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Improved catalyst for the preparation of polytetrahydrofuran

10 The present invention relates to a catalyst which comprises a catalytically active
amount of at least one oxygen-containing molybdenum and/or tungsten compound
on an oxidic support material and has been calcined at from 400°C to 900°C after
application of the precursor compounds of the catalytically active compounds to
the support material or a support material precursor, and to a process for preparing
it. This catalyst is used in processes for preparing polytetrahydrofuran,
15 tetrahydrofuran copolymers and diesters or monoesters of these polymers.

Polytetrahydrofuran, hereinafter referred to as PTHF for short, which is also
known as polyoxybutylene glycol, is used as a versatile intermediate in the plastics
and synthetic fibers industries and is employed, inter alia, for producing
20 polyurethane, polyester and polyamide elastomers. In addition, it is, possibly in the
form of its derivatives, a valuable auxiliary in a variety of applications, e.g. as
dispersant or in the deinking of waste paper.

Industrially, PTHF is advantageously prepared by polymerization of
25 tetrahydrofuran over suitable catalysts. Additional reagents whose addition makes
it possible to control the chain length of the polymer chains and thus to set the
mean molecular weight to the desired value are used in this process. Such chain
termination reagents are also known as "telogens". The chain length is controlled
via selection of type and amount of telogen. Choice of suitable telogens
30 additionally makes it possible to introduce functional groups at one end or at both
ends of the polymer chain. For example, use of carboxylic acids or carboxylic
anhydrides as telogens makes it possible to prepare the monoesters or diesters of
PTHF. Other telogens act not only as chain termination reagents but are also
incorporated into the growing polymer chain of PTHF and have the function of
35 both a telogen and a comonomer. Examples of such comonomers are telogens
having two hydroxy groups, for example dialcohols. Examples of such dialcohols

- are ethylene glycol, propylene glycol, butylene glycol, 1,4-butanediol, 2-butyne-1,4-diol, 1,6-hexanediol or low molecular weight PTHF. The use of such comonomers leads to tetrahydrofuran copolymers. In this way, the PTHF can also be chemically modified. An example is the use of the telogen 2-butyne-1,4-diol
- 5 whose addition leads to the presence of a proportion of C-C triple bonds in the polymer chains of the PTHF. PTHF which has been modified in this way can be further modified chemically at these points due to the reactivity of these triple bonds, e.g. by hydrogenation of the triple bonds to form double bonds, by subsequent grafting of other monomers to adjust the properties of the polymer,
- 10 crosslinking to form polymers having a comparatively rigid structure or other customary measures of polymer chemistry. Complete hydrogenation of the triple bonds present is likewise possible and generally leads to PTHF having a particularly low color number.
- 15 Industrially, the preparation of PTHF is predominantly carried out by two-stage processes in which tetrahydrofuran is, for example, polymerized in the presence of fluorosulfonic acid to form PTHF esters which are subsequently hydrolyzed to give PTHF. In a further two-stage process, tetrahydrofuran is polymerized with acetic anhydride in the presence of acid catalysts to form PTHF diacetate which is
- 20 subsequently transesterified with methanol to give PTHF. Apart from the fact that a two-stage process has to be carried out, a particularly disadvantageous factor is that by-products such as hydrofluoric acid or methyl acetate are formed.
- 25 Single-stage processes for preparing PTHF in which tetrahydrofuran is polymerized using diols, water or low molecular weight PTHF as telogen over acid catalysts are also known. The catalyst or the catalyst system can be homogeneous, i.e. present in solution in the reaction system, or it is possible to use heterogeneous, largely undissolved catalyst systems.
- 30 DE-A-44 33 606 describes a process for preparing PTHF, PTHF diesters of C_2 - C_{20} -monocarboxylic acids or PTHF monoesters of C_1 - C_{10} -monocarboxylic acids by polymerization of tetrahydrofuran over a heterogeneous catalyst in the presence of one of the telogens water, 1,4-butanediol, PTHF having a molecular
- 35 weight of from 200 to 700 dalton, a C_1 - C_{10} -monocarboxylic acid or a carboxylic anhydride of a C_2 - C_{20} -monocarboxylic acid or a mixture of these telogens, where

the catalyst is a supported catalyst comprising a catalytically active amount of an oxygen-containing tungsten or molybdenum compound or a mixture of these compounds on an oxidic support material and has been calcined at from 500°C to 1000°C after application of precursor compounds of the oxygen-containing molybdenum and/or tungsten compounds to the support material precursor.

DE-A-196 41 481 discloses a process for preparing polytetrahydrofuran, copolymers of tetrahydrofuran in 2-butyne-1,4-diol, diesters of these polymers with C₂-C₂₀-monocarboxylic acids or monoesters of these polymers with C₁-C₁₀-monocarboxylic acids by polymerization of tetrahydrofuran in the presence of one of the telogens water, 1,4-butanediol, 2-butyne-1,4-diol, polytetrahydrofuran having a molecular weight of from 200 to 700 dalton, a C₁-C₁₀-monocarboxylic acid or a carboxylic anhydride of a C₂-C₂₀-monocarboxylic acid or a mixture of these telogens over a heterogeneous support catalyst. The supported catalyst comprises a catalytically active amount of an oxygen-containing molybdenum and/or tungsten compound on an oxidic support material and has been calcined at from 500°C to 1000°C after application of precursor compounds of the oxygen-containing molybdenum and/or tungsten compounds to the support material precursor and activated by treatment with a reducing agent before being used as polymerization catalyst.

It is an object of the present invention to provide a catalyst which has a higher activity than the known catalysts so as to achieve higher polymer yields and/or space-time yields.

We have found that this object is achieved by a catalyst which comprises a catalytically active amount of at least one oxygen-containing molybdenum and/or tungsten compound on an oxidic support material and has been calcined at from 400°C to 900°C after application of precursor compounds of the catalytically active compounds to the support material or a support material precursor, wherein the catalyst has a porosity comprising transport pores having a diameter of > 25 nm and a volume of these transport pores of at least 50 mm³/g.

In a further embodiment, the catalyst of the present invention has a molybdenum and/or tungsten content, based on the finished catalyst, of x μmol of (tungsten and/or molybdenum)/m², where 10.1 < x < 20.9. This targeted setting of the ratio

of the tungsten and/or molybdenum content to the BET surface area enables a further significant increase in the catalyst activity to be obtained. This targeted setting is achieved by matching the calcination conditions to the tungsten and/or molybdenum content of the catalyst.

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In a preferred embodiment, the catalyst of the present invention has a BET surface area of $25\text{-}200\text{ m}^2/\text{g}$, preferably $40\text{-}150\text{ m}^2/\text{g}$ and particularly preferably $50\text{-}120\text{ m}^2/\text{g}$.

- 10 The oxidic material to which the catalytically active amount of the molybdenum and/or tungsten compound is applied has a BET surface area of $135\text{-}220\text{ m}^2/\text{g}$.

It has been found that catalysts whose oxidic support material has a BET surface area in the abovementioned range are particularly acidic and active. Laboratory
15 tests have revealed that both the acidity and the activity of the catalysts decreases when the BET surface area is below the abovementioned range. Astonishingly, the acidity and the activity of the catalysts also decrease again when the BET surface area of the oxidic support material used is above the range claimed here.

- 20 In a preferred embodiment of the catalyst of the present invention, the BET surface area of the acidic support material is $170\text{-}220\text{ m}^2/\text{g}$, particularly preferably $190\text{-}220\text{ m}^2/\text{g}$.

The catalyst of the present invention can have an acidity at pK_a of < -3 of at least
25 $10\text{ }\mu\text{mol/g}$. When such an oxidic support material whose BET surface area is in the abovementioned ranges is used, catalysts whose acidity at $\text{pK}_a < -3$ is at least $70\text{ }\mu\text{mol/g}$, preferably at least $90\text{ }\mu\text{mol/g}$, in the dry state are obtained. Here, the pK_a is determined by the Hammett titration method.

