhydride acid copolymer.
106. The pad of claim 96, wherein the pad comprises polypropylene.
107. The pad of claim 106, wherein the pad is substantially polypropylene.
108. The pad of claim 107, wherein the pad is 100% polypropylene.
109. The pad of claim 96 wherein, the pH of the solution is between about 3.5 and 5.5.
110. A gel comprising a pharmaceutically acceptable amount of glycopyrrolate tosylate.
111. The pad of claims 96, 99 or 100, wherein the glycopyrrolate tosylate is a racemic mixture of (R)-3-((S)-2-cyclopentyl-2-hydroxy-2-phenylacetoxy)-1,1-dimethylpyrrolidinium 4-methylbenzenesulfonate and (S)-3-((R)-2-cyclopentyl-2-hydroxy-2-phenylacetoxy)-1,1-dimethylpyrrolidinium 4-methylbenzenesulfonate.
112. The pad of claims 96 or 100, wherein the amount of glycopyrrolate tosylate solution is about 2.8 g

113. An absorbent pad containing an aqueous glycopyrrolate solution comprising about 0.15% by weight anhydrous citric acid, about 0.06% sodium citrate dihydrate by weight, between about 57 to about 59.5% by weight of dehydrated ethanol, and between about 1% and about 6% by weight glycopyrrolate tosylate.
114. The pad of claim 113, wherein the pad comprises polypropylene.
115. The pad of claim 114, wherein the pad is substantially polypropylene.
116. The pad of claim 105, wherein the pad is 100% polypropylene.
117. A process for making an absorbent pad containing an aqueous solution of glycopyrrolate tosylate absorbed onto the pad comprising treating solid glycopyrrolate tosylate in solution with water, one or more buffers and an alcohol, and mixing until dissolution; and wetting a pad with the glycopyrrolate tosylate solution so that a pharmaceutically acceptable amount of glycopyrrolate tosylate has been absorbed onto the pad.
118. The process of claim 117 wherein the pad is in a pouch prior to wetting.
119. The process of claim 118 further comprising heat-sealing the pouch after wetting.
120. The process of claim 117 wherein the one or more buffers are sodium citrate and citric acid.
121. The process of claims 117 or 120 wherein the alcohol is dehydrated ethanol.
122. The process of claims 117 or 120 wherein the pharmaceutically acceptable amount of glycopyrrolate tosylate solution is about 2.8g.
123. The process of claim 121 wherein the pharmaceutically acceptable amount of glycopyrrolate tosylate solution is about 2.8 g.
124. The process of claims 117 or 123 wherein the pad comprises polypropylene.
125. The process of claim 124 wherein the pad is substantially polypropylene.
126. The process of claim 123 wherein the pad is 100% polypropylene.
127. The process of claim 122 wherein the pad comprises polypropylene.
128. The process of claim 127 wherein the pad is 100% polypropylene.
129. An absorbent pad containing a pharmaceutically acceptable aqueous solution of glycopyrrolate tosylate made by the process of claims 117, 118, 119, 120, 126, 127, or 128.

