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DRY ANALYTICAL ELEMENT FOR LITHIUM ASSAY

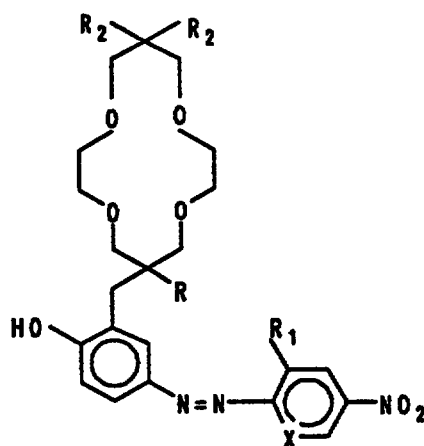
This invention relates to the field of clinical chemistry, particularly elements for assaying lithium in serum.

5 Lithium in the form of lithium carbonate is administered to manic-depressive patients. The therapeutic range of lithium ion in plasma is quite narrow, namely, 0.8 to 1.2 mM. It is important to monitor the lithium level in such patients because of
10 the toxic side effects that appear when the lithium level in blood exceeds the recommended level.

It is known to assay lithium quantitatively using solution assays and ion-selective electrodes. Dyes, such as 14-Crown-4-ether derivatives are known
15 for use in such assays. Japanese Kokai 62/72683 (1985) discloses a class of such dyes that act as colorimetric reagents in extracting lithium from solutions and as charge transport carriers for ion-selective electrodes. The problem is that none of the known 14-crown-4-ether
20 derivatives can be adapted for use in dry analytical elements. Moreover the above "wet" methods involving solution extractions or ion-selective electrodes suffer from severe ion and other interferant problems. It would be desirable to have a dry analytical element for
25 assaying lithium that obviated these problems.

This invention solves the above problems by providing a multilayer dry analytical element for quantitatively assaying lithium comprising a spreading layer, and a reagent layer characterized in that
30 sufficient buffer is present in a layer above the reagent layer to establish a pH of at least 9 and the reagent layer comprises a 14-Crown-4-ether derivative selected from Table 1 below:

TABLE 1
14-Crown-4-Ether Derivatives



5

Dye No.	R	R ₁	R ₂	X
1	n-C ₁₂ H ₂₅	SO ₂ CH ₃	CH ₃	CH
2	n-C ₁₂ H ₂₅	NO ₂	CH ₃	CH
3	CH ₃	SO ₂ N(n-C ₆ H ₁₃) ₂	CH ₃	CH
4	CH ₂ C ₆ H ₅	SO ₂ N(n-C ₆ H ₁₃) ₂	CH ₃	CH
5	CH ₂ C ₆ H ₅	SO ₂ N(CH ₃)-n-C ₁₈ H ₃₇	CH ₃	CH
6	CH ₃	NO ₂	n-C ₇ H ₁₅	CH
7	n-C ₁₂ H ₂₅	H	CH ₃	N

Using the element of this invention a colorometric assay for lithium can be performed directly on the sample without deproteinization. The use of the above selected dyes of Table 1 is unobvious, since not all 14-Crown-4-ether derivatives are useful even in solution assays of lithium.

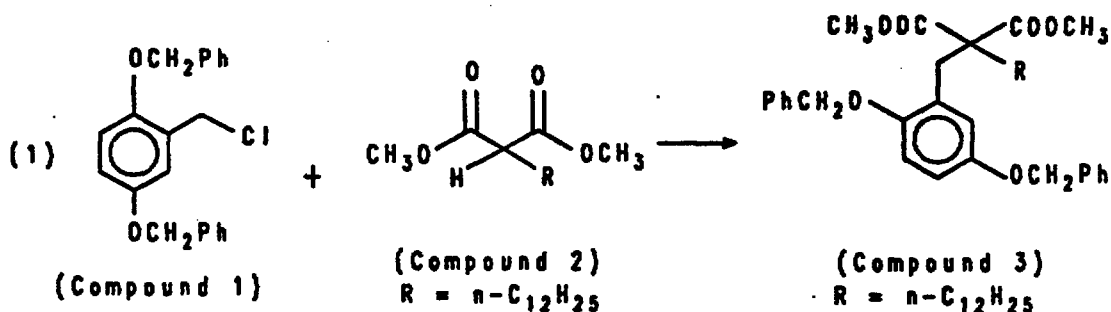
The following procedure was used to prepare dyes 2-7, 10-12 of Table 2, *infra*. The procedure comprises the steps of:

- a) reducing a 2-substituted-2-[2,5-bis-(alkoxy)benzyl]malonate ester with a strong reducing agent to produce a 2-substituted-2-[2,5-bis(alkoxy)benzyl]-1,3-propanediol, said alkoxy group being
5 benzyloxy or unsubstituted alkoxy of 1 to 10 carbon atoms,
- b) condensing the 1,3-propanediol of step a) with a 5,5-disubstituted-3,7-dioxanonane-1,9-diyl-bis(toluene p-sulfonate) to form a 6,13,13-trisubstituted-6-[2,5-bis(alkoxy)benzyl]-1,4,8,11-tetraoxa-
10 cyclotetradecane,
- c) hydrogenating the bis(alkoxy) compound of step b) to the hydroquinone, 6,13,13-trisubstituted-6-(2,5-dihydroxybenzyl)-1,4,8,11-tetraoxacyclotetradecane,
15 decane,
- d) oxidizing the hydroquinone of step c) to the quinone, 6,13,13-trisubstituted-6-(1,4-cyclohexadiene-3,6-dione-1-ylmethyl)-1,4,8,11-tetraoxacyclotetradecane, and
- 20 e) condensing the quinone of step d) with an N-(p-nitrophenyl)hydrazine to produce the 14-crown-4-ether azo dye product.