- 30 Suitable catalytically active compounds are molybdenum or tungsten compounds which, without addition of a support, are mainly present as MoO_3 or WO_3 after prolonged heating at 600°C in moist air, or mixtures of these compounds. Examples of such compounds are all types of tungstic and molybdic acids, e.g. WO_3 , $\text{WO}_3\cdot\text{H}_2\text{O}$, $\text{WO}_3\cdot 2\text{H}_2\text{O}$, metatungstic and paratungstic acids and the
35 isopolyacids and also the corresponding analogous molybdenum compounds. In this context, mention may also be made of heteropolyacids such as $\text{H}_3\text{PW}_{12}\text{O}_{40}$,

H₄SiW₁₂O₄₀, H₄GeW₁₂O₄₀ and so forth to H₆P₂W₁₈O₅₆ or the corresponding molybdenum compounds, and the salts of the abovementioned acids, e.g. ammonium orthotungstate, ammonium metatungstate and ammonium paratungstate, the corresponding molybdenum compounds, and reduced tungsten and molybdenum compounds and also organomolybdenum or organotungsten compounds.

Preference is given to the use of oxygen-containing tungsten compounds or of mixtures of oxygen-containing tungsten compounds with oxygen-containing molybdenum compounds in which the tungsten compounds predominate. Particular preference is given to using catalysts whose catalytically active compounds, apart from the usual impurities, consist virtually only of oxygen-containing tungsten compounds. Particular mention may here be made of WO₃, tungstic acid and ammonium paratungstate.

Suitable support materials are oxidic supports such as zirconium dioxide, titanium dioxide, hafnium oxide, yttrium oxide, iron(III) oxide, aluminum oxide, tin(IV) oxide or mixtures of these oxides. Preference is given to using zirconium dioxide or titanium dioxide, with titanium dioxide being particularly preferred and titanium dioxide having a proportion of > 65% in the anatase modification being particularly suitable. The support material can be used in freshly precipitated, dried or calcined form. The respective support material can be obtained by hydrolysis of appropriate halides, nitrates, sulfates or alkoxides, e.g. starting from titanium tetraisopropoxide, titanyl chloride, titanyl nitrate, titanium tetrachloride or titanyl sulfate.

For the purposes of the present invention, the figures given for the proportion of catalytically active compounds in the catalyst is based on the trioxides of tungsten and/or molybdenum. The catalysts generally contain from 0.1 to 70% by weight, preferably from 5 to 40% by weight and particularly preferably from 10 to 35% by weight, of the catalytically active compounds, in each case calculated as MoO₃ and/or WO₃ and based on the total weight of the finished, calcined and dry catalyst.

After application of the oxygen-containing molybdenum and/or tungsten compound, the BET surface area of the finished catalyst does not exceed a maximum specific surface area of

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$$\text{BET}[\text{cat, max}] = \text{BET}[\text{X}] / (1 - \text{loss on ignition}[\text{X}]) \cdot (1 - \text{Y}).$$

Here, X refers to the oxidic support material and Y to the oxygen-containing molybdenum and/or tungsten compound, each in g. In this calculation, it is assumed that both H₂O and the oxygen-containing molybdenum and/or tungsten
10 compound which has been used occupy the surface of the oxidic support material.

In addition, it is possible to use pore formers such as tartaric acid, oxalic acid, citric acid, ammonium nitrate, ammonium formate, ammonium oxalate, guanidinium salts, urotropin, proteins such as gelatin, polyvinyl alcohol,
15 polyethylene oxide, polymethylene oxide, carbohydrates such as glucose, sucrose and soluble starch, PTHF, surfactants and sulfonic acids. Preference is given to tartaric acid, citric acid and oxalic acid, especially oxalic acid.

With regard to the proportion of pore former present in the catalyst, the content of
20 pore former is, for the purposes of the present invention, given in % by weight based on the finished, calcined and dry catalyst. The sum of the percentages by weight of the compounds present in the catalyst therefore exceeds 100%. According to the present invention, the catalysts contain up to 200% by weight, preferably up to 100% by weight and particularly preferably up to 60% by weight,
25 of pore formers.

In addition, the catalyst may further comprise a sulfur and/or phosphorus compound. Since it is assumed that the sulfur in the catalysts of the present invention is present as sulfate and the phosphorus is present as phosphate, the
30 proportion of these compounds in the catalyst is calculated as sulfate or phosphate, based on the total weight of the finished, calcined and dry catalyst. In this way, the catalysts can contain up to 15% by weight, preferably up to 10% by weight and particularly preferably up to 8% by weight, of sulfate and/or phosphate.

35 Sources of the corresponding phosphorus and/or sulfur compounds are, for example, sulfuric acid, soluble sulfates such as ammonium sulfate or ammonium

hydrogensulfate or the corresponding sulfites or hydrosulfites, phosphoric acid, soluble phosphates such as ammonium phosphate or the ammonium hydrogen-phosphates or the corresponding phosphites.

- 5 As further additive, a binder such as silica sol, clays, boric acid or AlPO_4 may also be present in the catalysts.

10 The present invention also provides a process for preparing a catalyst comprising a catalytically active amount of at least one oxygen-containing molybdenum and/or tungsten compound as an oxidic support material, in which the following steps are carried out:

- precipitation of the oxidic support material,
- thermal treatment of the oxidic support material to set a BET surface area of $135\text{-}220\text{ m}^2/\text{g}$,
- 15 - application of the catalytically active compound(s) or its/their precursor compound(s) to the support material and
- calcination of the resulting catalyst at from 400°C to 900°C to give the finished catalyst.

20 It has been found that catalysts having the specified BET surface area can be prepared by the process of the present invention when freshly precipitated oxidic support material, e.g. preferably $\text{TiO}_2\cdot\text{aq}$, is not used directly as oxidic support material for the catalysts to be prepared, but is firstly subjected to thermal treatment. Freshly precipitated oxidic support material has a high specific surface area. It has been found that such a high specific surface area causes a drop in the acidity and activity of the finished catalyst. If the surface area of the oxidic support material is set in a targeted manner to BET values of $135\text{-}220\text{ m}^2/\text{g}$, preferably $170\text{-}220\text{ m}^2/\text{g}$ and particularly preferably $190\text{-}220\text{ m}^2/\text{g}$, by thermal treatment, e.g. by calcination at from about 300°C to about 550°C , it is possible to prepare catalysts which have a particularly high activity and an acidity at $\text{pK}_a < -3$ of at least $70\text{ }\mu\text{mol/g}$, preferably at least $90\text{ }\mu\text{mol/g}$, in the dry state. A further advantage is that a uniform support material attains a defined BET surface area. Support materials used are zirconium dioxide or titanium dioxide, preferably titanium dioxide. Since the process of the present invention is preferably employed for preparing the abovementioned catalyst of the present invention, what has been said above applies to the properties of the catalyst prepared in this way, e.g. in respect

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of the proportion of catalytically active compounds, any pore formers added, added sulfate and/or phosphate.

5 In a preferred embodiment of the preparative process of the present invention, it is possible to prepare catalysts which have a particularly good mechanical stability. This mechanical stability can be achieved by carrying out the thermal treatment of the freshly precipitated $\text{TiO}_2 \cdot \text{aq}$ for the targeted setting of a desired BET surface area under a water-containing or steam-containing atmosphere. In this case, the water-containing or steam-containing atmosphere should comprise more than 10%
10 by volume, preferably more than 20% by volume and particularly preferably more than 30% by volume, of water or steam. The process can also be carried out at a pressure other than atmospheric pressure.

When carrying out the process of the present invention, preference is given to
15 applying the catalytically active compound(s) in a molar ratio of support material to catalytically active compound of 5.0-7.5, preferably 6.0-7.0.

Subsequently, the calcination of the catalyst is preferably carried out stepwise, with the catalyst firstly being calcined at 100-600°C and then being subjected to final calcination at up to 900°C, preferably up to 700°C, to give the finished catalyst. In
20 the case of calcination with a stepwise increase in the temperature, it is possible, for example, to carry out calcination firstly at a low initial temperature, e.g. 200°C, then to increase the temperature to, for example, 500°C and subsequently to carry out a final calcination with a further increase in the temperature.