130. An absorbent pad containing a pharmaceutically acceptable aqueous solution of glycopyrrolate tosylate made by the process of claim 122.
131. The process of claim 117, wherein the glycopyrrolate tosylate is a racemic mixture of (R)-3-((S)-2-cyclopentyl-2-hydroxy-2-phenylacetoxy)-1,1-dimethylpyrrolidinium 4-methylbenzenesulfonate and (S)-3-((R)-2-cyclopentyl-2-hydroxy-2-phenylacetoxy)-1,1-dimethylpyrrolidinium 4-methylbenzenesulfonate.
132. The process of claim 123, wherein the glycopyrrolate tosylate is a racemic mixture of (R)-3-((S)-2-cyclopentyl-2-hydroxy-2-phenylacetoxy)-1,1-dimethylpyrrolidinium 4-methylbenzenesulfonate and (S)-3-((R)-2-cyclopentyl-2-hydroxy-2-phenylacetoxy)-1,1-dimethylpyrrolidinium 4-methylbenzenesulfonate.
133. The pad of claim 129, wherein the glycopyrrolate tosylate is a racemic mixture of (R)-3-((S)-2-cyclopentyl-2-hydroxy-2-phenylacetoxy)-1,1-dimethylpyrrolidinium 4-methylbenzenesulfonate and (S)-3-((R)-2-cyclopentyl-2-hydroxy-2-phenylacetoxy)-1,1-dimethylpyrrolidinium 4-methylbenzenesulfonate.
134. The pad of claim 130, wherein the glycopyrrolate tosylate is a racemic mixture of (R)-3-((S)-2-cyclopentyl-2-hydroxy-2-phenylacetoxy)-1,1-dimethylpyrrolidinium 4-methylbenzenesulfonate and (S)-3-((R)-2-cyclopentyl-2-hydroxy-2-phenylacetoxy)-1,1-dimethylpyrrolidinium 4-methylbenzenesulfonate.
135. The process of claim 117, wherein the absorbent pad contains an aqueous glycopyrrolate solution comprising about 0.15% by weight anhydrous citric acid, about 0.06% sodium citrate dihydrate by weight, between about 57 to about 59.5% by weight of dehydrated ethanol, and between about 1% and about 6% by weight glycopyrrolate tosylate.
136. The process of claim 135, wherein the pad comprises polypropylene.
137. The process of claim 136, wherein the pad is substantially polypropylene.
138. The process of claim 137, wherein the pad is 100% polypropylene.
139. A method of treating hyperhidrosis comprising topically administering a therapeutically effective amount of an aqueous glycopyrrolate tosylate solution from an absorbent pad to the skin of a mammal.
140. The method of claim 139 wherein the mammal is human.
141. The method of claim 140, wherein the solution is applied to the hands or axillae.

142. The method of claim 140, wherein the solution is applied to the feet.
143. The method of claim 140, wherein the solution is applied to the groin, face, back or abdomen.
144. The method of claims 139, 140, 141, 142 or 143, wherein the pad comprises polypropylene.
145. The method of claim 144, wherein the pad is substantially polypropylene.
146. The method of claim 145, wherein the pad is 100% polypropylene.
147. The method of claim 145, wherein the solution comprises one or more buffers and ethanol.
148. The method of claim 144, further comprising one or more of a binding agent and a film-forming agent.
149. The method of claim 148, wherein the binding agent is povidone and the film-forming agent is a butyl ester of a polyvinylmethylether/maleic anhydride acid copolymer.
150. The method of claim 144, wherein the glycopyrrolate tosylate is a racemic mixture of (R)-3-((S)-2-cyclopentyl-2-hydroxy-2-phenylacetoxy)-1,1-dimethylpyrrolidinium 4-methylbenzenesulfonate and (S)-3-((R)-2-cyclopentyl-2-hydroxy-2-phenylacetoxy)-1,1-dimethylpyrrolidinium 4-methylbenzenesulfonate.
151. The method of claim 139, wherein the amount of glycopyrrolate tosylate solution is about 2.8 g.
152. The method of claim 139, wherein the absorbent pad contains an aqueous glycopyrrolate solution comprising about 0.15% by weight anhydrous citric acid, about 0.06% sodium citrate dihydrate by weight, between about 57 to about 59.5% by weight of dehydrated ethanol, and between about 1% and about 6% by weight glycopyrrolate tosylate.
153. The method of claim 152, wherein the pad comprises polypropylene.
154. The method of claim 153, wherein the pad is substantially polypropylene.
155. The method of claim 154, wherein the pad is 100% polypropylene.
156. A solid dispersion comprising glycopyrrolate tosylate.

