

DESCRIPTION

CARBOXYLATE METAL COMPLEX AND CATALYST FOR OLEFIN POLYMERIZATION

5 TECHNICAL FIELD

[0001]

The present invention relates to a novel metal complex of group 10 elements of the periodic table, a catalyst mainly comprising the metal complex for olefin polymerization, and a method for producing a polymer using the catalyst.

BACKGROUND ART

[0002]

In recent years, a method for producing a polymer using complexes of late transition metals as a catalyst for olefin polymerization has become known.

Generally, metal complexes having a metal-hydrogen bond or metal-carbon bond have been used as a catalyst system used for coordination polymerization.

J. Am. Chem. Soc., 127, 5132 (2005) (Non-patent document 1) discloses a polymerization mechanism and a catalyst deactivation mechanism in polymerization of vinyl acetate using a metal complex of group 10 elements of the periodic table. Particularly, the document proposes β -acetoxy elimination after the vinyl acetate insertion into the metal-alkyl bond as one of the catalyst deactivation mechanisms. This shows that the metal-acetoxy complex generated by β -acetoxy elimination does not have polymerization activity and it was generally considered that the olefin insertion into the metal-oxygen bond cannot happen.

[0003]

Chem. Commun., 744 (2002) (Non-patent document 2) discloses copolymerization of ethylene and methyl acrylate using the product obtained by allowing the zero-valent late transition metal and ligand to react in advance as a catalyst. In this catalyst system, it is presumed that the generation of a metal-hydride complex initiates the polymerization reaction. However, it is considered that the decomposition of the catalyst species proceeds due to the low thermal stability of the metal-hydride complex, resulting in the low polymerization activity.

[0004]

J. Am. Chem. Soc., 131, 14606 (2009) (Non-patent document 3) discloses the copolymerization of ethylene and vinyl acetate using the isolated metal-methyl complex catalyst, and shows that the polymerization activity was noticeably improved in the case of using the isolated metal-methyl complex catalyst compared to the case of using a metal-hydride complex as generated in situ. However, metal-alkyl complexes generally have a problem in the stability such that they tend to decompose by the impurities during the polymerization and there was a challenge for the industrial use.

PRIOR ART

Non-patent Documents

[0005]

Non-patent Document 1: J. Am. Chem. Soc., 127, 5132 (2005)

Non-patent Document 2: Chem. Commun., 744 (2002)

Non-patent Document 3: J. Am. Chem. Soc., 131, 14606 (2009)

DISCLOSURE OF THE INVENTION

Problems to be Solved by the Invention

[0006]

An object of the present invention is to provide a metal complex of group 10 elements of the periodic table which is more stable than conventional late transition metal catalysts for olefin polymerization, a catalyst mainly comprising the metal complex for olefin polymerization, and a method for producing a polymer using the catalyst. In the present specification, "olefin" means a hydrocarbon compound having an ethylenic carbon-carbon double bond with or without a polar group.

Means to Solve the Problem

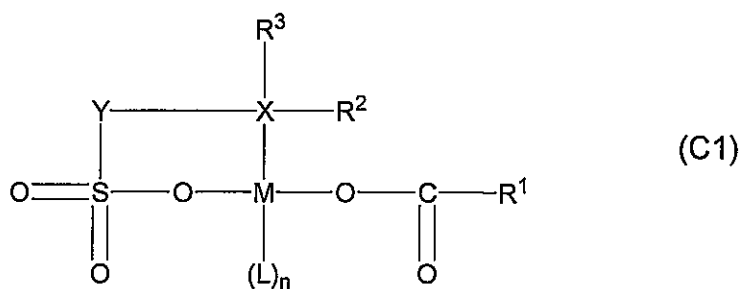
[0007]

As a result of intensive studies to solve the above-mentioned problems, the present inventors have developed a stable metal complex of group 10 elements of the periodic table, in which all of the coordinating atoms to the metal are a hetero atom, unlike the conventional late transition metal catalysts for olefin polymerization; have found that the olefin polymerization proceeds when the metal complex is used as a catalyst and accomplished the present invention.

[0008]

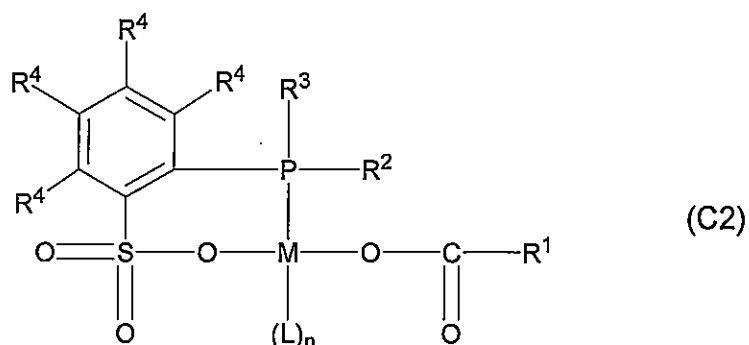
That is, the present invention relates to the following [1] to [11].

[1] A carboxylate metal complex represented by formula (C1)



(in the formula, M represents a metal atom of group 10 element in the periodic table; X represents a phosphorous atom (P) or an arsenic atom (As); R¹ represents a hydrogen atom or hydrocarbon group having 1 to 30 carbon atoms which may be substituted by one or more functional groups selected from a halogen atom, alkoxy group, aryloxy group, acyloxy group and amino group; R² and R³ independently represent a hydrogen atom, alkoxy group, aryloxy group, acyloxy group, silyl group, amino group or hydrocarbon group having 1 to 30 carbon atoms which may be substituted by one or more functional groups selected from a halogen atom, alkoxy group, aryloxy group, acyloxy group and amino group; R² and R³ may bond to each other to form a ring structure; Y represents divalent hydrocarbon group having 1 to 70 carbon atoms which may be substituted by one or more functional groups selected from a halogen atom, alkoxy group, aryloxy group, acyloxy group and amino group; L represents a neutral electron-donating ligand and n is 0, 1/2, 1 or 2.

[2] The carboxylate metal complex as described in [1] above, wherein the carboxylate metal complex represented by formula (C1) is represented by formula (C2)



(in the formula, four R⁴s independently represent a hydrogen atom, alkyl group having 1 to 8 carbon atoms, alkoxy group having 1 to 8 carbon atoms, aryloxy group having 6 to 18 carbon atoms or a halogen atom; and M, R¹, R², R³, L and n have the same meanings as in [1] above).

[3] The carboxylate metal complex as described in [2] above, wherein M in formula (C2) is Ni or Pd.

[4] The carboxylate metal complex as described in [2] or [3] above, wherein R¹ in formula (C2) is hydrocarbon group having 1 to 30 carbon atoms which may be substituted by a halogen atom.

[5] The carboxylate metal complex as described in [4] above, wherein R¹ in formula (C2) is methyl group or trifluoromethyl group.

[6] The carboxylate metal complex as described in [5] above, wherein R¹ in formula (C2) is methyl group.

[7] The carboxylate metal complex as described in any one of [2] to [6] above, wherein both of R² and R³ in formula (C2) are isopropyl group or cyclohexyl group.

[8] The carboxylate metal complex as described in any one of [2] to [7] above, wherein all of R⁴s in formula (C2) are a hydrogen atom.

[9] The carboxylate metal complex as described in any one of [2] to [8] above, wherein L in formula (C2) is 2,6-

dimethylpyridine and n is 1.

[10] A catalyst for olefin polymerization mainly comprising the carboxylate metal complex described in any one of [1] to [9] above.

5 [11] A method for producing a polymer characterized in polymerizing olefin represented by formula (1) singly or in combination of two or more kinds thereof, or copolymerizing olefin represented by formula (1) and polar group-containing olefin represented by formula (2)



(In the formula, R^5 represents a hydrogen atom or hydrocarbon group having 1 to 20 carbon atoms)



(In the formula, R^6 represents a hydrogen atom or methyl group,
15 R^7 represents $-\text{COOR}^8$, $-\text{CN}$, $-\text{OCOR}^8$, $-\text{CH}_2-\text{OCOR}^8$, $-\text{CH}_2\text{OH}$, $-\text{CH}_2-\text{N}(\text{R}^9)_2$ or $-\text{CH}_2-\text{Hal}$ (R^8 represents a hydrogen atom or hydrocarbon group having 1 to 5 carbon atoms, R^9 represents a hydrogen atom, hydrocarbon group having 1 to 5 carbon atoms or aromatic substituent having 6 to 18 carbon atoms or alkoxycarbonyl group,
20 and Hal represents a halogen atom).

EFFECTS OF THE INVENTION

[0009]

The novel metal complex of group 10 elements of the
25 periodic table of the present invention, in which all of the coordinating atoms are a hetero atom, is stable and useful as a catalyst component for olefin polymerization, and can be used for olefin homopolymerization or copolymerization of two or more kinds of olefin for a long period of time.

BRIEF DESCRIPTION OF DRAWINGS

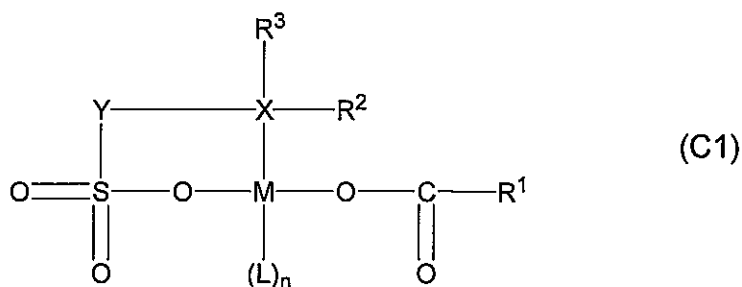
[0010] Fig. 1 shows an X-ray structure analysis of metal complex 1 prepared in Example 1.

Fig. 2 shows an X-ray structure analysis of metal
5 complex 2 prepared in Example 2.

MODE FOR CARRYING OUT THE INVENTION

[0011] [Metal Complex]

The metal complex of the present invention is
10 represented by formula (C1).



[0012]

In the formula, M represents a metal atom of group
15 10 element in the periodic table; X represents a phosphorous
atom (P) or an arsenic atom (As); R¹ represents a hydrogen atom
or hydrocarbon group having 1 to 30 carbon atoms which may be
substituted by one or more groups selected from a halogen atom,
alkoxy group, aryloxy group, acyloxy group and amino group; R²
20 and R³ independently represent a hydrogen atom, alkoxy group,
aryloxy group, acyloxy group, silyl group, amino group or
hydrocarbon group having 1 to 30 carbon atoms which may be
substituted by one or more groups selected from a halogen atom,
alkoxy group, aryloxy group, acyloxy group and amino group; R²
25 and R³ may bond to each other to form a ring structure; Y
represents divalent hydrocarbon group having 1 to 70 carbon

atoms which may be substituted by one or more functional groups selected from a halogen atom, alkoxy group, aryloxy group, acyloxy group and amino group; L represents a neutral electron-donating ligand and n is 0, 1/2, 1 or 2.

5 In the present specification, "hydrocarbon" includes saturated and unsaturated aliphatic hydrocarbon, alicyclic hydrocarbon and aromatic hydrocarbon.

[0013]

 The structure of the complex represented by formula
10 (C1) is described below.

M represents an element of group 10 in the periodic table. Examples of the element of group 10 in the periodic table include nickel (Ni), palladium (Pd) and platinum (Pt), Ni and Pd are preferable and Pd is more preferable.