25 The present invention also provides a process for preparing polytetrahydrofuran, tetrahydrofuran copolymers, diesters or monoesters of these polymers by polymerization of tetrahydrofuran in the presence of at least one telogen and/or comonomer using the abovementioned catalyst in one of its inventive embodiments, with the catalyst having been obtained, for example, by the
30 preparative process likewise claimed in one of its inventive embodiments.

Particular mention may be made of the preparation of polytetrahydrofuran and tetrahydrofuran copolymers and of polytetrahydrofurandicarboxylate, and also, in the preparation of polytetrahydrofuran, the preparation from tetrahydrofuran and
35 1,4-butanediol or H_2O as telogen or using short-chain PTHF.

Polytetrahydrofurandicarboxylate is prepared from THF and carboxylic anhydride, preferably acetic anhydride.

The invention is illustrated by the examples below.

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I. Measurement and determination methods used

I.1. Measurement of the BET surface area

- 10 The BET surface areas were determined by adsorption of N₂ in accordance with DIN 66131.

I.2. Measurement of the tungsten content

- 15 200 mg of the homogenized sample are mixed with about 25 g of borax/lithium fluoride as flux and fused in a platinum crucible at 1000°C. The fused mass is analyzed by X-ray fluorescence. Evaluation is carried out by means of external calibration using matrix-batched reference standards.

I.3. Measurement of the loss on ignition of the TiO₂ support

The loss on ignition was calculated as follows from the initial weight (m₂₀) and the weight after ignition at 900°C in air for five hours (m₉₀₀):

- 25 Loss on ignition [TiO₂] = 1 – m₉₀₀/m₂₀

I.4. Measurement of the Hammett acidity

- 30 The number of acid centers having a pK_a of < -3 ("acidity") was determined by titration with pyrazine as base in absolute toluene as solvent using dicinnamalacetone (CAS 622-21-9) as indicator.

- The finely milled catalyst is dried overnight at 200°C under reduced pressure (<1 mbar) and in the absence of light. About 50 ml of absolute toluene and 1 mg of dicinnamalacetone (dissolved in 1 ml of absolute toluene) are then added under
35 argon. The mixture is subsequently titrated while stirring with a 0.1% strength by weight solution of pyrazine in absolute toluene, with the mixture being allowed to

equilibrate for at least 5 hours after each addition of pyrazine. The end point of the titration can be recognized by a color change from blue to red, which becomes more distinct if another 1 mg of dicinnamalacetone is added shortly before the end point of the titration. In general terms, the titration was carried out as indicated, for example, in H. Benesi and B. Winquist, *Advances in Catalysis*, Vol. 27, 1978, Academic Press, ISBN 0-12-007827-9, and the references cited therein.

I.5. Measurement of the cutting hardness

The measurement of the cutting hardness was carried out as follows:
A blade having a thickness of 0.3 mm is pressed with increasing force onto each of 25 randomly chosen, visually crack-free extrudates having a mean length of 2-3 times the diameter until the respective extrudate has been cut through. The force required for this is the cutting hardness in N.

I.6. Measurement of the porosity

The porosity of the catalyst was determined by the Hg intrusion method in accordance with DIN 66133.

I.7. Setting of the BET surface area of the TiO₂ support

Precipitated TiO₂*aq having a BET surface area of 275 m²/g is calcined in air in a continuously operated rotary tube at from 350 to 500°C and a residence time of about 1 hour. This gives a uniform material having a defined BET surface area in the range from 135 to 220 m²/g.

II. Starting compounds

II.1. Titanium oxides and hydroxides

Possible titanium oxides and hydroxides which are suitable as support materials for preparing the catalysts of the present invention and were used as desired are ones which are commercially available under the name S 150, S 140 and S 240 from Kemira, under the name P 25 from Degussa, as VKR 611, Hombikat UV100 and

Hombifine N from Sachtleben, XT 25376 from Norton, DT 51 from Thann et Mulhouse and Bayoxide T A-DW-1, Bayoxide T A-K-1 from Bayer.

II.2. Zirconium oxides and hydroxides

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Zirconium oxides and hydroxides which may be mentioned as suitable for preparing the catalysts of the present invention are S97/718 from MEL, D9-89 from BASF, XZ 16052 from Norton and HC 15 from SEPR.

10 II.3. Tungsten and molybdenum compounds

As tungsten or molybdenum compounds, use is made of WO_3 , MoO_3 , H_2WO_4 and ammonium metatungstate and paratungstate from H.C. Starck.

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III. General method of preparing the catalyst

III.1a. Impregnation of the support with the catalytically active compound

20 To prepare the catalyst, the support is impregnated with the appropriate molybdenum and/or tungsten compound in water. If desired, a soluble pore former, as indicated above by way of example, can also be added to this impregnation solution. In addition, a binder in the form of a powder or in the form of a suspension may be added if desired.

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III.1b. Mixing the support with the catalytically active compound

30 As an alternative to this impregnation method, the support can also be mixed with the molybdenum and/or tungsten compound and water. This can be carried out, for example, in a pan mill or a kneader. In addition, a pore former which is soluble or insoluble in water and a binder in the form of a powder or a suspension may be added. This method is preferred over impregnation as described under 1a.

III.2. Spraying the mix obtained

An intermediate step which can be carried out if desired is spraying the mix obtained in this way, drying and precalcining the sprayed material at from about 80 to 600°C, preferably from 100 to 500°C, in an oxidizing, inert or reducing atmosphere which may in each case be either static or flowing. As oxidizing atmosphere, it is possible to use air or the like; as inert atmosphere, it is possible to use N₂, Ar or CO₂; as reducing atmosphere, it is possible to use organic compounds in inert gas, CO or H₂.

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III.3. Pan milling or kneading the material

As the next step in the preparation, the catalyst is pan milled or kneaded, with in each case a pore former which is soluble or insoluble in water and a binder being again able to be added if desired.

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III.4. Shaping the catalyst

The catalyst is subsequently shaped by, for example, extrusion or tableting to give shaped bodies such as cylinders, spheres, rings, spirals, rods or granules.

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III.5. Precalcination

If desired, a precalcination at from about 80 to 600°C, preferably from 100 to 500°C, under an oxidizing, inert or reducing atmosphere which may in each case be either static or flowing can now be carried out.

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III.6. Final calcination

The final calcination is subsequently carried out at from 400 to 900°C, preferably from 550 to 850°C and particularly preferably from 600 to 800°C, again under an oxidizing, inert or reducing atmosphere which may in each case be either static or flowing and using a gas as described by way of example under 2. Calcination at from 600 to 800°C produces a catalyst which gives a particularly high space-time yield.

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Exhausted catalysts which have been regenerated by customary methods, e.g. by treatment with solvents, steam or calcination at from 300°C to 600°C for from 0.1 to 10 hours in air, CO₂ or a reducing atmosphere such as H₂ or CO, can also be used. Furthermore, it is possible to reuse catalyst powders, shaped bodies and/or
5 granulated catalyst by addition to the above-described preparation steps.

If desired, this final calcination can be followed by targeted reduction, as has been described in DE-A-196 41 481.

10 The preparation of the catalyst is essentially carried out by methods analogous to those described in DE-A-44 33 606 or DE-A-196 49 803, which are hereby expressly incorporated by reference.

15 **IV. Preparation of PTHF**

Since the activities of each of the catalysts according to the present invention were tested in the preparation of PTHF, this preparative process will not be discussed at this point but the basic process variants used will be presented.

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IV.1. General

It is in principle possible to use any THF as monomer for preparing PTHF. However, preference is given to using commercial THF which has been
25 prepurified by acid treatment as described, for example, in EP-A 003 112 or by distillation.