157. The solid dispersion of claim 156 further comprising a monosaccharide, disaccharide, or a pharmaceutically acceptable polymer containing a cyclic ether moiety.
158. The solid dispersion of claim 156 further comprising a disaccharide.
159. The solid dispersion of claim 158 wherein the disaccharide is sucrose.
160. The solid dispersion of claim 156 further comprising a pharmaceutically acceptable polymer containing a cyclic ether moiety.
161. The solid dispersion of claim 160 wherein the cyclic ether moiety is a six-membered ring.
162. The solid dispersion of claim 161 wherein the polymer is HPMCAS.
163. The solid dispersion of claim 162 wherein the glass transition temperature is greater than about 40°C and the weight ratio of HPMCAS to glycopyrrolate tosylate is about 1 to 1.
164. The solid dispersion of claim 163 wherein the glass transition temperature is about 42°C.
165. The solid dispersion of claims 162, 163, or 164 wherein the infrared spectrum of said dispersion has a peak at about 1228 cm⁻¹.
166. The solid dispersion of claim 165 made by a process comprising cooling a solution comprising glycopyrrolate tosylate and HPMCAS until the solution has frozen; and drying the frozen solution to form a solid dispersion.
167. A process for preparing a solid dispersion of claim 165 comprising cooling a solution comprising glycopyrrolate tosylate and HPMCAS until the solution has frozen; and drying the frozen solution to form a solid dispersion.
168. The solid dispersion of claim 159 wherein the glass temperature is greater than about 60°C and the weight ratio of sucrose to glycopyrrolate tosylate is about 9 to 1.
169. The solid dispersion of claim 168 wherein the glass temperature is about 62°C.
170. The solid dispersion of claims 159 or 169 made by a process comprising cooling a solution comprising glycopyrrolate tosylate and sucrose until the solution has frozen; and drying the frozen solution to form a solid dispersion.
171. A process for preparing a solid dispersion of claims 159 or 169 comprising cooling a solution comprising glycopyrrolate tosylate and sucrose until the solution has frozen; and drying the frozen solution to form a solid dispersion.

172. The solid dispersion of claim 156 further comprising a pharmaceutically acceptable polymer containing a polyethylene glycol moiety.
173. The solid dispersion of claim 172 wherein the polymer is a polyvinyl alcohol – polyethylene glycol graft copolymer or a polyvinyl caprolactam-polyvinyl acetate-polyethylene glycol graft copolymer.
174. The solid dispersion of claim 156 further comprising a polyvinyl alcohol – polyethylene glycol copolymer.
175. The solid dispersion of claim 174, wherein the copolymer is Kollicoat® IR and the weight ratio of Kollicoat® IR to glycopyrrolate tosylate is about 1:1.
176. The solid dispersion of claim 175, wherein the infrared spectrum of said dispersion has at least one peak selected from 1099 cm^{-1} and 1324 cm^{-1} .
177. The solid dispersion of claims 175 or 176 wherein the glass transition temperature of the dispersion is about 32°C .
178. The solid dispersion of claim 174 wherein the copolymer is Kollicoat® IR and the weight ratio of Kollicoat® IR to glycopyrrolate tosylate is about 9:1.
179. The solid dispersion of claim 178, wherein the dispersion has a glass transition temperature of about 35°C .
180. The solid dispersion of claims 174, 175 or 179 prepared by the process comprising cooling a solution comprising glycopyrrolate tosylate and the polyvinyl alcohol – polyethylene glycol copolymer until the solution has frozen; and drying the frozen solution to form a solid dispersion.
181. A process for preparing a solid dispersion of claims 174, 175, or 179 prepared by the process comprising cooling a solution comprising glycopyrrolate tosylate and the polyvinyl alcohol – polyethylene glycol copolymer until the solution has frozen; and drying the frozen solution to form a solid dispersion.
182. The solid dispersion of claim 156 further comprising a polyvinyl caprolactam-polyvinyl acetate-polyethylene glycol graft copolymer.
183. The solid dispersion of claim 182 wherein the copolymer is Soluplus® and the weight ration of copolymer to glycopyrrolate tosylate is about 1:1.
184. The solid dispersion of claims 182 or 183 wherein the glass transition temperature is greater than about 30°C .