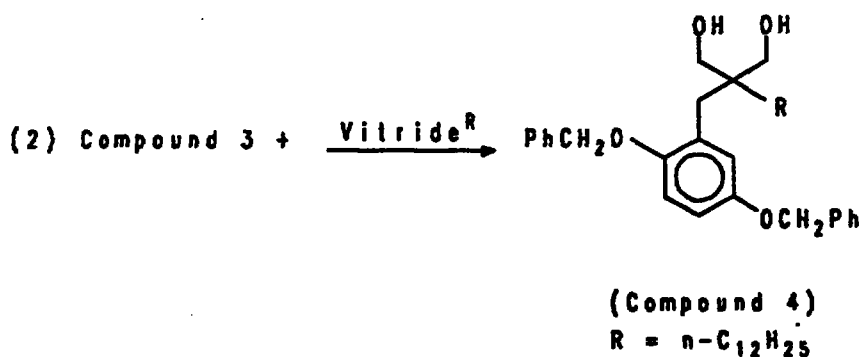
Preparation Sequence I

The preparation of the diol portion of the
25 dyes is shown in Sequence I, *infra*. A 2,5-dibenzyloxybenzyl chloride, compound 1, is condensed with dimethyl 2-alkylmalonate, compound 2, yielding a 2-substituted-2-[2,5-bis(alkoxy)benzyl]malonate ester, compound 3, which is subsequently reduced with a
30 strong reducing agent to produce a 2-substituted-2-[2,5-bis(alkoxy)benzyl]-1,3-propanediol, said alkoxy group being benzyloxy or unsubstituted alkoxy of 1 to 10 carbon atoms, compound 4.

SEQUENCE I



5



1. Preparation of dimethyl 2-dodecyl-2-[2,5-bis(benzyloxy)benzyl]malonate (Compound 3)

10 A solution of 121.4 g (404.7 mmol) of Compound 2 (dimethyl 2-dodecylmalonate) in 600 mL of dimethylformamide was added to 56.0 g (500 mmol) of potassium tert. butoxide . After 30 minutes the reaction mixture was cooled to 0°C and 137 g (404.7

15 mmol) of Compound 1, 2-(chloromethyl)-1,4-bis(phenylmethoxy)benzene, was added and the mixture allowed to warm to room temperature for 30 minutes. The mixture was next heated at 60°C for 90 minutes and an additional 4.5 g of Compound 2 and 4.5 g of

20 potassium tertiary butoxide was added. The reaction mixture was then heated at 60°C for 10 hours. Removal

of the solvent yielded a viscous oil. This oil was partitioned between ethyl acetate and dilute hydrochloric acid, and the organic phase dried over anhydrous sodium sulfate, filtered and concentrated to yield an impure viscous oil (Compound 3). Compound 3 was used as prepared in the next step.

2. Preparation of 2-dodecyl-2-(2,5-di(benzyloxy)-benzyl)-1,3-propanediol (Compound 4)

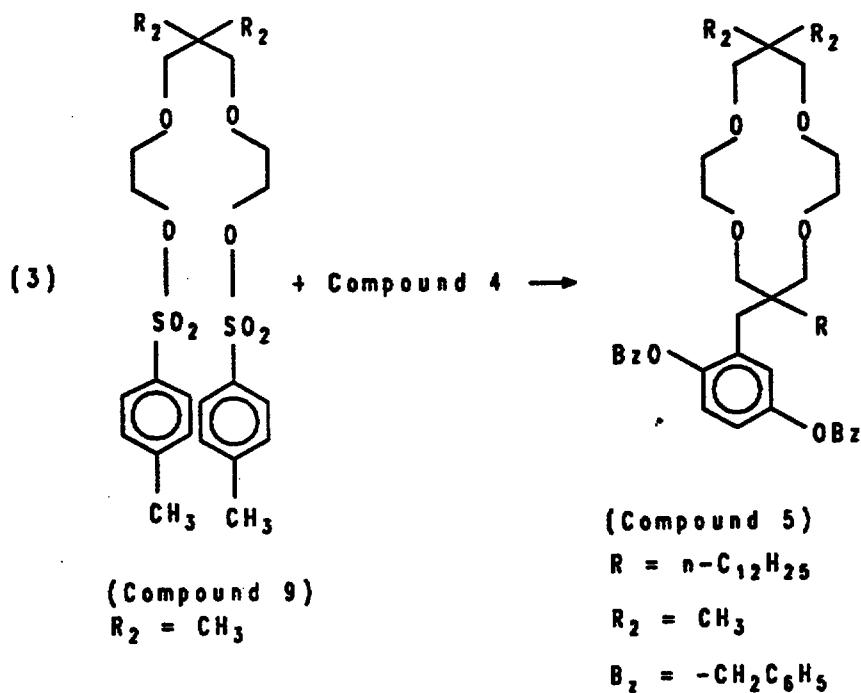
10 A solution of 302.4 g (502.3 mmol) of Compound 3 in 2 L of toluene was prepared. Then 200 mL of the toluene was distilled off to remove residual water. After cooling to room temperature 303.2 g (1.05 mol) of sodium bis(2-methoxyethoxy)aluminum hydride
15 (Vitride®) was added and the solution stirred for four hours. The reaction was quenched with 90 mL of water, 90 mL 15% sodium hydroxide and 275 mL water. The toluene phase was isolated and the aqueous phase extracted with 500 mL of toluene. The combined organic
20 extracts were dried over anhydrous sodium sulfate, filtered and the solvent removed. The resulting oily product was triturated with petroleum ether. The resulting solid was filtered and washed with petroleum ether yielding a compound having an ¹H NMR spectrum
25 consistent with the structure of compound 4 (melting point 64-65 °C).