15 [0014]

 Single electron of the high-electronegative oxygen atom of $Y-S(=O)_2-O-$ moiety coordinates to metal atom M. Since the bonding electron between $Y-S(=O)_2-O-M$ is transferred from M to $Y-S(=O)_2-O-$, $Y-S(=O)_2-O-$ and M may be indicated formally as
20 an anion state and a cation state, respectively.

[0015]

R^1 represents a hydrogen atom or hydrocarbon group having 1 to 30 carbon atoms which may be substituted by one or more groups selected from a halogen atom, alkoxy group, aryloxy group, acyloxy group and amino group. Examples of the
25 hydrocarbon group having 1 to 30 carbon atoms which may be substituted by one or more groups selected from a halogen atom, alkoxy group, aryloxy group, acyloxy group and amino group include alkyl group, aryl group and cycloalkyl group. A
30 preferred halogen atom in the hydrocarbon group having 1 to 30

carbon atoms which may be substituted by one or more groups selected from a halogen atom, alkoxy group, aryloxy group, acyloxy group and amino group is fluorine. Preferred alkoxy group is the one having 1 to 10 carbon atoms, including methoxy group, ethoxy group, propoxy group and isopropoxy group.

Preferred aryloxy group is the one having 6 to 24 carbon atoms, including phenoxy group. Preferred acyloxy group is the one having 1 to 10 carbon atoms, including acetoxy group and pivaloxy group. Preferred amino group is the one having 1 to 10 carbon atoms, including dimethylamino group, diethylamino group and diisopropylamino group. As R^1 , methyl group, trifluoromethyl group and phenyl group are more preferable due to the easy availability and from a cost point of view.

[0016]

L represents a neutral electron-donating ligand. Neutral electron-donating ligand L is a compound which contains an electron-donating group and is capable of stabilizing a metal complex by coordinating to metal atom M as a two-electron ligand to thereby isolate a metal complex.

Examples of the neutral electron-donating ligand include those containing a sulfur atom such as dimethyl sulfoxide (DMSO). Examples of those containing a nitrogen atom include trialkyl amine having 1 to 10 carbon atoms, dialkyl amine, pyridine, 2,6-dimethylpyridine, quinoline, 2-methylquinoline, aniline, 2,6-dimethylaniline, 2,6-diisopropylaniline, N,N,N',N'-tetramethylethylenediamine (TMEDA), 4-(N,N-dimethylamino)pyridine (DMAP), acetonitrile and benzonitrile. Examples of those containing a phosphorous atom include trimethylphosphine, triisopropylphosphine, tributylphosphine, tri(t-butyl)phosphine, triphenylphosphine,

tris(o-tolyl)phosphine and trifurylphosphine. Examples of those containing an oxygen atom include diethyl ether, tetrahydrofuran, 1,4-dioxane and 1,2-dimethoxyethane. From the viewpoint of a synthesis method, electron-donating ligands
5 containing a nitrogen atom and those containing a phosphorous atom are preferable, and those containing a nitrogen atom are particularly preferable.

[0017]

n represents the coordination number of the neutral
10 electron-donating ligand L to metal atom M and is 0, 1/2, 1 or 2. n of 1/2 means that one bivalent electron-donating ligand coordinates to two metal complexes. n is preferably 1/2 or 1 to stabilize a metal complex. n of 0 means that there is no ligand.

15 [0018]

X represents a phosphorous atom (P) or an arsenic atom (As), wherein two electrons coordinate to M. P is preferred as X for reasons of availability and the catalyst cost.

20 [0019]

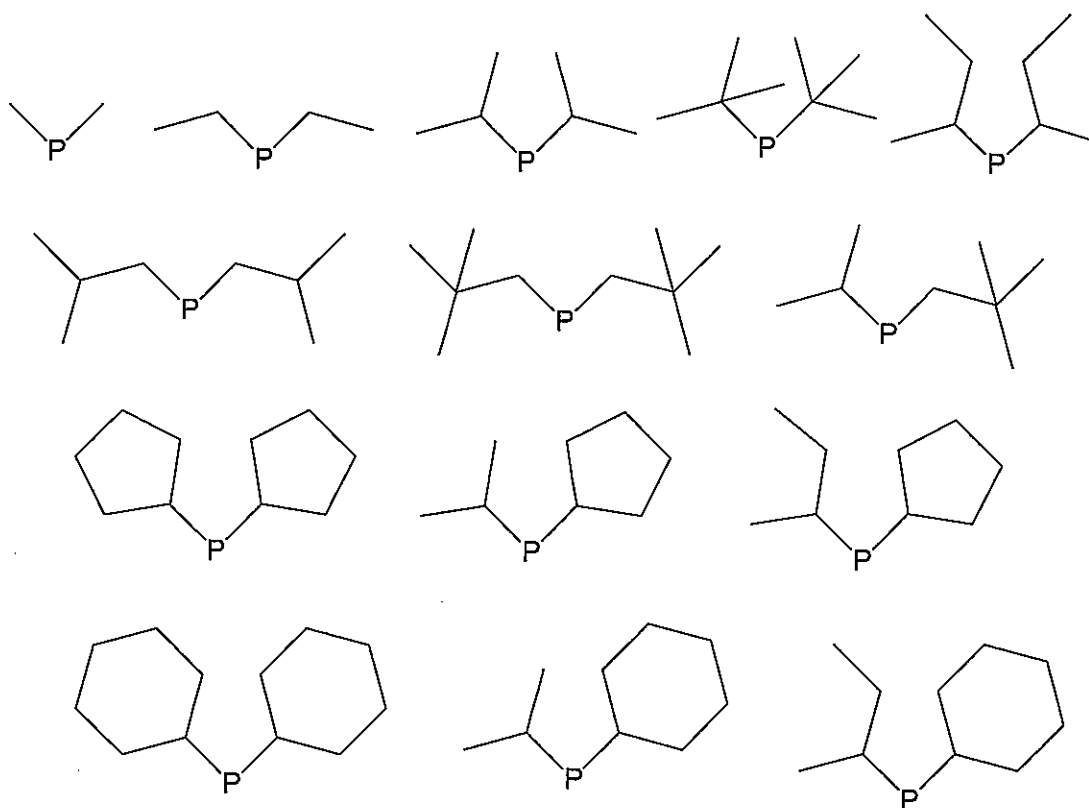
R² and R³ independently represent a hydrogen atom, alkoxy group, aryloxy group, acyloxy group, silyl group, amino group, or a hydrocarbon group having 1 to 30 carbon atoms which may be substituted by one or more groups selected from a halogen
25 atom, alkoxy group, aryloxy group, acyloxy group and amino group. Preferred alkoxy group is the one having 1 to 20 carbon atoms, including methoxy group, ethoxy group, propoxy group and isopropoxy group. Preferred aryloxy group is the one having 6 to 24 carbon atoms, including phenoxy group. Examples of silyl
30 group includes trimethylsilyl group and examples of amino group

include amino group, methylamino group and dimethylamino group.
R² and R³ may be the same or different from each other. R² and
R³ may bond to each other to form a ring structure. Examples of
the hydrocarbon group having 1 to 30 carbon atoms represented by
5 R² and R³ which may be substituted by one or more groups
selected from a halogen atom, alkoxy group, aryloxy group,
acyloxy group and amino group include alkyl group, aryl group,
cycloalkyl group and furyl group. Specific examples of the
halogen atom, alkoxy group, aryloxy group, acyloxy group and
10 amino group in the hydrocarbon group having 1 to 30 carbon atoms
which may be substituted by one or more groups selected from a
halogen atom, alkoxy group, aryloxy group, acyloxy group and
amino group include the same as those in the explanation of R¹.
As R² and R³, alkyl group, cycloalkyl group and aryl group are
15 preferable, and isopropyl group and cyclohexyl group are most
preferable.

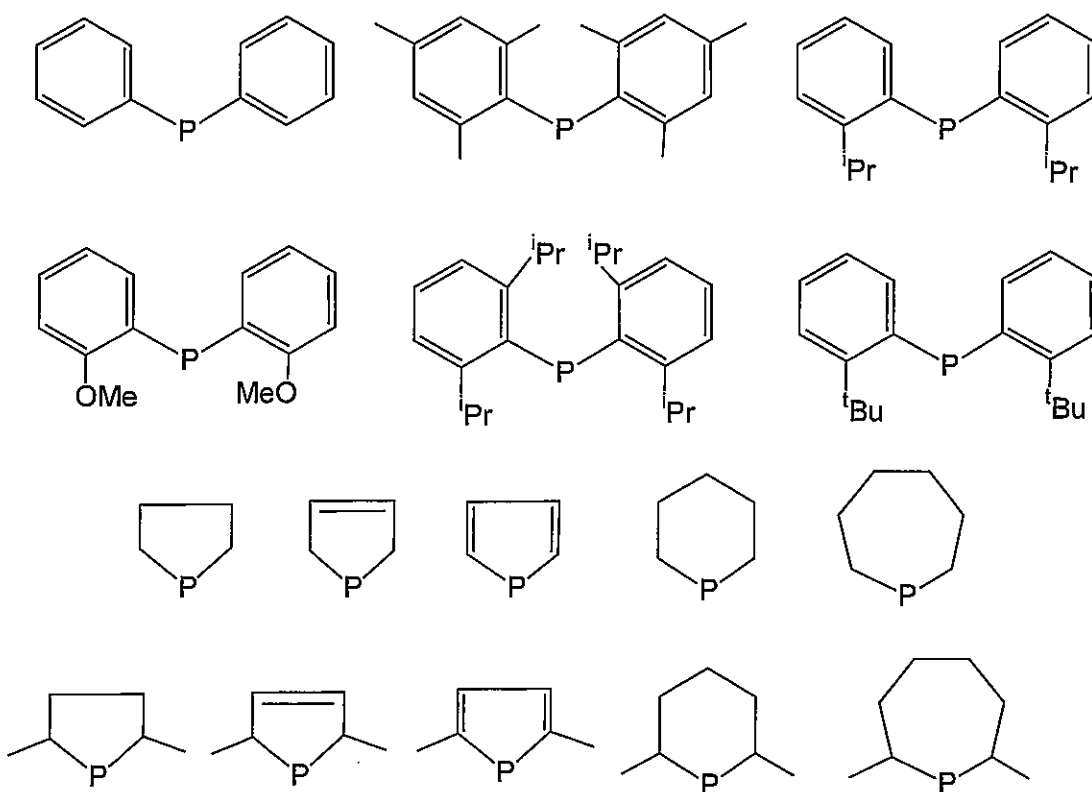
[0020]

Examples of [(R²)(R³)P] moiety when X is P (a
phosphorous atom) include the following structures. In the
20 following structure formulae, the bonds between P and M or Y are
not shown. In the structure formula, "Me" represents methyl
group, "iPr" represents isopropyl group, and "tBu" represents t-
butyl group.