185. The solid dispersion of claim 184 wherein the glass transition temperature is greater than about 40°C.
186. The solid dispersion of claims 182, 183, 184 or 185 wherein the infrared spectrum of the solid dispersion comprises one or more peaks at about 942cm⁻¹, about 1195 cm⁻¹, and about 1445 cm⁻¹.
187. The solid dispersion of claims 182 or 183 prepared by the process comprising cooling a solution comprising glycopyrrolate tosylate and the polyvinyl caprolactam-polyvinyl acetate-polyethylene glycol graft copolymer until the solution has frozen; and drying the frozen solution to form a solid dispersion.
188. A process for preparing a solid dispersion of claims 182 or 183 prepared by the process comprising cooling a solution comprising glycopyrrolate tosylate and the polyvinyl caprolactam-polyvinyl acetate-polyethylene glycol graft copolymer until the solution has frozen; and drying the frozen solution to form a solid dispersion.
189. The solid dispersion of claim 156 further comprising a pharmaceutically acceptable polymer containing a vinyl pyrrolidone moiety.
190. The solid dispersion of claim 156 further comprising a pharmaceutically acceptable polymer selected from polyvinyl pyrrolidone or a vinyl pyrrolidone-vinyl acetate copolymer.
191. The solid dispersion of claim 156 further comprising polyvinyl pyrrolidone having a glass transition temperature is at least about 25°C wherein the weight ratio of polyvinyl pyrrolidone to glycopyrrolate tosylate is between about 1:1 and about 8:1.
192. The solid dispersion of claim 191 wherein the glass transition temperature is greater than about 35°C.
193. The solid dispersion of claim 192 wherein the polymer is selected from PVP K29/32 and PVP K90.
194. The solid dispersion of claim 193, wherein the polymer is PVP K29/32, the weight ratio of PVP K29/32 to glycopyrrolate tosylate is about 1:1, and the infrared spectrum of the dispersion comprises one or more peaks at about 1288 cm⁻¹, about 1461 cm⁻¹ and about 1664 cm⁻¹.
195. The solid dispersion of claim 194, wherein the glass transition temperature of about 38°C.

196. The solid dispersion of claim 193, wherein the polymer is PVP K29/32, the weight ratio of PVP K29/32 to glycopyrrolate tosylate is about 8:1 and the glass transition temperature is about 26°C.
197. The solid dispersion of claim 193 wherein the polymer is PVP K90 the weight ratio of PVP K90 to glycopyrrolate tosylate is about 1:1, and the infrared spectrum of the dispersion comprises a peak at about 1664 cm⁻¹.
198. The solid dispersion of claim 197, wherein the glass transition temperature is about 36°C.
199. A solid dispersion of claims 191 or 193 comprising cooling a solution comprising glycopyrrolate tosylate and the polyvinyl pyrrolidone polymer until the solution has frozen; and drying the frozen solution to form a solid dispersion.
200. A process for preparing a solid dispersion of claims 191 or 193 comprising cooling a solution comprising glycopyrrolate tosylate and the polyvinyl pyrrolidone polymer until the solution has frozen; and drying the frozen solution to form a solid dispersion.
201. The solid dispersion of claim 156 further comprising a vinyl pyrrolidone – vinyl acetate copolymer.
202. The solid dispersion of claim 201 wherein the glass transition temperature is greater than about 60°C and the weight ratio of copolymer to glycopyrrolate tosylate is about 1:1.
203. The solid dispersion of claims 201 wherein the copolymer is Kollidon® VA 64.
204. The solid dispersion of claims 201, 202, or 203 wherein the infrared spectrum of the solid dispersion comprises one or more peaks selected from about 1287 cm⁻¹, about 1371 cm⁻¹ or about 1673 cm⁻¹.
205. The solid dispersion of claim 204, wherein the glass transition temperature is about 64°C.
206. The solid dispersion of claims 201, 202, or 203, wherein the glass transition temperature is about 64°C.
207. The solid dispersion of claims 201 or 203 prepared by a process comprising cooling a solution comprising glycopyrrolate tosylate and the vinyl pyrrolidone – vinyl

acetate copolymer until the solution has frozen; and drying the frozen solution to form a solid dispersion.