Suitable telogens in the single-stage process are saturated or unsaturated, unbranched or branched alpha,omega-C₂-C₁₂-diols, water, polytetrahydrofuran
30 having a molecular weight of from 200 to 700 dalton, cyclic ethers or mixtures thereof. Preference is given to water, ethylene glycol, 1,3-propanediol, 1,4-butanediol, 1,5-pentanediol, 1,6-hexanediol, polytetrahydrofuran having a molecular weight of from 200 to 700 dalton, 1,6-hexanediol, 1,8-octanediol, 1,10-decanediol, 2-butyne-1,4-diol and neopentyl glycol or mixtures thereof, with particular
35 preference being given to water, 1,4-butanediol and/or polytetrahydrofuran having a molecular weight of from 200 to 700 dalton.

Telogens used for preparing PTHF and THF copolymers by the two-stage process are carboxylic anhydrides and/or carboxylic anhydride/protic acid mixtures. Preference is given to carboxylic anhydrides which are derived from aliphatic or aromatic polycarboxylic and/or monocarboxylic acids and contain from 2 to 12, preferably from 1 to 8, carbon atoms. Particular preference is given to acetic anhydride. The protic acids are preferably organic and inorganic acids which are soluble in the reaction system. Preferred carboxylic acids are aliphatic or aromatic polycarboxylic and/or monocarboxylic acids which contain from 2 to 12, preferably from 1 to 8, carbon atoms. Examples of aliphatic carboxylic acids are acetic acid, lactic acid, propionic acid, valeric acid, caproic acid, caprylic acid and pelargonic acid. Examples of aromatic carboxylic acids are phthalic acid and naphthalenecarboxylic acid. Among these carboxylic acids, preference is given to using acetic acid.

Suitable comonomers are cyclic ethers which can be polymerized in a ring-opening polymerization, preferably three-, four- and five-membered rings such as the THF derivatives 2-methyltetrahydrofuran or 3-methyltetrahydrofuran, cyclic ethers such as 1,3-dioxolane, trioxane, oxetane, substituted oxetanes such as 3,3-dimethyloxetane, 1,2-alkylene oxides, for example ethylene oxide or propylene oxide, with particular preference being given to 2-methyltetrahydrofuran or 3-methyltetrahydrofuran.

The telogen is advantageously introduced into the polymerization as a solution in THF; preference is given to a telogen content of from 0.04 to 17 mol%, based on tetrahydrofuran. Comonomers are likewise advantageously introduced into the polymerization as a solution in THF; the comonomer content can be up to 30 mol%, preferably 20 mol%, based on tetrahydrofuran. However, it is also possible to introduce THF and the telogen and/or the comonomer separately into the polymerization reactor. Since the telogen effects termination of the polymerization, the mean molecular weight of the PTHF or THF copolymer can be controlled via the amount of telogen used. The more telogen present in the reaction mixture, the lower the mean molecular weight of the PTHF or the THF copolymers concerned. Depending on the telogen content of the polymerization mixture, it is possible to prepare PTHF and THF copolymers having mean molecular weights of

from 650 to 5000 dalton, preferably from 650 to 3000 dalton and particularly preferably from 1000 to 3000 dalton.

5 The polymerization is generally carried out at from 0 to 80°C, preferably from 25 to 75°C and particularly preferably from 40 to 70°C. The pressure employed is generally not critical to the result of the polymerization, which is why the polymerization is generally carried out at atmospheric pressure or under the autogenous pressure of the polymerization system.

10 To avoid formation of ether peroxides, the polymerization is advantageously brought about under an inert gas atmosphere. Inert gases which can be used are, for example, nitrogen, carbon dioxide or the noble gases; preference is given to using nitrogen.

15 The polymerization can be carried out in the presence of hydrogen at hydrogen pressures of from 0.1 to 10 bar.

20 The process of the present invention is preferably operated continuously. However, the polymerization step and/or one or more or all of the work-up steps of the process of the present invention can also be operated batchwise, but in this case it is preferred for at least the polymerization to be carried out continuously.

25 The reaction is preferably carried out in conventional reactors or reactor assemblies suitable for continuous processes in a suspension or fixed-bed mode, in the case of a suspension mode, for example, in loop reactors or stirred reactors or in the case of a fixed-bed mode in tube reactors or fixed-bed reactors. The fixed-bed mode is preferred.

30 In the fixed-bed mode, the polymerization reactor can be operated in the upflow mode, i.e. the reaction mixture is conveyed from the bottom upward, or in the downflow mode, i.e. the reaction mixture is passed through the reactor from the top downward. The starting material mixture (feed) of THF and telogen and/or comonomer is fed continuously to the polymerization reactor and the WHSV is from 0.05 to 0.8 kg of THF/(l h), preferably from 0.1 to 0.6 kg of THF/(l h) and particularly preferably from 0.15 to 0.5 kg of THF/(l h).
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Furthermore, the polymerization reactor can be operated in the single-pass mode, i.e. without product recirculation, or in the recycle mode, i.e. the polymerization mixture leaving the reactor is circulated. In the recycle mode, the ratio of recycle to fresh feed is $\leq 100:1$, preferably $< 40:1$ and particularly preferably $< 30:1$.

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IV.2. Continuous THF polymerization: Variant A

150 g of catalyst were installed in a fixed-bed reactor without recirculation and were subsequently dried at 200°C in a stream of N₂ for 24 hours. The reactor was then cooled to 60°C and started up at a throughput of 45 g/h of a mixture comprising 4000 ppm of 1,4-butanediol in THF (tetrahydrofuran). The output was collected daily and freed of THF and butanediol by heating under reduced pressure (0.3 mbar, 120°C). The number average molecular weight of the PTHF obtained in this way was then measured.

The butanediol content of the feed was subsequently varied until the mean molecular weight of the PTHF M_n was 2000 ± 200 on three successive days. The THF conversion was averaged over these three days.

IV.3. Continuous THF polymerization: Variant B

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200 g of catalyst were installed in a fixed-bed reactor without recirculation and were subsequently dried at 200°C in a stream of N₂ for 24 hours. The reactor was then cooled to 60°C and started up at a throughput of 50 g/h of a mixture comprising 4000 ppm of 1,4-butanediol in THF. The output was collected daily and freed of THF and butanediol by heating under reduced pressure (0.3 mbar, 120°C). The number average molecular weight of the PTHF obtained in this way was then measured. The butanediol content of the feed was subsequently varied until the mean molecular weight of the PTHF was 2000 ± 200 on three successive days. The THF conversion was averaged over these three days.

30

IV.4. Batchwise THF polymerization

The batchwise polymerization experiments were carried out under an argon atmosphere in 40 ml GC sample bottles. 5 g of catalyst extrudates which had been dried at 400°C in a stream of O₂/N₂ ("synthetic air") for 18 hours to remove adsorbed water before use were heated at 60°C in 20.75 g of butanediol-containing

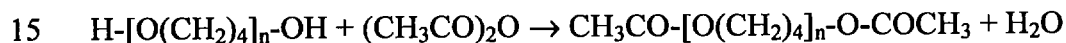
35

THF (butanediol concentration: 2000 ppm) under Ar for 6 hours while shaking. The reaction vessel was subsequently cooled and the catalyst was filtered off and washed three times with 10 g each time of THF. The filtrates were combined, evaporated at 70°C/20 mbar on a rotary evaporator and weighed.

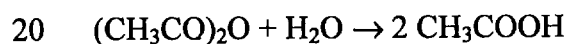
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IV.5. Measurement of the molecular weight of the PTHF

The molecular weight is determined by measurement of the hydroxyl number of polytetrahydrofuran. For the purposes of the present invention, the hydroxyl number is the amount of potassium hydroxide in mg which is equivalent to the amount of acetic acid bound in the acetylation of 1 g of a substance. The hydroxyl number is determined by esterification of the hydroxyl groups present with an excess of acetic anhydride.



After the reaction, the excess acetic anhydride is hydrolyzed with water according to the following reaction equation



and back-titrated as acetic acid using sodium hydroxide.

Catalysts – Examples

25

I. Examples of the setting of a transport pore size of > 25 nm and a pore volume of at least 50 mm³/g in the finished catalysts

Comparative Example 1

30

130 kg of TiO₂*aq (= 100 kg of TiO₂) and 26.7 kg of tungstic acid (= 25 kg of WO₃) were compounded in a pan mill with 1.52 kg of H₃PO₄ (85%) and 46.7 kg of H₂O, and the resulting paste was extruded and shaped to form extrudates having a diameter of 4 mm. These were dried at 170°C and subsequently calcined at 690°C for 2 hours.