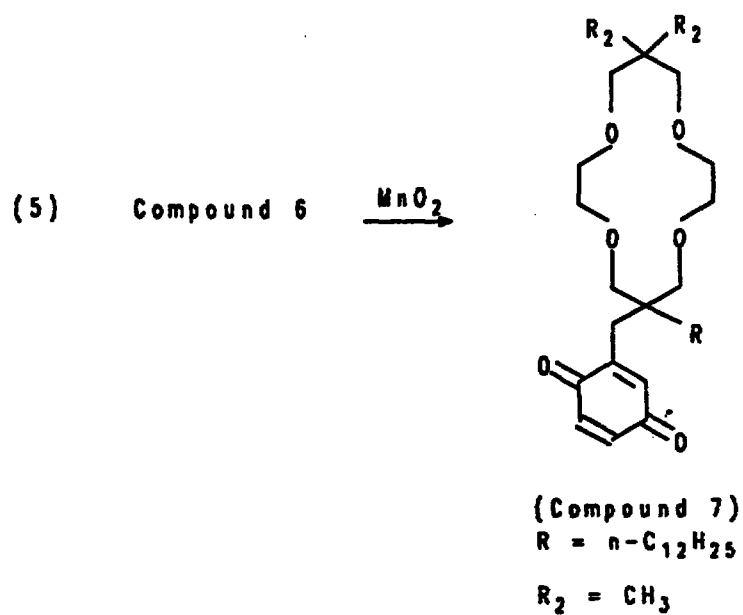
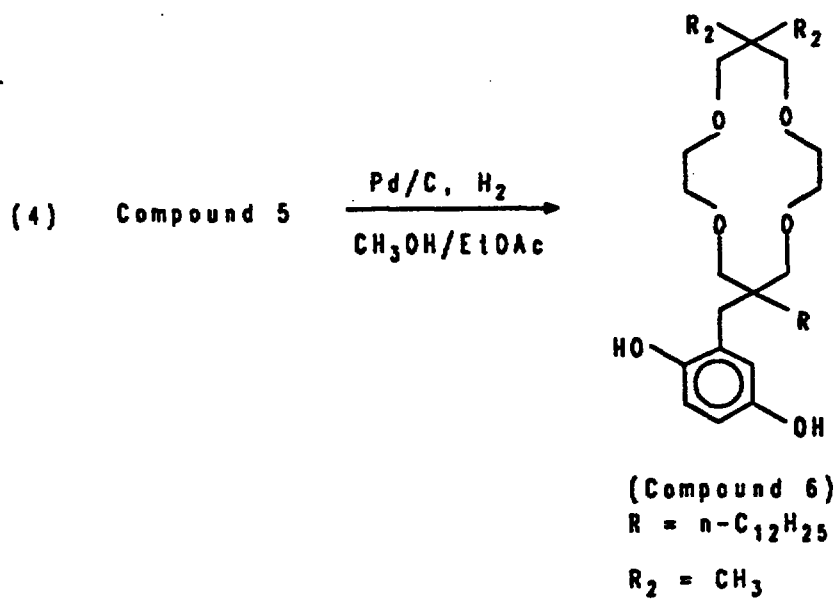
Preparation sequence II

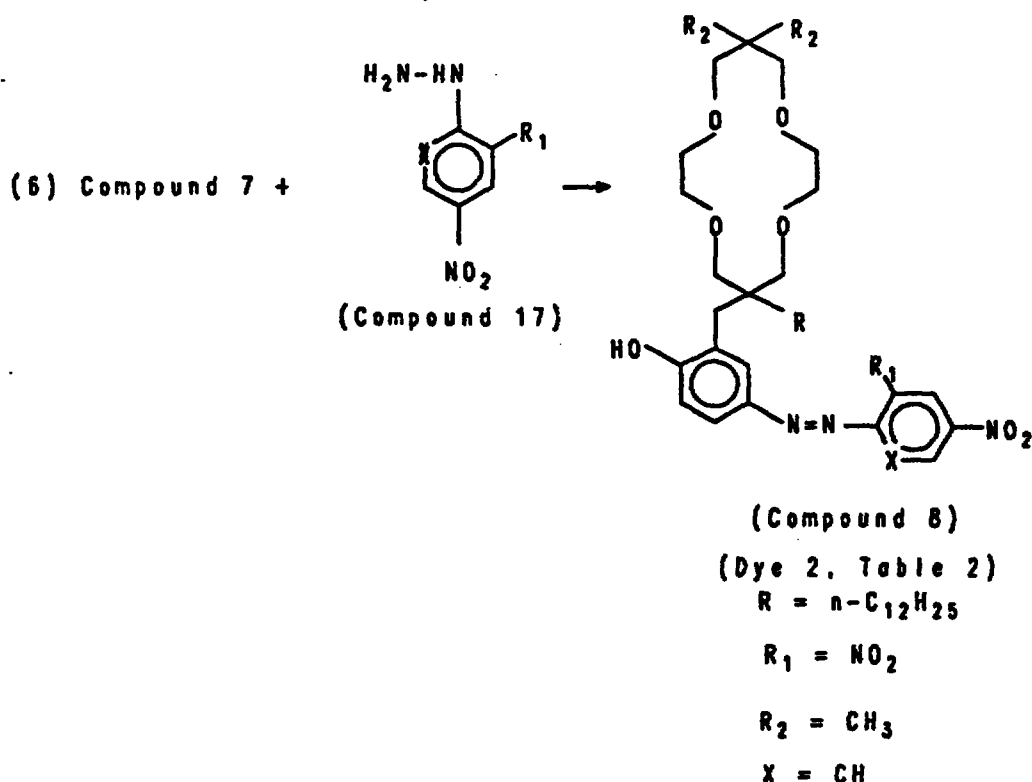
The final preparation steps for making the proton-ionizable 14-crown-4-ether dyes (Dyes 2-7 and
30 10-12 of Table 2) are shown in Sequence II. The initial ionophore is formed by condensation of 1,3-propanediol, compound 4, with a 5,5-disubstituted-3,7-dioxanonane-1,9-diylbis(toluene p-sulfonate), compound 9, a ditosylate, to form a 6,13,13-trisubstituted-6-
35 [2,5-bis(alkoxy)benzyl]-1,4,8,11-tetraoxa-