208. A process for preparing a solid dispersion of claims 201 or 203 comprising cooling a solution comprising glycopyrrolate tosylate and the vinyl pyrrolidone – vinyl acetate copolymer until the solution has frozen; and drying the frozen solution to form a solid dispersion.
209. A method for treating primary axillary hyperhidrosis with glycopyrrolate tosylate or a solvate thereof comprising administering topically administering a therapeutically effective amount of an aqueous glycopyrrolate tosylate solution to the skin of a mammal in need thereof.
210. A method for treating palmar or plantar hyperhidrosis with glycopyrrolate tosylate or a solvate thereof comprising administering topically administering a therapeutically effective amount of an aqueous glycopyrrolate tosylate solution to the skin of a mammal in need thereof.
211. Glycopyrrolate tosylate.
212. A topical comprising three glycopyrrolate tosylate.
213. The topical of claim 212 further comprising one or more buffers.
214. The topical of claim 212 or 213 wherein the topical is an aqueous solution.
215. A pharmaceutically effective amount of the topical of claim 214.
216. The topical of claim 213, wherein the one or more buffers are citric acid and sodium citrate and further comprising an alcohol.
217. The topical of claim 216, wherein the alcohol is ethanol.
218. The topical of claim 217 further comprising a binding agent.
219. The topical of claim 218, wherein the binding agent is povidone.
220. The topical of claim 219, wherein the binding agent is povidone K90.
221. The topical of claims 216, 217, 218, 219, or 220 further comprising a film-forming agent.
222. The topical of claim 221, wherein the film forming agent is a butyl ester of a polyvinylmethylether/maleic anhydride acid copolymer..
223. The topical of claim 214 wherein, the pH of the solution is between about 3.5 and 5.5.

224. A cream or ointment comprising a pharmaceutically acceptable amount of glycopyrrolate tosylate.
225. The topical of claims 212, 213, 215, wherein the glycopyrrolate tosylate is a racemic mixture of (R)-3-((S)-2-cyclopentyl-2-hydroxy-2-phenylacetoxy)-1,1-dimethylpyrrolidinium 4-methylbenzenesulfonate and (S)-3-((R)-2-cyclopentyl-2-hydroxy-2-phenylacetoxy)-1,1-dimethylpyrrolidinium 4-methylbenzenesulfonate.
226. The topical of claims 212, 213, or 215, wherein the amount of glycopyrrolate tosylate solution is about 2.8 g.
227. A process for making an aqueous formulation comprising threo glycopyrrolate tosylate comprising treating solid glycopyrrolate tosylate in solution with water, a buffer and an alcohol, and mixing until dissolution.
228. The process of claim 227 wherein the pH of the solution is between about 4.0 and 5.0 and the alcohol is ethanol.
229. A solution of threo glycopyrrolate tosylate prepared by the processes of claims 227 or 228.
230. The process of claims 227 or 228 wherein the solid glycopyrrolate tosylate is Form D.
231. The process of claims 227 or 228 wherein the solid glycopyrrolate tosylate is selected from Form C, a solid dispersion of glycopyrrolate tosylate, x-ray amorphous glycopyrrolate tosylate, and dehydrated Form D glycopyrrolate tosylate.
232. A solution of threo glycopyrrolate tosylate prepared by the processes of claim 230.
233. A solution of threo glycopyrrolate tosylate prepared by the processes of claim 231.
234. The process of claims 117-120 wherein the glycopyrrolate tosylate is crystalline.
235. The process of claims 227-228 wherein the glycopyrrolate tosylate is crystalline.
236. A solution of threo glycopyrrolate tosylate prepared by the process of claim 234.
237. A solution of threo glycopyrrolate tosylate prepared by the process of claim 235.
238. A process for making an aqueous formulation comprising glycopyrrolate tosylate comprising dissolving crystalline glycopyrrolate tosylate in solution comprising water.
239. A solution of glycopyrrolate tosylate prepared by the process of claim 238.