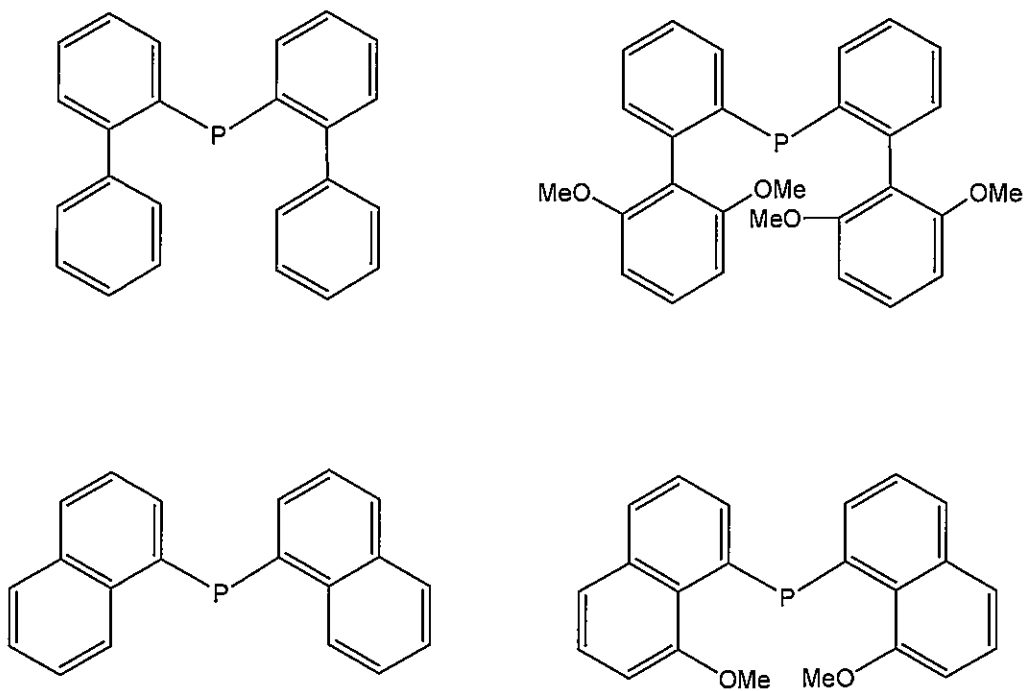
[Chem. 6]



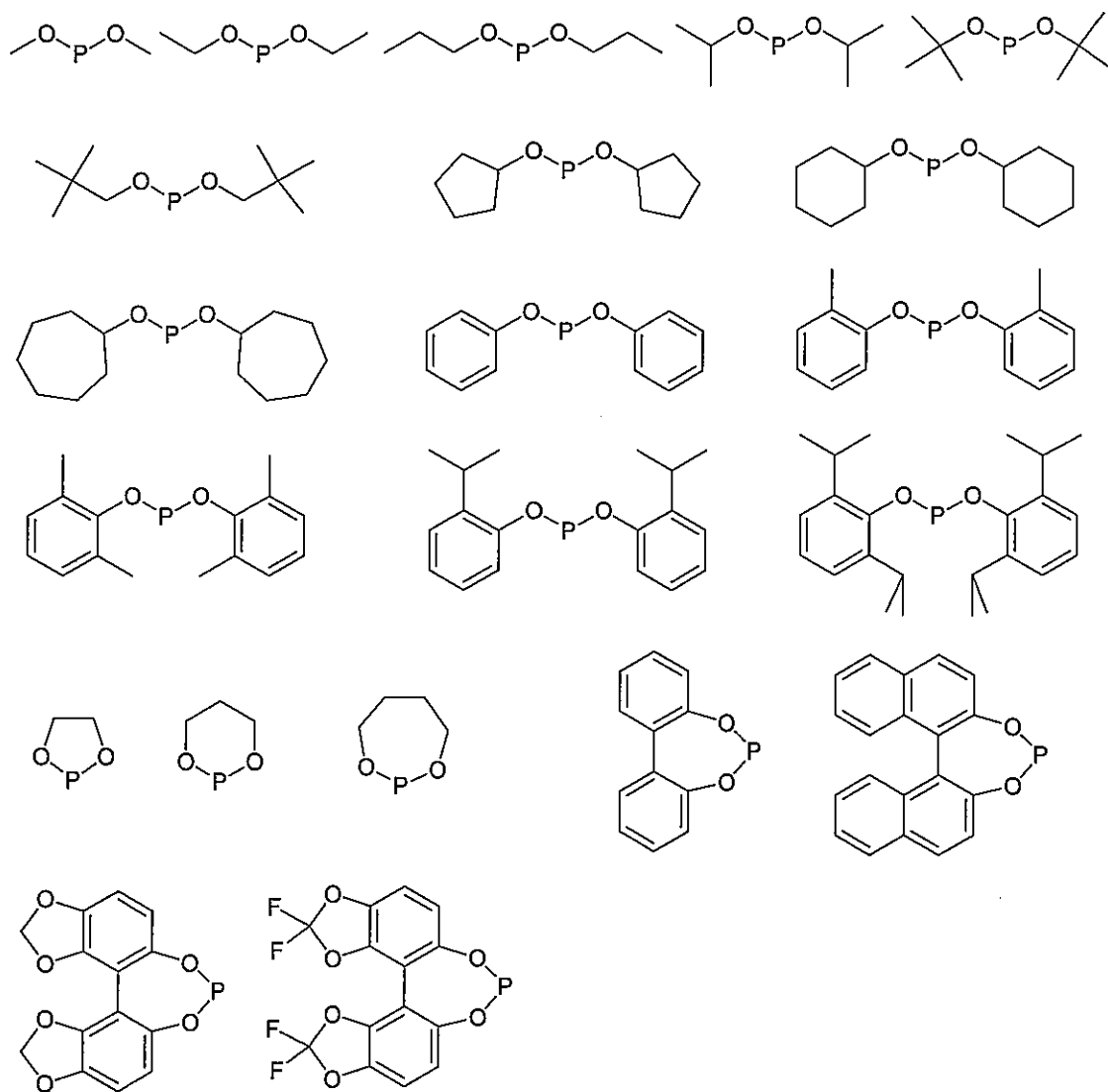
[Chem. 7]



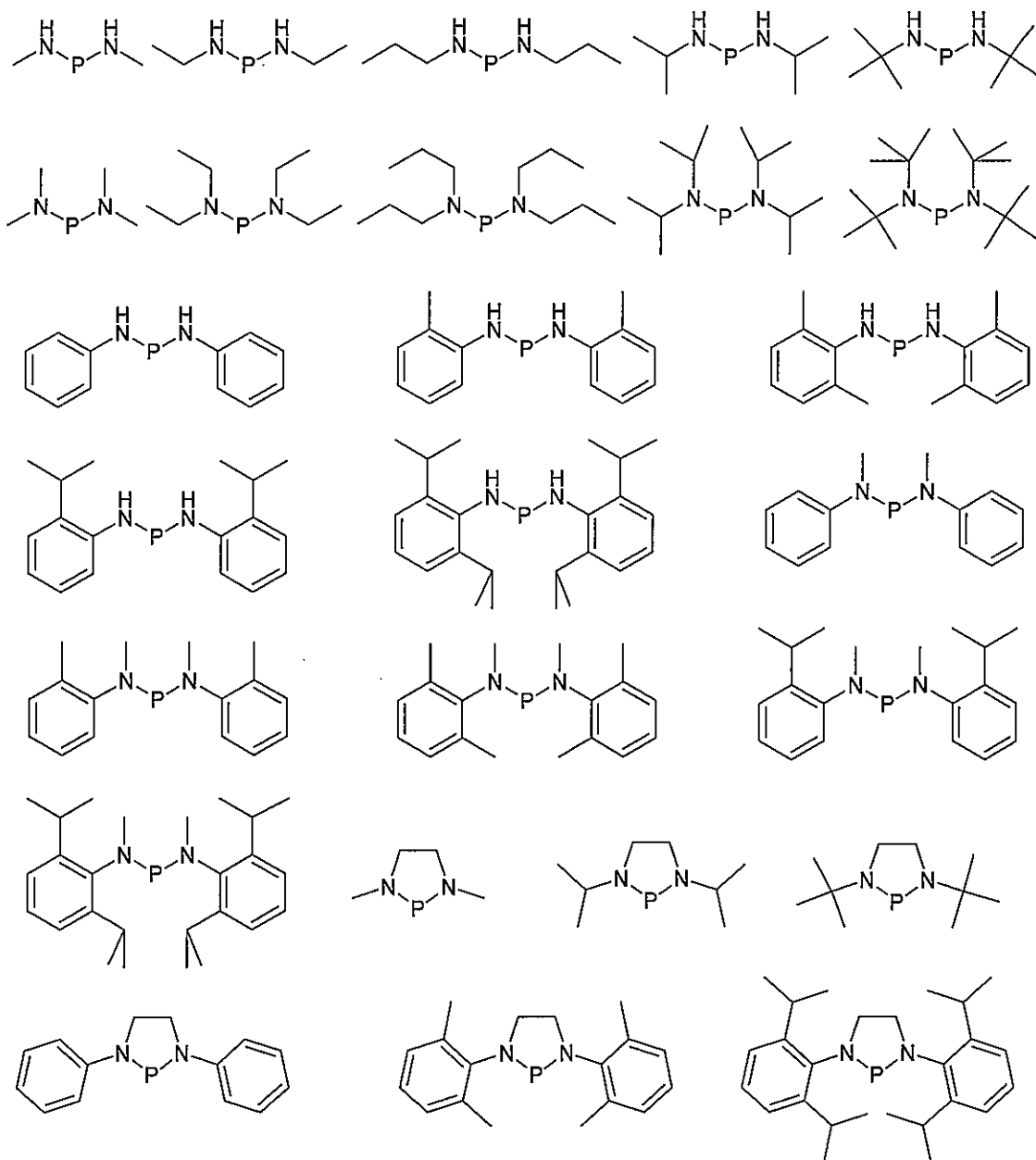
[Chem. 8]



[Chem. 9]



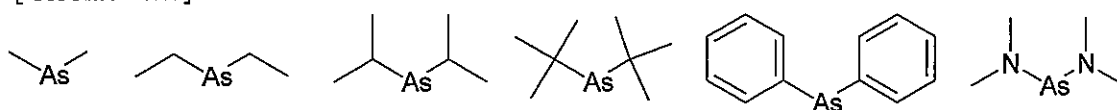
[Chem. 10]



[0021]

Examples of $[(R^2)(R^3)As]$ moiety when X is As (an arsenic atom) include the following structures. In the following structure formulae, the bonds between As and M or Y are not shown.