35

Example 1

119 kg of $\text{TiO}_2 \cdot \text{aq}$ (= 95.6 kg of TiO_2) and 25.6 kg of tungstic acid (= 23.9 kg of WO_3) were compounded in a pan mill with 1.46 kg of H_3PO_4 (85%) and 48 kg of H_2O , and the resulting paste was extruded and shaped to form extrudates having a diameter of 4 mm. These were dried at 170°C and subsequently calcined at 690°C for 2 hours.

Example 2

150.2 g of $\text{TiO}_2 \cdot \text{aq}$ (= 115.7 g of TiO_2) and 40.46 g of tungstic acid (= 37.63 g of WO_3) were kneaded with 94 g of stearic acid, 9.4 g of tartaric acid, 2.2 g of H_3PO_4 (85%) and 92 g of H_2O in a laboratory kneader, and the resulting paste was shaped to form extrudates having a diameter of 4 mm. These were dried at 400°C in a stream of N_2 for 4 hours and subsequently calcined at 650°C in air for 3 hours.

Example 3

200 g of $\text{TiO}_2 \cdot \text{aq}$ (= 160 g of TiO_2) and 43 g of tungstic acid (= 40 g of WO_3) were kneaded with 48.6 g of stearic acid, 12.2 g of tartaric acid, 38.6 g of aqueous silica sol (15.8% SiO_2) and 40 g of H_2O in a laboratory kneader, and the resulting paste was shaped to form extrudates having a diameter of 4 mm. These were subsequently dried at 350°C in air for 2 hours and calcined at 700°C in air for 3 hours.

Example 4

128.7 g of $\text{TiO}_2 \cdot \text{aq}$ (= 99 g of TiO_2), 35.4 g of tungstic acid (= 33 g of WO_3), 65.9 g of oxalic acid, 6.9 g of methylcellulose and 70 g of H_2O were kneaded in a laboratory kneader, and the resulting paste was shaped to form extrudates having a diameter of 4 mm. These were subsequently calcined at 650°C in air for 2 hours.

Example 4a

100 g of $\text{TiO}_2 \cdot \text{aq}$ (= 80 g of TiO_2), 31.3 g of ammonium paratungstate (= 28.2 g of WO_3), 91.9 g of oxalic acid, 4 g of polyvinyl alcohol, 4 g of polyethylene oxide

and 80 g of H₂O were kneaded in a laboratory kneader, and the resulting paste was shaped to form extrudates having a diameter of 4 mm. These were subsequently calcined at 650°C in air for 1.5 hours.

5 **Example 5**

143.2 g of TiO₂*aq (= 110 g of TiO₂), 29.6 g of tungstic acid (= 27.5 g of WO₃), 68.7 g of stearic acid, 6.9 g of tartaric acid, 60.8 g of guanidinium carbonate, 1.6 g of H₃PO₄ (85%) and 80 g of H₂O were kneaded in a laboratory kneader, and the
10 resulting paste was shaped to form extrudates having a diameter of 4 mm. These were dried at 400°C in a stream of N₂ for 4 hours and subsequently calcined at 650°C in air for 3 hours.

Example 6

15

217.5 g of TiO₂*aq (= 167.6 g of TiO₂), 44.9 g of tungstic acid (= 41.8 g of WO₃), 104.4 g of ammonium citrate, 10.4 g of tartaric acid, 2.5 g of H₃PO₄ (85%) and 41.5 g of 25% strength NH₃ solution were kneaded in a laboratory kneader, and the
20 resulting paste was shaped to form extrudates having a diameter of 4 mm. These were calcined at 650°C in air for 3 hours.

Example 7

143 g of TiO₂*aq (= 110 g of TiO₂), 29.6 g of tungstic acid (= 27.5 g of WO₃),
25 68.7 g of guanidinium carbonate, 35 g of formic acid, 6.9 g of tartaric acid, 1.6 g of H₃PO₄ (85%) and 15 g of H₂O were kneaded in a laboratory kneader, and the resulting paste was shaped to form extrudates having a diameter of 4 mm. These were calcined at 650°C in air for 3 hours.

30 **Example 8**

150 g of TiO₂*aq (= 115.7 g of TiO₂) and 40.46 g of tungstic acid (= 37.63 g of WO₃) were kneaded with 94 g of gelatin, 9.4 g of tartaric acid, 2.2 g of H₃PO₄ (85%) and 117 g of H₂O in a laboratory kneader, and the resulting paste was
35 shaped to form extrudates having a diameter of 4 mm. These were subsequently calcined at 700°C in air for 3 hours.

Example 9

150.2 g of $\text{TiO}_2 \cdot \text{aq}$ (= 115.7 g of TiO_2) and 40.46 g of tungstic acid (= 37.63 g of WO_3) were kneaded with 94 g of pure, soluble starch, 9.4 g of tartaric acid, 2.2 g of H_3PO_4 (85%) and 140 g of H_2O in a laboratory kneader, and the resulting paste was shaped to form extrudates having a diameter of 4 mm. These were subsequently calcined at 700°C in air for 3 hours.

Example 10

150.2 g of $\text{TiO}_2 \cdot \text{aq}$ (= 115.7 g of TiO_2) and 40.46 g of tungstic acid (= 37.63 g of WO_3) were kneaded with 93.9 g of cold-milled paraffin, 9.4 g of tartaric acid, 2.2 g of H_3PO_4 (85%) and 100 g of H_2O in a laboratory kneader, and the resulting paste was shaped to form extrudates having a diameter of 4 mm. These were subsequently dried at 400°C in a stream of N_2 for 4 hours and calcined at 700°C in air for 3 hours.

Example 11

160 g of $\text{TiO}_2 \cdot \text{aq}$ (= 128 g of TiO_2) and 34.4 g of tungstic acid (= 32 g of WO_3) were kneaded with 77.8 g of carbon black (Printex 75, from Degussa), 9.8 g of tartaric acid, 2.3 g of H_3PO_4 (85%), 20.4 g of alkali-free Lutensol and 120 g of H_2O in a laboratory kneader, and the resulting paste was shaped to form extrudates having a diameter of 4 mm. These were subsequently dried at 400°C in a stream of N_2 for 4 hours and calcined at 600°C in air for 8 hours.

Example 12

200 g of $\text{TiO}_2 \cdot \text{aq}$ (= 160 g of TiO_2) and 41.4 g of $\text{H}_3\text{PW}_{12}\text{O}_{40} \cdot \text{aq}$ (= 33.2 g of WO_3) were stirred with 520 g of H_2O in a glass beaker and evaporated at 60°C under reduced pressure. The material was subsequently heated to 500°C in a stream of N_2 and calcined at this temperature in air for 2 hours. The product was finely milled in an analytical mill, kneaded with 120.8 g of H_2O in a laboratory kneader, and the resulting paste was shaped to form extrudates having a diameter of 4 mm. These were subsequently calcined at 700°C in air for 3 hours.

Example 13

375 g of $\text{TiO}_2 \cdot \text{aq}$ (= 300 g of TiO_2) and 103.4 g of $\text{H}_3\text{PW}_{12}\text{O}_{40} \cdot \text{aq}$ (= 82.9 g of WO_3) were stirred with 1040 g of H_2O in a glass beaker and evaporated at 60°C under reduced pressure. The material was subsequently heated to 500°C in a stream of N_2 and calcined at this temperature in air for 2 hours. The product was milled in a spiral jet mill at 8 bar to give particles having a mean size of $0.8 \mu\text{m}$, kneaded with 120.8 g of H_2O in a laboratory kneader, and the resulting paste was shaped to form extrudates having a diameter of 4 mm. These were subsequently calcined at 650°C in air for 1 hour.