cyclotetradecane, Compound 5. Compound 9 was prepared according to methods described in the literature. Compound 5 is then hydrogenated to the hydroquinone, 6,13,13-trisubstituted-6-(2,5-dihydroxybenzyl)-1,4,8,11-tetraoxacyclotetradecane, compound 6. Compound 6 is then oxidized to yield 6,13,13-trisubstituted-6-(1,4-cyclohexadiene-3,6-dione-1-ylmethyl)-1,4,8,11-tetraoxacyclotetradecane, Compound 7. The desired dyes are then prepared by condensation of the quinone, Compound 7, with an N-(p-nitrophenyl)hydrazine to produce the 14-crown-4-ether azo dye product, Compound 8.

SEQUENCE II







3. Preparation of 6-Dodecyl-6-[2,5-di(benzyloxy)-benzyl]-13,13-dimethyl-1,4,8,11-tetraoxacyclo-tetradecane (Compound 5)

A suspension of 110 g (219.8 mmol) of compound 9, 5,5-dimethyl-3,7-dioxanonane-1,9-diyl-bis(toluene p-sulfonate), 120 g (219.8 mmol) of compound 4 and 12.2 g (220 mmol) of lithium bromide in 2 L of dry t-pentyl alcohol was reacted with 7.0 g of lithium hydride. After refluxing the reaction mixture for 7 days the solvent was removed and the residue dissolved in a mixture of dichloromethane and dilute hydrochloric acid. The dichloromethane phase was separated, dried over anhydrous magnesium sulfate, filtered and concentrated to a dark oil. The oil was dissolved in 10% dichloromethane/90% petroleum ether and eluted through a silica gel column (600 g) the fractions containing the product (silica gel

TLC/CH₂Cl₂, R_f 0.2) were combined to yield 83.2 g of impure product (Compound 5).

4. Preparation of 6-Dodecyl-6-(2,5-dihydroxybenzyl)-13,13-dimethyl-1,4,8,11-tetraoxacyclotetradecane (Compound 6)

A solution of 15.8 g (22.5 mmol) of Compound 5 in 20 mL of ethyl acetate and 20 mL of methanol was mixed with 0.4 g of palladium-on-carbon in a parr flask under a nitrogen atmosphere. The mixture was reacted at 50-60°C under 50 psi H₂ for 4 hours. After cooling, the product was filtered through Celite diatomaceous earth and the solvent removed to yield 12.7 g of impure product (Compound 6).

5. Preparation of 6-Dodecyl-6-(1,4-cyclohexadiene-3,6-dione-1-ylmethyl)-13,13-dimethyl-1,4,8,11-tetraoxacyclotetradecane (Compound 7)

A solution of 67.6 g (129.3 mmol) of Compound 6 in 600 mL of dichloromethane was mixed with 50.6 g (581.9 mmol) of activated manganese dioxide. The heterogeneous mixture was stirred at room temperature for 1 hour. The insolubles were removed by filtration through Celite diatomaceous earth and the solvent removed. The oily residue was triturated with 500 mL of petroleum ether and cooled. The volume of solvent was reduced to 1/2 and the yellow solid suction filtered and washed with cold pentane.

A first and second fraction of compound 7 was isolated, giving a ¹H NMR spectrum consistent with the assigned structure.

6. Preparation of 6-dodecyl-6-[2-hydroxy-5-(2,4-dinitrophenylazo)benzyl]-13,13-dimethyl-1,4,8,11-tetraoxacyclotetradecane [Compound 8 (dye 2, Table 2)]