[Chem. 11]



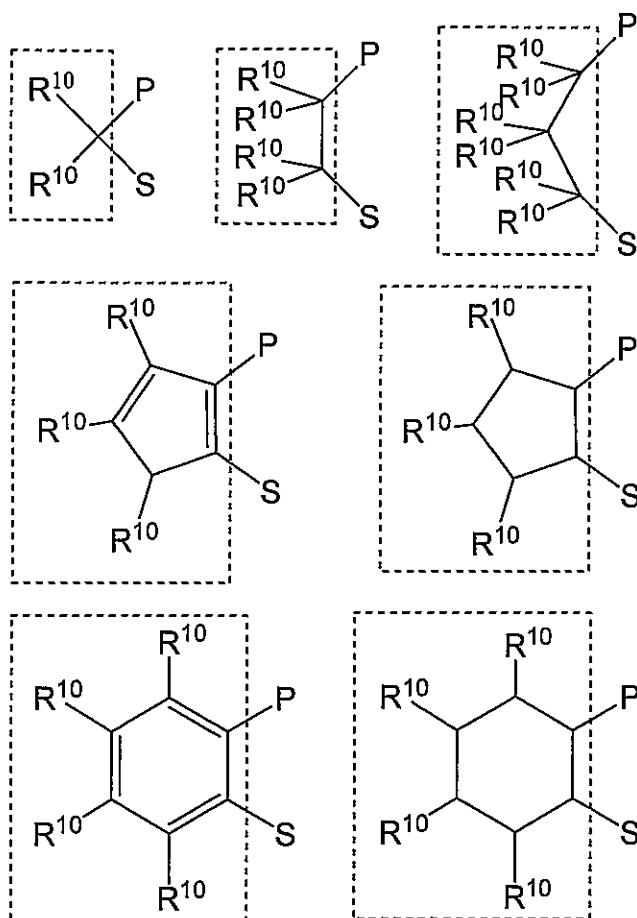
[0022]

Y represents a bivalent hydrocarbon group having 1 to 70 carbon atoms (crosslinked structure) which may be substituted by one or more functional groups selected from a halogen atom, alkoxy group, aryloxy group, acyloxy group and amino group. Examples of the bivalent hydrocarbon group having 1 to 70 carbon atoms which may be substituted by one or more functional groups selected from a halogen atom, alkoxy group, aryloxy group, acyloxy group and amino group include alkylene group and arylene group. Arylene group is particularly preferable.

[0023]

The cross-linked structure Y is the crosslinking moiety which binds X and a sulfur atom (S). Specific examples of the cross-linked structure Y in which X is represented by a P atom are shown below. Here, multiple R^{10} s may be the same or different to each other and represent a hydrogen atom, halogen atom, hydrocarbon group having 1 to 20 carbon atoms, or a hydrocarbon group having 1 to 20 carbon atoms substituted by a halogen atom.

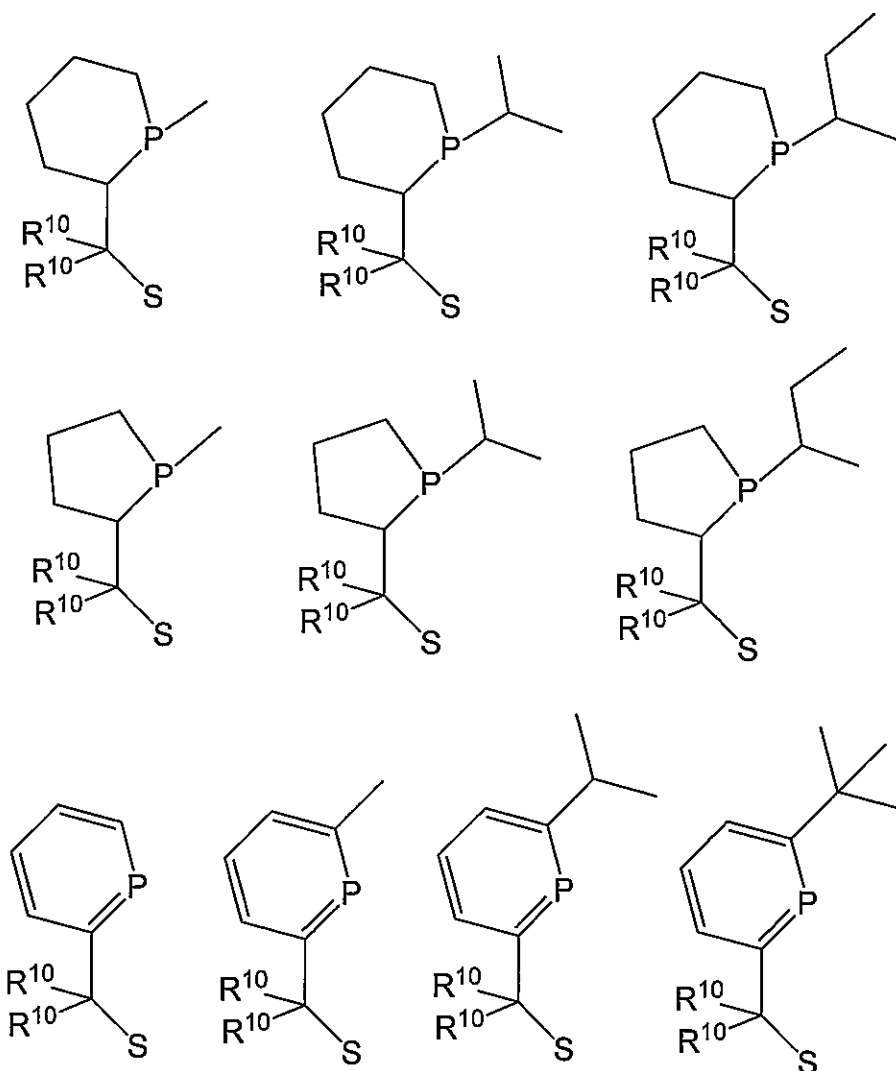
[Chem. 12]



[0024]

Substituents R^2 and R^3 may bond to Y moiety to form a
5 ring structure. Specific examples include the structures as
follows:

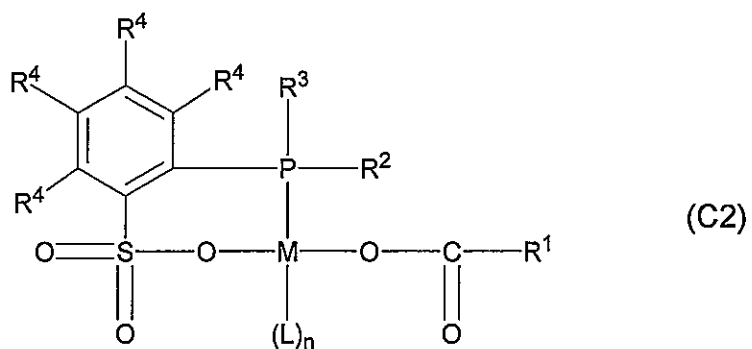
[Chem. 13]



[0025]

5 Among the metal complexes represented by formula (C1), those represented by the following formula (C2) are particularly preferable.

[Chem. 14]



[0026]

In the formula, four R⁴s independently represent a hydrogen atom, alkyl group having 1 to 8 carbon atoms, alkoxy group having 1 to 8 carbon atoms, aryloxy group having 6 to 18 carbon atoms or halogen atom; and M, R¹, R², R³, L and n have the same meanings as those in formula (C1). Examples of the alkyl group having 1 to 8 carbon atoms include methyl group, ethyl group, propyl group and octyl group. Examples of the alkoxy group having 1 to 8 carbon atoms include methoxy group, ethoxy group, propoxy group and isopropoxy group. Examples of the aryloxy group having 6 to 18 carbon atoms include phenoxy group. Examples of the halogen atom include chlorine and fluorine.

[0027]

In formula (C2), preferred R¹ is a methyl group. Both of R² and R³ are preferably a cyclohexyl group or isopropyl group. Each of the four R⁴s is preferably a hydrogen atom. L is preferably dimethyl sulfoxide, pyridine, 2,6-dimethylpyridine (lutidine), quinoline, 2-methylquinoline or N,N,N',N'-tetramethylethylenediamine, and 2,6-dimethylpyridine (lutidine) is particularly preferable. n is preferably 1/2 or 1, and 1 is particularly preferable. M is preferably Ni or Pd, and Pd is particularly preferable.

[0028]

There is no particular limitation on the method for synthesizing metal complexes (C1) and (C2) of the present invention and the complexes can be synthesized according to a known method or a combination thereof. That is, the complexes can be synthesized by reacting zerovalent or bivalent M source, a ligand in formula (C1) or (C2) and a neutral electron-donating

ligand L.

[0029]

Examples of zerovalent M source include tris(dibenzylidene acetone)dipalladium and tetrakis(triphenylphosphine)palladium as a palladium source and tetracarbonyl nickel, bis(1,5-cyclooctadiene)nickel and tetrakis(triphenylphosphine)nickel as a nickel source.

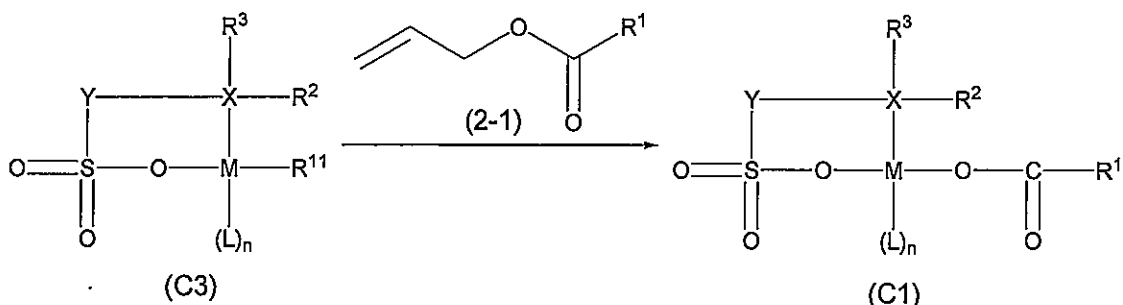
[0030]

Examples of bivalent M source include (1,5-cyclooctadiene)palladium chloride, (1,5-cyclooctadiene)(methyl)palladium chloride, (allyl)palladium chloride, palladium chloride, palladium bromide, palladium acetate, palladium nitrate, bis(acetonitrile)dichloropalladium, bis(benzonitrile)dichloropalladium, (N,N,N',N'-tetramethylethylenediamine)dichloro palladium(II), (N,N,N',N'-tetramethylethylenediamine)dimethyl palladium(II), bis(acetylacetonato)palladium(II), palladium(II) trifluoromethanesulfonate as a palladium source; and (allyl)nickel chloride, (allyl)nickel bromide, nickel chloride, nickel acetate, bis(acetylacetonato)nickel(II), (1,2-dimethoxyethane)dichloronickel(II), (1,2-dimethoxyethane)dibromonickel(II) and nickel(II) trifluoromethanesulfonate as a nickel source.

[0031]

Metal complexes represented by formula (C1) and formula (C2) can also be synthesized by reacting an M-hydride complex or an M-alkyl complex (the two are collectively referred to as "(C3)") corresponding to the metal complex with allyl ester (2-1).

[Chem. 15]

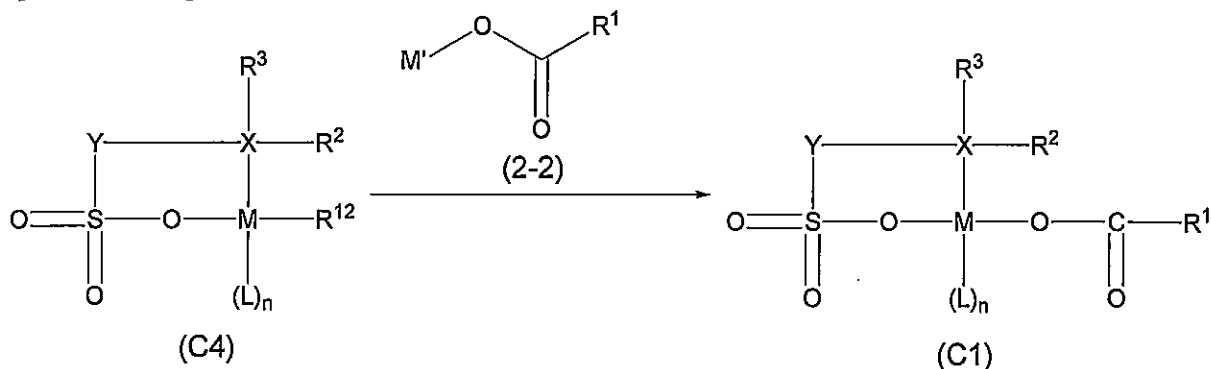


In the formula, R^{11} represents a hydrogen atom or alkyl group having 1 to 6 carbon atoms, and R^1 , R^2 , R^3 , X , Y , M , L and n have the same meanings as described above. While the above formula indicates the method for synthesizing metal complex (C1), metal complex (C2) can be synthesized in the same way.