Example 14

8.38 kg of $\text{TiO}_2 \cdot \text{aq}$ (= 6.44 kg of TiO_2) and 2.00 kg of $\text{H}_3\text{PW}_{12}\text{O}_{40} \cdot \text{aq}$ (= 1.61 kg of WO_3) were slurried with 138 kg of H_2O and sprayed in a spray dryer at 120°C to give a powder having a mean particle size of $2.7 \mu\text{m}$, and this powder was subsequently calcined at 500°C for 2 hours. 200 g of this spray-dried powder were kneaded with 10 g of methylcellulose and 120 g of H_2O in a laboratory kneader, and the resulting paste was shaped to form extrudates having a diameter of 4 mm. These were subsequently calcined at 700°C in air for 3 hours.

Example 15

7.86 kg of $\text{TiO}_2 \cdot \text{aq}$ (= 6.04 kg of TiO_2) and 2.50 kg of $\text{H}_3\text{PW}_{12}\text{O}_{40} \cdot \text{aq}$ (= 2.01 kg of WO_3) were slurried with 138 kg of H_2O and sprayed in a spray dryer at 120°C to give a powder having a mean particle size of $2.6 \mu\text{m}$, and this powder was subsequently calcined at 500°C for 2 hours. 200 g of this spray-dried powder were kneaded with 10 g of methylcellulose and 120 g of H_2O in a laboratory kneader, and the resulting paste was shaped to form extrudates having a diameter of 4 mm. These were subsequently calcined at 650°C in air for 3 hours.

Example 16

7.86 kg of $\text{TiO}_2 \cdot \text{aq}$ (= 6.04 kg of TiO_2) and 2.50 kg of $\text{H}_3\text{PW}_{12}\text{O}_{40} \cdot \text{aq}$ (= 2.01 kg of WO_3) were slurried with 138 kg of H_2O and sprayed in a spray dryer at 120°C to

give a powder having a mean particle size of 2.6 μm , and this powder was subsequently calcined at 500°C for 2 hours. 200 g of this spray-dried powder were kneaded with 6 g of polyethylene oxide and 116 g of H_2O in a laboratory kneader, and the resulting paste was shaped to form extrudates having a diameter of 4 mm.

5 These were subsequently calcined at 650°C in air for 2 hours.

Comparison of the catalysts from Comparative Example 1 and Examples 1 to 16 when used in the polymerization of THF is shown in Table 1. In this comparison, THF was polymerized batchwise, as described above under IV.4.

10

The experiments show that the use of catalysts having a high proportion of transport pores of >25 nm in the process of the present invention in all cases leads to an increase in activity. When using the process of the present invention, the capacity of existing plants for the preparation of PTHF can therefore be increased significantly without plant modification or future plants can be made correspondingly smaller and therefore be built at a correspondingly reduced capital cost.

15

Table 1: Comparison of the catalysts from the examples

Method of obtaining the porosity	Cat.	WO ₃ content of the batch [% by weight]	BET [m ² /g]	Acidity at PK _a < -3 [μmol/g]	Volume of pores > 25 nm [cm ³ /g]	STY [g _(pTHF) /kg _(cat) *h]
“Dry”-shaped	CE1	20	79	28	0.020	18
“Wet”-shaped	E1	20	76	36	0.146	33
Pore former	E2	24.5	93	45	0.754	104
Stearic acid	E3	20	71	25	0.311	69
Oxalic acid	E4	25	74	27	0.376	86
Oxalic acid	E4a	26	73	48	0.365	140
Carbonates	E5	20	65	37	0.779	79
Ammonium salts	E6	20	71	19	0.153	51
Guanidinium salts	E7	20	63	22	0.054	26
Proteins	E8	24.5	91	37	0.624	84
Carbohydrates	E9	24.5	107	42	0.106	66
Paraffin	E10	24.5	73	37	0.440	92
Carbon black	E11	25	75	30	0.477	70
Milling + shaping	E12	18	59	22	0.438	54
	E13	21	67	19	0.248	65
Spraying + shaping	E14	20	65	19	0.481	41
	E15	25	79	30	0.435	45
	E16	25	78	34	0.397	69

II. Examples of the setting of a tungsten content of from > 10.1 to < 20.9 $\mu\text{mol}/\text{m}^2$

Comparative Example II.1

5 100 g of TiO_2^*aq (= 78 g of TiO_2), 15.5 g of ammonium paratungstate (= 13.8 g of WO_3) and 80.9 g of anhydrous oxalic acid were compounded in a laboratory kneader with 75 g of H_2O , and the resulting mass was extruded and shaped to form extrudates having a diameter of 4 mm. These were calcined at 200°C for 2 hours,
10 at 500°C for 3 hours and at 600°C for 3 hours.

Comparative Example II.2

15 184.7 g of TiO_2^*aq (= 144 g of TiO_2), 29 g of tungstic acid (= 27 g of WO_3) and 67 g of H_2O were compounded in a laboratory kneader, and the resulting mass was extruded and shaped to form extrudates having a diameter of 4 mm. These were dried at 105°C and subsequently calcined at 775°C for 24 hours.

Comparative Example II.3

20 123.1 kg of TiO_2^*aq (= 96 g of TiO_2), 25.7 kg of tungstic acid (= 24 kg of WO_3) and 6.09 kg of tartaric acid were compounded in a pan mill with 2.95 kg of aqueous H_3PO_4 (= 1.764 kg of P_2O_5) and 45 kg of H_2O , and the resulting mass was extruded and shaped to form extrudates having a diameter of 4 mm. These were
25 dried at 105°C and subsequently calcined at 690°C for 2 hours.

Comparative Example II.4

30 180 g of TiO_2^*aq (= 144.4 g of TiO_2), 38.7 g of tungstic acid (= 36.1 g of WO_3) and 9.02 g of tartaric acid were compounded in a laboratory kneader with 2.21 g of aqueous H_3PO_4 (= 1.32 g of P_2O_5) and 73 ml of H_2O , and the resulting mass was extruded and shaped to form extrudates having a diameter of 4 mm. These were dried at 105°C and subsequently calcined at 690°C for 48 hours.

Comparative Example II.5

180 g of $\text{TiO}_2 \cdot \text{aq}$ (= 140 g of TiO_2), 69.9 g of ammonium paratungstate (= 62.2 g of WO_3) and 37.5 g of anhydrous oxalic acid were compounded in a laboratory kneader with 53 g of H_2O , and the resulting mass was extruded and shaped to form extrudates having a diameter of 4 mm. These were dried at 35°C and subsequently calcined at 200°C for 2 hours and at 500°C for 3 hours.

Comparative Example II.6

125 g of $\text{TiO}_2 \cdot \text{aq}$ (= 97.5 g of TiO_2), 48.5 g of ammonium paratungstate (= 43.2 g of WO_3) and 43.4 g of anhydrous oxalic acid were compounded in a laboratory kneader with 56 g of H_2O , and the resulting mass was extruded and shaped to form extrudates having a diameter of 4 mm. These were dried at 40°C and subsequently calcined at 200°C for 2 hours, at 500°C for 7 hours and at 670°C for 3 hours.

Comparative Example II.7

66.6 kg of $\text{TiO}_2 \cdot \text{aq}$ (= 51.5 kg of TiO_2), 31.1 kg of ammonium paratungstate (= 27.7 g of WO_3) were compounded in a pan mill with 20.4 kg of oxalic acid dihydrate and 15 kg of H_2O , and the resulting mass was extruded and shaped to form extrudates having a diameter of 4 mm. These were dried at 90°C and subsequently calcined at 200°C for 2 hours, at 500°C for 3 hours and at 510°C for 3 hours.

Comparative Example II.8

The catalyst from Example 7 was additionally calcined at 690°C for 12 hours.

Example II.1

184.7 g of $\text{TiO}_2 \cdot \text{aq}$ (= 144 g of TiO_2), 29 g of tungstic acid (= 27 g of WO_3) and 67 g of H_2O were compounded in a laboratory kneader, and the resulting mass was extruded and shaped to form extrudates having a diameter of 4 mm. These were dried at 105°C and subsequently calcined at 675°C for 16 hours.