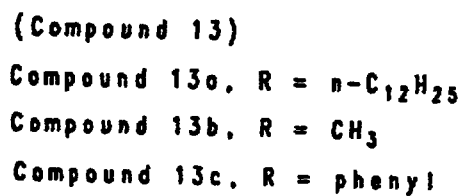
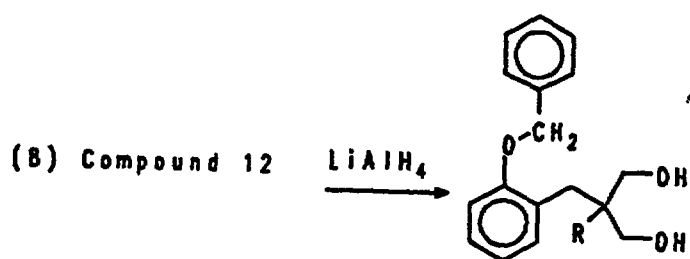
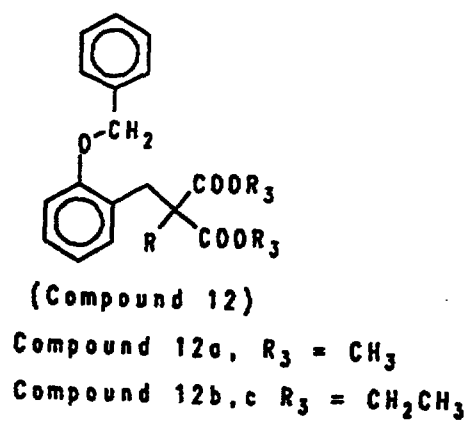
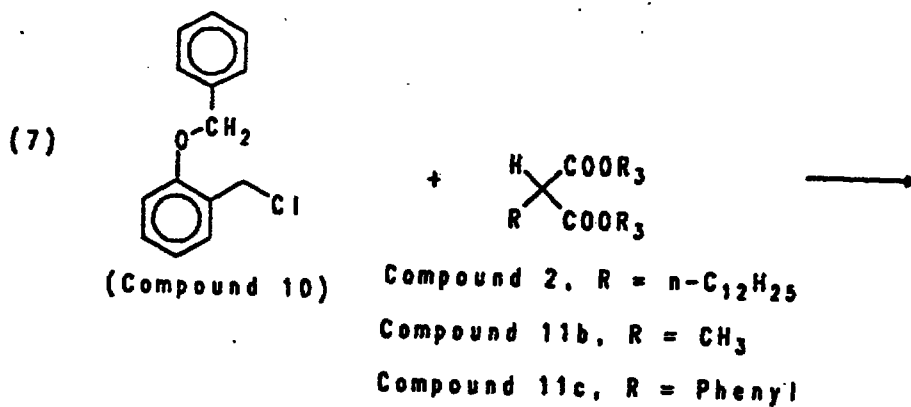
A solution of 31.7 g (60.9 mmol) of compound 7 and 14.2 g (60.9 mmol) of 2,4-dinitrophenylhydrazine (15% H₂O) in 300 cc of acetic acid was heated at 60°C for 90 minute with constant stirring. After removing 5 225 mL of acetic acid, the product was added to 500 mL of petroleum ether. Initially several g of product was recovered. After recrystallization from dichloromethane/petroleum ether compound 8 was obtained, mp 64-65°C. The ¹H NMR spectrum was 10 consistent with the assigned structure.

Preparation Sequence III

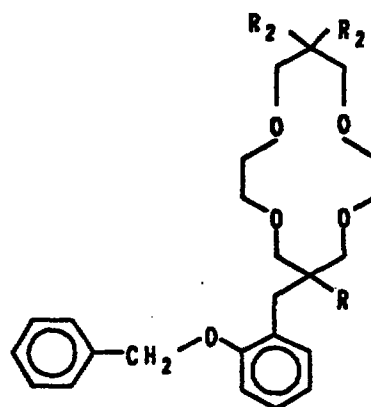
An alternative method was used to prepare Dyes 1, and 8-9 of Table 2. This method is illustrated in Sequence III. The diol used in the sequence was 15 prepared by the condensation of o-benzyloxybenzyl chloride, compound 10, with dialkyl 2-dodecyl- or 2-methyl- or 2-phenylmalonate, to form a substituted dialkyl malonate, compounds 2 or 11b-c. The substituted dialkyl malonate (compounds 12a-c, were 20 subsequently reduced to the diol (compounds 13a-c).

The reaction of the ditosylate, compound 9, and diols compound 13a, b, c) yield the corresponding 14-crown-4 ionophore compounds 14a, b and c. Subsequent hydrogenation yields the proton-ionizable 25 14-crown-4-ether compounds 15a, b and c which are then converted into dyes by coupling with a diazonium salt.

SEQUENCE III



(9) Compound 9 + Compound 13a $\longrightarrow$



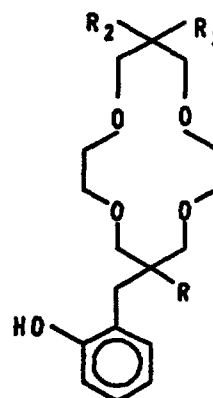
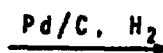
(Compound 14)

$R = n-C_{12}H_{25}$

Compound 14a,

$R_2 = CH_3$

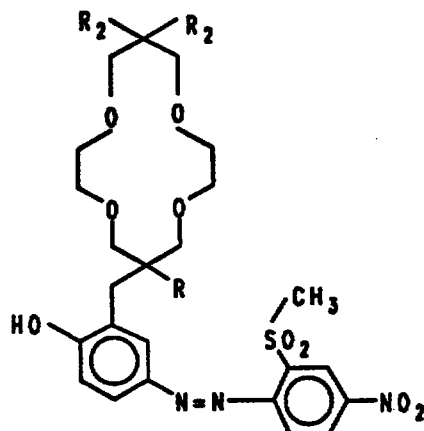
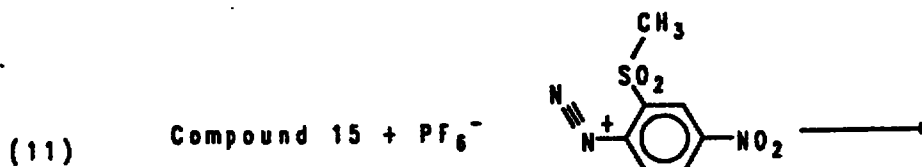
(10) Compound 14