[0032]

The metal complexes represented by formula (C1) and formula (C2) can also be synthesized by reacting an M-halogen complex (C4) corresponding to the metal complex with carboxylate metal salt (2-2).

[Chem. 16]



In the formula, R^{12} represents a halogen atom, M' represents a metal atom, and R^1 , R^2 , R^3 , X , Y , M , L and n have the same meanings as described above. While the above formula indicates the method for synthesizing metal complex (C1), metal complex (C2) can be synthesized in the same way.

Examples of the halogen atom represented by R^{12}

include a chlorine atom, a bromine atom or an iodine atom, and a chlorine atom or a bromine atom is preferable from a cost point of view.

[0033]

5 Specific examples of carboxylate metal salts (2-2) which can be used in the synthesis method of the present invention include silver acetate, silver trifluoroacetate, silver propionate, silver 2-ethylhexanoate, lithium acetate, lithium trifluoroacetate, lithium propionate, lithium 2-ethylhexanoate, sodium acetate, sodium trifluoroacetate, sodium propionate, sodium 2-ethylhexanoate, potassium acetate, potassium trifluoroacetate, potassium propionate and potassium 2-ethylhexanoate.

[0034]

15 The metal complex represented by formula (C1) or (C2) may be supported on a support to be used for polymerization. In this case, there are no particular limitations on the support and examples include an inorganic support such as silica gel and alumina and an organic support such as polystyrene, polyethylene and polypropylene. Examples of the method for depositing a metal complex on a support include a physical adsorption method of impregnating the support with a solution of the metal complex and drying it and a method of depositing the metal complex onto a support by chemically bonding the metal complex to a support.

[0035] [Monomer]

Olefin, which is the first monomer used in the method for producing the polymer of the present invention, is represented by formula (1).

30 [Chem. 17]



[0036]

In formula (1), R^5 represents a hydrogen atom or hydrocarbon group having 1 to 20 carbon atoms and is preferably
 5 a hydrogen atom, alkyl group having 1 to 3 carbon atoms or aryl group having 6 to 20 carbon atoms. Specifically, examples of olefin of formula (1) include ethylene, propylene, 1-butene, 1-hexene, 4-methyl-1-pentene, 1-octene and styrene. Among these, ethylene and propylene are particularly preferable. One of
 10 these compounds may be used independently or two or more of them may be used in combination.

[0037]

A polar group-containing olefin, which is the second monomer used for polymerization in the present invention, is
 15 represented by formula (2).

[Chem. 18]



[0038]

In formula (2), R^6 represents a hydrogen atom or
 20 methyl group. R^7 represents $-\text{COOR}^8$, $-\text{CN}$, $-\text{OCOR}^8$, $-\text{CH}_2-\text{OCOR}^8$, $-\text{CH}_2\text{OH}$, $-\text{CH}_2-\text{N}(\text{R}^9)_2$ or $-\text{CH}_2-\text{Hal}$ (R^8 represents a hydrogen atom, hydrocarbon group having 1 to 5 carbon atoms; R^9 represents a hydrogen atom, hydrocarbon group having 1 to 5 carbon atoms, aromatic substituent having 6 to 18 carbon atoms or
 25 alkoxy carbonyl group; and Hal represents a halogen atom). R^8 is preferably a hydrogen atom or alkyl group having 1 to 3 carbon atoms, and methyl group is particularly preferable. R^9 is preferably a hydrogen atom, alkyl group having 1 to 3 carbon atoms, phenyl group, benzyl group, t-butoxycarbonyl group or

benzyloxycarbonyl group. The halogen atom of Hal is preferably a chlorine atom or a bromine atom.

[0039]

Specific examples of the polar group-containing vinyl compound represented by formula (2) include vinyl acetate, allyl acetate, allyl alcohol, methyl methacrylate, methacrylic acid, methyl acrylate, acrylic acid, acrylonitrile, allyl amine, N-allylaniline, N-t-butoxycarbonyl-N-allylamine, N-benzyloxycarbonyl-N-allylamine, N-benzyl-N-allylamine, allyl chloride and allyl bromide. Among these, vinyl acetate, allyl acetate, methyl methacrylate, methyl acrylate and acrylonitrile are particularly preferable. One of these compounds may be used independently or two or more of them may be used in combination.

[0040]

Examples of the combination of olefin represented by formula (1) and the polar group-containing olefin represented by formula (2) include ethylene and vinyl acetate, ethylene and allyl acetate, ethylene and allyl alcohol, ethylene and methyl methacrylate, ethylene and methacrylic acid, ethylene and methyl acrylate, ethylene and acrylic acid, ethylene and acrylonitrile, ethylene and allylamine, ethylene and N-allylaniline, ethylene and N-t-butoxycarbonyl-N-allylamine, ethylene and N-benzyloxycarbonyl-N-allylamine, ethylene and N-benzyl-N-allylamine, ethylene and allyl chloride, ethylene and allyl bromide, propylene and vinyl acetate, propylene and allyl acetate, propylene and allyl alcohol, propylene and methyl methacrylate, propylene and methyl acrylate, propylene and acrylic acid, propylene and acrylonitrile, propylene and allylamine, propylene and N-allylaniline, propylene and N-t-butoxycarbonyl-N-allylamine, propylene and N-benzyloxycarbonyl-

N-allylamine, propylene and N-benzyl-N-allylamine, propylene and allyl chloride and propylene and allyl bromide. Among these, preferred are ethylene and vinyl acetate, ethylene and allyl acetate, ethylene and allyl alcohol, ethylene and methyl methacrylate, ethylene and methyl acrylate, ethylene and acrylonitrile, ethylene and allyl chloride, and ethylene and allylamine from the viewpoint of the polymer performance and economic efficiency.

[0041]

In the method for producing the (co)polymer of the present invention, in addition to the monomers represented by formulae (1) and (2), one or more kinds of a third monomer may be incorporated to be copolymerized.

The third monomers include ethyl vinyl ether, propyl vinyl ether, norbornene and carbon monoxide.

[0042]

[Polymerization method]

There are no particular limitations on the method of polymerizing monomers and the monomers can be polymerized by a widely-used method. That is, a process such as a solution polymerization method, a suspension polymerization method and a gas-phase polymerization method is available. Particularly preferred are a solution polymerization method and a suspension polymerization method. The polymerization mode may employ either of a batch mode and a continuous mode. Also, either of a single stage polymerization or a multistage polymerization can be employed.

[0043]

The molar ratio between the metal complex catalyst represented by formula (C1) or (C2) and the total amount of

monomers (monomers/metal complex) is within the range of from 1 to 10,000,000, preferably the range of from 10 to 1,000,000, more preferably the range of from 100 to 100,000.

[0044]

5 The blending ratio of the monomers represented by formulae (1) and (2) is to be appropriately adjusted depending on the composition ratio of the targeted copolymer. This includes a case where a third monomer is used.

[0045]

10 There are no particular limitations on the polymerization temperature. The polymerization is generally conducted at a temperature in the range of from -30 to 200 °C, preferably in the range of from 0 to 180 °C, more preferably in the range of from 20 to 150 °C.

15 The polymerization is conducted at a polymerization pressure, wherein the internal pressure consists mostly of the pressure of olefin represented by formula (1), in the range from normal pressure to 20 MPa, preferably in the range from normal pressure to 10 MPa.

20 [0046]

 The polymerization time can be appropriately adjusted depending on the processing mode and the polymerization activity of the catalyst, and can be as short as several minutes or as long as several thousand hours.

25 [0047]

 It is preferable to fill the atmosphere in the polymerization system with an inert gas such as nitrogen and argon to prevent components other than monomers to be polymerized such as air, oxygen and moisture being mixed into
30 the atmosphere to retain the catalyst activity. In the case of

the solution polymerization, an inert solvent may be used in addition to monomers to be polymerized. There are no particular limitations on the inert solvent, and examples include aliphatic hydrocarbon such as pentane, hexane and heptane; alicyclic hydrocarbon such as cyclopentane, cyclohexane and cycloheptane; aromatic hydrocarbon such as benzene, toluene and xylene; halogenated aliphatic hydrocarbon such as chloroform, methylene chloride, carbon tetrachloride, dichloroethane and tetrachloroethane; halogenated aromatic hydrocarbon such as chlorobenzene, dichlorobenzene and trichlorobenzene; aliphatic ester such as methyl acetate and ethyl acetate; and aromatic ester such as methyl benzoate and ethyl benzoate.

[0048]

After completion of the polymerization reaction, the (co)polymer as a reaction product is to be isolated by post-treatment using a known operation and treating method (e.g. neutralization, extraction with solvents, washing with water, liquid separation, distillation with solvents and reprecipitation).

The (co)polymer can be molded into the form of a pellet, film, sheet and the like under conditions for general thermoplastic resin.

[0049]

The obtained (co)polymer can be a product per se by molding such as injection molding, extrusion and film processing. Or the (co)polymer may be added to polyolefin and the like to be used as a modifier of the surface features such as adhesiveness and printing performance; a compatibility agent between nonpolar polyolefin and highly-polar other resins; and a dispersing agent of pigments and the like. The copolymer may

also be used for purposes such as paint, ink, adhesive agent, binder, plasticizer, lubricant, lubricant oil and surface active agent.

5 EXAMPLES

[0050]

Hereinafter, the present invention is described in greater detail by referring to Examples and Comparative Examples. The present invention is by no means limited thereto.

10 [0051] [Method for analyzing a metal complex compound]

The structures of Metal complex 1 and Metal complex 2 obtained in the examples were determined by X-ray crystal structure analysis using an single crystal X-ray diffractometer (Mercury CCD) manufactured by Rigaku Corporation. Mo-K α radiation ($\lambda=0.71069$ Å) was used as X-ray source, and radiation within the range of $2\theta \leq 50^\circ$ was used. Diffraction intensity was corrected using Lorentz factor. The structure was elucidated by the direct method via SIR-97, and optimized using SHELXL-97. The atoms other than hydrogen were anisotropically optimized and the hydrogen atoms were arranged using the AFIX instruction.

[0052]

[Method for analyzing the polymer structure]

The structure of the copolymers obtained in Examples was determined by various analysis of the NMR spectra using JNM-ECS400 manufactured by JEOL Ltd. The content of the monomer unit derived from the polar group-containing olefin represented by formula (2) and the terminal structure of the copolymer were determined by analyzing ^{13}C -NMR spectrum (90° pulse at 9.0 microseconds, spectrum width: 31 kHz, relaxation time: 10 seconds, acquisition time: 10 seconds, times of accumulating FID

signals: 5,000 to 10,000 times) through the inverse-gated decoupling method at 120°C using 1,2,4-trichlorobenzene (0.55 ml) as a solvent and Cr(acac)₃ (10 mg) as relaxation agent.