Example II.2

119.1 kg of $\text{TiO}_2\cdot\text{aq}$ (= 95.6 kg of TiO_2), 25.6 kg of tungstic acid (= 23.9 kg of WO_3) and 5.97 kg of tartaric acid were compounded in a pan mill with 1.46 kg of aqueous H_3PO_4 (= 0.872 kg of P_2O_5) and 48 kg of H_2O , and the resulting mass was extruded and shaped to form extrudates having a diameter of 4 mm. These were dried at 105°C and subsequently calcined at 690°C for 2 hours.

Example II.3

119.1 kg of $\text{TiO}_2\cdot\text{aq}$ (= 95.6 kg of TiO_2), 25.6 kg of tungstic acid (= 23.9 kg of WO_3) and 5.97 kg of tartaric acid were compounded in a pan mill with 1.46 kg of aqueous H_3PO_4 (= 0.872 kg of P_2O_5) and 48 kg of H_2O , and the resulting mass was extruded and shaped to form extrudates having a diameter of 4 mm. These were dried at 120°C and subsequently calcined at 400°C for 2 hours and at 750°C for 3 hours.

Example II.4

180 g of $\text{TiO}_2\cdot\text{aq}$ (= 140 g of TiO_2), 71.7 g of ammonium paratungstate (= 63.8 g of WO_3) and 37.8 g of anhydrous oxalic acid were compounded in a laboratory kneader with 48 g of H_2O , and the resulting mass was extruded and shaped to form extrudates having a diameter of 4 mm. These were dried at 90°C for 2 hours and subsequently calcined at 200°C for 2 hours, at 500°C for 5 hours and at 650°C for 3 hours.

Example II.5

180 g of $\text{TiO}_2\cdot\text{aq}$ (= 140 g of TiO_2), 71.7 g of ammonium paratungstate (= 63.8 g of WO_3) and 37.8 g of anhydrous oxalic acid were compounded in a laboratory kneader with 48 g of H_2O , and the resulting mass was extruded and shaped to form extrudates having a diameter of 4 mm. These were dried at 35°C and subsequently calcined at 200°C for 2 hours, at 500°C for 5 hours and at 620°C for 3 hours.

Example II.6

5 66.6 kg of $\text{TiO}_2 \cdot \text{aq}$ (= 51.5 kg of TiO_2) and 31.1 kg of ammonium paratungstate (= 27.7 kg of WO_3) were compounded in a pan mill with 20.4 kg of oxalic acid dihydrate and 15 kg of H_2O , and the resulting mass was extruded and shaped to form extrudates having a diameter of 4 mm. These were dried at 90°C and subsequently calcined at 200°C for 2 hours, at 500°C for 3 hours and at 610°C for 3 hours.

10

The catalysts prepared in this way were used in the continuous THF polymerization described as variant B under IV.3. above. The results achieved are listed in Table 2, whose last column reports the THF conversion in percent.

Table 2: Comparison of the catalysts from the examples

Catalyst	W[g/100 g]	BET	W[$\mu\text{mol}/\text{m}^2$]	Conversion
Comparative Example II.1	11.5	75	8.3	3.2%
Example II.1	11.2	54	11.3	4.4%
Comparative Example II.2	11.4	29.5	21.0	2.2%
Comparative Example II.3	15.2	86.4	9.6	3.0%
Example II.2	15.8	75.5	11.2	4.7%
Example II.3	15.6	47.2	18.0	4.1%
Comparative Example II.4	15.9	39.6	21.8	3.4%
Comparative Example II.5	24.0	172	7.6	3.6%
Example II.4	24.7	95	14.1	5.4%
Example II.5	24.3	84.5	15.6	5.7%
Comparative Example II.6	23.9	59.7	21.8	4.6%
Comparative Example II.7	27.8	151	10.0	4.1%
Example II.6	27.2	109	13.6	5.7%
Comparative Example II.8	27.8	68	22.2	4.2%

III. Examples of the setting of a defined BET surface area of the oxidic support material

Example III.1: Preparation of the catalyst

5

Variant 1

720 g of precalcined TiO_2 *aq having, in each case, the BET surface area indicated in Table 1, ammonium paratungstate in an amount corresponding to a molar ratio of Ti to W of 6.77 and 156.4 g of anhydrous oxalic acid were compounded in a laboratory kneader with an appropriate amount of H_2O , and the resulting mass was extruded and shaped to form extrudates having a diameter of 4 mm. The amount of water required depends on the type of kneader used and on the TiO_2 material; it has no influence on activity and acidity. It is typically from 240 to 340 g. The kneaded composition was dried overnight at 35°C and subsequently calcined stepwise at 15 200°C for 2 hours and at 500°C for another 3 hours in a stream of air. It was then finally calcined for one hour at the temperature indicated in Table 3 to give the finished catalyst. Table 3 reports the loss on ignition of the TiO_2 *aq determined in each case, the temperatures used for final calcination to give the finished catalyst, the maximum BET surface area and the BET surface area of the catalysts obtained 20 in this way together with the calculated BET drop, the acidity at $\text{pK}_a < -3$ and the conversion in the polymerization of PTHF. It can clearly be seen from the values reported in Table 3 that the reduced surface area of the thermally pretreated TiO_2 *aq is less susceptible to further sintering and leads to WO_3/TiO_2 -based catalysts which have a significantly increased number of centers having a pK_a of 25 < -3 . Maximum acidities and activities for acid catalysis are obtained, in particular, at a BET surface area of the treated TiO_2 *aq support in the region of $200 \text{ m}^2/\text{g}$. The PTHF polymerization was carried out by the continuous process described as Variant A under IV.2. above.

Table 3: Comparison of the catalysts

Precalcination of the precipitated TiO_2 *aq [$^{\circ}\text{C}$]	BET of the TiO_2 *aq after precalcination [m^2/g]	Loss on ignition of the TiO_2 *aq after precalcination	Temperature of the final calcination [$^{\circ}\text{C}$]	Max. BET of the catalyst [m^2/g]	BET of the catalyst [m^2/g]	BET drop [1-BET/BET (max)]	Acidity at $\text{pK}_a < 3$ [$\mu\text{mol/g}$]	Conversion in the THF polymerization $M_n(\text{pTHF}) = 2000$
500	142	9%	550	109	91	17%	95	
500	142	9%	575	109	97	11%	96	
500	142	9%	600	109	92	16%	100	
500	142	9%	630	109	87	20%	95	5.3%
430	195	9%	550	150	99	34%	95	
430	195	9%	575	150	121	19%	104	7.0%
430	195	9%	600	150	108	28%	104	
430	195	9%	630	150	86	43%	97	6.3%
410	214	9%	550	165	120	27%	87	
410	214	9%	575	165	116	30%	98	7.7%
410	214	9%	600	165	95	42%	95	
410	214	9%	630	165	96	42%	88	
390	225	9%	550	173	118	32%	85	3.7%
390	225	9%	575	173	105	39%	89	
390	225	9%	600	173	138	20%	90	3.0%
390	225	9%	630	173	104	40%	88	

200	275	10%	550	214	114	47%	59	3.0%
200	275	10%	575	214	112	48%	59	
200	275	10%	600	214	97	55%	72	
200	275	10%	630	214	100	53%	70	4.3%

Example III.2: Preparation of the catalyst under an atmosphere comprising water vapor

Variant 2a

5 2 kg of freshly precipitated $\text{TiO}_2 \cdot \text{aq}$ which had been dried at 100°C (78% TiO_2) were calcined at 400°C in a rotary tube through which air was not passed, i.e. under the water vapor atmosphere given off from the $\text{TiO}_2 \cdot \text{aq}$, for 3 hours in a batch mode. This resulted in a material having a BET surface area of $143 \text{ m}^2/\text{g}$ and a loss on ignition of 9%. 720 g of this material were subsequently kneaded with
10 322 g of ammonium paratungstate (88% WO_3), 156 g of oxalic acid and 290 g of H_2O and shaped to form extrudates having a diameter of 4 mm. These were subsequently calcined stepwise at 200°C for 2 hours and 500°C for 3 hours and finally calcined at 575°C for one hour to give the finished catalyst. The catalyst extrudates had a mean cutting hardness of 34 N and gave a conversion
15 of 5.2% in the continuous THF polymerization test.