(Compound 15)

Compound 15a, $R = n-C_{12}H_{25}$

$R_2 = CH_3$



(Compound 16)
Compound 16a, R = n-C₁₂H₂₅; R₂ = CH₃ (Dye 1)

7. Preparation of dimethyl 2-dodecyl-2-(2-benzyloxybenzyl)malonate (Compound 12a)

5 A solution of 162.6 g (541.3 mmol) of dimethyl 2-dodecylmalonate (Compound 2) in 800 mL of dimethylformamide was mixed with 67.2 g (600 mmol) of potassium tertiary butoxide under a nitrogen atmosphere. After cooling to room temperature 116.0 g
10 (541.3 mmol) of 2-(benzyloxy)benzyl chloride was added and the mixture heated at 75°C for 2 hr. The product was poured over ice/water, extracted twice with ethyl acetate, dried, filtered and isolated as an oily product. This was used as recovered in the next step.
15 The methyl and phenyl derivatives were prepared in a similar manner (See Sequence III).

8. Preparation of 2-dodecyl-2-(2-benzyloxybenzyl)-1,3-propanediol (Compound 13a)

A solution of 260 g (524 mmoles) of compound 12a in 1 L of tetrahydrofuran (THF) was added to a suspension of 29.9 g (786 mmol) of lithium aluminum hydride in 500 mL of dry THF. The reaction mixture was
5 refluxed for 2 hours and then quenched successively with isopropanol and dilute hydrochloric acid. The product was extracted into ether, dried, filtered and concentrated to dryness.

9. Preparation of 6-dodecyl-6-(2-benzyloxybenzyl)-
10 13,13-dimethyl-1,4,8,11-tetraoxacyclotetradecane
(Compound 14a)

A solution of 40 g (80 mmol) of Compound 9 and 35.2 g (80 mmol) of Compound 13a in 400 mL of dry dimethylformamide (DMF) was slowly added to a
15 suspension of 13.4 g (560 mmol) of sodium hydride and 21.4 g (160 mmol) of lithium iodide in 200 mL of DMF at 70°C. The reaction product was heated for 24 hours, cooled and the solvent removed. The ¹H nmr spectrum of the product is consistent with the assigned structure.
20 The product was extracted into ethyl acetate and washed with dilute hydrochloric acid, then twice with a saturated salt solution, dried over anhydrous magnesium sulfate, filtered and the solvent removed. The product was eluted through a silica gel column, yielding 13.8 g
25 (28.9%). The ¹H NMR spectrum of Compound 14a is consistent with the assigned structure.

10. Preparation of 6-dodecyl-6-(2-hydroxybenzyl)-
13,13-dimethyl-1,4,8,11-tetraoxacyclotetradecane
(Compound 15)

30 A solution of 13.3 g (22.3 mmol) of compound 14 in 200 cc of tetrahydrofuran and 150 cc of acetic acid was mixed with 8.0 g of Pd on carbon. The suspension was hydrogenated at 55 psi for 16 hours at room temperature. After filtration and removal of the

solvent Compound 15 was isolated. The ¹H nmr spectrum is consistent with the assigned structure.

11. Preparation of 6-Dodecyl-6-[2-hydroxy-5-(2-methylsulfonyl-4-nitrophenylazo)benzyl]-13,13-dimethyl-
5 1,4,8,11-tetraoxacyclotetradecane (Compound 16)

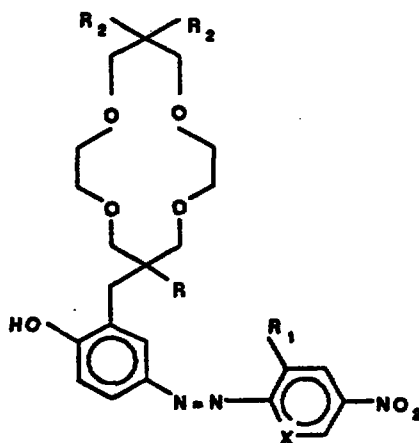
To a solution of 3.1 g (6.1 mmol) of Compound 15 in 50 cc acetic acid at 5°C was added a solution of 3.0 g (8.0 mmol) of 2-methylsulfonyl-4-nitrophenyl-diazonium hexafluorophosphate in 25 cc water and 75 cc
10 tetrahydrofuran. The suspension was stirred at room temperature for 16 hours. The product was extracted into ethyl acetate and isolated by silica gel chromatography yielding 0.22 g (4.7%) of Compound 16. The ¹H nmr spectrum is consistent with the assigned
15 structure.

The following screening procedure was used to identify those 14-Crown-4 ether derivatives that would be useful in formulating a dry analytical element for quantitatively assaying lithium. The data show that
20 some, but not all, derivatives are useful for conducting lithium assays.