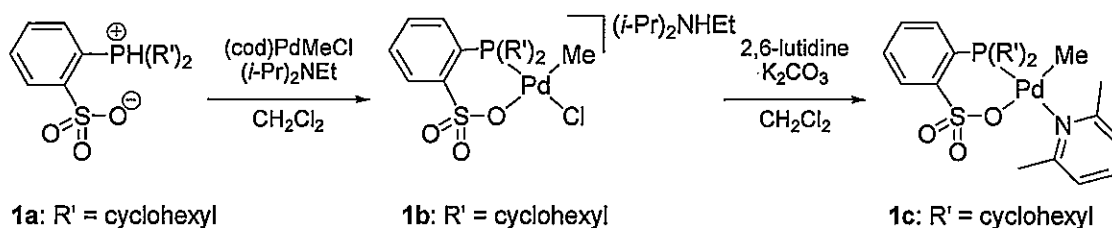
[0053]

5 A number average molecular weight and a weight average molecular weight were calculated by size exclusion chromatography in which polystyrene is employed as a standard substance using a high-temperature GPC apparatus, HLC-8121GPC/HT, manufactured by Tosoh Corporation, provided with
10 TSKgel GMHHR-H(S) HT column (two columns of 7.8 mm I.D. x 30 cm arranged in series) manufactured by Tosoh Corporation (solvent: 1,2-dichlorobenzene, temperature: 145°C).

[0054] Metal complex 1c:

Metal complex 1c was synthesized according to the
15 following reaction scheme:

[Chem. 19]



[0055]

(a) Synthesis of compound 1a

20 n-Butyllithium (manufactured by Kanto Chemical Co., Inc., 1.65 M hexane solution, 5.1 ml, 8.4 mmol) was added to a tetrahydrofuran (THF) solution (20 ml) of benzenesulfonic acid (manufactured by Tokyo Chemical Industry Co., Ltd., 0.66 g, 4.2 mmol) at 0°C under argon atmosphere and the mixture was stirred
25 for 2.5 hours at room temperature. After cooling the reaction container to -78°C, chlorodicyclohexylphosphine (manufactured by Sigma-Aldrich, 0.89 mg, 3.8 mmol) was added thereto at -78°C and

stirred for 24 hours at room temperature. After ceasing the reaction with trifluoroacetic acid (manufactured by Tokyo Chemical Industry Co., Ltd., 0.50 M THF solution, 8.4 ml, 4.2 mmol), the precipitate formed was collected by filtration and
5 dried under reduced pressure to obtain 2-(dicyclohexylphosphino)benzenesulfonic acid (compound 1a). The yield was 0.66 g (85%).

[0056]

$^1\text{H-NMR}$ (400MHz, CDCl_3): δ 0.98-0.27 (m, 4H), 1.30-1.58
10 (m, 6H), 1.62-1.78 (m, 4H), 1.88 (br s, 4H), 2.28 (br s, 2H), 3.33 (br s, 2H), 5.19 (br d, $^1J_{\text{PH}} = 370$ Hz, 1H), 7.48-7.58 (m, 2H), 7.80 (br s, 1H), 8.27 (br s, 1H);

$^{13}\text{C-NMR}$ (101MHz, CDCl_3): δ 25.0 (s), 25.6-26.2 (m), 28.8 (br), 30.3 (br), 34.6 (br d, $^1J_{\text{PC}} = 40$ Hz), 113.4 (br d, $^1J_{\text{PC}} = 87$ Hz), 128.8
15 (d, $J_{\text{PC}} = 9$ Hz), 130.1 (d, $J_{\text{PC}} = 9$ Hz), 135.4 (br), 137.1 (br), 150.5 (br);

$^{31}\text{P-NMR}$ (162MHz, CDCl_3): δ 52.8 (d, $^1J_{\text{PH}} = 370$ Hz) (90%), 20.8 (d, $^1J_{\text{PH}} = 530$ Hz) (10%);

Anal. calcd for $\text{C}_{18}\text{H}_{27}\text{O}_3\text{PS}$, C, 60.99; H, 7.68.

20 found: C, 60.90; H, 7.55.

[0057]

(b) Synthesis of metal complex 1b

A methylene chloride solution (6 ml) of (COD)PdMeCl (synthesized according to "Inor. Chem., 1993, 32, 5769-5778;
25 COD: 1,5-cyclooctadiene; 0.32g; 1.2 mmol) was added to a methylene chloride solution (16 ml) of 2-(dicyclohexylphosphino)benzenesulfonic acid; compound 1a) (0.43 g, 1.2 mmol) and diisopropylethylamine (manufactured by Wako Chemical Pure Chemical Industries Ltd.; 1.1 ml, 6.0 mmol) under
30 argon atmosphere and the mixture was stirred for one hour at

room temperature. After condensing the solvent, the precipitation was removed by filtration and the filtrate was added to hexane. The precipitate formed was collected by filtration, washed with hexane and then dried under reduced
5 pressure to obtain metal complex 1b. The yield was 0.66g (85%).
[0058]

^1H -NMR (500MHz, CDCl_3): δ 0.71 (d, $^3J_{\text{PH}} = 1.4$ Hz, 3H, PdCH_3), 1.11-1.35 (m, 8H), 1.45 (d, $J = 6.6$ Hz, 6H, $\text{HNCH}(\text{CH}_3)_2$), 1.57 (d, $J = 6.6$ Hz, 6H, $\text{HNCH}(\text{CH}_3)_2$), 1.57 (t, $J = 7.3$ Hz, 3H, HNCH_2CH_3), 1.60-1.70 (m, 6H), 1.72-1.84 (m, 4H), 2.12-2.28 (m, 4H), 3.29 (dq, $J = 7.3, 5.0$ Hz, 2H, HNCH_2CH_3), 3.92-4.01 (m, 2H, $\text{HNCH}(\text{CH}_3)_2$), 7.45 (dd, $J = 7.2, 7.2$ Hz, 1H), 7.49 (dd, $J = 7.6, 7.6$ Hz, 1H), 7.59 (dd, $J = 7.3, 7.3$ Hz, 1H), 8.21 (ddd, $J = 7.7, 3.6, 1.3$ Hz, 1H), 8.87 (br, 1H, NH);
10 ^{13}C -NMR (101MHz, CDCl_3): δ -7.3 (s, PdCH_3), 12.0 (s, HNCH_2CH_3), 17.9 (s, $\text{HNCH}(\text{CH}_3)_2$), 19.2 (s, $\text{HNCH}(\text{CH}_3)_2$), 26.0 (s), 26.9-27.4 (m), 28.7 (s), 29.4 (d, $J_{\text{PC}} = 4$ Hz), 35.6 (d, $^1J_{\text{PC}} = 25$ Hz), 42.4 (s, HNCH_2CH_3), 54.6 (s, $\text{HNCH}(\text{CH}_3)_2$), 125.5 (d, $^1J_{\text{PC}} = 33$ Hz), 128.3 (d, $J_{\text{PC}} = 7$ Hz), 128.9 (d, $J_{\text{PC}} = 6$ Hz), 130.3 (s), 132.5 (s),
15 150.9 (d, $^2J_{\text{PC}} = 11$ Hz);

^{31}P -NMR (162MHz, CDCl_3): δ 31.7;

Anal. calcd for $\text{C}_{27}\text{H}_{49}\text{ClNO}_3\text{PPdS}$, C, 50.62; H, 7.71; N, 2.19.

found: C, 50.49; H, 8.00, N, 2.12

[0059]

25 (c) Synthesis of metal complex 1c

A methylene chloride solution (4 ml) of metal complex 1b (0.19 g, 0.30 mmol) was added to a methylene chloride suspension (2 ml) of potassium carbonate (0.42 g, 3.0 mmol) and 2,6-lutidine (manufactured by Tokyo Chemical Industry Co., Ltd.,
30 0.33 g, 3.1 mmol) under argon atmosphere and stirred for one

hour at room temperature. Solid remained after distilling away the solvent under reduced pressure was washed with diethyl ether and extracted with methylene chloride. The extract was filtered through by Celite (dry diatom earth) and slowly added to hexane (40 ml). The precipitate formed was collected by filtration, washed with hexane and then dried under reduced pressure to obtain metal complex 1c. The yield was 0.12 g (70%).

[0060]

$^1\text{H-NMR}$ (400MHz, CDCl_3): δ 0.32 (d, $^3J_{\text{PH}} = 2.3$ Hz, 3H, PdCH_3), 1.12-1.47 (m, 8H), 1.60-1.94 (m, 10H), 2.22-2.33 (m, 4H), 3.18 (s, 6H, CH_3 of lutidine), 7.12 (d, $J = 7.7$ Hz, 2H), 7.47 (dd, $J = 7.6, 7.6$ Hz, 1H), 7.52 (dddd, $J = 7.6, 7.6, 1.4, 1.4$ Hz, 1H), 7.58 (dd, $J = 8.1, 8.1$ Hz, 1H), 7.60 (dd, $J = 7.5, 7.5$ Hz, 1H), 8.29 (ddd, 7.8, 3.9, 1.5 Hz, 1H);

$^{13}\text{C-NMR}$ (101MHz, CDCl_3): δ -9.4 (d, $^2J_{\text{PC}} = 4.8$ Hz, PdCH_3), 26.3 (s, CH_3 of lutidine), 26.9-27.5 (m), 28.6 (s), 29.6 (d, $J_{\text{PC}} = 3$ Hz), 35.5 (d, $^1J_{\text{PC}} = 26$ Hz), 122.5 (s), 122.5 (s), 124.6 (d, $^1J_{\text{PC}} = 35$ Hz), 128.9 (d, $J_{\text{PC}} = 7$ Hz), 129.0 (d, $J_{\text{PC}} = 6$ Hz), 130.7 (s), 132.4 (s), 138.1 (s). 151.0 (d, $^2J_{\text{PC}} = 12$ Hz), 159.0 (s);

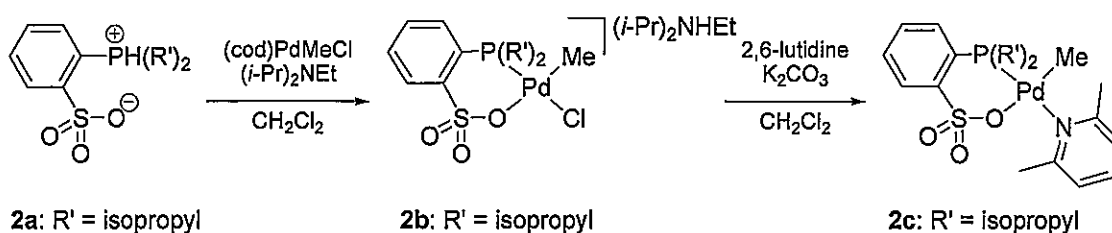
$^{31}\text{P-NMR}$ (162MHz, CDCl_3): δ 27.5;

Anal. calcd for $\text{C}_{26}\text{H}_{38}\text{NO}_3\text{PPdS}$, C, 53.65; H, 6.58; N, 2.51. found: C, 53.51; H, 6.74; N, 2.40

[0061] Synthesis of Metal complex 2c

Metal complex 2c was synthesized according to the following reaction scheme:

[Chem. 20]



[0062] (a) Synthesis of Compound 2a

2-(Diisopropylphosphino)benzenesulfonic acid

(Compound 2a) was obtained in the same way as in the above-mentioned synthesis of Compound 1a except that

5 chlorodiisopropylphosphine (manufactured by Sigma-Aldrich, 19.0 g, 125 mmol) was used instead of chlorodicyclohexylphosphine. The yield was 26.8 g (78%).