Example III.2: Preparation of the catalyst under an atmosphere comprising water vapor

Variant 2b

20 1.2 kg of freshly precipitated $\text{TiO}_2 \cdot \text{aq}$ which had been dried at 100°C (78% TiO_2) were precalcined at 400°C and a residence time of 20 minutes in a continuously operated rotary tube having an internal diameter of 4 cm over a period of 5 hours. 400 standard liters of air per hour were passed through the rotary tube. In addition,
25 60 ml per hour of liquid H_2O were metered into the stream of air. This resulted in a material having a BET surface area of $195 \text{ m}^2/\text{g}$ and a loss on ignition of 8%. 216 g of this material were subsequently kneaded with 96.6 g of ammonium paratungstate (88% WO_3), 46.8 of oxalic acid and 80 g of H_2O and shaped to form extrudates having a diameter of 4 mm. These were subsequently calcined stepwise
30 at 200°C for 2 hours and 500°C for 3 hours and finally calcined at 575°C for 1 hour. The catalyst extrudates had a mean cutting hardness of 20 N and gave a conversion of 6.0% in the continuous THF polymerization test.

35 The following comparative experiments demonstrate the influence of the addition of water during the precalcination of the support material on the cutting hardness

and also demonstrate the influence of the precalcination as such on the activity of the finished catalyst.

Comparative Example III.1 to examine the cutting hardness

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1.2 kg of freshly precipitated $\text{TiO}_2 \cdot \text{aq}$ which had been dried at 100°C (78% TiO_2) was precalcined as above, but no H_2O was metered into the rotary tube. This resulted in a material having a BET surface area of $211 \text{ m}^2/\text{g}$ and a loss on ignition of 9%. 720 g of this material were subsequently kneaded with 322 g of ammonium paratungstate (88% WO_3), 156 g of oxalic acid and 350 g of H_2O and shaped to form extrudates having a diameter of 4 mm. These were subsequently calcined stepwise at 200°C for 2 hours and at 500°C for 3 hours and finally calcined at 575°C for one hour.

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The catalyst extrudates had a mean cutting hardness of 8 N and gave a conversion of 7.5% in the continuous THF polymerization test.

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In this comparative example, the support material $\text{TiO}_2 \cdot \text{aq}$ was prepared according to the present invention, but was precalcined according to Variant 1 without addition of water. The catalysts prepared using this support were found to be mechanically unstable compared to the catalysts prepared according to the Variants 2a and 2b.

Comparative Example III.2 to examine the catalyst activity

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750 g of freshly precipitated $\text{TiO}_2 \cdot \text{aq}$ which had been dried at 100°C (78% TiO_2) were kneaded with 230 g of ammonium paratungstate (88% WO_3), 156 g of oxalic acid and 190 g of H_2O and shaped to form extrudates having a diameter of 4 mm. These were subsequently calcined stepwise at 200°C for 2 hours, at 500°C for 3 hours and at 575°C for 1 hour.

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The catalyst extrudates had a mean cutting hardness of 26 N and gave a conversion of 3.2% in the continuous THF polymerization test.

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In this comparative example, the support material $\text{TiO}_2 \cdot \text{aq}$ was not precalcined at all. Compared to the catalysts prepared according to the present invention using Variants 1 and 2a, 2b, a catalyst having a low activity was obtained.

We claim:

1. A catalyst which comprises a catalytically active amount of at least one oxygen-containing molybdenum and/or tungsten compound on an oxidic support material and has been calcined at from 400°C to 900°C after application of precursor compounds of the catalytically active compounds to the support material or a support material precursor, wherein the catalyst has a porosity comprising
 - transport pores having a diameter of > 25 nm, and
 - a volume of these transport pores of at least 50 mm³/g.
2. A catalyst as claimed in claim 1 which has a molybdenum and/or tungsten content, based on the finished catalyst, of:

$$x \text{ } \mu\text{mol of (tungsten and/or molybdenum)/m}^2$$
 where $10.1 < x < 20.9$.
3. A catalyst as claimed in claim 1 or 2 which has a BET surface area of from 25 to 200 m²/g, preferably from 40 to 150 m²/g and particularly preferably from 50 to 120 m²/g.
4. A catalyst as claimed in any of claims 1 to 3, wherein the oxidic support material has a BET surface area of from 135 to 220 m²/g.
5. A catalyst as claimed in claim 4, wherein the oxidic support material has a BET surface area of from 170 to 220 m²/g, preferably from 190 to 220 m²/g.
6. A catalyst as claimed in any of claims 1 to 5 which has an acidity at pK_a < -3 of at least 10 μmol/g.
7. A catalyst as claimed in any of claims 1 to 6 which has an acidity at at pK_a < -3 of at least 70 μmol/g, preferably at least 90 μmol/g, in the dry state.

8. A catalyst as claimed in any of claims 1 to 7 which comprises a catalytically active amount of at least one oxygen-containing tungsten compound.
- 5 9. A catalyst as claimed in any of claims 1 to 8, wherein the oxidic support material is zirconium dioxide or titanium dioxide, preferably titanium dioxide.
- 10 10. A catalyst as claimed in any of claims 1 to 9 which contains from 0.1 to 70% by weight, preferably from 5 to 40% by weight, particularly preferably from 10 to 35% by weight, of the catalytically active compound(s), calculated as MoO_3 and/or WO_3 and based on the total weight of the catalyst.
- 15 11. A catalyst as claimed in any of claims 1 to 10 which has a content of pore formers of up to 200% by weight, preferably up to 100% by weight, particularly preferably up to 60% by weight, in each case based on the catalytically active compound together with the support material.
- 20 12. A catalyst as claimed in any of claims 1 to 11 which contains up to 15% by weight, preferably up to 10% by weight, particularly preferably up to 8% by weight, of sulfate and/or phosphate.
- 25 13. A process for preparing a catalyst comprising a catalytically active amount of at least one oxygen-containing molybdenum and/or tungsten compound as an oxidic support material, which comprises the following steps:
- precipitation of the oxidic support material,
 - thermal treatment of the oxidic support material to set a BET surface area of 135-220 m^2/g ,
 - 30 - application of the catalytically active compound(s) or its/their precursor compound(s) to the support material and
 - calcination of the resulting catalyst at from 400°C to 900°C to give the finished catalyst.

14. A process as claimed in claim 13, wherein the thermal treatment of the
oxidic support material is carried out in a water-containing atmosphere
comprising more than 30% by weight, preferably more than 50% by weight
and particularly preferably more than 70% by weight, of water.
- 5 15. A process as claimed in claim 13 or 14, wherein the catalytically active
compound(s) is/are applied in a molar ratio of support material to
catalytically active compound of 5.0 - 7.5, preferably 6.0 - 7.0.
- 10 16. A process as claimed in any of claims 13 to 15, wherein the catalyst is
calcined firstly at 100-600°C and subsequently at up to 900°C, preferably
up to 700°C, to give the finished catalyst.
- 15 17. A process as claimed in any of claims 13 to 16, wherein the calcination is
carried out with a stepwise increase in the temperature.
18. A process for preparing polytetrahydrofuran, tetrahydrofuran copolymers,
diesters or monoesters of these polymers by polymerization of
tetrahydrofuran in the presence of at least one telogen and/or comonomer
using a catalyst as claimed in any of claims 1 to 12.
- 20 19. A process for preparing polytetrahydrofuran, tetrahydrofuran copolymers,
diesters or monoesters of these polymers by polymerization of
tetrahydrofuran in the presence of at least one telogen and/or comonomer
using a catalyst which has been obtained by a process as claimed in any of
claims 13 to 17.
- 25 20. A process as claimed in claim 18 or 19 for preparing polytetrahydrofuran.
- 30 21. A process as claimed in any of claims 18 to 20 for preparing
polytetrahydrofuran from tetrahydrofuran and 1,4-butanediol or H₂O as
telogen.
- 35 22. A process as claimed in claim 18 or 19 for preparing polytetrahydrofuran
from tetrahydrofuran and carboxylic anhydride.