Three mL of a 10% dimethyl sulfoxide-water solution containing 3.42×10^{-5} mole of a 14-crown-4-ether dye, 50 μ L of a 1M pH 11 buffer (4-Aminobutyric
25 acid or phosphate or 3-[cyclohexylamino]-1-propane-sulfonic acid [CAPS]) and 50 μ L of aqueous lithium chloride solution was mixed in a disposable cuvette. Lithium solutions were added at the levels 0.0195, 0.0975 and 0.1952 mM which produced mixtures containing
30 0.32, 1.61 and 3.2 mM lithium respectively. Absorbance was measured immediately at $\gamma=600$ nm and again after 5 minutes to monitor the progress of lithium complexing by the 14-crown-4-ether dye. The spectrum of each solution was measured between 300 and 800 nm after each
35 600 nm absorbance measurement. This measurement allows

monitoring of the 14-crown-4-ether absorbance maximum decrease at 400 nm as the lithium is complexed with the dye and a concurrent increase in the absorbance maximum at 600 nm. The delta absorbance values for a
 5 representative number of the dyes tested between 0.32 and 3.2 nm are listed in Table 2. Crown ether dye-lithium complexes possessing delta absorbance values $\geq .12$ are considered satisfactory for use in a dry analytical element for lithium assay provided by the
 10 present invention.

TABLE 2
 DELTA D_T OF 14-CROWN-4 ETHER DERIVATIVES



15

					Delta D_T 3.2- .32 nM
Dye No.	R	R ₁	R ₂	X	
1	n-C ₁₂ H ₂₅	SO ₂ CH ₃	CH ₃	CH	0.57
2	n-C ₁₂ H ₂₅	NO ₂	CH ₃	CH	0.46
3	CH ₃	SO ₂ N(n-C ₆ H ₁₃) ₂	CH ₃	CH	0.31
4	CH ₂ C ₆ H ₅	SO ₂ N(n-C ₆ H ₁₃) ₂	CH ₃	CH	0.12
5	CH ₂ C ₆ H ₅	SO ₂ N(CH ₃)-n- C ₁₈ H ₃₇)	CH ₃	CH	0.18
6	CH ₃	NO ₂	n-C ₇ H ₁₅	CH	0.22

7	n-C ₁₂ H ₂₅	H	CH ₃	N	0.22
8	CH ₃	SO ₂ CH ₃	CH ₃	CH	< 0.1
9	C ₆ H ₅	SO ₂ CH ₃	CH ₃	CH	< 0.1
10	C ₆ H ₅	NO ₂	CH ₃	CH	< 0.1
11	CH ₂ C ₆ H ₅	NO ₂	CH ₃	CH	< 0.1
12	n-C ₁₂ H ₂₅	H	CH ₃	CH	0.09

Phosphate Buffer pH 11.7

The data of Table 2 show that, of the dyes presented, only dyes 1-7 meet DT criteria for use in a dry

5 analytical element.

The element of this invention can be used to assay lithium qualitatively and quantitatively in biological fluids in animals or humans, but preferably in humans. Such fluids include, but are not limited to, whole blood, plasma, sera, lymph, bile, urine, spinal fluid, sputum, perspiration and the like as well as stool secretions. It is also possible to assay fluid preparations of human or animal tissue such as skeletal muscle, heart, kidney, lungs, brains, bone marrow, skin and the like.

Elements of the invention can be configured in a variety of forms, including elongated tapes of any desired width, sheets, slides or chips.

The elements can be used in manual or automated assay techniques. In general, in using the elements, lithium determination is made by taking the element from a supply roll, chip packet or other source and physically contacting it with a sample (for example, up to 200μL) of the liquid to be tested so that the sample and reagents interact sequentially within the element and become mixed. Such contact can be accomplished in any suitable manner, for example, by dipping or immersing the element into the sample or, preferably, by spotting the element by hand or machine

with a drop of the sample with a suitable dispensing means.

After sample application, the element is incubated, for a period of up to 5 minutes, to facilitate color development. By incubation, we simply mean that the reagents are maintained in contact with each other for a period of up to 5 minutes before color measurements are made.

The dry analytical elements of this invention are multilayered. In one embodiment the element includes a spreading layer, a subbing layer which contains buffer and an additional reagent layer having at least two distinct zones. All of the foregoing layers are coated on a support.

Another embodiment includes a combined reagent/spreading layer along with a subbing layer and an additional reagent layer having at least two distinct zones. In this element buffer is incorporated into the spreading layer by means of either coating on the complete element or ball milling buffer into the spreading layer prior to coating.

In all embodiments the layers are generally in fluid contact with each other, meaning that fluids (reagents for example buffer) can be transported between superposed regions of adjacent zones. In other words, when the element is contacted with an aqueous fluid, all reagents of the analytical composition of this invention are mixed sequentially as stated hereinbefore.

Useful spreading layers can be prepared using fibrous materials, either mixed with a suitable binder material or woven into a fabric, as described in U.S. Patent 4,292,272, polymeric compositions or particulate materials, for example a blush polymer such as disclosed in U.S. Patent 3,992,158, beads bound together with or without binding adhesives, as

described in U.S. Patents 4,258,001 and 4,430,436 and Japanese Patent Publication 57(1982)-101760.

Particularly useful spreading layers comprise barium sulphate or titanium dioxide. Since the sample is
5 generally applied directly to the spreading layer, it is desirable that the spreading layer be isotropically porous, meaning that the porosity is the same in each direction in the layer as caused by interconnected spaces or pores between particles, fibers or polymeric
10 strands.

The elements of this invention can also contain one or more other addenda commonly put in the elements for various manufacturing or operational advantages. Such addenda include surfactants,
15 bacteriostats, buffers, solvents, hardeners and other materials known in the art.

The layers can be coated on transparent supports such as polyethylene terephthalate. Other supports are well known in the art.

20 Besides the references noted above, suitable element components are described also, for example, in U. S. Patents 4,042,335; 4,132,528; and 4,144,306 .

Examples of the Invention

Dye 2 of Table 2 was tested in the following
25 dry analytical element. The element was prepared according to procedures known in the dry analytical art as described in the aforementioned patents.