[0063]

$^1\text{H-NMR}$ (400MHz, CDCl_3): δ 1.25 (dd, $J = 21.6, 7.0$ Hz, 6H), 1.53 (dd, $J = 21.8, 7.2$ Hz, 6H), 3.45 (m, 2H), 5.42 (br d, $^1J_{\text{PH}} = 380$ Hz), 7.58 (tdd, $J = 7.6, 2.8, 1.1$ Hz, 1H), 7.69 (ddd, $J = 15.1, 7.7, 0.7$ Hz, 1H), 7.83 (dd, $J = 7.6, 7.6$ Hz, 1H), 8.27 (dd, $J = 7.5, 4.4$ Hz, 1H);

15 $^{13}\text{C-NMR}$ (101MHz, CDCl_3): δ 19.4 (s), 24.5-27.7 (m), 114.4 (br d, $J = 93$ Hz), 129.1 (d, $J = 8.6$ Hz), 130.3 (d, $J = 12.5$ Hz), 134.7-137.1 (m), 150.7 (br s);

$^{31}\text{P-NMR}$ (162MHz, CDCl_3): δ 62.5 (d, $^1J_{\text{PH}} = 380$ Hz) (83%), 31.0 (d, $^1J_{\text{PH}} = 460$ Hz) (17%)

[0064] (b) Synthesis of Metal complex 2b

20 Metal complex 2b was obtained in the same way as in the above-mentioned synthesis of Compound 1b except that Compound 2a (2-(dicyclohexylphosphino)benzenesulfonic acid, 16.3 g, 59.3 mmol) was used instead of Compound 1a. The yield was 33.7 g (99%).

25 [0065]

$^1\text{H-NMR}$ (500MHz, CDCl_3): δ 0.73 (d, $J = 1.8$ Hz, 3H), 1.20-1.29 (m, 12H), 1.45-1.59 (m, 15H), 2.54 (dt, $J = 21.9, 7.0$ Hz, 2H), 3.29 (d, $J = 7.1$ Hz, 2H), 3.97 (t, $J = 6.9$ Hz, 2H), 7.43-7.52 (m, 2H), 7.57 (t, $J = 7.0$ Hz, 1H), 8.22-8.24 (m, 1H);

30 $^{13}\text{C-NMR}$ (101MHz, CDCl_3): δ -7.79 (s), 12.16 (s), 17.91 (s), 18.54

(s), 19.21-19.28 (m), 25.93 (d, $J = 25.9$ Hz), 42.58 (s), 54.69 (s), 125.25 (s), 125.58 (s), 128.8 (d, $J = 6.7$ Hz), 129.13 (d, $J = 5.8$ Hz), 130.58 (s), 132.48 (s), 150.89 (d, $J = 10.5$ Hz);

^{31}P -NMR (162MHz, CDCl_3): δ 38.4 (s)

5 [0066] (c) Synthesis of Metal complex 2c

Metal complex 2c was obtained in the same way as in the above-mentioned synthesis of Metal complex 1c except that Metal complex 2b (33.7 g, 58.5 mmol) was used instead of Metal complex 1b. The yield was 27.6 g (94%).

10 [0067]

^1H -NMR (400 MHz, CDCl_3): δ 0.34 (d, $J = 2.3$ Hz, 3H), 1.32 (ddd, $J = 49.9, 16.0, 7.0$ Hz, 12H), 2.58 (dt, $J = 22.3, 7.2$ Hz, 2H), 3.18 (s, 6H), 7.12 (d, $J = 7.8$ Hz, 2H), 7.46 (t, $J = 7.4$ Hz, 1H), 7.53 (t, $J = 7.6$ Hz, 1H), 7.58 (t, $J = 7.7$ Hz, 2H), 8.29-

15 8.32 (m, 1H) ;

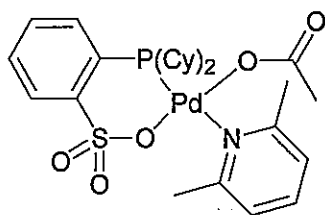
^{13}C -NMR (101 MHz, CDCl_3): δ -10.10 (d, $J = 4.8$ Hz), 18.44 (s), 19.29 (d, $J = 4.8$ Hz), 25.91 (d, $J = 25.9$ Hz), 26.20 (s), 122.72 (d, $J = 3.8$ Hz), 124.56 (d, $J = 35.5$ Hz), 129.19 (t, $J = 6.7$ Hz), 131.03 (d, $J = 1.9$ Hz), 132.39 (s), 138.30 (s), 151.13 (d, $J = 10.5$ Hz), 159.17 (s) ;

^{31}P -NMR (162MHz, CDCl_3): δ 34.4 (s);

Anal. Calcd for $\text{C}_{20}\text{H}_{30}\text{NO}_3\text{PPdS}$, C, 47.86; H, 6.02; N, 2.79. Found: C, 47.92; H, 5.98; N, 2.51

[0068] Example 1: Synthesis of Metal complex 1

25 [Chem. 21]



Cy = cyclohexyl

The toluene (10 ml)/allyl acetate (10 ml) suspension of Metal complex 1c (0.12 g, 0.20 mmol) was heated to 80°C and stirred for one hour under nitrogen atmosphere. After concentrating the reaction solution, methylene chloride (10 ml) was added to the residue and the particulate was removed by filtration using a syringe filter. Hexane (15 ml) was slowly added to the filtration to perform recrystallization. The precipitated crystals were recovered by filtration, washed with hexane, and then dried under reduced pressure to obtain Metal complex 1. The yield was 0.083 g (66%). Furthermore, single crystals of the obtained Metal complex 1 were produced from a methylene chloride/hexane solution to determine the structure by X-ray crystal structure analysis. The X-ray structure analysis of Metal complex 1 is shown in Fig. 1.

[0069]

¹H-NMR (400MHz, CDCl₃): δ 1.04–2.44 (m, 25H), 3.42 (s, 6H), 7.06, (d, J = 7.6 Hz, 2H), 7.53–7.62 (m, 4H), 8.17 (m, 1H).

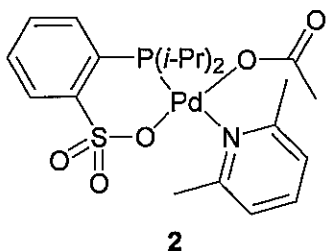
¹³C-NMR (101MHz, CDCl₃): δ 22.7 (s), 24.7 (s), 26.1 (s), 27.1–27.3 (m), 28.1 (d, ¹J_{pc} = 19 Hz), 34.7 (d, ¹J_{pc} = 25 Hz), 121.7 (s),

122.1 (s), 122.8 (d, J_{pc} = 3 Hz), 128.7 (d, J_{pc} = 7 Hz), 130.1 (d, J_{pc} = 6 Hz), 131.9 (s), 132.1 (s), 138.7 (s), 148.9 (d, J_{pc} = 10 Hz), 160.1 (s), 176.7 (s).

³¹P-NMR (162MHz, CDCl₃): δ 17.0 (s)

[0070] Example 2: Synthesis of Metal complex 2

[Chem. 22]



Metal complex 2 was synthesized using Metal complex 2c in the same way as in Example 1. That is, toluene (94 ml)/allyl acetate (94 ml) suspension of Compound 2c (1.2 g, 2.4 mmol) was heated to 80°C and stirred for two hours under nitrogen atmosphere. After concentrating the reaction solution, methylene chloride (10 ml) was added to the residue and the particulate was removed by filtration using a syringe filter. Hexane (15 ml) was slowly added to the filtrate to perform recrystallization. The precipitated crystals were collected by filtration, washed with t-butyl methyl ether and hexane, and then dried under reduced pressure to obtain Metal complex 2. The yield was 1.1 g (83%). Furthermore, single crystals of the obtained Metal complex 2 were produced from acetone/hexane to determine the structure by X-ray crystal structure analysis. The X-ray structure analysis of Metal complex 2 is shown in Fig. 2.

[0071]

¹H-NMR (400MHz, CDCl₃): δ 1.36-1.52 (m, 12H), 1.75 (s, 3H), 2.51-2.60 (m, 2H), 3.42 (s, 6H), 7.06 (d, 2H, J = 7.2 Hz), 7.55-7.64 (m, 4H), 8.20 (br, 1H);

¹³C-NMR (101MHz, CDCl₃): δ 17.7 (s), 17.8 (s), 22.5 (s), 24.5 (s, CH₃ of lutidine), 25.0 (d, ¹J_{PC} = 26.5 Hz), 121.4 (d, ¹J_{PC} = 39.8 Hz), 122.6 (s), 128.5 (d, J_{PC} = 5.7 Hz), 129.8 (d, J_{PC} = 5.8 Hz), 131.6 (s), 131.8 (s), 138.5 (s), 148.5 (d, ²J_{PC} = 8.3 Hz), 159.8 (s), 176.7 (s)

[0072] Reference Example 1: Comparison of stability of Metal complex 2c and Metal complex 2 in the presence of acetic acid

An experiment for comparing the stability of Metal complex 2c and Metal complex 2 in the presence of acetic acid

was conducted. That is, a deuterated toluene solution of acetic acid (0.062 M, 0.6 ml, 0.037 mmol) was added to an NMR tube (inner diameter of 5 mm) filled with Metal complex 2c (0.0037 mmol) and Metal complex 2 (0.0037 mmol) respectively, for the ¹H-NMR measurement. After heating the NMR tubes at 80°C for ten hours, ¹H-NMR was again measured. In addition to the peak derived from Metal complex 2c, the peak derived from Metal complex 2 (δ 1.74 ppm, 3.44 ppm, 8.22 ppm) and the peak derived from 2,6-dimethylpyridine (lutidine) (δ 2.38 ppm) decomposed and dissociated from the center metal were observed in the reaction solution containing Metal complex 2c. The abundance ratio of these was calculated to be 81.7 % of Metal complex 2c, 10.6% of Metal complex 2 and 7.7% of dissociated lutidine by the integral value of lutidine. In contrast, a peak derived from a complex other than Metal complex 2 was not observed in the reaction solution containing Metal complex 2. The peak of acetic acid was observed at 1.59 ppm and at 9.80 ppm, which can be distinguished from the above-mentioned peaks of Metal complex 2, Metal complex 2c and lutidine.

Protonic acid such as acetic acid and acrylic acid exists as an impurity in the polar-group containing olefin, which is a monomer to be polymerized. The results of Reference Example 1 show that the metal complex of the present invention is remarkably stable against the impurity and difficult to be decomposed by the impurity even in the polymerization reaction.

[0073] Example 3: Ethylene homopolymerization (preparation of Polymer 1)

A toluene solution (75 ml) of metal complex catalyst 2 (0.50 mmol) was added into a 120 ml-volume autoclave under nitrogen atmosphere. After filling the autoclave with ethylene

so that the inner pressure is 3.0 MPa, the content of the autoclave was stirred at 80°C for 15 minutes. After cooling the autoclave to room temperature, the reaction solution was added to methanol (400 ml). The copolymer formed was collected by
5 filtration, washed with methanol and dried under reduced pressure to obtain polymer 1. The yield was 0.24 g. The number average molecular weight and weight average molecular weight of the copolymer were 19,000 and 42,000, respectively, and Mw/Mn was 2.2.

10 [0074] Example 4: Copolymerization of allyl acetate and ethylene (preparation of copolymer 2)

A toluene solution (37.5 ml) of metal complex catalyst 2 (0.10 mmol) was added into a 120 ml-volume autoclave containing allyl acetate (37.5 ml, 34.9 g, 350 mmol) under
15 nitrogen atmosphere. After filling the autoclave with ethylene so that the inner pressure is 3.0 MPa, the content of the autoclave was stirred at 80°C for five hours. After cooling the autoclave to room temperature, the reaction solution was added to methanol (400 ml). The copolymer formed was collected by
20 filtration, washed with methanol and dried under reduced pressure to obtain Polymer 2. The yield was 2.7 g. The number average molecular weight and weight average molecular weight of the copolymer were 12,000 and 25,000, respectively, and Mw/Mn was 2.1. The allyl acetate content in Polymer 2 was determined
25 to be 3.1 % by molar fraction by ¹³C-NMR spectrum using the inverse-gated decoupling method.

[0075] Example 5: Copolymerization of allyl acetate and ethylene (preparation of Polymer 3)

A toluene solution (37.5 ml) of metal complex
30 catalyst 2 (0.010 mmol) was added into a 120 ml-volume autoclave

containing allyl acetate (37.5 ml, 34.9 g, 350 mmol) under nitrogen atmosphere. After filling the autoclave with ethylene so that the inner pressure is 3.0 MPa, the content of the autoclave was stirred at 80°C for 25 hours. After cooling the autoclave to room temperature, the reaction solution was added to methanol (400 ml). The copolymer formed was collected by filtration, washed with methanol and dried under reduced pressure to obtain Polymer 3. The yield was 2.2 g. The number average molecular weight and weight average molecular weight of the copolymer were 12,000 and 28,000, respectively, and Mw/Mn was 2.3. The allyl acetate content in Polymer 3 was determined to be 3.4 % by molar fraction by ¹³C-NMR spectrum using the inverse-gated decoupling method.

[0076] Comparative Example 1: Copolymerization of allyl acetate and ethylene using a Pd-hydride complex as a catalyst

Methylene chloride (3.75 ml) was added into a 50 ml-volume autoclave containing Compound 2b (0.12 mmol) and Pd₂(DBA)₃·CHCl₃ (DBA: dibenzylideneacetone, 0.10 mmol) and stirred for ten minutes under argon atmosphere. It is presumed that a Pd-hydride complex is formed as a catalyst species (Reference document: J. Am. Chem. Soc., 131, 14605 (2009)).

Subsequently, toluene (3.75 ml) and allyl acetate (7.5 ml, 70 g, 70 mmol) were added to the catalyst solution. After filling the autoclave with ethylene so that the inner pressure is 3.0 MPa, the content of the autoclave was stirred at 80°C for 15 hours. After cooling the autoclave to room temperature, methanol (about 20 ml) was added into the autoclave. The copolymer formed was collected by filtration, washed with methanol and dried under reduced pressure to obtain a polymer. The yield was 1.7 g. The number average molecular weight and weight average molecular

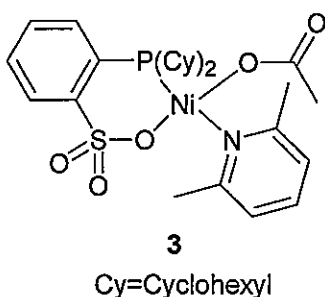
weight of the copolymer were 4,000 and 9,000, respectively, and Mw/Mn was 2.7. The allyl acetate content in the copolymer was determined to be 2.7 % by molar fraction by ^{13}C -NMR spectrum using the inverse-gated decoupling method.

5 [0077] Example 6: Copolymerization of allyl acetate and ethylene (preparation of Polymer 4)

A toluene solution (188 ml) of metal complex catalyst 2 (0.050 mmol) was added into a 500 ml-volume autoclave containing allyl acetate (188 ml, 174 g, 1740 mmol) under
10 nitrogen atmosphere. After filling the autoclave with ethylene so that the inner pressure is 3.0 MPa, the content of the autoclave was stirred at 80°C for 50 hours. After cooling the autoclave to room temperature, the reaction solution was added to methanol (1,000 ml). The copolymer formed was
15 collected by filtration, washed with methanol and dried under reduced pressure to obtain Polymer 4. The yield was 23 g. The number average molecular weight and weight average molecular weight of the copolymer were 12,000 and 28,000, respectively, and Mw/Mn was 2.3. The allyl acetate content in Polymer 4 was
20 determined to be 3.4 % by molar fraction by ^{13}C -NMR spectrum using the inverse-gated decoupling method.

[0078] Example 7: Synthesis of Metal complex 3

[Chem. 23]



25

Bis(1,5-cyclooctadiene)nickel (0.14 g, 0.50 mmol),

2-(dicyclohexylphosphino)benzenesulfonic acid (compound 1a, 0.21 g, 0.60 mmol), 2,6-dimethylpyridine (0.11 g, 1.0 mmol) and THF (30 ml) were stirred at room temperature for five hours under argon atmosphere. Allyl acetate (1 ml, 0.93 g, 9.3 mmol) was added to the reaction solution to allow the reaction to proceed further. After filtering the reaction solution, the filtrate was concentrated under reduced pressure to obtain 0.12 g (43%) of orange powder. It was confirmed by the ^{31}P -NMR spectrum that the powder was Metal complex 3.

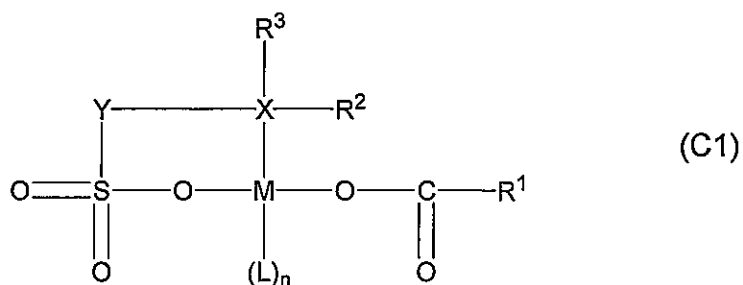
^{31}P -NMR (162 MHz, CD_2Cl_2): δ 16.6 (s)

[0079] Example 8: Ethylene homopolymerization (Preparation of Polymer 5)

A toluene solution (6 ml) was added into a 50 ml-volume autoclave containing metal complex catalyst 3 (0.050 mmol) under nitrogen atmosphere. After filling the autoclave with ethylene so that the inner pressure is 3.0 MPa, the content of the autoclave was stirred at 80°C for seven hours. After cooling the autoclave to room temperature, methanol (20 ml) was added into the autoclave. The polymer formed was collected by filtration, washed with methanol and dried under reduced pressure to obtain Polymer 5. The yield was 0.12 g. The number average molecular weight and weight average molecular weight of the polymer were 2,800 and 9,800, respectively, and M_w/M_n was 3.5.

CLAIMS

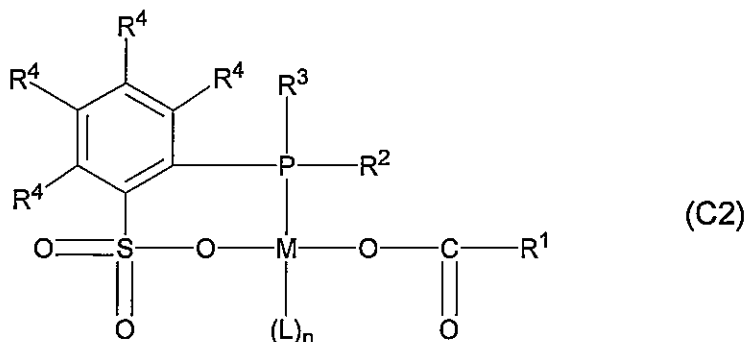
[1] A carboxylate metal complex represented by formula
(C1)



(in the formula, M represents a metal atom of group 10 element in the periodic table; X represents a phosphorous atom (P) or an arsenic atom (As); R¹ represents a hydrogen atom or hydrocarbon group having 1 to 30 carbon atoms which may be substituted by one or more functional groups selected from a halogen atom, alkoxy group, aryloxy group, acyloxy group and amino group; R² and R³ independently represent a hydrogen atom, alkoxy group, aryloxy group, acyloxy group, silyl group, amino group or hydrocarbon group having 1 to 30 carbon atoms which may be substituted by one or more functional groups selected from a halogen atom, alkoxy group, aryloxy group, acyloxy group and amino group; R² and R³ may bond to each other to form a ring structure; Y represents divalent hydrocarbon group having 1 to 70 carbon atoms which may be substituted by one or more functional groups selected from a halogen atom, alkoxy group, aryloxy group, acyloxy group and amino group; L represents a neutral electron-donating ligand and n is 0, 1/2, 1 or 2.

2. The carboxylate metal complex as claimed in claim 1, wherein the carboxylate metal complex represented by formula

(C1) is represented by formula (C2)



(in the formula, four R⁴s independently represent a hydrogen atom, alkyl group having 1 to 8 carbon atoms, alkoxy group having 1 to 8 carbon atoms, aryloxy group having 6 to 18 carbon atoms or a halogen atom; and M, R¹, R², R³, L and n have the same meanings as in claim 1).

3. The carboxylate metal complex as claimed in claim 2, wherein M in formula (C2) is Ni or Pd.

4. The carboxylate metal complex as claimed in claim 2 or 3, wherein R¹ in formula (C2) is hydrocarbon group having 1 to 30 carbon atoms which may be substituted by a halogen atom.

5. The carboxylate metal complex as claimed in claim 4, wherein R¹ in formula (C2) is methyl group or trifluoromethyl group.

6. The carboxylate metal complex as claimed in claim 5, wherein R¹ in formula (C2) is methyl group.

7. The carboxylate metal complex as claimed in any one of claims 2 to 6, wherein both of R² and R³ in formula (C2) are

isopropyl group or cyclohexyl group.

8. The carboxylate metal complex as claimed in any one of claims 2 to 7, wherein all of R⁴s in formula (C2) are a hydrogen atom.

9. The carboxylate metal complex as claimed in any one of claims 2 to 8, wherein L in formula (C2) is 2,6-dimethylpyridine and n is 1.

10. A catalyst for olefin polymerization mainly comprising the carboxylate metal complex claimed in any one of claims 1 to 9.

11. A method for producing a polymer characterized in polymerizing olefin represented by formula (1) singly or in combination of two or more kinds thereof, or copolymerizing olefin represented by formula (1) and polar group-containing olefin represented by formula (2)



(In the formula, R⁵ represents a hydrogen atom or hydrocarbon group having 1 to 20 carbon atoms)



(In the formula, R⁶ represents a hydrogen atom or methyl group,

R⁷ represents -COOR⁸, -CN, -OCOR⁸, -CH₂-OCOR⁸, -CH₂OH, -CH₂-N(R⁹)₂ or -CH₂-Hal (R⁸ represents a hydrogen atom or hydrocarbon group having 1 to 5 carbon atoms, R⁹ represents a hydrogen atom, hydrocarbon group having 1 to 5 carbon atoms or aromatic substituent having 6 to 18 carbon atoms or alkoxy carbonyl group, and Hal represents a halogen atom).