LITHIUM ELEMENT

Reagent/Spreading Layer	Coverage g/m2	
	Useful	Actual
	Range	Range
Barium Sulfate	75-120	100
Cellulose Acetate	1-20	8
Surfactant TX-100	0.2-2	0.7

Estane	1-5	3
o-Xylene	59-68	65
Subbing Zone (pH 11)		
Polyvinylpyrrolidone	2-12	4
Buffer	2-10	4
Reagent Layer Zone 1 (pH 7.0)		
Hardened Gelatin	4-24	10
Surfactant F-35	0.005-0.1	0.01
Reagent Layer Zone 2 (pH 7.0)		
Unhardened Gelatin	4-24	12
Surfactant F-35**	0.005-0.1	0.01
14-crown-4-ether Dye	0.2-1.0	0.6
2,4-Di-n-amylphenol*	2-10	3.2

Buffers: The buffer used is specified in tables 3 and 4.
CAPS pH 11, 4-aminobutyric acid pH 11,
Phosphate pH 11.7 arginine pH 12.3 can be
substituted.

*Other Organic Solvents such as 2,4-di-tert-amyl
phenol; N-n-butylacetanilide or
diethylauramide can be used.

**Other Surfactants such as the Alkanol-XC sodium
naphthalene sulfonate mixture from DuPont can
be used.

The elements were kept at 25°C and 15%
relative humidity until tested. Then they were spotted
with solutions containing human serum spiked with
lithium chloride. The serum contained actual lithium
concentrations varying between 0.27 and 10.59 mM.
After 5 minutes incubation at 37°C, the reflectance
density was measured on a recording reflectance
densitometer at $\lambda=600$ nm. The lithium concentrations
obtained from these reflectance densities are the
predicted concentrations for the various spiked lithium

levels. The results for each dye tested is presented in Tables 3 and 4.

TABLE 3
ASSAY WITH DYE 2 OF TABLE 1

DR	ACTUAL	PREDICTED
0.68915	0.28	0.22
0.69625	0.50	0.37
0.71936	0.80	0.88
0.72882	1.00	1.11
0.75111	1.52	1.68
0.78255	2.50	2.57
0.81521	3.72	3.59
0.85550	5.00	5.00
0.98290	10.60	10.50

5

Buffer; CAPS is pH 11, 2.0 g/m²

TABLE 4
ASSAY USING DYE 2 OF TABLE 1

DR	ACTUAL	PREDICTED
0.81035	0.28	-
0.82587	0.50	0.49
0.83864	0.80	0.89
0.85304	1.00	1.36
0.86021	1.52	1.59
0.88924	2.50	2.58
0.91160	3.72	3.39
0.95331	5.00	4.97
1.08228	10.60	10.62

10

Buffer; 4-Aminobutyric Acid pH 11, 4.0 g/m²
Buffer; 4-Aminobutyric Acid pH 11, 2.0 g/m²

Comparison of the actual values to the
predicted values in Tables 3 and 4 shows that the dyes
selected for use in the analytical elements of the

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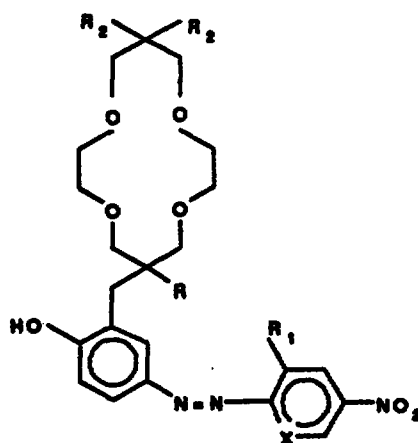
invention provide quantitative analysis of lithium ion in serum containing interferants such as sodium, potassium, calcium , magnesium, iron and proteins. Dye 2 was tested in a similar element and gave similar results. Moreover it is expected that each of the dyes of table 1 having a delta absorbance of at least 0.12 will provide similar results.



CLAIMS:

1. A multilayer dry analytical element for quantitatively assaying lithium comprising a spreading layer, and a reagent layer characterized in that sufficient buffer is present in a layer above the reagent layer to establish a pH of at least 9 and the reagent layer comprises a 14-Crown-4-ether derivative selected from Table 1 below:

TABLE 1
14-Crown-4-Ether Derivatives



Dye				
No.	R	R ₁	R ₂	X
1	n-C ₁₂ H ₂₅	SO ₂ CH ₃	CH ₃	CH
2	n-C ₁₂ H ₂₅	NO ₂	CH ₃	CH
3	CH ₃	SO ₂ N(n-C ₆ H ₁₃) ₂	CH ₃	CH
4	CH ₂ C ₆ H ₅	SO ₂ N(n-C ₆ H ₁₃) ₂	CH ₃	CH
5	CH ₂ C ₆ H ₅	SO ₂ N(CH ₃)-n-C ₁₈ H ₃₇	CH ₃	CH
6	CH ₃	NO ₂	n-C ₇ H ₁₅	CH
7	n-C ₁₂ H ₂₅	H	CH ₃	N



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2. The element of claim 1 wherein the reagent layer has two distinct zones.

3. The element of claim 1 or 2 wherein the buffer is incorporated in a subbing layer between the spreading layer and the reagent layer.

4. The element of claim 1 or 2 wherein the buffer is incorporated in the spreading layer.

5. The element of claim 1 or 2 wherein dye 1 or 2 of table 1 is used in the element.

10 6. The element of claim 1 or 2 wherein the buffer is present in amount sufficient to establish a pH of at least 10.